

 Open access • Journal Article • DOI:10.1107/S1600536807026815

(2E)-3-(2-Chlorophenyl)-1-(3-nitrophenyl)prop-2-en-1-one — [Source link](#)

Balladka Kunhanna Sarojini, Hemmige S. Yathirajan, K. Lakshmana, Badiadka Narayana ...+1 more authors

Institutions: P A College of Engineering, University of Mysore, Mangalore University, Goethe University Frankfurt

Published on: 01 Jul 2007 - Acta Crystallographica Section E-structure Reports Online (International Union of Crystallography (IUCr))

Related papers:

- [\(2E\)-1-\(3-Chlorophenyl\)-3-\(4-chlorophenyl\)prop-2-en-1-one](#)
- [4-Chloro-3-nitrophenyl methyl sulfone](#)
- [\(E\)-3-\(2,3-Dichlorophenyl\)-1-\(2-methoxyphenyl\)prop-2-en-1-one](#)
- [\(2E\)-3-\(5-bromo-2-thienyl\)-1-\(4-hydroxyphenyl\)prop-2-en-1-one](#)
- [N-\[3-\(2-Nitrophenyl\)allylidene\]naphthalen-1-amine](#)

Share this paper:    

View more about this paper here: <https://typeset.io/papers/2e-3-2-chlorophenyl-1-3-nitrophenyl-prop-2-en-1-one-nm2muwbj7f>

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

(2E)-3-(2-Chlorophenyl)-1-(3-nitrophenyl)prop-2-en-1-one**B. K. Sarojini,^a H. S. Yathirajan,^b K. Lakshmana,^c B. Narayana^c and Michael Bolte^{d*}**

^aDepartment of Chemistry, P. A. College of Engineering, Nadupadavu, Mangalore 574 153, India, ^bDepartment of Studies in Chemistry, University of Mysore, Manasagangotri, Mysore 570 006, India, ^cDepartment of Studies in Chemistry, Mangalore University, Mangalagangotri 574 199, India, and ^dInstitut für - Anorganische Chemie, J. W. Goethe-Universität Frankfurt, Max-von-Laue-Strasse 7, 60438 Frankfurt/Main, Germany
Correspondence e-mail: bolte@chemie.uni-frankfurt.de

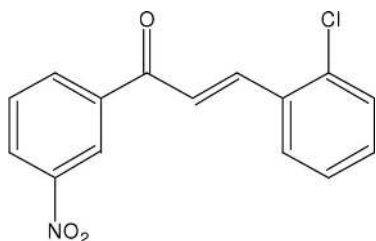
Received 29 May 2007; accepted 1 June 2007

Key indicators: single-crystal X-ray study; $T = 173$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.041; wR factor = 0.112; data-to-parameter ratio = 13.4.

Geometric parameters of the title compound, $\text{C}_{15}\text{H}_{10}\text{ClNO}_3$, a chalcone derivative, are in the usual ranges. The molecules are almost planar and crystallize in stacks with an interplanar distance of approximately 3.4 Å. No classical hydrogen bonds were found.

Related literature

For related structures, see: Yathirajan, Mayekar, Narayana *et al.* (2007); Fischer *et al.* (2007); Yathirajan, Mayekar, Sarojini *et al.* (2007). For related literature, see: Di Carlo *et al.* (1999); Dimmock *et al.* (1999); Go *et al.* (2005); Fichou *et al.* (1988); Goto *et al.* (1991); Uchida *et al.* (1998); Sarojini *et al.* (2006).

**Experimental***Crystal data* $\text{C}_{15}\text{H}_{10}\text{ClNO}_3$ $M_r = 287.69$ Monoclinic, $P2_1/n$ $a = 7.3643$ (10) Å $b = 12.8295$ (11) Å $c = 13.9452$ (14) Å $\beta = 93.898$ (9)° $V = 1314.5$ (2) Å³ $Z = 4$ Mo $K\alpha$ radiation $\mu = 0.30$ mm⁻¹ $T = 173$ (2) K

0.43 × 0.40 × 0.37 mm

Data collection

Stoe IPDSII two-circle diffractometer

Absorption correction: multi-scan (*MULABS*; Spek, 2003; Blessing, 1995) $T_{\min} = 0.883$, $T_{\max} = 0.898$

7978 measured reflections

2433 independent reflections

1977 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.055$ *Refinement* $R[F^2 > 2\sigma(F^2)] = 0.041$ $wR(F^2) = 0.112$ $S = 1.04$

2433 reflections

182 parameters

H-atom parameters constrained

 $\Delta\rho_{\text{max}} = 0.26$ e Å⁻³ $\Delta\rho_{\text{min}} = -0.28$ e Å⁻³

Data collection: *X-AREA* (Stoe & Cie, 2001); cell refinement: *X-AREA*; data reduction: *X-AREA*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1990); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *SHELXL97*.

KL thanks Mangalore University for permission to carry out the research work and UGC, New Delhi, for a fellowship under FIP.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2310).

References

- Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
Di Carlo, G., Mascolo, N., Izzo, A. A. & Capasso, F. (1999). *Life Sci.* **65**, 337–353.
Dimmock, J. R., Elias, D. W., Beazely, M. A. & Kandepu, N. M. (1999). *Curr. Med. Chem.* **6**, 1125–1149.
Fichou, D., Watanabe, T., Takeda, T., Miyata, S., Goto, Y. & Nakayama, M. (1988). *Jpn. J. Appl. Phys.* **27**, 429–430.
Fischer, A., Yathirajan, H. S., Ashalatha, B. V., Narayana, B. & Sarojini, B. K. (2007). *Acta Cryst.* **E63**, o1351–o1352.
Go, M. L., Wu, X. & Liu, X. L. (2005). *Curr. Med. Chem.* **12**, 483–499.
Goto, Y., Hayashi, A., Kimura, Y. & Nakayama, M. (1991). *J. Cryst. Growth*, **108**, 688–698.
Sarojini, B. K., Narayana, B., Ashalatha, B. V., Indira, J. & Lobo, K. G. (2006). *J. Cryst. Growth*, **295**, 54–59.
Sheldrick, G. M. (1990). *Acta Cryst.* **A46**, 467–473.
Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.
Stoe & Cie (2001). *X-AREA*. Stoe & Cie, Darmstadt, Germany.
Uchida, T., Kozawa, K., Sakai, T., Aoki, M., Yoguchi, H., Abduryim, A. & Watanabe, Y. (1998). *Mol. Cryst. Liq. Cryst.* **315**, 135–140.
Yathirajan, H. S., Mayekar, A. N., Narayana, B., Sarojini, B. K. & Bolte, M. (2007). *Acta Cryst.* **E63**, o2200–o2201.
Yathirajan, H. S., Mayekar, A. N., Sarojini, B. K., Narayana, B. & Bolte, M. (2007). *Acta Cryst.* **E63**, o424–o425.

supplementary materials

Acta Cryst. (2007). E63, o3211 [doi:10.1107/S1600536807026815]

(2E)-3-(2-Chlorophenyl)-1-(3-nitrophenyl)prop-2-en-1-one

B. K. Sarojini, H. S. Yathirajan, K. Lakshmana, B. Narayana and M. Bolte

Comment

Chalcones are one of the major classes of natural products with widespread distribution in fruits, vegetables, spices, tea and soy based foodstuff have been recently subjects of great interest for their interesting pharmacological activities. Reviews on the bioactivities of varieties of chalcones are published in the literature recently. Recently, it has been noted that, among many organic compounds reported for their second harmonic generation, chalcone derivatives are known for their excellent blue light transmittance and good crystallizability. In continuation of our broad programme on chalcones, the present paper reports the crystal structure of a newly synthesized chalcone.

Geometric parameters of the title compound are in the usual ranges. The molecules are almost planar (r.m.s. deviation for all non-H atoms 0.113 Å) and crystallize in stacks. There is no classical hydrogen bonds.

Experimental

To a thoroughly stirred solution of 1-(3-nitrophenyl)ethanone (1.65 g, 0.01 mol) and 2-chlorobenzaldehyde (1.40 g, 0.01 mol) in 25 ml methanol, 5 ml of 30% KOH solution was added. The solution was stirred overnight and filtered. The product was crystallized from acetone by slow evaporation (m.p.: 410–412 K). Analysis for C₁₅H₁₀ClNO₃: Found (Calculated): C 62.56 (62.62), H 3.41 (3.50), N 4.81 (4.87).

Refinement

H atoms were found in a difference map, but they were refined using a riding model with C—H = 0.95 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Figures



Fig. 1. Perspective view of the title compound with the atom numbering. Displacement ellipsoids are at the 50% probability level.

(2E)-3-(2-Chlorophenyl)-1-(3-nitrophenyl)prop-2-en-1-one

Crystal data

C₁₅H₁₀ClNO₃

$M_r = 287.69$

Monoclinic, $P2_1/n$

$F_{000} = 592$

$D_x = 1.454 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation

$\lambda = 0.71073 \text{ \AA}$

supplementary materials

Hall symbol: -P 2yn
 $a = 7.3643$ (10) Å
 $b = 12.8295$ (11) Å
 $c = 13.9452$ (14) Å
 $\beta = 93.898$ (9)°
 $V = 1314.5$ (2) Å³
 $Z = 4$

Cell parameters from 7582 reflections
 $\theta = 3.5$ – 25.7 °
 $\mu = 0.30$ mm⁻¹
 $T = 173$ (2) K
Block, brown
 $0.43 \times 0.40 \times 0.37$ mm

Data collection

STOE IPDS II two-circle diffractometer
Radiation source: fine-focus sealed tube
Monochromator: graphite
 $T = 173$ (2) K
 ω scans
Absorption correction: multi-scan (MULABS; Spek, 2003; Blessing, 1995)
 $T_{\min} = 0.883$, $T_{\max} = 0.898$
7978 measured reflections

2433 independent reflections
1977 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.055$
 $\theta_{\max} = 25.6$ °
 $\theta_{\min} = 3.5$ °
 $h = -7 \rightarrow 8$
 $k = -13 \rightarrow 15$
 $l = -16 \rightarrow 16$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.041$
 $wR(F^2) = 0.112$
 $S = 1.04$
2433 reflections
182 parameters
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.0678P)^2 + 0.1649P]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.26$ e Å⁻³
 $\Delta\rho_{\min} = -0.28$ e Å⁻³
Extinction correction: SHELXL,
 $F_c^* = kF_c[1 + 0.001 \times F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.009 (2)

Special details

Experimental. ;

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculat-

ing R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.20045 (8)	0.15573 (4)	0.60287 (3)	0.03658 (19)
N1	0.5177 (2)	0.91050 (12)	0.71148 (11)	0.0321 (4)
O1	0.3613 (3)	0.52709 (11)	0.68321 (10)	0.0594 (6)
O2	0.5581 (2)	1.00014 (10)	0.68882 (11)	0.0470 (4)
O3	0.5469 (2)	0.87440 (12)	0.79277 (9)	0.0418 (4)
C1	0.3232 (3)	0.55836 (14)	0.60200 (13)	0.0303 (5)
C2	0.2583 (3)	0.48688 (15)	0.52359 (13)	0.0314 (5)
H2	0.2269	0.5149	0.4616	0.038*
C3	0.2423 (3)	0.38477 (14)	0.53709 (12)	0.0275 (4)
H3	0.2738	0.3591	0.5999	0.033*
C11	0.3414 (3)	0.67301 (13)	0.58022 (12)	0.0245 (4)
C12	0.4164 (3)	0.73743 (14)	0.65350 (12)	0.0253 (4)
H12	0.4546	0.7093	0.7146	0.030*
C13	0.4339 (3)	0.84219 (13)	0.63547 (12)	0.0247 (4)
C14	0.3776 (3)	0.88744 (15)	0.54758 (13)	0.0288 (4)
H14	0.3895	0.9603	0.5377	0.035*
C15	0.3042 (3)	0.82363 (15)	0.47535 (13)	0.0299 (4)
H15	0.2652	0.8526	0.4147	0.036*
C16	0.2868 (3)	0.71654 (14)	0.49089 (12)	0.0276 (4)
H16	0.2377	0.6730	0.4404	0.033*
C21	0.1804 (3)	0.30800 (14)	0.46390 (12)	0.0244 (4)
C22	0.1591 (3)	0.20200 (14)	0.48547 (12)	0.0261 (4)
C23	0.0989 (3)	0.12985 (15)	0.41637 (14)	0.0326 (5)
H23	0.0856	0.0587	0.4333	0.039*
C24	0.0579 (3)	0.16236 (16)	0.32188 (14)	0.0356 (5)
H24	0.0165	0.1134	0.2741	0.043*
C25	0.0778 (3)	0.26644 (16)	0.29803 (14)	0.0349 (5)
H25	0.0495	0.2890	0.2339	0.042*
C26	0.1392 (3)	0.33786 (14)	0.36791 (13)	0.0305 (4)
H26	0.1537	0.4088	0.3503	0.037*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C11	0.0563 (4)	0.0228 (3)	0.0303 (3)	0.0022 (2)	0.0010 (2)	0.00355 (17)
N1	0.0376 (11)	0.0251 (8)	0.0332 (9)	-0.0014 (8)	-0.0014 (7)	-0.0066 (7)
O1	0.1133 (16)	0.0240 (8)	0.0363 (8)	-0.0050 (9)	-0.0284 (9)	0.0049 (6)
O2	0.0688 (12)	0.0213 (8)	0.0491 (9)	-0.0094 (8)	-0.0087 (8)	-0.0055 (6)
O3	0.0591 (11)	0.0380 (8)	0.0269 (7)	-0.0061 (8)	-0.0064 (6)	-0.0040 (6)
C1	0.0378 (12)	0.0213 (9)	0.0306 (10)	-0.0005 (8)	-0.0053 (8)	0.0002 (7)
C2	0.0416 (13)	0.0222 (9)	0.0297 (10)	-0.0025 (9)	-0.0034 (8)	0.0006 (7)
C3	0.0321 (11)	0.0231 (9)	0.0266 (9)	0.0005 (8)	-0.0026 (7)	0.0004 (7)

supplementary materials

C11	0.0257 (10)	0.0211 (9)	0.0263 (9)	-0.0003 (8)	-0.0003 (7)	-0.0018 (7)
C12	0.0279 (10)	0.0243 (9)	0.0236 (9)	0.0016 (8)	0.0002 (7)	-0.0001 (7)
C13	0.0269 (10)	0.0226 (9)	0.0245 (9)	-0.0018 (8)	0.0018 (7)	-0.0043 (7)
C14	0.0331 (11)	0.0217 (9)	0.0317 (9)	-0.0003 (8)	0.0033 (8)	0.0018 (7)
C15	0.0359 (11)	0.0275 (10)	0.0257 (9)	0.0011 (8)	-0.0032 (8)	0.0030 (7)
C16	0.0320 (11)	0.0249 (10)	0.0251 (9)	-0.0008 (8)	-0.0029 (8)	-0.0032 (7)
C21	0.0230 (10)	0.0229 (9)	0.0270 (9)	-0.0008 (7)	0.0006 (7)	-0.0007 (7)
C22	0.0268 (10)	0.0240 (9)	0.0274 (9)	0.0008 (8)	0.0021 (7)	-0.0011 (7)
C23	0.0356 (12)	0.0235 (9)	0.0392 (11)	-0.0049 (9)	0.0058 (8)	-0.0041 (8)
C24	0.0348 (12)	0.0356 (11)	0.0359 (11)	-0.0054 (9)	-0.0023 (8)	-0.0119 (8)
C25	0.0397 (12)	0.0368 (11)	0.0273 (9)	-0.0003 (10)	-0.0048 (8)	-0.0008 (8)
C26	0.0348 (11)	0.0256 (9)	0.0306 (10)	-0.0019 (9)	-0.0012 (8)	0.0004 (7)

Geometric parameters (Å, °)

C11—C22	1.7486 (18)	C14—C15	1.380 (3)
N1—O3	1.230 (2)	C14—H14	0.9500
N1—O2	1.234 (2)	C15—C16	1.398 (3)
N1—C13	1.478 (2)	C15—H15	0.9500
O1—C1	1.216 (2)	C16—H16	0.9500
C1—C2	1.481 (2)	C21—C22	1.404 (3)
C1—C11	1.510 (2)	C21—C26	1.406 (2)
C2—C3	1.330 (3)	C22—C23	1.387 (3)
C2—H2	0.9500	C23—C24	1.395 (3)
C3—C21	1.468 (2)	C23—H23	0.9500
C3—H3	0.9500	C24—C25	1.386 (3)
C11—C12	1.399 (2)	C24—H24	0.9500
C11—C16	1.399 (2)	C25—C26	1.391 (3)
C12—C13	1.375 (3)	C25—H25	0.9500
C12—H12	0.9500	C26—H26	0.9500
C13—C14	1.394 (2)		
O3—N1—O2	123.84 (16)	C14—C15—C16	120.36 (16)
O3—N1—C13	118.42 (15)	C14—C15—H15	119.8
O2—N1—C13	117.71 (15)	C16—C15—H15	119.8
O1—C1—C2	121.70 (17)	C15—C16—C11	120.44 (16)
O1—C1—C11	119.39 (16)	C15—C16—H16	119.8
C2—C1—C11	118.90 (15)	C11—C16—H16	119.8
C3—C2—C1	122.19 (17)	C22—C21—C26	116.61 (16)
C3—C2—H2	118.9	C22—C21—C3	122.33 (16)
C1—C2—H2	118.9	C26—C21—C3	121.06 (17)
C2—C3—C21	126.06 (17)	C23—C22—C21	122.20 (17)
C2—C3—H3	117.0	C23—C22—C11	116.80 (14)
C21—C3—H3	117.0	C21—C22—C11	120.96 (13)
C12—C11—C16	119.34 (16)	C22—C23—C24	119.66 (18)
C12—C11—C1	117.70 (15)	C22—C23—H23	120.2
C16—C11—C1	122.96 (15)	C24—C23—H23	120.2
C13—C12—C11	118.77 (16)	C25—C24—C23	119.66 (17)
C13—C12—H12	120.6	C25—C24—H24	120.2
C11—C12—H12	120.6	C23—C24—H24	120.2

C12—C13—C14	122.83 (16)	C24—C25—C26	120.08 (18)
C12—C13—N1	119.22 (16)	C24—C25—H25	120.0
C14—C13—N1	117.94 (15)	C26—C25—H25	120.0
C15—C14—C13	118.25 (17)	C25—C26—C21	121.78 (18)
C15—C14—H14	120.9	C25—C26—H26	119.1
C13—C14—H14	120.9	C21—C26—H26	119.1
O1—C1—C2—C3	-1.7 (3)	C13—C14—C15—C16	0.2 (3)
C11—C1—C2—C3	178.7 (2)	C14—C15—C16—C11	0.9 (3)
C1—C2—C3—C21	-179.60 (19)	C12—C11—C16—C15	-1.1 (3)
O1—C1—C11—C12	5.3 (3)	C1—C11—C16—C15	178.94 (19)
C2—C1—C11—C12	-175.09 (18)	C2—C3—C21—C22	-176.8 (2)
O1—C1—C11—C16	-174.8 (2)	C2—C3—C21—C26	3.1 (3)
C2—C1—C11—C16	4.9 (3)	C26—C21—C22—C23	-0.6 (3)
C16—C11—C12—C13	0.2 (3)	C3—C21—C22—C23	179.33 (18)
C1—C11—C12—C13	-179.82 (18)	C26—C21—C22—C11	-178.37 (14)
C11—C12—C13—C14	0.9 (3)	C3—C21—C22—C11	1.6 (3)
C11—C12—C13—N1	-178.30 (17)	C21—C22—C23—C24	0.1 (3)
O3—N1—C13—C12	-9.7 (3)	C11—C22—C23—C24	178.00 (16)
O2—N1—C13—C12	168.70 (19)	C22—C23—C24—C25	0.0 (3)
O3—N1—C13—C14	171.00 (18)	C23—C24—C25—C26	0.3 (3)
O2—N1—C13—C14	-10.5 (3)	C24—C25—C26—C21	-0.8 (3)
C12—C13—C14—C15	-1.1 (3)	C22—C21—C26—C25	0.9 (3)
N1—C13—C14—C15	178.10 (17)	C3—C21—C26—C25	-179.01 (19)

Fig. 1

