

Durham Research Online

Deposited in DRO:

24 October 2012

Version of attached file:

Published Version

Peer-review status of attached file:

Peer-reviewed

Citation for published item:

Clark, S.J. and Robertson, J. (2009) 'Energy levels of oxygen vacancies in BiFeO(3) by screened exchange.', Applied physics letters., 94 (2). 022902.

Further information on publisher's website:

<http://dx.doi.org/10.1063/1.3070532>

Publisher's copyright statement:

© 2009 American Institute of Physics. This article may be downloaded for personal use only. Any other use requires prior permission of the author and the American Institute of Physics. The following article appeared in Clark, S.J. and Robertson, J. (2009) 'Energy levels of oxygen vacancies in BiFeO(3) by screened exchange.', Applied physics letters., 94 (2). 022902 and may be found at <http://dx.doi.org/10.1063/1.3070532>

Additional information:

Use policy

The full-text may be used and/or reproduced, and given to third parties in any format or medium, without prior permission or charge, for personal research or study, educational, or not-for-profit purposes provided that:

- a full bibliographic reference is made to the original source
- a [link](#) is made to the metadata record in DRO
- the full-text is not changed in any way

The full-text must not be sold in any format or medium without the formal permission of the copyright holders.

Please consult the [full DRO policy](#) for further details.

Energy levels of oxygen vacancies in BiFeO₃ by screened exchange

S. J. Clark^{1,a)} and J. Robertson^{2,b)}

¹*Department of Physics, University of Durham, South Road, Durham DH1 3LE, United Kingdom*

²*Department of Engineering, Cambridge University, Cambridge CB2 1PZ, United Kingdom*

(Received 29 October 2008; accepted 20 December 2008; published online 13 January 2009)

The oxygen vacancy in BiFeO₃ is calculated to be a double donor with states 0.6 eV below the conduction band edge, consistent with cathodoluminescence and electronic conductivity data. The atomic configurations were relaxed using the local density approximation plus Hubbard U (LDA+ U) to the electron-correlation energy for each defect charge state to ensure that the oxide had a nonzero band gap. The defect formation energies were calculated using the screened exchange (sX) functional. © 2009 American Institute of Physics. [DOI: 10.1063/1.3070532]

There is increasing interest in materials that display simultaneous ferroelectric and magnetic properties, the multiferroics,^{1–4} of which BiFeO₃ (BFO) is the best known example. BFO has both its Néel temperature and Curie temperature well above room temperature. It has a very large remnant polarization of $\sim 90 \mu\text{C}/\text{cm}^2$ of interest for ferroelectric nonvolatile memories. However, BFO can show rather high electrical leakage currents, which mask the measurement of the ferroelectric polarization curve and could short circuit ferroelectric storage capacitors.^{5–7} It is therefore important to know if the leakage is intrinsic or extrinsic.

We previously calculated the band gap of BFO and its Schottky barrier heights and concluded that the leakage currents were extrinsic in origin.⁸ This is consistent with experiment data, which shows that leakage can be reduced by optimizing growth conditions and aliovalent doping.⁶ The leakage current has been attributed to oxygen vacancies, which is equivalent to a valence change of Fe from 3+ to 2+.⁹ We calculate here the energy levels of the oxygen vacancy and find they are 0.6 eV below the conduction band edge and consistent with them being the cause of leakage.

The antiferromagnetism of BFO arises from the partial filled Fe 3d band, and its band gap arises from the ferroelectric lattice distortion and the on-site correlation energy of the Fe 3d electrons.¹⁰ This causes certain difficulties in accurate band structure calculations. The standard method to calculate band structures uses the local spin density approximation (LSDA) or generalized gradient approximation functional to represent the valence electron's exchange-correlation energy.^{10,11} This functional provides an accurate description of ground state properties such as lattice constants. However, it does not give band gaps correctly due to a discontinuity in the derivative of the exchange-correlation energy versus particle number across the gap. The incorrect band gap is often worse in correlated electron systems (a classic example being NiO, where the simple LDA gives a metal, rather than a 4.2 eV gap semiconductor). There are ways to fix this problem at varying computational cost. GW is one of the most accurate methods but it is very expensive and hence difficult to perform calculations of defect properties in materials.

The calculation of defect levels requires the use of a supercell of 50–200 atoms containing the defect, in which the atomic coordinates are then relaxed to minimize the total

energy. This requires the use of an efficient but accurate functional. The usual method of calculating defect levels while giving a good band gap is to relax the atomic structure in LSDA and then apply a more accurate but expensive method as postprocessing to that LSDA atomic structure.^{12–15} This method is dangerous for BFO, as the LSDA gives a small band gap, so it could give the incorrect defect structure. We need a functional that is efficient enough to treat large supercells and give the right band gap. The simplest of these is LDA+ U . This method supplements the LDA functional by an on-site correlation energy U for the Fe 3d electrons.¹⁶ The LDA+ U method is appropriate for fixing band gaps in open shell systems such as NiO or BFO. LDA+ U was previously used by Neaton *et al.*¹⁰ to calculate the band structure of BFO. Another way is to relax the structure using a local orbital basis set and a hybrid functional such as B3LYP.¹⁷ This is suitable for localized deep levels. However, a more complete basis set of plane waves is more appropriate for near shallow levels. This paper represents a calculation of defect geometries in large supercells using the LDA+ U method.

In the present calculation, we use the plane wave, total energy pseudopotential code CASTEP.¹⁸ The electron-ion interactions are described by nonlocal norm-conserving pseudopotentials. The valence electron wave functions are expanded in a plane-wave basis set to a kinetic energy cutoff of 800 eV, which converges total energy differences to better than 1 meV/atom. Integrations over the Brillouin zone used the k -point sampling method of Monkhorst and Pack with a grid that converges the energies to a similar accuracy. The U is an empirical parameter in the LDA+ U scheme and is fitted to the band gap and average valence band width of our screened exchange band structure⁸ of bulk BFO, we find $U = 4.7$ eV. Geometry optimizations are performed self-consistently using a minimization scheme and the Hellmann–Feynman forces. Geometries are converged when forces are below 0.01 eV/Å.

An oxygen vacancy is introduced into a 120 atom supercell of BFO. The lattice parameters of this $3 \times 2 \sqrt{2} \times 2 \sqrt{2}$ cell are fixed to that of the calculated bulk cell. The internal atomic positions are then relaxed separately in LDA+ U for the three possible charge states $q=0, +1,$ and $+2$ of the vacancy (other charge states do not occur for the defect). For the supercell we use a single k -point (1/4,1/4,1/4) for Brillouin zone integrations, which converges the quantities faster than the Γ point with respect to the supercell size.¹⁹ The total energy (E_q) is calculated for the defect cell of charge q , the

^{a)}Electronic mail: s.j.clark@durham.ac.uk.

^{b)}Electronic mail: jr@eng.cam.ac.uk.

perfect cell (E_H) of charge q , and a perfect cell of charge 0. This allows us to calculate the defect formation energy, H_q , as a function of the relative Fermi energy (ΔE_F) from the valence band edge E_V and the relative chemical potential ($\Delta\mu$) of element α ,²⁰

$$H_q(E_F, \mu) = [E_q - E_H] + q(E_V + \Delta E_F) + \sum_{\alpha} n_{\alpha}(\mu_{\alpha}^0 + \Delta\mu_{\alpha}),$$

where qE_V is the change in energy of the Fermi level when charge q is added. Essentially, this is the shift in the average electrostatic potential due to the change in charge of the system with respect to the uncharged system. The oxygen chemical potential (μ^0) is referred to that of the O_2 molecule, taken as zero, and n_{α} is the number of α atoms.

The total energy of the three cells is then also calculated by screened exchange on the structures found by LDA+ U . These values are used, as these have more physical basis than those from LDA+ U . The screened exchange method^{8,21,22} is a generalized Kohn–Sham functional, where the single particle orbitals are used to construct the nonlocal exchange operator while introducing an exponential screening term. This improves the physical description of electron exchange and correlation beyond LDA and gives accurate total energies and correct band gaps at moderate cost.

Antiferromagnetic (AF) BFO has the $R3c$ structure, a distortion of the cubic perovskite lattice, in which the FeO_6 octahedra are tilted and the Fe and Bi ions move toward each other along $\langle 111 \rangle$ so that the O site becomes fourfold coordinated by two Fe and two Bi. Figure 1(a) shows the standard LSDA band structure of the AF $R3c$ BFO. We see that there is a small band gap of 0.43 eV in LSDA. Figure 1(b) shows the band structure of AF $R3c$ BFO calculated with by screened exchange. The band gap is calculated to be 2.8 eV. This agrees well with subsequent measurements of the optical gap of 2.6–2.8 eV.^{23,24} For reference, we show in Fig. 1(c) the band structure of the cubic BFO in the screened exchange method. This verifies that there is no band gap in the cubic polytype even with a method that is capable of giving it. Note that the cubic phase corresponds to the metallic phase well above the Néel temperature, where we assume that there is still some local order, so that spins along $\langle 111 \rangle$ still tend to lie antiparallel, so we show the band structure for a cell of two cubic cells along $\langle 111 \rangle$.

Figure 2 shows the O vacancy formation energy from screened exchange as a function of the Fermi energy for oxygen-rich conditions, where the O chemical potential is that of the O_2 molecule. Figure 2(b) shows the formation energy for oxygen-poor conditions of $\mu(O) = -1.97$ eV corresponding to the Bi_2O_3/Bi equilibrium. A formation energy of 2.6 eV for the neutral vacancy in O-poor conditions is rather low compared to $SrTiO_3$, and it accounts for the ease of forming vacancies at high temperatures.

We see that the oxygen vacancy gives rise to $+2+$ and $0/+$ transitions at 0.6 eV below the conduction band edge. The energy level is borderline deep. The energy level is the same for the $0/+$ and the $+2+$ transitions. This equality does not mean that the level is effective masslike, as the $+2+$ level would then be twice as deep. It means that the level shift due to a slight lattice relaxation balances the extra attraction in the $2+$ state. The ions around the vacancy relax

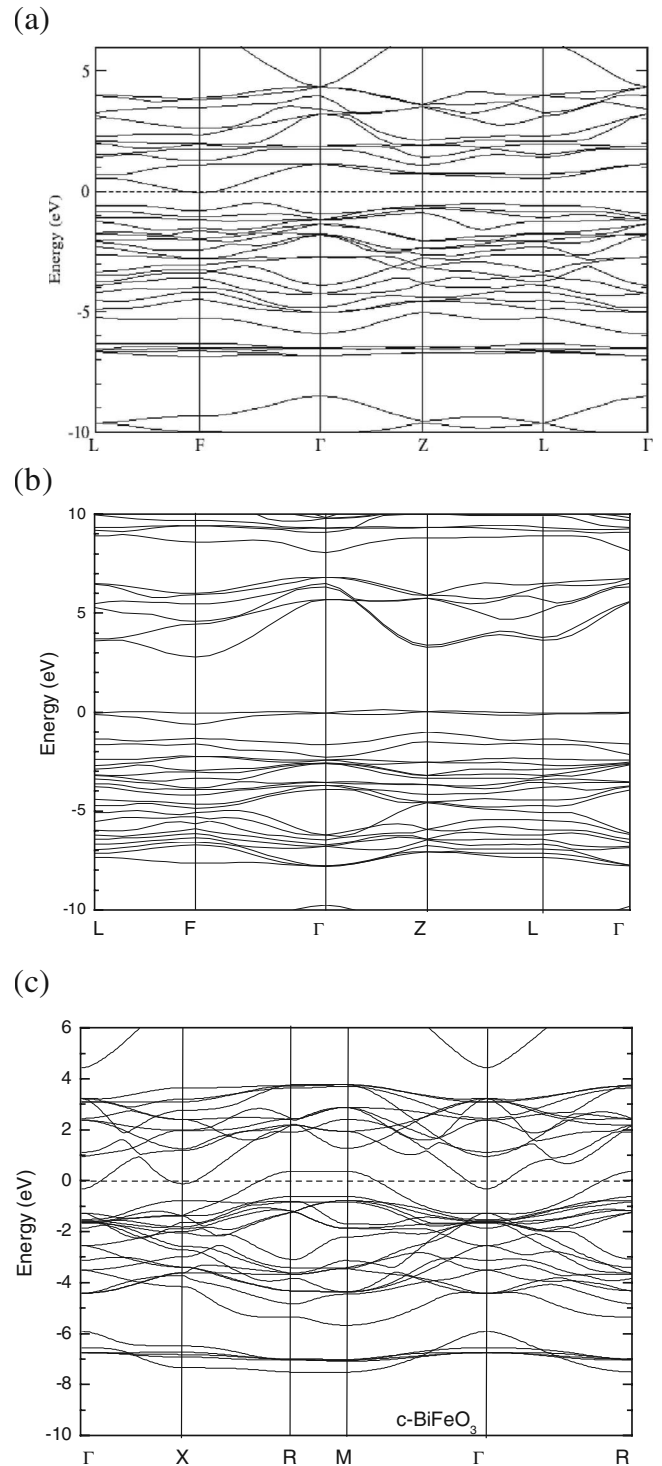


FIG. 1. Band structure of AF BFO in $R3c$ structure (a) in LSDA, (b) in screened exchange, and in the (c) band structure of cubic BFO in screened exchange, for reference.

outward by 0.092 Å for the $2+$ state. The vacancy state eigenvalues lie ~ 0.4 eV below the conduction band, Fig. 2(c). These are consistent with the formation energies.

Figure 3(a) shows the charge density for the defect state in the Fe–O–Fe plane. The white ball represents the oxygen vacancy site. The defect state is seen to be mainly localized on the d states of the two Fe atoms adjacent to the vacancy, rather than inside the vacancy itself, and rather delocalized onto adjacent cells with a Bohr radius of ~ 4.3 Å. Figure 3(b) shows the charge density in the plane of the vacancy and its Bi neighbors. We see that the defect state is not lo-

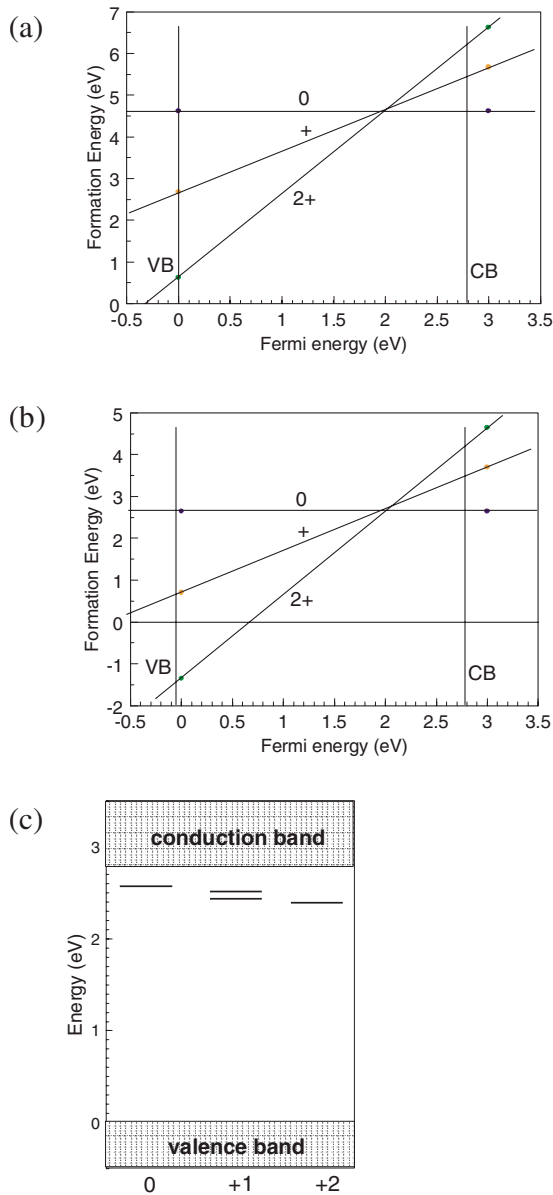


FIG. 2. (Color online) Formation energy of the O vacancy vs Fermi energy for (a) O-rich conditions ($\mu_{\text{O}}=0$ eV) and (b) O-poor conditions ($\mu_{\text{O}}=-1.97$ eV). (c) Eigenvalues of vacancy levels for each charge state 0, +1, and +2.

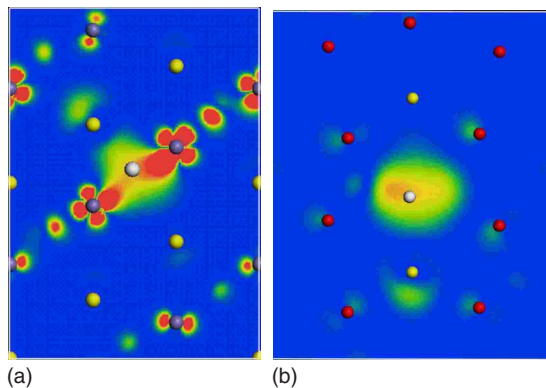


FIG. 3. (Color online) Charge density of the O vacancy defect level (a) through a plane containing Fe atoms and (b) a plane containing the vacancy and Bi atoms. White sphere is the vacancy, Fe=gray, Bi=yellow, O=red.

calized on Bi states. This is consistent with the bulk band structure where the conduction band minimum is formed mainly of Fe states and Bi states lie higher.

The presence of a moderately shallow oxygen vacancy state is consistent with experimental observations.^{24–28} Hauser *et al.*²⁴ found a transition in cathodoluminescence at ~ 0.3 eV below the conduction band edge, whose properties are consistent with the oxygen vacancy. A calculated level depth of 0.6 eV is similar to a trap depth of 0.65–0.8 eV found by Pabst *et al.*²⁶ and 0.68 eV from the temperature dependence of the conductivity.²⁷ The ability to compensate self-doping and oxygen deficiency by aliovalent metals is consistent with a moderately shallow defect.^{6,25–28}

In summary, we calculated the formation energy of oxygen vacancy in BFO and found it to create a marginally deep level 0.6 eV below the conduction band edge. We used the LDA+ U method to give the atomic geometry and screened exchange to give the formation energies.

- ¹J. Wang, J. B. Neaton, H. Zheng, V. Nagarajan, S. B. Ogale, B. Liu, D. Viehland, V. Vaithyanathan, D. G. Schlom, U. V. Waghare, N. A. Spaldin, K. M. Rabe, M. Wuttig, and R. Ramesh, *Science* **299**, 1719 (2003).
- ²R. Ramesh and N. Spaldin, *Nature Mater.* **6**, 21 (2007).
- ³N. A. Hill, *J. Phys. Chem. B* **104**, 6694 (2000).
- ⁴R. Palai, R. S. Katiyar, H. Schmid, P. Tissot, S. J. Clark, J. Robertson, S. A. T. Redfern, G. Catalan, and J. F. Scott, *Phys. Rev. B* **77**, 014110 (2008).
- ⁵K. Y. Yun, D. Ricinchi, T. Kanashima, and M. Okuyama, *Appl. Phys. Lett.* **89**, 192902 (2006).
- ⁶S. K. Singh and H. Ishiwara, *Jpn. J. Appl. Phys., Part 2* **44**, L734 (2005); S. K. Singh, H. Ishiwara, and K. Maruyama, *Appl. Phys. Lett.* **88**, 262908 (2006).
- ⁷D. Lebeugle, D. Colson, A. Forget, and M. Viret, *Appl. Phys. Lett.* **91**, 022907 (2007).
- ⁸S. J. Clark and J. Robertson, *Appl. Phys. Lett.* **90**, 132903 (2007).
- ⁹C. Ederer and N. A. Spaldin, *Phys. Rev. B* **71**, 224103 (2005).
- ¹⁰J. B. Neaton, C. Ederer, U. V. Waghmare, N. A. Spaldin, and K. M. Rabe, *Phys. Rev. B* **71**, 014113 (2005).
- ¹¹P. Ravindran, R. Vidya, A. Kjekshus, H. Fjellvag, and O. Eriksson, *Phys. Rev. B* **74**, 224412 (2006).
- ¹²K. Xiong, J. Robertson, M. C. Gibson, and S. J. Clark, *Appl. Phys. Lett.* **87**, 183505 (2005).
- ¹³P. Broqvist and A. Pasquarello, *Appl. Phys. Lett.* **89**, 262904 (2006).
- ¹⁴K. Xiong, J. Robertson, and S. J. Clark, *J. Appl. Phys.* **102**, 083710 (2007).
- ¹⁵F. Oba, A. Togo, I. Tanaka, J. Paier, and G. Kresse, *Phys. Rev. B* **77**, 245202 (2008).
- ¹⁶V. I. Anisimov, J. Zaanen, and O. K. Andersen, *Phys. Rev. B* **44**, 943 (1991).
- ¹⁷C. Di Valentin, G. Pacchioni, and A. Selloni, *Phys. Rev. Lett.* **97**, 166803 (2006).
- ¹⁸M. D. Segall, P. J. D. Lindan, M. J. Probert, C. J. Pickard, P. J. Hasnip, S. J. Clark, and M. C. Payne, *J. Phys.: Condens. Matter* **14**, 2717 (2002).
- ¹⁹M. I. J. Probert and M. C. Payne, *Phys. Rev. B* **67**, 075204 (2003).
- ²⁰S. Lany and A. Zunger, *Phys. Rev. B* **78**, 235104 (2008).
- ²¹B. M. Bylander and L. Kleinman, *Phys. Rev. B* **41**, 7868 (1990).
- ²²J. Robertson, K. Xiong, and S. J. Clark, *Phys. Status Solidi B* **243**, 2054 (2006).
- ²³S. R. Basu, L. W. Martin, Y. H. Chu, M. Gajek, R. Ramesh, R. C. Rai, X. Xu, J. L. Musfeldt, *Appl. Phys. Lett.* **92**, 091905 (2008).
- ²⁴A. J. Hauser, J. Zhang, L. Meier, R. A. Ricciardo, P. M. Woodward, T. L. Gustafson, L. J. Brillson, and F. Y. Yang, *Appl. Phys. Lett.* **92**, 222901 (2008).
- ²⁵C. Wang, T. Hidetoshi, H. Fujino, X. Zhao, E. Kume, T. Horiuchi, and S. Sakai, *J. Appl. Phys.* **99**, 054104 (2006).
- ²⁶G. W. Pabst, L. W. Martin, Y. H. Chu, and R. Ramesh, *Appl. Phys. Lett.* **90**, 072902 (2007).
- ²⁷H. Yang, M. Jain, N. Suvorova, H. Zhou, H. M. Luo, D. M. Feldman, P. C. Dowden, R. F. DePaula, S. R. Foltyn, and Q. X. Jia, *Appl. Phys. Lett.* **91**, 072911 (2007).
- ²⁸H. W. Jang, S. H. Baek, C. B. Eom, and R. Ramesh, *Appl. Phys. Lett.* **92**, 062910 (2008).