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Ronald G. McLaren, Ravi Naidu, J. Smith, K. G. Tiller

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# Fractionation and Distribution of Arsenic in Soils Contaminated by Cattle Dip

R. G. McLaren,\* R. Naidu, J. Smith, and K. G. Tiller

## ABSTRACT

Soils surrounding cattle dips in Australia are known to be highly contaminated with As and are potentially of major concern to the environment and human health. A study of 11 dip sites in northern New South Wales (NSW) has revealed considerable surface soil (0–10 cm) contamination with As (37–3542 mg As kg<sup>-1</sup> soil). In addition it was shown that considerable movement of As down through the soils had occurred with concentrations at 20 to 40 cm ranging from 57 to 2282 mg As kg<sup>-1</sup> soil. At one particular site, an As concentration of above 14 000 mg As kg<sup>-1</sup> soil was determined at a depth of 40 to 45 cm. A sequential fractionation scheme, based on a soil P fractionation, was developed to assess the chemical nature, and thus the potential bioavailability and mobility of As, at the sites. Soil As is separated into six fractions with (i) anion exchange resin, (ii) NaHCO<sub>3</sub>, (iii) NaOH, (iv) NaOH following sonication, (v) HCl, and (vi) HCl/HNO<sub>3</sub>. Although substantial differences between sites and soil depths were present, some general trends were apparent. Most sites contained substantial concentrations of As in the two most labile fractions, indicating high potential for phytotoxicity and leaching. The bulk of the contaminant As at the sites seemed to be associated with soil amorphous Fe and Al minerals.

FROM the turn of the century until the early 1950s, sodium arsenite was used extensively in cattle dips in Queensland and NSW, Australia, for controlling cattle ticks. Between 1920 and 1980, more than 1600 cattle dips were constructed in northeastern NSW alone (Vaughan, 1993; Barzi et al., 1996). The dips are constructed of concrete and are deep enough to completely immerse cattle as they are driven through the dip via entrance and exit ramps. During passage of stock through the dip, considerable splashing of the dip contents occur and the areas immediately surrounding the dips are often highly contaminated with As, with concentrations as high as 21 000 mg As kg<sup>-1</sup> soil (Vaughan, 1993). Population growth in northeastern NSW is increasing

R.G. McLaren, Department of Soil Science, P.O. Box 84, Lincoln University, Canterbury, New Zealand; R. Naidu, J. Smith and K.G. Tiller, CSIRO Division of Soils and the Cooperative Research Centre for Soil and Land Use Management, Private Mail Bag no. 2, Glen Osmond, S.A. 5064. Received 21 May 1997. \*Corresponding author (mclaren@lincoln.ac.nz).

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rapidly, and many former dip sites are now being encroached upon by urban developments, including residential housing. Such developments have raised concerns about possible harmful effects resulting from the residual As at these sites, including, for example, ingestion of contaminated soil by children or household pets, As toxicity to plants, and the potential release of As into surface and groundwaters. At some urban sites, the NSW Department of Agriculture is considering digging and storing contaminated soils in closed sheds. The Tweed Valley City Council recently expressed the possibility of storing 20 000 m<sup>3</sup> of soil in their region (Tweed Valley Council, 1996, personal communication).

Contamination of orchard soils with As-containing pesticides such as lead arsenate and various organoarsenical compounds has been studied in some detail (Woolson et al., 1971a; Merry et al., 1983; Veneman et al., 1983). However, there has been no comparable study of As contamination at cattle dip sites. In contrast to orchard locations, As contamination at dip sites is much more localized, and the source was predominantly sodium arsenite. The objective of this study was to examine the nature of As contamination at a range of former cattle dip sites in northeastern NSW, and develop techniques that might be useful for assessing the potential bioavailability and mobility of the As present in the contaminated soils.

It is generally recognized that total As concentration in the soil is not necessarily a good indicator of potential bioavailability or mobility. Several previous studies of As-contaminated soils have attempted to assess bioavailability using different extractants (Woolson et al., 1971b; Johnston and Barnard, 1979; Lee and Low, 1990). For other soil contaminants, particularly trace metals, detailed characterizations are often carried out using sequential extraction techniques. However, the feasibility of using sequential extraction schemes designed for transition metals for fractionating soil As has been questioned (Grubel et al., 1988). An alternative approach for As has been to use sequential extraction schemes

**Abbreviations:** AEM, anion exchange membrane; AAS, atomic absorption spectrophotometry; CCA, copper, chromium, and arsenic.

designed to study the forms of P in soils. The chemistry of As in soils has apparent similarities with that of P in that they both commonly form oxyanions in the +V oxidation state (O'Neill, 1990). In the 1970s, researchers used adaptations of the Chang and Jackson soil P fractionation procedure (Chang and Jackson, 1957) to examine As in contaminated soils (Jacobs et al., 1970; Woolson et al., 1971a, 1973; Johnston and Barnard, 1979). However, it is recognized that the Chang and Jackson procedure presents many interpretational problems for soil P (Tiessen and Moir, 1993), and, by analogy, for soil As. An alternative P-fractionation scheme (Hedley et al., 1982) is now widely used in soil P studies. In this paper an adaptation of the Hedley et al. (1982) scheme is used to fractionate As in soils from contaminated cattle dip sites. It was assumed that although As was originally added to the dips as As(III), under the predominantly aerobic soil conditions at the sites, oxidation to As(V) would have occurred during the 40 yr since As was last used.

## MATERIALS AND METHODS

### Soil Sampling and Preparation

Soils were sampled adjacent to cattle dip installations in northern NSW. At each site, two soil pits were located 1 m from the edge of the cattle dip, with the first pit 1.5 m from the point at which cattle would have entered the dip and the second pit 2 to 3 m further along the dip. Surface (0–10 cm) and subsurface (20–40 cm) soil samples were taken from each pit and then bulked to give single surface and subsoil samples from each site. Sampling depths were varied at some sites, to avoid inclusion of material from more than one recognizable soil horizon in a single sample. In some cases, more than two samples were taken to include material from other horizons, including some samples taken at depths >40 cm. The bulk samples were air-dried and ground to pass a 2 mm sieve. Some relevant properties of the samples are listed in Table 1.

### Total, Water-Extractable, and Acid Oxalate-Extractable Arsenic

For total soil As, a 1-g sample was digested with 10 mL aqua regia (5:3 HCl/HNO<sub>3</sub>) in a Milestone microwave digestion unit. Using 2 M HCl, the resulting digest was transferred to a 25-mL volumetric flask and made up to volume. The digest was then filtered through Whatman no. 42 filter paper.

For water-extractable As, a 5-g soil sample was weighed into a polypropylene centrifuge tube and shaken with 25 mL deionized water for 16 h. The suspension was then centrifuged, and the supernatant solution filtered through a 0.45 µm nylon filter.

For acid oxalate-extraction, a 1-g soil sample was weighed into a centrifuge bottle and 100 mL of oxalate reagent was added (0.2 M ammonium oxalate, 0.2 M oxalic acid, pH 3). The suspension was shaken for 4 h in the dark and then filtered through no. 42 Whatman filter paper.

### Sequential Fractionation of Soil Arsenic

**Step 1: Anion exchange membrane-extractable As.** A sample of soil (2 g) and a 62.5 mm by 25 mm strip of anion exchange membrane (AEM) (BDH Product 55164 2S) were suspended in 30 mL of deionized water in a polypropylene centrifuge tube and shaken for 24 h. The AEM strip was then carefully removed from the suspension, washed with deionized water to remove any adhering soil particles and then placed in another tube containing 30 mL of 0.1 M HCl. The tube containing the strip was then shaken for 1 h and As determined in the resulting solution.

Following removal of the AEM strip, the soil suspension was centrifuged and the supernatant solution decanted from the soil residue.

**Step 2: Sodium bicarbonate-extractable As.** The soil residue from Step 1 was resuspended in 30 mL of NaHCO<sub>3</sub> (0.5 M, pH 8.5) and shaken for 16 h. The sample was then centrifuged and the supernatant solution filtered through a 0.45 µm nylon filter.

**Step 3: Sodium hydroxide-extractable As.** The soil residue from Step 2 was resuspended in 30 mL of NaOH (0.1 M) and

Table 1. Properties of soils at cattle dip sites.

Site	Depth cm	pH†	Org. C‡ %	Oxalate		Resin-P¶ mg kg <sup>-1</sup>	Clay# %	Sand#
				Fe§ g kg <sup>-1</sup>	Al§ g kg <sup>-1</sup>			
Geraghty	0–10	6.05	5.59	8.0	1.3	75.9	24.1	55.0
	20–40	6.99	1.25	10.5	1.4	57.0	38.7	43.1
	40+	6.87	2.14	19.7	2.7	82.9	31.7	40.2
Lindendale	0–6	6.16	14.73	9.3	6.5	169.5	17.5	33.3
	6–10	5.71	1.56	3.8	4.3	151.2	30.5	48.3
	20–40	5.93	2.82	5.7	4.4	27.1	23.7	51.1
Sandy Flat	0–8	4.83	7.03	8.4	1.8	32.1	23.1	52.2
	20–40	4.97	1.54	1.5	0.2	45.0	8.1	87.9
	40–45	12.16	3.97	1.6	1.3	9.3	11.7	29.5
Willow Bank	0–10	5.65	5.39	14.3	3.5	26.9	42.3	22.0
	20–40	5.17	3.24	7.3	3.2	44.6	43.3	22.5
Lynwood	0–10	5.70	6.69	11.1	4.5	89.2	30.9	25.7
	20–40	5.83	2.75	4.5	3.2	12.5	27.3	33.4
Hetherington	0–10	5.48	1.11	1.1	0.3	105.2	6.9	88.3
	20–40	5.43	0.47	0.8	0.7	79.7	9.3	89.9
Wagner	0–10	5.57	5.75	9.8	1.0	72.1	27.9	48.6
	15–25	6.48	2.11	12.9	2.3	53.5	42.9	21.1
	25–40	6.33	1.35	7.7	2.1	22.3	46.7	23.2
Ridges	0–10	5.46	2.98	7.0	3.6	35.6	37.7	47.2
	20–40	5.93	3.30	6.4	6.0	19.7	36.9	30.5
Laurel	0–10	5.85	5.30	8.5	2.9	41.9	31.9	41.8
	20–40	6.61	1.83	8.7	3.8	51.9	ND	ND
Benaud	0–10	5.72	2.57	1.1	0.6	29.5	6.1	83.3
	20–40	6.18	0.53	0.4	0.0	26.7	8.5	87.6
Zeus	0–10	5.45	4.47	5.1	1.1	40.6	ND	ND
	20–40	6.57	0.98	1.8	1.1	60.3	ND	ND

† 1:5 water; ‡ Leco CNS analyzer; § Blakemore et al., 1987; ¶ Step 1 in fractionation; # Pipette method; ND = not determined.

shaken for 16 h. The suspension was then centrifuged and the supernatant solution filtered through Whatman no. 42 filter paper.

**Step 4: Sonicated/sodium hydroxide-extractable As.** The soil residue from Step 3 was resuspended in 30 mL of NaOH (0.1 M) and was then dispersed for 3 min with an ultrasonic probe (Branson sonifier, model B-30, operated at 20 kHz, meter reading 20). The suspension was then shaken for 16 h, centrifuged, and filtered through Whatman no. 42 filter paper.

**Step 5: Hydrochloric acid-extractable As.** The soil residue from Step 4 was resuspended in 30 mL of HCl (1 M) and shaken for 16 h. The suspension was then centrifuged and the supernatant solution filtered through a 0.45  $\mu\text{m}$  nylon filter. The soil residue was oven-dried at 60°C for 48 h and then finely ground in an agate mortar.

**Step 6: Residual Arsenic.** Arsenic in the dried and ground residue from Step 5 was determined as for total soil As.

### Analytical Methods

All the above extractions involved shaking for the required time on an end-over-end shaker, and all centrifuging was carried out at 10 000 g for 10 min. If possible, centrifuged extracts were filtered through 0.45  $\mu\text{m}$  nylon filters, however, with some extracts this was not practical and so Whatman no. 42 paper filters were used. Arsenic in the filtrates was determined by direct atomic absorption spectrophotometry (AAS) using a nitrous oxide-acetylene flame, or after hydride generation using an air-acetylene flame. Iron and aluminium in the oxalate filtrates were also determined by AAS (Blakemore et al., 1987). All analyses were carried out in duplicate and all results shown are mean values.

## RESULTS AND DISCUSSION

Total As concentrations in uncontaminated surface soils worldwide range from <0.1 to 95 mg kg<sup>-1</sup> with mean values generally <20 mg kg<sup>-1</sup> (Kabata-Pendias and Pendias, 1992). There are very few data available for

Australian soils, although Merry et al. (1983) reported a mean As concentration of 3.9 mg kg<sup>-1</sup> for 15 South Australian surface soils and 0.60 mg kg<sup>-1</sup> for six soils from Tasmania. In comparison to such values, surface and subsoils at all the cattle dip sites sampled are heavily contaminated with As (Table 2). Surface soil concentrations range from 29.8 to 3542 mg As kg<sup>-1</sup> soil and subsoil concentrations from 56.7 to 2282 mg kg<sup>-1</sup>. At all sites there has been an apparent substantial downward movement of As, although the pattern of distribution with depth varies between sites. Because of the permanent and waterproof nature of dip construction at these sites, it is extremely unlikely that any of the subsoil As contamination has resulted underground lateral flow from the dips. Seven of the sites (Geraghty, Sandy Flat, Willow Bank, Lynwood, Hetherington, Benaud, and Zeus) show much higher concentrations in the subsoil compared to the surface (Table 2). With the exception of Geraghty and Sandy Flat, these sites have relatively uniform soil profiles and in the case of the Hetherington and Benaud sites are extremely sandy in texture and have very low oxide concentrations (Table 1). At the Geraghty site, the major accumulation of As occurs below 40 cm, coinciding with a marked increase in iron (and Al) oxides at this depth in the soil profile (Table 1). The Ridges and Laurel sites are the only sites where As concentrations are highest in the surface horizons. These two sites also have relatively uniform soil profiles, but compared to Hetherington, Benaud and Zeus sites in particular, are much less sandy and have much higher Fe and Al oxide concentrations (Table 1). At the Wagner and Lindendale sites there were peak accumulations of As in the middle horizons sampled (Wagner, 15–25 cm; Lindendale 6–10 cm, Table 2). In both cases, these accumulations correspond with an increase in ox-

Table 2. Total and extractable soil As concentrations (mg As kg<sup>-1</sup> soil).

Site depth	cm	As fractionation							Total As	Water-extractable As	Oxalate-extractable As
		Resin	NaHCO <sub>3</sub>	NaOH(1)	NaOH(2)	HCl	Residual	Sum of fractions			
Geraghty	0–10	95	87	687	189	371	259	1 690	1 920	22.4	1 210
	10–40	133	92	499	145	95	335	1 300	1 220	22.4	883
	40+	220	287	1 460	366	342	827	3 500	3 550	36.7	2 540
Lindendale	0–6	4	10	27	71	54	109	275	380	1.0	133
	6–10	46	84	556	261	0	566	1 510	1 480	10.0	671
	20–40	1	0	43	19	0	8	71	103	0.0	24
Sandy Flat	0–8	4	15	208	71	26	88	412	446	0.9	267
	20–40	17	20	182	38	13	39	309	310	7.6	265
	40–45	52	10 400	974	249	1 950	458	14 100	14 800	1.2	4 310
Willow Bank	0–10	5	22	277	124	0	177	605	658	0.2	304
	20–40	6	153	1 110	382	0	702	2 350	2 230	1.1	1 160
Lynwood	0–10	3	3	90	45	2	90	233	276	0.3	128
	20–40	47	127	1 070	353	0	575	2 170	2 100	7.2	830
Hetherington	0–10	1	0	12	3	2	0	18	30	0.6	12
	20–40	3	0	36	9	3	9	60	57	1.5	33
Wagner	0–10	58	55	778	242	276	109	1 520	1 440	13.1	1 180
	15–25	180	186	1 020	179	209	357	2 130	2 050	35.7	1 840
	25–40	73	80	413	74	28	167	835	844	14.6	707
Ridges	0–10	33	127	869	247	38	548	1 860	1 630	2.6	1 080
	20–40	3	5	336	109	0	130	583	629	0.1	238
Laurel	0–10	104	130	1 050	540	49	1 530	3 400	3 540	8.0	1 410
	20–40	102	144	936	405	19	942	2 550	2 280	2.2	1 070
Benaud	0–10	2	13	3	5	5	3	31	37	0.5	15
	20–40	5	13	38	2	2	6	66	88	1.8	47
Zeus	0–10	7	9	92	17	21	27	173	165	1.7	189
	20–40	12	7	139	14	30	13	215	232	2.1	222

ide iron and/or clay concentrations in the profile (Table 1).

The most contaminated soil occurred at Sandy Flat. Although the 0 to 8 cm and 20 to 40 cm samples showed moderate contamination compared to some of the other sites (Table 2), a sample taken from 40 to 45 cm had an As concentration of 14 760 mg kg<sup>-1</sup> (Table 2). The upper part of the soil profile at this site is very sandy and the 20 to 40 cm horizon in particular has very low Fe and Al oxide concentrations (Table 1). However, the 40 to 45 cm horizon has some very unusual properties that could account for the marked As accumulation in this layer. The horizon is far less sandy than the horizons above and has an extraordinary high pH of 12.16 (Table 1). The 21% CaCO<sub>3</sub>, present in this horizon could not on its own account for the high pH, although the presence of Ca(OH)<sub>2</sub> could. Apart from CaCO<sub>3</sub>, the mineralogy appears to be dominated by poorly crystalline kaolin with traces of quartz. Analysis of soil solution from this horizon and speciation modelling using MINTEQA2 (Allison and Brown, 1995) suggested that As in solution (300 mg As L<sup>-1</sup>) may be in equilibrium with calcium arsenate. This particular horizon remains something of an enigma.

Downwards movement of As through the soil profile has been recorded in orchard soils contaminated with lead arsenate pesticide (Merry et al., 1983; Peryea and Creger, 1994; Elfving et al., 1994) or sodium arsenite herbicide (Tammes and de Lint, 1969). In such situations both the level of As contamination and degree of leaching were much less than at the cattle dip sites. Comparable levels of As contamination and leaching to those found at the dip sites have been recorded in coarse-textured soil contaminated with copper, chromium, and arsenic (CCA) timber preservative solution (Armishaw et al., 1994).

The substantial downwards movement of As at the cattle dip sites results from several factors. During cattle dipping, the soil surrounding the dip quickly becomes saturated with fluid from the dip as the cattle splash through. In such a situation, As is likely to be leached downwards via preferential flow, particularly in coarse-textured or well-structured soils (McLaren et al., 1994; Carey et al., 1996). McLaren et al. (1994) also demonstrated that As leaching is enhanced by short contact periods with the soil. This would certainly be the case during cattle dipping where the continuous input of As-containing liquid from the dip would ensure optimum leaching conditions. Although use of As at the dip sites was discontinued in the early 1950s, cattle dipping at many of the sites has continued using a variety of organic pesticides. This would have ensured that large volumes of water would still have passed through the soil at frequent intervals, with the potential to cause further As leaching during the past 40 yr. In addition, some of the organic chemicals used since the 1950s have been stabilized in the dips by the addition of substantial amounts of phosphate (Kerryn McDougall, 1995, personal communication). Addition of phosphate can enhance As leaching in soils (Davenport and Peryea, 1991; Melamed et al., 1995; Peryea and Kammereck, 1997).

At many of the dip sites, soil concentrations of resin-extractable phosphate were present at concentrations much higher than those found in normal agricultural soils (Table 1). Subsequent laboratory leaching column studies with phosphate solutions have shown significant leaching of As from the contaminated soils (R. Naidu, 1996, unpublished data).

Although total soil As concentration gives a good indication of the extent of soil contamination, it is not a reliable indicator of potential toxicity to plants (Vaughan, 1993), or presumably of potential for As leaching. Sequential extraction of soil As helps to differentiate between As that is readily labile and that which is bound strongly by the soil. The fractionation used in this study consists of six separate steps and thus, as with any such technique, has the potential for considerable combined experimental and analytical errors. To check the procedure, the sums of the As concentrations determined in the six individual fractions were compared with the single determinations of total As for the dip site samples (Table 2). There was extremely good agreement between the two sets of values ( $As_{sum} = 1.02As_{total} - 9.03$ ,  $r^2 = 0.99$ ).

Although no fractionation scheme can be considered 100% effective in distinguishing between different forms of an element, by analogy with the P fractionation scheme of Hedley et al. (1982) as discussed by Tiessen and Moir (1993), the forms of As present in the six fractions are designated as follows:

1. Resin-extractable As: freely exchangeable As.
2. NaHCO<sub>3</sub>-extractable As: nonexchangeable but readily labile As associated with soil mineral surfaces.
3. NaOH-extractable As: As held more strongly by chemisorption to Fe and Al components of soil surfaces.
4. Sonication/NaOH-extractable As: As held at the internal surfaces of soil aggregates.
5. HCl-extractable As: Ca-associated As.
6. Residual As: Highly recalcitrant As.

It is assumed that As is associated predominantly with inorganic soil constituents. Organic forms of As do not appear to compose an appreciable proportion of total soil As (Johnson and Hiltbold, 1969).

Considerable variation occurs between the sites and soil depths in the amounts and proportions of As present in individual fractions (Tables 2 and 3). However, in spite of this variation, some general trends are apparent. The most labile As fraction (resin-extractable As) accounts for <10% of the total soil As (mean 3.9%, Table 3). In most soils, relatively small proportions of total As are also extracted in the second fraction (NaHCO<sub>3</sub>-extractable As). However the 40 to 45 cm sample from Sandy Flat has nearly 74% of its total As in this fraction. The unusual properties of this sample have been commented on above. The Benaud site samples also have a high proportion of their total As in the NaHCO<sub>3</sub>-extractable fraction (20–40%, Table 3). This soil at this site is extremely sandy and has very low Fe and Al oxide concentrations and thus probably has a much lower As-

**Table 3. Proportions of As in individual fractions (% sum of all fractions).**

Site	Depth	Resin As	NaHCO <sub>3</sub> As	NaOH(1) As	NaOH(2) As	HCl As	Residual As
	cm						
Geraghty	0-10	5.6	5.2	40.7	11.2	22.0	15.3
	20-40	10.2	7.1	38.4	11.2	7.4	25.8
	40+	6.3	8.2	41.7	10.5	9.8	23.6
Lindendale	0-6	1.4	3.6	9.9	25.9	19.5	39.7
	6-10	3.0	5.6	36.7	17.3	0.0	37.4
	20-40	0.9	0.0	60.5	26.6	0.5	11.4
Sandy Flat	0-8	0.9	3.7	50.5	17.3	6.3	21.4
	20-40	5.6	6.4	58.8	12.2	4.3	12.6
	40-45	0.4	73.9	6.9	1.8	13.8	3.3
Willow Bank	0-10	0.8	3.6	45.8	20.6	0.0	29.2
	20-40	0.3	6.5	47.2	16.2	0.0	29.8
Lynwood	0-10	1.1	1.2	38.7	19.5	0.7	38.8
	20-40	2.2	5.8	49.3	16.3	0.0	26.5
Hetherington	0-10	5.8	0.0	69.1	15.4	9.6	0.0
	20-40	4.2	0.0	61.2	14.7	5.3	14.6
Wagner	0-10	3.8	3.6	51.3	15.9	18.2	7.2
	15-25	8.4	8.7	48.0	8.4	9.8	16.7
	25-40	8.7	9.6	49.5	8.9	3.3	20.0
Ridges	0-10	1.8	6.8	46.7	13.3	2.0	29.5
	20-40	0.5	0.8	57.7	18.7	0.0	22.3
Laurel	0-10	3.1	3.8	30.8	15.9	1.4	45.0
	20-40	4.0	5.6	36.7	15.9	0.7	37.0
Benaud	0-10	4.9	42.3	9.7	16.2	16.1	10.8
	20-40	9.1	20.2	57.8	2.6	2.7	8.8
Zeus	0-10	3.8	5.4	53.7	9.7	11.9	15.4
	20-40	5.7	3.5	64.7	6.5	13.7	6.0
Mean values		3.9	9.3	44.7	14.2	6.9	21.1

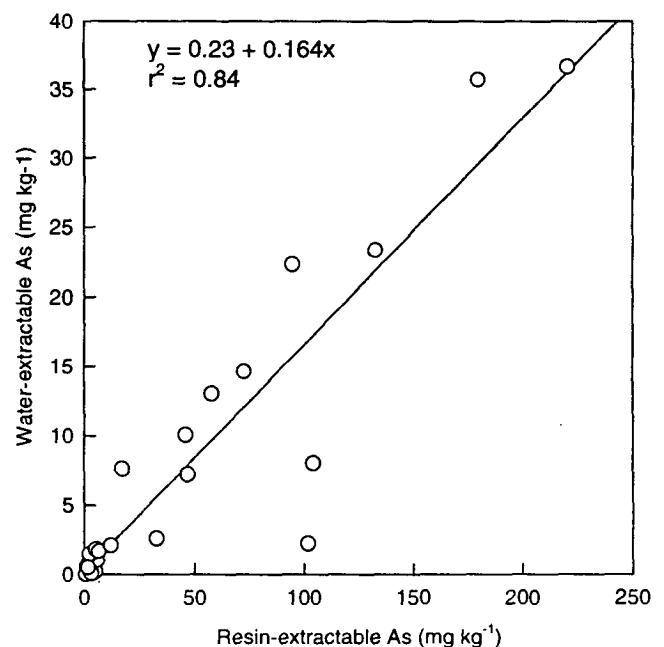
retention capacity than soil from most other sites. The relatively low proportions of resin and NaHCO<sub>3</sub>-extractable As at most sites does not mean that As toxicity or mobility is unlikely to be a problem. The high total As concentrations ensure that the labile forms of As are often well above the levels likely to cause phytotoxicity (Sheppard, 1992). In addition, the relatively high amounts of As in labile forms at depth in some soil profiles, suggests a significant potential for further leaching.

Resin-extractable As was highly correlated with water-soluble As (Fig. 1), with water-soluble As concentrations being approximately 15% of resin-extractable As concentrations (Table 2). Resin-extractable As may have value as an indicator of potential bioavailability and/or mobility of contaminant As in soil. The higher concentrations of As extracted by the resin compared to water can be a distinct advantage as far as analysis of As is concerned.

At most of the sites, the bulk of the As was present in the NaOH-extractable fractions with a mean 44.7% in the first extraction and a further 14.2% after sonication. The As in these fractions is likely to be associated with Fe and Al oxide materials present in the soils. Arsenic is known to be strongly sorbed by such materials (Huang, 1975; Anderson et al., 1975; Goldberg, 1986; Vaughan, 1993), and previous fractionation studies of As-contaminated soils have also found a predominance of Fe- and Al-bound forms of As (Jacobs et al., 1970; Woolson et al., 1971a; Woolson et al., 1973). Relatively small proportions of As were found in the HCl-extractable fraction overall (mean 6.9%). However, for some individual samples, up to 22% of the total As was present in this fraction (Table 3). The final fraction contained on average 21.1% of the total As (range 0-45%).

The relative difficulty in extracting this fraction of As, suggests that it should be considered to be extremely tightly bound and thus very unlikely to be bioavailable or mobile. It should be noted, that at some sites, for example Wagner (Table 3), the proportion of As associated with this fraction increases with soil depth. At such sites, it would appear that downward movement of mobile forms of As has been followed by transformation into relatively recalcitrant forms.

Further examination of the fractionation data showed



**Fig. 1. Relationship between water-extractable As and resin-extractable As.**

**Table 4. Linear correlation coefficients (*r*) between As concentrations in individual fractions and some soil properties.**

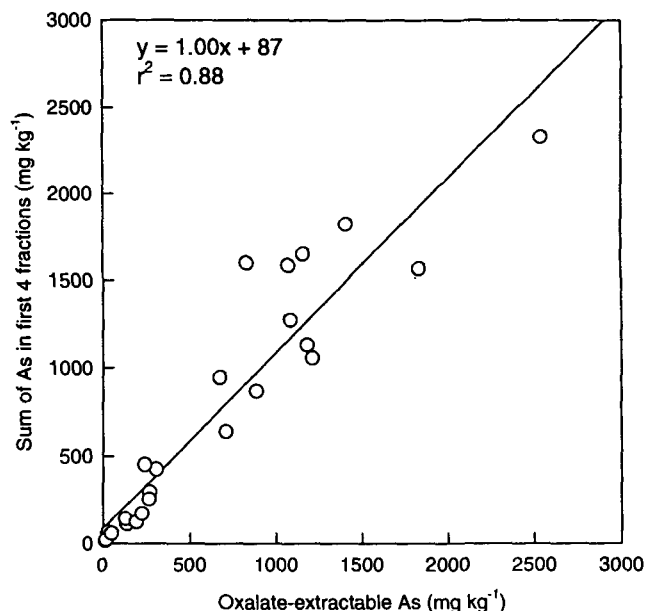
	NaHCO <sub>3</sub> -As	NaOH(1)-As	NaOH(2)-As	HCl-As	Residual As	Total As	Oxalate-Fe	Oxalate-Al	Soil pH	Organic C
Resin-As	0.85***	0.74***	0.56**	0.71***	0.56**	0.77***	0.65***	-0.03	0.66***	-0.18
NaHCO <sub>3</sub> -As		0.94***	0.79***	0.53**	0.75***	0.91***	0.61**	0.12	0.43*	-0.18
NaOH(1)-As			0.90***	0.51**	0.80***	0.96***	0.57**	0.17	0.28	-0.11
NaOH(2)-As				0.29	0.95***	0.94***	0.45*	0.30	0.17	0.02
HCl-As					0.16	0.52**	0.56**	-0.16	0.37	0.12
Residual As						0.91***	0.39	0.28	0.23	-0.03
Total As							0.59**	0.20	0.33	-0.03

\*, \*\*, \*\*\* *P* < 0.05, 0.01, 0.001.

that the amounts of As in the individual fractions were generally strongly correlated with each other and with total As concentrations (Table 4). (As a result of its anomalously high As concentration and the unusual properties of the Sandy Flat 40 to 45 cm sample, the data for this horizon was not included in the correlations.) Of all the fractions, HCl-extractable As showed the weakest correlations with other fractions. In spite of the large variation in the degree of As contamination between sites, total As and the amounts of As in individual fractions were significantly correlated with oxalate-extractable Fe (with the exception of residual As). There were no significant correlations with soil organic C or oxalate-extractable Al. The significant correlations between the two most labile fractions and soil pH (Table 4) may reflect reduced As sorption with increasing pH, and consequent increases in the more weakly bound forms of soil As. Multiple regression analysis indicated that a high proportion of the variation in resin-extractable As concentrations could be explained by the regression:

$$\text{Resin-extractable As} = 0.035 (\text{Total As}) + 50.4 (\text{pH}) - 289 \quad (r^2 = 0.78)$$

In addition to the As determination in the six fractions described above, the oxalate extracts used for soil oxide



**Fig. 2. Relationship between oxalate-extractable As and the sum of As concentrations in fractions 1 to 4.**

determinations were also analyzed for As, thus giving a direct measure of As associated with the extracted Fe and Al (Table 2). There was an extremely close correlation (almost 1:1) between oxalate-extractable As and the sum of the As in the first four fractions of the sequential extraction scheme (Fig. 2). The oxalate reagent used is considered to extract only short-range order (amorphous) materials, thus the As extracted can be assumed to be associated predominantly with these types of Fe (and Al) minerals. It seems likely that the As present in the first four fractions as determined by the sequential extraction scheme are not really completely separate pools, but represent a continuum of Fe- and Al-associated As with increasing strength of bonding with the mineral surfaces.

## CONCLUSIONS

The results obtained from this study demonstrate a considerable degree of soil contamination with As at the cattle dip sites examined. Although the actual areas of land contaminated are relatively small, restricted to the immediate area surrounding the cattle dip, the high level of contamination and downwards movement of As are significant causes for concern. Redevelopment of the sites for other uses will require the removal or remediation of substantial volumes of soil. The high concentrations of As in the most labile fractions, resin-extractable and NaHCO<sub>3</sub>-extractable, are of particular concern. The sequential fractionation scheme for As, based on the P fractionation, appears to be a useful method for investigating the chemical nature of contaminant As in soil. Resin-extractable As in particular might provide a good index of potential As bioavailability and mobility. The results from the study suggest that the bulk of the As present at the cattle dip sites is associated with the amorphous Fe (and Al) minerals in the soils.

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