



J. Chem. Pharm. Res., 2010, 2(5):626-641

ISSN No: 0975-7384
CODEN(USA): JCPRC5

Adsorption dynamics of methylene blue by acid activated carbon

S Arivoli¹, M. Hema², S Parthasarathy³ and N Manju⁴

¹*Department of Chemistry, Thiru Vi Ka Government Arts College, Thiruvarur, India*

²*Department of Chemistry, Anna University of Technology, Tiruchirappalli, Thirukkuvallai Campus, India*

Department of Chemistry, CARE School of Engineering, Thiruchirappalli, India

⁴*Department of Chemistry, CARE School of Engineering and Technology, Thiruchirappalli, India*

ABSTRACT

A carbonaceous adsorbent prepared from an indigenous waste by acid treatment was tested for its efficiency in removing Methylene blue. The parameters studied include agitation time, initial dye concentration, carbon dose, pH and temperature. The adsorption followed first order reaction equation and the rate is mainly controlled by intra-particle diffusion. Freundlich and Langmuir isotherm models were applied to the equilibrium data. The adsorption capacity (Q_m) obtained from the Langmuir isotherm plots were 60.61, 59.95, 62.11 and 62.50 mg/g respectively at an initial pH of 7.0 at 30, 40, 50 and 60^oC. The temperature variation study showed that the Methylene blue adsorption is endothermic and spontaneous with increased randomness at the solid solution interface. Significant effect on adsorption was observed on varying the pH of the Methylene blue solutions. Almost 68% removal of Methylene blue was observed at 60^oC. The Langmuir and Freundlich isotherms obtained, positive ΔH^0 value, pH dependent results and desorption of dye in mineral acid suggest that the adsorption of Methylene blue on MCC involves physisorption mechanism.

Key words: Activated carbon (MCC), Methylene blue (MB), Adsorption isotherm, Equilibrium, Kinetic and Thermodynamic parameters, Intraparticle diffusion, Regeneration pattern.

INTRODUCTION

The concentration and effective utilization of activated carbon generated from natural plant material have attracted worldwide attention in a view of the large disposal problem without detriment to environment. Many investigators have studied the feasibility of using inexpensive alternative materials like pearl millet husk, date pits, saw dust buffing dust of leather industry, coir pith, crude oil residue tropical grass, olive stone and almond shells, pine bark, wool waste, coconut shell etc., as carbonaceous precursors for the removal of dyes from water and wastewater [1,2,3].

The present study undertaken to evaluate the efficiency of a carbon adsorbent prepared from acid activated **morinda coreia buch-ham bark carbon (MCC)** for the removal of Methylene blue dye in aqueous solution. In order to design adsorption treatment systems, knowledge of kinetic and mass transfer processes is essential. In this paper, we have reported the applicability of kinetic and mass-transfer models for the adsorption of Methylene blue onto acid activated carbon.

EXPERIMENTAL SECTION

Materials

2.1 Adsorbent Materials

The dried **morinda coreia buch-ham bark** was carbonized with concentrated sulphuric acid in the weight ratio of 1:1 (w/v). Heating for six hours in a furnace at 400°C has completed the carbonization and activation. The resulting carbon was washed with distilled water until a constant pH of the slurry was reached. Then the carbon was dried for four hours at 120°C in a hot air oven. The dried material was ground well to a fine powder and sieved.

Adsorption dynamic experiments

3.1 Batch equilibration method

The adsorption experiments were carried out in a batch process at 30, 40, 50 and 60° C temperatures. The known weight of adsorbent material was added to 50 ml of the dye solutions with an initial concentration of 20 mg/L to 100 mg/L. The contents were shaken thoroughly using a mechanical shaker rotating with a speed of 120 rpm. The solution was then filtered at preset time intervals and the residual dye concentration was measured.

3.2 Effect of variable parameters

3.2.1 Dosage of adsorbents

The various doses of the adsorbents are mixed with the dye solutions and the mixture was agitated in a mechanical shaker. The adsorption capacities for different doses were determined at definite time intervals by keeping all other factors constant.

3.2.2 Initial concentration of dye

In order to determine the rate of adsorption, experiments were conducted with different initial concentrations of dyes ranging from 20 to 100 mg/L. All other factors have kept constant.

3.2.3 Contact time

The effect of period of contact on the removal of the dye on adsorbent in a single cycle was determined by keeping particle size, initial concentration, dosage, pH and concentration of other ions constant.

3.2.4 pH

Adsorption experiments were carried out at pH 3 - 10. The acidic and alkaline pH of the media was maintained by adding the required amounts of dilute hydrochloric acid and sodium hydroxide solutions. The parameters like particle size of the adsorbents, dye concentration, dosage of the adsorbent and concentration of other ions have kept constant while carrying out the experiments. The pH of the samples was determined using a portable pH meter, Systronics make. The pH meter was calibrated with 4.0 and 9.2 buffers.

3.2.5 Chloride

The experiments were done in the presence of varying chloride environments using various sodium chloride solutions. While doing the experiments, the absence of other anions has ensured.

3.2.6 Temperature

The adsorption experiments were performed at four different temperatures viz., 30, 40, 50 and 60°C in a thermostat attached with a shaker, Remi make. The constancy of the temperature was maintained with an accuracy of $\pm 0.5^\circ\text{C}$.

3.2.7 Zero point charge

The pH at the potential of zero charge of the carbon (pH_{ZPC}) was measured using the pH drift method [4]. The pH of the solution was adjusted by using 0.01 M sodium hydroxide or hydrochloric acid. Nitrogen was bubbled through the solution at 25°C to remove the dissolved carbon dioxide. 50 mg of the activated carbon was added to 50 ml of the solution. After stabilization, the final pH was recorded. The graphs of final pH versus initial pH used to determine the zero point charge of the activated carbon.

3.2.8 Titration studies

According to Boehm [4] only strong acidic carboxylic acid groups are neutralized by sodium bicarbonate, where as those neutralized by sodium carbonate are thought to be lactones, lactol and carboxyl group. The weakly acidic phenolic groups only react with strong alkali, sodium hydroxide. Therefore, by selective neutralization using bases of different strength, the surface acidic functional group in carbon can be characterized both quantitatively and qualitatively. Neutralization with hydrochloric acid characterizes the amount of surface basic groups that are, for example, pyrones and chromenes. The basic properties have described to surface basic groups and the pi electron system of carbon basal planes. The results indicate that the activated carbon used may possess acidic oxygen functional group on their surface and this is supported well by their respective zero point charge values. The results obtained from the above characterization studies are given in the Table 1.

3.2.9 Desorption studies

Desorption studies help to elucidate the nature of adsorption and recycling of the spent adsorbent and the dye. The effect of various reagents used for desorption are studied.

RESULTS AND DISCUSSION

4.1 Characterization of the adsorbent

Activated carbons are a widely used adsorbent due to its high adsorption capacity, high surface area, micro porous structure and high degree of surface respectively. The wide usefulness of carbon is a result of their specific surface area, high chemical and mechanical stability. The chemical nature and pore structure usually determines the sorption activity. The physico chemical properties of the chosen adsorbent are listed in Table 1

Table 1-Characteristics of the adsorbent

Properties	MCC
Particle size (mm)	0.055
Density (g/cc)	0.3575
Moisture content (%)	2.25
Loss on ignition (%)	82
Acid insoluble matter (%)	2.5
Water soluble matter (%)	0.32
pH of aqueous solution	6.85
pH _{zpc}	6.35
Surface groups (m equiv/g)	
i) Carboxylic acid	0.325
ii) Lactone, lactol	0.049
iii) Phenolic	0.061
iv) Basic (pyrones and chromenes)	0.036

4.2 Effect of contact time and initial dye concentration

The experimental results of adsorptions of at various concentrations (20-100 mg/L) with contact time are shown in representative Figure 1. The equilibrium data were collected in Table 2 reveals that, percent adsorption decreased with increase in initial dye concentration, but the actual amount of Methylene blue dye adsorbed per unit mass of carbon increased with increase in dye concentration. It means that the adsorption is highly dependent on initial concentration of dye. It is because of that at lower concentration, the ratio of the initial number of Methylene blue dye molecules to the available surface area is low subsequently the fractional adsorption becomes independent of initial concentration. However, at high concentration the available sites of adsorption becomes fewer and hence the percentage removal of Methylene blue dye is dependent upon initial concentration [5,6]. Equilibrium have established at 40 minutes for all concentrations. Figure 1 reveals that the curves are single, smooth, and continuous, leading to saturation, suggesting the possible monolayer coverage of the Methylene blue dye on the carbon surface.

Table 2- Equilibrium parameters for the adsorption of Dye onto activated Carbon

[MB] ₀	C _e (mg/L)				Q _e (mg/g)				Dye removed (%)			
	30°	40°	50°	60°	30°	40°	50°	60°	30°	40°	50°	60°
20	2.2578	2.0142	1.8529	1.6434	17.7422	17.9858	18.1470	18.3560	88.7	89.9	90.7	91.7
40	10.5752	10.1849	9.7487	9.3544	29.4248	29.8150	30.2500	30.6450	73.5	74.5	75.6	76.6
60	23.1572	22.5127	21.9092	21.3157	36.8428	37.4870	38.0900	38.6840	61.4	62.4	63.4	64.4
80	34.5942	363.7849	32.9802	32.1799	45.4000	46.4870	47.0190	47.8200	56.7	58.1	58.7	59.7
100	46.3415	45.2987	44.3014	43.2892	53.6500	53.6580	55.6980	56.7100	53.6	53.7	55.6	56.7

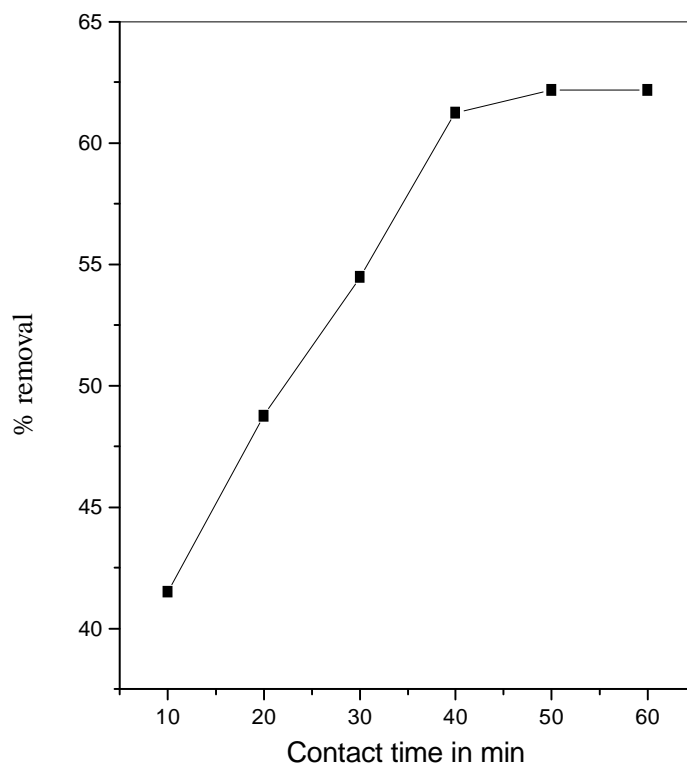


Fig.1-Effect of contact time on the adsorption of methylene blue onto MCC
 [MG]=60mg/L; Adsorbent dose=50 mg/50ml; pH=7

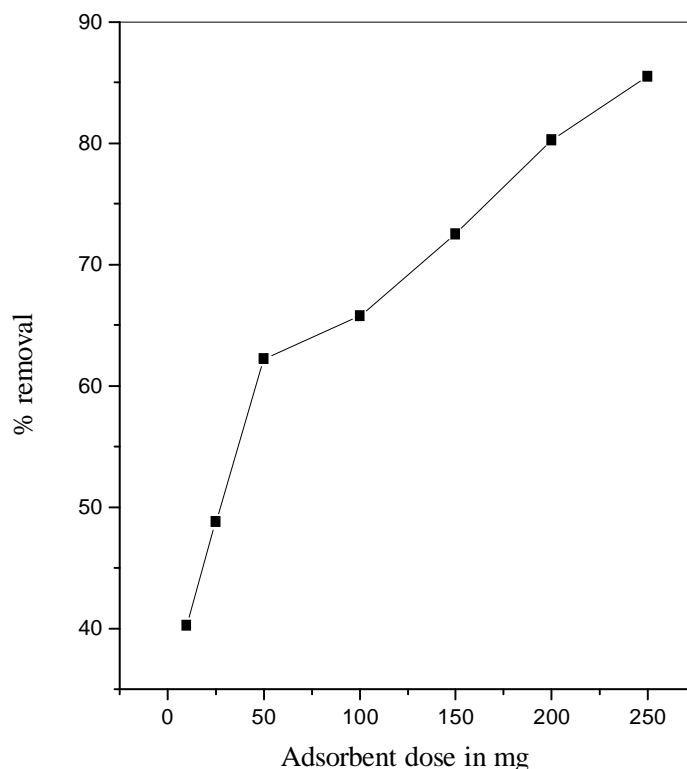


Fig.2-Effect adsorbent dose on the adsorption of methylene blue onto MCC
[MG]=60mg/L;Contact time=60 min;pH=7

4.3 Effect of carbon concentration:

The adsorption of the Methylene blue dye on carbon was studied by varying the carbon concentration (10-250 mg/50ml) for 60 mg/L of dye concentration. The percent adsorption increased with increase in the carbon concentration (Figure 2). This was attributed to increased carbon surface area and availability of more adsorption sites [5,6]. Hence the entire studies are carried out with the adsorbent dose of 50 mg of adsorbent /50 ml of the adsorbate solution.

4.4 Adsorption isotherm

The experimental data analyzed according to the linear form of the Langmuir and Freundlich isotherms [7,8]. The Langmuir isotherm represented by the following equation

$$C_e/Q_e = 1/Q_m b + C_e/Q_m$$

Where C_e is the equilibrium concentration (mg/L), Q_e is the amount adsorbed at equilibrium (mg/g) and Q_m and b is Langmuir constants related to adsorption efficiency and energy of adsorption, respectively. The linear plots of C_e/Q_e versus C_e suggest the applicability of the Langmuir isotherms (Figure 3). The values of Q_m and b were determined from slope and intercepts of the plots and are presented in Table 3. From the results, it is clear that the value of adsorption efficiency Q_m and adsorption energy b of the carbon increases on increasing the

temperature. From the values we can conclude that the maximum adsorption corresponds to a saturated monolayer of adsorbate molecules on adsorbent surface with constant energy and no transmission of adsorbate in the plane of the adsorbent surface. The observed b values shows that the adsorbent prefers to bind acidic ions and that speciation predominates on sorbent characteristics, when ion exchange is the predominant mechanism takes place in the adsorption of Methylene blue, it confirms the endothermic nature of the process involved in the system [9,10,11]. To confirm the favourability of the adsorption process, the separation factor (R_L) was calculated and presented in Table 4. The values were found to be between 0 and 1 and confirm that the ongoing adsorption process is favorable [12].

The Freundlich equation was employed for the adsorption of Methylene blue on the adsorbent. The Freundlich isotherm was represented by

$$\log Q_e = \log K_f + 1/n \log C_e$$

Where Q_e is the amount of Methylene blue dye adsorbed (mg/g), C_e is the equilibrium concentration of dye in solution (mg/L) and K_f and n are constants incorporating the factors affecting the adsorption capacity and intensity of adsorption, respectively. Linear plots of $\log Q_e$ versus $\log C_e$ shows that the adsorption of Methylene blue dye obeys the Freundlich adsorption isotherm (Figure 4). The values of K_f and n given in the Table 5 shows that the increase in negative charges on the adsorbent surface that makes electrostatic force like Vanderwaal's between the carbon surface and dye ion. The molecular weight, size and radii either limit or increase the possibility of the adsorption of the dye onto adsorbent. However, the values clearly show the dominance in adsorption capacity. The intensity of adsorption is an indicative of the bond energies between dye and adsorbent and the possibility of slight chemisorptions rather than physisorption [10,11]. However, the multilayer adsorption of Methylene blue through the percolation process may be possible. The values of n is greater than one indicating the adsorption is much more favorable [12].

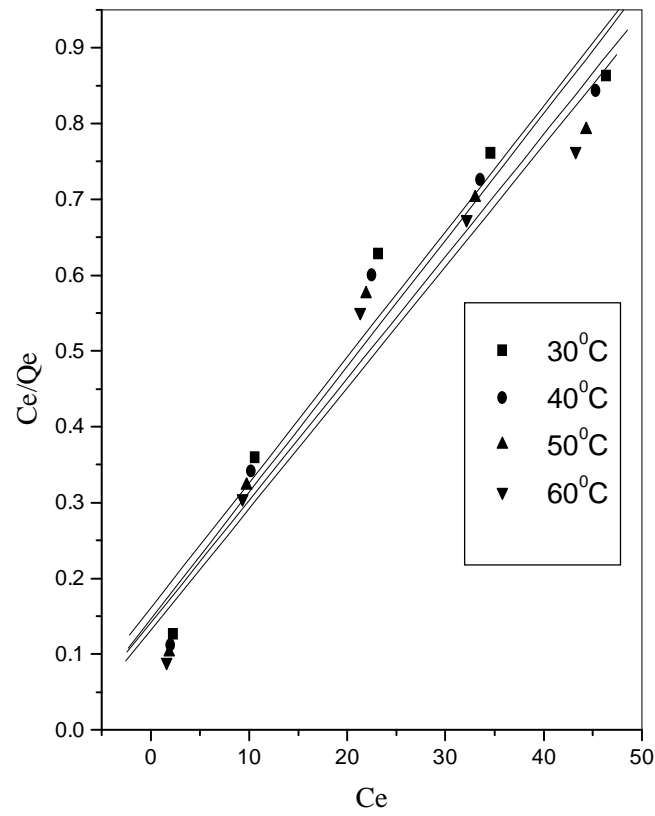


Fig.3-Langmuir isotherm for the adsorption of methylene blue by MCC

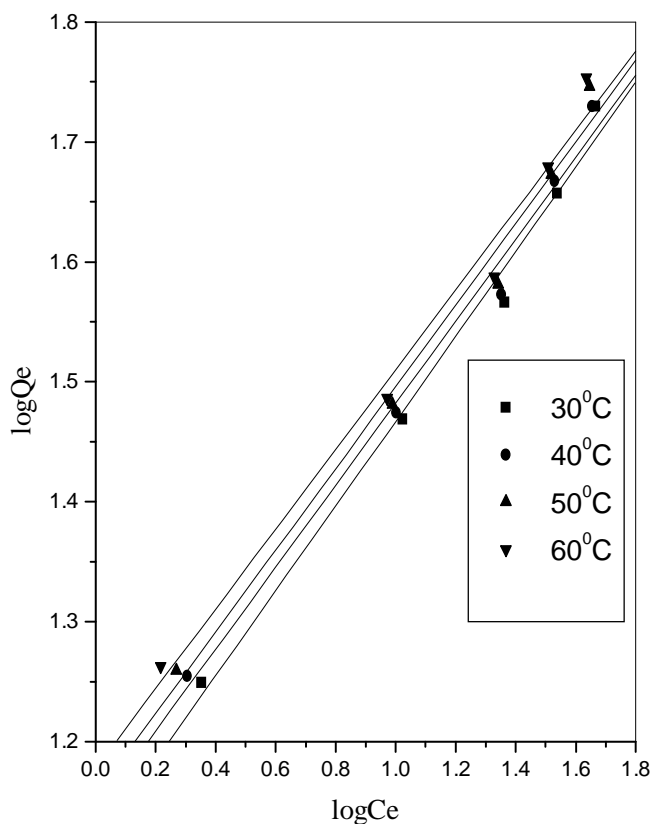


Fig.4-Freundlich isotherm for the adsorption of methylene blue onto MCC

Table 3 - Langmuir isotherm results

Dye	Temp °C	Statistical parameters/constants		
		r^2	Q_m	b
MB	30	0.9758	60.61	0.1078
	40	0.9787	59.95	0.1141
	50	0.9756	62.11	0.1137
	60	0.9737	62.50	0.1218

Table 4 Dimensionless Separation factor (R_L)

[MB] ₀ (mg/L)	Temperature (°C)			
	30	40	50	60
20	0.316	0.304	0.305	0.290
40	0.189	0.179	0.181	0.171
60	0.134	0.127	0.128	0.120
80	0.104	0.098	0.099	0.093
100	0.084	0.080	0.081	0.076

Table 5 - Freundlich isotherm results

Dye	Temp °C	Statistical parameters/constants		
		r ²	K _f	n
MB	30	0.9937	1.4239	2.8296
	40	0.9941	1.4070	2.9282
	50	0.9925	1.4055	2.9377
	60	0.9916	1.3971	2.9904

4.5 Effect of temperature

The adsorption capacity of the carbon increased with increase in the temperature of the system from 30°-60°C. Thermodynamic parameters such as change in free energy (ΔG°) (kJ/mol), enthalpy (ΔH°)(kJ/mol) and entropy (ΔS°)(J/K/mol) were determined using the following equations

$$K_0 = C_{\text{solid}}/C_{\text{liquid}}$$

$$\Delta G^\circ = -RT \ln K_0$$

$$\log K_0 = \Delta S^\circ / (2.303RT) - \Delta H^\circ / (2.303RT)$$

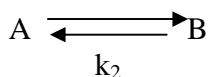
Where K_0 is the equilibrium constant, C_{solid} is the solid phase concentration at equilibrium (mg/L), C_{liquid} is the liquid phase concentration at equilibrium (mg/L), T is the temperature in Kelvin and R is the gas constant. The ΔH° and ΔS° values obtained from the slope and intercept of Van't Hoff plots have presented in Table 6. The values ΔH° are with in the range of 1 to 93 KJ/mol indicates the physisorption. From the results we could make out that physisorption is much more favorable for the adsorption of Methylene blue. The positive values of ΔH° show the endothermic nature of adsorption and it governs the possibility of physical adsorption [11,13]. Because in the case of physical adsorption, while increasing the temperature of the system, the extent of Methylene blue dye adsorption increases, this rules out the possibility of chemisorption [13]. The low ΔH° value depicts dye is physisorbed onto adsorbent MCC.

The negative values of ΔG° (Table 6) shows the adsorption is highly favorable and spontaneous. The positive values of ΔS° (Table 6) shows the increased disorder and randomness at the solid solution interface of Methylene blue with MCC adsorbent, while the adsorption there are some structural changes in the dye and the adsorbent occur. The adsorbed water molecules, which have displaced by the adsorbate species, gain more translational entropy than is lost by the adsorbate molecules, thus allowing the prevalence of randomness in the system. The enhancement of adsorption capacity of the activated carbon at higher temperatures was attributed to the enlargement of pore size and activation of the adsorbent surface [12,13,14].

4.6 Kinetics of adsorption

The kinetics of sorption describes the solute uptake rate, which in turn governs residence time or sorption reaction. It is one of the important characteristics in defining the efficiency or sorption. In the present study, the kinetics of the dye removal was carried out to understand the behaviour of these low cost carbon adsorbents. The adsorption of dye from an aqueous follows reversible first order kinetics, when a single species are considered on a heterogeneous surface. The heterogeneous equilibrium between the dye solutions and the activated carbon are expressed as

$$k_1$$



Where k_1 is the forward rate constant and k_2 is the backward rate constant. A represents Methylene blue dye remaining in the aqueous solution and B represent Methylene blue dye adsorbed on the surface of activated carbon. The equilibrium constant (K_0) is the ration of the concentration adsorbate in adsorbent and in aqueous solution ($K_0=k_1/k_2$).

In order to study the kinetics of the adsorption process under consideration the following kinetic equation proposed by Natarajan and Khalaf as cited in literature has been employed [1].

$$\log C_0/C_t=(K_{ad}/2.303)t$$

Where C_0 and C_t are the concentration of the dye in (in mg/L) at time zero and at time t, respectively. The rate constants (K_{ad}) for the adsorption processes have been calculated from the slope of the linear plots of $\log C_0/C_t$ versus t for different concentrations and temperatures. The determination of rate constants as described in literature given by

$$K_{ad}=k_1+k_2=k_1+(k_1/K_0)=k_1[1+1/K_0]$$

The overall rate constant k_{ad} for the adsorption of dye at different temperatures are calculated from the slopes of the linear Natarajan-Khalaf plots. The rate constant values are collected in Table 7 shows that the rate constant (k_{ad}) increases with increase in temperature suggesting that the adsorption process in endothermic in nature. Further, k_{ad} values decrease with increase in initial concentration of the dye. In cases of strict surface adsorption a variation of rate should be proportional to the first power of concentration. However, when pore diffusion limits the adsorption process, the relationship between initial dye concentration and rate of reaction will not be linear. Thus, in the present study pore diffusion limits the overall rate of dye adsorption. The over all rate of adsorption is separated into the rate of forward and reverse reactions using the above equation. The rate constants for the forward and reverse processes are also collected in Table 7 indicate that, at all initial concentrations and temperatures, the forward rate constant is much higher than the reverse rate constant suggesting that the rate of adsorption is clearly dominant [1,11,13].

Table 6 Equilibrium constant and thermodynamic parameters for the adsorption of Methylene blue dyes onto carbon

[M B] ₀	K ₀				ΔG°				ΔH°	ΔS°
	Temperature (°C)									
	30°	40°	50°	60°	30°	40°	50°	60°		
20	7.85	8.92	9.79	11.16	-5.70	-4.39	-6.12	-6.68	13.72	68.27
40	2.78	2.92	3.10	3.27	-2.57	-2.79	-3.04	-3.28	9.51	48.51
60	1.59	1.66	1.74	1.81	-1.17	-1.32	-1.48	-1.64	6.20	33.65
80	1.31	1.37	1.42	1.48	-0.68	-0.83	-0.95	-1.09	4.73	26.51
100	1.15	1.18	1.25	1.31	-0.36	-0.44	-0.62	-0.74	4.39	24.01

Table 7 - Rate constants for the adsorption of Methylene blue dye ($10^3 k_{ad}, \text{min}^{-1}$) and the constants for forward ($10^3 k_1, \text{min}^{-1}$) and reverse ($10^3 k_2, \text{min}^{-1}$) process

[D] ₀	Temperature (⁰ C)											
	k_{ad}				30		40		50		60	
	30 ^o	40 ^o	50 ^o	60 ^o	k_1	k_2	k_1	k_2	k_1	k_2	k_1	k_2
20	369.	454.	528.	653.	328.	41.7	408.	45.7	479.	48.9	599.	53.
	86	63	07	29	11	5	85	8	15	2	57	72
40	74.8	80.8	87.9	95.1	55.0	19.8	60.2	20.5	66.5	21.4	72.8	22.
	9	5	2	3	9	0	7	8	0	2	9	24
60	36.1	38.3	40.1	42.4	22.2	13.9	23.9	14.3	25.5	14.6	27.3	15.
	3	2	6	6	0	3	5	7	0	6	8	08
80	19.1	30.4	31.5	33.5	16.5	12.5	17.6	12.8	18.5	12.9	20.0	13.
	0	9	3	0	2	8	6	3	4	9	3	47
100	25.4	26.5	27.7	28.9	13.6	11.7	14.4	12.1	15.4	12.2	16.4	12.
	2	7	5	4	4	8	1	6	6	9	2	52

4.7 Intraparticle diffusion

The most commonly used technique for identifying the mechanism involved in the sorption process is by fitting the experimental data in an intraparticle diffusion plot. Previous studies by various researchers showed that the plot of Q_t versus $t^{0.5}$ represents multi linearity, which characterizes the two or more steps involved in the sorption process. According to Weber and Morris, an intraparticle diffusion coefficient K_p is defined by the equation:

$$K_p = Q/t^{0.5}$$

Thus the $K_p(\text{mg/g min}^{0.5})$ value can be obtained from the slope of the plot of $Q_t(\text{mg/g})$ versus $t^{0.5}$ for Methylene blue. From figure 5, it was noted that the sorption process tends to be followed by two phases. The two phases in the intraparticle diffusion plot suggest that the sorption process proceeds by surface sorption and intraparticle diffusion [15,16]. The initial curved portion of the plot indicates a boundary layer effect while the second linear portion is due to intraparticle or pore diffusion. The slope of the second linear portion of the plot has been defined as the intraparticle diffusion parameter $K_p(\text{mg/g min}^{0.5})$. On the other hand, the intercept of the plot reflects the boundary layer effect. The larger the intercept, the greater the contribution of the surface sorption in the rate limiting step. The calculated intraparticle diffusion coefficient K_p value was given by 0.300, 0.342, 0.389, 0.423 and 0.477 $\text{mg/g min}^{0.5}$ for initial Methylene blue dye concentration of 20,40,60,80 and 100 mg/L at 30^oC.

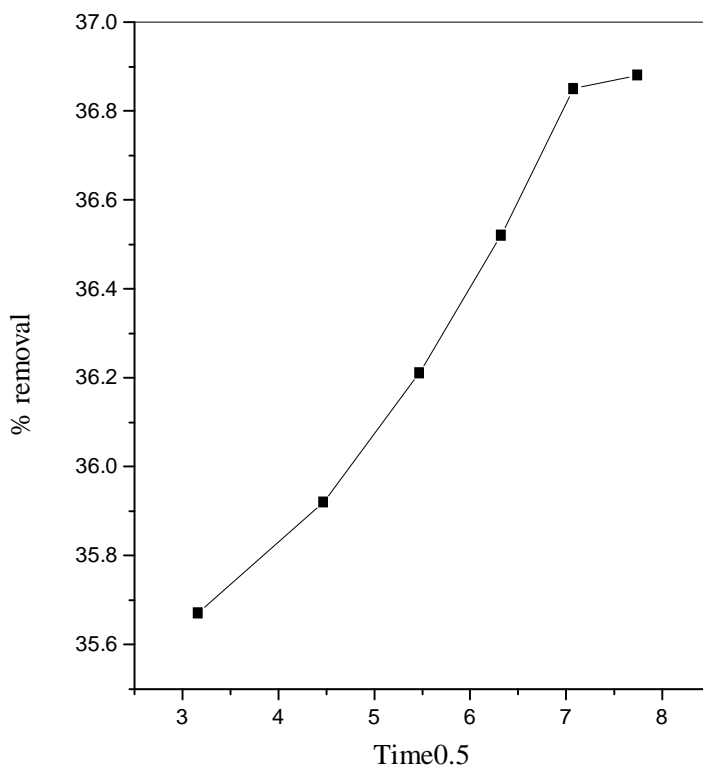


Fig.5-Intraparticle diffusion effect
on the adsorption of methylene blue onto MCC
[MG]= 60 mg/L; Contact time = 60 min; Adsorbent dose=50mg/50ml

4.8 Effect of pH

pH is one of the most important parameters controlling the adsorption process. The effect of pH of the solution on the adsorption of Methylene blue ions on MCC was determined. The result is shown in Fig. 6. The pH of the solution was controlled by the addition of HCl or NaOH. The uptake of Methylene blue ions at pH 8.0 was the minimum and a maximum in uptake was obtained at pH 3.00. However, when the pH of the solution was increased (more than pH 9), the uptake of Methylene blue ions was increased. It appears that a change in pH of the solution results in the formation of different ionic species, and different carbon surface charge. At pH values lower than 5, the Methylene blue ions can enter into the pore structure. At a pH value higher than 7, the zwitterions form of Methylene blue in water may increase the aggregation of Methylene blue to form a bigger molecular form (dimer) and become unable to enter into the pore structure of the carbon surface. The greater aggregation of the zwitterionic form is due to the attractive electrostatic interaction between the ionic groups of the monomer.

At a pH value higher than 9, the existence of MCC surface OH⁻ creates a competition between ionic dye and it will decrease the aggregation of Methylene blue, which causes an increase in the adsorption of Methylene blue ions on the carbon surface. The effect of the charge on the carbon surface and the electrostatic force of attraction and repulsion between the carbon surface and the Methylene blue ions cannot explain the outcome [12,17].

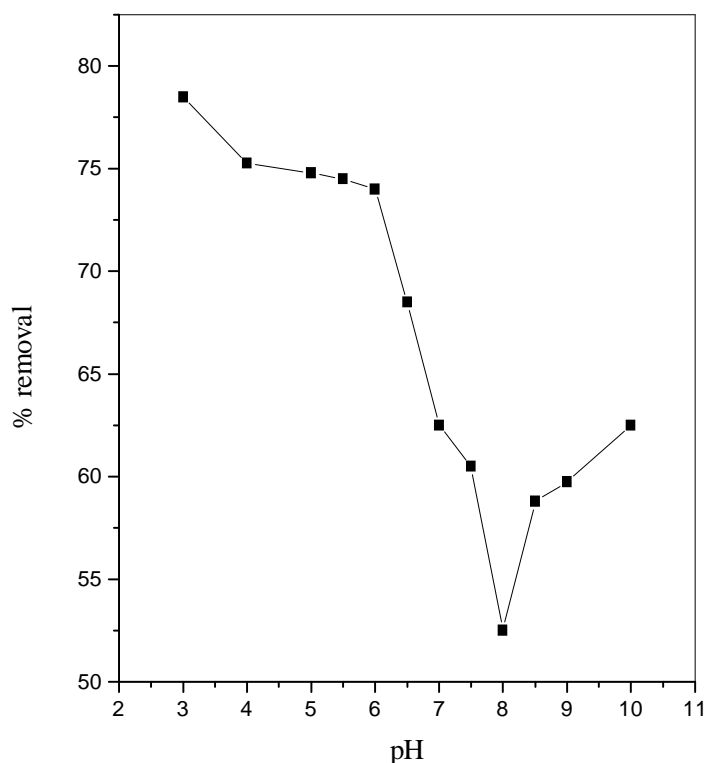


Fig.6-Effect of pH on the adsorption of methylene blue onto MCC

[MG]=60mg/L; Contact time=60 min; Adsorbent dose=50 mg/50 ml

4.9 Effect of the ionic strength on the adsorption of Methylene blue on MCC

The effect of sodium chloride on the adsorption of Methylene blue on MCC is shown in Fig. 7. In a low solution concentration NaCl had little influence on the adsorption capacity. At higher ionic strength the adsorption Methylene blue will be increased due to the partial neutralization of the positive charge on the carbon surface and a consequent compression of the electrical double layer by the Cl⁻ anion. The chloride ion can also enhances adsorption of Methylene blue ion onto MCC by pairing of their charges and hence reducing the repulsion between the Methylene blue molecules adsorbed on the surface. This initiates carbon to adsorb more of positive Methylene blue ions [1,17].

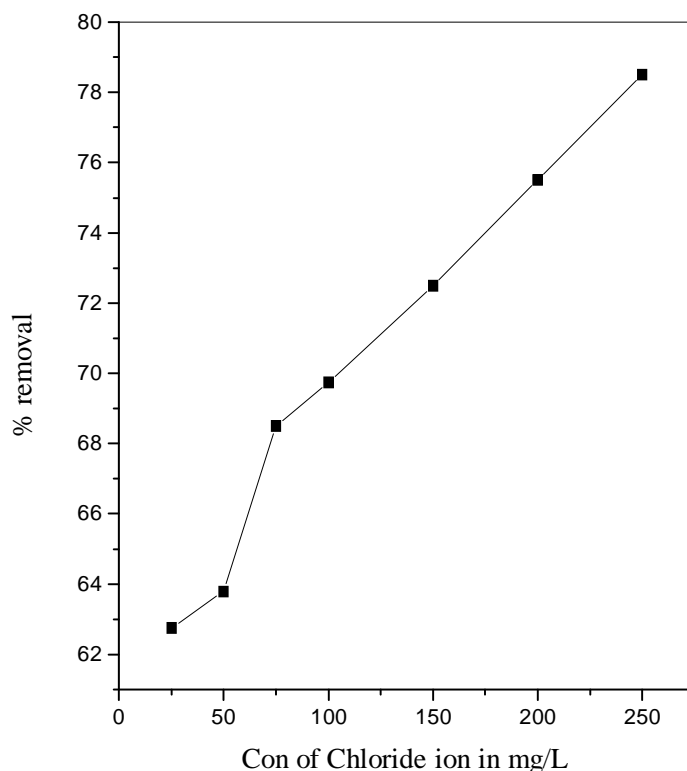


Fig.7-Effect of other ion on the adsorption of methylene blue by MCC
[MG]=60 mg/L; Contact time=60 min; Adsorbent dose=50 mg/50 ml

4.10 Desorption studies

Desorption studies help to elucidate the nature of adsorption and recycling of the spent adsorbent and the dye. If the adsorbed dyes can be desorbed using neutral pH water, then the attachment of the dye of the adsorbent is by weak bonds. If sulphuric acid or alkaline water desorb the dye then the adsorption is by ion exchange. If organic acids, like acetic acid can desorb the dye, then the dye has held by the adsorbent through chemisorption. The effect of various reagents used for desorption studies indicate that hydrochloric acid is a better reagent for desorption, because we could get more than 68% removal of adsorbed dye. The reversibility of adsorbed dye in mineral acid or base is in agreement with the pH dependent results obtained. The desorption of Methylene blue dye by mineral acids and alkaline medium indicates that the dyes were adsorbed onto the activated carbon through by physisorption mechanisms [12,18].

CONCLUSION

The experimental data correlated reasonably well by the Langmuir and Freundlich adsorption isotherms and the isotherm parameters were calculated. The low as well high pH value pay the way to the optimum amount of adsorption of the dye. The amount of Methylene blue adsorbed increased with increasing ionic strength and increased with increase in temperature. The dimensionless separation factor (R_L) showed that the activated carbon could be used for the

removal of Methylene blue from aqueous solution. The values of ΔH° , ΔS° and ΔG° results shows that the carbon employed has a considerable potential as an adsorbent for the removal of Methylene blue.

Acknowledgement

The authors acknowledge sincere thanks to Mrs. Mala Arivoli, The Principal, M.R.Government Arts College, Mannargudi and The Director of Collegiate Education, Chennai for carrying out this research work successfully.

REFERENCES

- [1] Arivoli S, Kinetic and thermodynamic studies on the adsorption of some metal ions and dyes onto low cost activated carbons, Ph D., Thesis, Gandhigram Rural University, Gandhigram, **2007**.
- [2] Sekaran G, Shanmugasundaram K A, Mariappan M and Raghavan K V, *Indian J Chem Technol*, **1995**, 2, 311.
- [3] Selvarani K, Studies on Low cost Adsorbents for the removal of organic and Inorganics from Water, Ph D., Thesis, Regional Engineering College, Thiruchirapalli, **2000**.
- [4] Jia Y F and Thomas K K, *Langmuir*, **2002**, 18, 470-478.
- [5] Namasivayam C, Muniasamy N, Gayathri K, Rani M and Renganathan K, *Biores Technol*, **1996**, 57, 37.
- [6] Namasivayam C and Yamuna R T, *Environ Pollut*, **1995**, 89, 1.
- [7] Langmuir I, *J Amer Chem Soc*, **1918**, 40, 1361.
- [8] Freundlich H, *Phys Chemie*, **1906**, 57, 384.
- [9] Krishna D G and Bhattacharyya G, *Appl Clay Sci*, **2002**, 20, 295.
- [10] Arivoli S, Viji Jain M and Rajachandrasekar T, *Mat Sci Res India*, **2006**, 3, 241-250.
- [11] Arivoli S and Hema M, *Intern J Phys Sci.*, **2007**, 2, 10-17.
- [12] Arivoli S, Venkatraman B R, Rajachandrasekar T and Hema M, *Res J Chem Environ*. **2007**, 17, 70-78.
- [13] Arivoli S, Kalpana K, Sudha R and Rajachandrasekar T, *E J Chem*, **2007**, 4, 238-254.
- [14] Renmin Gong, Yingzhi Sun, Jian Chen, Huijun Liu, Chao yang, *Dyes and Pigments*, **2005**, 67, 179.
- [15] Vadivelan V, Vasanthkumar K, *J Colloid Interf Sci*, **2005**, 286, 91.
- [16] Weber W J, Principle and Application of Water Chemistry, edited by Faust S D and Hunter J V Wiley, New York, **1967**.
- [17] Yupeng Guo, Jingzhu Zhao, Hui Zhang, Shaofeng Yang, Zichen Wang and Hongding Xu, *Dyes and Pigments*, **2005**, 66, 123-128.
- [18] Sreedhar M K and Anirudhan T S, *Indian J Environ Protect*, **1999**, 19, 8.