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Hierarchical porous carbon with ordered straight micro-channels templated by continuous filament glass fiber arrays for high performance supercapacitors†

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Hierarchical porous carbon with highly ordered straight micro-channels was prepared through a facile melt vacuum infiltration method using continuous filament glass fiber arrays as the template and glucose as the precursor. The as-prepared carbon material shows high specific surface areas up to 1880 m² g⁻¹ profited from the unique structure of straight micro-channels. A fine pore structure is formed in the channel wall through KOH activation after the removal of the glass fiber array. Electrochemical evaluation of the carbon material indicates that the hierarchical porous carbon exhibits a high specific capacitance of 283 F g⁻¹ at a current density of 0.25 A g⁻¹ with an alkaline electrolyte (6 M KOH) in a three-electrode system. It also demonstrates excellent cycling stability with a capacity retention of 88.5% over 10 000 cycles at a high current density of 6 A g⁻¹. These exciting results demonstrate a very simple and low-cost method for large-scale preparation of electrode materials for supercapacitors.

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Introduction

In response to resource depletion and environmental pollution, researchers have been making enormous efforts to search for high-performance, low-cost, sustainable and environmentally friendly energy storage solutions.¹ Besides lithium-ion batteries, supercapacitors have attracted tremendous attention due to their high power density, excellent charge and discharge properties and efficient electrical energy delivery in heavy-duty applications.^{2–5} The functional mechanism for supercapacitors relies on two different phenomena: (1) double layer charging at the interface of the electrode/electrolyte, referred to as electrochemical double layer capacitance, and (2) reversible redox reactions at the electrode surface, known as pseudocapacitance.⁶ For pseudo-capacitors, the energy is stored through a reversible redox reaction of the electroactive species. Electrically conducting polymers⁷ and transition metal oxides (binary or ternary)^{8–12} with high energy density are the widely explored electrode materials for pseudo-capacitors. However, these materials display either poor stability or poor electrical properties, which

limit their practical application.¹³ The commercially available supercapacitors are electrochemical double layer capacitors (EDLCs) which are conventionally based on carbon materials, such as activated carbon,^{14–17} carbon nanotubes,¹⁸ graphenes,¹⁹ carbon aerogels²⁰ and carbide-derived carbons.²¹

To achieve high electrochemical performance of supercapacitors, research efforts have been focused on synthesis of nanostructured carbon materials with a high surface area, good chemical stability, appropriate wettability and high electronic conductivity.^{22–26} In order to enhance the porosity and specific surface area, the carbon materials are usually activated to introduce more active sites. In general, the activation of carbon materials can be performed in two ways: (1) physical activation with different oxidizing gases, such as air, O₂, CO₂, steam or their mixture, and (2) chemical activation with KOH, NaOH, H₃PO₄, ZnCl₂, or other chemicals at high temperature.¹ In fact, the performance of supercapacitors is not only related to the specific surface area, but also to the pore size and the pore distribution of the carbon materials. Previous research studies have showed that pores with size variation often exist simultaneously in active carbon, including macropores (larger than 50 nm), mesopores (2 to 50 nm) and micropores (smaller than 2 nm).²⁷ These pores normally have random size distribution and small pores that are often not permeable for electrolytes. As a result, only a part of the surface is effectively utilized and most of them do not contribute to the double-layer capacitance of the electrode.^{28–30} Accordingly, ordered mesoporous carbon with a well-defined pore size distribution and connected pore channels is considered to be a more effective electrode material for supercapacitors instead of

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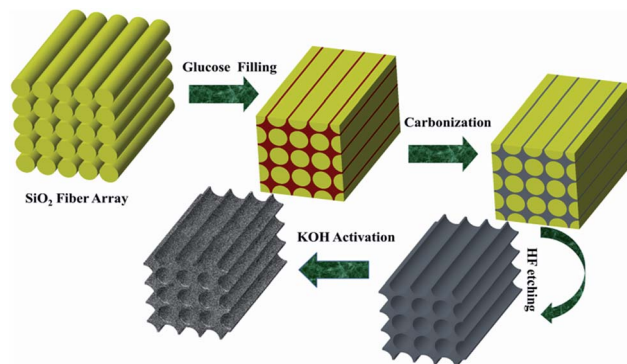
conventional active carbon.^{31–34} For instance, mesoporous SiO₂ templates are applied to synthesize structurally fine-tuned mesoporous carbon, such as CMK-3 derived from SiO₂ (SBA-15).³¹ Prepared through this way, the carbon materials with a high surface area often have micropores and mesopores, which are beneficial to electrolyte transport but unfavourable for retaining the electrolyte. I. Nongwe *et al.* applied SiO₂ spheres as templates for the preparation of macroporous carbon, which could be beneficial for holding more electrolyte.³⁵ However, these macropores that are formed through etching away the SiO₂ spheres are normally not continuously connected. The separated pores could be favourable for retaining the electrolyte but cannot provide enough continuous pathways for the electrolyte transport. Therefore, it is necessary to synthesize a carbon with continuous channels which can provide highways for the electrolyte transport. In addition, there are some disadvantages which cannot be ignored about those hard templates mentioned above, such as the complicated preparation steps and enormous time and energy consumption during the synthesis. Moreover, the high cost and complex process make the preparation of such porous carbon on a large-scale unrealistic and limit its further application. In order to accelerate the commercialization of supercapacitors for energy applications, it is necessary to develop simplified methods for fabrication of structurally ordered porous carbon and replace the expensive SiO₂ templates with low-cost substitutive materials.

Continuous filament glass fibers are inexpensive and promising alternative templates for further development of cost-effective functional carbon material manufacturing techniques. Normally, the diameter of continuous filament glass fibers is about 5–7 μm, which could be suitable for preparing templated hierarchical porous carbon with ordered straight micro-channels. Herein, we report a facile and novel melt vacuum infiltration method for the preparation of such hierarchical porous carbon materials, in which a bundled continuous filament glass fiber array is applied as the template and glucose is used as the precursor. The ordered straight micro-channels in the prepared carbonized material are favourable for the transportation and retaining of the electrolyte. Through KOH activation, micropores and small mesopores are formed on the wall of micro-channels, forming a hierarchical porous carbon material.^{25,36} Tested with a three-electrode configuration, the as-prepared carbon material demonstrates excellent gravimetric capacitance up to 283 F g⁻¹ in 6 M KOH electrolyte. It is remarkable that both the template (continuous filament glass fibers) and the precursor (glucose) are inexpensive commercialized products for many decades. Based on the facile and low-cost melt vacuum infiltration synthesis process, this work may bring a feasible solution for mass production of high-performance carbon-based electrode materials for supercapacitors in the near future.

Experimental section

Chemicals

All materials were purchased from commercial sources and used as received without further purification. The continuous filament glass fibers (88 g km⁻¹, 5–7 μm in diameter) were purchased from Tianjin Zhongtian Junda Glassfiber Pros Co.



Scheme 1 Schematic of the synthesis process of GTSDC-900 in this work.

Ltd. and D-glucose was purchased from Sinopharm Chemical Reagent Co. Ltd.

Material fabrication

The continuous filament glass fibers were cut into 5 cm in length and tightly bundled together to form a glass fiber array. Mixed with an appropriate amount of glucose powder, the fiber array was placed in a covered glass beaker and heated up to 180 °C in a vacuum drying chamber. In the process, molten glucose infiltrated into the gaps of the glass fiber array due to the capillary force. After two hours, the sample was taken out and cooled down to form a rigid solid hybrid precursor composite, which consisted of the glass fiber array and solidified glucose. The glucose on the glass fiber array was carbonized through a heat treatment at 800 °C for 3 hours with a heating rate of 5 °C min⁻¹ under N₂ gas flow. Then the samples were immersed in 20 wt% HF solution for 5 days to etch away the glass fibers completely to obtain a carbon material with ordered straight microchannels. For simplicity, this carbon material is abbreviated as GTSDC (glucose templated by SiO₂-fiber derived carbon) in this paper. The GTSDC samples were chemically activated in order to get enhanced electrochemical properties. The samples were immersed into 6 M KOH solution with a weight ratio of 1 : 3 (GTSDC : KOH) and dried at 100 °C to evaporate the water. Subsequently, the dried samples were thermally treated at activation temperature for 2 hours under N₂ gas flow with a heating rate of 5 °C min⁻¹ in a tubular furnace. The activation temperature was chosen as 700 °C, 800 °C, 900 °C and 1000 °C (corresponding samples were abbreviated as GTSDC-700, GTSDC-800, GTSDC-900 and GTSDC-1000, respectively). After cooling down to room temperature, the KOH residue was quickly removed with 1 M HCl solution, and then was washed repeatedly with deionized water until the pH value reached 7. Finally, the samples were dried at 80 °C in an ambient atmosphere for 12 hours. The preparation process of GTSDC-900 is illustrated in Scheme 1. As a comparison, two reference samples were also prepared. One reference sample was prepared by carbonizing the condensed glucose directly at 800 °C, abbreviated as GDC (glucose derived carbon). Another reference sample was GDC activated at 900 °C with the same KOH activation process used for GTSDC, abbreviated as GDC-900.

Material characterization

Scanning electron microscopy (SEM) images were obtained on a SEM-a HITACHI S-8020 at 5 kV for surface morphology investigation. High-resolution transmission electron microscopy (HRTEM) images were acquired from a JEOL JEM 2100 microscope with an operating voltage of 200 kV. N_2 adsorption-desorption analysis was carried out on a Micrometrics ASAP 2020 HD88 to evaluate the porosity and pore size distribution. The degas condition was set to 4 hours at 250 °C under a vacuum of 500 μmHg (~ 0.67 mbar), and all adsorption-desorption measurements were carried out at liquid nitrogen temperature. The X-ray diffraction (XRD) measurements were carried out on a Bruker D8 Advance powder X-ray diffractometer with Cu-K α radiation ($\lambda = 0.15406$ nm). The 2θ angular region was between 10° and 85°, and the scanning rate was set at 6° min^{-1} with a step of 0.02° for data acquisition. Raman spectra were acquired from a Dilor XY microspectrometer with 532 nm laser excitation.

Electrode fabrication and electrochemical measurements

The electrochemical behaviour of all carbon samples was investigated in a three-electrode cell configuration and a two-electrode cell configuration. In the three-electrode configuration, the prepared electrode with a size of 1 cm \times 1 cm was used as the working electrode, an Ag/AgCl electrode functioned as the reference electrode, and a platinum foil with a size of 2 cm \times 2 cm served as the counter electrode. The working electrode was prepared by mixing the as-prepared carbon material sample, carbon black and poly(vinylidene fluoride) (PVDF) with a weight ratio of 80 : 10 : 10. PVDF was previously dissolved in *N*-methyl pyrrolidinone to form 0.02 g mL^{-1} solution. After stirring for 12 hours, the mixture was coated on a nickel foil with a size of 1 cm \times 1 cm, and dried at 120 °C for 10 hours to remove NMP. Then the dried material was slightly pressed, and the mass loading of the active material on the nickel foil was weighed. The three-electrode system was tested in 6 M KOH electrolyte solution. The two-electrode fastener-type symmetric supercapacitor cells were assembled with a pair of stainless steel disks as the current collector and a piece of glassy fiber paper as the separator. A solution of 6 M KOH was used as the electrolyte.

The cyclic voltammetry (CV), electrochemical impedance spectroscopy (EIS) measurements (potential: 0 V, amplitude: 5 mV, frequency: 100 mHz to 100 KHz) and the galvanostatic charge-discharge measurements were performed by using a CHI 660 Electrochemical Workstation (CH Instruments, China). The capacitance from CV curves was calculated by the equations $C^* = \int I dt / mV$ for the three-electrode cells and $C = 4 \int I dt / mV$ for the two-electrode cells, where I is the constant current, V is the voltage window, and m is the mass of the active electrode material. The specific capacitance from galvanostatic charge-discharge was calculated by using the formulas $C^* = It / mV$ for the three-electrode cells and $C = 4It / mV$ for the two-electrode cells, where I is the constant current, V is the voltage window, m is the total mass for both carbon

electrodes, and t is the discharge time. For the two-electrode system, the energy density was calculated by the formulas $E = (CV^2)/2$ and $P = E/t$, where C is the specific capacitance from galvanostatic charge-discharge testing and t is the discharge time.

Results and discussion

The microstructure of the as-synthesized samples at each step is characterized by SEM. The continuous filament glass fibers with a uniform diameter about 5–7 μm have a smooth surface (Fig. S1a–d \dagger). Through the melt infiltration process under vacuum at 180 °C, the apertures of the bundled glass fiber array can be completely filled with glucose (Fig. S1e–h \dagger), and form a hybrid precursor of the glucose filled bundled bulk glass fiber array. After the carbonization of the glucose, the sample transformed into a hybrid composite of carbon filled bundled glass fibers without much morphological change (Fig. S1i–l \dagger). The glass fibers in the hybrid composite were removed through the subsequent chemical etching process, and the carbon material with uniform and straight ordered microchannels was obtained, as shown in Fig. S1m–p \dagger . It is obvious that the microchannels are well-aligned and parallel to each other forming a highly ordered array. From Fig. S1o and p \dagger we can find that the diameter of the channels in the porous carbon is approximately 5 μm , which is identical to the morphology and size of the continuous filament glass fiber template. The thickness of the channel wall is about 1 μm , which is determined by the gap between adjacent fibers. The surface of inner and outer channel walls is smooth and uniform without obvious pores or pits. Fig. S2 \dagger shows the images of samples at different synthesis steps, *i.e.* the precursor composite of glucose impregnated bundled glass fibers (Fig. S2a \dagger), hybrid material of carbon filled bundled glass fibers (Fig. S2b \dagger), and porous carbon with ordered channels (GTSDC) (Fig. S2c \dagger). As seen in Fig. S2 \dagger the sample always maintains its cylindrical macroscopic shape during the preparation process.

SEM images of the KOH activated carbon material (GTSDC) at 700–1000 °C are shown in Fig. 1. Fig. 1a is the top view of the sample of GTSDC-700. After activation, the microchannels of the carbon material keep the ordered characteristics and a similar diameter without collapse or distortion (Fig. 1a and b). In the images with higher resolution (Fig. 1c and d), some pits on the surface of the channel wall are found in the area with defects. When the temperature increases to 800 (Fig. 1g and h), 900 (Fig. 1k and l) and 1000 °C (Fig. 1o and p), the surface of the channel wall becomes more rough, and is full of open and interconnected pores. However, it is difficult to observe the micropores distributed on the channel wall due to the resolution limitation of the SEM. Additionally, fractures are observed on the channels of the samples activated at high temperature (Fig. 1e and f for GTSDC-800, i and j for GTSDC-900, m and n for GTSDC-1000). It can be deduced that the reaction between carbon and KOH may become more intensive as the activation temperature increases. As expected, hierarchical structured carbon materials with a high specific surface area could be obtained through the filament glass fiber templated KOH

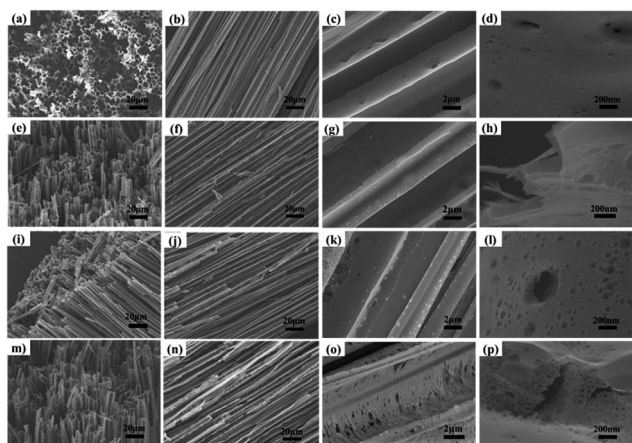


Fig. 1 SEM images of GTSDC activated at different temperatures: GTSDC-700 (a–d), GTSDC-800 (e–h), GTSDC-900 (i–l), and GTSDC-1000 (m–p).

activation process. The highly ordered straight microchannels are derived from the glass fiber array, and fine pores formed on the wall of the micro-channels originate from the KOH activation process. As a comparison, the morphology of carbon derived from the glucose direct carbonization process (GDC) and then activated by KOH at 900 °C (GDC-900) was also characterized by SEM (Fig. S3†). It is found that the surface of GDC is rough and full of pores after KOH activation (GDC-900), which is similar to the morphology of the surface on the channel wall of GTSDC-900. However, no ordered micro-channels can be found without using a glass fiber template.

The structure and morphology of crushed powders of all four carbon samples activated at different temperatures were further investigated by high-resolution transmission electron microscopy (HRTEM). The TEM image of GTSDC-700 with lower resolution shows that the sample has a uniform morphology (Fig. 2a). From the HRTEM micrograph (Fig. 2b), it can be observed that numerous nanoparticles about 2–3 nm in size homogeneously distribute in the crushed amorphous carbon. When the activation temperature increases to 800 °C (Fig. 2c and d), the size and distribution of the micropores in GTSDC-800 have no obvious difference compared with those of the sample activated at 700 °C. With further increase of the activation temperature to 900 °C (Fig. 2e and f), the pore size of

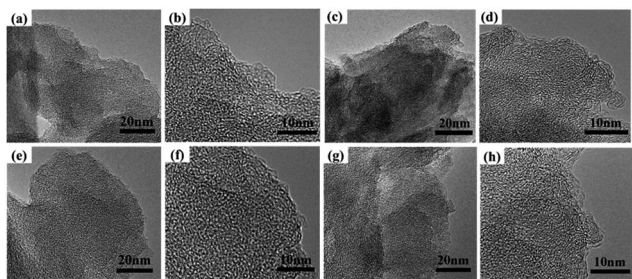


Fig. 2 HRTEM images of GTSDC activated at different temperatures: GTSDC-700 (a and b), GTSDC-800 (c and d), GTSDC-900 (e and f), and GTSDC-1000 (g and h).

GTSDC-900 also does not change obviously, but most of the pores are connected together. As the activation temperature increases to 1000 °C, the sample becomes more transparent under irradiation of electrons, and some lattice fringes can be observed in HRTEM images (Fig. 2g and h), indicating the formation of graphene at this temperature. The following XRD and Raman analysis could reach the same conclusion that higher temperature can lead to higher graphitization, which is also supported by the appearance of a 2D peak of GTSDC-1000 in Raman spectra. The HRTEM image of GTSDC without KOH activation is shown in Fig. S4† for comparison. It is mostly characterized by disordered carbon and few crystalline areas which is similar to GTSDC-700 and GTSDC-800.

To assess the phase transformation of the samples during the preparation process, X-ray diffraction (XRD) patterns of all samples are shown in Fig. 3a. The characteristic peaks located at 2θ of 25° can be assigned to the (002) diffraction of graphitic carbon, and the peak at about 43° corresponds to (101) diffraction of solid carbon materials.^{37,38} Particularly, for GDC and GTSDC carbon materials without KOH activation, the broad and low intensity peaks around 25° suggest the poor graphitic crystallinity. After activation by KOH, the peaks with higher intensity are observed for GDC-900 and GTSDC-900, indicating that the degree of graphitization enhances by KOH activation at higher temperature. In the XRD patterns of GTSDC-700, GTSDC-800, GTSDC-900 and GTSDC-1000, it can be found that the intensity of the peak at a 2θ position of around 25° gradually increases with increasing activation temperature, indicating that the graphitization improves.

To further characterize the graphitization of the samples, Raman spectra of all the carbon materials before and after activation were recorded. Fig. 3b depicts the Raman spectra of all the samples. All the spectra show very similar patterns with two broad peaks at about 1580 cm^{-1} and 1350 cm^{-1} . The peak at 1580 cm^{-1} (G-band) is attributed to the vibration of sp^2 -bonded carbon atom in a 2D hexagonal lattice, namely, the

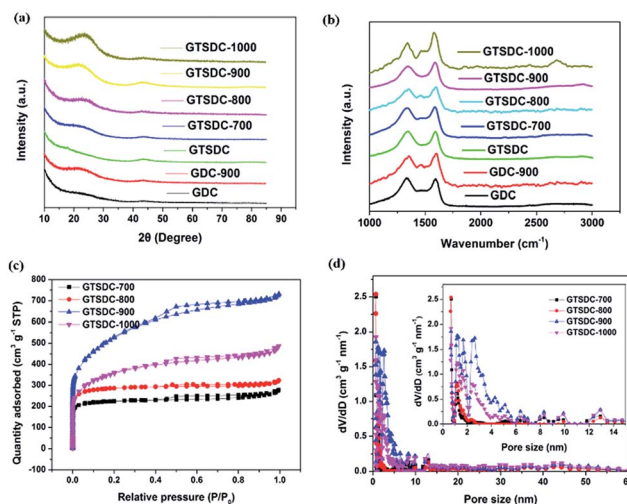


Fig. 3 The XRD patterns (a), Raman spectra (b), nitrogen adsorption–desorption isotherms (c) and DFT pore-size distribution curves (d) of carbon samples.

Table 1 The intensity ratio of the D-band to the G-band and surface area and pore structure of the as-prepared carbon materials

Samples	I_D/I_G^a	S_{BET}^b [$\text{cm}^2 \text{g}^{-1}$]	Micropore volume [$\text{cm}^3 \text{g}^{-1}$]	Mesopore volume [$\text{cm}^3 \text{g}^{-1}$]	Total pore volume ^c [$\text{cm}^3 \text{g}^{-1}$]
GTSDC-700	0.97	759	0.36	0.12	0.43
GTSDC-800	0.91	971	0.29	0.13	0.50
GTSDC-900	0.88	1880	0.18	0.87	1.13
GTSDC-1000	0.70	1259	0.20	0.50	0.75

^a I_G and I_D are the integral intensities of the D and G band. ^b Specific surface area calculated by the Brunauer–Emmett–Teller (BET) equation. ^c The total pore volume was determined from the amount of nitrogen adsorbed at a relative pressure of 0.99.

stretching modes of C=C bonds of typical graphite, while the peak at 1350 cm^{-1} (D-band) is associated with vibration of carbon atoms with dangling bonds in plane terminations of the disordered graphite and is related to the defects and disorders of the structures in carbon material.^{39,40} The relative intensities of these two lines (I_D/I_G) depend on the type of graphitic material and reflect the degree of graphitization. The intensity ratios of the D and G bands for all the carbon samples are listed in Table 1. For GDC and GTSDC, the intensity of the D band is almost as high as the G band, and the I_D/I_G ratio is about 1.03 and 1.01, which indicate the poor graphitization concentration. On the other hand, the similar Raman bands for GDC and GTSDC show that the continuous filament glass fiber array template has negligible impact on the crystal structures of carbon. For GDC-900 and GTSDC-900, the lower intensity at 1350 cm^{-1} (D band) and the peak with high intensity at 1580 cm^{-1} (G band) were observed, reflecting a higher graphitization concentration after activation by KOH at $900 \text{ }^\circ\text{C}$. As shown in Table 1, the I_D/I_G ratio of GDC-900 and GTSDC-900 decreases to 0.91 and 0.88, respectively. As the activation temperature increases from 700 to $1000 \text{ }^\circ\text{C}$, the I_D/I_G ratio of GTSDC-700, GTSDC-800, GTSDC-900 and GTSDC-1000 decreases to 0.97, 0.91, 0.88 and 0.70, respectively. This indicates that higher temperature can enhance the degree of graphitization in the carbon materials. Impressively, the GTSDC-1000 not only shows the D and G band, but also displays the 2D peak at about 2700 cm^{-1} , which is believed to be the characteristic peak of graphene.⁴¹ It can be expected that disordered carbon can provide more active sites for charge storage, and the graphene framework can improve the conductivity. The optimal carbon based supercapacitor electrode should keep a balance of the ratio of disordered carbon and ordered graphitized carbon.

The pore-size characterization of GTSDC activated at different temperatures was carried out by measuring the nitrogen adsorption/desorption isotherms. From the isotherms (Fig. 3c), all of them show a type I curve at relative pressure $P/P_0 < 0.90$ according to IUPAC classification, demonstrating adsorption in micropores.⁴² The steep increase in the adsorbed volume at very low relative pressure ($P/P_0 < 0.05$) results from the abundant micropores in solids.⁴³ For GTSDC-900 and GTSDC-1000, the curve exhibits a small rise at $0.5\text{--}0.9 P/P_0$, showing the existence of mesopores. This result can also be demonstrated by the pore size distribution (PSD) curves as shown in Fig. 3d. As for samples GTSDC-900 and GTSDC-1000, the PSD curves

exhibit obvious peaks at about 0.7, 1.2, 1.6, 3, 13 and 44 nm. It indicates that these two samples have both mesopores with a size of 3, 14 and 44 nm and ultramicropores with a size of 0.7 nm and other micropores (1.2 and 1.6 nm).⁴⁴ Compared with GTSDC-900, sample GTSDC-1000 has more micropores than mesopores. As for samples GTSDC-700 and GTSDC-800, the most obvious peaks of PSD curves lie in 0.7 and 1.2 nm, indicating that these two carbon samples have ultramicropores with a size of about 0.7 nm and other micropores (1.2 nm) but with very few mesopores. Therefore, sample GTSDC-900 has the ideal size distribution compared with the three other samples. It is well known that ideal electrochemical capacitor materials should have a hierarchical porous structure containing macropores for ion-buffering, mesopores for ion transport and micropores for enhancement of charge storage.⁴⁵ It could be predicated that GTSDC-900 should possess the best electrochemical performance among the prepared samples. In addition, the specific surface area and pore volumes of the carbon materials are summarised in Table 1. As the activation temperature varies from 700 , 800 to $900 \text{ }^\circ\text{C}$, the corresponding specific surface area enhances from 759.1 , 971.2 to $1880 \text{ m}^2 \text{g}^{-1}$. With further increase of the activation temperature to $1000 \text{ }^\circ\text{C}$, the specific area of the sample GTSDC-1000 drops to $1259 \text{ m}^2 \text{g}^{-1}$. It can be seen that the PSD curve of GTSDC-1000 is much lower than that of GTSDC-900 in the range of 1–5 nm, indicating that a large number of pores with the size of 1–5 nm disappear when the temperature increases from 900 to $1000 \text{ }^\circ\text{C}$. Therefore, we believed that at $1000 \text{ }^\circ\text{C}$, too much carbon reacted with KOH and led to part of small pores collapsing, hence resulting in the decrease of the surface area.⁴⁴ Therefore, the optimum activation temperature for GTSDC is $900 \text{ }^\circ\text{C}$, and GTSDC-900 not only possesses the highest specific surface area, but also the highest total pore volume ($1.13 \text{ m}^2 \text{g}^{-1}$), which makes it to be the potential electrodes with excellent electrochemical performance.

The performance of the supercapacitors based on all the carbon material samples was measured in 6 M KOH electrolyte using a three-electrode configuration. Cyclic voltammetry (CV) is utilized to estimate the electrochemical behaviour of all the carbon materials, and the results are depicted in Fig. 4. The CV curves of GDC, GDC-900, GTSDC and GTSDC-900 are acquired at a scan rate of 100 mV s^{-1} and compared in Fig. 4a. Small area and highly distorted CV curves of GDC and GTSDC indicate poor electrochemical performances. After KOH activation at $900 \text{ }^\circ\text{C}$, the area of CV curves of GDC-900 and GTSDC-900 is

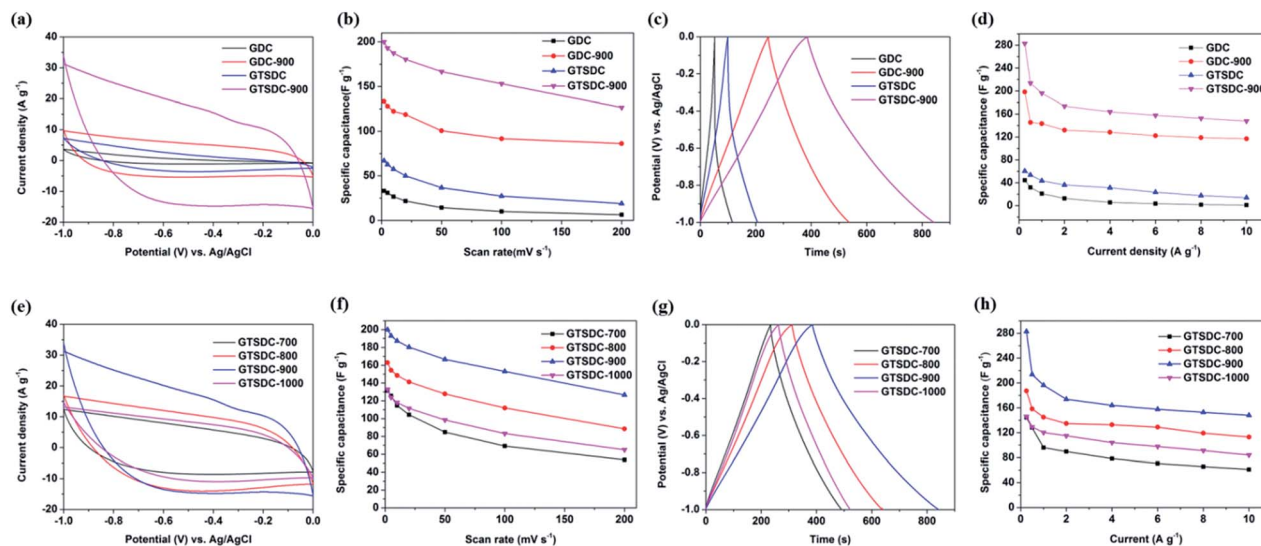


Fig. 4 CV curves of all carbon electrodes at 100 mV s^{-1} rates (a and e); specific capacitance of carbon electrodes at various scan rates (b and f); galvanostatic charge–discharge curves under 0.5 A g^{-1} (c and g); specific capacitance of carbon electrodes at different current densities (d and h).

much larger than that of GDC and GTSDC, respectively, and the CV curves display a nearly rectangular shape, which indicates the ideal capacitance performance and high energy storage efficiency due to the KOH activation. It is noteworthy that current density of the GTSDC and GTSDC-900 is much higher than that of GDC and GDC-900, respectively. The results demonstrate that the GTSDC and GTSDC-900 have a higher charge storage capacity than GDC and GDC-900. Considering the completely different structure between GTSDC-900 and GDC-900, as well as between GTSDC and GDC, it is deduced that the highly ordered straight channels existing in GTSDC and GTSDC-900 would be beneficial for the electrolyte osmosis and its diffusion in the whole electrode material. Fig. 4b shows the gravimetric specific capacitance of the four carbon samples, as a result of CV testing with the different scan rates in the range of 2 to 200 mV s^{-1} . According to the results, the values of the specific capacitance of unactivated carbon GDC and GTSDC are 33.2 and 66.9 F g^{-1} at a low scan rate of 2 mV s^{-1} , while only 6.4 and 18.9 F g^{-1} of the capacitance remain when the scan rate increases to 200 mV s^{-1} , indicating that the GDC and GTSDC not only have poor electrochemical performance, but also exhibit undesirable rate capability. For GDC-900 and GTSDC-900, the specific capacitances at a scan rate of 2 mV s^{-1} are 200 and 133.5 F g^{-1} , respectively, and keep 126.6 and 86.2 F g^{-1} at 200 mV s^{-1} . It indicates that the KOH activation should have improved the rate capability. In order to further investigate the electrochemical behaviour of the four carbon specimens, the galvanostatic charge–discharge curves at a current density of 0.5 A g^{-1} are recorded, and shown in Fig. 4c. The electrodes of GDC and GTSDC display distorted triangular shapes, which indicate low coulombic efficiency. Obviously, the electrodes of GDC-900 and GTSDC-900 present a nearly isosceles triangular shape with good symmetry, illustrating high coulombic efficiency. The calculated specific capacitances of GDC and GTSDC are only 31.9 and 53.6 F g^{-1} , while higher values of 145.6 and 213.8 F g^{-1} are observed for GDC-900 and GTSDC-900,

respectively. It indicates that the straight ordered channels derived from glass fibers and KOH activation can significantly improve the capacitive performance of the as-prepared carbon material. Fig. 4d summarizes the gravimetric specific capacitance of four samples as a result of galvanostatic charge–discharge curves. For instance, notably, the specific capacitance value of 283 F g^{-1} at a current density of 0.25 A g^{-1} is observed, which is much higher than that of the electrode made from other carbon materials as reported in previous literature, such as carbons derived from human hair (about 264 F g^{-1}),⁴⁷ carbon xerogels obtained from glucose–graphene oxide hybrids (143 – 239 F g^{-1}),⁴⁸ carbons derived from carbon dioxide activated glucose (158 F g^{-1}),⁴⁹ and carbon and graphene composites (150 – 210 F g^{-1}).^{50,51} In addition, it is found that the specific capacitances of GTSDC-900 and GTSDC are always higher than those of GDC-900 and GDC at various current densities in the range of 0.25 to 10 A g^{-1} . The results obtained from the galvanostatic charge–discharge curve are in complete accordance with the results obtained from CV testing. The effect of KOH activation temperature on the capacitance performance is also studied in detail. Fig. 4e displays the CV curves of the GTSDC activated at different temperatures at a scan rate of 100 mV s^{-1} . All four samples, *i.e.* GTSDC-700, GTSDC-800, GTSDC-900 and GTSDC-1000, exhibit a quasi-rectangular shape, which indicates excellent electrochemical reversibility. The current density of KOH activated GTSDC increases with increasing activation temperature from 700 to $900 \text{ }^\circ\text{C}$, and drops at $1000 \text{ }^\circ\text{C}$. The specific capacitances are shown in Fig. 4f. Activated with KOH at temperatures of 700 and $1000 \text{ }^\circ\text{C}$, the specific capacitances of GTSDC-700 and GTSDC-1000 are only 131 and 132.9 F g^{-1} at the scan rate of 2 mV s^{-1} . By contrast, the GTSDC-900 possesses a much higher specific capacitance of 200 F g^{-1} , suggesting that the suitable activation temperature is critical for the improvement of specific capacitance. As the scan rate increases to 100 mV s^{-1} , the capacitance of GTSDC-900 decreases to 153.1 F g^{-1} . The same phenomenon occurs for other carbon samples

activated at different temperatures, *i.e.* GTSDC-700, GTSDC-800 and GTSDC-1000. The specific capacitance decreased in an acceptable level with the increase of scan rate, and it suggests good rate capability for all the activated carbon materials. The galvanostatic charge–discharge curves at a current density of 0.5 A g^{-1} and the specific capacitances calculated from the galvanostatic charge–discharge curves at different current densities are shown in Fig. 4g and h, respectively. The obtained specific capacitance values are 128.1, 162.5, 213.8, and 129.3 F g^{-1} at a current density of 0.5 A g^{-1} for GTSDC-700, GTSDC-800, GTSDC-900, and GTSDC-1000, respectively. Obviously, the GTSDC-900 shows the best specific capacitive performance compared to the samples activated at other temperatures, which is consistent with the result calculated from the CV curve. The supercapacitive performance of GTSDC-900 is further investigated, and Fig. S4† shows the CV curves at different scan rates and galvanostatic charge–discharge curves at different

current densities. With the scan rate varying from 2 to 200 mV s^{-1} , CV curves of GTSDC-900 always keep the rectangular shape, demonstrating a nearly ideal capacitive behaviour. At the same time, the charge and discharge part of GTSDC-900' galvanostatic charge–discharge curves is symmetric at different current densities ranging from 0.25 to 10 A g^{-1} . In general, the electrochemical performance of all porous carbon samples was tested by CV at different scan rates and galvanostatic charge/discharge at different current densities. Therefore, the specific capacitance calculated from CV curves and galvanostatic charge–discharge curves is listed in Tables 1S and 2S,† respectively.

The performance of the supercapacitor based on the GTSDC-900 sample was measured by using a two-electrode symmetrical system in 6 M KOH aqueous electrolyte. Fig. 5a shows the CV curves at different scan rates ranging from 2 to 1000 mV s^{-1} . The rectangular shape of the CV is observed even

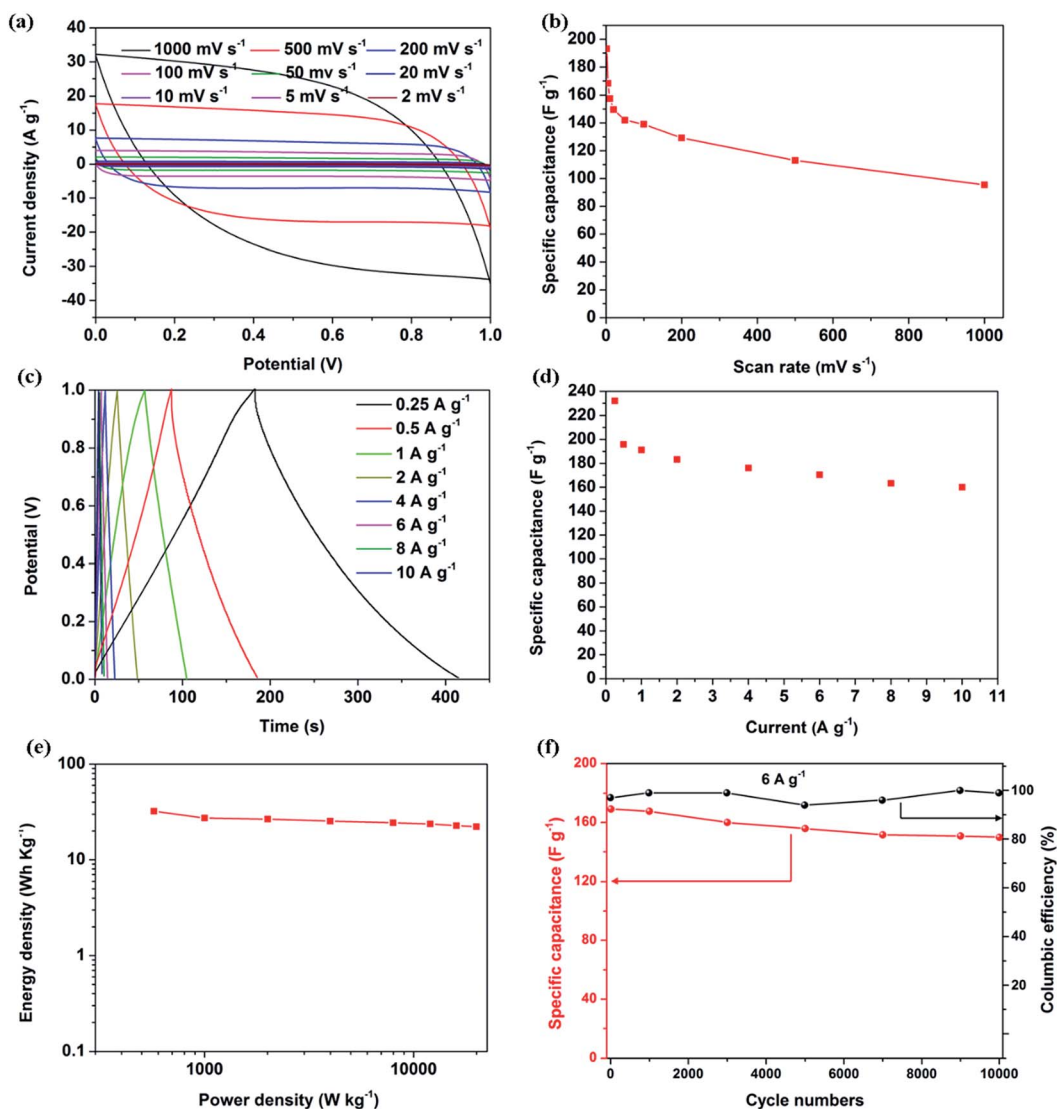


Fig. 5 CV curves of the GTSDC-900 symmetric supercapacitor (a) and calculated specific capacitance at various scan rates (b); galvanostatic charge/discharge curves of the symmetric supercapacitor (c) and calculated specific capacitance at various current densities (d); Ragone plots (e); specific capacitance (red) and coulombic efficiency (black) of GTSDC-900 symmetric supercapacitor at 6 A g^{-1} over 10 000 cycles (f).

at 500 mV s^{-1} , indicating a very fast and efficient charge transfer.⁵⁰ As shown in Fig. 5b, the highest specific capacitance value reaches 193.2 F g^{-1} at 2 mV s^{-1} , and the value remains at 113 F g^{-1} as the scan rate increases to 200 mV s^{-1} , suggesting good rate capability of GTSGC-900. As displayed in Fig. 5c, the IR drop for this symmetric supercapacitor is relatively low, indicating that the two-electrode cell has low internal resistance due to fast charge transfer and ion diffusion. Calculated by the galvanostatic charge–discharge curves, the specific capacitances *versus* different current density are summarized in Fig. 5d. The specific capacitance is measured with a high value of 202.3 F g^{-1} at 0.25 A g^{-1} , and still maintains 160 F g^{-1} , even when the current density increases to 10 A g^{-1} , *i.e.* 79.2% capacitance retention as the current density increased by 40 times. The symmetrical supercapacitor made from GTSDC-900 exhibits excellent rate capability. Moreover, the power density and energy density of the symmetric cell are shown in a Ragone plot (Fig. 5e). The energy density can reach as high as 32.2 W h kg^{-1} at a power density of 0.5 kW kg^{-1} and 22.2 W h kg^{-1} at a power density of 20 kW kg^{-1} , which is higher than those of previously reported carbon symmetric supercapacitors in aqueous electrolytes, such as graphene-based carbon aerogels (20 to 27.4 W h kg^{-1}),⁵³ templated active carbon (2.7 to 4.2 W h kg^{-1}),⁵⁴ petroleum-coke-derived active carbon (8 to 10 W h kg^{-1}),⁵⁵ sunflower-seed-shell-derived active carbon (3 to 6 W h kg^{-1}),⁵⁶ mesophase-based active carbon (0.45 to 2 W h kg^{-1}),⁵⁷ cherry-stone-derived active carbon (0.9 to 2 W h kg^{-1}),⁵⁸ waste-coffee-ground-derived active carbon (5 to 40 W h kg^{-1})⁵⁹ and polyfurfuryl-alcohol-derived active carbon (32 W h kg^{-1}).⁶⁰ The energy density of our device compared to the values of other reported devices is shown in Fig. S5† in detail.

In order to assess the cycling stability of GTSDC-900 for supercapacitor applications, the galvanostatic charge–discharge was conducted at a constant current density of 6 A g^{-1} for 10 000 cycles. Fig. 5f shows that the capacitance retention can reach 88.5% with a moderate decrease from 170.4 to 150.8 F g^{-1} , indicating the excellent cycling stability of GTSDC-900 even at high current density. Additionally, the coulombic efficiency is about 97% (10th cycle), 99% (1000th cycle), 99% (3000th cycle), 94% (5000th cycle), 96% (7000th cycle), 100% (9000th cycle), and 99% (10 000th cycle), respectively. Therefore, the symmetric supercapacitor always has nearly perfect coulombic efficiency over 10 000 cycles. Electrochemical impedance spectroscopy (EIS) was employed for further study of the electrochemical performance of the carbon specimens (Fig. 6). Generally, linear

curves in the low frequency region can indicate the capacitive performance, while the semicircles in the high frequency region represent the charge transfer resistance. Fig. 6a shows the impedance curves obtained from four different samples GDC, GDC-900, GTSDC and GTSDC-900 at different frequencies ranging from 0.01 to 100 kHz . Apparently, the GTSDC-900 possesses the steepest slope of line (the phase angle about 87°) and the smallest diameter of the semicircle (0.64Ω), indicating a nearly ideal capacitance and the smallest charge transfer resistance. As a comparison, the phase angle values of the other samples are much smaller than those of GTSDC-900. Additionally, GTSDC-900 and GDC-900 display a relatively low equivalent series resistance (ESR) of about 0.81Ω and 0.79Ω , which are lower than those of GTSDC (1.07Ω) and GDC (0.83Ω). These results can be explained by the disordered carbon content in the carbon specimens. When the glucose was only carbonized at 800°C , the main phase of the material is amorphous carbon. Thus the electrode made from GTSDC and GDC shows higher internal resistance. Part of amorphous carbon was etched away *via* the KOH activation process; therefore, the internal resistance of GTSDC-900 and GDC-900 are smaller than those of the unactivated carbon materials. The impedance curves of the samples activated at different temperatures can be seen in Fig. 5b. The equivalent series resistance decreases in the order of GTSDC-1000 (0.80) < GTSDC-900 (0.81) < GTSDC-800 (1.01) and GTSDC-700 (1.01). Hence, it is believed that with an increase of activation temperature, the relative content of graphene increases, leading to improved conductivity, which is consistent with the XRD and Raman results. Therefore, the highest electrochemical performance of GTSDC-900 should originate from: (1) hierarchical porous structure with well-defined pore distribution and ordered straight micro-channels; (2) large surface area and active sites derived from KOH activation; (3) improved conductivity related to the graphene content formed inside carbon materials.

Conclusions

In this work, a precursor made of glucose, templated by continuous filament glass fiber arrays, has been processed through a method combining melt vacuum infiltration, carbonization and KOH activation process. A unique hierarchical porous carbon with highly ordered straight micro-channels and well-defined porosity on the channel-wall has been synthesised. The carbon material shows an ultrahigh specific surface area of $1880 \text{ m}^2 \text{ g}^{-1}$ and exhibits an excellent specific capacitance of 283 F g^{-1} at 0.25 A g^{-1} . The symmetric supercapacitor employing two electrodes and the best performed carbon materials synthesized in this work displays high cycling stability with 88.5% capacitance retention over 10 000 cycles. Moreover, the energy densities of 32.2 W h kg^{-1} at a power density of 0.5 kW kg^{-1} and 22.2 W h kg^{-1} at a power density 20 kW kg^{-1} are also demonstrated. Compared with the reference specimen, outstanding electrochemical performance delivered by the synthesized carbon could be attributed to the unique microstructure with highly ordered straight micro-channels and well-defined porosity on the channel wall, which

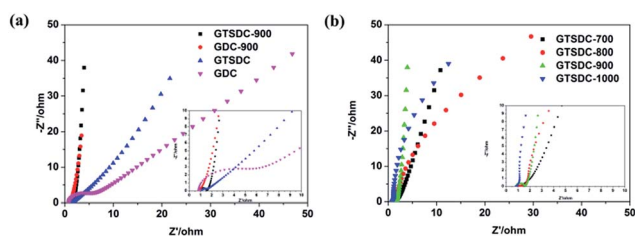


Fig. 6 The Nyquist plots of all the carbon specimens showing imaginary part *versus* real part of impedance in 6 M KOH aqueous solution. The inset magnifies its expanded high-frequency region.

is beneficial for retaining and transportation of the electrolyte. More importantly, both the glass fiber template and the glucose precursor are inexpensive industrial products. Therefore, the current work provides a promising route for mass production of high performance and low-cost supercapacitor materials.

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