



SPECIAL ISSUE: Optical Gain Materials towards Enhanced Light-Matter Interactions

# Enhanced nonlinear optical functionality in birefringence and refractive index dispersion of the deep-ultraviolet fluorooxoborates

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**ABSTRACT** As a promising candidate, the fluorooxoborate has enkindled new explorations of nonlinear optical materials to meet the deep-ultraviolet criteria. However, big challenges and open questions still remain facing this exciting new field, especially the birefringence and dispersion of refractive index which are fundamental parameters for determining the phase-matching second harmonic generation wavelength. Here we designed possible anionic groups in fluorooxoborates, and analyzed the optical anisotropy to check its influence on birefringence, which was proved further by the response electronic distribution anisotropy approximation. The functional modules modulating birefringence in fluorooxoborates were explored systematically. We developed an approach for evaluating the behavior of the refractive index dispersions and found that the fluorooxoborates had small refractive index dispersions owing to the introduction of fluorooxoborate modules. Our results demonstrate that fluorooxoborates can be utilized to realize short phase-matching wavelength markedly and offer a path toward novel performance-driven materials design.

**Keywords:** nonlinear optical crystal, deep-ultraviolet, borate, birefringence, phase-matching

## INTRODUCTION

Nonlinear optical (NLO) materials are the vital components of future photoelectric technologies as they can achieve the tunable laser output by the frequency-conversion technology [1–9]. As for deep-ultraviolet (DUV, wavelength  $\lambda < 200$  nm) region, the NLO crystals, the unique materials capable of generating DUV coherent

light through frequency conversion, are of urgent demands in ultrahigh resolution photolithography, photochemical synthesis, and high-precision micro processing [10–13]. Till now, only  $\text{KBe}_2\text{BO}_3\text{F}_2$  (KBBF) can certainly generate DUV lasers by direct second harmonic generation (SHG), but its application is limited because of the adverse layer growth habit and toxicity issue of containing beryllium [10]. Therefore, the exploration of new DUV NLO materials is in great demand. However, the primary challenge is to balance the three crucial and correlated requirements in one NLO material [12,13]: (i) wide DUV transparency window (a cutoff edge far below 200 nm, or band gap  $> 6.2$  eV); (ii) relatively large second-order nonlinear coefficient ( $d_{ij} > 1 \times$  commercial  $\text{KH}_2\text{PO}_4$ , KDP); (iii) sufficient birefringence to satisfy the phase-matching (PM) condition in DUV wavelength range.

A promising DUV NLO material should achieve a subtle balance of the correlated criteria. In view of the structure-property relationship, fundamental building units (FBUs) form the ‘backbone’ of inorganic materials, which, combining their microscopic functionality and arrangement information, greatly influence or even determine their performances [14–20]. For NLO systems, a superior FBU with optimized spatial arrangement, namely a functional module, featured by large HOMO-LUMO (the highest occupied molecular orbital and the lowest unoccupied molecular orbital) gap, apparent polarizability anisotropy ( $\delta$ ) and high hyperpolarizability, will bring benefits to crystals for band gaps, SHG responses, birefringences. In borate systems, a boron atom usually forms a triangle  $[\text{BO}_3]$  with B- $\text{sp}^2$  hybridization or

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a tetrahedron  $[\text{BO}_4]$  with  $\text{B-sp}^3$  hybridization by linking either three oxygen atoms or four oxygen atoms. As the FBUs in borates, the  $[\text{BO}_3]$  and  $[\text{BO}_4]$  units may further form rings, chains, layers and three-dimensional (3D) networks by sharing oxygen atoms [21,22]. Because of  $\pi$  electron, the planar  $[\text{BO}_3]$  unit can be easily polarized, which may render borates a large birefringence or SHG effect. However, considering that the HOMO of the isolated  $[\text{BO}_3]$  unit is controlled by the lone pair electrons, one can extend the boron-oxygen framework to increase the polymerization of  $[\text{BO}_3]$  to remove the lone pair electrons [23]. Introducing fluorine into borates as a terminal atom is also proved to be one effective way to enlarge the band gap. By introducing fluorine into borates, one can obtain two kinds of compounds: borate fluorides with the fluorine connecting the metal cations and fluorooxoborates with the fluorine also connecting the boron atom. The most typical borate fluorides are  $\text{KBBF}$  [10,24],  $\text{RbBe}_2\text{BO}_3\text{F}_2$  (RBBF) [25],  $\text{BaBe}_2\text{BO}_3\text{F}_3$  [26],  $\text{BaAlBO}_3\text{F}$  [27],  $\text{Ba}_4\text{B}_{11}\text{O}_2\text{F}$  [28],  $\text{Rb}_3\text{Al}_3\text{B}_3\text{O}_{10}\text{F}$  [29],  $\text{K}_3\text{Sr}_3\text{Li}_2\text{Al}_4\text{B}_6\text{O}_{20}\text{F}$  [30],  $\text{Ca}_5(\text{BO}_3)_3\text{F}$  [31],  $\text{NaSr}_3\text{Be}_3\text{B}_3\text{O}_9\text{F}_4$  [32],  $\text{NH}_4\text{Be}_2\text{BO}_3\text{F}_2$  [33],  $\gamma\text{-Be}_2\text{BO}_3\text{F}$  [33], and most of them have low UV cutoff edges below 200 nm. However, the KBBF family with the capability to directly generate DUV coherent light by an SHG process can hardly grow large single crystal due to the layer habit, which hinders the further application. Fluorooxoborates with the  $[\text{BO}_x\text{F}_{4-x}]$  ( $x = 1, 2, 3$ ) groups (represented by [BOF] for simplification), on the other hand, have recently been proved as a new class of promising NLO or optical crystals [34–37]. Among them,  $\text{AB}_4\text{O}_6\text{F}$  ( $\text{A} = \text{NH}_4, \text{Cs}$ ) has the ability to generate DUV coherent light [34,35]. Therefore, the [BOF] groups exhibit promising advantages in the HOMO-LUMO gap and hyperpolarizability, and can further optimize the performance of crystals when combining high-polymerized  $[\text{BO}_3]$ . As we know, birefringence reflects optical anisotropy induced by the structurally anisotropic confinement [38], which is also one vital factor to determine the PM wavelength [39–42]. However, how the fluorine influences the birefringence and the corresponding chromatic dispersion is still unknown. In this work, we analyzed the optical anisotropy distribution of fluorine-containing anionic groups with the aid of molecular design. And the origin of birefringence for the fluorooxoborates was clarified by the response electron distribution anisotropy (REDA) method [43]. Furthermore, the behavior of chromatic dispersion in refractive index was studied. The role of chromatic dispersion in influencing PM wavelength was explored, which can prove the superiority of introducing

fluorine to the B–O framework.

## METHODS

Polarizability anisotropy  $\delta$  was calculated by the density functional theory (DFT) implemented by Gaussian09 package at 6-31G level. We kept the default values of the Gaussian09 code [44] on the aspect of the other calculation parameters and convergence criteria.

The REDA method we proposed previously [43] was used to analyze the contribution of groups to the birefringence. The birefringence is proportional to the REDA index  $\xi$ :

$$\Delta n \sim \sum_i [N_c Z_a \Delta \rho_b]_i / (n_1 E_g), \quad (1)$$

where  $\xi = \sum_i [N_c Z_a \Delta \rho_b]_i / (n_1 E_g)$ ,  $N_c$  is the coordination number of the nearest neighbor cations to the central anion,  $Z_a$  is the formal chemical valence of the anion,  $\Delta \rho_b$  is the difference of covalent electron density of the covalent bond  $i$  on the optical principal axes of a crystal,  $n_1$  is the minimal refractive index,  $E_g$  is the optical band gap.

Linear optical properties were calculated by the first-principles method based on DFT with generalized gradient approximation (GGA) performed by a plane-wave pseudopotential calculation package CASTEP [45]. We kept the default values of the CASTEP code on the aspect of the calculation parameters and convergence criteria. The linear optical properties were examined *via* calculating the dielectric function  $\varepsilon(\omega) = \varepsilon_1(\omega) + i\varepsilon_2(\omega)$ , where  $\varepsilon_1(\omega)$  and  $\varepsilon_2(\omega)$  are the real and imaginary part of dielectric function, respectively. The  $\varepsilon_2(\omega)$  can be obtained by

$$\varepsilon_2(\omega) = \frac{4\pi^2 e^2}{\Omega} \lim_{q \rightarrow 0} \frac{1}{q^2} \times \sum_{c,v,k} 2\omega_k \delta(E_c - E_v - \omega) \left| \langle c | e \cdot q | v \rangle \right|^2, \quad (2)$$

where  $|\langle c | e \cdot q | v \rangle|$  is the integrated optical transitions from the valance states ( $v$ ) to the conduction states ( $c$ ), and the  $e, q$  denote the polariozation direction of the photon and electron momentum operator. The real part of the dielectric function  $\varepsilon_1(\omega)$  can be calculated from  $\varepsilon_2(\omega)$  based on the Kramers-Kroning transform. And then the refractive indices  $n(\omega)$  and the birefringence ( $\Delta n$ ) can be calculated accordingly.

## RESULTS AND DISCUSSION

### Functionality of fluorine-containing anionic groups

To date, there are 22 alkali-metal fluorooxoborates and

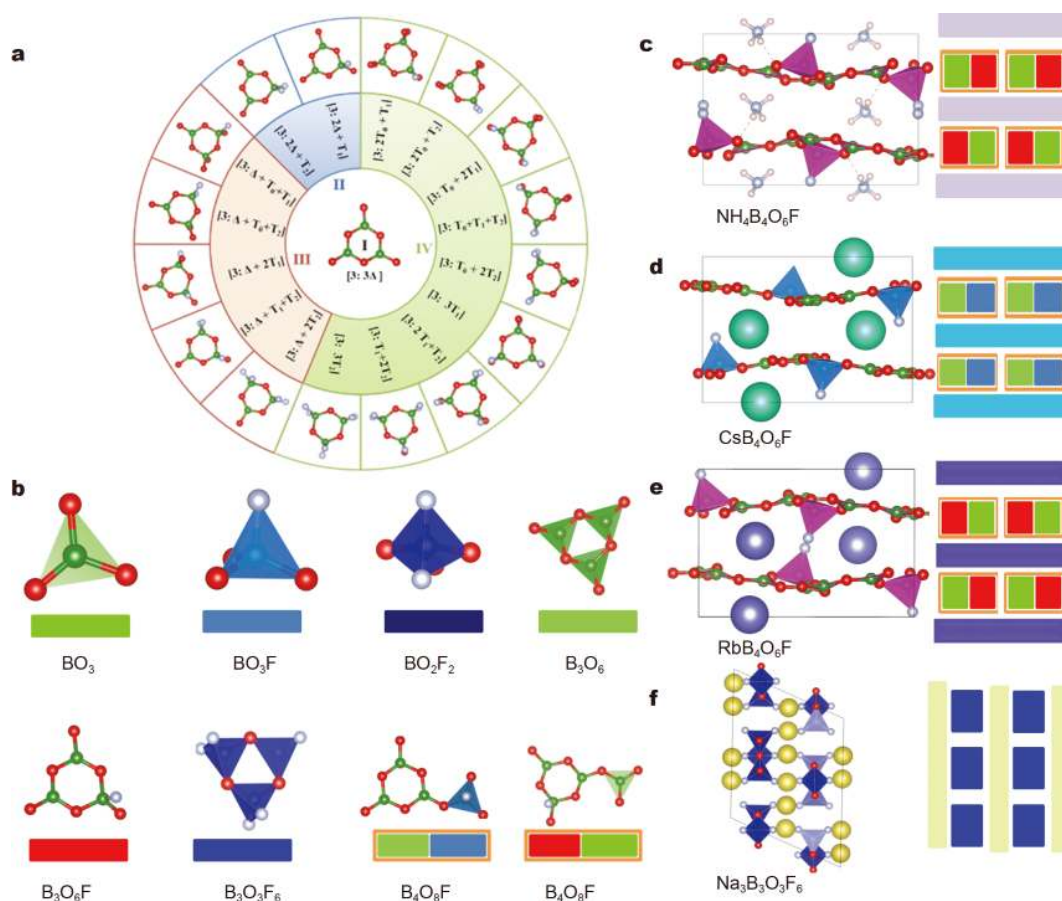
one ammonia-fluorooxoborate synthesized mostly by the standard solid-state reaction in sealed silica tubes. And 11 of them take the formula of  $A_n(B_2O_3)_mF_n$  ( $m, n = 1, 2, 3, \dots$ ):  $LiB_6O_9F$  [46,47],  $A_2B_6O_9F_2$  ( $A = Li, Na$ ) [48,49],  $NaRbB_6O_9F_2$  [50],  $K_3B_6O_9F_3$  [51],  $AB_4O_6F$  ( $A = Na, NH_4, Rb, Cs$ ) [34,35,52,53],  $CsKB_8O_{12}F_2$  [53],  $CsRbB_8O_{12}F_2$  [53] with the symmetry of  $Pna2_1, Cc, P2_1/c, P2_1/n, P2_1/c, Pna2_1, Pna2_1, Pna2_1, P321, P\bar{6}2c$ . And others are  $Li_2B_3O_4F_3$  [54],  $Li_2Na_{0.9}K_{0.1}B_5O_8F_2$  [55],  $A_3B_3O_3F_6$  ( $A = Na, K$ ) [56,57],  $K_{0.42}Rb_{2.58}B_3O_3F_6$  [58],  $Na_{0.76}Rb_{2.24}B_3O_3F_6, K_2-RbB_3O_3F_6, K_{1.66}Rb_{1.34}B_3O_3F_6, Rb_3B_3O_3F_6, KCs_2B_3O_3F_6, K_{2.64}Cs_{0.36}B_3O_3F_6,$  and  $Cs_{1.29}Rb_{1.71}B_3O_3F_6$  [59], with the symmetry of  $P2_12_12_1, Pbcn, C2/c, P2_1/n, Pbcn, P2_1/c, Pbcn, P2_1/c, P2_1/n, P2_1/c, P2_1/c, P2_1/c, P2_1/c$ . We can see that the symmetry of such fluorooxoborates is various including monoclinic, orthorhombic, trigonal and hexagonal systems. In the known alkali-metal fluorooxoborate family, the FBUs have marked diversity consisting of the combination of  $[BO_3]$ , or  $[BO_4]$  with  $[BO_3F]$ , or  $[BO_2F]$ , and Table S1 in the Supplementary Information shows the structural information of typical fluorooxoborates. And to date, only one reported structure consists of “[BOF<sub>3</sub>]” [60], which is hard to be inserted in a six-membered ring (6MR) owing to the terminal fluorine. According to the classification for the FBUs [61,62], the FBUs can be described as 3 : [(3 : T<sub>2</sub>)] in  $Na_3B_3O_3F_6$  and  $K_3B_3O_3F_6$ , 4 : [(3 : 2Δ + T<sub>1</sub>) + (1 : Δ)] in  $NH_4B_4O_6F$  and  $Rb_4B_4O_6F$ , 4 : [(3 : 3Δ) + (1 : T<sub>1</sub>)] in  $CsB_4O_6F$ , and 6 : [(3 : 3Δ) + (3 : 2Δ + T<sub>1</sub>)] in  $LiB_6O_9F$ , where Δ, T, T<sub>0</sub>, T<sub>1</sub>, and T<sub>2</sub> refer to  $[BO_3]$ , a general tetrahedra,  $[BO_4]$ ,  $[BO_3F]$ , and  $[BO_2F_2]$ , respectively [23]. It is found that the FBUs in the known alkali-metal fluorooxoborate family contain B–O or B–O–F 6MR, therefore it is a good maternal microscopic-structure to explore its functionality. As we know, fluorine is the most electronegative and reactive element, so introducing fluorine into the B–O groups may provide the advantage of blue-shift of the UV cutoff which is crucial for DUV NLO materials. If we introduce [BOF] into borates, there are three categories of 6MRs containing [BOF]: [3 : 2Δ + T]-type (II), [3 : Δ + 2T]-type (III), and [3 : 3T]-type (IV). Fig. 1a shows a classification of the configurations of 6MRs without considering the direction of fluorine, with [3 : 3Δ] as type I for comparison. To date, the type II-[3 : 2Δ + T<sub>2</sub>], type III-[3 : Δ + 2T<sub>2</sub>] and type III-[3 : Δ + T<sub>0</sub> + T<sub>2</sub>] configurations have not been found in the reported fluorooxoborates. And in type-IV, only [3 : T<sub>2</sub>] was found in the reported fluorooxoborates. And it is noticed that [BOF<sub>3</sub>] is hard to be embedded into the rings because of the terminal character of fluorine. According to the relationship between local

atom groups and NLO properties, NLO functional materials can be characterized by functional modules and fillers [63]. According to the module description, as shown in Fig. 1, in  $NH_4B_4O_6F$ , the  $[B_3O_6F]$  unit with the configuration of [3 : Δ + 2T<sub>1</sub>] and the  $[BO_3]$  compose  $[B_4O_8F]$  and further form the B–O–F layers, and  $RbB_4O_6F$  has a similar module structure. While for  $CsB_4O_6F$ , the  $[B_4O_8F]$  unit consists of  $[B_3O_6]$  and  $[BO_3F]$  and also forms the B–O–F layers. The module structures in  $NH_4B_4O_6F$  and  $CsB_4O_6F$  can increase the degree of polymerization, which guarantees the short UV cutoff edges (156 and 155 nm). In  $Na_3B_3O_3F_6$ , three  $[BO_2F_2]$  units form the  $[B_3O_3F_6]$  6MR with the configuration of [3 : 3T<sub>2</sub>]. Comparatively, the B–O–F layers with concurrent-parallel  $[BO_3]$  empower the large SHG effect and strong optical anisotropy.

#### Polarizability characters and microscopic contribution

To reveal the influence of the possible [BOF] 6MR on birefringence, we investigated their electronic structure and polarizability characteristics using the DFT method implemented by Gaussian09 package at 6-31G level. The  $\delta$  of the designed 6MRs was investigated, which can reflect the macroscopic birefringence to a certain extent. Fig. 2a shows the  $\delta$  along with the HOMO-LUMO gaps. Considering the DUV transparency criterion, the region with the HOMO-LUMO gap larger than 6.2 eV is screened out. And the  $\delta$  of 6MRs in the blue region, larger than  $2 \times \delta(BO_3)$ ,  $\delta(B_2O_5)$ ,  $3 \times \delta(BO_3)$  is highlighted, where the  $[BO_3]$  and its derived structures are birefringence-preferential structures. In particular, the  $[B_2O_5]$  group can be as the benchmark owing to its dominant role in the typical birefringent material  $Li_2Na_2B_2O_5$  with a large birefringence of 0.095 measured *via* the prism coupling method [64]. It is indicated that the 6MRs with  $\delta(6MRs)$  larger than  $\delta(B_2O_5)$  have the capability to produce large birefringences and possess large HOMO-LUMO gaps as screened out in Fig. 2a.

A hierarchical feature in polarizability anisotropy from I to IV types is shown in Fig. 2b. It can be seen that the groups only containing tetrahedra have the smallest  $\delta$ . The  $\delta$  of [3 : 3T<sub>2</sub>]-FBU, for example, only has 1/3 (or even smaller) times as [3 : 3Δ], which may indicate that materials only with [3 : 3T<sub>2</sub>]-configuration may have a small birefringence [65]. But we still expect that new materials only with tetrahedral T<sub>2</sub>-configuration possess a large birefringence. It is feasible because tetrahedral microscopic structures with some optimizing arrangements [63] can lead to large birefringence of materials. When combining Δ and T, namely, introducing [BOF] forms

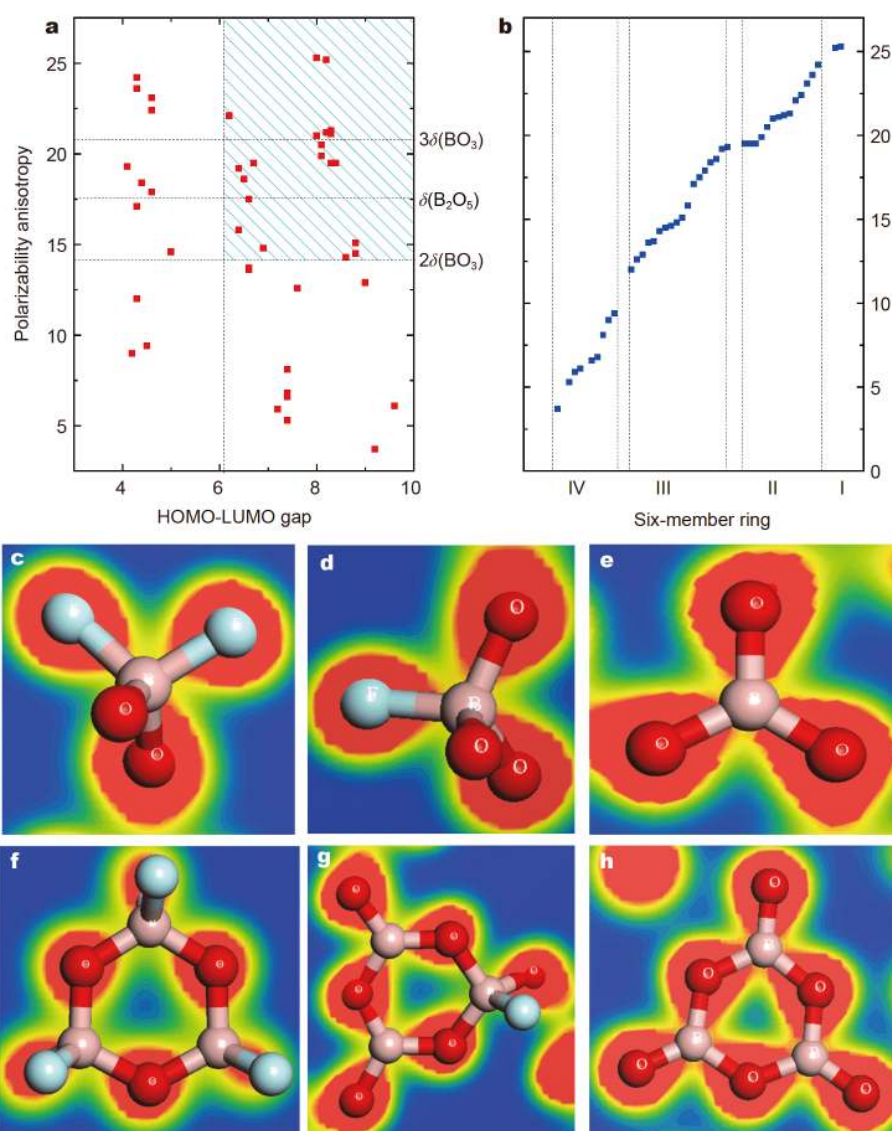


**Figure 1** Fluorine-containing 6MRs, functional groups and assembling optical materials with functional groups. (a) Designed 6MR containing [BOF] groups based on the maternal FBU [3 : 3 $\Delta$ ] (labelled as I configuration), II refers to [3 : 2 $\Delta$  + T], III refers to [3 :  $\Delta$  + 2T] and IV refers to [3 : 3T]. In the [3 : 2 $\Delta$  + T]-type, only two microstructure configurations exist: [3 : 2 $\Delta$  + T<sub>1</sub>], and [3 : 2 $\Delta$  + T<sub>2</sub>]; in the [3 :  $\Delta$  + 2T]-type, there are five kinds of rings: [3 :  $\Delta$  + 2T<sub>1</sub>], [3 :  $\Delta$  + 2T<sub>2</sub>], [3 :  $\Delta$  + T<sub>1</sub> + T<sub>2</sub>], [3 :  $\Delta$  + T<sub>0</sub> + T<sub>1</sub>], [3 :  $\Delta$  + T<sub>0</sub> + T<sub>2</sub>]; and in the [3 : 3T]-type, there are nine possible rings: [3 : 3T<sub>1</sub>], [3 : 3T<sub>2</sub>], [3 : 2T<sub>1</sub> + T<sub>2</sub>], [3 : T<sub>1</sub> + 2T<sub>1</sub>], [3 : 2T<sub>0</sub> + T<sub>1</sub>], [3 : T<sub>0</sub> + 2T<sub>1</sub>], [3 : 2T<sub>0</sub> + T<sub>2</sub>], [3 : T<sub>0</sub> + 2T<sub>2</sub>], and [3 : T<sub>0</sub> + T<sub>1</sub> + T<sub>2</sub>]. (b) Typical functional groups in borates, fluorooxoborates used in DUV region. (c) A modular description of NH<sub>4</sub>B<sub>4</sub>O<sub>6</sub>F. The [B<sub>3</sub>O<sub>6</sub>F] units with the configuration of [3 :  $\Delta$  + 2T<sub>1</sub>] and BO<sub>3</sub> compose the B–O–F layers which guarantee the large SHG effect and strong optical anisotropy. (d) A modular description of CsB<sub>4</sub>O<sub>6</sub>F. The planar [B<sub>3</sub>O<sub>6</sub>] and [BO<sub>3</sub>F] compose B–O–F layers in crystal structures, which render its strong optical anisotropy and large SHG response. (e) A modular description of RbB<sub>4</sub>O<sub>6</sub>F, a similar structure with NH<sub>4</sub>B<sub>4</sub>O<sub>6</sub>F. (f) A modular description of Na<sub>3</sub>B<sub>3</sub>O<sub>3</sub>F<sub>6</sub>. Three [BO<sub>2</sub>F<sub>2</sub>] form [B<sub>3</sub>O<sub>3</sub>F<sub>6</sub>] 6MR with the configuration of [3 : 3T<sub>2</sub>]. Alkali- and alkaline-earth metals are fillers and described in one block with colors consistent with the corresponding cation.

[3 :  $\Delta$  + 2T] and [3 : 2 $\Delta$  + T], one can obtain appropriate birefringence. Interestingly, some of 6MRs with the configuration of [3 : 2 $\Delta$  + T] can have comparable  $\delta$  to that of [3 : 3 $\Delta$ ], the later may produce a large birefringence around 0.12 as in  $\beta$ -BaB<sub>2</sub>O<sub>4</sub>. Therefore, the 6MRs can have a high value of  $\delta$  and keep a large HOMO-LUMO gap. Moreover, in the known fluorooxoborates, the 6MRs connect  $\Delta$  or T to form polymerized 2D B–O–F framework which further enhances the birefringence. The highlighted electronic density of microscopic units also shows the bonding electron distributions from T<sub>2</sub> to  $\Delta$  and from IV to I in Fig. 2b–g. So introducing [BOF] is a beneficial to obtaining a large

birefringence that can meet the DUV criterion, while also possessing a wide band gap.

To detect birefringence controlled by the fluorooxoborate units, the REDA method was employed [43]. As shown in Equation (1), it is illustrated that the optical anisotropy is proportional to the REDA index  $\xi$ . Usually, the changes of the minimal refractive index  $n_1$  and the band gap  $E_g$  bring tiny changes for alkali/alkaline-earth metal borates. Taking  $E_g$  for example, a band gap changing from 5.5 to 7.5 eV brings a 5% difference. Therefore, the birefringence is mainly related to the difference of electron density of anionic groups, namely,  $\Delta\rho_b$ . Fig. 3 shows the birefringence obtained by the first-principles



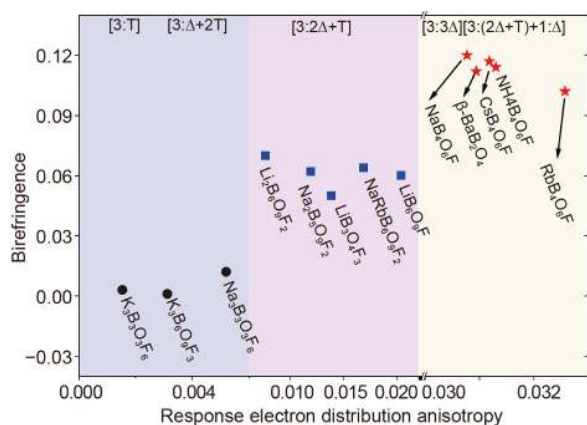
**Figure 2** Polarizability anisotropy and highlighted electronic density of microscopic units. (a)  $\delta$  of the designed 6MRs as mentioned in Fig. 1a along with the HOMO-LUMO gap. (b) Divided polarizability anisotropy along with different configurations of 6MRs, where I refers to [3 : 3 $\Delta$ ], II refers to [3 : 2 $\Delta$  + T], III refers to [3 :  $\Delta$  + 2T] and IV refers to [3 : 3T]. Highlighted electronic density of [BO<sub>2</sub>F<sub>2</sub>] extracted from Na<sub>3</sub>B<sub>3</sub>O<sub>3</sub>F<sub>6</sub> (c), [BO<sub>3</sub>F] extracted from CsB<sub>4</sub>O<sub>6</sub>F (d), [BO<sub>3</sub>] extracted from RbB<sub>4</sub>O<sub>6</sub>F (e). Highlighted electronic density of typical 6MR, IV-[B<sub>3</sub>O<sub>3</sub>F<sub>6</sub>] extracted from Na<sub>3</sub>B<sub>3</sub>O<sub>3</sub>F<sub>6</sub> (f), II-[B<sub>3</sub>O<sub>6</sub>F] extracted from RbB<sub>4</sub>O<sub>6</sub>F (g), and I-[B<sub>3</sub>O<sub>6</sub>] extracted from  $\beta$ -BaB<sub>4</sub>O<sub>6</sub> (h).

calculation or experiment of some typical fluorooxoborate materials as well as the difference of electron density of functional modules. One can see that K<sub>3</sub>B<sub>3</sub>O<sub>3</sub>F<sub>6</sub> with the [3 : 3T] configuration has a relatively small  $\Delta\rho_b$  corresponding to a small birefringence. And Li<sub>2</sub>B<sub>3</sub>O<sub>4</sub>F<sub>3</sub> with [3 :  $\Delta$  + 2T] has the birefringence about 0.04, and the compound with [3 : 2 $\Delta$  + T] has a relatively larger birefringence about 0.07 in Na<sub>2</sub>B<sub>6</sub>O<sub>9</sub>F<sub>2</sub>. And  $\beta$ -BaB<sub>2</sub>O<sub>4</sub> with [3 : 3 $\Delta$ ], NH<sub>4</sub>B<sub>4</sub>O<sub>6</sub>F with [(3 : 2 $\Delta$  + T<sub>1</sub>) + (1 :  $\Delta$ )] and CsB<sub>4</sub>O<sub>6</sub>F with [(3 : 3 $\Delta$ ) + (1 : T<sub>1</sub>)] have comparable and

large birefringences [66]. Therefore, it proves that different birefringences in the discussed borates and fluorooxoborates are mainly attributed to the different configurations in the anionic groups such as 6MR configurations (I–IV). And introducing fluorine can adjust the birefringence into a suitable region.

#### Dispersion of refractive index in affecting PM condition

As an NLO material, the ability of PM is crucial because it determines the applied shortest PM SHG wavelength and



**Figure 3** Birefringences obtained by the first-principles calculation or experiment of typical fluorooxoborate materials with respect to the response electron distribution anisotropy  $\Delta\rho^b$  of fluorooxoborate units.

SHG efficiency of the crystal [39]. Dispersion of refractive index is a vital factor for PM in SHG process. Small dispersion is preferred as it will give a short SHG wavelength. We define the dispersion characteristic factor as

$$D(\lambda) = \frac{n_{\min}(\lambda/2)}{n_{\max}(\lambda)}. \quad (3)$$

Therefore, when the dispersion characteristic factor  $D(\lambda) = 1$ , we can get the shortest PM SHG wavelength  $\lambda_s^{\text{PM}}$ . According to our previous research [43,63],  $D(\lambda)$  is related to,

$$D(\lambda) = \sqrt{1 + \frac{(n_{\max}(\lambda) - \Delta n)^2 - 1}{\lambda^2 - 4R\lambda_c^2}} / n_{\max}(\lambda), \quad (4)$$

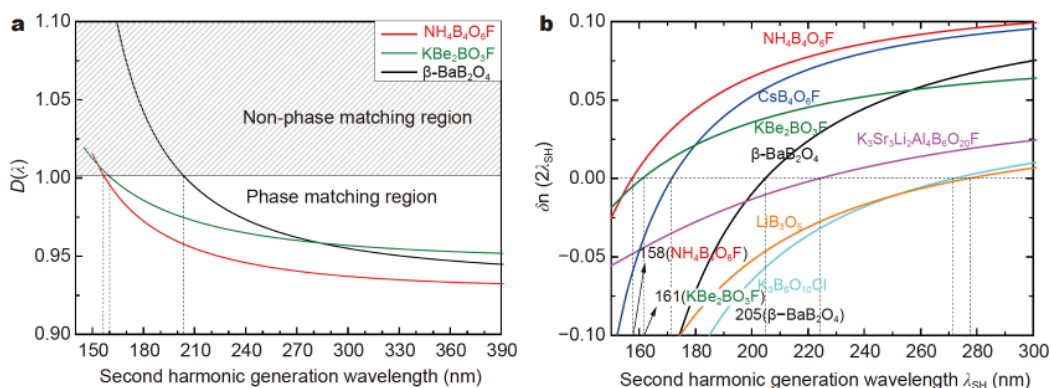
where,  $\lambda_c$  is determined by natural resonant frequency. Fig. 4a, b and Fig. S1 give the consistent results from Equations (3) and (4). At the PM wavelength,  $D(\lambda) = 1$ , from which the shortest PM wavelength  $\lambda_s^{\text{PM}}$  can be

obtained. In the PM region,  $D(\lambda) < 1$  while in the non-phase-matching (NPM) region,  $D(\lambda) > 1$ . And we can also see that a larger band gap or a shorter UV cutoff edge is beneficial to a flatter dispersion at a given wavelength. It can be checked directly from the Sellmeier equations [20,30,34,35,66], as shown in Fig. 4b and Fig. S1.  $\text{NH}_4\text{B}_4\text{O}_6\text{F}$  has a shorter  $\lambda_s^{\text{PM}}$ , although it has a slightly smaller birefringence as compared with  $\beta\text{-BaB}_2\text{O}_4$ . In fact, the factors inducing a shorter  $\lambda_s^{\text{PM}}$ , besides the large band gap and birefringence, including a small chromatic dispersion of refractive index, are also crucial.

A small dispersion is beneficial to realizing the angle PM to a shorter SHG wavelength as described above. Based on the dispersion equation, the birefringence, the refractive index and band gap have influences on the dispersion. One can see that the variation of the refractive index is small. Therefore, the dispersion is mainly determined by birefringence and band gap. This is also verified by the results shown in Fig. S2. In general, the materials with large band gap or birefringence always have smaller dispersion at an identical wavelength. Although KBBF has relatively small birefringence, the large band gap facilitates its PM in DUV region. It is because  $\text{NH}_4\text{B}_4\text{O}_6\text{F}$  has a short PM wavelength owing to its large birefringence and wide band gap, which is proved by the experimental results as shown in Fig. S3. The angle PM can be realized by the optical anisotropy as well as the frequency dispersion.

## CONCLUSIONS

In summary, we studied the influencing factors to control the birefringence and the dispersion of refractive index of recent DUV fluorooxoborates by estimating a dispersion characteristic factor. It reveals that the fluorooxoborates possess strong optical anisotropy due to large covalent



**Figure 4** (a) Refractive index dispersion of different NLO borates, from the borate  $\beta\text{-BaB}_2\text{O}_4$  to the borate fluoride  $\text{KBe}_2\text{BO}_3\text{F}_2$  and the fluorooxoborate  $\text{NH}_4\text{B}_4\text{O}_6\text{F}$ . (b) The shortest PM SHG wavelength based on the corresponding Sellmeier equation.

response electronic distribution induced by the B–O or B–O–F rings and the corresponding quasi-layer arrangement. By exploring the 6MRs (including newly designed ones), it is found that introducing different numbers of fluorine atoms into the ring can modulate the polarizability anisotropy and accordingly influence the birefringence. Furthermore, introducing fluorine into B–O groups can reduce the chromatic dispersion of refractive indices which is beneficial to a shorter PM wavelength. It further proves our previous assumption that the [BOF] group is one kind of superior basic units. Under the analysis of the structures, the designed possible six-membered [BOF] rings can be regarded as new fundamental units to form diverse materials. We hope that this work will be useful in guiding the novel performance-driven materials design.

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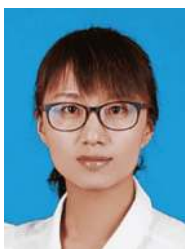
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**Conflict of interest** The authors declare that they have no conflict of interest.

**Supplementary information** Electronic supplementary information (ESI) are available in the online version of the paper.





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### 深紫外氟化硼酸盐双折射率和折射率色散的性能增益研究

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**摘要** 作为紫外/深紫外非线性光学材料的潜在体系, 氟化硼酸盐已引起了该领域的广泛关注. 鉴于该类体系中双折射和折射率色散的影响因素尚未明确, 我们设计了可能的氟化硼酸盐基团, 分析了光学各向异性以探索其对双折射率的影响. 通过响应电荷分布各向异性近似, 我们进一步证明和筛选了有利于双折射率的微观基团, 并系统地探讨了可有效调节双折射率的功能模块. 基于发展的折射率色散分析方法, 我们发现氟化硼酸盐基团的引入有利于降低该类体系在深紫外区的折射率色散, 从而获得较短的相位匹配波长. 该研究为功能驱动的材料设计提供了一条途径.