

RESEARCH ARTICLE

# Influence of La<sup>3+</sup> Substitution on Structure, Morphology and Magnetic Properties of Nanocrystalline Ni-Zn Ferrite

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## Abstract

Lanthanum substituted Ni-Zn ferrite nanoparticles ( $\text{Ni}_{0.5}\text{Zn}_{0.5}\text{La}_x\text{Fe}_{1-x}\text{O}_4$ ;  $0.00 \leq x \leq 1.00$ ) synthesized by sol-gel method were presented. X-ray diffraction patterns reveal the typical single phase spinel cubic ferrite structure, with the traces of secondary phase for lanthanum substituted nanocrystals. In addition, the structural analysis also demonstrates that the average crystallite size varied in the range of 21–25 nm. FTIR spectra present the two prominent absorption bands in the range of 400 to 600  $\text{cm}^{-1}$  which are the fingerprint region of all ferrites. Surface morphology of both substituted and unsubstituted Ni-Zn ferrite nanoparticle samples was studied using FESEM technique and it indicates a significant increase in the size of spherical shaped particles with La<sup>3+</sup> substitution. Magnetic properties of all samples were analyzed using vibrating sample magnetometer (VSM). The results revealed that saturation magnetization ( $M_s$ ) and coercivity ( $H_c$ ) of La<sup>3+</sup> substituted samples has decreased as compared to the Ni-Zn ferrite samples. Hence, the observed results affirm that the lanthanum ion substitution has greatly influenced the structural, morphology and magnetic properties of Ni-Zn ferrite nanoparticles.

## OPEN ACCESS

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## Introduction

Ferrite materials of spinel structural formula  $\text{AB}_2\text{O}_4$  are mainly composed of about 70% iron oxide ( $\text{Fe}_2\text{O}_3$ ) and about 30% of other metal oxides including MgO, MnO, NiO, CuO, and or FeO [1, 2]. Metal oxide nanoparticles have been intensively studied both theoretically and practically due to the electric, dielectric, optical and magnetic properties that are significantly different as compared to the bulk materials [3, 4]. Generally, the physical properties of ferrite materials are governed by processing techniques, stoichiometry, and also the dispersion of cations among tetrahedral-A sites and octahedral-B sites. The preference of cation to occupy either A, or B sites are dependent on their ionic radius, crystal field, electronic configuration and also the ionic polarization [5–9].

Ni-Zn spinel ferrites nanoparticles found to be one of the versatile ferrites from the view of technological applications with high resistivity and low eddy current loss [10]. Zinc ferrite

bulk material possesses normal spinel structure with their divalent and trivalent cations are located on the tetrahedral and octahedral sites respectively. On the other hand, Nickel ferrite bulk materials have a spinel structure, where the trivalent cations occupy the A and B sites equally while all the divalent cations move to the octahedral site. Generally, this type of ferrites is called as a hopping semiconductor, where their conducting properties are the thermally activated hopping of electrons from one to another cation [10]. There is numerous attempt have been made in order to enhance the qualities of ferrites by incorporating the same suitable non-magnetic/diamagnetic impurities with different valence state at the A and B sites includes Copper [11, 12], Manganese [13], Praseodymium [14], Lanthanum [15], Neodimium [16] ion and etc [17]. Lanthanum is known as the second most abundant and lightest rare earth element (REE) in the lanthanide series. This silvery white mineral found in monazite and bastnasite ores. Lanthanum possesses distinct quality as compared to other REE such as simple electronic spectra which is helpful for experimental analysis; it has the highest boiling point and lowest vapour pressure at its melting point; and at atmospheric pressure lanthanum is the only superconducting REE [18]. Therefore, lanthanum is demand for some important application includes a) used in the manufacture of expensive glasses as lanthanum imparts a high refractive index to the glass [19]; b) used in NiMH batteries that are currently used in almost all hybrid-electric vehicles [20, 21]; c) Lanthanum rich compounds are used in alloy and fluid cracking catalyst for petroleum refinery industry [22]

Rare earth materials are known to possess good electrical insulation properties with high electrical resistivity. Therefore, the substitution of these rare earth ions into spinel ferrites could alter the electrical and magnetic properties. Moreover, these rare earth ions have a huge influence on the magnetic anisotropy of the system making the spinel ferrite as promising materials replacing the hexaferrite or garnets [16, 23]. The partial substitution of Fe<sup>3+</sup> ion by rare earth materials ion into the spinel ferrites structure leads to a structural distortion and induce strain, hence modifies the magnetic and electrical properties remarkably [24–26]. Generally, the rare earth ions occupy the octahedral B-sites with limited solubility in the spinel lattice structure due to their large ionic radius. However, based on the previously reported work the behavior of different rare earth ions are varied in spinel different ferrites and yield different structural properties [27].

Various methods have been employed in order to synthesize the soft spinel ferrite materials Including chemical co-precipitation [28], hydrothermal method [29], mechano-chemical method [30], microemulsion method [28], rheological phase reaction method [31], and also sol-gel method [7]. Wu et al. [32] synthesized Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>2-x</sub>O<sub>4</sub> using solid-state reaction at low temperatures for the first time. The results showed that the calcination temperature affects the magnetic properties of lanthanum substituted Ni-Zn ferrite, while the highest coercivity value of 120.09 Oe obtained at 800°C of calcination temperature. Ahmed et al. [10] have examined the structural and electrical properties of La<sup>3+</sup> substituted Ni-Zn ferrite prepared by standard ceramic method. The obtained results showed an appearance of small peaks which represents the secondary phase due to the addition of lanthanum ion. Wang et al. [15] prepared Ni<sub>0.8</sub>Zn<sub>0.2</sub>La<sub>x</sub>Fe<sub>2-x</sub>O<sub>4</sub> nanomaterial by sol gel method. The structural and morphological analysis of obtained samples shows that the addition of lanthanum ion has reduced the grain size. Further, lanthanum ion substitution has decreased the saturation magnetization (M<sub>s</sub>) from 69.81 to 46.75 emu/g. Among these techniques, sol-gel synthesis has been receiving much attention as they can be applied to an extremely wide variety of materials and also they offer the possibility of controlling the size, shape, and distribution of particles [33].

To the best of our knowledge, there aren't many research works focused on the detailed structural and morphological analysis of the lanthanum ion substitution into the spinel ferrite materials. Therefore, the current work is intended on investigating the effect of rare earth

lanthanum ion substitution and distribution within the A and B sublattices on the structural, magnetic (super exchange interactions and spin alignment) and morphological properties of Ni-Zn ferrite nanoparticles. The structural, magnetic and morphological properties of synthesized nanoparticles were examined using XRD, FTIR, FESEM and VSM analysis.

## Experimental

All chemicals used were of analytical grade with purity  $\geq 99\%$  and all the stock solution was prepared with distilled water. Ferric nitrate  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ , Citric acid  $\text{C}_6\text{H}_8\text{O}_7$ , Nickel nitrate  $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , Zinc nitrate  $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , Lanthanum nitrate  $\text{La}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  were used as the starting materials. La<sup>3+</sup> substituted Ni-Zn ferrite having the general chemical formula of  $\text{Ni}_{0.5}\text{Zn}_{0.5}\text{La}_x\text{Fe}_{1-x}\text{O}_4$ , where  $x = 0.0, 0.1, 0.2, 0.4, 0.6$ , was prepared through a typical sol-gel method as previously reported [34–36]. The nitrate salts of all the necessary elements for the final sample was used in stoichiometric amounts and were dissolved in 1M citric acid ( $\text{C}_6\text{H}_8\text{O}_7$ ) solution. The mixture was stirred with heating at 80°C for 24 hours in order to facilitate the gelation process. The resulting thick gel was then further dried in an oven at 110°C for another 24 hours. The resulting dried gel was then crushed into fine powders before being subjected to calcination at 900°C for 3 hours to complete the ferrite synthesis process. Citric acid acts as an organic fuel during the combustion by providing a platform for initiating the redox reactions between the reactants owed to its high heat of combustion [37].

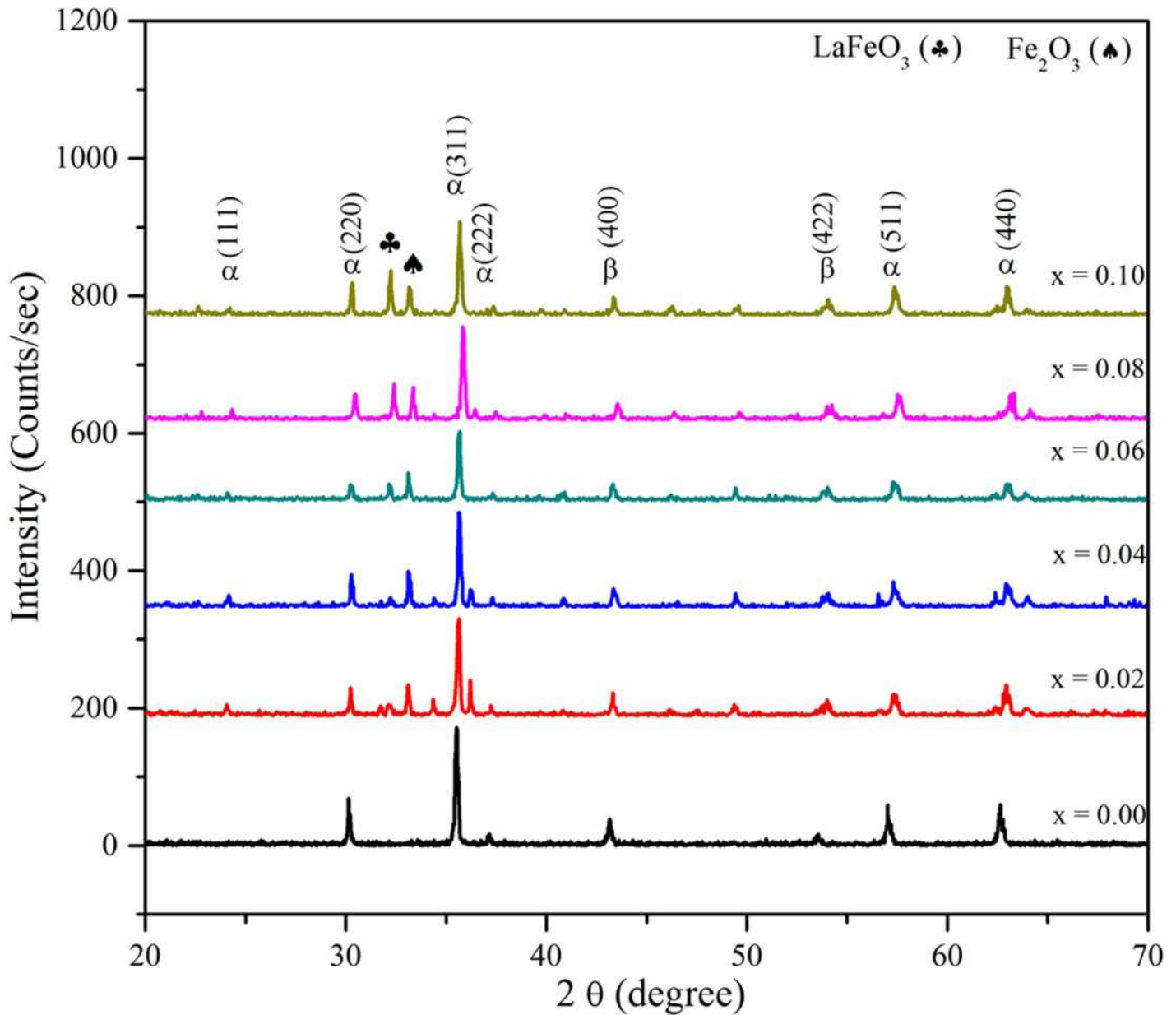
The crystal structure of  $\text{Ni}_{0.5}\text{Zn}_{0.5}\text{La}_x\text{Fe}_{1-x}\text{O}_4$  ferrite powder was characterized by x-ray diffraction (XRD) analysis by using a Bruker D8 powder x-ray diffraction system with Cu-K $\alpha$  radiation ( $\lambda = 1.5418 \text{ \AA}$ ) at scanning angles 10°–90°. Structural information such as the present phases, lattice parameter, and crystallite size were able to be extracted from the samples' diffraction patterns. The microstructure and morphology of substituted ferrite nanoparticles was examined using field emission scanning electron microscopy (FESEM model Zeiss Supra VP55) with an acceleration voltage of 20 kV. Further, fourier transform infrared spectroscopy (FTIR) was taken on Perkin-Elmer infrared spectrometer within the wave number range of 400–4000  $\text{cm}^{-1}$ . Magnetic measurements were performed by employing the vibrating sample magnetometer (VSM) with an applied field of 20 kOe to reach saturation and the magnetic hysteresis loops were measured at room temperature.

## Results and Discussion

A  $\text{Ni}_{0.5}\text{Zn}_{0.5}\text{La}_x\text{Fe}_{1-x}\text{O}_4$ , ( $x = 0.0, 0.1, 0.2, 0.4, 0.6$ ) nanocrystals was synthesized through a simple sol-gel method. The structural, morphological, and magnetic properties of Ni-Zn-La ferrite nanocrystals were examined for different composition of lanthanum in Ni-Zn ferrite nanocrystals. Further, the lanthanum ion substitution and distribution within the A and B sublattices of Ni-Zn ferrite nanocrystals was determined based on these observations.

### XRD analysis

XRD patterns of  $\text{Ni}_{0.5}\text{Zn}_{0.5}\text{La}_x\text{Fe}_{1-x}\text{O}_4$  nanocrystals, for all the samples with  $x = 0.00, 0.02, 0.04, 0.06, 0.08, \text{ and } 0.1$  are shown in Fig 1. The X-ray diffraction patterns reveal a single phase cubic spinel structure with few traces of secondary phase. Furthermore, the observed diffraction peaks could be assigned to the reflection plane of (111), (220), (311), (400), (422), (511) and (440) which could be indexed to a single-phase Ni-Zn ferrite nanocrystal ( $\alpha$ -phase). Besides that, the iron oxide ( $\beta$ -phase) was also detected in XRD pattern [28, 38, 39]. Meanwhile, the peak corresponding to  $2\theta = 32.21^\circ$  (200) is attributed to secondary phase at the grain boundaries for  $\text{LaFeO}_3$  ( $\clubsuit$ ) (ICDD PDF #74–2203) except for the cubic spinel phase. The intensity of  $\text{LaFeO}_3$  peak has increased with the increase in La<sup>3+</sup> ion concentration. Furthermore, another secondary phase of



**Fig 1. Structural and grain size analysis of La<sup>3+</sup> substituted Ni-Zn ferrite nanoparticle.** XRD diffraction patterns for the sample of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystalline (α: Ni-Zn ferrite; β: iron oxide) at various La<sup>3+</sup> substituted concentration.

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Fe<sub>2</sub>O<sub>3</sub> (♠) 2θ = 33.22° (ICDD PDF #86-2368) was detected with an increase in lanthanum ion concentration [40–43]. The appearance of secondary phase of a) LaFeO<sub>3</sub> (♣) in the Ni-Zn-La ferrite nanoparticles demonstrates that La<sup>3+</sup> ion substitution has a solubility limit in the spinel lattice; b) Fe<sub>2</sub>O<sub>3</sub> (♠) in the Ni-Zn-La ferrite nanoparticles shows that there was more than fifty mole percentage of the standard iron oxide content in the formula of spinel ferrite [38].

The lattice parameter of individual composition was investigated by using the following relation:

$$a = d_{hkl} \sqrt{h^2 + k^2 + l^2} \tag{1}$$

**Table 1. Lattice parameters, interplane distance and crystallite size of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystalline.**

Samples	Lattice constants	Interplane Distance	Crystallite Size (nm)
	a (Å)	(d Å)	
X = 0.00	8.557	2.528	22.42
X = 0.02	8.357	2.520	23.30
X = 0.04	8.347	2.517	24.22
X = 0.06	8.351	2.518	24.41
X = 0.08	8.308	2.505	22.76
X = 0.10	8.345	2.516	21.05

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Where, ‘a’ is the lattice constant, d is the inter-planar distance and (h, k, l) are the Miller indices [44].

Whereas the average crystallite diameter of the powder estimated from the intense (311) peak of XRD diffraction pattern using the Scherrer’s Eq (2):

$$D = \frac{0.9\lambda}{B \cos \theta} \tag{2}$$

Where D is the average crystalline dimension perpendicular to the reflecting phases, λ is the wavelength of the X-ray source which is taken as 1.5406 Å, and B is the full width at half maximum (FWHM) and θ is the Bragg’s angle [45, 46]. The calculated values for lattice constant, planner distance and crystallite size at (311) peaks are tabulated in Table 1.

The calculated crystallite size of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystals for this work is found to be smaller than the previously reported work where the crystallite size for the lanthanum substituted nanocrystals ranged from 22 nm to 24 nm. The calculated values in Table 1, show that the crystallite size of prepared ferrite samples increases with an increase in the amount of La<sup>3+</sup> ions concentration up to X = 0.06, due to the larger ionic radii that that of Fe<sup>3+</sup> ions which induces the grain growth [47, 48].

However, Table 1 shows the gradual decrease in the crystallite size value with an increase in La<sup>3+</sup> substitution at above X ≥ 0.06 which is in agreement with the results observed in the literature [49]. There are two possible explanations to account for this decrease in crystallite size with an increase in La<sup>3+</sup> substitution concentration, a) Firstly, the large size mismatch between La<sup>3+</sup> and Fe<sup>3+</sup> induces the crystalline anisotropy during the substitution of ions which creates the strain inside the volume of the crystals with an increase in La<sup>3+</sup> substitution. Furthermore, by balancing the crystal anisotropy and volume strain to each other the present system could remain in stable equilibrium. Therefore, the crystallite size decreases with an increase in La<sup>3+</sup> substitution concentration in order to reduce the volume strain; b) Secondly, in the crystal structure of La<sup>3+</sup> substituted Ni-Zn ferrite, Zn<sup>2+</sup> and Ni prefers tetrahedral (A-site) and octahedral (B-site) sites respectively. On the other hand, Fe<sup>3+</sup> occupies either tetrahedral or octahedral site while La<sup>3+</sup> ions have a tendency to enter into the octahedral site. However, it is difficult for lanthanum ion to substitute Fe<sup>3+</sup> as La<sup>3+</sup> has much larger ionic radius than that of Fe<sup>3+</sup> and requires more activation energy to enter octahedral sites as the bond energy of La-O is higher as compared to Fe-O. Therefore, instead of occupying the Fe<sup>3+</sup> sites in the lattice the La<sup>3+</sup> enter the lattice at interstitial sites due to extra stress coming from partial lanthanum ion which suppresses the crystallization of Ni-Zn ferrites [50, 51].

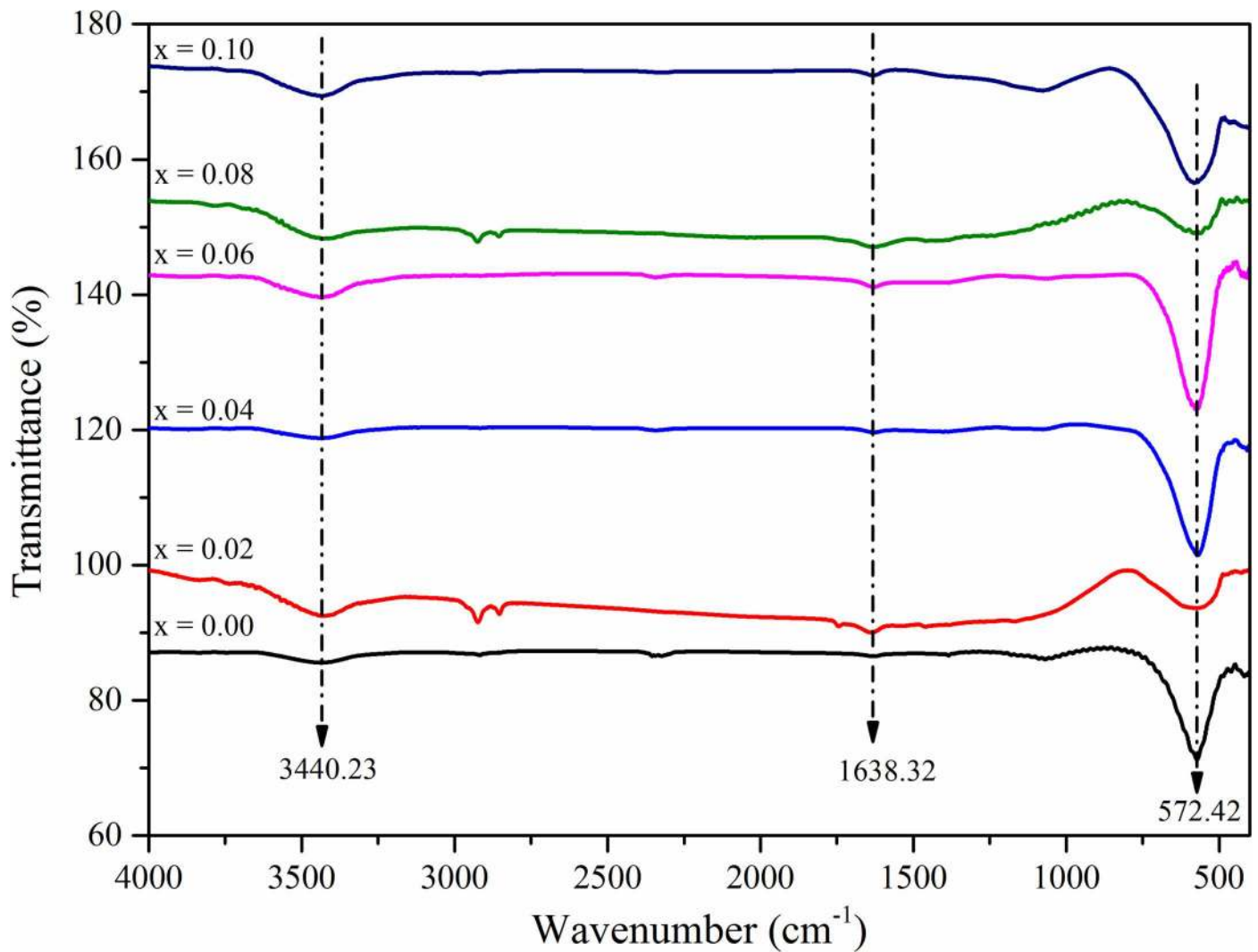
There is slight changes were observed for lattice constant ‘a’ which could be associated with the increasing La<sup>3+</sup> content where the ionic radius of La<sup>3+</sup> (1.6061 Å) is bigger than that of Fe<sup>3+</sup> ion (0.645 Å) replacing iron ions on octahedral B-site which causes asymmetry in the structure. Hence, the lattice constant should be aggrandized with the increasing content of the La<sup>3+</sup> during



the substitution process [15, 52]. Besides that, the decrease in lattice constant with the increase of La<sup>3+</sup> content could be attributed to the compression of spinel lattice induced by the secondary phases due to the difference in thermal expansion coefficients [42, 53].

### FTIR analysis

Fig 2. Show the FTIR spectra of the investigated Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystalline samples at increasing La<sup>3+</sup> concentration. FTIR spectra demonstrate the presence of two fundamental absorption bands which are attributed to the normal and inverse cubic spinel ferrites. Generally the higher frequency absorption band 'ν<sub>1</sub>' appears in the range of 500 to 600 cm<sup>-1</sup> represents the intrinsic vibration of tetrahedral group, while the octahedral groups are represented by the lower frequency absorption band 'ν<sub>2</sub>' in the range of 350–490 cm<sup>-1</sup> [54, 55]. Since the tetrahedral site dimensions are less as compared to the octahedral, the band frequency of 'ν<sub>1</sub>' is higher than the 'ν<sub>2</sub>' absorption band [56]. The characteristic bands that appear at wavenumber 572.42 cm<sup>-1</sup> attributed to the metal-oxygen stretching vibration of Fe<sup>3+</sup>-O<sup>2-</sup> [57]. Further, the



**Fig 2. FTIR spectroscopic analysis to identify the types of chemical bonding in La<sup>3+</sup> substituted Ni-Zn ferrite nanoparticle.** FTIR spectra for the samples of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystalline at various La<sup>3+</sup> substitution concentrations.

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spectrum shows absorption bands at 1638.32 cm<sup>-1</sup> corresponding to NO<sub>3</sub><sup>-1</sup> ions, and also due to the presence of carboxyl group (COO<sup>-</sup>) [46]. The broadband that appears at 3440.23 cm<sup>-1</sup> affiliated to hydrogen bonded O-H stretching vibration [58]. The difference in intensity of 'ν<sub>1</sub>' and 'ν<sub>2</sub>' absorption bands is due to the changes in the bond length of Fe<sup>3+</sup>-O<sup>2-</sup> at tetrahedral A-sites and octahedral B-sites [2, 4].

It can be observed from the IR spectrum that the absorption band 'ν<sub>1</sub>' intensity increases with an increase in La<sup>3+</sup> ion substitution concentration, while band 'ν<sub>2</sub>' does not show any changes. The variation in the intensity of 'ν<sub>1</sub>' is caused by the nature of lanthanum ion which prefers to occupy the octahedral B-site and forces the iron ion into A-site with an increase in the La<sup>3+</sup> concentration, which increases the radius of B-site. However, with an increase in lanthanum ion concentration above X ≥ 0.06 leads to decrease in the 'ν<sub>1</sub>'. Additionally, it is known that the variation in the position of IR spectra peaks is mainly dependent on the amount of substituent, method of preparation, grain size and also the density of particles. Moreover, the intensity is the ratio of change of dipole moment with the inter-nuclear distance (dμ/dr) [4, 59].

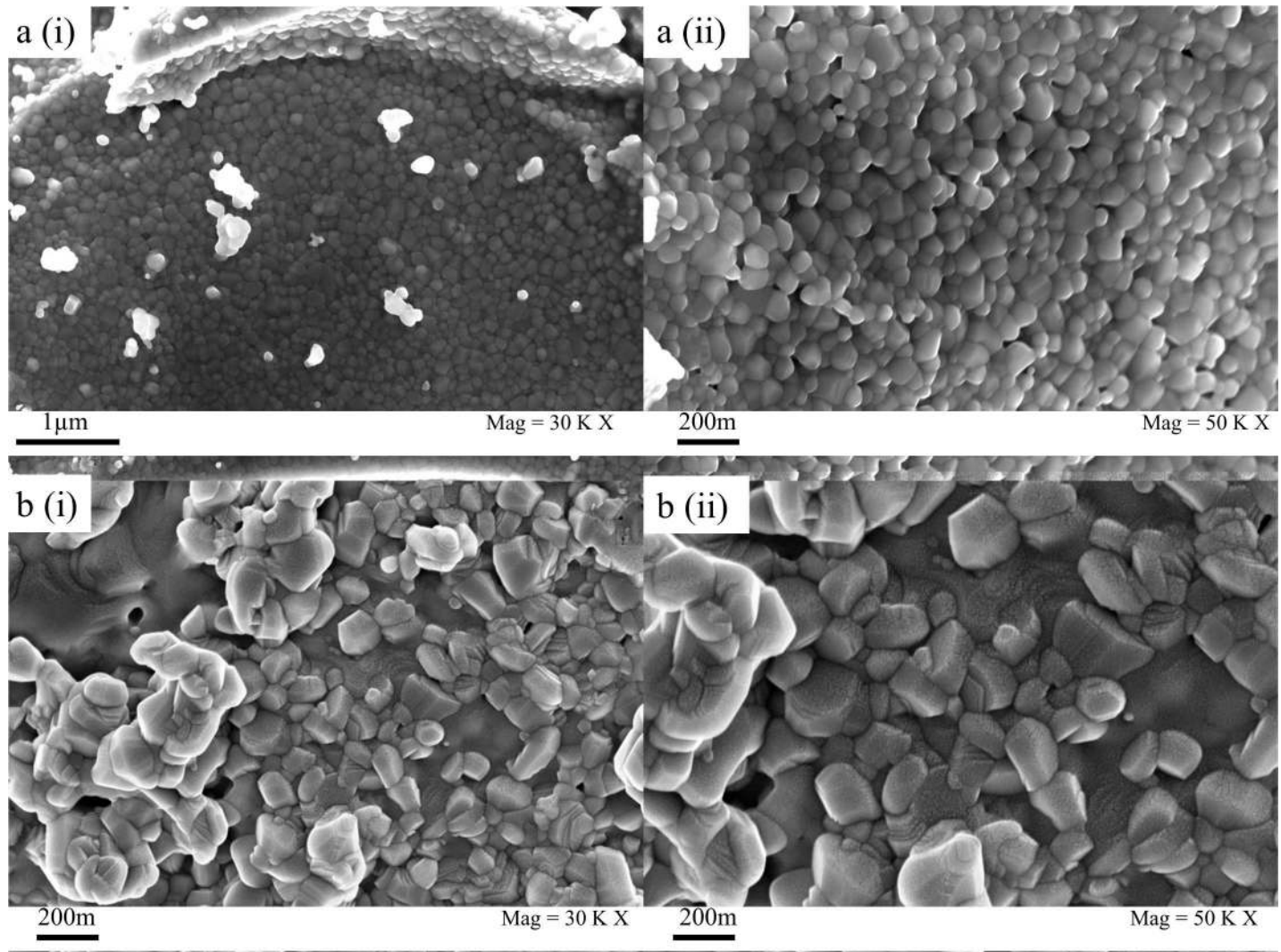
### Morphology analysis

The FESEM images of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanoparticles (0.0 ≤ X ≤ 0.10) are shown in Fig 3. Microstructure of Ni-Zn ferrite (X = 0.00) exhibits an angular morphology. Meanwhile, the microstructure of the particles became irregular in size and shape with La<sup>3+</sup> substitution [60]. The particles tend to agglomerate as they encounter a permanently magnetized moment proportional to their volume. Moreover, the high calcination temperature can also attribute to the agglomeration of particles due to emerging forces including Van-der Waals, capillary and electrostatic forces which create mutual interaction between particles [61, 62]. It can be observed from Fig 3 that the agglomeration is increasing with an increasing lanthanum ion concentration [48, 63]. Besides that, the larger ionic radius of lanthanum ion could attribute to the increase in grain size which occupy on the grain boundary on the outlay of cations vacancies in the lattice [64]. However, the average crystallite sizes of the substituted ferrite samples obtained from XRD analysis are significantly smaller than the values determined by FESEM analysis. The differences in average crystallite size of nanoparticles are associated with the fact that the value observed by FESEM analysis has the size of the secondary particles, where they are composed of several crystallites by the soft reunion. Moreover, the X-ray line broadening analysis reveals only the size of a single crystallite [32].

The composition of elements (%) and their atomic percentage of substituted ferrite nanoparticles were measured by elemental diffraction spectrum (EDS). Table 2 confirms the presence of Fe, Ni, Zn, La, and O in the sample which is used to further validate that lanthanum doped into Ni-Zn ferrite nanocrystals. Furthermore, the recorded values in Table 2 shows a gradual increase in Lanthanum composition which is in line with the synthesis condition. On the other hand, the detected traces of Carbon are mainly due to the coating applied to the sample prior to testing [65].

### Magnetic property

The magnetic hysteresis loop of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystals (0.0 ≤ X ≤ 0.10) is given in Fig 4. The measurements for magnetic parameters such as magnetization saturation (M<sub>s</sub>), coercivity (H<sub>c</sub>) and remanence (M<sub>r</sub>) were analyzed based on this hysteresis loop and presented in Fig 5. It can be observed from the Fig 4, that the value of saturation magnetization (M<sub>s</sub>) for Ni-Zn ferrite sample is very high as compared to the La<sup>3+</sup> ion substituted samples due to high crystallinity and uniform morphologies which can be observed from Fig 3. This is also due to the substitution of nonmagnetic lanthanum ion which leads to the formation of the nonmagnetic



**Fig 3. Morphological analysis of La<sup>3+</sup> substituted Ni-Zn ferrite nanoparticle.** FESEM micrographs of typical a) Ni-Zn ferrite (i) 30KX (ii) 50KX and b) Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> at (x = 0.02) (i) 30KX (ii) 50KX.

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spinel structure. In addition to this, the substitution of La<sup>3+</sup> ion causes the transformation of the collinear ferromagnetic order into non-collinear arrangement and disintegration of ferromagnetic order that reduces the M<sub>s</sub> [50].

As the lanthanum ion concentration increases (X = 0.02–0.06) in Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystals, the saturation magnetization shows a gradual increase due to the increase in the crystallite size and minimization of the surface effect [66]. However, the saturation magnetization decreases with further increase in lanthanum concentration at above X ≥ 0.06. The decrease in saturation magnetization can be explained based on two factors: (a) crystallite size and (b) site occupancy of the cations plus the modification in the super exchange effects caused by substituting La<sup>3+</sup> [66]. As the crystallite size decreases (X ≥ 0.06), the ratio of surface to volume atom increases which results in the prominent surface effect. Surface atoms are under the effect of the strain due to the distorted surface structure thus creates vacancies, a variety of the interatomic spacing and low coordination numbers. As a consequence, these factors could induce broken exchange bonds for the surface atoms which results in spin disorder. The



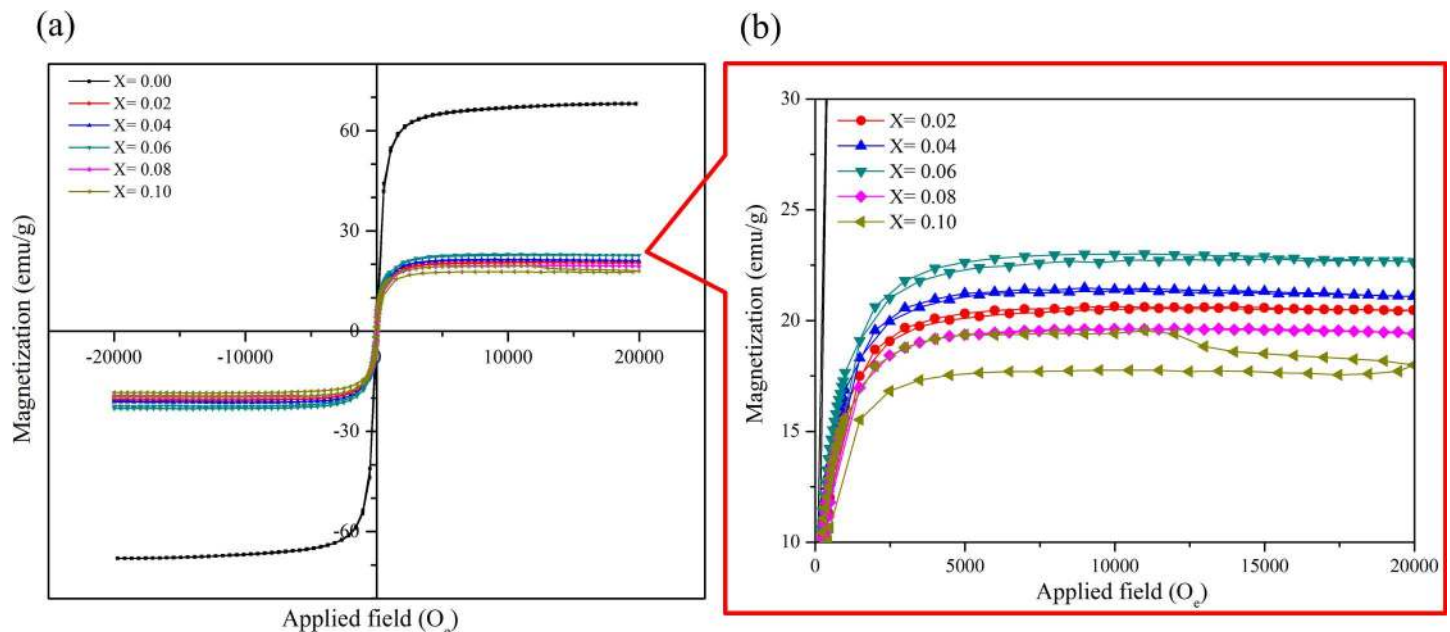
**Table 2. The composition of elements of Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystals at various La<sup>3+</sup> substitution concentrations.**

Elements (K)		X = 0.02	X = 0.04	X = 0.06	X = 0.08	X = 0.10
C	Weight (%)	5.90	-	-	-	-
	Atomic (%)	16.07	-	-	-	-
O	Weight (%)	20.99	20.70	21.65	26.75	21.85
	Atomic (%)	42.94	49.08	50.56	57.89	51.52
Fe	Weight (%)	49.05	51.01	51.73	42.90	44.20
	Atomic (%)	28.74	34.64	34.61	26.60	29.86
Ni	Weight (%)	9.93	11.11	13.55	14.82	13.53
	Atomic (%)	5.54	7.18	8.62	8.74	8.69
Zn	Weight (%)	12.77	14.36	8.89	10.34	14.34
	Atomic (%)	6.39	8.33	5.08	5.47	8.28
La	Weight (%)	1.35	2.82	4.18	5.20	6.08
	Atomic (%)	0.32	0.77	1.12	1.30	1.65

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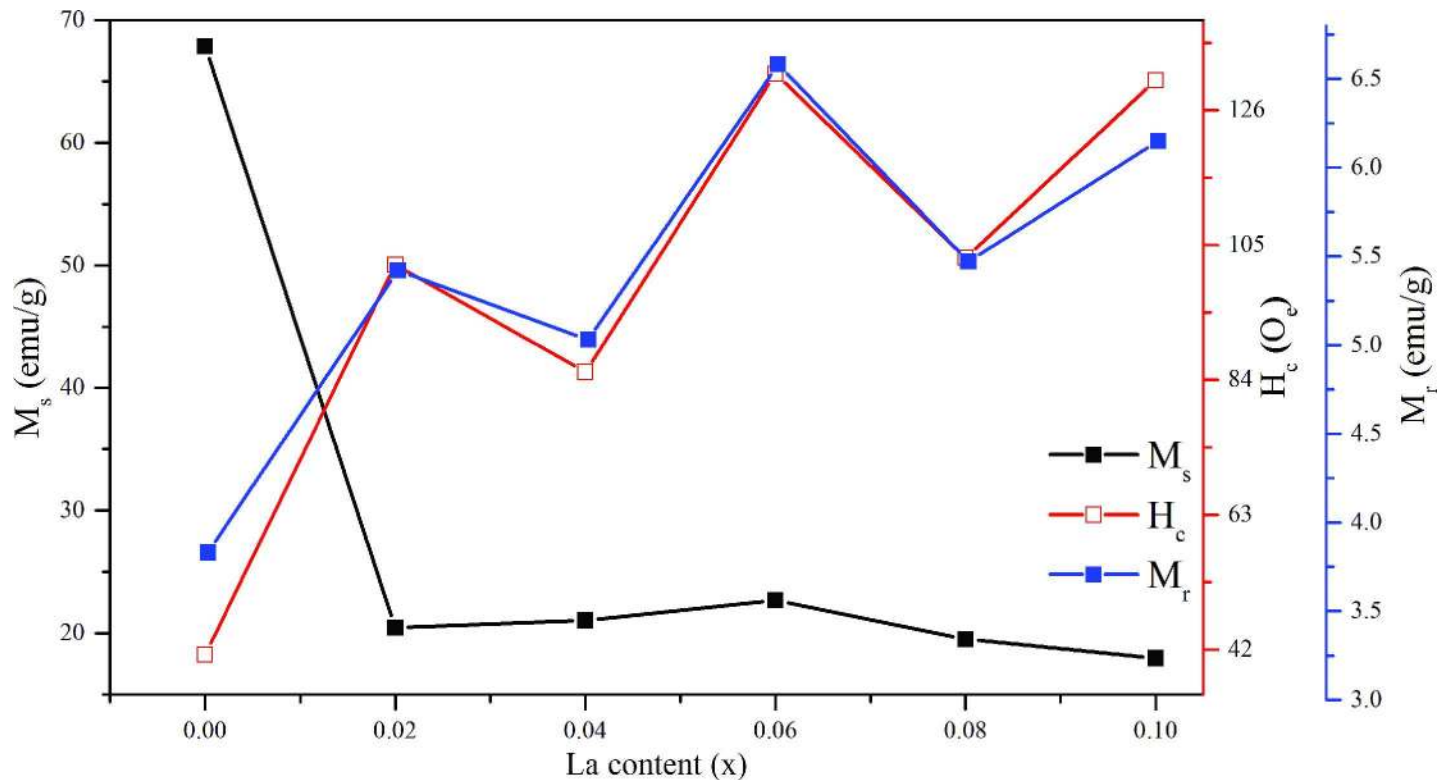
disordered spins at the surface demonstrate low magnetization. Furthermore, the disordered spins increase with the decrease in crystallite size, hence reduces the saturation magnetization [66, 67].

In addition, the observed reduction of the saturation magnetization can be described based on the cation dispersion and exchange interaction between iron and between lanthanum ions at tetrahedral A and octahedral B sites [68]. Generally, for spinel ferrite, the saturation magnetization is influenced by magnetic super exchange interaction of cations between tetrahedral A and octahedral B sites. The iron ions from B-sites are the dominant contributor to the magnetization. When lanthanum ions are introduced at the expense of iron ion, some of the Fe<sup>3+</sup> migrates from B- to A-sites in view of the site preferences for different ions which lead to the increase of Fe<sup>3+</sup> concentration at A-sites. The extended radius of lanthanum ion and the small



**Fig 4. Magnetic properties of La<sup>3+</sup> substituted Ni-Zn ferrite nanoparticle hysteresis loops.** Hysteresis loops of (a) Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanocrystals at various La<sup>3+</sup> substitution concentrations (b) Enlarged view of hysteresis loop.

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**Fig 5. Magnetic properties of La<sup>3+</sup> substituted Ni-Zn ferrite nanoparticle p.** Variation of saturation magnetization ( $M_s$ ) coercivity ( $H_c$ ) and remanence ( $M_r$ ).

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tetrahedral A site gap pressurize the La<sup>3+</sup> to enter the octahedral B site. The B-B exchange interaction at the octahedral site reduces as a result of decreasing iron ion concentration which will eventually reduce the A-B super-exchange interaction too. The weakening of A-O-B super exchange interactions attributes to the decrease of saturation magnetization. Moreover, La<sup>3+</sup> ion has no electrons in 4f orbit which makes it paramagnetic. Therefore, when the paramagnetic La<sup>3+</sup> ion replaces the Fe<sup>3+</sup> ion at B-site it leads to the decrease in the net overall magnetic moment. Further, the substitution of paramagnetic La<sup>3+</sup> ion leads to distortion of the spinel lattice thus changes the spin magnetic moment between the A-B sublattices into nonlinear antiferromagnetic coupling [15, 69, 70]. The above discussion confirmed the reduction in the magnetization of prepared ferrite nanocrystals samples at lanthanum ion concentration at above  $X \geq 0.06$ .

Generally, the coercivity value portrays the strength of the magnetic field that is required to surpass the anisotropy barrier and allow the magnetization of the Ni<sub>0.5</sub>Zn<sub>0.5</sub>La<sub>x</sub>Fe<sub>1-x</sub>O<sub>4</sub> nanoparticles following the magnetic field orientation. From the Fig 3 it can be observed that the coercivity and remanence values have the same trend with respect to the lattice constants (Å) where these values are not consistent with the increase in La<sup>3+</sup> ion concentration. The coercivity governs by factors such as magneto crystalline anisotropy, microstrain, magnetic particle morphology, size distribution, shape anisotropy, and magnetic content [68, 71].

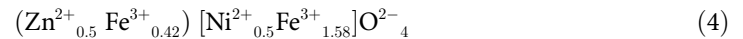
On the basis of the above explanation, the predicted distribution of ions for the system under investigation is proposed in the following outline:

Generally, Ni<sup>2+</sup> ion has a strong tendency to occupy octahedral site (B-site), and Zn<sup>2+</sup> ion occupy a tetrahedral site (A-site) only. On the contrary, Fe<sup>3+</sup> ions are distributed between the

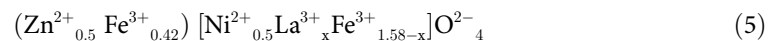
two sites. Therefore, the cation distribution of Ni-Zn spinel ferrite can be represented by the following formula [4].



where the brackets () and [] demonstrates A and B sites respectively. The preference of ions of basic composition Ni<sub>0.5</sub>Zn<sub>0.5</sub>Fe<sub>2</sub>O<sub>4</sub> into the A and B sites are influenced by theoretical formula as follows:



Rare earth ions are known to occupy the octahedral B-sites. Therefore, lanthanum ion has an initial preference for the octahedral site, where it replaced Fe<sup>3+</sup> ions in the B-sites as per the formula below.



The total magnetization decreases with an increasing La<sup>3+</sup> concentration, thus it proves that the substitution of paramagnetic lanthanum ion into ferromagnetic iron ions in spinel does not improve the magnetic properties of the synthesized nanoparticles [72].

## Conclusion

The present study demonstrates the effect of La<sup>3+</sup> substitution for Fe<sup>3+</sup> in Ni-Zn ferrite nanocrystals morphology, structure, and magnetic properties. The substitution of lanthanum ion was found to affect the structure, morphology and magnetic properties of Ni-Zn ferrite nanocrystals. Structural analysis of La<sup>3+</sup> substituted nanoparticles showed a cubic spinel ferrite with the presence of small peaks attributed to secondary phase due to La<sup>3+</sup> substitution. The crystallite size of prepared nanoparticles has increased with an increase in lanthanum ion concentration of X ≤ 0.06 due to large ionic radii of La<sup>3+</sup> ion. Further, the saturation magnetization of substituted samples decreases with an increase of lanthanum ion concentration. This is due to dilution of the net magnetic field by paramagnetic La<sup>3+</sup> ion. While the spectroscopic analysis of these materials demonstrates the changes in the frequency intensities due to the addition of La<sup>3+</sup> ion which affects the bond length between Fe<sup>3+</sup>-O<sup>2-</sup> at both the A and B sites. Finally, the observed results have indicated that the rare earth lanthanum ions prefer to occupy the octahedral B-site of spinel ferrites.

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