

Erratum: Viscosity of Water Substance— New International Formulation and Its Background

Cite as: J. Phys. Chem. Ref. Data **7**, 1755 (1978); <https://doi.org/10.1063/1.5145552>
Published Online: 06 July 2020

A. Nagashima



View Online



Export Citation

ARTICLES YOU MAY BE INTERESTED IN

[Wide-Ranging Reference Correlations for Dilute Gas Transport Properties Based on Ab Initio Calculations and Viscosity Ratio Measurements](#)

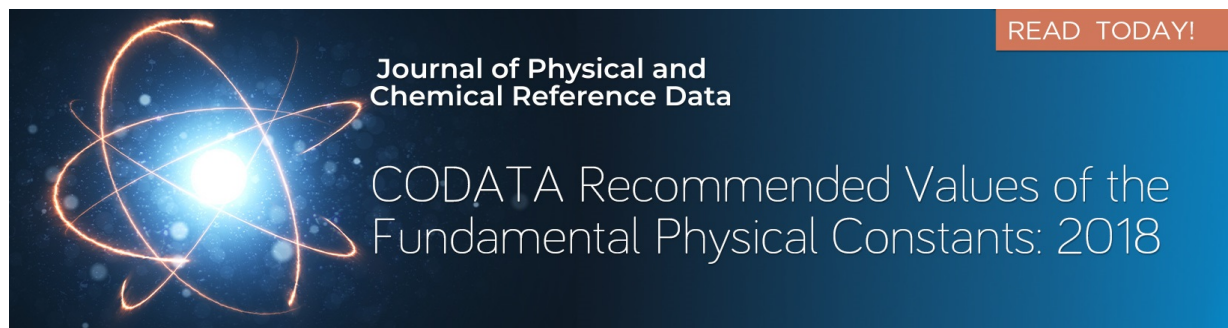
Journal of Physical and Chemical Reference Data **49**, 013101 (2020); <https://doi.org/10.1063/1.5125100>

[Publisher's Note: "Wide-Ranging Reference Correlations for Dilute Gas Transport Properties Based on Ab Initio Calculations and Viscosity Ratio Measurements" \[J. Phys. Chem. Ref. Data 49, 013101 \(2020\)\]](#)

Journal of Physical and Chemical Reference Data **49**, 029901 (2020); <https://doi.org/10.1063/5.0004137>

[Equations of State for the Thermodynamic Properties of Binary Mixtures for Helium-4, Neon, and Argon](#)

Journal of Physical and Chemical Reference Data **49**, 023101 (2020); <https://doi.org/10.1063/1.5142275>



Journal of Physical and
Chemical Reference Data

READ TODAY!

CODATA Recommended Values of the
Fundamental Physical Constants: 2018

Errata

Erratum: Viscosity of Water Substance—New International Formulation and Its Background

[J. Phys. Chem. Ref. Data 6, 1133 (1977)]

A. Nagashima

Keio University, Yokohama, 223, Japan

On page 1158, equation (1) should read:

$$\mu = \mu_0 \exp \left[\frac{\rho}{\rho^*} \sum_{i=0}^5 \sum_{j=0}^4 b_{ij} \left(\frac{T^*}{T} - 1 \right)^i \left(\frac{\rho}{\rho^*} - 1 \right)^j \right],$$

and the constant a_1 should read:

$$a_1 = 0.0177624.$$

Erratum: JANAF Thermochemical Tables: 1978 Supplement

[J. Phys. Chem. Ref. Data 7, 793 (1978)]

M. W. Chase, Jr., J. L. Curnutt, R. A. McDonald, and A. N. Syverud

Thermal Research, The Dow Chemical Company, Midland, Michigan 48640

On page 886 the tabulated values of ΔH_f° and $\Delta H_f^\circ_{298.15}$ for $ZrCl_4$ (c) should be -234.60 ± 0.40 and -234.35 ± 0.40 kcal/mol, respectively. The heat of formation discussion and the numerical tabulation from 0-1000 K agrees with these values. On page 887, the

first sentence in the heat of formation discussion should read "The heat of formation of gaseous $ZrCl_4$, $\Delta H_f^\circ_{298} = -207.93 \pm 0.5$ kcal/mol, is calculated from the heats of formation and sublimation of the crystal (1)."