



# Biosynthesis of silver nanocomposite with *Tarragon* leaf extract and assessment of antibacterial activity

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## Abstract

The aqueous extract of *Tarragon*, as a reducing agent, was used to synthesize silver–montmorillonite (MMT) nanocomposite (Ag–MMT-NPs) in the batch method. The leaf extract and metal solution concentrations were optimized to improve Ag–MMT-NPs synthesis in 48 h. For characterizing the nanocomposite, powder X-ray diffraction (XRD), Fourier transform infrared spectroscopy (FTIR), Transmission electron microscopy (TEM), and UV–Vis spectroscopy were performed. The peak was observed at 437 nm on the UV–Vis spectrum, showing the surface plasmon resonance of Ag–MMT-NPs. Using XRD analysis, the crystalline nature and purity of Ag–MMT-NPs were confirmed. FTIR was used to evaluate specific functional groups, causing a reduction in silver nitrate during Ag–MMT-NPs formation. According to TEM, the average particle size was 25.12 nm in AgNPs. The nanocomposite showed antibacterial properties against Gram-negative and Gram-positive bacteria (*Escherichia coli* and *Staphylococcus aureus*).

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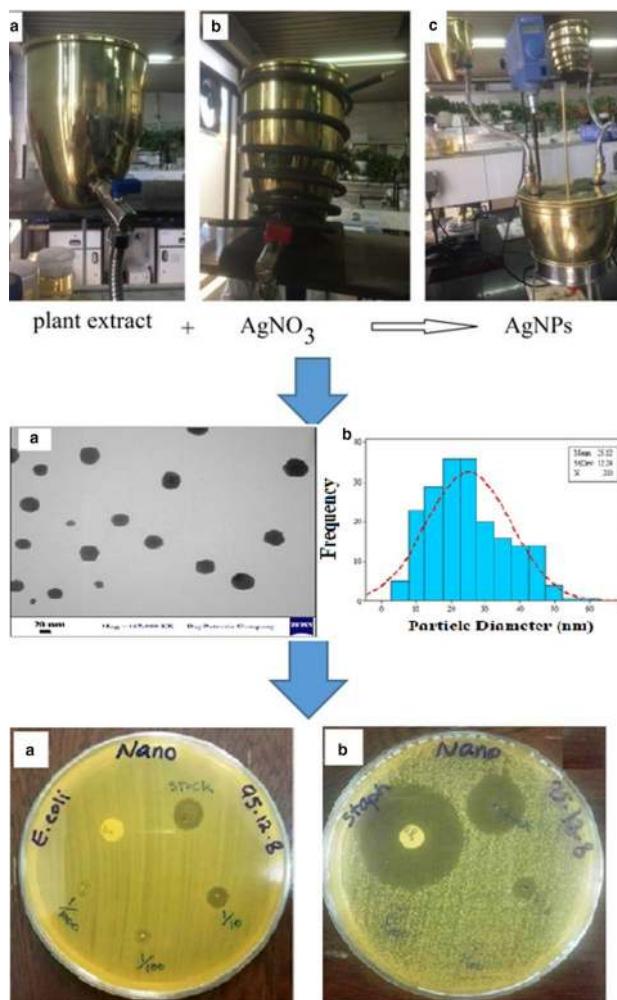
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## Graphical Abstract



**Keywords** Biosynthesis · Silver nanoparticles · Montmorillonite · *Tarragon* · SEM

### Abbreviations

XRD	X-ray diffraction
FTIR	Fourier transform infrared spectroscopy
TEM	Transmission electron microscopy
SPR	Surface plasmon resonance
<i>E. coli</i>	<i>Escherichia coli</i>
<i>S. aureus</i>	<i>Staphylococcus aureus</i>
AgNPs	Silver nanoparticles
UV–Vis	Ultraviolet–visible
FWHM	Full width half maxima
MBC	Minimum bactericidal concentration
MIC	Minimum inhibitory concentration

### Introduction

Nanotechnology research and development has been an area of rapid growth worldwide [1]. Silver nanoparticles (AgNPs) are one of the most extensively used varieties of NPs [2], with a large number of applications. AgNPs are used as coating in solar energy absorption. Moreover, they are used as optical receptors, biological labels, and intercalating materials for electrical batteries [3]. So far, metals, such as Ag, Au, Pd, and CdS, have been used for synthesizing metal NPs. Thus, the synthesis of Ag–NPs onto MMT supports with swelling and ion exchange properties is a good way to control the particle size [4, 5]. AgNPs exhibit important chemical, physical, and biological characteristics among metal NPs [6] and have potential applications in antimicrobial, anticancer, cosmetic, paint, food packing, and textile



industries [7]. Many techniques have been investigated for synthesizing AgNPs, including physical, chemical, and, most recently, biological methods [8]. The chemical synthesis of NPs is rapid, but requires the use of toxic, hazardous chemicals such as sodium borohydride, hydrazine, hydroxylamine, and ethanol [9]. Additionally, with chemical synthesis, it is difficult to control the stability, growth, and aggregation of particles, and capping agents are required for stabilization of NP size [2]. Recently, great attention has been directed towards plant extracts for NP [10], and Ag/nanocomposites synthesis [11]. Use of biosynthetic green metal NPs is gaining increasing approval owing to its simplicity, non-toxicity, and amenability to large-scale production [12].

AgNP synthesis has been evaluated in different extracts, including *Medicago sativa* [13], *Ulva flexuosa* [14], *Achillea biebersteinii* [15], *Moringa oleifera* [16], *Calendula officinalis* [17], *Peganum harmala* [18], Green Tea [19], *Pistacia atlantica* [20], *olive* [21], *Aloe vera* [22], and *Coriandrum Sativum* [23]. We developed a simple, rapid, and green method to synthesize AgNPs, using MMT and *Tarragon* leaf extract as both a reducing and capping agent under normal atmospheric conditions in the batch method in this study. This method is easily adopted for large-scale synthesis of NPs, without requiring any extra compounds or physical processes [24].

Simple collection and widespread availability of *Tarragon*, coupled with its remarkable biological activity, have led to its use as a medicine in many countries [25]. *Tarragon* has been recently shown to have wide-ranging antimicrobial activities [26]. Based on some researchers, *Tarragon* has been shown to have antimicrobial activities against two bacteria *Staphylococcus aureus* and *Escherichia coli* [27]. The size of synthesized AgNPs with *Tarragon* extract is smaller as compared to the many earlier green synthesis reports [1, 2]. Reaction of *Tarragon* extract with aqueous metal solution ( $\text{AgNO}_3$ ) yields only spherical nanoparticles. The leaf extract and metal solution ( $\text{AgNO}_3$ ) concentrations were optimized to improve AgNP synthesis. AgNPs were characterized by methods, including powder X-ray diffraction, UV–Vis spectroscopy, Transmission electron microscopy, and Fourier transform infrared spectroscopy. The disk diffusion method was applied to evaluate the antibacterial potential of AgNPs.

## Materials and methods

### Preparing of silver nanoparticles with *Tarragon* extract

Green *A. dracunculus* leaves were collected from north east of Iran in May 2017. After washing the *Tarragon* leaves, they were dried under direct light for 3 days at a temperature

of 38 °C.  $\text{AgNO}_3$  (99.80%) was provided by Merck Co. Montmorillonite powder (MMT), used as the solid support, was purchased from Fluka Chemical Co. 0.01 g of MMT powder was dispersed with vigorous stirring in certain amount of double-distilled water for 1 h. MMT suspension was added to 100 mL of 0.01 (mol/L)  $\text{AgNO}_3$  solutions for the synthesis of Ag/MMT nano composite. The mixture was then added to 20 mL of *A. dracunculus* water extract at room temperature while sonicating and then vigorously stirred for 48 h. The color changed from yellow to dark brown at room temperature, and AgNPs were gradually obtained during the reaction.

### Purification of silver nanoparticles

To remove silver colloid residues, the solution was first washed with double-distilled water via centrifugation for 15 min at 4000 rpm, and then, washed three times with deionized water. After incubation at 65 °C for 2 h, the dried powder of Ag–MMT-NPs was collected for further characterization [28]. In other previous similar investigations, the maximum volume used for synthesis of AgNPs has been 50 mL [18], but in our work the volume was increased to 2000 cc which resulted in synthesizing 2 g of silver nanoparticles. due to this information, we can claim that this method could be easily adopted for large-scale synthesis of NPs.

### UV–Vis spectroscopy

Considering the UV–Vis spectrum of the reaction medium, AgNP formation was monitored at a wavelength range of 300–800 nm. A UV–Vis spectrophotometer (BioTek) was used to record the spectra. The samples were appropriately diluted with water before each measurement.

### Fourier transform infrared (FTIR) spectroscopy

For determining the functional groups in the synthesis of AgNPs, spectroscopy analyses were performed to describe the functional groups [29], bound distinctively to the surface of AgNPs. Samples were prepared by vacuum-drying the *Tarragon* extract before and after synthesis of AgNPs. After mixing the dried samples with potassium bromide, they were pressed into a sheer slice. The samples were examined using an FTIR spectrometer (PerkinElmer, USA).

### X-ray diffraction (XRD)

The crystal structure of AgNPs was determined and confirmed via XRD analysis. After placing the air-dried NPs on an XRD grid, they were assessed in terms of AgNP formation, using an X-ray diffractometer (Cu  $K\alpha$  radiation,  $\theta$ – $2\theta$  configuration; PANalytical) with an X'Pert Pro generator



(30 mA; 40 kV). For measuring the crystallite domain size, the XRD peak width was determined under the assumption that non-uniform strains are absent.

### Transmission electron microscopy (TEM)

Morphological analysis of Ag–MMT-NPs was carried out on a Zeiss EM 10C/CR TEM microscope (Zeiss, Germany) at 100 kV. The suspension was drop-casted on a carbon-coated copper grid and left to dry overnight at 25 °C.

### Antibacterial activity

For determining the antibacterial activity of Ag–MMT-NPs, the agar disk diffusion method was applied, as previously reported [30]. Briefly, AgNPs at 5, 0.5, and 0.05 µg/mL were used to prevent *E. coli* (ATCC 25922) and *S. aureus* (ATCC 25923) growth. Bacterial cultures at a concentration of  $1.5 \times 10^8$  CFU/mL were inoculated with different concentrations of AgNPs [31]. After incubating the culture plates at 37 °C for 24 h, the bacterial growth inhibition zone was measured in millimeters [32]. For each bacterial strain, three independent experiments were performed.

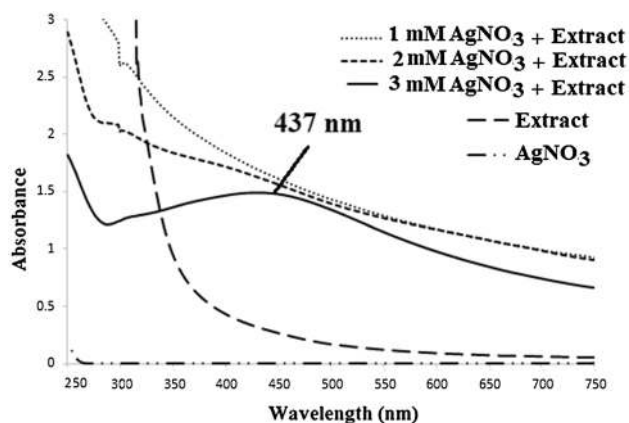
## Results and discussion

To investigate whether an aqueous extract of *Tarragon* can be used in the green synthesis of AgNPs, we combined the extract with 1, 2 and 3 mM of AgNO<sub>3</sub> for 24 h. The slow change of color in the solution (from yellow to dark brown) was indicative of AgNP formation. The dark brown color might be related to the surface plasmon resonance (SPR) effect and AgNO<sub>3</sub> reduction [33].

### UV–Vis spectroscopy

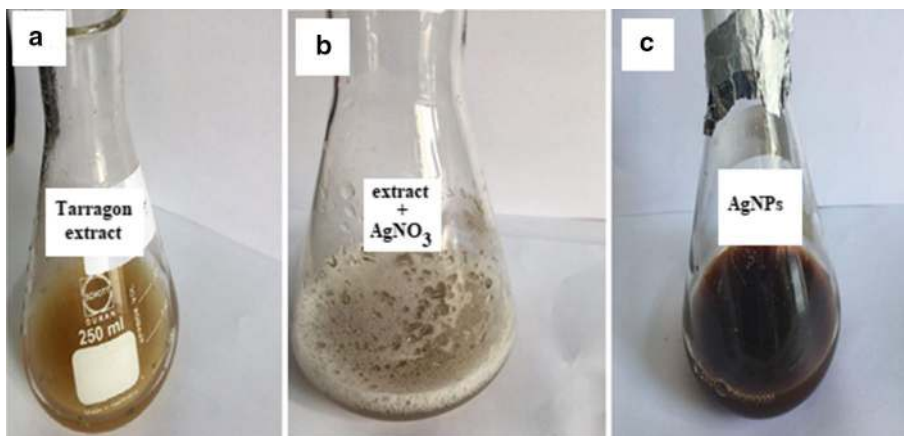
UV–Vis spectroscopy is an efficient technique for NP analysis, particularly for determining the stability of metal NPs in aqueous solutions. Different concentration of AgNO<sub>3</sub> (1, 2 and 3 mM) was reacted with the aqueous extract, and the spectra were recorded at a range of 350–850 nm [34]; then, the wavelength of maximum absorption was identified. Biosynthesis of AgNPs by the *Tarragon* extract was confirmed based on changes in color after the addition of AgNO<sub>3</sub> (Fig. 1).

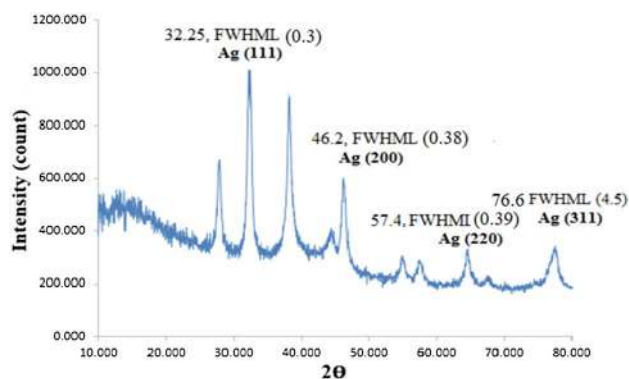
Different concentrations of AgNO<sub>3</sub> were used for optimization. The analysis showed a sharp plasmon resonance of 3 mM AgNO<sub>3</sub> at nearly 437 nm (Fig. 2), consistent with previous reports of peaks in the absorption spectrum of 400–500 nm due to SPR in AgNPs [35, 36]. Without any physical or chemical capping agents, the NPs were dispersed. The *Tarragon* synthesized AgNPs were quite stable in the solution.



**Fig. 2** UV–Vis spectra of synthesized silver nanoparticles using *Tarragon* extract and AgNO<sub>3</sub>

**Fig. 1** **a** Extract of *Tarragon*, **b** aqueous solutions of silver nitrate (AgNO<sub>3</sub>) and extract of *Tarragon* (*Artemisia dracunculifolia*), **c** Ag–MMT-NPs after 24 h of incubation





**Fig. 3** XRD pattern of Ag–MMT-NPs, indicating crystal structure

### XRD analysis

Using the XRD pattern of synthesized Ag–MMT-NPs, the peaks were attributed to the crystalline planes of Ag–MMT-NPs nanocomposite (Fig. 3). The peaks were very sharp, indicating the crystalline nature of synthesized Ag–MMT-NPs [18].

The diffraction profile had intense peaks at  $2\theta$  of  $38.14^\circ$ ,  $46.2^\circ$ ,  $64.51^\circ$ , and  $76.6^\circ$ , corresponding to the (111), (200), (220), and (311) planes, respectively. The mean particles size of AgNPs was evaluated using the Debye–Scherer Eq. (1):

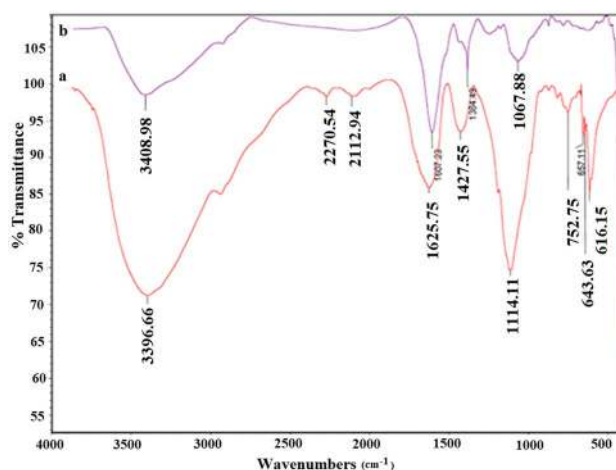
$$d = \frac{K\lambda}{\beta \cos\theta}, \quad (1)$$

where  $\lambda$  denotes the X-ray wavelength ( $1.540560 \text{ \AA}$ ),  $\beta$  represents the width of XRD peak,  $\theta$  shows the Bragg angle,  $d$  refers to the size of particle, and  $K$  describes the Scherer constant (0.9) [7]. The average size of AgNPs was almost 28 nm.

### FTIR analysis

In FTIR spectrum, the biomolecules resulting in Ag<sup>0</sup> reduction and stabilization can be shown (Fig. 4).

FTIR spectroscopy was carried out to determine the potential biomolecules responsible for the reduction and capping of the AgNPs synthesized. The FTIR spectra of plant extract shows absorption bands characteristics of functional groups such as alcohol, phenol, amine and carbonyl group. In FTIR absorption bands at 3396.66, 1625.75, and 1427.55/cm (Fig. 4a). The vibrational bands corresponded to bonds, such as alcohols (–O–H), alkenes



**Fig. 4** a FTIR spectra of *Tarragon* extract, b Ag–MMT-NPs synthesized using the *Tarragon* extract

(>C=C), amines (=N–H), flavonoids, and amines (–NH<sub>2</sub>), which are all in the range of 800–3442/cm. The absorption band at 3396.66/cm in the spectrum might be attributed to the stretching vibrations of secondary amines of N–H bond and bonded hydroxyl (–OH) groups of phenols and carboxylic acids [18]. Bonded hydroxyl (–OH) groups of carboxylic acids and phenols could produce absorption bands at 3396.66/cm (Fig. 4a). In the AgNP spectrum, the new band at 1607.23/cm might suggest a new C=O group (a ketone or an aldehyde) (Fig. 4b). FTIR spectroscopy was carried out to determine the potential biomolecules responsible for the reduction and capping of the AgNPs synthesized. The FTIR spectra of plant extract shows absorption bands characteristics of functional groups such as alcohol, phenol, amine, and carbonyl group.

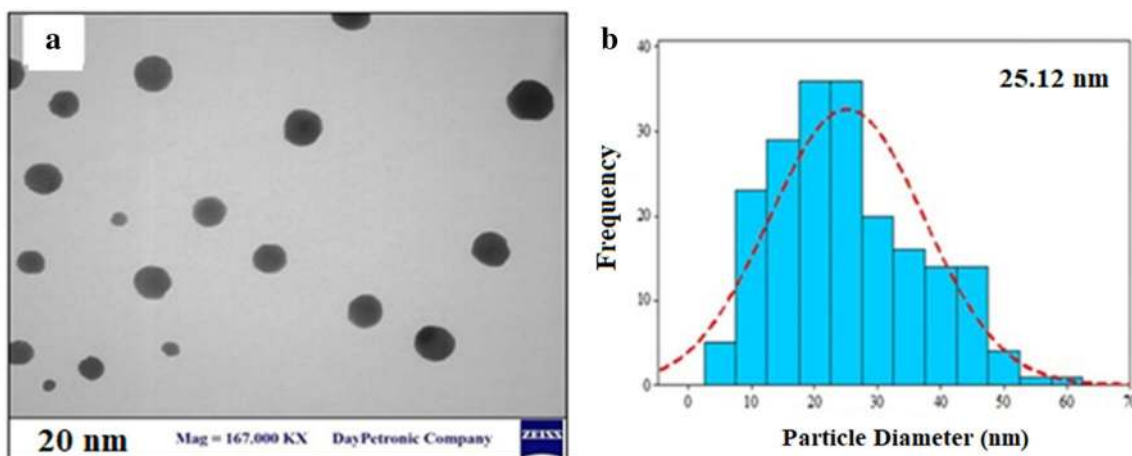
In this manner, they essentially act as a capping agent, providing stability and preventing agglomeration for AgNPs. This phenomenon has been observed in the green synthesis of AgNPs using other plant extracts [37]. These findings confirm that phenolic agents and proteins in the *Tarragon* (*Artemisia dracunculus*) extract are responsible for the reduction of AgNO<sub>3</sub> [22].

### TEM analysis

As shown in (Fig. 5a), the TEM image shows the average size was about 25.12 nm with interring particle distance (Fig. 5b). The AgNPs are quasi-spherical in shape and no aggregation of AgNPs was observed.

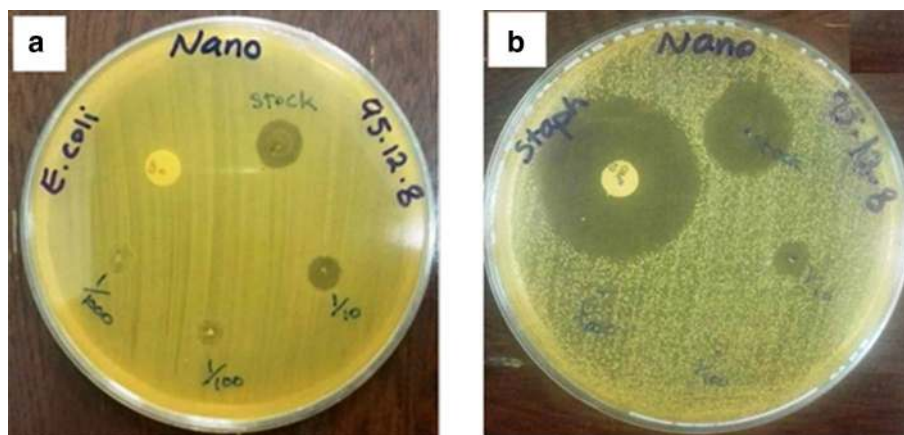






**Fig. 5** **a** TEM image of the synthesized Ag–MMT-NPs with *Tarragon* extract; **b** particle size distribution of biosynthesized Ag–MMT-NPs

**Fig. 6** Antibacterial activity of Ag–MMT-NPs against both **a** Gram-negative (*E. coli*) and **b** Gram-positive (*S. aureus*) in different solutions



### Antibacterial activity

For evaluating the antibacterial activity of AgNPs, both Gram-positive (*S. aureus*) and Gram-negative (*E. coli*) bacteria were used (Fig. 6). The inhibition zone of Ag–MMT-NPs is listed in Table 1.

### Conclusion

In the present research, a novel process was used for the synthesis of Ag–MMT-NPs with *Tarragon* extract as a reducing, nontoxic reagent and capping agent. Additionally, this method is easily adopted for the mass production

of nanoparticles without requiring any extra compounds or physical processes and have well-defined size and morphology. The green synthesis of Ag–MMT-NPs with the extract is a more cost-effective and eco-friendly strategy, compared to physical and chemical synthesis methods. Optimization of AgNO<sub>3</sub> and *Tarragon* extract concentration was achieved to obtain large amount of AgNPs in batch process. UV–Vis, FTIR, TEM, and XRD studies were performed to confirm the Ag–MMT-NPs formation. For AgNPs, UV–Vis peak was observed at 437 nm. The crystalline structure of synthesized Ag–MMT-NPs was revealed on XRD. The Ag–MMT-NPs had an anisotropic nature and a quasi-spherical shape (average size around, 25.12 nm).



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