

Immiscible silicate liquid partition coefficients: implications for crystal-melt element partitioning and basalt petrogenesis

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Received: 20 September 2005 / Accepted: 5 July 2006 / Published online: 10 August 2006
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Abstract This study investigates partitioning of elements between immiscible aluminosilicate and borosilicate liquids using three synthetic mixtures doped with 32 trace elements. In order to get a good spatial separation of immiscible liquids, we employed a high-temperature centrifuge. Experiments were performed at 1,050–1,150°C, 1 atm, in sealed Fe and Pt containers. Quenched products were analysed by electron microprobe and LA ICP-MS. Nernst partition coefficients (D 's) between the Fe-rich and Si-rich aluminosilicate immiscible liquids are the highest for Zn (3.3) and Fe (2.6) and the lowest for Rb and K (0.4–0.5). The plots of D values against ionic potential Z/r in all the compositions show a convex upward trend, which is typical also

for element partitioning between immiscible silicate and salt melts. The results bear upon the speciation and structural position of elements in multicomponent silicate liquids. The ferrobasalt–rhyolite liquid immiscibility is observed in evolved basaltic magmas, and may play an important role in large gabbroic intrusions, such as Skaergaard, and during the generation of unusual lavas, such as ferropicrites.

Introduction

The petrogenetic role of silicate liquid immiscibility has been debated since the start of modern igneous petrology. A popular hypothesis of magma differentiation in early years (Scrope 1825; Daly 1914), immiscibility fell into disrepute for almost half a century in the late 1920s (Greig 1927; Bowen 1928), and the interest in it was revived only in the 1950s and 1970s after the discoveries of low-temperature liquid immiscibility in the system leucite–fayalite–silica (Roedder 1951), and unequivocal evidence of immiscibility in groundmass glasses and melt inclusions of lunar and terrestrial basalts (Philpotts 1976, 1982; Roedder and Weiblen 1970, 1971). At that time, the geologic observations and experimental studies were summarised by Philpotts (1976, 1982) and Roedder (1979), and the main conclusion of petrogenetic relevance was that the most common magmas, such as tholeiitic and alkaline basalts, (and probably also some rhyolites, Naumov et al. 1993), may encounter liquid immiscibility at some, most likely advanced, stage of crystallisation. It seemed that silicate liquid immiscibility was about to regain its status as an important petrogenetic process. However,

Communicated by J. Hoefs

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very few studies have further examined the petrogenetic implications. Until now immiscibility remains virtually ignored in models of terrestrial magma evolution. In fact, this process is much more debated for lunar rocks (e.g., Longhi 1990; Shearer et al. 2001) than in the vast literature on terrestrial basalts and granites. In exceptional cases, immiscibility has been involved (but as a side issue) in the debates about magma evolution in the Skaergaard intrusion (McBirney and Nakamura 1974; McBirney and Naslund 1990); it was discussed in connection to the origin of silicic rocks in oceanic crust (Dixon and Rutherford 1979), and in recent years, several studies investigated possible cases of liquid immiscibility in xenoliths and rare types of alkaline plutonic rocks (Hurai et al. 1998; Markl 2001; Rajesh 2003; Sørensen et al. 2003).

Revisiting the chemical aspects of silicate liquid immiscibility is important and timely for several reasons. First, one can now exploit the advantages of modern experimental and analytical tools not available 20 or 30 years ago. For example, the introduction of centrifuge phase separation (Dorfman et al. 1996; Veksler et al. 2002; Schmidt et al. 2006) now enables the collection of small droplets of immiscible liquids into larger, analysable pools, and also distinguishes between stable and metastable immiscibility. The progress in analytical methods, and especially the employment of LA ICP-MS, has improved the detection limits and the precision of trace element analyses. Recent years have also seen a significant progress in theoretical models of silicate liquid immiscibility (e.g., Hudon and Baker 2002a, b; Schmidt et al. 2006, and references therein). Finally, new evidence of immiscible silicate liquids in melt inclusions from classical suites of magmatic rocks, such as the Skaergaard intrusion of layered gabbros, East Greenland (Jakobsen et al. 2005), have shed new light on the extent and petrogenetic importance of liquid immiscibility in tholeiitic and other common magmas. This evidence will hopefully revive the interest of igneous petrologists to the causes and geochemical consequences of silicate liquid immiscibility.

The experimental study presented in this paper aims at examining the regularities of element partitioning between multicomponent immiscible aluminosilicate and borosilicate liquids. The study constitutes a part of our long-term research programme of immiscibility between salt, silicate and aqueous liquids (Veksler et al. 1998, 2002) with implications to natural magmas (Veksler 2004). This paper deals solely with silicate liquid immiscibility, which means that both conjugate liquids are essentially silicate melts, and non-silicate acidic components are either insignificant, or absent.

Experimental and analytical methods

The establishment of liquid–liquid element partitioning and the determination of precise values of partition coefficients were the main objectives of our study. We wanted to cover several cases of low-temperature silicate liquid immiscibility, with an emphasis on relative variations in partitioning behaviour between the elements, and thus the strategy was to analyse as many elements as possible in products of a single run. Three synthetic starting mixtures were prepared (Table 1): two ferrobasaltic aluminosilicate compositions and a multicomponent borosilicate composition. The ferrobasaltic mixtures SF-1 and SF-2 (Table 1) were based on previously published analyses of synthetic and natural basaltic–rhyolitic immiscible liquid pairs (Longhi 1990; Philpotts 1982; McBirney and Nakamura 1974); the borosilicate composition SB was designed on the basis of phase diagrams of borosilicate ternaries featuring miscibility gaps (Levin et al. 1964), and our previous work on immiscible multicomponent borosilicate mixtures (Veksler et al. 2002).

Preparation of starting mixtures

The starting mixtures were synthesised by fusion of carefully weighed and mixed reagent-grade chemicals (oxides, carbonates and H_3BO_3) in Pt crucibles. In order to avoid strong oxidation and/or Fe losses from the ferrobasaltic mixtures, Fe-free glasses were first prepared by repeated fusions in an electric furnace at 900 and 1,350°C and then the glasses were crushed into powder and mixed with reagent-grade FeO (99.9%, manufactured by “Aldrich”). Finally, the mixtures

Table 1 Starting mixtures, wt.% oxides

	SF-1	SF-2	SB
SiO_2	56.5	54.5	40.0
TiO_2	2.0	4.8	0.1
Al_2O_3	6.0	7.3	2.0
FeO	24.0	20.0	0.1
MgO	1.4	4.3	3.0
CaO	6.0	7.6	3.0
Na_2O	0.5	1.0	1.0
K_2O	2.0	1.7	1.0
B_2O_3	0.0	0.0	40.0
SrO	0.0	0.0	3.0
BaO	0.0	0.0	3.0
Rb_2O	0.0	0.0	0.9
Cs_2O	0.0	0.0	0.9
Li_2O	0.0	0.0	0.8
Total	98.5	98.5	98.8

were spiked with trace elements (Li, Rb, Cs, Be, Sr, Ba, REE, Zr, Hf, Nb, Ta, Zn, Cu, Ga, In, Sn, Mo, and W), at about 200–300 ppm of each. Most of trace elements were taken for spiking in the form of chemical grade oxides, but hydroxides and concentrated hydroxide solutions were used for Li, Rb and Cs, and natural beryl was used as a spike for Be. The oxides were first mixed with each other, and certain amounts of oxide mixtures were then added and thoroughly mixed with the glass–FeO mixtures. Trace element oxide mixtures used for the spiking were not identical for all the starting compositions. Thus, the ferrobaltic mixtures SF-1 and SF-2 were not spiked with Sn; the borosilicate mixture SB was not spiked with Cu and Mo. The borosilicate synthetic glass SB contained major amounts of Li, Rb, Cs, Sr and Ba (thousands parts per million to weight percents) and was prepared by three subsequent fusions of the reactant mixture of chemicals at 600, 950 and 1,300°C. Before and after the fusions the glassy products were crushed into a fine powder in an agate mortar under acetone.

High-temperature centrifuge runs

High-temperature centrifuge phase separation has been successfully employed in our previous studies of liquid immiscibility (Veksler et al. 1998; 2002) and the same technique was used in this study. The experimental equipment has been described in detail by Dorfman et al. (1996). Briefly, the runs were performed in a small cylindrical wire-wound electric furnace mounted on a powerful centrifuge, which enabled centrifuge acceleration of up to 1,000g ($g = 9.8 \text{ m/s}^2$) and maximal run durations of 3.5 h. In case of highly viscous aluminosilicate and borosilicate liquids, like in this study, phase separation is usually incomplete (Veksler et al. 2002), but the centrifuge technique nevertheless greatly increases the size of the immiscible liquid droplets and thus assists the chemical analyses of run products.

Experimental run conditions are listed in Table 2, and the Fe–Pt containers used in the experiments with ferrobaltic mixtures are presented in Fig. 1. The inner Fe containers shown in Fig. 1 prevented Fe losses into Pt and served as an oxygen buffer. Experiments with the Fe-free borosilicate mixture SB were performed in single Pt containers of about the same size as those in Fig. 1. About 200 mg of fine-grained powders of the starting mixtures were loaded into the containers, and centrifuged at the conditions listed in Table 2. Unfortunately, the equipment does not allow run durations longer than a few hours. To approach the

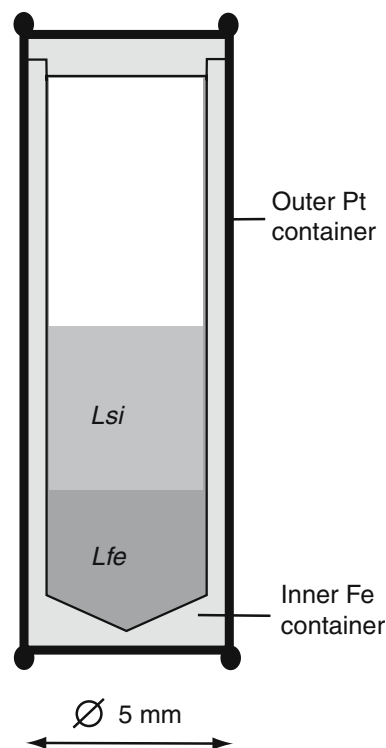


Fig. 1 A sketch of the container used in the experiments with ferrobaltic compositions. Inner iron container is inside an outer Pt container, which is welded shut. *Lsi* and *Lfe* are immiscible silica- and Fe-rich liquids

chemical equilibrium as close as possible, experimental charges were first heated without spatial separation of the phases at the minimal rotation rates for about 30 min, and then the phases were separated at the highest rotation rates corresponding to the acceleration of 1,000g (Table 2). Quenching was achieved in air by turning the electric furnace off, while the centrifuge rotation continued. Typically 2 min were required to cool the furnace and the charge down to room temperature.

Electron microprobe analyses

Run products were mounted in epoxy, ground, polished, examined optically in reflected light under microscope, and analysed using Cameca SX-50 electron microprobes (EMP) at the GFZ Potsdam and the Department of Mineralogy, Natural History Museum, London. Analyses were performed in WDS mode at 15 nA beam current and accelerating voltage of 15 kV. Counting time for all the other elements was set to 20 s on peak and 10 s on background. The following synthetic and natural standards were used for the calibration: orthoclase (Al and K), rutile (Ti), wollastonite

Table 2 Run conditions and phase composition of the products

Run no.	Starting mixture	<i>t</i> (°C)	Duration (h)	Acceleration (g)		Phases in run products
				Initial	Final	
C-36	SF-1	1,100	0.5 + 2.5 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr + ol + Fe–Ti
C-37	SF-1	1,075	0.5 + 2.5 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr + ol + Fe–Ti
C-38	SF-1	1,050	0.5 + 2.5 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr + ol + Fe–Ti
C-39	SF-1	1,125	0.5 + 2.5 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr
C-40	SF-1	1,150	0.5 + 2.5 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr
C-137	SF-2	1,160	1 + 2 ^a	160	1,000	<i>Lfe</i> + <i>Lsi</i> + tr
C-20	SB	1,100	1.5	–	1,000	<i>Lfe</i> + <i>Lsi</i>

Abbreviations for phases: *Lfe* ferrobaltic melt, *Lsi* silicic melt, *Lb* boron-rich melt, *tr* tridymite, *ol* olivine, Fe–Ti oxide solid solutions (wüstite and ulvöspinel)

^aSlow rotation at 160 g for half an hour followed by fast rotation at 1,000g for two and half hours

(Si and Ca), albite and jadeite (Na), diopside and periclase (Mg), celestine (Sr) and barite (Ba). Glasses were analysed with a defocused beam (beam diameter of 5–10 μm); 5–20 point analyses were performed for each phase to obtain statistically representative averages listed in Tables 2 and 3.

Laser ablation ICP-MS analyses

Trace element analyses by LA ICP-MS technique were performed using a Hewlett Packard HP4500 ICP-MS in combination with a New Wave UP-213 laser ablation system in the Centre for Ore Deposit Research, University of Tasmania. For this study, analyses were performed in He atmosphere by ablating ~60-μm diameter spots at a rate of 5 shots/s using a laser power of ~12 J/cm². The instrument was optimised for sensitivity on mid- to high-mass isotopes (in the range 80–240 a.m.u.) and for minimal molecular oxide species (i.e., ²³²Th/¹⁶O/²³²Th < 0.3%) and doubly charged ion species (i.e., ¹⁴⁰Ce²⁺/¹⁴⁰Ce⁺ < 0.4%) production. The low level of molecular oxide and doubly charged ion

production precludes the need to correct any of the analyte signal intensities for interfering species. Each analysis was performed in the time-resolved mode which involves sequential peak hopping through the mass spectrum, with the accumulation of ~130 measurements of ~20 ms duration (2.6 s total) for each of the 31 analyte isotopes. The analysis time for each sample was 90 s, comprising a 30-s measurement of background (laser off) and a 60 s analysis with laser-on. Instrument calibration was performed by ablating the NIST612 glass standard (Pearce et al. 1997). Data reduction was undertaken according to standard methods (Longerich et al. 1996) using the NIST612 glass as a primary reference material and Ca as the internal standard. All analyses were normalised to the USGS BCR-2g glass, which was used as a secondary reference material throughout the analytical sessions. For further control, Ti, Fe and Mg in ferrobaltic samples, and Si and Mg in borosilicate glasses were analysed by both LA ICP-MS and electron microprobe, and the agreement between the two methods was better than ±10% of measured values for all the samples.

Table 3 Electron microprobe analyses of crystal phases in wt. % oxides

Sample phase	C-36			C-37			C-38		
	ol	wu	usp	ol	wu	usp	ol	wu	usp
SiO ₂	31.02	0.46	0.11	31.19	0.42	0.12	30.69	0.36	0.05
TiO ₂	0.19	4.52	25.54	0.28	2.95	22.18	0.27	2.57	23.26
Al ₂ O ₃	0.02	0.78	4.79	0.05	0.57	4.08	0.03	0.50	4.55
MgO	5.40	0.70	0.70	5.57	0.68	0.61	4.88	0.62	0.58
FeO	62.47	91.04	65.89	62.11	93.90	69.45	62.64	93.64	67.73
CaO	0.65	0.05	0.07	0.82	0.05	0.08	0.86	0.04	0.09
Na ₂ O	0.01	0.01	0.01	0.01	0.01	0.02	0.01	0.01	0.00
K ₂ O	0.02	0.03	0.01	0.03	0.01	0.02	0.03	0.02	0.00
Total	99.77	97.59	97.12	100.05	98.59	96.55	99.40	97.76	96.26

Abbreviations for phases: *ol* olivine, *wu* wüstite, *usp* ulvöspinel

Results

Phase composition and textures of run products

All the experimental samples listed in Table 2 contain two distinct immiscible silicate liquids quenched to glasses, and in most cases also crystalline liquidus phases. Experiments with the ferrobaltic starting compositions SF-1 and SF-2 produced a dense Fe-rich liquid *Lfe*, a light silica-rich liquid *Lsi*, fayalitic olivine (Table 3), tridymite and Fe–Ti oxides. The borosilicate composition produced a lighter silica-rich liquid *Lsi*, denser B-rich liquid *Lb* and no crystal phases.

Typical textures of run products of the ferrobaltic compositions are shown in Fig. 2. As evident from the textures, run durations were insufficient for complete spatial separation of the immiscible liquid phases. The starting composition SF-1 produced approximately equal volumes of the immiscible liquids, but because of the high viscosity of *Lsi* numerous irregular droplets of *Lfe* remained trapped within the *Lsi* layer after 3 h of centrifuge separation (Fig. 2a). The second starting composition SF-2 produced very little *Lsi*, no more than 5–10 vol.%, and after the separation, the silica-rich liquid *Lsi* accumulated at the very top of the sample forming relatively large, homogenous droplets enclosing tridymite crystals, but also diffuse streaks resembling plumes of smoke in air or flows of a dye in water (Fig. 2b). The interfaces between the *Lfe* and *Lsi* phases of the ferrobaltic samples appear blurred and never as sharp as liquid–liquid interfaces in borosilicate samples. In all the samples one could nevertheless find areas of homogenous conjugate glasses large enough for electron microprobe and LA ICP-MS analyses. The bottom layer of the *Lfe* quenched to brownish glass, which looked homogenous (although somewhat turbid) optically, and in backscattered electron (BSE) images. Detailed microprobe analyses of the *Lfe* layers revealed however significant vertical gradients in some samples. The strongest variations were observed for SiO₂ and FeO, which showed complimentary trends with silica content increasing, and FeO concentration decreasing towards the top of the *Lfe* layer (Fig. 3). The gradients might be due to incomplete phase separation of sub-micron liquid emulsions and/or diffusion in homogenous liquid phase, but most likely they formed during quench, which was apparently slow enough for the secondary vertical stratification of the low-viscosity *Lfe* before it quenched to glass. Crystals of Fe–Ti oxides (presumably wüstite and ulvöspinel) and olivine sank to the very bottom of the container (Fig. 2c), while

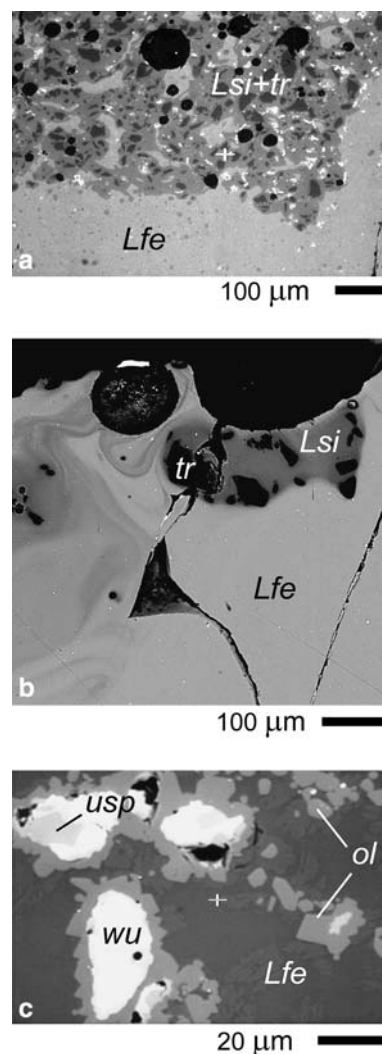


Fig. 2 Backscattered electron images of run products. **a** Interface between the iron-rich (*Lfe*) and silica-rich (*Lsi*) immiscible liquids in the sample C-38. The upper layer of the silica-rich liquid contains numerous crystals of tridymite (*tr*, dark grey), and unseparated droplets of the iron-rich liquid (*Lfe*, light grey) **b** Upper part of the sample C-137 with the droplets and plumes of the silica-rich liquid and tridymite crystals (*tr*). **c** Olivine and Fe–Ti oxide crystals in the iron-rich liquid at the bottom of the sample C-38. White areas of the oxide crystals are wüstite (*wu*); light grey zones are composed of ulvöspinel (*usp*)

tridymite grains all floated up surrounded by *Lsi* (Fig. 2a, b). Olivine typically formed euhedral crystals dispersed in *Lfe*, and also corona overgrowths around oxide grains. Tiny spherical metal particles, no more than 1–2 μm in size, were found dispersed in the glasses, and larger oval droplets of pure Fe were found in some samples, in contact with wüstite crystals and along the container walls. Some of the metal may have exsolved during quench or resulted from redox reaction between the silicate melts and the Fe container.

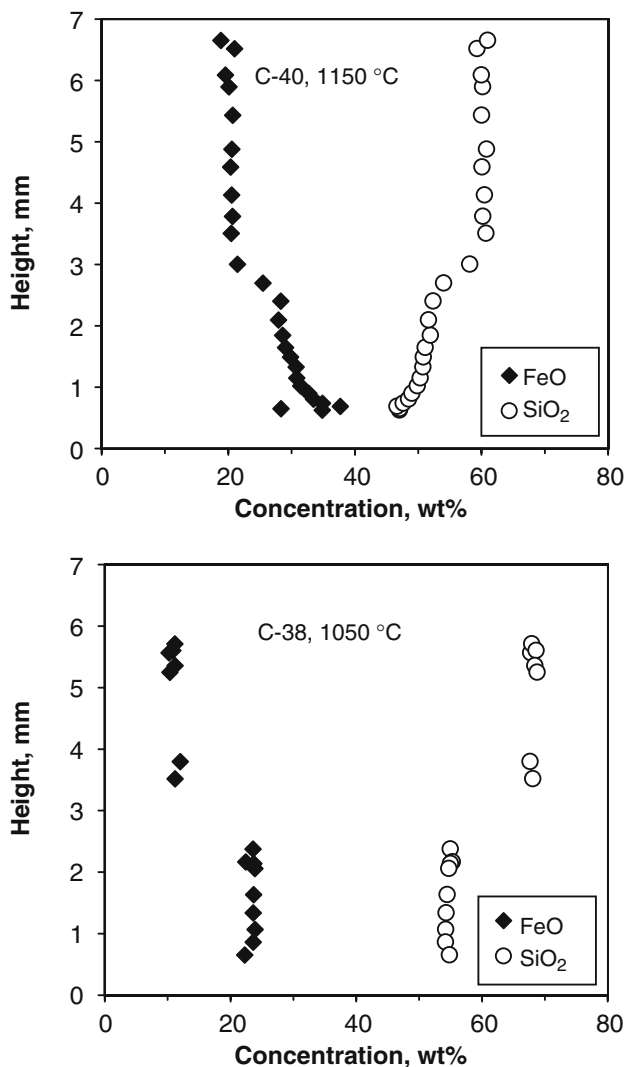


Fig. 3 Vertical compositional gradients of FeO and SiO₂ in quench glasses of the ferrobaltic samples. See text for discussion

Both immiscible liquids of the borosilicate composition quenched to whitish, semitransparent glasses, similar to those formed in our previous experiments with borosilicate melts (Veksler et al. 2002). Centrifuge separation of the liquids was also incomplete and resulted in three layers of approximately equal height: homogenous *Lb* glass at the bottom, homogenous *Lsi* glass at the top, and a medium layer of intermingling spherical droplets of one immiscible melt in another. Electron microprobe study of the borosilicate sample C-20 did not reveal any significant compositional gradients in the homogenous layers, and the chemical compositions of the *Lsi* and *Lb* glasses were found to be uniform in the layers and dispersed droplets throughout the sample.

Chemical composition of coexisting phases

Crystal phases

All the ferrobaltic run products contain tridymite, although in the sample C-40 run at 1,150°C it is present only in trace amounts, and probably is a quench phase. Olivine and Fe–Ti-oxides crystallised in the SF-1 starting composition below 1,120°C. Electron microprobe analyses of the crystal phases are presented in Table 3. The analyses of tridymite are not included, but they were found to correspond to pure SiO₂. The composition of olivine varies slightly with falling temperature from Fo₁₄ to Fo₁₂, but the variation is about the same as the extent of compositional zoning in a single sample. Oxide crystals are Fe–Ti solid solutions composed of two distinct inter-grown phases. The predominant phase is interpreted as wüstite with 2.5–4.6 wt.% TiO₂ and less than 1 wt.% MgO and Al₂O₃. The second, less abundant phase contains more than 22 wt.% TiO₂, 4–4.8 wt.% Al₂O₃, and approaches the composition of ulvöspinel.

Immiscible liquids

Electron microprobe and LA ICP-MS analyses of the glasses (quenched liquids) are listed in Table 4. The conjugate immiscible liquids of the ferrobaltic compositions (*Lfe*) are similar to those produced in previous experimental studies (McBirney and Nakamura 1974; Dixon and Rutherford 1979; Longhi 1990) and found in natural volcanic glasses (Roedder and Wiebelen 1971; Roedder 1979; Philpotts 1982; Jakobsen et al. 2005). The hallmarks of the *Lfe* are very high contents of iron oxides totalling to 28–35 wt.%, and high CaO/Al₂O₃ values usually exceeding 1 in mass units. Silica content in the *Lfe* is at about 48–54 wt.% in this study, but it can be as low as 38–40 wt.% according to some previously published experimental and natural compositions. The compositions of the conjugate silica-rich liquid *Lsi* are similar to those of natural rhyodacites, with SiO₂ around 60–75 wt.% and high feldspar norms. Thus, in terms of liquid–liquid element partitioning the silica-rich conjugate liquid is enriched in alkalis (except Li), Al and Si. All the other elements with a possible exception of Mo concentrate in the *Lfe*.

The immiscible silica-rich liquid *Lsi* in the borosilicate composition is composed mostly of SiO₂ and B₂O₃. The latter was not analysed directly by electron microprobe, but according to the low totals, the combined concentration of both oxides amounts to more than 90 wt.% (Table 3). All the other major oxides are at concentration levels below 2 wt.%. The conjugate boron-rich

Table 4 Compositions of immiscible liquids based on electron microprobe and LA ICP-MS analyses of quenched glasses

Sample	C-36		C-37		C-38		C-39		C-40		C-137		C-20	
<i>T</i> (°C)	1,100		1,075		1,050		1,125		1,150		1,160		1,100	
Phase	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lfe</i>	<i>Lsi</i>	<i>Lb</i>
SiO ₂	69.36	49.20	68.94	52.58	68.18	54.42	60.11	48.25	60.09	49.63	67.06	51.42	55.66	35.03
TiO ₂	1.33	2.34	1.37	2.22	1.49	2.50	2.18	2.18	2.14	2.27	2.46	4.84	0.08	0.14
Al ₂ O ₃	7.63	5.60	7.78	6.14	8.46	6.27	6.36	5.39	6.30	5.51	6.44	6.92	1.53	2.17
MgO	0.60	1.26	0.59	1.09	0.48	0.93	1.03	1.25	1.02	1.19	1.52	2.92	1.12	2.91
FeO	12.03	31.13	11.86	26.64	10.98	23.58	20.43	32.65	20.38	31.33	10.81	20.94	0.007	0.015
CaO	3.01	6.86	3.25	7.12	3.55	7.82	5.37	6.60	5.41	7.03	3.73	7.48	1.41	3.86
SrO	0.08	0.11					0.11	0.12					1.14	3.05
BaO	0.00	0.02					0.01	0.02					0.60	1.47
Na ₂ O	0.64	0.41	0.66	0.49	0.69	0.50	0.51	0.41	0.51	0.41	1.05	0.98	0.72	1.05
K ₂ O	4.59	1.95	4.75	2.45	4.93	2.48	3.17	1.87	2.84	1.79	4.15	2.34	0.83	1.04
Rb ₂ O	0.20	0.10	0.18	0.12	0.18	0.11	0.14	0.09					0.98	1.15
Cs ₂ O													1.02	1.07
Total	99.47	98.97	99.39	98.84	98.94	98.61	99.42	98.81	98.70	99.14	97.21	97.84	65.09	52.95
Li	188	345			260	399	333	333	337	358	167	180	1,523	3,012
Be	63	119			74	129	105	107	116	125	68	96	37	75
Cu	66	147			67	112	81	96	104	121	26	56	5.5	4.8
Zn	105	378			193	428	321	372	147	366	51	141	3.5	3.0
Ga	153	266			227	296	260	254	215	233	151	213	7.8	16.6
Rb	471	480			852	643	847	505	806	574	690	498	5,333	6,041
Sr	659	1,439			811	1,603	1,191	1,353	1,205	1,422	538	1,158	6,376	17,898
Y	70	143			79	170	128	135	128	142	65	133	77	213
Zr	228	379			275	455	412	370	406	395	237	384	280	523
Nb	191	366			209	411	338	338	342	366	174	390	258	534
Mo	87	169			199	162	173	93	158	160	101	96	4.4	3.8
In	179	442			283	517	467	450	272	469			d. l.	d. l.
Sn	4.1	7.6			5.1	8.3	8.1	8.5					187	386
Cs											689	477		
Ba	696	1,371			945	1,594	1,263	1,317	1,258	1,376	572	1,07	4,749	12,306
La	71	151			81	179	131	143	131	149	57	141	77	224
Ce	63	134			70	156	114	125	116	133	40	104	68	191
Pr	70	150			79	176	128	141	127	146	50	128	73	213
Nd	72	154			81	182	132	146	130	149	57	143	77	222
Sm	73	155			84	187	136	149	133	153	57	146	80	229
Eu	71	162			84	181	130	152	135	165	56	141	81	230
Gd	75	157			83	183	135	146	133	151	62	149	78	226
Tb	76	159			85	184	136	148	139	157	60	143	76	217
Dy	75	156			84	186	136	148	138	156	66	156	83	236
Ho	76	158			85	185	138	148	139	155	67	154	84	236
Er	77	157			88	191	142	153	140	156	66	154	84	233
Tm	76	155			86	185	139	149	137	152	69	154	86	236
Yb	77	155			92	194	147	157	145	161	72	165	91	250
Lu	79	159			91	192	146	155	143	158	74	160	90	245
Hf	240	406			299	468	416	389	426	425	258	420	319	627
Ta	247	450			281	503	421	412	451	468	239	481	366	647
W	177	392			227	446	305	367	290	373	94	200	241	642
Pb											1.6	4.5		
Th	0.12	0.22			0.14	0.25	0.22	0.20	0.23	0.20	0.13	0.25	0.06	0.14
U	0.08	0.14			0.09	0.19	0.14	0.14	0.15	0.14	0.06	0.00	0.03	0.08

Oxide concentrations are in wt.%, trace elements, in ppm

liquid *Lb* is also enriched in network-modifying elements and low in SiO₂ (35 wt.%). The compositional features of the immiscible borosilicate liquids are in good agreement with the immiscibility patterns in borosilicate ternaries (Levin et al. 1964) and our previous study of multicomponent systems (Veksler et al. 2002).

Liquid–liquid partition coefficients

The simplest conventional way to quantify element partitioning is by Nernst partition coefficients *D*'s, calculated as oxide or elemental ratios:

$$D_i = \frac{C_i^{Lfe}}{C_i^{Lsi}}$$

where C 's are weight concentrations, subscripts denote components, and superscripts denote phases (*Lfe*, Fe-rich melt; *Lsi*, silica-rich melt). In case of the borosilicate conjugate liquids, weight concentrations in *Lb* are divided by the concentrations in the conjugate *Lsi*. The

liquid–liquid D values calculated from the concentrations in Table 4 are presented in Table 5. Elements are grouped according to their ionic charge Z , and for each ion the table also lists coordination number CN , ionic radius r , and ionic potential Z/r . Some elements having multiple oxidation states (Sn, Mo and W) are presented by two alternative cations (see discussion below). The coordination numbers are mostly based on

Table 5 Ionic properties and liquid–liquid D values

	Z	N	r (pm)	Z/r ($\times 10^3$)	Nernst partition coefficients (D)							
					C-36	C-37	C-38	C-39	C-40	C-137	C-20	
Cs	1+	8	188	5.3							0.69 (0.21)	1.05 (0.08)
Rb	1+	6	166	6.0	0.48 (0.15)		0.59 (0.11)	0.60 (0.05)	0.71 (0.03)	0.72 (0.13)	0.72 (0.13)	1.17 (0.04)
K	1+	6	152	6.6	0.42 (0.03)	0.52 (0.08)	0.50 (0.04)	0.59 (0.06)	0.63 (0.12)	0.56 (0.03)	0.56 (0.03)	1.25 (0.06)
Na	1+	6	116	8.6	0.64 (0.12)	0.74 (0.09)	0.72 (0.07)	0.80 (0.10)	0.80 (0.15)	0.94 (0.08)	0.94 (0.08)	1.45 (0.09)
Li	1+	5	82	12.2	1.84 (0.16)		1.53 (0.18)	1.00 (0.05)	1.06 (0.04)	1.07 (0.12)	1.07 (0.12)	1.98 (0.10)
Cu	1+	6	91	11.0	2.22 (0.13)		1.67 (0.36)	1.18 (0.05)	1.16 (0.08)	2.13 (0.13)	2.13 (0.13)	0.88 (0.23)
Ba	2+	6	149	13.4	1.97 (0.10)		1.69 (0.11)	1.04 (0.03)	1.09 (0.02)	1.93 (0.10)	1.93 (0.10)	2.45 (0.16)
Sr	2+	6	132	15.2	2.18 (0.07)		1.98 (0.07)	1.14 (0.01)	1.18 (0.03)	2.15 (0.08)	2.15 (0.08)	2.68 (0.08)
Ca	2+	6	114	17.5	2.28 (0.26)	2.19 (0.46)	2.21 (0.17)	1.23 (0.03)	1.30 (0.28)	2.41 (0.12)	2.41 (0.12)	2.74 (0.10)
Mg	2+	6	86	23.3	2.10 (0.25)	1.85 (0.38)	1.93 (0.41)	1.21 (0.10)	1.16 (0.16)	1.92 (0.15)	1.92 (0.15)	2.60 (0.08)
Be	2+	5	50	40.0	1.89 (0.05)		1.74 (0.12)	1.02 (0.02)	1.08 (0.03)	1.42 (0.07)	1.42 (0.07)	2.05 (0.14)
Zn	2+	6	88	22.7	3.59 (0.31)		2.22 (0.90)	1.16 (0.04)	2.49 (0.17)	2.80 (0.27)	2.80 (0.27)	0.87 (0.30)
Fe	2+	6	92	21.7	2.59 (0.28)	2.25 (0.40)	2.15 (0.13)	1.60 (0.09)	1.54 (0.16)	1.94 (0.12)	1.94 (0.12)	1.97 (0.31)
Sn	2+	6	123	16.3	1.86 (0.16)		1.63 (0.41)	1.04 (0.10)				
Pb	2+	6	133	15.0						2.77 (0.32)		
Eu	2+	6	131	15.3	2.29 (0.08)		2.15 (0.05)	1.17 (0.02)	1.22 (0.04)	2.49 (0.13)	2.49 (0.13)	
Y	3+	7	110	27.3	2.03 (0.05)		2.16 (0.09)	1.06 (0.02)	1.11 (0.03)	2.03 (0.19)	2.03 (0.19)	2.77 (0.10)
La	3+	7	123	24.4	2.12 (0.06)		2.22 (0.11)	1.09 (0.02)	1.14 (0.03)	2.50 (0.24)	2.50 (0.24)	2.91 (0.13)
Ce	3+	6	115	26.1	2.13 (0.08)		2.24 (0.11)	1.10 (0.02)	1.15 (0.04)	2.58 (0.23)	2.58 (0.23)	2.80 (0.10)
Pr	3+	6	113	26.5	2.14 (0.07)		2.23 (0.10)	1.10 (0.02)	1.15 (0.03)	2.56 (0.22)	2.56 (0.22)	2.92 (0.14)
Nd	3+	6	112.3	26.7	2.14 (0.08)		2.24 (0.09)	1.10 (0.02)	1.15 (0.03)	2.51 (0.24)	2.51 (0.24)	2.88 (0.09)
Sm	3+	6	109.8	27.3	2.13 (0.07)		2.23 (0.10)	1.10 (0.02)	1.14 (0.03)	2.55 (0.24)	2.55 (0.24)	2.88 (0.11)
Eu	3+	6	108.7	27.6								2.83 (0.11)
Gd	3+	6	107.8	27.8	2.10 (0.07)		2.21 (0.10)	1.09 (0.02)	1.13 (0.03)	2.38 (0.26)	2.38 (0.26)	2.89 (0.15)
Tb	3+	6	106.3	28.2	2.09 (0.07)		2.18 (0.10)	1.09 (0.03)	1.13 (0.03)	2.38 (0.26)	2.38 (0.26)	2.86 (0.12)
Dy	3+	6	105.2	28.5	2.08 (0.06)		2.20 (0.10)	1.09 (0.03)	1.13 (0.03)	2.37 (0.24)	2.37 (0.24)	2.85 (0.13)
Ho	3+	6	104.1	28.8	2.07 (0.06)		2.17 (0.10)	1.08 (0.02)	1.12 (0.03)	2.30 (0.24)	2.30 (0.24)	2.82 (0.12)
Er	3+	6	103	29.1	2.05 (0.06)		2.17 (0.11)	1.08 (0.02)	1.11 (0.03)	2.32 (0.26)	2.32 (0.26)	2.76 (0.10)
Tm	3+	6	102	29.4	2.04 (0.06)		2.15 (0.10)	1.07 (0.02)	1.11 (0.02)	2.24 (0.21)	2.24 (0.21)	2.75 (0.10)
Yb	3+	6	100.8	29.8	2.03 (0.05)		2.12 (0.11)	1.07 (0.03)	1.11 (0.02)	2.29 (0.28)	2.29 (0.28)	2.74 (0.11)
Lu	3+	6	100.1	30.0	2.02 (0.05)		2.11 (0.10)	1.07 (0.02)	1.10 (0.02)	2.17 (0.20)	2.17 (0.20)	2.73 (0.09)
Ga	3+	6	76	39.5	1.74 (0.09)		1.30 (0.21)	0.97 (0.06)	1.08 (0.03)	1.41 (0.10)	1.41 (0.10)	2.15 (0.25)
In	3+	6	94	31.9	2.47 (0.23)		1.83 (0.65)	0.96 (0.04)	1.73 (0.06)			
Al	3+	4	53	56.6	0.73 (0.02)	0.79 (0.05)	0.74 (0.04)	0.85 (0.02)	0.87 (0.05)	1.07 (0.03)	1.07 (0.03)	1.42 (0.05)
Sn	4+	6	83	48.2								2.06 (0.11)
Th	4+	8	119	33.6	1.84 (0.23)		1.74 (0.59)	0.93 (0.14)	0.90 (0.06)	1.85 (0.17)	1.85 (0.17)	2.25 (0.45)
U	4+	8	114	35.1	1.73 (0.11)		2.19 (0.44)	1.02 (0.16)	0.98 (0.09)			2.23 (0.68)
Zr	4+	8	98	40.8	1.66 (0.06)		1.66 (0.10)	0.90 (0.03)	0.97 (0.03)	1.62 (0.06)	1.62 (0.06)	1.87 (0.10)
Hf	4+	6	85	47.1	1.69 (0.07)		1.56 (0.10)	0.94 (0.03)	1.00 (0.02)	1.63 (0.08)	1.63 (0.08)	1.97 (0.10)
Ti	4+	5	66	60.6	1.76 (0.16)	1.61 (0.24)	1.68 (0.30)	1.00 (0.03)	1.06 (0.10)	1.97 (0.20)	1.97 (0.20)	1.89 (0.36)
Si	4+	4	40	100	0.71 (0.03)	0.76 (0.05)	0.80 (0.01)	0.80 (0.02)	0.83 (0.03)	0.77 (0.01)	0.77 (0.01)	0.63 (0.01)
Mo	4+	6	79	50.6	1.94 (0.96)		0.82 (0.37)	0.54 (0.25)	1.01 (0.26)	0.96 (0.52)	0.96 (0.52)	
W	4+	6	80	50.0	2.21 (0.12)		1.97 (0.32)	1.20 (0.08)	1.29 (0.08)	2.12 (0.38)	2.12 (0.38)	
Nb	5+	6	78	64.1	1.92 (0.05)		1.97 (0.13)	1.00 (0.03)	1.07 (0.02)	2.24 (0.18)	2.24 (0.18)	2.07 (0.09)
Ta	5+	6	78	64.1	1.82 (0.05)		1.79 (0.07)	0.98 (0.03)	1.04 (0.02)	2.01 (0.18)	2.01 (0.18)	1.77 (0.12)
Mo	6+	6	73	82.2								0.86 (0.29)
W	6+	6	74	81.1								2.67 (0.08)

Standard deviations are in brackets

the spectroscopic evidence reviewed by Hudon and Baker (2002a), and ionic radii are from Shannon (1976), assuming $rO^{2-} = 121$ pm, and with the high-spin value for Fe^{2+} . Standard deviations of the D values are given in brackets in Table 5. They were calculated using error propagation formulae from standard deviations of the average concentrations. In Fig. 4 the liquid–liquid D values of three most representative experimental samples are plotted against Z/r . In the series of experiments with the starting composition SF-1 (samples from C-36 to C-40, Table 2), D values in

two-liquid field at temperatures above olivine and oxide liquidus generally decrease, as one would expect, with rising temperature, as the compositions of conjugate melts become closer while consolute (critical) temperature is approached and the miscibility gap closes up. Judging by the low D values in the sample C-40 (Table 4), the consolute temperature for the starting mixture SF-1 is not far above 1,150°C.

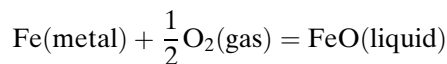
Discussion

Attainment of equilibrium

Vertical chemical gradients in the bottom *Lfe* layers (Fig. 3), diffuse streaks and intermingling of *Lfe* and *Lsi* instead of sharp liquid–liquid interfaces in the ferrobasaltic charges (Fig. 2b) suggest incomplete separation of the immiscible liquids, possible heterogeneity on a sub-micron scale, and are obvious signs of thermodynamic disequilibrium. The attainment of chemical equilibrium between phases is a crucial issue, which bears upon the reliability of the key liquid–liquid D values in Table 5. Centrifuge separation is a new technique, which enabled us to obtain large, analysable pools of immiscible liquids, but limited run durations may have hampered chemical equilibration of the phases by convection and diffusion. In this section we examine the problem of equilibrium, drawing comparisons with other previous studies of silicate liquid immiscibility, and using thermodynamic criteria based on regression of experimental data.

Oxygen fugacity

Oxygen fugacity in the ferrobasaltic charges was buffered by the dissolution reaction of metal iron (the inner container) in silicate melts:



Notably, this reaction of liquid(s) with metal iron container buffers not only oxygen fugacity, but also the activity of FeO component. Ariskin et al. (1997) calibrated the reaction using a set of 396 one-atmosphere melting experiments, in which natural or synthetic silicate liquids had been equilibrated with pure metal iron. They processed the data by least-squares fitting and presented as an expression

$$\log X_{FeO}^L = k \log fO_2 + \frac{h}{T} + \sum d_j X_j^L$$

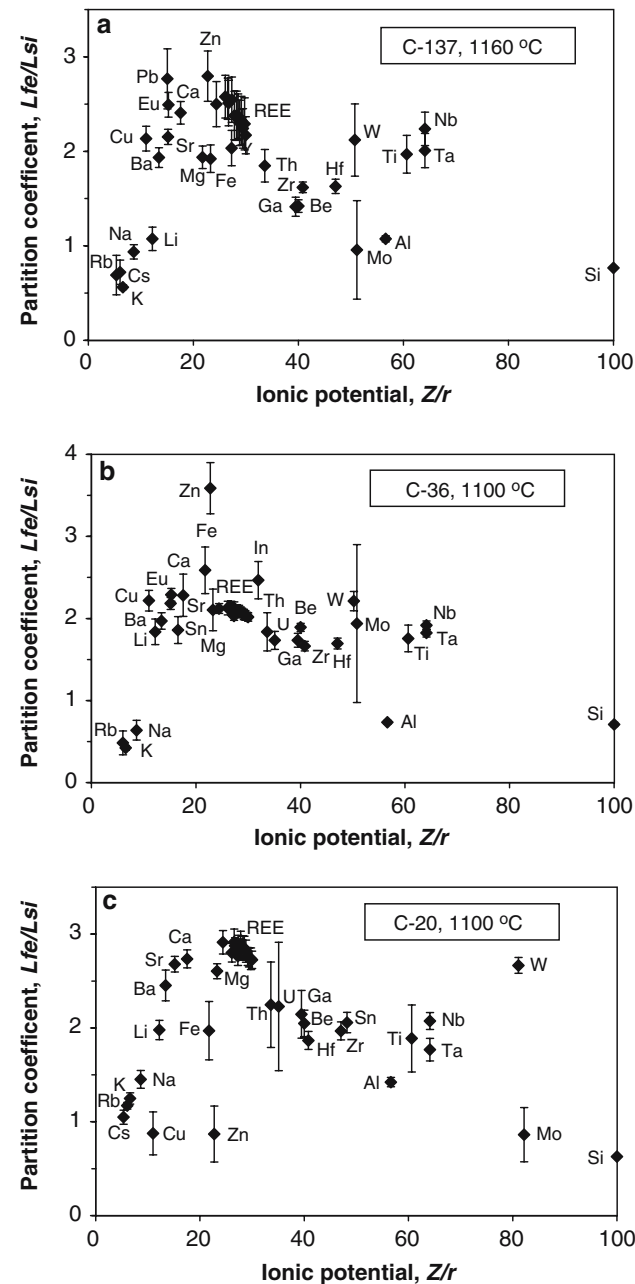


Fig. 4 Liquid–liquid Nernst partition coefficients D plotted against ionic potential Z/r

where X_{FeO}^L is the molar percent concentration of the ferrous Fe in the melt, X_j^L are molar percent concentrations of seven other major oxides (SiO_2 , TiO_2 , Al_2O_3 , MgO , CaO , Na_2O and K_2O), T is absolute temperature, and k , h and d_j are regression parameters (Ariskin et al. 1997). Using this expression and liquid compositions in Table 4, we calculated $\log f\text{O}_2$ values for the conjugate *Lfe* and *Lsi*. Because the proportion of the ferric Fe in equilibrium with metal Fe is low, we assumed X_{FeO}^L to be equal to FeO total. If the regression equation accounts for the compositional effects well, and the equilibrium was reached, the pairs of immiscible liquids from the same run should give the same, within the analytical and statistical uncertainty, $\log f\text{O}_2$ values close to those of the Fe–FeO buffer. Thus, the $f\text{O}_2$ calculation can be used as one test of chemical equilibrium.

The results are presented in Table 6. All the calculated $\log f\text{O}_2$ values are approximately 0.5–1 log-units lower than those of the Fe–FeO buffer, exactly as expected as the FeO activities in the liquids (and also the oxide solid solutions, when they are present) are less than 1. The fugacities obtained for the Si-rich liquids *Lsi* are however systematically lower than those for the conjugate Fe-rich *Lfe* by 0.36–0.76 log-units. Compared to the estimated accuracy of the calculated $\log f\text{O}_2$ values (± 0.1 at 1σ level, Ariskin et al. 1997), the difference appears to be significant, and may indicate chemical disequilibrium. It should be noted though, that the equation was calibrated against experiments at higher temperatures and more basic compositions, and may not work well for rhyolitic liquids. Our further calculations showed that not only our compositions, but all the previously published analyses of *Lsi* and *Lfe* pairs kept in Fe containers (Longhi 1990; Roedder and

Weiblen 1970; Watson 1976) gave systematically lower $\log f\text{O}_2$ values for *Lsi*, and the *Lsi*–*Lfe* difference in the calculated $\log f\text{O}_2$ values often exceeded 1 log-unit. One may thus conclude that either the extrapolation of the equation to lower temperatures and more silicic compositions does not work well, or rhyolitic liquids never equilibrate with metal iron and conjugate *Lfe* at experimentally realistic run durations. In any case, our liquid compositions seem to pass the equilibration test reasonably well, no worse than the previously published results.

Fe–Mg distribution between olivine and the immiscible liquids

In olivine-saturated experiments chemical equilibrium is traditionally checked by compliance of Fe–Mg distribution between olivine and melt with the experimentally calibrated equilibrium constant K_D . Roeder and Emslie (1970) showed that the distribution constant K_D calculated as

$$K_D = \frac{X_{\text{Fe}}^{\text{ol}}/X_{\text{Mg}}^{\text{ol}}}{X_{\text{Fe}}^L/X_{\text{Mg}}^L}$$

in melt compositions similar to natural tholeiites does not vary much with temperature and is close to the value of 0.3 ± 0.03 . The K_D values in our olivine-saturated runs calculated from the data in Tables 3 and 4 are much greater than the classical basaltic value: 0.46–0.5 for the olivine–*Lfe* distribution, and 0.55–0.58 for the olivine–*Lsi* distribution (Table 6). The experimental immiscible liquids are however compositionally far from common tholeiitic basalts. Two recent studies, which revisited the K_D calibration, reported a strong

Table 6 Oxygen fugacity calculated for the immiscible liquids *Lfe* and *Lsi* using the equation by Ariskin et al. (1997), and the experimental K_D values of the olivine-melt equilibria compared to the models by Kilinc and Gerke (2003) and Toplis (2005)

Sample	Liquid	Equilibrium crystals	T (°C)	Calculated $\log f\text{O}_2$	K_D olivine-melt		
					Measured	Kilinc and Gerke (2003)	Toplis (2005)
C-137	<i>Lfe</i>	tr	1,160	–13.4			
C-137	<i>Lsi</i>	tr	1,160	–13.9			
C-40	<i>Lfe</i>	tr	1,150	–13.2			
C-40	<i>Lsi</i>	tr	1,150	–13.6			
C-39	<i>Lfe</i>	tr	1,125	–13.4			
C-39	<i>Lsi</i>	tr	1,125	–13.8			
C-36	<i>Lfe</i>	ol + Fe–Ti + tr	1,100	–13.8	0.47	0.29	0.45
C-36	<i>Lsi</i>	ol + Fe–Ti + tr	1,100	–14.5	0.58	0.63	0.39
C-37	<i>Lfe</i>	ol + Fe–Ti + tr	1,075	–14.2	0.47	0.30	0.45
C-37	<i>Lsi</i>	ol + Fe–Ti + tr	1,075	–14.8	0.56	0.60	0.39
C-38	<i>Lfe</i>	ol + Fe–Ti + tr	1,050	–14.6	0.52	0.31	0.46
C-38	<i>Lsi</i>	ol + Fe–Ti + tr	1,050	–15.2	0.57	0.58	0.38

Abbreviations for the liquid and crystal phases the same as in Table 2

compositional dependence of the constant. Kilinc and Gerke (2003) showed that K_D values could be as low as 0.18 in silica-undersaturated alkali basalt compositions, and up to 0.68 in silica-rich andesitic–rhyolitic melts. The K_D variations in the experimental data set used by Toplis (2005) ranged from 0.17 to 0.45. It follows that silica activity and alkali content in the melt strongly affect the K_D . All our ferrobaltic run products were saturated in pure silica (tridymite), and consequently silica activity in the immiscible liquids was fixed at 1. The elevated K_D values may thus reflect the very high silica activity in the liquids.

We also used regression equations proposed by Toplis (2005) and Kilinc and Gerke (2003), which calculated K_D 's from melt composition and experimental conditions. The results are presented in Table 6, and they are mixed. On one hand, the equations by Toplis (2005) reproduce our olivine–*Lfe* K_D values very well, but they do not give the same excellent results for *Lsi*. On the other hand, the compositional regression by Kilinc and Gerke (2003) is consistent with our experimental values for *Lsi*, but does not work well for the *Lfe* compositions. One should also keep in mind that the K_D estimations have large error margins because of large analytical uncertainties for MgO at low concentrations, especially in *Lsi*. All in all, it may be concluded that our liquid compositions in ferrobaltic runs pass the olivine K_D test satisfactory, and further improvements are probably needed in the regression equations to better account for the compositional effects.

Previous reports of olivine composition in equilibrium with the immiscible liquids are limited, to our knowledge, to a single analysis of olivine Fo₅ at 1,045°C reported by Roedder and Weiblen (1970), which gives K_D value of 0.87 for the Fe-rich liquid and 0.54 for the silica-rich liquid. The value 0.87 is way too high, but the value for the silica-rich liquid is in good agreement with our study.

Effects of incomplete phase separation

Blurred *Lfe*–*Lsi* interfaces (Fig. 2b), and vertical chemical gradients in ferrobaltic samples (Fig. 3) imply that the quench glasses may be heterogeneous at sub-micron scale, and the heterogeneity may have arisen from incomplete separation of fine emulsions in the liquids before quench. Incomplete phase separation would primarily affect the absolute D values, but should have much smaller effect upon the relationship between the D values of individual elements in the same sample. If we are sampling incompletely separated emulsions, D values will be converging to 1, but

all the elements will be affected proportionately, and inter-elemental comparisons will be still valid. Because of the possibility of incomplete phase separation and mutual contamination of the liquids the absolute D values in Table 4 should be treated with caution and considered as minimal estimates. The contamination will not affect however the general patterns of elemental plots, like those presented in Fig. 4.

Effects of variable diffusion rates

If liquid–liquid chemical equilibration is controlled by diffusion, and run durations were not long enough to achieve complete equilibration, one may expect a situation when equilibrium is reacted for elements with rapidly diffusing cations (e.g., Li and other alkalis), and not reached for slowly diffusing components (e.g., HFSE). In contrast to incomplete phase separation, variable diffusion rates will affect not only absolute D values, but also elemental ratios of slowly and rapidly diffusing elements. Although such a possibility cannot be completely ruled out in our runs, comparisons with previous kinetic studies (Martin and Kusiro 1991) and partitioning data obtained in experiments at longer durations (Watson 1976; Ryerson and Hess 1978; Vicenzi et al. 1994; Schmidt et al. 2006) suggest that the diffusion effects are likely to be minor at the timescale of our experiments. The effects of variable cooling rates upon phase separation and liquid–liquid element partitioning have been studied by Martin and Kushiro (1991). Their study demonstrated a fast kinetics of liquid–liquid chemical equilibration and small variations of the D values in response to variable cooling rates. In our experiments, phase separation and chemical equilibration should be enhanced in the sample C-137 where the proportion of the viscous *Lsi* liquid is very low (Fig. 2b). It is notable that the D values in C-137, although in general more contrasting than those in the C-40 run at similar T , show the same overall pattern of element partitioning. Finally, within certain limits imposed by variations in melt composition, pressure and temperature, our D values are in good agreement with previous estimations, and this is a further indication of the validity of our data. Generally, a good agreement with recent partitioning results by Schmidt et al. (2006) is most encouraging in this regard. Schmidt et al. (2006) also used centrifuge phase separation, but their samples were pre-equilibrated for 12 h at static conditions, and P and T of interest. The overall span of the absolute D values by Schmidt et al. (2006) appear to be greater, but the pattern of the D – Z/r plot is very similar to that presented in Fig. 4.

Relationships between partition coefficients and ionic potentials

Before looking at individual elements, let us consider first the overall pattern of the D values plotted against ionic potential Z/r (Fig. 4). The plots show that with increasing Z/r the D values in all the compositions first increase, reach a maximum somewhere between Ca and La, and then decrease slowly. Highly charged cations with Z/r greater than 40 show a very large scatter of D values, but even in that region many elements are distributed in a similar way in all the three compositions. In fact, such general pattern of the D – Z/r plots is characteristic not only for the aluminosilicate and borosilicate immiscible liquids, but other cases of immiscibility including silicate melts and various molten salts (Veksler 2004; Veksler et al. 2005). Let us now proceed to specific chemical and structural groups of elements.

Alkalis and alkaline earths

Alkalis from K to Li and alkaline earths from Ba to Ca show gradual increase in the liquid–liquid D values consistent with the increasing Z/r of the cations, and in good agreement with broadening miscibility gaps in the oxide-silica binaries (Hess 1995; Hudon and Baker 2002a). In the binaries, the tendency to liquid immiscibility increases in the order $K < Na < Li < Ba < Sr < Ca < Mg$ as do, in general, the liquid–liquid D values (Table 5; Fig. 4). In our experiments D_{Mg} is however systematically lower than D_{Ca} , but this drop can be explained by the amphoteric properties of Mg^{2+} cation (Hudon and Baker 2002a; see also Schmidt et al. 2006). The tendency develops even further as we move to Be^{2+} , which is more amphoteric than Mg^{2+} , and accordingly, has a lower D value.

Oxide–silica binaries of Rb and Cs show no signs of liquid immiscibility, and in the ferrobasic compositions the elements concentrated in *Lsi*. In our ferrobasic samples D values of Rb and Cs are same as D_K within the analytical uncertainty. Other studies (Watson 1976; Vicenzi et al. 1994; Schmidt et al. 2006) demonstrated a systematic and significant variation in compliance with the general D – Z/r trend: $D_{Cs} < D_{Rb} < D_K$. The trend is also revealed in our borosilicate sample C-20 (Fig. 4c; Table 5).

In Fig. 5 the D values of the univalent alkalis and some divalent cations are plotted against r and fitted with parabolic curves. Such plots are similar to the Onuma diagrams, which are broadly used for plotting crystal–liquid partition coefficients (e.g., Blundy and

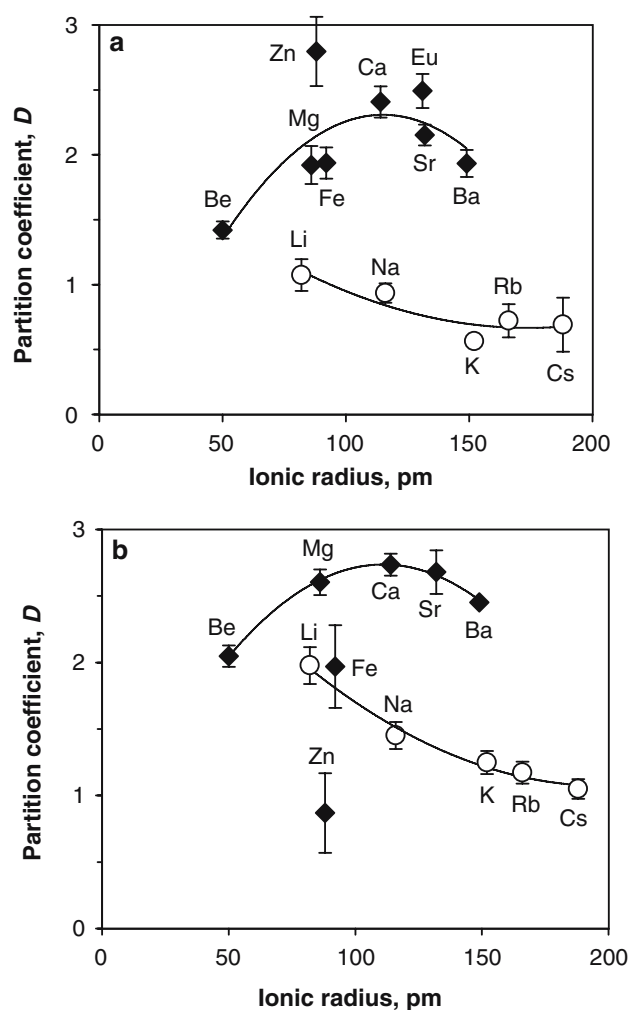


Fig. 5 Partition coefficients plotted against ionic radii. **a** Ferrobasic sample C-137. **b** Borosilicate sample C-20

Wood 1994). The parabolic fits are tight and convex for the divalent cations, and slightly concave for the alkalis. The implications of Fig. 5 for crystal–melt partitioning are discussed in one of the following sections in more detail.

Transition elements Cu, Fe, and Zn

In the ferrobasic compositions, D values of these elements usually plot above the main trend (Figs. 4, 5). Slightly anomalous D_{Fe} values can well be explained in light of the classification proposed by Hudon and Baker (2002a, b), as Fe^{2+} belongs to the VCFSE (variable field strength stabilisation energy) cations. Such an explanation is not valid, though, for Cu^+ and Zn^{2+} , which in the ferrobasic composition also show significantly higher D values than the homovalent cations with a similar Z/r . In fact, Zn in our ferrobasic samples shows the highest D value among all the

studied elements. In the binary systems, Zn^{2+} is classified as an amphoteric cation (Hudon and Baker 2002a), and that implies, if anything, anomalously low D_{Zn} values. Systematically lower D_{Zn} values were demonstrated, in accordance with the theoretical predictions, in experiments by Schmidt et al. (2006). No comparisons can be done for Cu, because this element has never been included, to the best of our knowledge, in liquid–liquid partitioning studies before.

One possible explanation for the unusual behaviour of Cu and Zn in our runs might be that in the reduced conditions set by the Fe–FeO buffer Cu and Zn are partly reduced to metal (see, for example, O’Neil and Pownceby (1993) for the data on the Cu–Cu₂O equilibrium), and form tiny alloy droplets sinking down and contaminating the *Lfe* layer. However, the experiments by Schmidt et al. (2006), which demonstrated low D_{Zn} values, were also at similarly reducing conditions set by the Mo–MoO₂ buffer. Thus, it seems that no simple explanation can be offered for Zn and Cu at the moment, and the matter should be resolved in future experiments carried out at more oxidising conditions. Notably, some studies of natural immiscible glasses report high ZnO contents, at about 0.2 wt.% in Fe-rich globules (Naumov et al. 1993; Hurai et al. 1998), but the globules are too high in total Fe to be directly comparable with our glass compositions.

In the borosilicate composition, which was run in air (sample C-20), all the transition elements have, on the contrary, anomalously low D values. The reason for such partitioning is again not clear, but apparently the D values result from chemical interactions of Cu^{2+} , Zn^{2+} and Fe cations with borate and borosilicate species. For instance, the anomalously low D values of the transition elements may indicate that these elements do not bond well with the trigonal $(BO_3)^{3-}$ anions, which are likely to be more abundant in the *Lb* immiscible liquid than in the *Lsi* according to the previous studies of speciation of borosilicate immiscible glasses by Raman spectroscopy (Veksler et al. 2002).

Rare earth elements

With the exception of Eu^{2+} , REE form trivalent cations with r gradually decreasing from 123 pm (La^{3+}) to 100.1 pm (Lu^{3+}). The D values of the REE trivalent cations slightly decrease with the increasing atomic number, as the heavier REE become increasingly amphoteric (Hudon and Baker 2002a). Europium has two oxidation states: 2+ and 3+, and Eu^{2+} has r and Z/r values close to those of Sr^{2+} (Table 5). In the reduced ferrobaltic compositions, Eu is predominantly in the

divalent form, and accordingly D_{Eu} in them is close to D_{Sr} . In contrast, borosilicate composition was run at oxidised conditions, in which Eu should form Eu^{3+} , and D_{Eu} accordingly plot in-between D_{Sm} and D_{Gd} (Fig. 4c). In general, REE are among the most contrastingly distributed elements, and they strongly enrich the depolymerised *Lfe* and *Lb* liquids. Other authors (Watson 1976; Ryerson and Hess 1978; Vicenzi 1994; Schmidt et al. 2006) usually reported higher absolute D_{REE} values, up to about 5 or 10, with D_{REE} higher than D_{Fe} (like in our sample C-137).

High field strength elements

In the group of high field strength elements (HFSE) we include Ti, Zr, Hf, Nb, Ta, Mo and W, which form relatively small cations with the nominal charges from 4+ to 6+. Most of the elements have multiple oxidation states, and thus the observed D values listed in Table 5 may in fact represent some effective mixtures of several ions with different Z . The situation is particularly complex for Mo and W, which are likely to form 4+ cations in reduced ferrobaltic samples, and 6+ cations in the oxidised borosilicate sample C-20 (see Holzheid et al. 1994; Ertel et al. 1996; O’Neil and Eggins 2002 for the data on Mo and W oxidation states in silicate melts). The exact proportion of the 4+ and 6+ cations in our samples are not known. In Table 5 and Fig. 5 we simply assumed all Mo and W as 4+ in ferrobaltic samples and 6+ in the borosilicate sample. The choice of the r values is also arbitrary because in silicate crystals and melts the cations can have more than one coordination number.

The scatter of the D values in the high field strength part of the plots in Fig. 4 is in any case so large that it defies an easy structural interpretation. With the exception of Mo, the D values of the HFSE seem to form a sequence: $D_{Zr} = D_{Hf} < D_{Ti} = D_{Ta} < D_{Nb} < D_W$, with a rough positive correlation with Z , and also Z/r . As one moves from the 4+ cations to 5+ and 6+, the chemical properties of the oxides change from amphoteric to increasingly acidic, and the tendency of the cations to form oxyanions, such as $(WO_4)^{2-}$, also increases. The growing D values may reflect a growing tendency of the oxyanions to bond with typical network-modifying cations, such as alkalis and alkaline earths, rather than insert themselves into polymerised aluminosilicate or borosilicate networks. The interactions with the cations in the depolymerised *Lfe* and *Lb* seem to be more important for the partitioning behaviour of the HFSE than their compatibility with the polymerised network structure of *Lsi*. Comparisons with the oxide-silica binaries and ternaries are

instructive in this respect. For instance, Stemprok (1974) reported no liquid immiscibility along the $\text{WO}_3\text{--SiO}_2$ and $\text{MoO}_3\text{--SiO}_2$ border joins, but drew broad ternary miscibility gaps in the $\text{Na}_2\text{O--WO}_3\text{--SiO}_2$ and $\text{Na}_2\text{O--MoO}_3\text{--SiO}_2$ systems at 1,200°C. An interesting analogy is also provided by phosphorus. Apparently, there is no liquid immiscibility in the $\text{P}_2\text{O}_5\text{--SiO}_2$ binary (Tien and Hummel 1962), but in ferrobaltic systems P_2O_5 strongly enhances silicate liquid immiscibility, strongly concentrates on the Fe-rich liquid, and greatly increases the D values of network-modifying cations, such as alkaline earths and REE (Watson 1976; Ryerson and Hess 1978). Because of those effects, we deliberately excluded P from our starting compositions.

The decoupling of the Zr–Hf, Nb–Ta and W–Mo pairs is a notable feature of liquid–liquid equilibria in salt melt–silicate melt element partitioning (Veksler et al. 1998; Veksler 2004). As a rule, Zr, Nb and W have greater affinity to depolymerised, ionic liquids than their chemical “twins”, so that $D_{\text{Zr}} > D_{\text{Hf}}$, $D_{\text{Nb}} > D_{\text{Ta}}$ and $D_{\text{W}} > D_{\text{Mo}}$. The decoupling defies an explanation in the frame of the simple ionic model, because in any system of ionic radii, the size and Z/r values of the “twins” are very close. That is probably where the simple model fails, and the explanation should be sought in more sophisticated theories of chemical bonds. For some reason, Mo and W show a very large scatter in ferrobaltic samples, and the decoupling of the Mo–W, Zr–Hf and Nb–Ta pairs is insignificant in the *Lfe–Lsi* partitioning, when standard deviations (Table 5) are taken into account.

Network formers

Within the group, which includes Si^{4+} , Al^{3+} , Ga^{3+} , and B^{3+} , Si^{4+} is a perfect network former, while Al^{3+} and Ga^{3+} are somewhat too large and consequently amphoteric, and B^{3+} is too small and thus fitting into trigonal and tetrahedral sites. Indium lies below Al and Ga in the periodic table, but In^{3+} is too large to be a network former. In our ferrobaltic compositions, In^{3+} tends to plot slightly above the main trend of the D values (Fig. 4), but we cannot offer any explanation for such behaviour. Si^{4+} and Al^{3+} obviously concentrate in *Lsi*, while B^{3+} in the borosilicate composition preferentially partitions to the depolymerised *Lb* (see also Veksler et al. 2002). As with the HFSE, the liquid–liquid partitioning of Al^{3+} and B^{3+} appears to be strongly affected by chemical interactions with network-modifying cations, such as alkalis and alkaline earths. For instance, the partitioning of Al^{3+} to *Lsi* seems to be enhanced by the formation of strongly

polymerised feldspathoid species, which mix well with silica networks. That is probably why Al^{3+} in Fig. 4a and b tends to plot below the main down-slope trend formed by Be^{2+} , Ga^{3+} and Si^{4+} . Because of the ambiguity of the proportion between the trigonal and tetrahedral B^{3+} and uncertainty of the effective r value, B^{3+} is not included in the Z/r systematics and not plotted in Fig. 4c.

Implications to crystal-melt element partitioning

The effects of melt composition upon crystal-melt element partitioning have been a matter of extensive discussion in recent years (e.g., Mysen 2004; O’Neil and Eggins 2002; Prowatke and Klemme 2005; Schmidt et al. 2006). Watson (1976) and Ryerson and Hess (1978) were probably the first to propose that studies of immiscibility and two-liquid element distribution can help to better understand and quantify these effects, and Schmidt et al. (2006) developed the approach much further both experimentally and in their theoretical model. The liquid–liquid partitioning results presented here (Table 5; Fig. 4) imply that the melt composition effects should be the strongest for the most strongly fractionated elements. Transition metals are likely to be decoupled from other cations of a similar Z/r , as evidenced, for example, by the already discussed strong compositional dependence of the olivine-melt K_D (see also Kushiro and Mysen 2002). Strong and very complex effects are likely for the HFSE, and their crystal-melt partitioning may depend not only on degree of polymerisation of the silicate melt network, but also on the type of predominant network-modifying cations.

It is also important to note that the melt effects upon cations are not a linear function of Z/r . Non-linear patterns have been established in binary systems (Hudon and Baker 2002a, b), and are confirmed in the complex multicomponent compositions presented here. Despite the absence of lattice strain from liquids, the liquid–liquid D values of di- and trivalent network formers and amphoteric cations appear to form convex parabolic fits (see some examples in Fig. 5), similar to those calculated for crystal–liquid element partitioning using the lattice strain model developed by Blundy and Wood (1994). The lattice strain model, which basically ignores the melt effects, may work so well partly because of the similarity between the “parabolic” distribution of cations between different species in the melt, and between melt and crystal. The similarity does not however hold for the highly charged HFSE. In our runs, W and Mo having almost identical Z and r are

strongly fractionated, but $D_{\text{Nb}}/D_{\text{Ta}}$ does not exceed 1.1 in ferrobaltic compositions, and 1.17 in the borosilicate composition, and is within the limits of analytical uncertainty. In the study by Vicenzi et al. (1994), the Nb–Ta fractionation between the iron-rich and silicate-rich immiscible liquids was however significant, with $D_{\text{Nb}}/D_{\text{Ta}}$ values up to 1.35. The significant liquid–liquid fractionation proves that the difference in the crystal–liquid D_{Nb} and D_{Ta} values (e.g., Horng and Hess 2000; Prowatke and Klemme 2005) cannot be due to effects of crystal lattices only, and the speciation and local environments of Nb^{5+} and Ta^{5+} cations in liquid phase should play an important role.

Silicate liquid immiscibility in basaltic magmas

Recent findings of the conjugate Fe-rich and Si-rich melt inclusions in evolved gabbroic rocks from the Upper Zone of the Skaergaard intrusion, East Greenland (Jakobsen et al. 2005) confirm the viability of silicate liquid immiscibility as a petrogenetic process. The melt inclusion study proved that silicate liquid immiscibility proposed for Skaergaard (McBirney 1975; McBirney and Nakamura 1974) had taken place indeed, and because the Skaergaard intrusion has been considered for decades as the classical example of a tholeiitic magma chamber, liquid immiscibility may be not unique to Skaergaard but could have played a role, as it was proposed before, in other mafic intrusions such as the Stillwater complex (Loferski and Arculus 1993) and the Pleasant Bay intrusion (Wiebe 1997). In a series of ongoing experiments, we are now revisiting the extent and petrogenetic significance of silicate liquid immiscibility in natural basaltic magma. A detailed account of the results will be published elsewhere, and only some general considerations are presented below.

Conditions for silicate liquid immiscibility in basaltic magma

The timing of silicate liquid immiscibility is crucial, and the earlier immiscibility starts, the greater its potential geochemical and petrogenetic role can be. Previous experimental studies, which used tholeiitic starting compositions (Philpotts 1979; Dixon and Rutherford 1979; McBirney and Nakamura 1974; McBirney and Naslund 1990), usually reported the onset of silicate liquid immiscibility at temperatures around 1,010–1,040°C, and only 5–10% of residual liquid. The evolved Upper Zone of the Skaergaard intrusion is believed to form at 1,000–1,085°C (e.g., McBirney and Naslund 1990; Ariskin 2002) and the degree of

crystallisation 0.8 (Nielsen 2004). The occurrence of the immiscible liquids in melt inclusions from the Upper Zone (Jakobsen et al. 2005) is thus in good agreement with the experiments. This study documents liquid immiscibility at temperature as high as 1,160°C (sample C-137) and FeO concentration in *Lfe* of around 21 wt.%, well achievable during crystallisation of the Skaergaard magma (Toplis and Carroll 1995). It should be noted however that the bulk composition SF-2 used in the experiment C-137 is not a perfect analogue of the Skaergaard magma, and the immiscible liquids in C-137 precipitate tridymite, but not olivine, plagioclase, pyroxenes and Fe–Ti oxides as should be in Skaergaard. Crystallisation experiments with the Skaergaard bulk composition (Toplis and Carroll 1995) did not report liquid immiscibility. Nevertheless, the occurrence of high-temperature unmixing in the sample C-137 demands that the compositional and temperature extent of liquid immiscibility in natural ferrobaltic liquids should be revisited.

Oxygen fugacity in our experiments in iron containers (Table 6) was too low compared to the conditions expected in vast majority of terrestrial tholeiitic magmas. Previous experimental studies (Naslund 1983; Philpotts 1979; Vicenzi et al. 1994) have shown however that more oxidised conditions favour liquid immiscibility, broaden the miscibility gaps, and increase the liquid–liquid D values. Thus, it is not obvious, which of the two opposing factors—increased stability of magnetite, which prevents strong Fe enrichment in melt (Toplis and Carroll 1995), or enhanced liquid immiscibility—would prevail at the evolved stages of crystallisation in natural ferrobaltic magmas of the Skaergaard type.

The ferrobalt–rhyolite silicate liquid immiscibility appears to be very sensitive to subtle changes in melt composition, and minor components, such as P_2O_5 , TiO_2 , S or CO_2 may determine the onset of immiscibility. The effects of volatile components may prove to be important in future experiments. In volatile-free compositions, pressure as such does not seem however to be an important factor. For instance, experiments at pressures up to 0.7 GPa did not show significant effects on liquid immiscibility in the model $\text{KAlSi}_2\text{O}_6\text{–Fe}_2\text{SiO}_4\text{–SiO}_2$ system (Schmidt et al. 2006).

Geochemical effects of silicate liquid immiscibility

Silicate liquid immiscibility should impose distinct geochemical imprints upon conjugate liquids, and the geochemical traces of the process may be retained in lavas and crystallised igneous rocks. Silica-rich immiscible liquids *Lsi* are similar to natural rhyolites, dacited

and granites, and it would be not easy to distinguish them from the silicic products of fractional crystallisation of basaltic magmas. Silicic rocks produced by immiscibility should be very depleted in all incompatible trace elements except Rb and Cs, and the strongest depletion is expected for light REE and HFSE. Elevated Ba/La values, or anomalously high Ta/Nb may be good indicators of liquid immiscibility in silicic rocks. Dixon and Rutherford (1979) proposed that silicic rocks produced by liquid immiscibility should also show high K/Ba, K/P and K/REE values. Much depends however on the effectiveness of separation of Fe-rich liquid from the pools of viscous silicic immiscible melt. The separation is difficult to achieve in centrifuge experiments, and good separation may prove to be rare in natural magmas. Any remnants of the iron-rich liquid in the silicic counterpart will obscure the geochemical indicators and the origin of the rock by immiscibility.

It has been pointed out in many previous works that contrary to the Si-rich liquid, the compositions of Fe-rich immiscible melts do not resemble any known igneous rocks, except exotic nelsonites or rare lava flows, such as the enigmatic Fe-rich lavas at the El Laco volcano, Chile (Naslund et al. 2003). This should not be considered however as a proof of rareness or insignificance of silicate liquid immiscibility in natural magmas. Dense Fe-rich liquids are not likely to erupt. Therefore, they would stay behind in magma chambers, or migrate downward in magmatic plumbing systems, and there may mix with new portions of hot, primitive mantle magmas. High FeO content and CaO/Al₂O₃ values, strong enrichment in P, Ti, REE and HFSE are the geochemical hallmarks of the Fe-rich immiscible liquids. The liquids may also have anomalous, non-chondritic Nb/Ta values, and high Zn and Cu contents. Because of the enrichment and the extreme composition, even small additions of such liquids should alter geochemical characteristics of primitive basalts. The crystallisation–immiscibility–mixing cycles in long-lasting and highly productive magmatic systems, such as mid-ocean ridges and mantle plumes, could eventually produce rare hybrid magmas enriched in the elements, which partition into the immiscible iron-rich melt. For example, it is notable that in rare and unusual ferropicrites, interpreted as products of heterogeneous plume mantle sources with iron-rich streaks (Gibson et al. 2000, 2002), high FeO contents correlate with elevated CaO/Al₂O₃ values, enrichment in P, Ti, and a number of incompatible trace elements. Such correlation is not readily explained by the mantle melting models, but mixing of evolved, immiscible, iron-rich liquid into picritic magma could be, in our view, a viable alternative hypothesis.

Acknowledgments Dieter Rhede, Oona Appelt (GFZ Potsdam), and John Spratt (NHM London) are thanked for the help with electron microprobe analyses. Experimental work was supported by the DFG SPP 1055 grant to DBD (Di 431/23-1). IVV visits to NHM London were financed by CERCAMS, and a grant from EU SYNTHESYS programme. Comments of two anonymous reviewers helped us to improve the original version of the manuscript.

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