

Research Article

Formation of Intermetallic Phases in Ion Implantation

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This paper presents a model for the formation of intermetallic phases in the modified nickel ions in the surface layer of aluminum. It is shown that the absorption of ions in the bulk of the qualitative difference between the models with and without the relaxation of the mass flux is reduced to a difference in the characteristic scales. It was shown that the concentration distribution depends on the relation between time scales of various physical processes. We have extended the existing model to a unique simple model describing the formation of a new phase at the initial stage of ion implantation. The parameters containing in the model were evaluated using literature data. The known problem is a special case for our model.

1. Introduction

Intermetallics—chemical compounds of two or more metals—are widely used in modern science and technology, as they retain their structure and properties at high temperatures. The volume of intermetallic alloys is of great interest based on intermetallic compounds, which are applied in different ways. For example, a coating based on NiTi, obtained by laser ablation [1] significantly increases the wear resistance of the alloy Ti–6Al–4V. Coating the Ti–Cu system significantly improves the surface properties of copper [2]. Intermetallic Al–Cu system as coating deposition using CVD method [3] increases the corrosion resistance of steels and other compounds. Improved surface properties due to the formation of intermetallic phases, including intermetallic nanostructures, were possibly due to the implantation of metal ions in the surface layers of other metals. Reports of studies in this area have appeared for a long time [4]. These authors reported the formation of different phases in surface layer of steels and alloy layers depending on the conditions of

implantation and subsequent annealing temperature. Formation of multilayer nanostructures in the Al–Ti ion implantation of Ar⁺ [5] is also accompanied by the formation of intermetallic compounds. The authors of [6, 7] have undertaken targeted research of formation of intermetallic phases in the surface layers, leading to changes in the mechanical properties, depending on the mode of implantation. These papers argue that the formation of chemical compounds is an important modification component of surface properties. The first model of the process of ion implantation with the chemical steps was proposed in [8], where the phase sequence obtained surface layer, similar to the experiments [6]. However, the nonequilibrium nature of the evolution of the surface layer that is associated with the nonequilibrium mass transfer in these studies was not considered. Wave processes were at the initial stage of the study of ion implantation, for example, in [9, 10]. A simplified model of the process with a single chemical reaction in the surface layer, but considering the finite relaxation time of the mass flow, is proposed in [11, 12].

In this paper, continuing research, we take a model for nonequilibrium conditions of three-stage reaction for the formation of intermetallic compounds.

2. Statement of the Problem

We consider a model for nonequilibrium conditions of three-stage reaction for the formation of intermetallic compounds. Consider the flat nickel layer. The flow of aluminum particles is distributed uniformly along the surface to be treated. Suppose that the implantation of aluminum ions into the nickel surface may occur as follows:

- (a) $\text{Ni} + \text{Al} \longrightarrow \text{NiAl}$
- (b) $3\text{Ni} + \text{Al} \longrightarrow \text{Ni}_3\text{Al}$
- (c) $\text{Ni}_3\text{Al} + 2\text{Al} \longrightarrow 3\text{NiAl}$

The molar concentration distribution of Ni, Al, NiAl, and Ni_3Al follows from equations. For simplicity, we assume that $Y_1 = [\text{Al}]$; $Y_2 = [\text{Ni}]$; $Y_3 = [\text{NiAl}]$; and $Y_4 = [\text{Ni}_3\text{Al}]$, and the reaction rate can be written in accordance with the law of mass action

$$\begin{aligned}\phi_1 &= k_1 Y_1 Y_2, \\ \phi_2 &= k_2 Y_1 Y_2^3, \\ \phi_3 &= k_3 Y_1^2 Y_4.\end{aligned}\quad (1)$$

Then, the concentration will vary in accordance with the following equation:

$$\begin{aligned}\frac{\partial Y_1}{\partial t} &= -\frac{\partial J_1}{\partial x} + \omega_1 + q_m F(x, t), \\ \frac{\partial Y_2}{\partial t} &= -\frac{\partial J_2}{\partial x} + \omega_2, \\ \frac{dY_3}{dt} &= \omega_3, \\ \frac{dY_4}{dt} &= \omega_4,\end{aligned}\quad (2)$$

where $F(x, t) = F_1(x)F_2(t)$, q_m denotes maximum flux density of particles, and

$$\omega_k = \sum_{i=1}^r \nu_{ki} \phi_i, \quad r = 3; i = 1, 2, 3; k = 1, 2, 3, 4. \quad (3)$$

Particle flows ($\text{mol}/(\text{m}^2\text{s})$) satisfy the generalized law

$$\begin{aligned}J_1 &= -D_1 \frac{\partial Y_1}{\partial x} - t_1 \frac{\partial J_1}{\partial t}, \\ J_2 &= -D_2 \frac{\partial Y_2}{\partial x} - t_2 \frac{\partial J_2}{\partial t},\end{aligned}\quad (4)$$

where D_1 and D_2 are diffusion coefficients and t_1 and t_2 are the relaxation times.

The boundary conditions are as follows:

$$J_1 = J_2 = 0, x = 0, \text{ and } x = L. \quad (5)$$

At the initial time,

$$\begin{aligned}Y_1 &= Y_3 = Y_4 = 0, \\ Y_2 &= Y_{20}.\end{aligned}\quad (6)$$

Numerically, we use double-sweep method and implicit-difference scheme. The computations are concentration distributions of elements and phases at various times while varying the model parameters.

3. Results Analysis

For illustrations taken,

$$F(x, t) = \exp(-Ax^2)H(t - t_{\text{imp}}), \quad (7)$$

where $H(t - t_{\text{imp}})$ is a unit step function; t_{imp} is the pulse duration; and $A = 50^{-2}\text{cm}^{-2}$. Diffusion coefficients and reaction rates depend on temperature by Arrhenius law. The activation energy of reactions is taken from [13] (Table 1).

The diffusion coefficients are the same, as in [11, 12].

If $t_1 = t_2 = 0$, this model is useful for slow process, when the reaction rate is controlled by slow diffusion. The pre-exponential factor can adapt with ideal conditions for a chemical reaction when there is no kinetic troubles as shown in Table 1. The mass concentration can be determined utilizing the following formula:

$$C_k = \frac{Y_k m_k}{\sum_{i=1}^4 Y_i m_i}, \quad (8)$$

where m_k denotes molar masses of substances. From Figure 1 it is clear that, at a temperature $T = 600\text{K}$ phase, Ni_3Al is practically not formed. Appearing as an intermediate, it is quickly spent on the formation of phase NiAl. Practically, the diffusion at this temperature is absent. An injected item is consumed where it has been injected. After exhausting Al, in this area, the reaction is almost stopped. At a temperature $T = 800\text{K}$, the ratio between the reaction rate changes. As a result, there is a region where we have a finite portion or part of the phase of Ni_3Al .

With increasing temperature to $T = 1000\text{K}$, the diffusion zone is activated. Evidently, in the treating zone, the phase is preferably NiAl. This can be followed by a zone containing predominantly phase Ni_3Al . Zone sizes depend on the temperature of the mass flux density, which is in qualitative agreement with the data [6, 7].

Model for $t_1 \neq 0$ and $t_2 \neq 0$ is of interest for treating conditions of short pulses with a high particle density. To ensure that the observed reaction under these conditions, it is necessary to take into account the possible acceleration of the reactions and diffusion, for example, by activation [14]. Qualitatively different effects in a model taking into account the relaxation compared to the traditional version of the model are not found. For example, when $t_1 = 1 \cdot 10^{-4}$ and $t_2 = 1 \cdot 10^{-3}$ at the initial stage of the processing flow particle density (see equation for the first component) $q_m = 1.5 \cdot 10^5 (\text{mol}/\mu\text{m}^3\text{s})$, the chemical reaction is observed, if the speed increases from $(10^4 \div 10^8)$.

Figure 2 shows a steady slow increase in the occupied phase area NiAl and a mixture of the two phases

TABLE 1: Formal-kinetic parameters of reactions.

-	$\text{Ni} + \text{Al} \longrightarrow \text{NiAl}$	$3\text{Ni} + \text{Al} \longrightarrow \text{Ni}_3\text{Al}$	$\text{Ni}_3\text{Al} + 2\text{Al} \longrightarrow 3\text{NiAl}$
$E_a, \text{J}\cdot\text{mol}^{-1}$	86128.77	169148.544	148715.212
k_0	$8.994 * 10^4$	$1.517 * 10^9$	$0.853 * 10^2$

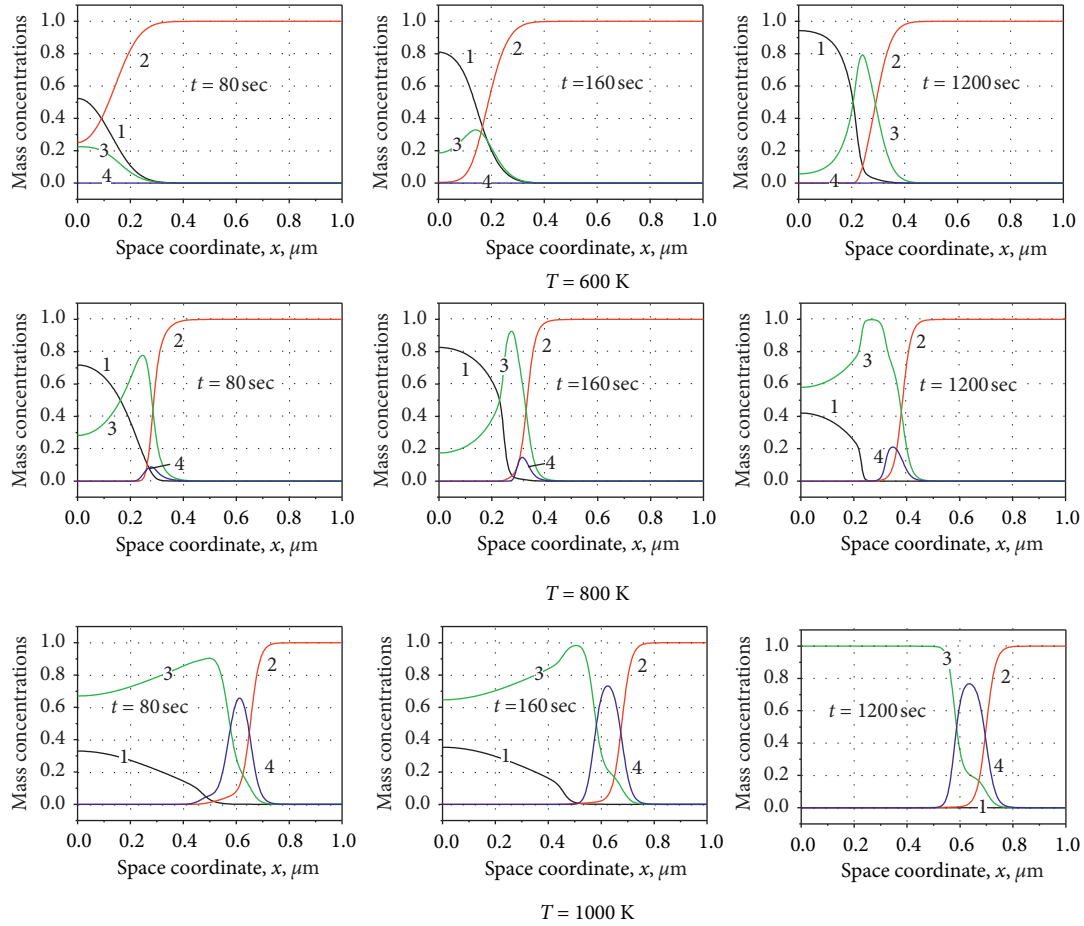


FIGURE 1: The distribution of phases and elements at various times and at various temperatures are given for the slow process, $q_m = 0.3 \text{ (mol}/\mu\text{m}^3\text{s)}$; $t_{\text{imp}} = 400 \text{ sec}$.

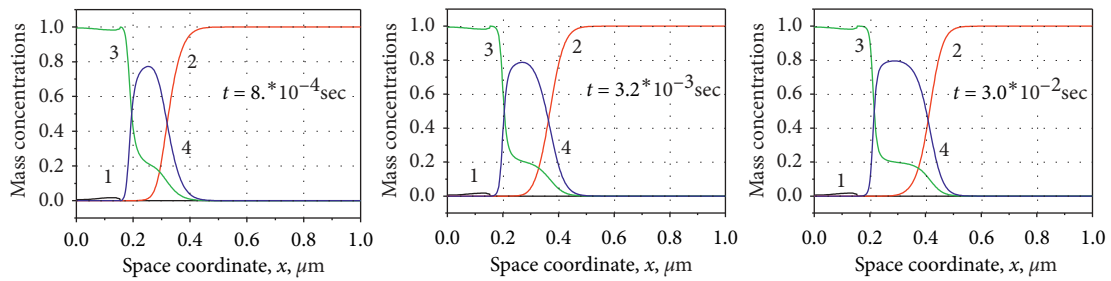


FIGURE 2: Distribution of phases and elements at various times at a temperature $T = 1000 \text{ K}$ for a “fast” process: a steady stream.

NiAl + Ni₃Al. This feature is only for a richer separation of the two zones.

4. Novelty of the Work [15]

The paper, for the first time, formulated and investigated related models on thermal elastic diffusion with chemical reactions. The algorithms were developed excluding the possibility of inconsistency appearance in terms of physical solutions. Based on the results of numerical modeling, the new effects were revealed. The work presents a new knowledge in the thermal elastic diffusion theory with chemical reactions. Results of the study can be used to develop models and methods of composition modification with acceptable treated samples. Models admit their further development by taking into account the different parameters and conditions of ion deposition.

5. Conclusion

Thus, in this paper, a model has been constructed in order to describe the intermetallic formation on surface layers during ion implantation, with the assumption of isothermal condition. The implicit-difference scheme has been suggested for the solution of diffusion kinetic problem describing ion implantation by intermetallic phase formation. For further interesting models and methods, we refer the readers to [16–32]. We actually suggest a model of the surface modification of nickel-aluminum ions with the relaxation of mass flows. The model corresponds to irreversible conditions and includes finiteness of relaxation times for mass fluxes. The results illustrate the convergence of difference schemes at variation of its parameters. It has also been shown that the finiteness of relaxation time changes the concentration distribution in the diffusion zone in comparison with usual diffusion problems with high temperature.

Data Availability

The data used to support the findings of this study are available from the corresponding author upon request.

Conflicts of Interest

The authors declare that they have no conflicts of interest.

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