

# Electrochemical synthesis of oxide nanotubes on Ti6Al7Nb alloy and their interaction with the simulated body fluid

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**Abstract** The aim of the present work was to investigate electrochemical behavior of the Ti6Al7Nb alloy in the simulated body fluid (SBF) containing  $\text{Ca}^{2+}$ ,  $\text{HCO}_3^-$ , and  $\text{HPO}_4^{2-}$  ions. At first, optimal conditions necessary for oxide nanotube formation were determined. The experiments were conducted in the 1 M  $(\text{NH}_4)_2\text{SO}_4$  with 0.5 wt%  $\text{NH}_4\text{F}$  electrolyte at room temperature. Anodization of the alloy samples was carried out under variable external voltage  $U$  in the range from 10 to 40 V at room temperature. Obtained surface morphology was examined by SEM and X-ray techniques. Nanotube diameter was calculated and correlated with the imposed voltage. Having control over the size of nanotubes, samples with the obtained nanostructures of a chosen diameter were immersed into SBF solution with  $\text{pH} = 7.4$  for a fixed period of time. Then, they were removed from the fluid and subjected to the electrochemical investigation. Corrosion current and corrosion potential were determined, and it was found that the best anticorrosion properties were obtained for heat-treated nanotube layer:  $i_{\text{corr}} = 39 \text{ nA/cm}^2$  and  $E_{\text{corr}} = -0.236 \text{ V vs Ag/AgCl}$ . Finally, the interaction between the oxide surface and the solution was studied using polarization and electrochemical impedance spectroscopy (EIS) techniques.

**Keywords** Ti6Al7Nb alloy · Oxide nanotubes · Electrochemical measurements · Apatite formation

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## Introduction

Due to aging of our society, there is an increased need for implant manufacturing, and consequently, various materials like stainless steels, CoCr alloys, and recently, titanium (Ti)-based alloys have been tested as biomaterials in orthopedic and dental applications [1]. Since these materials must be biocompatible, i.e., non-toxic and they should not cause any inflammatory or allergic reactions [2], titanium alloys seem to be the best candidate for the majority of these applications.

It has been already demonstrated that titanium and Ti-based alloys are highly corrosion resistant in different environments [3] including body fluids [4]. In most aqueous solutions, they form a stable passive oxide layer, which is responsible for this particular property. Titanium itself, however, has high bulk modulus (~100 GPa) which makes this metal not compatible with mechanical properties of the bones. Therefore, other non-toxic elements like Ta, Zr, Nb, and Pd were considered as the components of engineered biomedical alloys which exhibit reduced bulk modulus. First Ti alloy implemented as a biomaterial was Ti6Al4V alloy originally developed for aerospace applications. It was soon discovered that a multicomponent alloy Ti35Nb7Zr5Ta has bulk modulus as low as 55 GPa, which is close to that of the bone (30 GPa) [5]. However, the possibility of adjusting mechanical properties by varying alloy composition raised the question of their electrochemical behavior, since it may not be obvious whether passive layer of  $\text{TiO}_2$  formed on the Ti surface remains unchanged after alloying.

To answer that question, a corrosion behavior of Ti alloys in artificial bioliquids was investigated in a number of studies. Assis and Costa [6] compared electrochemical behavior of Ti13Nb13Zr, Ti6Al4V, and Ti6Al7Nb alloys in Hank's solution and found that their behavior is very similar. They used long-term immersion tests and employed both, electrochemical impedance spectroscopy (EIS) and polarization methods.

Experiments showed that these alloys exhibit very high corrosion resistance in this electrolyte, which is maintained by developed dense, passive oxide layer. Popa et al. [7] correlated microstructure and corrosion behavior of Ti6Al7Nb alloy in Ringer and Ringer–Brown solutions with different pH: 2.49, 6.9, and 8.9. All samples showed spontaneous passivity with passive current densities lower than pure Ti. Looking for vanadium-free alloys, Robin et al. investigated corrosion behavior of a series of Ti13Nb13Zr alloys with variable Nb content in Ringer solution [8]. Comparing the results of electrochemical studies with those obtained for Ti and Ti6Al4V alloys, they found that the resistance of a passivating film is higher for Ti13Nb13Zr alloy, which indicates higher corrosion resistance. In the similar study, Al-Mobarak et al. [9] compared corrosion behavior of titanium alloy Ti6Al7Nb with that determined for Ti6Al4V alloy. For tests in Hank's solution, they found that the alloy with niobium has higher corrosion resistance. They concluded that Ti6Al7Nb alloy can be used instead of the alloy containing vanadium, if the problems connected with vanadium release have to be avoided.

The influence of a microstructure on corrosion characteristics of Ti13Nb13Zr alloy in Ringer solution was investigated by Cvijovic–Alagic et al. [10]. They found that this alloy exhibits spontaneous passivity; however, the protective layer is formed more rapidly on Ti6Al4V alloy. The corrosion resistance of Ti13Nb13Zr alloy is very similar to that of martensitic Ti6Al4V alloy. Simka et al. [11] pointed out that surface treatment by electrolytic polishing influences corrosion resistance. They studied this influence on Ti13Nb13Zr alloy and found that after electropolishing, the corrosion resistance increased significantly. Apparently smooth surface stimulated oxide layer formation. As far as corrosion resistance is concerned, the addition of noble metals to Ti may also have beneficial effect. We found that Ti–Pd alloy in phosphate-buffered saline (PBS) solution with H<sub>2</sub>O<sub>2</sub> addition simulating the inflammatory state has better corrosion resistance than pure Ti [12].

The formation of the oxide layer on the alloy surface is not enough to conclude whether this material will bond to the living bone. It is known that biological integration requires good bone formability. One may suggest that bioactivity of the material can be predicted from the apatite formation on its surface in a simulated body fluid (SBF) [13]. It was also recently demonstrated that to assure good osseo-integration, TiO<sub>2</sub> nanotubes formed on the Ti surface enhance hydroxyapatite formation, which in turn bonds to a living bone [14, 15]. Tsuchiya et al. [16] showed that self-organized tubular oxide layer can be formed on TiAl alloys. Adhesion of hydroxyapatite coating on Ti6Al4V implant was extensively analyzed by Mohsendi et al. [17]. They found that using an interfacial layer made of TiO<sub>2</sub> can enhance the bonding strength. Consequently, the aim of the present study is twofold:

-first, to find proper conditions for the formation of the tubular self-organized oxide layer on the Ti6Al7Nb alloy surface, and to determine its corrosion behavior in SBF solution; and

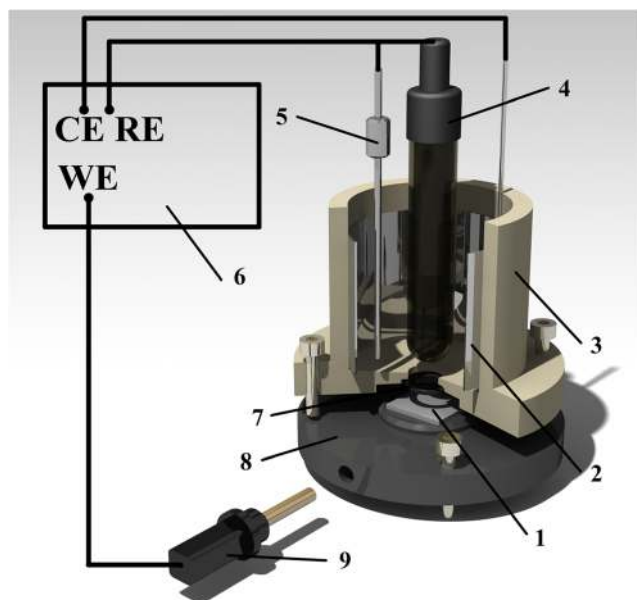
-second, to check if this oxide layer assures enhancement of apatite formation on its surface.

## Experimental

Titanium foils (99.7 % purity 1.0-mm thick, Alfa Aesar) and Ti6Al7Nb alloy (ISO-5832-11, extruded rod 22 mm, Bibus Metals) were cut into slices on the cutting wheel. They were mechanically polished on cloth to 1 μm and ultrasonically cleaned using deionized water and acetone. The samples having area of 1.13 cm<sup>2</sup> were anodized in electrolyte containing 1 M (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (POCH) and 0.5 wt% NH<sub>4</sub>F (POCH) [18]. During all anodization processes, the electrolyte was stirred using magnetic stirrer. Anodization was carried out at room temperature in two-electrode electrochemical cell with platinum sheet as counter electrode (CE) and Ti6Al7Nb alloy as a working electrode (WE). As a power supply, programmable DC power supply Agilent N5751A with digital multimeter Keithley 2100 was used. This procedure was similar to that reported in our previous paper [19] on the synthesis of ZrO<sub>2</sub> nanotubes. After anodization, samples were rinsed with deionized water and ultrasonically cleaned and dried in air stream. Part of the samples was annealed in air at 200 °C for 8 h to remove residual traces of fluorine and moisture.

Polarization, open circuit potential (OCP), and EIS measurements were performed using an AUTOLAB Potentiostat, PGSTAT 128N with FRA2 (frequency response analyzer) module controlled by NOVA 1.8 electrochemical interface. Usually, each experiment was repeated three times. Three-electrode electrochemical cell with Ag/AgCl as reference electrode (RE) was used (Fig. 1). The working electrodes consisted finally of disks with active surface area of 0.50 cm<sup>2</sup> isolated by O-ring. All the electrochemical investigations were carried out in SBF solution at pH 7.4, prepared according to the recipe given by Kokubo and Takadama [13]. It contains the mixture of NaCl (8.035 g), KCl (0.225 g), NaHCO<sub>3</sub> (0.355 g), K<sub>2</sub>HPO<sub>4</sub>·3H<sub>2</sub>O (0.231 g), MgCl<sub>2</sub>·6H<sub>2</sub>O (0.311 g), 39 ml 1 M HCl, CaCl<sub>2</sub> (0.292 g), Na<sub>2</sub>SO<sub>4</sub> (0.072 g), and tris-hydroxymethyl aminomethane ((HOCH<sub>2</sub>)<sub>3</sub>CNH<sub>2</sub>) (6.118 g) dissolved in 1000 ml demineralized water. The electrolyte was maintained at constant temperature 36.6 ± 0.2 °C and not stirred during all experiments and measurements. All electrolytes were prepared using analytical purity grade chemicals and deionized water. Open circuit potential was measured for 12 h with sampling rate 10 s.

The linear polarization method was used to determine the corrosion current densities ( $i_{\text{corr}}$ ) and corrosion potentials ( $E_{\text{corr}}$ ). Polarization potential ±500 mV vs. OCP was applied after 24-h sample immersion, (in order to obtain stabilization of



- 1 - Working electrode
- 2 - Counter electrode - Pt sheet
- 3 - PTFE cell body
- 4 - Reference electrode - Ag/AgCl
- 5 - Pt rod in series with capacitor
- 6 - Potentiostat
- 7 - O-ring
- 8 - SS ANSI 316Ti base
- 9 - Connector

**Fig. 1** Three-electrode electrochemical cell used in the experiments

the potential and wetting of the tubes surface by SBF) at the scan rate of 1 mV/s. Both,  $E_{\text{corr}}$  and  $i_{\text{corr}}$ , were calculated using NOVA 1.8 software. The parameter which determined the approach towards stationary potential was the rate of its change recorded with time. For a rate smaller than 0.001 mV/s, it was assumed that stationary potential has been reached.

The electrochemical impedance test at OCP after 24 h of immersion was performed, and experiments were repeated every 7 days up to end of immersion test. EIS spectra were acquired in the frequency range of 100 kHz–0.01 Hz for 10 points per decade using single sine wave of 10-mV peak to peak amplitude. The results were analyzed with the fitting program ZVIEW version 2.0.

The immersion test was conducted in SBF at human body temperature of 36.6 °C for 28 days. SBF was exchanged every 24 h by pouring out the solution and pouring in the fresh one. After the immersion test was completed, the sample was gently rinsed with demineralized water and dried in air stream.

The morphology of the anodized sample was investigated by using scanning electron microscope (SEM) (Hitachi SU-70). Structural analysis of a nanotube layer was carried out with an X-ray diffractometer (Rigaku MiniFlex II) using monochromatic Cu K $\alpha$  radiation (0.15416 nm). All samples were studied by point analyses using EDS analysis (Thermo Scientific) attached

to the SEM to obtain chemical composition of the nanotubular layer. The X-ray photoelectron spectroscopy (XPS) spectra were registered with the use of photoelectron spectrometer Versa Probe II (PHI Electronics). To check compatibility of the initial sample composition with the equilibrium phase diagram, an isothermal section of Ti–Al–Nb system was calculated at room temperature with ThermoCalc software.

## Results

### Ti6Al7Nb alloy phase composition

Investigation of the alloy samples by SEM back-scattered electrons (BSE) and X-ray diffractometry showed that the alloy used in our experiments is composed of two phases: a hexagonal close-packed (hcp)  $\alpha$  phase and a body-centered cubic (bcc)  $\beta$  phase. This is demonstrated in Fig. 2a–e. These results are compatible with the similar alloy analysis reported by Sitting et al. [20].

A comparison of this result with the equilibrium phase diagram calculated for Ti–Al–Nb system at room temperature (Fig. 3) indicates that this alloy is in fact in nonequilibrium state. At this composition ( $X_{\text{Ti}} = 0.86$ ,  $X_{\text{Al}} = 0.10$ ), and at room temperature,  $\alpha$  hcp phase should be in equilibrium with the ternary  $\text{O}_2\text{--Ti}_2\text{NbAl}$  phase. More detailed analysis of the phase diagram shows that this ternary phase is formed in eutectoid reaction from bcc beta phase at 1260 K. It seems that identified in the sample beta phase is the result of the alloy preparation process, in which ternary phase did not manage to precipitate yet. Pandat Software [21] and the data optimized by Witusiewicz et al. [22] were used to complete these calculations.

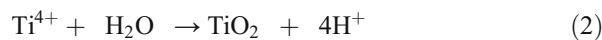
### Nanotubes formation

After preparation of the alloy samples, they were anodized to form oxide nanotube layer on the alloy surface.

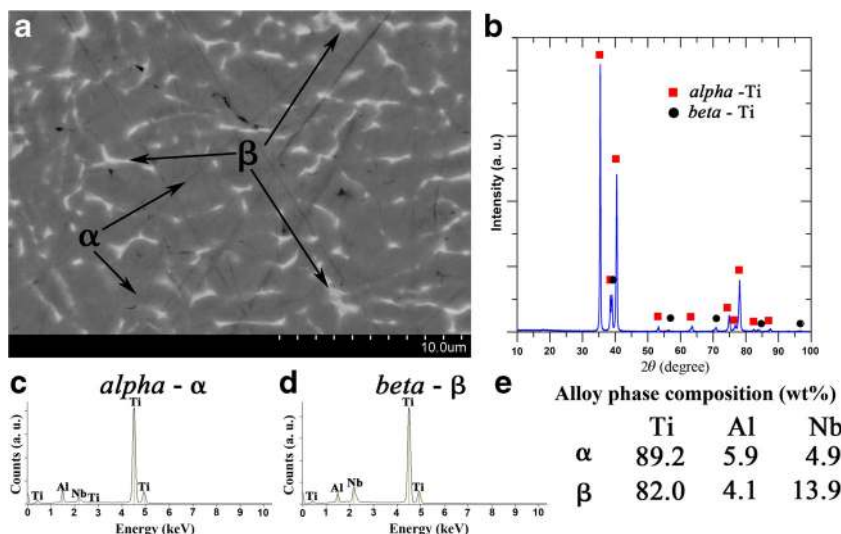
Anodization of the samples was carried out in the voltage range from 10 to 40 V, and nanotubes morphology was obtained in the range from 10 to 35 V.

The process of nanotube formation on titanium-rich alloys has been already described in the literature by Macak et al. [23], Crawford et al. [24], Roy et al. [25], and Rafieerad et al. [26]. It consists of several steps which can be described as follows:

1. field-assisted dissolution of titanium takes place due to the very high electric field across the electrode:

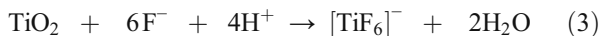


**Fig. 2** Investigation of the Ti6Al7Nb alloy: **a** SEM-BSE image, **b** XRD spectrum, **c** EDS spectrum for  $\alpha$  phase, **d** EDS spectrum for  $\beta$  phase, and **e** alloy composition based on EDS analysis results



During this stage,  $\text{TiO}_2$  is formed at the metal/electrolyte interface.

- in the fluoride ions containing acidic electrolyte, dissolution of the oxide layer takes place. As field-assisted growth slows down due to increasing oxide thickness, competing chemical reaction proceeds:



This localized dissolution is responsible for pores formation.

- the transfer of  $\text{Ti}^{4+}$  ions from metal to the oxide surface is accompanied by the transfer of  $\text{O}^{2-}$  and  $\text{F}^-$  ions to the metal/oxide interface. It was shown by Habazaki et al. [27] that the fluoride ions migrate inwards at a rate twice that of  $\text{O}^{2-}$  ions. They accumulate at the interface between

the alloy and anodic oxide where a thin  $\text{TiF}_4$  layer is formed between the metal and the oxide as well as on the outer sides of the pores.

- Volume expansion of the oxide and fluoride may result in the plastic flow, which pushes the fluoride-rich layer upwards. Such a model was suggested by Garcia-Vergara et al. [28]. It is the beginning of the tube formation. When rates of oxide formation and oxide dissolution became equal, the steady state is reached inside each pore resulting in steady tube growth.

Figure 4 shows SEM top view of the nanotube layer formed during the experiments after 2 h of anodization at room temperature

In this voltage range, the outer tube diameter varied approximately in the range from 50 to 180 nm, while the inner tube diameter varied in the range from 20 to 100 nm. In both cases, the diameter increases approximately linearly with an increasing anodization voltage, but at a different rate. These results of diameter measurements are shown in Fig. 5.

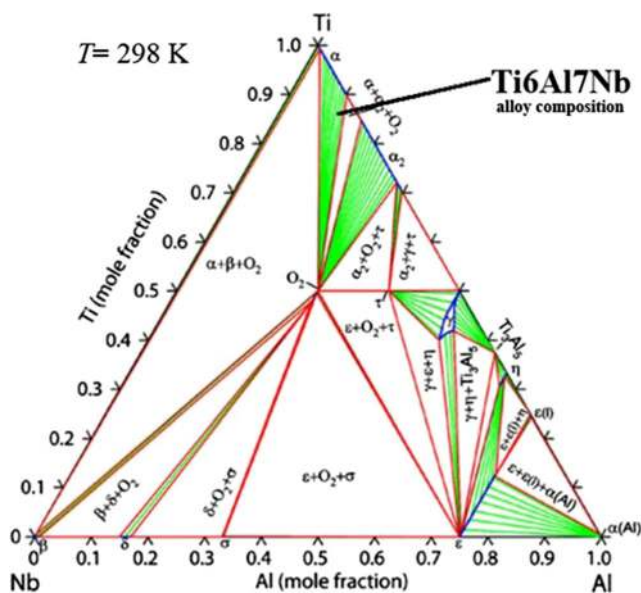
We found that the increase of the diameter can be described with the following equations:

$$d_{\text{inner}} = 3.75 \times U (V) - 14.33 \text{ (nm)} \quad (4)$$

$$d_{\text{outer}} = 5.55 \times U (V) - 12.91 \text{ (nm)} \quad (5)$$

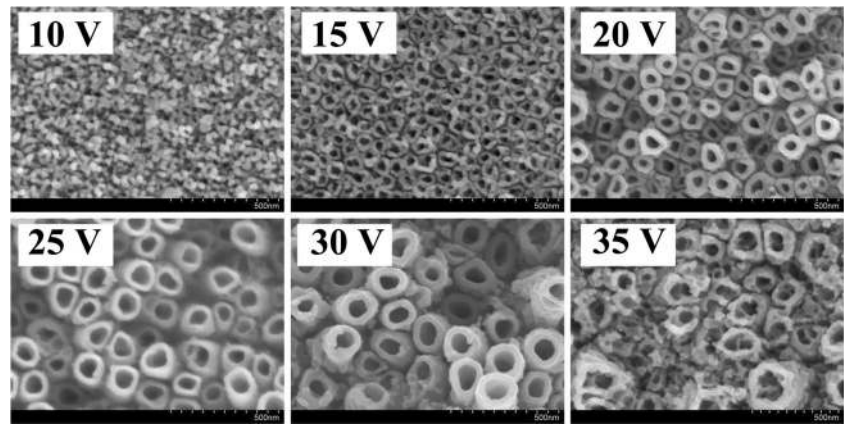
Another characteristic feature visible after anodization is the oxide surface morphology which shows the patches of another oxide formed on the alloy surface. The comparison of this result with the microstructure of the alloy before anodization suggests that another oxide has grown on  $\beta$  phase grains (Fig. 6a, b) enriched in Nb.

XPS spectra of as-synthesized nanotube array and chemical composition show (Fig. 6c) that the oxide is composed predominantly of  $\text{TiO}_2$ , but Al and Nb oxides are also present. The results indicate that Ti is present in the  $\text{Ti}^{4+}$  state. The Ti



**Fig. 3** Calculated Ti–Al–Nb phase diagram

**Fig. 4** SEM top views of TiO<sub>2</sub> nanotube layers formed after 2 h of anodization



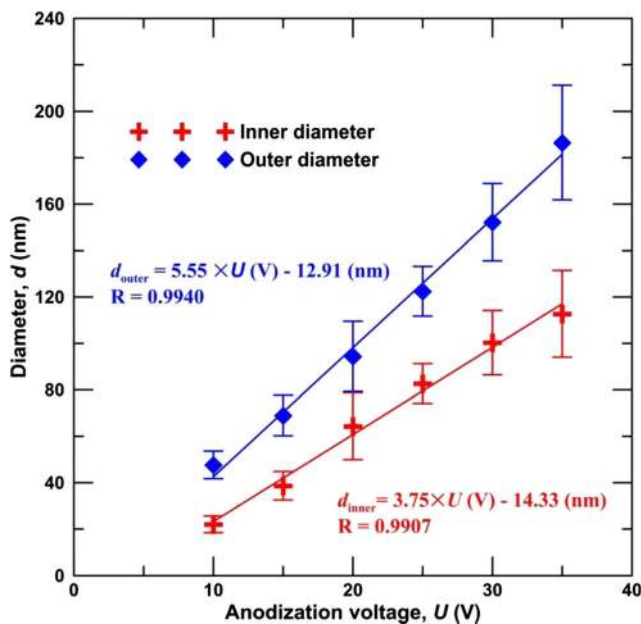
2p doublet structure with 2p<sub>3/2</sub> peak centered at 458.5 eV confirms the presence of TiO<sub>2</sub> [29].

**OCP measurements**

One of the ways to study the corrosion behavior of the materials is to check the open circuit potential (OCP) as a function of time. OCP variation versus time in SBF solution was recorded for bare Ti and bare Ti6Al7Nb and for Ti6Al7Nb samples covered with the oxide nanotube layers without heat treatment (Ti6Al7Nb–NT) and after heat treatment (Ti6Al7Nb–NT–HT). Potential evolution with time was monitored for 12 h. The results of these measurements are shown in Fig. 7.

It is seen that the potential of untreated samples (Ti and Ti6Al7Nb) changed in positive direction approaching after 12 h the value of –0.116 V vs. Ag/AgCl for Ti, and –0.040 V vs. Ag/AgCl for Ti6Al7Nb alloy. In this case, the shift of the potential towards the nobler values reveals that on

the surface of metals, a protective oxide film is being formed. It is known that after immersion into fluoride-free electrolyte, compact oxide layer is formed on the surface. To produce oxide nanotubes layer, the samples were anodized in fluoride-containing electrolyte at one potential of 30 V. This potential was chosen for further study since; on one hand, well-developed round shape of nanotubes is formed between 20 and 30 V. On the other hand, bigger diameter assures better electrolyte penetration. It can be seen in the same Fig. 7 that for porous nanotube surface on Ti6Al7Nb–NT obtained by anodization process at 30 V, the potential decreased slightly from 0.057 V vs. Ag/AgCl after immersion to –0.007 V vs. Ag/AgCl at the end of this experiment. This trend of potential falling can be attributed either to progressive surface wettability inside the tubes or reaction with the residue of fluoride. Samples of Ti6Al7Nb–NT–HT after heat treatment exhibit almost constant potential vs. Ag/AgCl on 0.000 V level. For all the investigated electrodes, the potential changes (rapidly for samples without any treatment) during the first 30 min and then, is slowly approaching constant value. The results demonstrate that the alloy covered with oxide nanotubes exhibits the most positive OCP potential in the environment under investigation. This may mean that NT and NT–HT covers improve corrosion resistance of Ti6Al7Nb alloy, as compared with pure titanium and bare alloy.

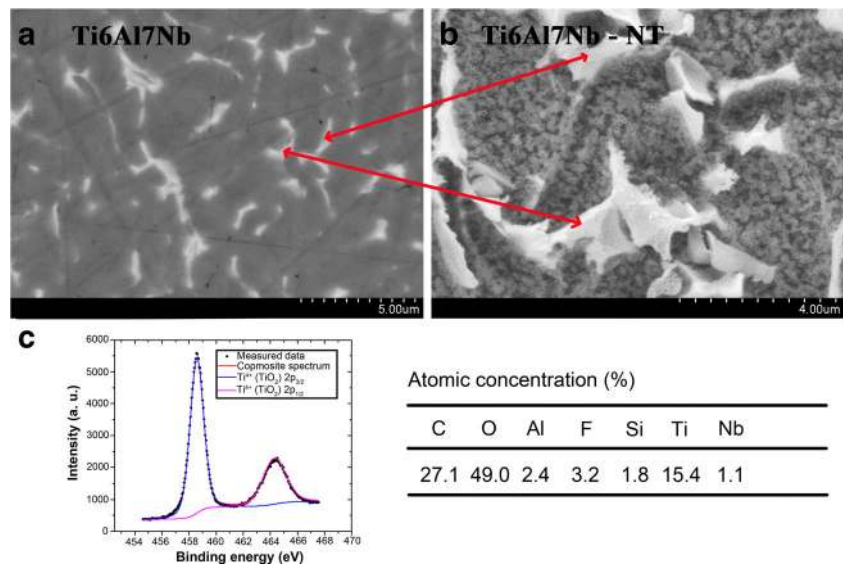


**Fig. 5** The change of nanotube diameter with anodization voltage

**Polarization measurements**

Recorded polarization curves are shown in Fig. 8. Determined corrosion currents and corrosion potentials based on Tafel plots are gathered for all the samples in Table 1. The results show that the poorest corrosion resistance exhibits the alloy covered with as-synthesized nanotubes without heat treatment. It is known that except for Zr, porous amorphous oxide layers are formed during anodization of Ti, Nb, Ta, and W [30]. Similar amorphous structure was recently observed after anodization of Ti6Al7Nb alloy [26]. However, in case of the alloy, Ghicov et al. [31] demonstrated that the presence of another oxide in the oxide matrix enhances the stability of

**Fig. 6** Alloy surface morphology before (a) and after anodization (b). *White fields* correspond to  $\beta$  phase enriched in Nb. **c** XPS spectra and chemical composition of as-synthesized nanotube array



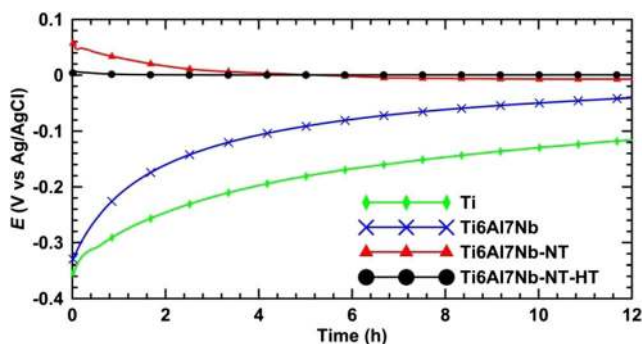
the crystallite phase. In our case, X-ray patterns obtained for as-anodized samples revealed also an amorphous oxide, which transformed into anatase at about 450 °C. Thus, our polarization experiments indicate that the formation of amorphous oxide nanotubes on the investigated alloy increases the corrosion current.

Samples of Ti, bare alloy, and alloy with heat treatment nanotubes show similar corrosion potentials and also similar corrosion current densities (Table 1). However, the lowest corrosion current density was recorded for the alloy with heat-treated nanotubes. It demonstrates the best corrosion resistance. The polarization curves obtained for the alloy covered with nanotubes after heat treatment at 200 °C show positive shift as compared to pure Ti indicating better corrosion resistance. However, the alloy samples covered with nanotube layer without the heat treatment, as it was already mentioned, indicate poor corrosion resistance. It may mean that porous oxide layer does not shield the transport of electrons and ions between the metallic substrate and the solution.

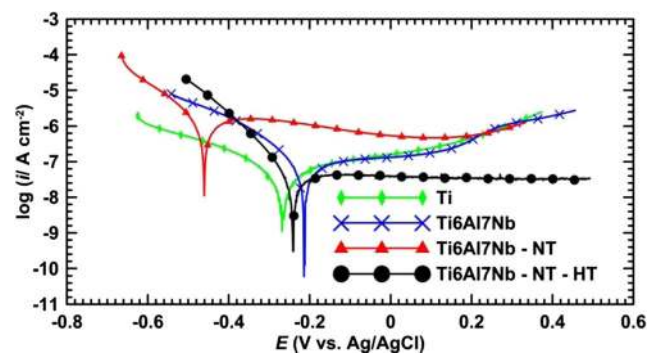
These results are similar with the results obtained for Zr by Wang and Luo [32]. They pointed out that the nanotubes

increase oxide surface area significantly and provide more channels for the electrolyte to reach the barrier layer. It results in a significant decrease of anticorrosion properties of the obtained oxide film and explains high value of the corrosion current for the alloy sample after anodization.

The fact that the heat treatment changes significantly the corrosion behavior of oxide nanotubes must have something to do with changes brought about by the annealing. It is known that due to the nanotube formation in fluoride-containing electrolyte, as-prepared oxides show the presence of fluorine. There are two ways in which fluoride ions can be trapped in the oxide matrix. One way results from the mechanism of nanotube formation. Due to faster  $F^-$  ion migration through the oxide [27], very thin  $TiF_4$  fluoride layer is formed at the metal/oxide interface. This layer can be later removed by dissolving in water-containing electrolyte. This process was observed by Fang et al. [33] who studied the nanotube evolution on Zr with time. They demonstrated that the compact layer rich in fluorine is formed between the metal and upper tubular layer after 1 h of anodization, and then, it disappears after 3 h. This experiment shows that the outer layer of



**Fig. 7** OCP vs. time plots obtained in SBF solution during 12-h immersion



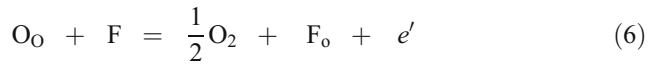
**Fig. 8** Polarization curves recorded for Ti, Ti6Al7Nb, Ti6Al7Nb-NT, and Ti6Al7Nb-NT-HT

**Table 1** Corrosion current and corrosion potentials obtained from polarization test

Sample	$i_{\text{corr}}$ (nA cm <sup>-2</sup> )	$E_{\text{corr}}$ (V vs Ag/AgCl)
Ti	69 ± 4	-0.262 ± 0.030
Ti6Al7Nb	91 ± 14	-0.212 ± 0.021
Ti6Al7Nb-NT	968 ± 58	-0.497 ± 0.049
Ti6Al7Nb-NT-HT	39 ± 4	-0.236 ± 0.020

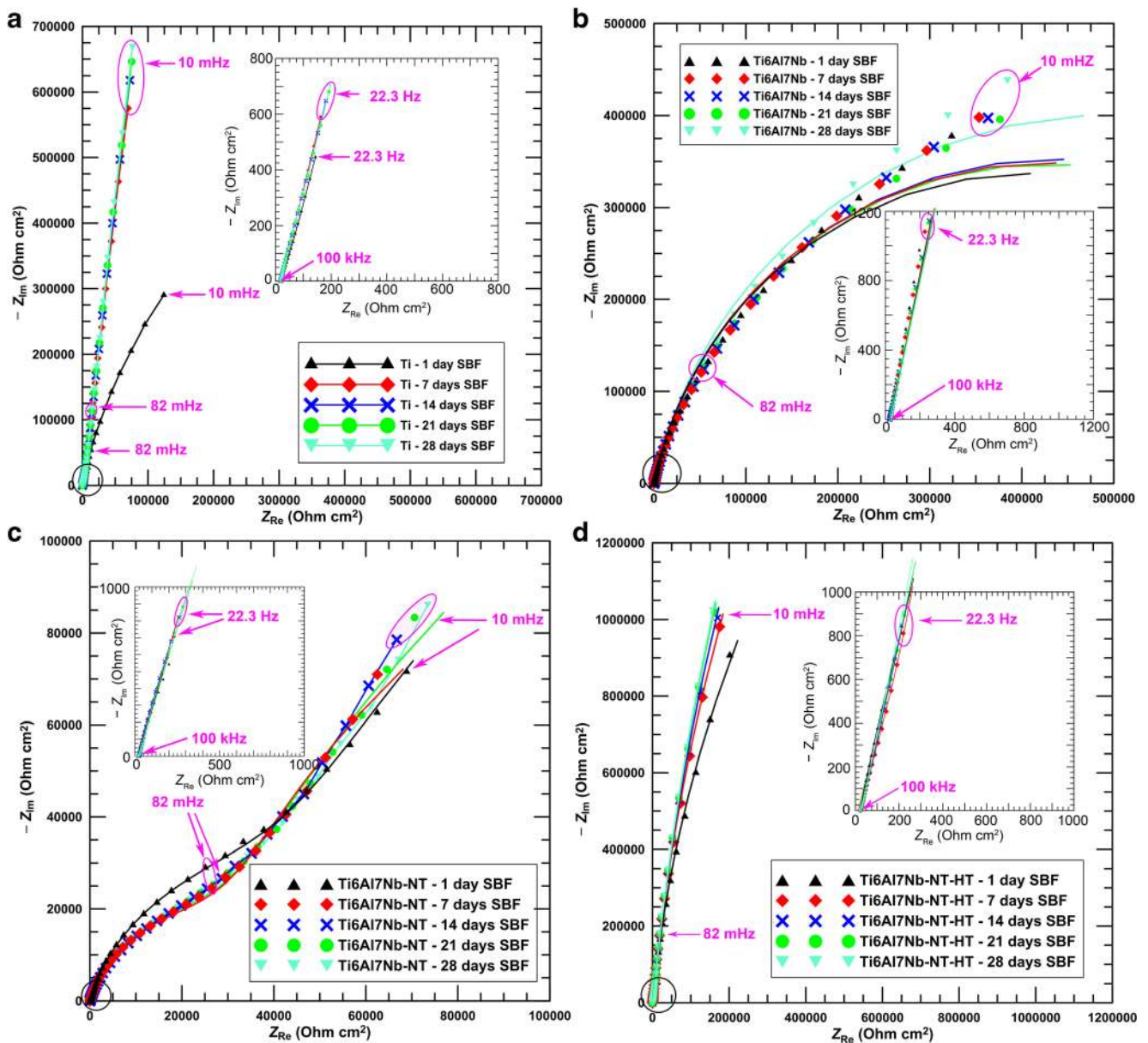
nanotubes is fluoride-rich and it can be dissolved in the electrolyte. During annealing, it can be removed with the moisture as HF [34]. Another way is possible substitution of oxygen atoms with fluorine atoms in the bulk oxide. The presence of F

on oxygen sites generates (in Kroger–Vink notation) excess electrons:



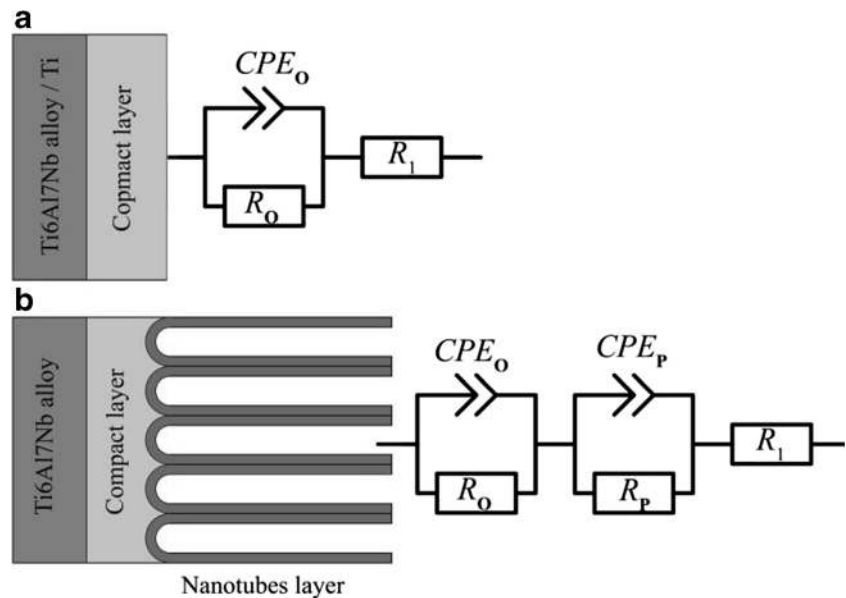
enhancing n-type conductivity. Annealing in the oxygen-containing atmosphere removes both F and electrons decreasing the conductivity. This effect was observed by Vacandio et al. [34] during investigations of electrical properties of zirconia nanotubes before and after annealing.

Independently of the way of the fluoride ion removal, there can be a parallel process responsible for the enhancement of anticorrosion properties of NT-HT alloy. In our case, heat



**Fig. 9** Impedance diagrams representation (Nyquist) of examined surface during soaking in SBF solution for Ti (a), Ti6Al7Nb (b), Ti6Al7Nb-NT (c), and Ti6Al7Nb-NT-HT (d)

**Fig. 10** Equivalent circuits representing the impedance results for compact oxide layer and nanotube oxide layer



treatment process of anodized Ti6Al7Nb sample was carried out at 200 °C in air. This temperature is sufficient to produce an additional oxide barrier at the metal/oxide interface. The overall effect of  $F^-$  removal and oxide formation is reflected in corrosion behavior of the oxide nanotubes. The obtained result of corrosion current measurements for Ti6Al7Nb samples show a decrease from 967 nA/cm<sup>2</sup> to the level below 40 nA/cm<sup>2</sup> after heat treatment. Thus, the alloy covered with oxide nanotubes exhibits the best corrosion resistance after heat treatment.

### EIS measurements

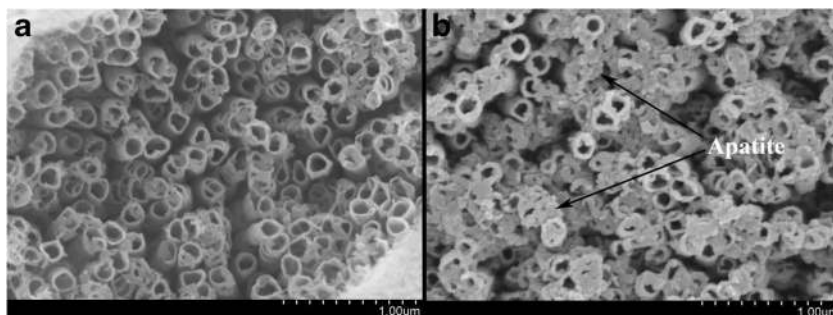
In order to get deeper insight into the passivation phenomena at the electrode/electrolyte interface, and to follow the spontaneous apatite formation, electrochemical impedance spectroscopy (EIS) was used. In our study, impedance measurements were carried out under the conditions which are described in the experimental part. The results of impedance studies obtained after 24 h, 7, 14,

**Table 2** Parameters obtained from impedance data by non-linear least-squares fits

Soaking and measurements in SBF (day)	$R_1$ ( $\Omega$ )	$R_p$ ( $\Omega$ cm <sup>2</sup> )	$\alpha_p$	$Q_p$ ( $\mu$ F cm <sup>-2</sup> s <sup><math>\alpha_p</math></sup> )	$R_o$ ( $\Omega$ cm <sup>2</sup> )	$\alpha_o$	$Q_o$ ( $\mu$ F cm <sup>-2</sup> s <sup><math>\alpha_o</math></sup> )	$\chi^2 \cdot 10^{-3}$
Ti6Al7Nb bare alloy								
1	59.62	–	–	–	822,380	0.8737	13.7	6.72
7	55.24	–	–	–	856,590	0.8694	12.6	6.11
14	50.65	–	–	–	873,660	0.8641	12.2	5.08
21	57.67	–	–	–	854,020	0.8687	12.1	5.03
28	62.36	–	–	–	989,580	0.8665	11.9	3.70
Ti6Al7Nb with nanotubes								
1	47.38	42,616	0.8273	39.4	428,380	0.8381	129.2	3.83
7	45.51	28,739	0.8371	28.0	301,120	0.8360	11.3	2.32
14	45.88	28,539	0.8328	25.3	367,580	0.8181	94.5	2.92
21	44.43	27,097	0.8413	24.5	459,950	0.7794	81.0	3.45
28	47.20	24,936	0.8591	23.8	514,360	0.7506	71.4	4.06
Ti6Al7Nb with nanotubes after heat treatment								
1	44.75	$7.14 \cdot 10^6$	0.9394	13.9	533	0.7969	79.9	0.11
7	45.77	$1.19 \cdot 10^7$	0.9411	13.2	283	0.7576	91.8	0.11
14	46.58	$1.47 \cdot 10^7$	0.9393	13.0	445	0.7159	131.2	0.23
21	45.57	$1.78 \cdot 10^7$	0.9393	12.9	704	0.7189	122.5	0.27
28	45.57	$1.83 \cdot 10^7$	0.9393	12.8	1163	0.6915	148.0	0.22



**Fig. 11** SEM image of Ti6Al7Nb alloy covered with oxide nanotubes after heat treatment before (a) and after (b) exposure for 28 days to SBF



21, and 28 days of soaking in SBF fluid are shown in Fig. 9a–d.

For the interpretation of the experimental data, we used equivalent circuits available in the literature [12, 35–37], which have been often applied to dense and porous oxide layer formed on the metal surface immersed in the electrolyte. Both models are based on Randles circuit and are shown in Fig. 10a, b. In both models,  $R_1$  denotes the resistance of the electrolyte. In turn,  $R_O$  and  $CPE_O$  denote polarization resistance and constant phase element, respectively, and describe the behavior of compact oxide layer, while  $R_p$  and  $CPE_p$  are related to the behavior of porous layer of oxide nanotubes.

Constant phase element described by the equation:

$$Z_{CPE} = \frac{1}{(j\omega)^\alpha \cdot Q} \quad (7)$$

enables to observe the changes of heterogeneity and porosity of the electrode. For  $\alpha = 1$ ,  $Z_{CPE}$  corresponds to the capacitor, for  $\alpha = 0$ , it takes part of the resistor, and for  $\alpha = -1$ , it corresponds to the inductor [36, 38, 39].

In all the cases, the simplest equivalent circuits were chosen in order to achieve lower confidence intervals for the estimated parameters and to capture all qualitative features of the impedance spectra [35].

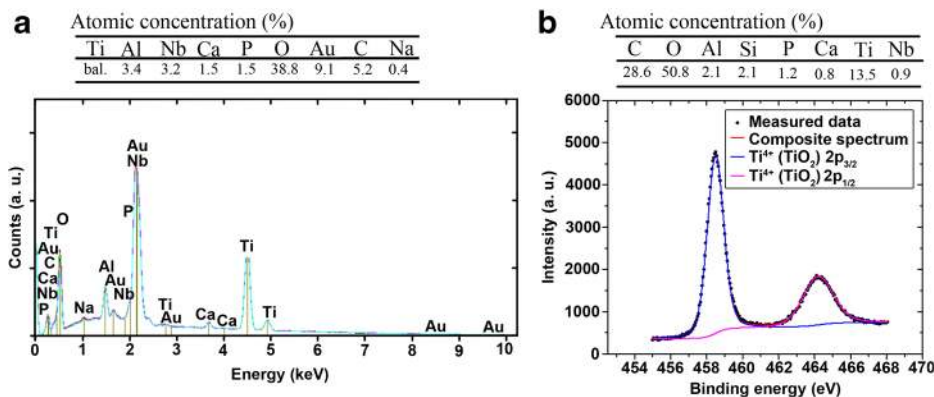
In the case of bare alloy, the results were fitted with model shown in Fig. 10a. It corresponds to metal covered with compact oxide layer. The model is similar to that one used in our

previous study on corrosion resistance of Ti and Ti–Pd alloy in phosphate-buffered saline solution [12, 40]. Experimental data obtained for Ti6Al7Nb “bare” alloy are well described by the chosen circuit. It can be seen that as time flows, resistance is slightly increasing while capacity is insignificantly decreasing. It may suggest that on the electrode surface, apatite layer is not formed, and the observed change can be attributed to the slow growth of compact oxide layer.

In the case of the alloy covered with oxide nanotubes (without and after the heat treatment), for the interpretation of impedance data, the model shown in Fig. 10b was chosen. It incorporates two time constants which can be deduced from Bode plots. In both the cases, there is a decrease of  $\alpha_O$  which means deviation from pure capacitance due to surface inhomogeneity [35]. It may be a result of apatite deposition on the surface of nanotubes. After heat treatment, nanotubes exhibit almost constant  $\alpha_p$  (which is close to that for pure capacitor) during the time of experiment (i.e., up to 28 days) with simultaneous increase of resistance. It may reflect the oxide growth due to annealing in air. Nanotubes without heat treatment show an increase of  $\alpha_p$  and decrease of resistance with time of soaking. It may be a result of residual fluoride dissolution in the electrolyte.

The results interpreted in a way described above are gathered in Table 2. The chi square test indicates that the fit of model to the data is good ( $10^{-3}$ – $10^{-4}$ ), which speaks for the applicability of both models. In order to verify the correctness of EIS measurements, for all the obtained spectra, the

**Fig. 12** EDS and XPS spectra obtained for heat-treated nanotube layer after soaking experiments (28 days)



Kramer–Kronig test was performed. It consists in the transformation of real axis into the imaginary axis and the imaginary axis into the real axis. This procedure was recently used by Feng et al. [41] in the study of passivity of 316L stainless steel in borate buffer solution. Its detailed explanation is given in the book of Orazem and Tribollet [38] and by Macdonald [42]. For all the measurements, the test result ranged from  $10^{-5}$  to  $10^{-6}$ , and it means that EIS measurements were performed very well. It is also clear from EIS results that the best corrosion resistance exhibits Ti6Al7Nb alloy covered with oxide nanotubes after heat treatment. The results are in good agreement with the results of polarization studies.

Soaking experiments performed with the alloy covered with oxide nanotubes after heat treatment took 28 days. The comparison of the oxide surface before and after soaking experiments is shown in Fig. 11a, b. EDS and XPS (Fig. 12a, b) analysis showed after experiments that apatite is formed on the nanotube surface. This oxide nanotube layer was synthesized at the voltage of 30 V.

It is seen that apatite has grown on the top of nanotubes. We have no evidence that its penetration inside nanotubes occurred. The similar result was obtained for Ti6Al7Nb alloy covered with oxide nanotubes without any heat treatment. In the case of pure Ti and bare alloy, we observed that the apatite deposition did not occur. It means that for the deposition of apatite, porous oxide is needed.

## Conclusions

In the present study, the conditions of oxide nanotube formation on Ti6Al7Nb were investigated. It was found that, similarly to ZrO<sub>2</sub> nanotubes, the size of nanotubes is a linear function of applied potential, and in the range of potentials from 10 to 35 V, the inner diameter can vary from 30 to 110 nm.

Corrosion resistance study conducted by polarization and EIS measurements gave similar results. Electrochemical measurements were performed on the samples anodized at constant voltage of 30 V. Polarization tests revealed that an alloy covered with oxide nanotubes subjected to the heat treatment exhibits the best corrosion resistance. EIS results were interpreted with two models based on Randles' circuit. These experiments also showed that an alloy covered with oxide nanotubes after heat treatment exhibits the best corrosion resistance, which polarization studies confirmed. The deposition of apatite on investigated samples occurred only on alloy samples after anodization, with and without heat treatment. Formation of the apatite and corrosion studies indicates that the alloy covered with oxide nanotubes formed under potential of 30 V after heat treatment is well suited for biological integration.

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