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COMMUNICATION

Connecting chloride solvation with hydration in deep eutectic systems

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The Deep Eutectic Solvents/Systems (DESs) choline chloride:urea ($x_{\text{ChCl}} = 0.33$) and choline chloride:glycolic acid ($x_{\text{ChCl}} = 0.5$) were investigated using viscosity-corrected ^{35}Cl NMR spectroscopy and molecular dynamics simulations to probe the role of chloride as a function of water content. Three Cl⁻ solvation regimes are revealed, with high-symmetry environments for pure and highly dilute DES, and an unusual low-symmetry interstitial region where the primary coordination sphere is most disordered.

Eutectic mixtures show a melting temperature depression upon mixing, which increases the accessibility of their liquid window.¹ Deep Eutectic Systems (DESs) are widely applied as a subclass of eutectic mixtures,² defined by a eutectic temperature significantly below the ideal eutectic point.³ DES literature is mainly focused on extensively H-bonded mixtures,⁴ comprising a hydrogen bond acceptor (HBA) and hydrogen bond donor (HBD). Among other hygroscopic quaternary ammonium salts, choline chloride (ChCl), a non-toxic, widely available salt, is one of the most widely studied HBAs. The equilibrium water content of DESs ranges from ppm to significant mass fractions, making trace water unavoidable.^{5,6} This water acts as an additional H-bond donor and acceptor, modifying interactions between DES components,⁵ and strongly affecting crucial physicochemical properties and phase behaviour.

Developing fundamental understanding of intermolecular interactions in DESs will facilitate the design of task-specific systems. The structure of DES is complex and disordered, with competing strong and weak H-bonds between the components,⁷ whose contributions are poorly understood due to heavy system-dependence. A recurrent finding is that the anion, usually chloride, plays a primary role by defining the H-

bond network.⁷⁻¹⁴ In the archetypal DES ChCl:urea, the local structure sees chloride in constant H-bond flux with choline and urea, at the centre of a local cluster. The cation and HBD have different bonding affinities and short-range diffusion dynamics.⁹ The interaction between chloride and HBD also defines the structure of other DES such as ChCl:propylene glycol,¹⁵ ChCl:malic acid,¹⁶ and ChCl:oxalic acid.¹⁷ The anion therefore defines physicochemical properties of DESs,^{18,19} and affects the DES-solute interactions.^{20,21} Naturally, the distribution of species bound to chloride will vary as water content changes, but this is currently poorly understood.^{8,10,22}

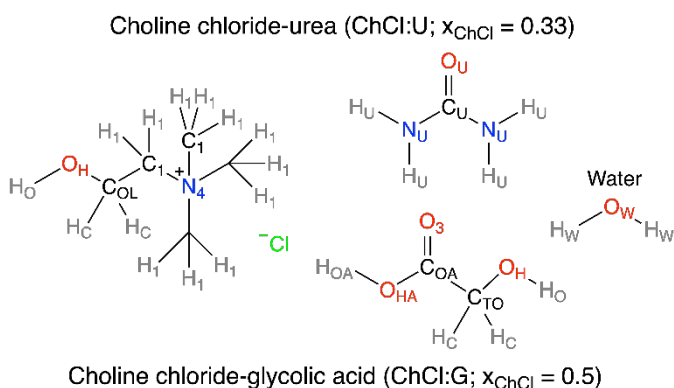


Figure 1. Chemical structure of the ChCl:U and ChCl:G DES studied in this work, and atom type labels for each molecule used in MD simulations.

Various works have investigated the H-bond network of pure and hydrated DESs,^{16,23,32,24-31} but not specifically the environment of chloride, and how it changes with hydration. A single paper has applied chlorine NMR to some choline chloride:glycol DESs,³³ which suggested weakened DES-DES H-bonds upon dilution with D₂O, from trends in ^{35}Cl linewidth. This study aims to directly explore how the key role of the anion is affected by the mixture's water content through connecting viscosity-corrected ^{35}Cl NMR and MD simulations. Two model ChCl-based DESs were chosen (Fig 1), choline chloride : urea (ChCl:U, $x_{\text{ChCl}} = 0.33$), and choline chloride : glycolic acid (ChCl:G, $x_{\text{ChCl}} = 0.5$); some critical discussion on DES nomenclature is included in the ESI †.

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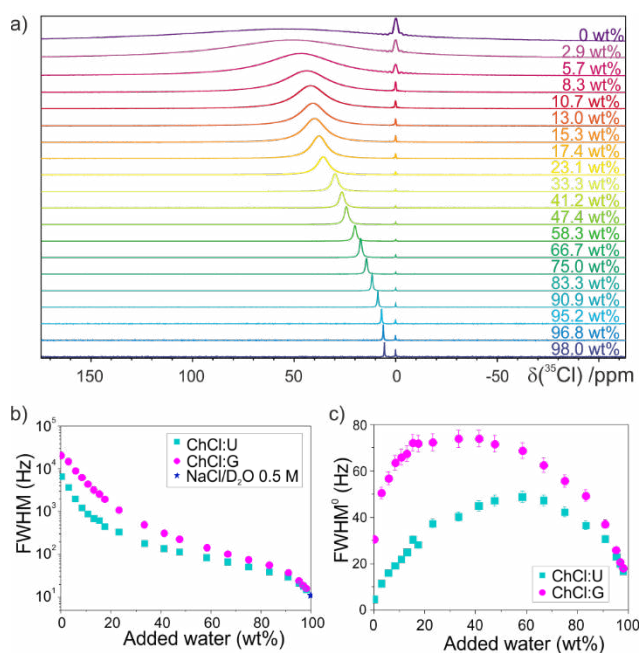


Figure 2. (a) Representative ^{35}Cl NMR spectra of ChCl:U samples as a function of added water content (increasing from pure DES at the top to 98.0 wt% H_2O at the bottom). The reference peak at 0 ppm corresponds to Cl^- of a 0.5 M NaCl solution in D_2O in the coaxial insert. (b) Experimental log-scaled linewidth (FWHM = full width at half maximum) and (c) viscosity-corrected linewidth (FWHM 0) for the ^{35}Cl signal of ChCl:U and ChCl:G samples as a function of the water content. FWHM of ChCl:U was corrected with viscosity data from Agieienko and Buchner.³⁴

^{35}Cl NMR spectroscopy has been used before to investigate H-bonding and Cl^- binding in a multitude of systems, with Cl^- chemical shift, linewidth and relaxation time characterising ion dynamics.^{35–43} Twenty samples of ChCl:U and ChCl:G were prepared at several dilutions, up to 98 wt% H_2O (experimental details in the ESI)†. Fig. 2a shows representative spectra of ChCl:U samples, with a large ^{35}Cl upfield shift of ~ 50 ppm for the most diluted sample relative to the pure DES (more details in Figs. S1–S3 and Table S5). This shift indicates substantially greater shielding of Cl^- upon dilution, consistent with the stronger $\text{Cl}\cdots\text{H}$ interaction with water molecules.⁴⁴ The ^{35}Cl linewidth is symptomatic of the Cl^- bond character and local environment.⁴⁵ A small quadrupolar coupling constant and thus narrow lineshapes are observed for highly symmetric scenarios, such as Cl^- in isolation or with a homogeneous ligand population, due to a minimised electric field gradient.³⁵ Decreasing the symmetry around Cl^- , such as by adding a different HBD to the ligand shell, increases the quadrupolar coupling constant and broadens the NMR linewidth.^{42,46} Experimental linewidths at half height (FWHM) of the ^{35}Cl peaks and corresponding relaxation times (T_2) and rates (R_2) were calculated and are shown in the ESI.† A marked drop in linewidth and relaxation rate are observed with increasing water content (Figs. 2b, S4 and S6). The log-scaled FWHM trend shows a steep initial decrease, a second region with lesser gradient, and a third regime with steeper gradient at high dilution; these findings are consistent with stronger aqueous H-bonding at high dilutions, where the asymmetrical network in the DES is replaced by spherically-symmetric hydrated Cl^- .^{39,47} Pioneering studies of long-chain alkylammonium chlorides showed a marked effect of the cations on the relaxation rate of

Cl^- , due to asymmetric hydration of the anion in the aqueous cage around the hydrophobic cation tail.⁴⁶

Because ion relaxation is closely related to the system viscosity,⁴⁸ we applied a viscosity-correction to the linewidth for these viscous DESs, shown in Fig. 2c as a function of water content for both ChCl:U and ChCl:G, using experimental viscosity data:⁴⁵

$$\text{FWHM}^0 = \text{FWHM}/(\eta/\eta^0) \quad (1)$$

where η is the viscosity and $\eta^0 = 1$ cP (1 mPa s). For ChCl:U, linewidth corrections were derived from interpolated viscosities reported at 298.15 K by Agieienko and Buchner.³⁴ We experimentally verified their data, as well as other datasets,^{49,50} at two ChCl:U–water dilutions (Fig. S7 and S10). Without literature examples, we measured the viscosity of ChCl:G at 298.15 K over the full hydration range, and interpolated the resulting data to correct the corresponding linewidths.

ChCl:G partially (auto-)esterifies,⁵¹ the resulting water release and presence of esters can significantly affect the viscosity and other physicochemical properties.^{51,52} Dilution with water decreases this deviation, by shifting the reaction equilibrium towards the DES components. A small ester peak was distinguishable in ^1H -NMR spectra, under the same conditions as the viscosity and ^{35}Cl -NMR measurements, for ChCl:G at low water contents (esterification yield of 1.5–4%). The initial latent water content (0.983–1.029 wt% for ChCl:U, 1.022–1.314 wt% for ChCl:G) slightly shift the viscosity corrections at low water contents. This was not considered for the ^{35}Cl NMR experiments, but we show that this effect on the FWHM 0 would be limited in Figs. S11 and S12.

The viscosity-corrected linewidths in Fig. 2c show a remarkable re-entrant behaviour as water content grows. The FWHM 0 first increases up to a maximum value (ca. 40 wt% for ChCl:G and ca. 60 wt% for ChCl:U), followed by a decrease at higher dilutions. In the most diluted samples, the linewidth is ~ 15 and 16 Hz for ChCl:U and ChCl:G, respectively (respective FWHM 0 of 17 and 18 Hz). As Cl^- approaches the Debye–Hückel limit of completely dissociated ions, i.e. predominantly hydrated Cl^- at these higher dilutions,⁵³ these FWHM match more closely the linewidths of ‘free ions’ (FWHM = 11 Hz for 0.5 M NaCl). Here, the ^{35}Cl coordination shell is symmetrical, leading to the narrowest possible linewidths. An artefact of our correction is a slightly narrower FWHM 0 for neat ChCl:U than for the most diluted aqueous samples (5 vs 17 Hz). This is due to the proportionality between ionic relaxation rate and macroscopic viscosity. Future correction strategies for such viscous systems, which do not over-correct pure DES due to direct normalisation, constitute an interesting avenue.

Molecular dynamics (MD) trajectories of 10 ns of both DES, from 0–90.9 wt% H_2O , were analysed to determine the Cl^- coordination environment. Interactions between DES components (i.e. Ch^+-Cl^- , $\text{HBD}-\text{Cl}^-$, and $\text{HBD}-\text{HBD}$) decrease, as expected, when water is added,⁵⁴ until the DES–DES hydrogen bonding network is disrupted, with onset around a volume fraction of 50 % H_2O .⁸ Further MD details, including densities, diffusion coefficients and methodology, are given in the ESI.† The trends in diffusivity agree with previous computations.⁵⁴

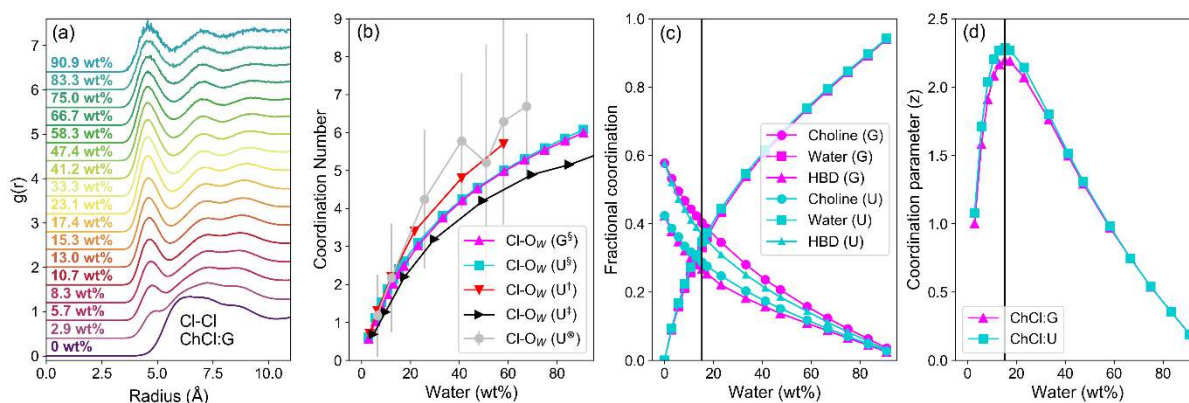


Figure 3. Data from MD simulations of ChCl:U and ChCl:G aqueous mixtures; (a) Cl–Cl RDFs for the ChCl:G DES as a function of water content (increasing from bottom to top); (b) Cl–O_w (water oxygen) coordination number as a function of water concentration for both systems, and compared with 3 literature studies of ChCl:U as a function of water content: ⁵Data from this study, ⁷Data from Kumari *et al.*,¹³ ⁸Data from Harries *et al.*,¹⁴ ⁹Data from Hammond *et al.*; (c) Fractional coordination of each species around Cl[−] in both DES as a function of water content. Fractional coordination is the proportion of each species relative to the sum of all, omitting Cl[−]–Cl[−] because it does not form contact pairs and thus should have minimal effect on the NMR environment; (d) calculated coordination parameter (*z*) as a function of water content for both DES, where the maximum shows the point at which the ligand environment is most varied. Note that the lowest water content used for the *z*-parameter is 2.9 wt%. Coordination numbers were calculated by integrating RDFs of atom types proximal to the molecular geometric centres; N_o for cholinium, O_w for water, C₇₀ for glycolic acid. A tie line is drawn on both (c) and (d) for the concentration in both DES (15.3 wt%) where *z*, the coordination parameter, is maximised; lines serve otherwise as a guide to the eye.

The models show a strong contraction in the Cl–Cl RDF as water content increases, which is most prominent in ChCl:G (Fig. 3a). The primary Cl–Cl interaction peak is diffuse in pure ChCl:G, spanning from *ca.* 4.8–10 Å. Once water is added, the Cl–Cl peak contracts to respective onset and first minima of *ca.* 4 and 6 Å. The closely-binned water contents studied here reveal this change even for the lowest water content of 2.9 wt% H₂O, which is easily achieved if DES are prepared without control over water content.⁵⁵ Other studies have recently noted similar phenomena in DES such as ChCl:U and ChCl:Ethylene Glycol,^{13,14,54} though this effect appears stronger in ChCl:G.

Even at very low water contents, water is a strong ligand for chloride, binding strongly and outcompeting DES species due to its H-bonding capability and small molecular volume. Water therefore occupies the space around Cl[−] anions, which helps to reduce Cl–Cl separation *via* bridging in solvent-separated pairs. At just 25 wt% H₂O, the Cl–H₂O coordination number is *≈*3, (Fig. 3b), so half-hydrated relative to [Cl(H₂O)₆].⁵⁶ The Cl–O_w coordination number follows the same trends for the two DESs, despite different structures and eutectic ratios. OPLS parameters with 0.8*q* scaling of ionic species gave good agreement with previous DES Cl–O_w coordination from MD^{13,14} and neutron scattering studies.⁸

The strong hydration of Cl[−] is therefore key, but the contribution from other components is considered *via* the fractional coordination (Fig. 3c), i.e. the proportion of each species as a fraction of the sum of coordination numbers, excluding chloride, which is too distant to occupy the primary sphere: The linewidth of Cl[−] under dilute aqueous conditions was shown only to be affected by nearest neighbours, and two layers of water molecules screen this effect.⁴⁵ ChCl:G has *x*_{ChCl} = 0.5 as compared to *x*_{ChCl} = 0.33 for ChCl:U. Hence, ChCl:G has significantly stronger Ch⁺–Cl[−] interaction than ChCl:U, whereas for the latter the HBD–Cl[−] interaction is more dominant. As a result, the fractional Cl[−] coordination of ChCl and HBD are inverted (Fig. 3c). In both cases the Cl–water fractional coordination follows the same trend, with the structure around Cl[−] rearranging to compensate for the systematic differences. In

both systems, a crossover point is then reached where water begins to dominate. We therefore developed a simple coordination parameter, *z*, to describe the solvation asymmetry around Cl[−], shown in Eq. 2 for our 3-component system:

$$z = \frac{2(\sum_{i \neq j} n_i)}{\sum_{i \neq j} \frac{n_i}{n_j}} = \frac{2(n_{cho} + n_w + n_{HBD})}{\left(\frac{n_w + n_{HBD}}{n_{cho}}\right) + \left(\frac{n_{cho} + n_{HBD}}{n_w}\right) + \left(\frac{n_{cho} + n_w}{n_{HBD}}\right)} \quad (2)$$

where *z* is the calculated coordination parameter, and *n_i* are the coordination numbers for each component in the shell, *n_{cho}*, *n_w* (water), *n_{HBD}*, derived from the models. *z* is a non-normalised numerical descriptor for the local environment of a ligated species. When one ligand dominates (such as water) *z* is minimised, and when every ligand is a different species, *z* is maximised. *z* is pertinent due to the difficulties in directional structural analysis around a spherically symmetric solute. This parameter is a proxy for the asymmetry around Cl[−], but the peak in *z* is lower in water concentration than the ‘true’ experimental (NMR) asymmetry due to multiple possible ligand orientations, the effect of which is more profound for concentrated DES.

The calculated *z* (Fig. 3d) follows the same re-entrant trend shown by ³⁵Cl NMR measurements. *z* peaks at 15.3 wt% H₂O for both DES, or a water fractional coordination of 0.33 (Fig. 3c). The environment becomes more symmetrical as water becomes dominant. This ‘intermediate’ water content, with the highest asymmetry of the Cl[−] environment, corresponds with other observations of large water–chloride ‘pearl-on-string’ clusters in ChCl:U,¹⁴ and the minimum melting temperature of ChCl:U:water ternary systems.⁵⁷ Though it is a bulk measure, the key role of Cl[−] is highlighted by the minimum in mixing enthalpy as a function of water content,⁵⁸ which corresponds with maximum Cl[−] asymmetry. The mixing enthalpy then increases to a maximum at 60–70 wt% H₂O. Initial added water is therefore energetically favourable, but this ceases when it has to directly compete with the HBD, Ch⁺, and Cl[−] interactions. At 25–35 wt%, Δ_{mix}*H* = 0 coincides with a fractional water coordination around the chloride of about 0.5 (Fig. 3c). Finally, the competition between the water and DES components weakens after reaching maximum mixing enthalpy and conductivity.³⁴

The unified NMR and MD analysis therefore highlights three distinct solvation regimes in DES/water mixtures:

1. In the pure/low-hydrated DES regime, the Cl⁻ environment is mostly symmetrical, with a remarkably homogeneous solvation shell of the HBD and Ch⁺.
2. Increasing water content to maximum z/FWHM⁰ (ca. 40 wt% for ChCl:G and 60 wt% for ChCl:U), the asymmetry around Cl⁻ is maximised, as water competes with the HBD and Ch⁺.
3. Above the maximum in z/FWHM⁰, water dominates the ³⁵Cl local environment, maximising symmetry. This reduces the net field gradient at the Cl⁻ nucleus, minimising linewidth.

In conclusion, viscosity-corrected ³⁵Cl NMR spectroscopy and MD simulations elucidated three hydration regimes of Cl⁻ in DES upon aqueous dilution. Adding water initiates asymmetric hydration around the Cl⁻, as it displaces DES components in the primary solvation shell, distorting the symmetry. This causes the initial rise in FWHM⁰, as the field gradient upon the nucleus increases. The asymmetry then decreases as water becomes the dominant ligand. Overall, these findings are in line with the debate suggesting structural heterogeneity on molecular scales in certain hydrated DESs,^{59,60} and the idea of different solvation regimes in the mixtures, where chloride (and the solvent as a whole) becomes partially and then fully hydrated.

Conflicts of interest

There are no conflicts to declare.

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