

# Catalytic non-thermal plasma reactor for the decomposition of a mixture of volatile organic compounds

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**Abstract.** The decomposition of mixture of selected volatile organic compounds (VOCs) has been studied in a catalytic non-thermal plasma dielectric barrier discharge reactor. The VOCs mixture consisting *n*-hexane, cyclo-hexane and *p*-xylene was chosen for the present study. The decomposition characteristics of mixture of VOCs by the DBD reactor with inner electrode modified with metal oxides of Mn and Co was studied. The results indicated that the order of the removal efficiency of VOCs followed as *p*-xylene > cyclo-hexane > *n*-hexane. Among the catalytic study, MnOx/SMF (manganese oxide on sintered metal fibres electrode) shows better performance, probably due to the formation of active oxygen species by *in situ* decomposition of ozone on the catalyst surface. Water vapour further enhanced the performance due to the *in situ* formation of OH radicals.

**Keywords.** Non-thermal plasma; mixture of VOCs; ozone; CO<sub>2</sub>; oxidation.

## 1. Introduction

The emission of volatile organic compounds (VOCs) by various industrial and agricultural processes is a significant source of air pollution and may cause a problem to human health due to their toxicity (some of the VOCs are carcinogenic or responsible for respiratory diseases), hence has a negative impact on the environment.<sup>1–4</sup> Some of the well-established technologies for VOC abatement are thermal and catalytic incineration, adsorption, condensation, bio-filtration, membrane separation or ultra-violet oxidation.<sup>5–9</sup> These techniques may demand thermal energy in the range of 700–800°C (for thermal incineration) and 200–600°C (for catalytic oxidation). However, these techniques may not be effective and energetically expensive for the treatment of low VOC concentrations. In addition, the catalyst poisoning followed by deactivation represents a serious problem for these technologies.<sup>10</sup> In this context non-thermal plasma (NTP) is an emerging technology for the removal of dilute concentrations of air pollutants.<sup>11,12</sup> The major advantages of NTP technology include operation under ambient conditions, moderate capital cost, compact system, easy operations and short reaction time.<sup>13–17</sup>

Dielectric barrier discharge (DBD) technique is one of the commonly used methods for producing electrical discharge plasma. The main principle of operation in NTP-DBD reactor is to apply energetic electric field between two electrodes that were separated by a barrier to generate primary electrons with high kinetic energy.<sup>14–17</sup> These high energy electrons (1–10 eV) may collide and excite the background gas (O<sub>2</sub> and N<sub>2</sub>) that lead to the formation of highly reactive species.<sup>18</sup> The critical parameter in the use of this technique in pollution control devices is the energy cost to remove toxic molecules. The energy cost mainly depends on energy transfer from a power source to plasma reactor, configuration of electrodes and efficiency of chemical reactions.<sup>19</sup> However, its application is greatly restricted by low energy efficiency and CO<sub>2</sub> selectivity as well as suppression of undesirable by-products such as ozone.<sup>20–23</sup> An attempt to overcome these limitations is to combine NTP with catalysts.<sup>4,23–25</sup>

The selected catalysts (MnOx and CoOx) were known for ozone decomposition and several researchers reported the plasma catalytic effect of MnOx and CoOx for ozone decomposition.<sup>4,24,26–32</sup> However, majority of these studies were dealt with single component VOCs. As industrial emission contain VOCs mixture, the present work focuses on the synergetic effect of DBD plasma with catalyst during the removal of mixture

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of VOCs. The inner electrode was made of sintered metal fibres (SMF) which acted as electrode as well as a catalyst after modification with transition metal (Mn or Co) oxides. A typical VOCs mixture consisting aliphatic compound (*n*-hexane), alicyclic compound (cyclo-hexane) and aromatic compound (*p*-xylene) was chosen to understand the influence of various parameters like input energy and water vapour on mitigation efficiency.

## 2. Experimental

The detailed schematic representation of the DBD reactor used in the present study was given elsewhere<sup>4</sup> and the experimental set-up is illustrated in figure 1. Briefly, the dielectric barrier discharge was generated in a cylindrical quartz tube with an inner diameter of 19.5 mm. SMF (sintered metal fibres) made of Fe–Cr alloy (in the form of cylindrical tube with a diameter of 12.5 mm) was used as inner electrode after modification with transition metal oxides. Deposition of 5 wt% MnO<sub>x</sub> and CoO<sub>x</sub> on SMF was achieved by impregnation with metal nitrate solutions followed by drying and calcination at 773 K for 5 h. One end of SMF filter was

connected to AC high voltage through a stainless steel rod whereas the other end was connected to the gas stream outlet. Silver paste painted on the outer surface of the tube acted as an outer electrode. The discharge length was 10 cm and the discharge gap was fixed at 3.5 mm. The AC high voltage in the range 14–22 kV (peak–peak) with the frequency of 50 Hz was applied between the two electrodes. Conversion of mixture of VOC at each voltage was measured after 30 min.

The mixture of VOCs was introduced through a motor driven syringe pump and mixed with ambient air (300 ml/min at STP) in a mixing chamber. Air flow was regulated by pre-calibrated mass flow controllers (Aalborg, USA). The input concentration of total VOCs was fixed at 250 ppm (hexane = 100 ppm, cyclo-hexane = 75 ppm and *p*-xylene = 75 ppm) and was fed into the plasma reactor with a Teflon tube. Reactor outlet was connected to gas chromatography (Varian-450) equipped with a flame ionization detector (FID) detector and a capillary column for monitoring VOC concentration. The formation of CO<sub>2</sub>, CO and NO<sub>x</sub> was simultaneously monitored with an infrared gas analyzer (Siemens Ultramat 22), whereas, ozone formed in the plasma reactor was measured with an UV absorption ozone monitor (API-450 NEMA). The global selectivity of CO<sub>2</sub>, CO and CO<sub>x</sub> for represented as

$$\begin{aligned} & \text{Global selectivity of CO}_2 \% (S_{\text{CO}_2}) \\ &= \frac{\text{Total concentration of [CO}_2\text{]}}{6 (\text{Conv. of } n\text{-hexane}) + 6 (\text{Conv. of cyclo-hexane}) + 8 (p - \text{xylene})} \times 100. \end{aligned} \quad (1)$$

$$\begin{aligned} & \text{Global Selectivity of CO}\% (S_{\text{CO}}) \\ &= \frac{\text{Total concentration of [CO]}}{6 (\text{Conv. of } n\text{-hexane}) + 6 (\text{Conv. of cyclo-hexane}) + 8 (p - \text{xylene})} \times 100. \end{aligned} \quad (2)$$

$$S_{\text{CO}_x} = S_{\text{CO}} + S_{\text{CO}_2}. \quad (3) \quad (\text{SIE}) \text{ of the discharge was calculated by following relation}$$

The electrical power applied to the discharge in the DBD reactor was measured by using the V-Q Lissajous diagram,<sup>4</sup> where the charge Q (i.e., time integrated current) was recorded by measuring the voltage across the capacitor ( $C = 1 \mu\text{F}$ ) connected in a series to the ground. Applied voltage was measured with a 1000:1 high voltage probe (Agilent 34136A) and the V–Q wave forms were monitored by a digital oscilloscope (Tektronix TDS 2014 B). The area of Lissajous figure characterizes the energy dissipated during the one period of voltage. Specific input energy

$$\text{SIE (J/l)} = \frac{\text{Power (W)}}{\text{gas flow rate (l/s)}}. \quad (4)$$

Figure 2 shows that the energy (SIE) dissipated in the gas per pulse increases linearly with the applied voltage. Then one may suppose that a high electric field at the anode modifies electron mean energy, increases production of active species in the streamers, and then improves VOCs destruction and removal efficiency.<sup>18</sup>

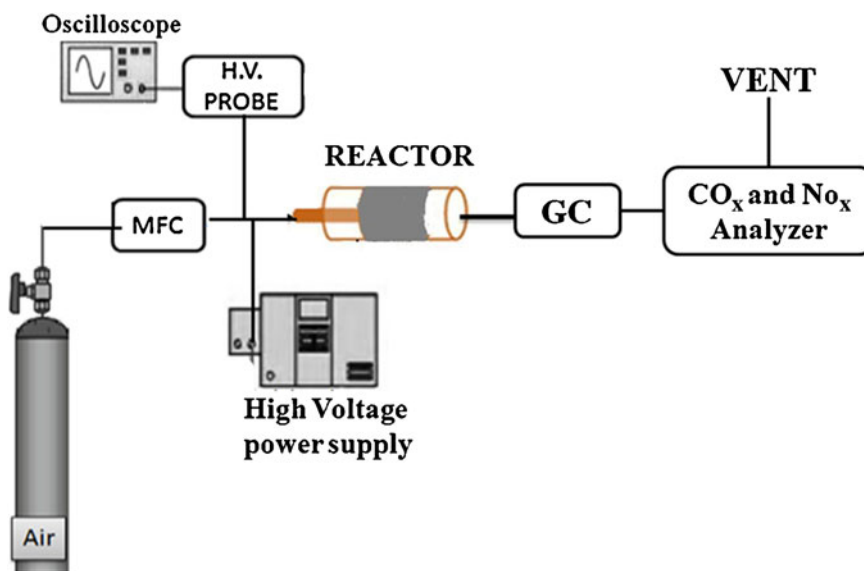


Figure 1. Schematic representation of experimental set-up.

### 3. Results and discussion

#### 3.1 Destruction of the mixture of VOCs in DBD reactor

The pollutant concentration at the inlet of the DBD reactor is fixed at 250 ppm. The carrier gas is either dry air or humid air (2.3 vol% humidity was introduced by bubbling air through the water at 298 K), at a flow rate of 300 ml/min. The effect of the specific input energy on the VOCs decomposition with SMF, CoOx/SMF, MnOx/SMF and MnOx/SMF with humid air was presented in figure 3. This figure illustrates that the decomposition of VOCs increased

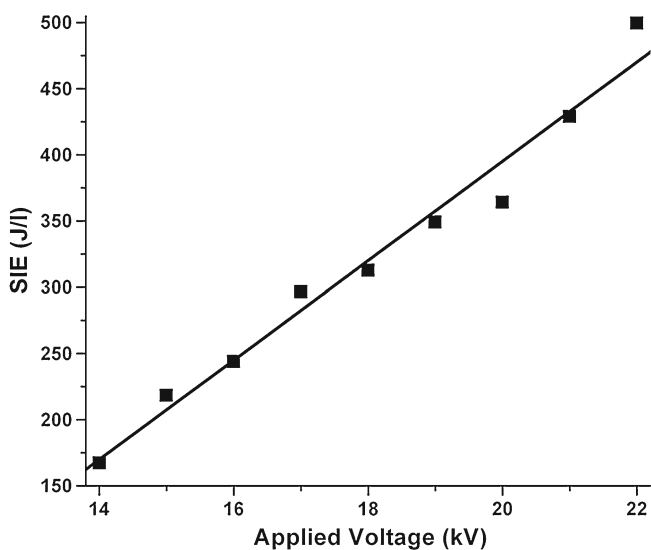


Figure 2. Variation of SIE as a function of voltage.

with SIE and the concentration of *p*-xylene and cyclo-hexane were decreased from 75 to 0 ppm at SIE of 350 and 425 J/l, respectively, whereas *n*-hexane was decreased from 100 ppm to 0 ppm at 500 J/l when only pure SMF used as an inner electrode. While modified SMF (CoOx/SMF, MnOx/SMF) used as an inner electrode it showed better activity compared to pure SMF. *P*-xylene, cyclo-hexane and *n*-hexane were completely degraded at 300, 350 and 365 J/l, respectively, when CoOx/SMF as an inner electrode and 245, 310 and 350 J/l, respectively, when MnOx/SMF used as an inner electrode. Here, SMF modified MnOx catalytic electrode showed better activity compare to other catalytic electrodes may be due to manganese oxide catalyst have better activity to transform the ozone to active oxygen species<sup>24,33</sup> which were more reactive to

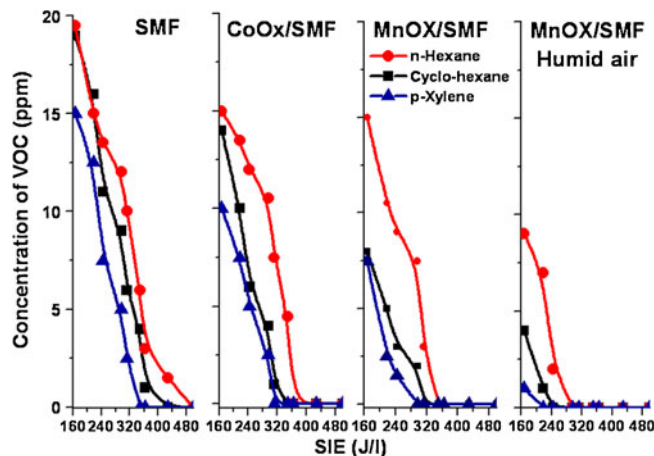


Figure 3. VOCs decomposition on different catalytic electrodes as a function of SIE.

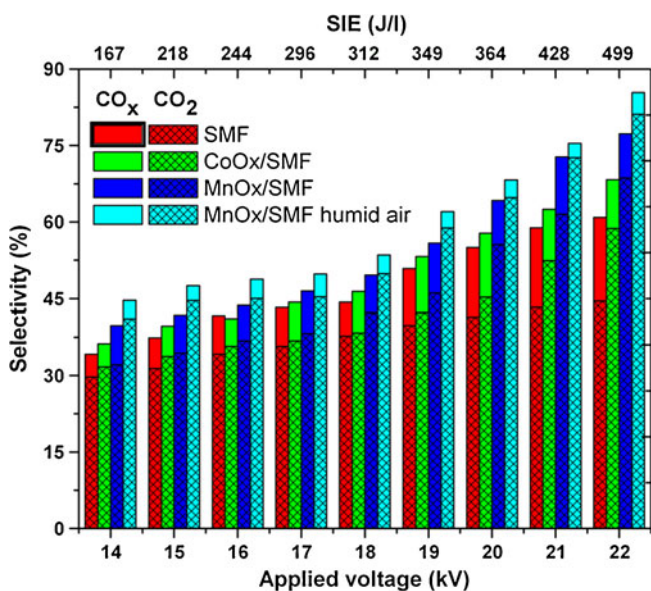
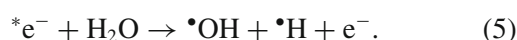
oxidize the organic components to  $\text{CO}_2$ . *p*-xylene is the easiest compound to decompose and *n*-hexane is much more resistant to NTP treatment than the two other compounds in target VOCs present in the mixture. It showed that the order of removal efficiency of mixed VOCs followed *p*-xylene > cyclo-hexane > *n*-hexane. This phenomenon could be explained by the competition and interaction of each component VOC molecule by high energy electrons (6~10 eV) and active radicals (O, OH, N) that may play an important role in the decomposition of gas pollutants.<sup>34</sup> The chemical bond strength and molecule stability may also determine the removal rates of VOCs in the plasma process.<sup>35</sup> Previous studies demonstrated that the removal efficiency of organic compound largely depends on its chemical structure.<sup>36-39</sup> Among these VOCs, the bond energy of carbon-carbon bond is the strongest in *n*-hexane than cyclo-hexane, where as methyl group in *p*-xylene is comparatively unstable that can be broken easily.

The selectivity to gaseous products  $\text{CO}_x$  ( $\text{CO} + \text{CO}_2$ ) formed in the decomposition of mixture of VOCs is given in figure 4. As seen from figure 4, increasing SIE leads to higher selectivity and it is clearly observed that at 22 kV, for SMF as an inner electrode 45%  $\text{CO}_2$  selectivity and 59 and 69% for CoOx/SMF, MnOx/SMF electrodes, respectively. With CoOx and MnOx/SMF electrodes at low input energy, a significant increase in the selectivity to total oxidation products was achieved when compared to unmodified electrode. Within the plasma catalytic reactor, synergy was due to the utilization of the short-lived oxidizing species like oxygen

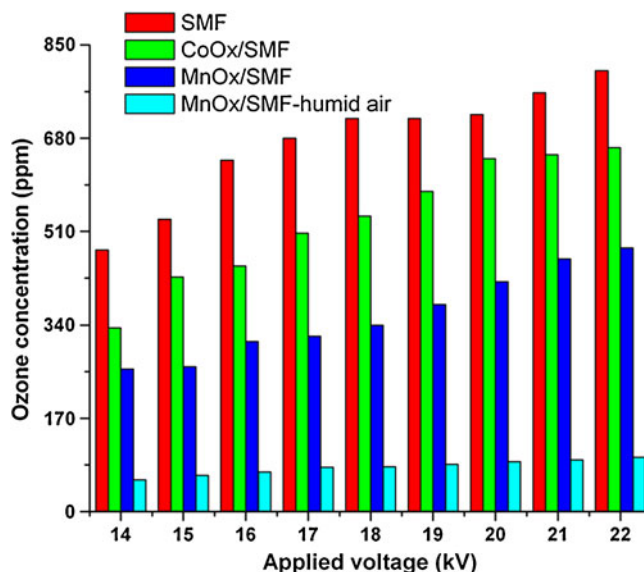
radical anion ( $\text{O}_n^{-1}$ ) and electronically activated oxygen molecule ( $\text{O}_2^*$ ). The DBD reactor containing catalytic electrode, it is possible to overcome the limitations in-plasma reactor and avoiding the carbon deposition.

### 3.2 Influence of water vapour

Influence of water vapour on decomposition of mixture of VOCs was studied by SMF modified with MnOx electrode which showed better activity compared to remaining catalytic electrodes. The mixture of VOCs removal by DBD has been enhanced as a function of the presence of the water vapour (2–3 Vol %) in air. For water vapour production, air was passing through the bubbler. It can be observed that mixture of VOCs removal by DBD in humid air is higher than that in dry air (figure 3), at the same value of the specific energy. From figure 3, it is observed that at 220, 245 and 295 J/l SIE was sufficient for the complete decomposition of *P*-xylene, cyclo-hexane and *n*-hexane VOCs, respectively and figure 4 illustrates that 82%  $\text{CO}_2$  selectivity was observed which was better than dry condition. In the presence of water vapour, the hydroxyl radical may be produced by dissociative electron collisions with  $\text{H}_2\text{O}$  (Eq. 5) that may lead to a higher oxidation rate of VOCs for a given value of the specific energy.<sup>40,41</sup> On the other hand, it also has an adverse effect on removal of target compounds due to electronegative characteristics.<sup>41,42</sup>



**Figure 4.** Selectivity of  $\text{CO}_x$  and  $\text{CO}_2$  with different catalytic electrode during the abatement of 250 ppm mixture of VOC as a function of SIE and applied voltage.

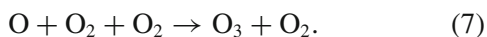
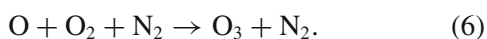


**Figure 5.** Ozone concentration variation with different catalytic electrodes as a function of applied voltage.

It is well-known that O<sub>3</sub> can be abundantly generated during the discharge in DBD reactor.<sup>4,24</sup> However, its formation was suppressed due to the extinguishing of energetic electrons by water vapour.<sup>43</sup> It can be found that the O<sub>3</sub> concentration dropped in the presence of water vapour (figure 5). O<sub>3</sub> is a strong oxidant and can be decomposed into O, which is highly reactive and can efficiently oxidize the VOCs components. Thus, water vapour can impose negative effect on VOCs removal by inhibiting O<sub>3</sub> formation and subsequent catalytic ozonation. However, the rate constants O and ·OH are much more active than O<sub>3</sub> for decomposition of VOCs<sup>43</sup> hence the decomposition rate of VOCs increases. Due to OH radicals and O species, the selectivity of CO<sub>2</sub> also increases when compared to dry air.

### 3.3 Formation of undesirable products

In general, during the removal of VOCs, formation of undesirable products is one of the major problems.<sup>24</sup> Some of these molecules are a source of olfactory pollution or environmental hazards. Therefore the by-products identification and quantification is a key step to appreciate the viability of the plasma depollution process. The dissociation of air may lead to N and O production, which is the main precursor of NO<sub>x</sub>. The formation of NO<sub>x</sub> occurs when the temperature of the gas is significantly increased by the plasma and reaches about 373 K.<sup>23</sup> However, no traces of nitrogen oxides such as NO and NO<sub>2</sub> detected among the numerous by-products. One of the applications of NTP is ozone generation. Large amounts of ozone are formed in the DBD plasma while VOCs are removed. Ozone formation occurs when following reactions take place between atomic oxygen and molecular oxygen.



Ozone is a long-lived species and for both in-plasma and post-plasma catalytic reactors, it plays an important role during the oxidation of VOCs. However, the atomic oxygen formed due to *in situ* decomposition of ozone may effectively interact with partially oxygenated species only when catalyst placed in discharge zone.<sup>44</sup> It is also interesting to note that ~ 800 ppm of O<sub>3</sub> was observed when VOCs mixture decomposition on SMF electrode at 22 kV that decreased to ~ 660 ppm with CoOx/SMF and ~ 480 ppm with MnOx/SMF and finally < 100 ppm MnOx/SMF in humid air condition, as shown in figure 5. As observed in figure 5, the best performance of the plasma reactor was achieved with catalytic MnOx/SMF electrode under humid air

condition, which shifted the product distribution towards total oxidation, whereas conversion remains nearly the same on all the catalysts. These results suggest that the better performance of MnOx/SMF electrode under dry condition may be due to the formation of atomic oxygen by *in situ* decomposition of ozone,<sup>33</sup> whereas under humid conditions, it may be primarily due to oxidation by OH radicals.

## 4. Conclusions

The experimental results during the oxidation of low concentration of mixture of VOCs indicated that the removal efficiency of VOCs significantly enhanced in the presence of metal oxide catalysts, especially MnOx/SMF. The removal efficiency followed the order, *p*-xylene > cyclo-hexane > *n*-hexane. Water vapour may facilitate the formation of strong oxidant hydroxyl radicals and active O species by *in situ* decomposition of ozone, hence high conversion and selectivity to total oxidation of VOCs.

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## References

1. Monks P S, Granier C, Fuzzi S, Stohl A, Williams M L, Akimoto H, Amann M, Baklanov A, Baltensperger U, Bey I, Blake N, Blake R S, Carslaw K, Cooper O R, Dentener F, Fowler D, Fragkou E, Frost G J, Generoso S, Ginoux P, Grewet V, Guenther A, Hansson H C, Hennew S, Hjorth J, Hofzumahaus A, Huntrieser H, Isaksen I S A, Jenkin M E, Kaiser J, Kanakidou M, Klimont Z, Kulmala M, Laj P, Lawrence M G, Lee J D, Liou S C, Maione M, McFiggans G, Metzger A, Mieville A, Moussiopoulos N, Orlando J J, ODowd C D, Palmer P I, Parrish D D, Petzold A, Platt U, Poschl U, Prevot A S H, Reeves C E, Reimann S, Rudich Y, Sellegri K, Steinbrecher D, Simpson H, ten Brink J, Theloke G R, van der Werf R, Vautard V, Vestreng R, Vlachokostas C and von Glasow R 2009 *Atmos. Environ.* **43** 5268
2. Atkinson R 2000 *Atmos. Environ.* **34** 2063
3. Odum J R, Jungkamp T P W, Griffin R J, Forstner H J L, Flagan R C and Seinfeld J H 1997 *Environ. Sci. Technol.* **31** 1890
4. Subrahmanyam C, Magureanu M, Renken A and Minsker-Kiwi L 2006 *Appl. Catal. B: Environ.* **65** 150
5. Choma J, Jaroniec M and Spikelet M 2007 *Environ. Protec.* **29** 3

6. Li W B and Gong H 2010 *Acta Phys-Chim. Sin.* **26** 885
7. Kumar A, Dewulf J and Van Langenhove H 2008 *Chem. Eng. J.* **136** 82
8. Mudliar S, Giri B, Padoley K, Satpute D, Dixit R, Bhatt P, Pandey R, Juwarkar A and Vaidya A 2010 *J. Environ. Manag.* **91** 1039
9. Demeestere K, Dewulf J and Van Langenhove H 2007 *Crit. Rev. Environ. Sci. Technol.* **37** 489
10. Than Quoc An H, Pham Huu T, Le Van T, Cormier J M and Khacef A 2011 *Catal. Today* **176** 474
11. Chang M B and Tseng T D 1996 *J. Environ. Eng.* **122** 41
12. Nunez C M, Ramsey G H, Ponder W H, Abbott J H, Hamel L E and Kariher P H 1993 *Air Waste* **43** 242
13. Mizuno A 2007 *Plasma Phys. Contr. Fusion* **49** A1
14. Kogelschatz U 2003 *Plasma Chem. Plasma Processing* **23** 1
15. Kogelschatz U, Eliasson B and Egli W 1997 *J. De Phys. IV* **7** C4-47
16. Fridman A, Chirokov A and Gutsol A 2005 *J. Phys. D: Appl. Phys.* **38** R1
17. Fridman A 2008 *Plasma chemistry*. New York: Cambridge University Press
18. Jarrige J and Vervisch P 2006 *J. Appl. Phys.* **99** 113303(1)
19. Khacef A, Cormier J M and Pouvesle J M 2006 *Eur. Phys. J. Appl. Phys.* **33** 195
20. Chen H L, Lee H M, Chen S H, Chang M B, Yu S J and Li S N 2009 *Environ. Sci. Technol.* **43** 2216
21. Oda T 2003 *J. Electrostat.* **57** 293
22. Delagrangé S, Pinard L and Tatibouet J M 2006 *Appl. Catal. B: Environ.* **68** 92
23. Fan X, Zhu T, Wang M and Li X 2009 *Chemosphere* **75** 1301
24. Subrahmanyam Ch, Renken A and Kiwi-Minsker L 2006 *Appl. Catal. B: Environ.* **65** 157
25. Kim H H, Oh S M, Ogata A and Futamura S 2005 *Appl. Catal. B: Environ.* **56** 213
26. Roland U, Holzer F and Kopinke F D 2005 *Appl. Catal. B: Environ.* **58** 217
27. Roland U, Holzer F and Kopinke F D 2002 *Catal. Today* **73** 315
28. Ayrault C, Barrault J, Blin-Simiand N, Jorand F, Pasquiers S, Rousseau A and Tatibouet J M 2004 *Catal. Today* **89** 75
29. Einaga H and Ogata A 2009 *J. Hazardous Mater.* **164** 1236
30. Karuppiyah J, Karvembu R and Subrahmanyam Ch 2012 *Chem. Eng. J.* **180** 39
31. Karuppiyah J, Sivachrandiran L, Karvembu R and Subrahmanyam Ch 2010 *Chem. Eng. J.* **165** 194
32. Karuppiyah J, Sivachrandiran L, Karvembu R and Subrahmanyam Ch 2011 *Chinese J. Catal.* **32** 795
33. Futamura S, Einaga H, Kabashima H and Hwan L Y 2004 *Catal. Today* **89** 89
34. Ogata A, Shintani N, Mizuno A, Kushiyama S and Yamamoto T 1999 *IEEE Trans. Ind. Appl.* **35** 753
35. Wang H, Li D, Wu Y, Li J and Li G 2009 *J. Electrostat.* **67** 547
36. Mok Y S, Nam C M, Cho M H, Nam I S 2002 *IEEE Trans. Plasma Sci.* **30** 408
37. Yamamoto T 1997 *J. Electrostat.* **42** 227
38. Ogata A, Ito D, Mizuno K, Kushiyama S, Gal A and Yamamoto T 2002 *Appl. Catal. A: General* **236** 9
39. Koutsospyros A D, Yin S M, Christodoulatos C and Becker K 2005 *IEEE Trans. Plasma Sci.* **33** 42
40. Bo Z, Yan J H, Li X D, Chi Y, Cen K F and Cheron B G 2007 *Plasma Chem. Plasma Processing* **27** 546
41. Guo Y F, Ye D Q, Chen K F and Tian Y F 2006 *Plasma Chem. Plasma Processing* **26** 237
42. Zhu T, Li J, Jin Y, Liang Y and Ma G 2008 *Int. J. Environ. Sci. Technol.* **5** 375
43. Huang H, Ye D and Leung D Y C 2011 *IEEE Trans. Plasma Sci.* **39** 576
44. Roland U, Holzer F, Poppl A and Kopinke F D 2005 *Appl. Catal. B: Environmental* **58** 227