

CORROSION RESISTANCE OF AISI 316Ti STAINLESS STEEL SUBJECTED TO SHOT PEENING WITH AND WITHOUT PICKLING IN VARIOUS CHLORIDE ENVIRONMENTS

Summary

Surface conditions of materials affect the chemical, physical and mechanical properties of stainless steels. The main interest of this paper is to focus on corrosion properties of the AISI 316Ti stainless steel treated by shot peening and shot peening and pickling. In the last years shot peening has been widely used as a method improving fatigue properties of components under cyclical load. Because surface finishing has strong effect on corrosion susceptibility of stainless steels, e. g. pitting and created pits can initialize cracks, it is important to analyse corrosion behaviour of the material in question. In this study, the tested surfaces were characterized by SEM, EDX analyses, measurement of surface free energy and roughness. Corrosion properties were investigated in various environments by exposure, potentiodynamic cyclic and EIS tests.

Key words: shot peening, stainless steel, corrosion, pitting

1. Introduction

Stainless steels are very important construction materials because of their mechanical properties and corrosion resistance to various environments. Therefore, they are used in areas where high safety, reliability and lifetime are demanded (energy industry, medical instruments and implants, pharmaceutical and food processing industry etc.) [1,2]. In spite of the fact that stainless steels exhibit very good corrosion resistance in conventional conditions they are susceptible to local corrosion attacks (pitting) in specific environments, especially if chlorides are present. Their corrosion behaviour is affected by many factors, such as chemical composition, microstructure and surface condition of the material [3 - 5]. According to the authors [6] small dimensions of corrosion pits, short timescale of corrosion pits in metastable state, dynamic interplay between heterogeneous surface and changing potential and solution concentration gradients complicate the development of a complete understanding of the phenomena of local corrosion. Our study focuses on an investigation into the corrosion behaviour of the AISI 316Ti stainless steel with the surface treated by shot peening. The tests were carried out by applying several experimental methods in solutions with a constant Cl⁻ concentration with and without oxidizing agents.

The surface of stainless steels prepared by a mechanical or chemical method influences the corrosion properties by affecting passive layers properties, changing the surface reactivity and altering the near surface residual stress. Surface roughness resulting from the applied finishing processes can also be an important factor in the subsequent corrosion behaviour. As a result, surface treatment operations can alter the resistance of steel to the initiation and propagation of pitting. The mentioned problems have been studied extensively by many authors over the last 50 years on the basis of different laboratory approaches [7 - 10]. In the last years, surface treatment of stainless steels (e.g. shot peening) is investigated with the aim of improving their fatigue properties [11 - 13]. Shot peening changes the surface properties, therefore it is necessary to study its influence on susceptibility to a local form of corrosion where pits can operate as crack initializers [14]. Wang and Yu [15] found that higher corrosion resistance to pitting in chloride solutions can be achieved through surface nanocrystallization induced by high energy shot peening of the 1Cr18Ni9Ti stainless steel. Onizawa et al. [16] observed an increase in corrosion and fatigue resistance of the high nitrogen austenitic stainless steel (RS561) after shot peening was carried out. According to authors [17] a fine particle peening treatment on the 316L stainless steel prior to gas nitriding resulted in higher fatigue life. In this study, changes in the surface parameters of the AISI 316Ti stainless steel finished by shot peening and shot peening + pickling are evaluated as well as their effect on the corrosion properties in various solutions (constant content of chlorides but different concentration of oxidizing ions). Various experimental methods were used to compare the corrosion behaviour of the tested surfaces during different control phases of the corrosion process.

2. Experiments and results

To investigate the effect of shot peening and shot peening and pickling on corrosion properties of the AISI 316Ti stainless steel, corrosion tests with different mechanisms were used. Thus it was possible to compare corrosion resistance of specimens with different surface states in various environments.

2.1 Characterization of experimental material

The specimens of the experimental material were produced of the AISI 316Ti austenitic stainless steel with the chemical composition shown in Table 1. The microstructure of the original material investigated by using the optical metallographic microscope can be seen in Figure 1. The microstructure of the AISI 316Ti stainless steel is composed of polyedric austenitic grains with deformation twins. In Figure 1 δ -ferrite and Cr and Mo carbides were identified.

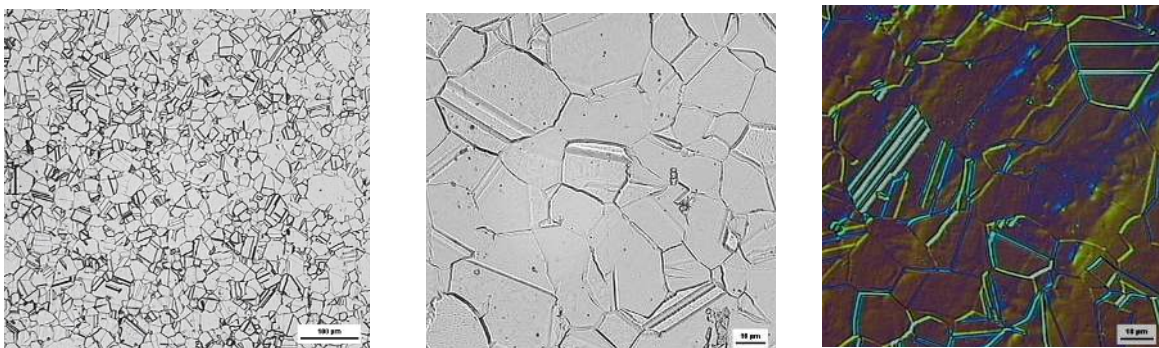


Fig. 1 Microstructure of cross section of AISI 316Ti steel

Table 1 Chemical composition of the AISI 316Ti steel

Element content [wt.%]	Cr	Ni	Mo	Mn	N	Ti	C	Si	P	S	Fe
	16.5	10.6	2.12	1.69	0.012	0.41	0.04	0.43	0.026	0.002	balance

2.1.1 Surface treatment of experimental specimens

The surfaces of specimens were prepared by shot peening with steel balls (diameter from 0.20 to 0.7 mm) of the steel with nearly the same chemical composition as the experimental material, the pressure was 0.4 MPa, the distance of the jet 220 mm. One half of the shot peened specimens was finished by pickling (in the solution of 1 portion of the hydrofluoric (HF) acid 40%, 20 portions of the 63% HNO₃ solution and 79 portions of distilled water), the exposure time was 30 minutes at a temperature of 22 ± 1°C. The character of the surfaces treated by shot peening (S) and shot peening and pickling (S+P) was studied by a scanning electron microscope (SEM) (Fig. 2). At the point of the steel ball impact the surfaces were plain and their surrounding had a very non-homogenous broken topography (Fig. 2a). According to the authors [11] the parameters of the atomic lattice are changed by this surface treatment and an increase in the residual stress in the surface is caused. It is also expected that the surface oxide layers are pressed under the surface by shot peening, some crevices are created and this can cause a danger of a local corrosion attack. Pickling commonly removes impurities from the surface and increases homogeneity and integrity of passive oxide layers. The effect of pickling on the surface character is evident and it can be seen in Fig 2b. In this figure, the crevices are deeper and the surface topography is more difficult. Changes in the chemical composition (S and S+P) made by the EDX analysis are presented in Table 2. The main change has been recorded in the content of oxygen (homogeneity of the passive layer).

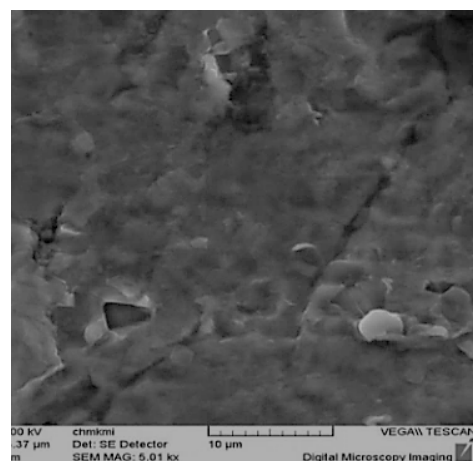
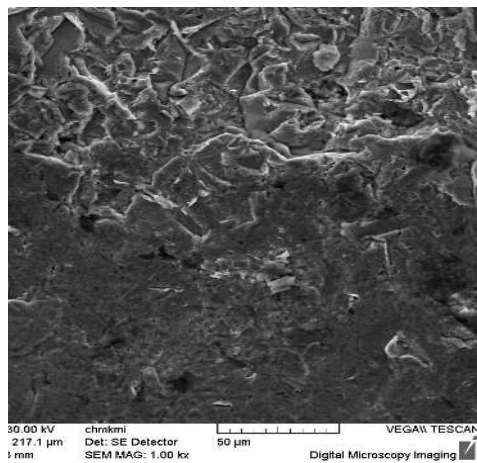


Fig. 2a SEM images of shot peened surface

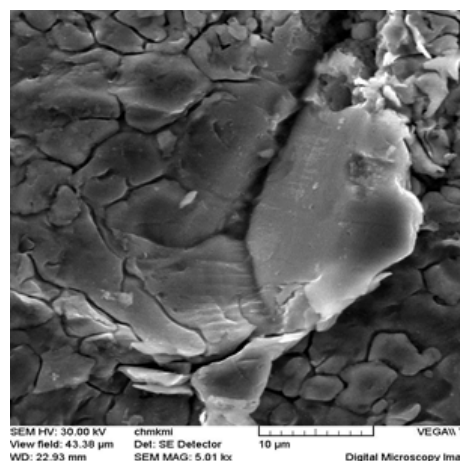
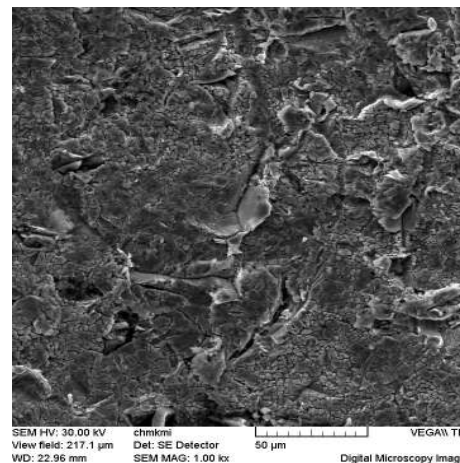


Fig. 3b SEM images of shot peened + pickled surface

Table 2 Chemical composition of the surfaces S and S+P

S			S+P		
Element	C norm. [wt. %]	C atom. [at.%]	Element	C norm. [wt.%]	C atom. [at.%]
O	6.90	20.12	O	2.45	7.89
Al	1.05	1.82	Al	0.25	0.48
Si	1.59	2.65	Si	0.63	1.15
Ti	0.40	0.39	Ti	0.42	0.45
Cr	13,40	12.02	Cr	15.65	15.52
Mo	1.85	0.90	Mo	1.57	1.48
Fe	65.43	54.65	Fe	68.19	62.96
Ni	9.37	7.45	Ni	10.11	8.88

The geometry of the examined surfaces was evaluated by measuring the mean values of roughness Ra (arithmetical mean deviation of the profile) and Rz (maximum height of the profile) and by comparing both ground and pickled surfaces, see Table 3. It can be seen that Ra of the surfaces finished by shot peening and pickling have higher values. The profiles of the tested surfaces are shown in Fig 3. The segmentation of the surface means also that the real surface area is larger. This fact can also affect the kinetics of the corrosion process.

Table 3 Roughness of the tested surfaces in the longitudinal and transverse direction

[µm]	G	G+P	S	S+P
Long. Ra	0.151	0.391	4.072	8.757
Trans.Ra	0.185	0.299	3.931	9.084
Long. Rz	1.69	3.46	35.1	52.8
Trans.Rz	1.89	2.44	38.3	55.2

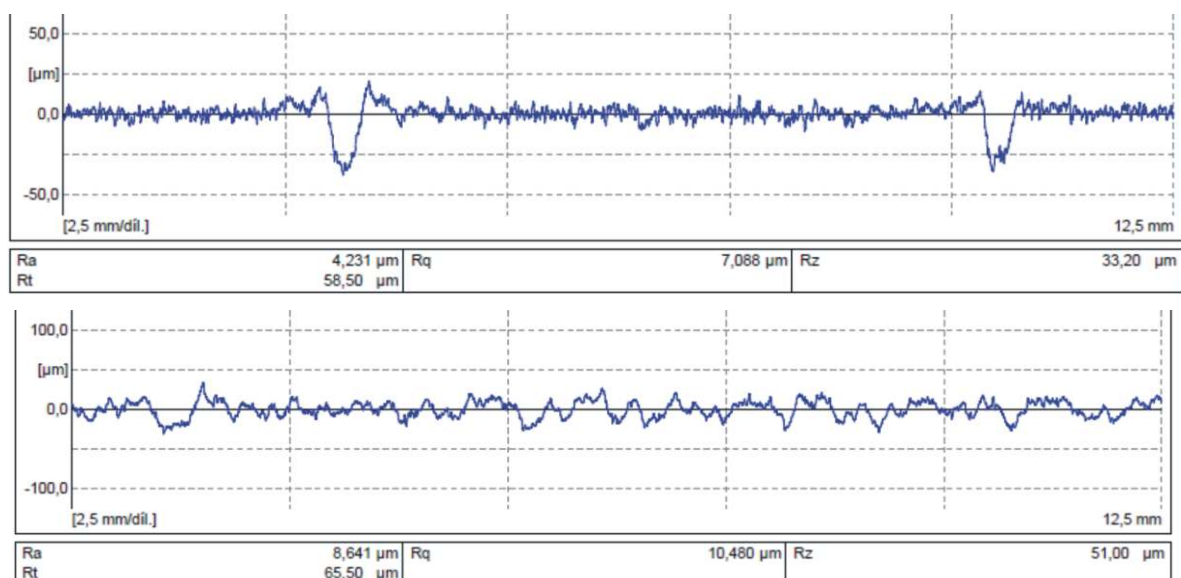


Fig. 3 Comparison of representative surface roughness determined on specimens: S and S+P surfaces in longitudinal direction

For the specimens S and S+P the surface free energy (SFE) was measured by applying the Owen-Wendt method because of its influence on the adsorption of reactants contained in the environment on the metal surface [18]. The mean value of the SEF of the shot peened specimens was 42.43mJ/m² and after pickling it grew to 95.45mJ/m². Comparing the SFE with a standard finished surface of the same material (grinding), the measured SFE of the ground surface was 43.92mJ/m² and after pickling it also rose to the value of 59.36mJ/m².

2.2 Corrosion testing

The corrosion properties of the variously treated AISI 316Ti stainless steel were studied by using three types of corrosion tests (exposure, potentiodynamic cyclic test, EIS). All the used tests differ with respect to the evaluation of the corrosion process in the same conditions and because of that it is possible to identify which parameters are effective for various surface states in various environments [19,20].

2.2.1 Exposure test

The tests were performed in a 5% NaCl solution with redox potential of 291 mV (solution 1)), in a 4.6% FeCl₃ solution with redox potential of 674 mV (solution 2) and in a solution which is a mixture of 50% of 4.6% FeCl₃ and 50% of 5% NaCl solution with redox potential of 671 mV (solution 3). The redox potential was read relative to a redox electrode SHE with $E = +207$ mV at a temperature of 23°C. The pH of solution 1 was 6.8, of solution 2 it was 1.12 and of solution 3 it was 1.51 due to a strong acidic reaction of FeCl₃. The reason why these solutions were chosen is their different oxidizing ability and pH at the same concentration of Cl⁻ ions. The experimental temperature during the test was 21±2 °C. The exposure time in the solutions was 72 hours. After exposure the specimens were cleaned in distilled water and dried so no corrosion products were left on the surfaces. The weight losses were determined with an accuracy of 10⁻⁵ g. The calculated main corrosion rates in solution 1, 2 and 3 are given in Table 4. The corrosion rates of specimen S in the NaCl solution are very low and after having undergone pickling (S+P) they were reduced almost five times. The corrosion rates in solution 2 are very similar for both tested surfaces, and pickling decreased the corrosion resistance of shot peened specimens negligibly. The differences in solution 3 after pickling have been determined to be higher. The size and density of the corrosion pits were determined according to the standard STN ISO 11463. On specimens S and S+P after exposure in solution 3 no differences in the pit size (2B) and density (2A) were established. In solution 2 on specimens S the pit size was 2B but of a higher density 3A, and on the pickled samples no change in size was identified but only in the density of corrosion pits. No corrosion pits were found in solution 1. In comparison with the corrosion rates of the AISI 316Ti steel treated by grinding and grinding + pickling in solution 1 negligible differences were found, but in solutions 2 and 3 the corrosion rates of ground and pickled specimens were about two times higher than that of the ground specimens. In both types of mechanical treatment pickling increased the corrosion rates in the environments containing oxidizing ions [21, 22]. The effect of surface roughness and the value of the SFE after pickling become evident in the oxidizing solutions.

Table 4 Corrosion rates of the tested specimens

Specimen/corr. rate (g.cm⁻².day⁻¹)	S	S+P
Solution 1	0.232 × 10 ⁻⁴	0.052 × 10 ⁻⁴
Solution 2	0.0180	0.0184
Solution 3	0.0116	0.0145

2.2.1 Potentiodynamic cyclic test

The potentiodynamic cyclic tests were carried out on a VoltaLab (Radiometer Analytical) device according to ASTM G61-86 (reapproved 2009) “Standard Test Method for Conducting Cyclic Potentiodynamic Polarization Measurements for Localized Corrosion Susceptibility of Fe, Ni, Co-based alloys”. Using this method the pitting potential E_{dp} (potential of depassivation) and the repassivation potential E_{rp} of the tested specimens were

measured. The measurements were made in solution 1. The conditions of the potentiodynamic cyclic test were: the initial potential was -200 mV, the potential maximum was 900 mV, and the rate of potential change was 10 mV/min measured relative to a saturated calomel electrode (SCE). The time for free potential stabilization of specimens was 5 minutes. The results are presented in Table 5 and the potentiodynamic curves can be seen in Fig. 4.

Table 5 Results of potentiodynamic test

specimen	E_{dp} (mV)	E_{rp} (mV)	E_{corr} (mV)	i_{corr} (mA/cm ²)
S	182.2	-241.3	-272.1	0.003
S+P	418.8	-158.7	-193	0.007

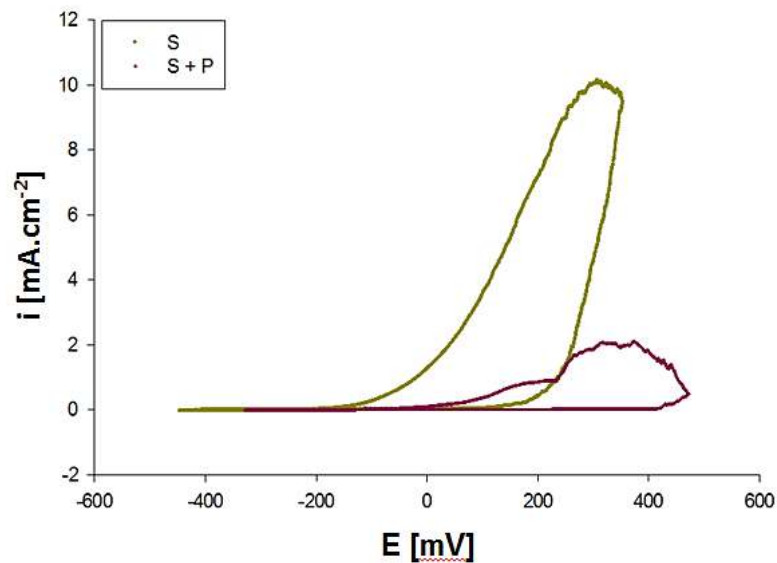


Fig. 4 Potentiodynamic curves of the tested AISI 316Ti steel with various finishes

Potential E_{dp} of the shot peened surfaces is evidently shifted after pickling to more positive values, which means that the shot peened surfaces exhibit higher electrochemical stability. This proves that pickling affects the protective effectivity of the passive layer. In comparison with the ground surface according to [23, 24], the measured electrochemical characteristics were worse. According to the potentiodynamic cyclic test results a positive effect of pickling is visible on the tested surfaces. In this type of test the character of a passive layer (purity, homogeneity) is very important.

2.2.2 Electrochemical impedance spectroscopy

The electrochemical impedance spectroscopy (EIS) experiments were carried out on a VoltaLab (Radiometer Analytical) device. The conditions of experiments were: the time for potential stabilization between specimen and electrolyte (solutions 1, 2, 3) was 10 minutes, the frequency was in a range from 100 kHz to 1 mHz with 10-times frequency change per decade, the amplitude of the alternating current AC voltage was 20 mV. The direct current (DC) voltage, which polarizes the specimen during the test, was set to the measured value of free potential after 10 minutes of stabilization. The EIS tests were performed at a laboratory temperature (21 ± 1 °C) by using a conventional three-electrode cell system with a reference SCE (Hg/Hg_2Cl_2 ; $KCl_{saturated}$) and a platinum auxiliary electrode (Pt). The impedance Z of the equivalent circuit was calculated for a typical simplified Randles cell [25, 26]. The values of the specimen polarization resistances R_p were determined from the Nyquist diagrams

(Figure 5-7) and the values of the chosen characteristic in solutions 1, 2 and 3 are given in Table 6. The diagrams show differences caused by different redox potential of the testing solutions with the same Cl⁻ concentration.

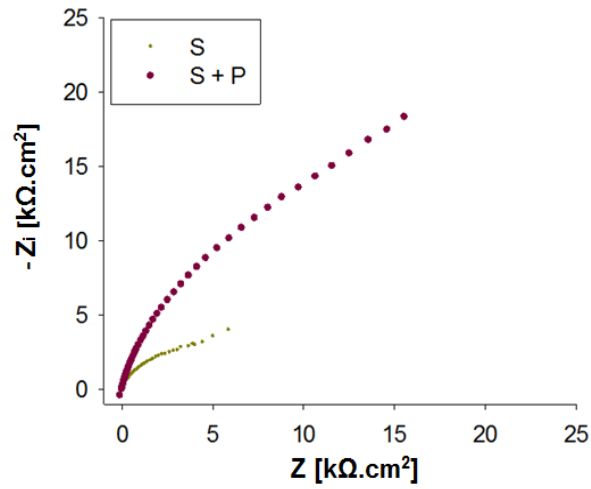


Fig. 5 Results of EIS test in solution 1

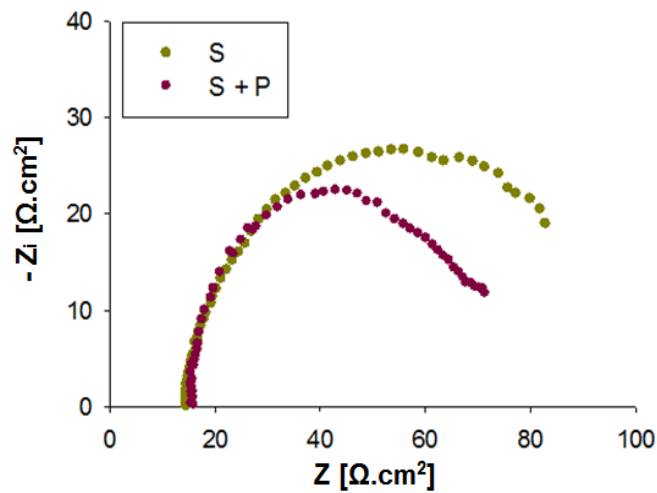


Fig. 6 Results of EIS test in solution 2

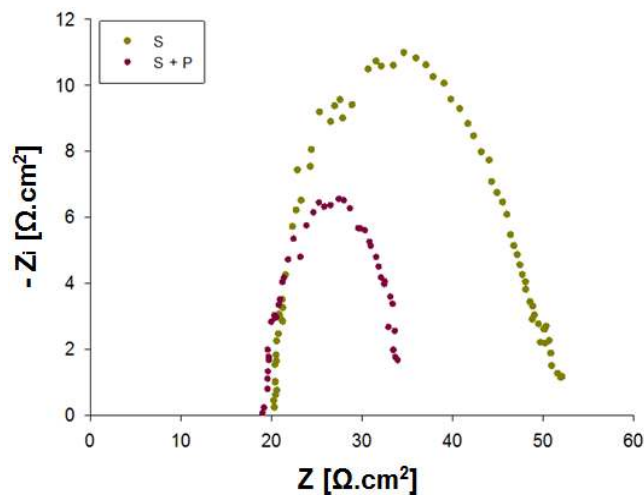


Fig. 7 Results of EIS test in solution 3

Table 6 Polarization resistance in solutions 1, 3 and 2

specimen\solution	NaCl/1 Rp [$\Omega \cdot \text{cm}^{-2}$]	FeCl ₃ /2 Rp [$\Omega \cdot \text{cm}^{-2}$]	NaCl+ FeCl ₃ /3 Rp [$\Omega \cdot \text{cm}^{-2}$]
S	11.31×10^3	31.38	66.79
S+P	112.6×10^3	15.06	55.13

The results of the EIS tests in the testing solutions suggest a positive influence of pickling in chloride solution 1 without oxidizing ions Fe⁺³, but in solutions 2 and 3 corrosion resistance falls as the amount of oxidizing ions increases. The results correspond with the results of the exposure tests.

3. Conclusions

- Shot peening in an aggressive chloride solution reduces corrosion resistance of the AISI 316Ti stainless steel to pitting in comparison with conventional ways of finishing (e.g. grinding). The difference in corrosion behaviour can be caused by changes in the surface characteristics e.g. bigger real surface area and different surface energy.
- The effect of the surface treatments of shot peening and shot peening and pickling of specimens on the corrosion behaviour in an only chloride solution was not dependent on the test method. In the solutions with the same chloride concentration but different oxidizing power and lower value of pH the pickling of the shot peened surface had a negative influence on corrosion resistance according to the results of the exposure and EIS tests. The different concentration of oxidizing ions (solution 2 and 3) did not affect substantially the corrosion behaviour of the specimens with different surface treatment.
- The application of tests with different control phases of the corrosion process (potentiodynamic, EIS and immersion tests) has enabled a more precise determination of corrosion behaviour in various environments and a comparison of the surface parameters effect.
- The AISI 316Ti stainless steel with the surface treated by shot peening is not especially susceptible to pitting corrosion in an environment containing chlorides; however, the present oxidizing reactants evidently increase susceptibility to pitting corrosion.

ACKNOWLEDGEMENT

The research is supported by the European regional development fund and the Slovak state budget within the project “Research center of University of Žilina – Second Phase”, ITMS 26220220183 and projects APVV No. APVV-14-0284 and APVV-15-0120.

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Submitted: 19.02.2016

Accepted: 02.02.2017

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