

 Open access • Journal Article • DOI:10.1116/1.569972

Abstract: Chemisorption of oxygen on the (110) surface of iridium — [Source link](#)

D.E. Ibbotson, J. L. Taylor, W. H. Weinberg

Published on: 01 Mar 1979 - Journal of Vacuum Science and Technology (American Vacuum Society)

Topics: Chemisorption, Auger electron spectroscopy, Thermal desorption, Chemical state and Iridium

Related papers:

- [The chemisorption of oxygen on the \(110\) surface of iridium](#)
- [Auger electron spectroscopic study of chemisorption of oxygen to \(111\)-oriented gold surfaces](#)☆
- [Chemisorption of H₂O on the Si\(111\) 7 × 7 surfaces](#)
- [Dissociative chemisorption of oxygen on the Ru\(001\) surface: Spectroscopic identification of precursor intermediates at low surface temperatures](#)
- [Thermal Desorption of 15 NO Chemisorbed on Polycrystalline Rhenium Surfaces](#)

Share this paper:    

View more about this paper here: <https://typeset.io/papers/abstract-chemisorption-of-oxygen-on-the-110-surface-of-35pwdjmex1>

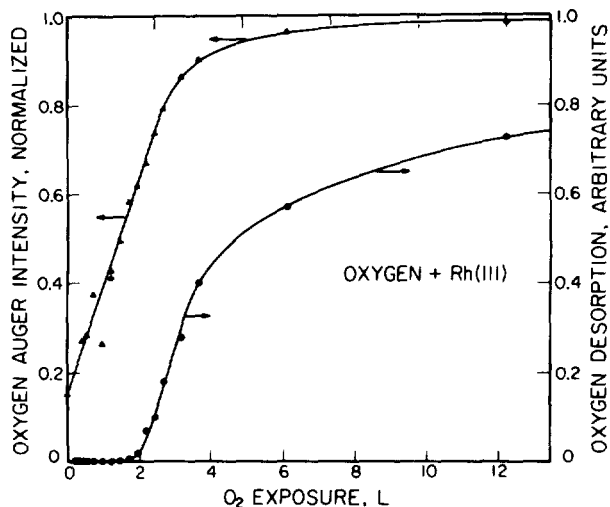


FIG. 3. Oxygen thermal desorption peak areas and oxygen differential Auger peak intensities (510 eV) as a function of O_2 exposure at $T \leq 170$ K.

On the basis of the Auger data of Fig. 2, at least 40% of the saturation coverage of oxygen must adsorb before O_2 desorption is observed. Thus, it appears that competition between two modes of oxygen depletion occurs during thermal desorption of oxygen: Dissolution into the bulk and desorption as molecular oxygen, with dissolution being the predominant depletion mode at low coverages. Conrad *et al.*³ have also

reported evidence for dissolution of oxygen into the bulk of Pd(111).

Analysis of the oxygen thermal desorption data, assuming second order kinetics, yields values of 56 ± 2 kcal/mol and 2.5×10^{-3} cm² s⁻¹ for the activation energy of desorption and the preexponential factor of the desorption rate coefficient, respectively, in the limit of zero coverage. At nonzero coverages the desorption data are complicated by contributions from multiple states and/or interadsorbate interactions. A value for the initial sticking probability of 0.2 was determined from the Auger data shown in Fig. 3, where the constant rate of adsorption at exposures below 2 L implies a mobile precursor model for oxygen adsorption at $T \leq 170$ K.

Acknowledgment: Research was supported by the Army Research Office under Grant No. DAHC04-75-0170.

^aComplete paper to appear in *Surface Science*.

^bNational Science Foundation Predoctoral Fellow.

^cSherman Fairchild Distinguished Scholar. Permanent Address: National Bureau of Standards, Washington, D.C. 20234.

^dAlfred P. Sloan Foundation Fellow, and Camille and Henry Dreyfus Foundation Teacher-Scholar.

¹V. P. Ivanov, G. K. Boreskov, V. I. Savchenko, W. F. Egelhoff, Jr. and W. H. Weinberg, *Surf. Sci.* **61**, 25 (1976).

²R. Klein and A. Shih, *Surf. Sci.* **69**, 403 (1977).

³H. Conrad, G. Ertl, J. Küppers, and E. E. Latta, *Surf. Sci.* **65**, 245 (1977).

Abstract: Chemisorption of oxygen on the (110) surface of iridium

D. E. Ibbotson, J. L. Taylor,^a and W. H. Weinberg^b

Division of Chemistry and Chemical Engineering, California Institute of Technology, Pasadena, California 91125

(Received 27 September 1978; accepted 6 November 1978)

PACS numbers: 82.65.My, 82.65.Jv, 73.40.Cg, 61.14.Hg

The chemisorption of oxygen on Ir(110) has been investigated under ultrahigh vacuum conditions with thermal desorption mass spectrometry (TDS), contact potential difference measurements, Auger electron spectroscopy, and LEED. Oxygen may adsorb in three distinct chemical states: a molecularly chemisorbed species that is stable below 100 K, a dissociatively chemisorbed species, and a surface oxide that forms rapidly above 700 K.

The oxide layer saturates at a coverage of 0.25 ML (1 ML = 9.6×10^{14} atoms/cm²) and forms with an estimated activation energy of 45 kcal/mol, assuming first order kinetics and a preexponential factor of 10^{13} s⁻¹. Two forms of the surface oxide are observed, a disordered phase and an ordered phase which corresponds to a (1 × 4) LEED pattern. Both forms exhibit identical chemical properties. Formation of the

oxide layer through thermal activation causes the total saturation coverage to increase from 1 ML on the clean surface to 1.25 ML on the oxidized surface, as evidenced from TDS.

Different LEED patterns of dissociatively chemisorbed oxygen are observed on clean and oxidized Ir(110). A $p(2 \times 2)$ structure forms on the clean surface near 0.25 ML coverage, whereas a $c(2 \times 2)$ structure forms on the oxidized surface near 0.5 ML. Thermal annealing is necessary to order the $c(2 \times 2)$ in the [001] azimuth with an energy of diffusion of 47 kcal/mol, assuming first order kinetics and a preexponential factor of 10^{12} s⁻¹. Slow ordering of the $p(2 \times 2)$ in the [001] azimuth at 300 K is observed on the clean surface with an estimated diffusional energy of 20 kcal/mol, again assuming the same kinetics as above.

Oxygen desorbs molecularly from Ir(110) with an activation energy of 45–70 kcal/mol, decreasing continuously with increasing coverage. The adsorption kinetics are described by a second-order precursor model for surface temperatures between 300 and 700 K. The difference between the activation energies for the desorption and diffusion of the precursor in this model is 2.1 kcal/mol. Oxygen chemisorption is not activated since the initial sticking probabilities on the clean and oxidized Ir(110) surfaces are equal to 0.28 and 0.4, irrespective of the surface temperature. The dipole moment and polarizability of dissociatively chemisorbed oxygen have been calculated from contact potential difference measurements and adsorption kinetics for clean and oxidized Ir(110). The dipole moment and polarizability change at 0.25 ML coverage on the clean surface and at 0.5 ML coverage on the oxidized surface. These changes are associated with the completion of overlayer structures, noted above, and a shift in binding site from chemisorption on top of the rows at low coverages to chemisorption between the rows at high coverages on the clean and oxidized surfaces. Although the dipole moment for any coverage is independent of temperature, the polarizability is inversely proportional to temperature.

A third chemical state of oxygen that is stable below 100

K is inferred from contact potential difference (CPD) measurements. If an overlayer of oxygen that is chemisorbed at 90 K is annealed to 300 K, the CPD changes irreversibly indicating that oxygen may chemisorb molecularly below 100 K. No evidence for desorption is observed. The activation energy for the dissociation of molecular oxygen is 8 kcal/mol, whereas the desorption energy is at least 10 kcal/mol. This cannot be associated with diffusion since the activation energy for diffusion is 20 kcal/mol in the [001] direction and 1 kcal/mol in the [110] direction. On clean and oxidized Ir(110), the dipole moment (per atom) of molecular oxygen is equal to that of dissociatively chemisorbed oxygen in the low coverage state on the respective surfaces. This suggests that molecular oxygen chemisorbs with the molecular axis parallel to the surface and occupies two adjacent sites above the top-most row of surface atoms.

Acknowledgment: Work was supported by the National Science Foundation under Grant No. DMR77-14976.

^aComplete paper *Surf. Sci.* **79**, 349 (1979).

^bNational Science Foundation Predoctoral Fellow.

^cAlfred P. Sloan Foundation Fellow, and Camille and Henry Dreyfus Foundation Teacher-Scholar.