Corrigenda

Complexes of Aminobenzylamines. Part I. Complexes of *o*-Aminobenzylamine with Copper(II), Cobalt(II) and Nickel(II)

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Experimental

Materials and methods

Anhydrous $CuCl_2$, $CuBr_2$, $NiCl_2$, $NiBr_2$, $CoCl_2$, $CoBr_2$ together with $NiCl_2 \cdot 6H_2O$, 2-aminobenzenonitrile and $Li[AlH_4]$ were purchased from Fluka, ROC/RK Merck and used without further purification.

I.r. spectra were recorded on Perkin-Elmer 577 and 597 spectrophotometers, in KBr pellets or Nujol mulls on KBr plates. Diffuse reflectance spectra were recorded in a Varian 634 spectrophotometer by diluting the samples with MgO. U.v.-vis. spectra were recorded on Bosch-Lomb-Shimadji Spectronic 210 W, Hitachi 100-70 and Cary 17 spectrophotometers using 1 cm rectangular quartz cells. E.s.r. spectra were recorded on a Varian Y-4502 Spectrometer using DPPH as internal reference. Molar conductance measurements were performed on a Metrohm Ltd., Herisau E-527 conductoscope. Magnetic susceptibility measurements were made by the Faraday method using $Hg[Co(CNS)_4]_2$ as reference. M.ps. are uncorrected. C, H and N elemental analyses were performed by Dr. Mantzos of the National Hellenic Research Foundation, in Athens. Halogens were determined potentiometrically or by precipitation as AgX (X = Cl or Br). Metal analyses was made by following the appropriate method either titrimetrically or by precipitation[48].

Preparation of the complexes

All the complexes described above, were first dried at room temperature in a vacuum desiccator containing $CaCl_2$. The drying was continued to constant weight at 90 °C under vacuum over P_2O_5.

The deuterated *o*-aba and its complexes were prepared by shaking small quantities with D_2O for 24 h.

The ligand *o*-aba was prepared according to the method of Amundsen and Nelson[49].

The preparation of the complexes is briefly described in the Table.

Complex	Starting materials	Solvents	Reaction time and temperature	Washing the precipitate	Yield %
$CuLCl_2$	$CuCl_2$ (anhydrous) (4 mmols) + Ligand (2 mmols)	EtOH + EtOAc, 1 : 1 proportion (20 cm^3)	5 h (25 °C)	EtOAc + petroleum ether	90
CuL_2Cl_2	$CuCl_2$ (anhydrous) (2 mmols) + L (4.4 mmols)	EtOH + EtOAc, 1 : 2 proportion (30 cm^3)	5 h (25 °C)	EtOAc + petroleum ether	90
CuL_2Br_2	$CuCl_2$ (anhydrous) (2 mmols) + L (4.4 mmols)	MeOH · EtOAc, 1 : 2 proportion	24 h (25 °C)	MeOH + petroleum ether	90
$CoLX_2$	CoX_2 (anhydrous) (4–6 mmols) + L (2 mmols)	EtOH + EtOAc (20 cm^3)	24 h (25 °C)	EtOAc + petroleum ether	70
CoL_2X_2	CoX_2 (anhydrous) (2 mmols) + L (4.2 mmols)	EtOH + EtOAc, 1 : 1 proportion (20 cm^3)	24 h (25 °C)	EtOAc + petroleum ether	75–85
CoL_3X_2	CoX_2 (anhydrous) (2 mmols) + L (6.3 mmols)	EtOH (20 cm^3)	24 h (25 °C)	EtOAc + petroleum ether	70–80
$NiLCl_2$	$NiCl_2$ (anhydrous) (4 mmols) + L (4 mmoles)	EtOH (20 cm^3)	24 h (25 °C)	EtOAc + petroleum ether	90
$NiLBr_2$	$NiBr_2$ (anhydrous) (3 mmoles) + L (3 mmoles)	EtOH (20 cm^3)	24 h (25 °C)	EtOH + petroleum ether	90
NiL_2Cl_2	$NiCl_2 \cdot 6H_2O$ (2 mmoles) + L (4.3 mmoles)	EtOH (20 cm^3)	12 h (25 °C)	EtOH + petroleum ether	80
NiL_2Br_2	$NiBr_2$ (anhydrous) (3 mmoles) + L (6.2 mmoles)	MeOH (20 cm^3)	12 h (reflux)	MeOH + petroleum ether	85
NiL_3Cl_2	$NiCl_2 \cdot 6H_2O$ (2 mmoles) L (6.5 mmoles)	EtOH + EtOAc (20 cm^3)	24 h (25 °C)	[b]EtOAc	90
NiL_3Br_2	$NiBr_2$ (2 mmoles) L (6.4 mmoles)	EtOH + EtOAc (20 cm^3)	24 h (25 °C)	EtOAc + petroleum ether	90
$Ni_2L_3Cl_2$	$NiCl_2$ (2 mmoles) L (4 mmoles)	EtOH (20 cm^3)	24 h (reflux)	EtOH + petroleum ether	70