

A Green Protocol for Efficient Synthesis of 1,8-Dioxo-Octahydroxanthenes Using Ionic Liquid

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ABSTRACT

A clean, simple, efficient synthesis of 1,8-dioxo-octahydroxanthenes derivatives using ionic liquid 1-butyl, 3-methylimidazolium perchlorate ([bmim] ClO₄) has been developed. The method provides several advantages such as simple work up, environmental friendliness and excellent yields in short reaction time. The ionic liquid [bmim] ClO₄ used was recovered and reused for three times.

Keywords: 1,8-Dioxo-Octahydroxanthenes; Ionic Liquid; Aldehydes; Dimedone

1. Introduction

1,8-dioxo-octahydroxanthenes are important class of oxygen heterocycles in which a phenyl substituted pyran ring is fused on either side with two cyclohexanone rings. Presence of conjugated bis-dienone functionality makes these compounds sensitive to be attacked by nucleophiles and light energy. In the past decade, synthesis of xanthenes derivatives has been of considerable interest to organic chemists because they possess various biological and pharmaceuticals activities such as antiviral [1], antibacterial [2] and anti-inflammatory [3] properties. These are being utilized as antagonists for paralyzing action of zoxazolamine [4] and in photodynamic therapy [5]. Xanthenes and benzoxanthenes derivatives are the parent frame works found in a large number of naturally occurring as well as synthetic products possessing prominent position in medicinal chemistry [6,7]. Xanthenediones are likewise special structural units constituting various natural products [8] and being used as versatile synthons, because of inherent reactivity of their inbuilt pyran ring [9].

Apart from their use as valuable synthetic precursors for many organic compounds [10] and dyes [11], they are also used in laser technologies [12] and pH—sensitive fluorescent materials for the visualization of biomolecules assemblies [13].

There are a wide variety of methods for the preparation of the xanthenes having been reported and they are classified according to starting compounds which includes synthesis by cyclization of polycyclic aryltiffate esters [14], intra-molecular trapping of benzynes by phenols [15] and reaction of aryloxy magnesium halides with triethyl orthoformate [16] as well as cyclo-condensation of 2-hydroxyl aromatic aldehyde with 2-tetralone [17].

One of the commonly used methods reported for the synthesis of xanthenediones involves the condensation of aromatic aldehydes with 1, 3-cyclohexanedione or 5,5-dimethyl-1, 3-cyclohexanedione. The various catalysts have been reported for synthesis of 1,8-dioxo-octahydroxanthenes: NSPVC [18], [Et₃NH][HSO₄] [19], SiCl₄ [20], [Et₃NC₄SO₃H] [HSO₄]/Al₂O₃ [21], polyaniline p-toluenesulfonate [22], alumina-sulfuric acid [23], CAN [24], [BMIm][BF₄]-Mg(BF₄) [25], I₂ [26], [Et₃N-SO₃H]Cl [27], DABCO [28], SmCl₃ [29], cellulose sulfonic acid [30], heteropolyacid-MCM-41 *i.e.* (H₃PW₁₂O₄₀) [31], [bmim] HSO₄ [32], Fe(HSO₄)₃ [33], [Hbmim]BF₄ [34], TBAHS [35] tetramethylguanidium trifluoroacetate [36], [DDPA] [HSO₄] [37], [TMPSA] HSO₄ [38], [Hmim] TFA [39], trichloro isocyanuric acid [40], SbCl₃/SiO₂ [41], SiO₂-R-SO₃H [42], InCl₃/P₂O₅ [43], p-dodecyl benzenesulfonic acid [44], triethyl benzyl ammonium chloride (TEBA) [45], diammonium hydrogen phosphate [46], PPA-SiO₂ [47], montmorillonite K-10 [48], Fe³⁺-Montmorillonite

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[49], trimethyl silyl chloride [50], $\text{NaHSO}_4\text{-SiO}_2$ [51], $\text{ZnO-Acetyl chloride}$ [52], PMA-SiO_2 [53], $\text{ZrOCl}_2\cdot 8\text{H}_2\text{O}$ [54], Amberlyst-15 [55], Dowex-50W [56]. Most of these methods have some advantages and disadvantages, taking the weakness into consideration such as long reaction time, low yields, use of hazardous catalysts and solvents, tedious work up processes and difficulty in recovery and reusability of catalyst. To overcome these problems we have developed the green protocol for efficient synthesis of 1,8-dioxo-octahydroxanthenes using [bmim]ClO₄ ionic liquid (**Scheme 1**).

In recent years, ionic liquids have emerged as a useful alternative to conventional organic solvents and catalysts due to their particular properties, such as negligible vapor pressure, chemical stability, excellent solvent power for organic and inorganic compounds, as well as the ease of recovery [57]. Recently, much attention has been paid to organic reactions carried by using ionic liquids greater than ever importance in the context of green synthesis. Although ionic liquids were primarily introduced as an alternative green reaction medium, today they have progressed extremely beyond that, playing a significant role in controlling reactions as catalysts.

2. Experimental Procedure

All the chemicals were purchased from Sigma Aldrich and Merck chemicals. Melting points recorded on Veego melting point apparatus and were uncorrected. ¹H NMR spectra were recorded on a Bruker Avance II 400 MHz spectrometer with TMS as internal standard. IR spectra were determined on Perkin Elmer Spectrum RX FTIR spectrometer. The ionic liquid was synthesized according to ref. [58].

General Procedure for Synthesis of 1,8-Dioxo-Octahydroxanthenes Derivatives

To a mixture of aldehyde (1 mmol) and 5,5-dimethyl-1,3-cyclohexanedione (2 mmol) and [bmim]ClO₄ (4 mmol) was added in 50 ml round bottom flask and was stirred at 100°C on heating magnetic stirrer (**Scheme 1**). The progress of the reaction was monitored by TLC after completion of the reaction the reaction mixture was cooled to room temperature and water (5 ml) was added, solid separated was filtered. The crude product was recrystallized from ethanol to give the pure product. The [bmim]ClO₄ was recovered by distillation and reused 3 times.

The comparison of various reported catalyst is summarized in (**Table 1**).

3,3,6,6-tetramethyl-9-(4-chloro-phenyl)-1,8-dioxo-octahydroxanthene (Entry 1) IR (cm⁻¹): 3030, 2963, 2952, 1679, 1661, 1469, 1361, 1198, 1166, 1003, 852. ¹H NMR: 7.17 - 7.24 (dd, 4H), 4.71 (s, 1H), 2.46 (s, 4H),

2.18 (q, 4H), 1.10 (s, 6H), 0.98 (s, 6H). ¹³C NMR: 191.10, 157.17, 137.45, 126.79, 124.53, 122.97, 110.03, 45.45, 35.61, 26.96, 26.22, 24.03, 22.05 m/z: 385.2 (M + 1), 273.2.

3,3,6,6-tetramethyl-9-(4-methyl-phenyl)-1,8-dioxo-octahydroxanthene (Entry 2) IR (cm⁻¹): 3036, 2959, 2873, 1663, 1623, 1359, 1197, 1164, 1139, 1000, 842. ¹H NMR: 7.17 (d, 2H), 7.01(d, 2H), 4.71(s, 1H), 2.45 (s, 4H), 2.24 (s, 3H), 2.18 (q, 4H), 1.09 (s, 6H), 0.99 (s, 6H). ¹³C NMR: 196.4, 162.0, 141.1, 135.7, 128.7, 128.2, 115.7, 50.7, 40.8, 32.2, 31.4, 29.2, 27.3, 21.0. m/z: 365.3 (M + 1), 273.2.

9,9'-(1,4-phenylene)bis(3,3,6,6-tetramethyl-3,4,5,6,7,9-hexahydro-1H-xanthene-1,8 (2H)-dione (Entry 4) IR (cm⁻¹): 2957, 2871, 1665, 1620, 1365, 1201, 1167, 1144, 1005, 590. ¹H NMR: 7.07 (s, 4H), 4.70 (s, 2H), 2.42 (s, 8H), 2.17 (s, 8H), 1.07 (s, 12H), 0.97 (s, 12H). m/z: ¹³C NMR: 196.4, 162.4, 141.7, 127.9, 115.7, 50.8, 40.8, 32.2, 30.1, 28.9, 27.7. m/z: 645.4 (M + Na), 273.2.

3,3,6,6-tetramethyl-9-(4-hydroxy-3-methoxy-phenyl)-1,8-dioxo-octahydroxanthene (Entry 5) IR (cm⁻¹): 3412, 3029, 2954, 1666, 1622, 1514, 1431, 1359, 1278, 1229, 1197, 1135, 1028, 624, 571. ¹H NMR: 7.01 (s, 1H), 6.73 (d, 1H), 6.56 (dd, 1H), 5.49 (bs, 1H, -OH), 4.66 (s, 1H), 3.89 (s, 3H), 2.45 (s, 4H), 2.20 (q, 4H), 1.00 (s, 12H). ¹³C NMR: 196.6, 162.0, 145.8, 144.0, 136.4, 120.0, 115.8, 113.9, 112.2, 55.8, 50.7, 40.8, 32.2, 31.3, 29.3, 27.2. m/z: 396.47, 419.3 (M + Na), 273.2.

3,3,6,6-tetramethyl-9-(4-methoxy-phenyl)-1,8-dioxo-octahydroxanthene (Entry 7) IR (cm⁻¹): 3059, 2958, 2876, 1665, 1626, 1511, 1462, 1357, 1260, 1193, 1109, 1031, 841, 569. ¹H NMR: 7.18 (d, 2H), 6.75 (d, 2H), 4.69 (s, 1H), 3.73 (s, 3H), 2.45 (s, 4H), 1.09 (s, 6H), 0.99 (s, 6H). ¹³C NMR: 196.5, 162.0, 157.9, 136.4, 129.3, 115.7, 113.4, 55.1, 50.7, 40.8, 32.2, 30.9, 29.2, 27.3. m/z: 381.2, 273.2.

3. Result and Discussion

To a mixture of 4-chloro benzaldehyde (1 mmol) and 5,5-dimethyl-1,3-cyclohexanedione (2 mmol) and [bmim]ClO₄ (4 mmol) was added in 50 ml round bottom flask and was stirred at 100°C on heating magnetic stirrer. The progress of the reaction was monitored by TLC After completion of the reaction the reaction mixture was cooled to room temperature and water (5 ml) was added, solid separated was filtered and product was characterized by IR, NMR, ¹³C-NMR and mass.

The reaction between dimedone and 4-chloro benzaldehyde in the presence of [bmim]ClO₄ was studied as a model reaction. To optimize the reaction conditions of reaction temperature we have tried the study of above model reaction from temperature 60 - 110°C in the gap of 10°C. The best result in terms of time and yield for the reaction has been observed at 100°C (**Table 2**).

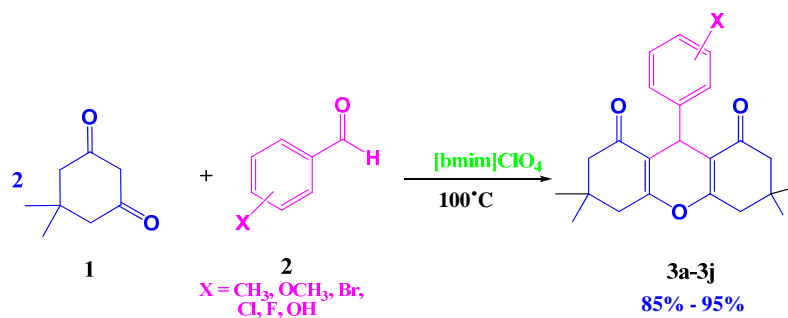
Scheme 1. Synthesis of 1,8-dioxo-octahydroxanthene using [bmim]ClO₄ ionic liquid.

Table 1. Comparison table with various reported catalysts.

Entry	Catalyst used	Time	Solvent	Temperature	Yield (%)	Ref.
1	Tetrachloro silane	3 hrs	CH ₂ Cl ₂	60 - 70°C	92	[20]
2	SmCl ₃	8 hrs	No solvent	120°C	98	[29]
3	Cellulose sulfonic acid	6 hrs	No solvent	110°C	95	[30]
4	Heteropoly acid supported MCM-41	5 hrs	Ethanol	90°C	90	[31]
5	[Bmim]HSO ₄	3.5 hrs	No solvent	80°C	76 - 94	[32]
6	[Hbim]BF ₄	45 min.	Methanol	Ultrasonic 25 - 30°C	85	[34]
7	Trimethyl silyl chloride	8 - 10 hrs	Acetonitrile	Reflux	72 - 84	[50]
8	[Bmim]ClO ₄	40 - 90 min.	----	100°C	85 - 95	Present work

Table 2. Optimization of reaction temperature for synthesis of 1,8-dioxo-octahydroxanthenes.

Entry	Temperature (in °C)	Time (in hrs.)	Yields (in %)
1	60	8 hrs	55
2	70	6 hrs	70
3	80	5 hrs	75
4	90	2.5 hrs	85
5	100	1 hr	94
6	110	1 hr	94

To evaluate the scope of this reaction a range of 1, 8-dioxo-octahydroxanthenes were prepared by the reaction of dimedone and aromatic aldehydes under optimized reaction conditions. The results are summarized in the table (Table 3). Various aromatic aldehydes with electron donating substituent's reacted efficiently and quickly with dimedone to give cyclocondensation products in high yields over short reaction times. Similarly aldehydes with electron withdrawing substituent's reacted efficiently to yield the corresponding 1,8-dioxo-octahydroxanthenes.

Study of reuse of the ionic liquid can be easily done up to 3 times with excellent yields without loss in the reactivity (as shown in graph/Figure 1). After 3rd cycle we observed sudden decrease in the 4th reuse of ionic liquid thus the ionic liquid used was recycled for 3 times. The structure of isolated products was assigned based on their

Table 3. Green protocol for efficient synthesis of 1, 8-dioxo-octahydroxanthenes using [bmim]ClO₄.^a

Entry	Aldehyde	Product ^b	Time (min.)	Yield ^c (%)	Mp. (°C)
1	4-Chloro benzaldehyde	3a	60	94	215 - 217
2	4-Methyl benzaldehyde	3b	70	95	217 - 218
3	Benzaldehyde	3c	40	92	205 - 206
4	Terephthaldehyde	3d	90	88	>300
5	Vanillin	3e	50	85	225 - 227
6	4-Fluro benzaldehyde	3f	50	92	232 - 233
7	4-Methoxy benzaldehyde	3g	60	95	241 - 243
8	3-Chloro benzaldehyde	3h	60	90	183 - 184
9	4-Bromo benzaldehyde	3i	60	90	222 - 224
10	4-Hydroxy benzaldehyde	3j	70	87	228 - 230

^aReaction condition: aromatic aldehydes (2 mmol), 5,5-dimethyl-1-3-cyclohexanedione (4 mmol), [bmim]ClO₄ (4 mmol) at 100°C. ^bAll the product were characterized by IR spectral data and comparison of their melting point with those of the authentic samples. Also the structures of the some products were confirmed by ¹H NMR spectral data. ^cIsolated yield.

spectral analyses as well as by matching with their melting point with reported analogous.

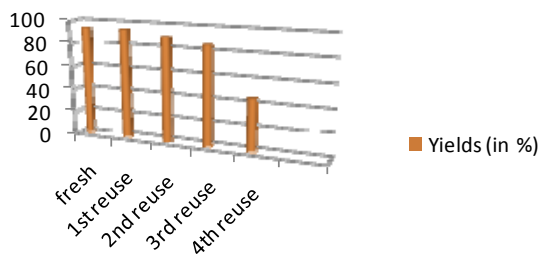


Figure 1. Study of reuse of ionic liquid.

4. Conclusion

In summary, we have developed a high-yielding simple, convenient, straight-forward and practical one-pot procedure for the synthesis of different types of 1,8-dioxo-octahydroxanthenes derivatives using [bmim]ClO₄. The ionic liquid used can be easily recovered and reused 3 times without significant decrease in the yield of the product.

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