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Investigation of the $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$ system at 25°C by thermodynamic calculation

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The phase diagram of the $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$ system relevant to cement has been calculated at 25°C and 1 bar pressure. It is composed of 331 stable phase assemblages in equilibrium with the aqueous phase and contains 30 invariant points with 15 solids. Thaumasite appears to be a stable phase with a wide stability domain. Compared to ettringite, thaumasite is stable to lower pH and generally demands higher sulphate concentrations to be stable.

Introduction

Thaumasite ($\text{Ca}_6\text{Si}_2(\text{SO}_4)_2(\text{OH})_{12}(\text{CO}_3)_2 \cdot 24\text{H}_2\text{O}$) and ettringite ($\text{Ca}_6\text{Al}_2(\text{SO}_4)_3(\text{OH})_{12} \cdot 26\text{H}_2\text{O}$) formation in mortars and concrete are related to the interaction of aqueous sulphates and carbonates with the solid components of cement matrices. However, the exact conditions under which the thaumasite form of sulphate attack (TSA) occurs are still not fully known, mainly due to the lack of basic knowledge on thaumasite such as its temperature range of stability, its pH stability and the nature of other phases compatible with thaumasite. Thaumasite is generally considered to need low temperatures to form at appreciable reaction rates;¹ however, field occurrences of thaumasite have been reported in areas known for their mild climates.²

The main objective of this work consists of a thermodynamic investigation of the $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$ system at 25°C . This temperature has been selected because thermodynamic data are either available for 25°C or, as for thaumasite, can be extrapolated to 25°C using literature data. Also, several important constituent subsystems have already been investigated at 25°C : $\text{CaO}-\text{Al}_2\text{O}_3-\text{CaSO}_4-\text{H}_2\text{O}$,³ $\text{CaO}-$

$\text{Al}_2\text{O}_3-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$,⁴ $\text{CaO}-\text{SiO}_2-\text{CaSO}_4-\text{H}_2\text{O}$,⁵ $\text{CaO}-\text{SiO}_2-\text{CaCO}_3-\text{H}_2\text{O}$,⁵ and $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{H}_2\text{O}$.⁶ More complex systems have been partially studied such as the eight-component system $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{Fe}_2\text{O}_3-\text{MgO}-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$.⁷ However, in this latter case the model did not include such important stable phases as hydrogarnet and hydrotalcite and, as a consequence, the results are incomplete.

Method of calculation

Solubility determination

It is necessary to know the solubility product of thaumasite in order to determine the stability conditions of thaumasite by thermodynamic calculations. Thaumasite was first synthesised and then redispersed in water at different temperatures ranging from 5 to 30°C , as reported by Edge and Taylor.⁸ Analyses of constituent ion concentrations after several redispersions of thaumasite were obtained by standard analytical chemical techniques. Thaumasite and ettringite can each undergo ionic substitutions and a partial solid solution exists in the thaumasite-ettringite system. Although ettringite and thaumasite are partially miscible with up to 50 mol % of the thaumasite component soluble in ettringite, much less ettringite is soluble in thaumasite. The authors have therefore considered only the 'pure' thaumasite end member in this study. Data for the solid solution, showing negligible solid solution of ettringite in thaumasite but substantial solid solution in ettringite, will be reported in the near future.⁹

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Thermodynamic database

In order to be consistent with the previous calculations obtained on the subsystems, the same methodology and the same computer program, PHRQPITZ,¹⁰ which calculates the activity coefficients using Pitzer's relations,¹¹ has been used. The method is based on the minimisation of free energy and the calculation outputs include, among other data, the invariant points of the system, tabulation of solids present and the composition of the aqueous coexisting phase. The accuracy of the results mainly depends on the reliability of the equilibrium constants and inclusion of all relevant solids; missing solids not included in calculations obviously have an important effect on the results.

The CaO–SiO₂–H₂O system is certainly one of the most difficult to model due to the variable composition and metastability of C–S–H relative to crystalline calcium silicate hydrates. Although recent progress has been made in a physico-chemical treatment of C–S–H¹², the authors have kept to the approach used to calculate the CaO–Al₂O₃–SiO₂–H₂O system which models C–S–H in terms of two fixed compositions, designated C–S–H(SI) and C–S–H(SII).¹³ C–S–H(SI) has a low C/S ratio (fixed C/S corresponding to an average value of 1.1) and is stable in solution having lime concentrations typically lower than 20 mmol/l at 25°C. On the other hand, C–S–H(SII) has a higher C/S ratio, fixed at 1.8 and is stable in solutions with lime concentrations greater than 20 mmol/l at 25°C. Similarly, the authors have modelled the hydrogrossularite–grossularite garnet solid solutions by designating two hydrogarnet compositions:¹⁴ a low SiO₂-containing hydrogarnet (HG Si low; C₃AS_{0.3}H_{5.4}); and a higher SiO₂-containing hydrogarnet (HG Si high; C₃AS_{0.8}H_{4.4}).

Knowledge of the thermodynamic properties of the

constituent phases of cement systems is still far from perfect and the present set of calculations cannot be more accurate than the numerical values of the database used in the calculation. However, the authors have taken account of the variation in composition of the C–S–H phase as well as that of the hydrogarnet phase insofar as data are presently available with which to assess the relevant thermodynamic values. The consequences of the simplifications employed are discussed subsequently.

The values of the solubility products are not given here because they mainly depend on the method used to calculate them; for example, the species used to write the equation of dissolution, methods used to calculate activity coefficients. However, the composition of the solution in equilibrium with each solid used in this work is given in Table 1 and with these data the reader can, if desired, calculate a solubility product appropriate to a particular system by defining solubility equations and activity corrections.

Results

The calculated CaO–Al₂O₃–SiO₂–CaSO₄–CaCO₃–H₂O system at 25°C contains 30 invariant points (Table 2). It is difficult to represent graphically this phase diagram due to its high dimensionality. In order to describe it as accurately as possible, the stable phase assemblages deriving from the invariant points have been listed. The system has 331 stable phase assemblages in equilibrium with the aqueous phase, comprising 99 assemblages of four solids (Table 3), 121 assemblages of three solids (Table 4), 66 assemblages of two solids (Table 5) and finally 15 stable solids. It is also of interest to know which phases are *not* stable together. Key assemblages of two phases that are not compatible are listed in Table 6. Phase assemblages of

Table 1. Calculated composition and pH of the aqueous solution in equilibrium with selected solids in water, 25°C, 1 bar pressure

	Ca: mM/kg	Al: mM/kg	Si : mM/kg	SO4: mM/kg	CO3: mM/Kg	pH
C–S–H(SI)*	1.228	—	1.116	—	—	11.03
C–S–H(SII)*	5.212	—	2.895	—	—	11.69
SH amorphous	—	—	1.413	—	—	6.37
C ₂ ASH ₈	0.769	0.769	0.384	—	—	10.59
Hydrogarnet Si low*	2.926	1.951	0.292	—	—	11.49
Hydrogarnet Si high*	2.207	1.472	0.588	—	—	11.29
C ₃ AH ₆	3.711	2.474	—	—	—	11.64
CH	21.95	—	—	—	—	12.54
AH ₃ (gibbsite)	—	3.25E-6	—	—	—	6.99
Calcium monocarboaluminate†	2.561	1.280	—	—	0.640	11.37
Calcium hemicarboaluminate†	4.110	2.055	—	—	0.514	11.65
Calcite	0.146	—	—	—	0.146	9.96
Ettringite	1.647	0.549	—	0.824	—	11.00
Gypsum	15.26	—	—	15.26	—	7.07
Thaumasite	0.7144	—	0.2382	0.2382	0.2382	10.50

* See text for definition of the formulae of these phases.

† Calcium hemicarboaluminate, 4CaO·Al₂O₃·0.5CO₃·12H₂O and calcium monocarboaluminate, 4CaO·Al₂O₃·CO₃·11H₂O.

Table 2. Composition of the aqueous phase at the invariant points sorted on increasing aqueous $[SO_4]$ (concentration in mol/kg)

No.	Phase Assemblages	Ca	Al	Si	CO ₃	SO ₄	pH
1	PORTLAND C ₃ AH ₆ SIHGlow ETTRINGI HEMICARB	2.20×10^{-2}	7.59×10^{-6}	1.16×10^{-5}	1.63×10^{-6}	7.60×10^{-6}	12.52
2	C ₃ AH ₆ SIHGlow ETTRINGI MONOCARB HEMICARB	1.73×10^{-2}	1.41×10^{-5}	6.95×10^{-6}	2.86×10^{-6}	7.81×10^{-6}	12.43
3	PORTLAND SIHGlow CSH(II) ETTRINGI HEMICARB	2.20×10^{-2}	6.69×10^{-6}	2.70×10^{-5}	2.70×10^{-6}	8.27×10^{-6}	12.52
4	SIHGlow CSH(II) ETTRINGI MONOCARB HEMICARB	2.02×10^{-2}	7.89×10^{-6}	3.21×10^{-5}	4.08×10^{-6}	8.64×10^{-6}	12.49
5	SIHGlow CSH(I) CSH(II) ETTRINGI MONOCARB	1.97×10^{-2}	8.28×10^{-6}	3.38×10^{-5}	4.21×10^{-6}	8.76×10^{-6}	12.48
6	PORTLAND CSH(II) ETTRINGI MONOCARB HEMICARB	2.20×10^{-2}	5.75×10^{-6}	2.70×10^{-5}	4.95×10^{-6}	9.15×10^{-6}	12.52
7	C ₃ AH ₆ SIHGlow GIBBSITE ETTRINGI MONOCARB	5.62×10^{-3}	3.47×10^{-4}	6.29×10^{-7}	3.14×10^{-6}	1.04×10^{-5}	11.97
8	CALCITE PORTLAND CSH(II) ETTRINGI MONOCARB	2.20×10^{-2}	4.28×10^{-6}	2.70×10^{-5}	8.93×10^{-6}	1.11×10^{-5}	12.52
9	CALCITE CSH(I) CSH(II) ETTRINGI MONOCARB	1.97×10^{-2}	5.68×10^{-6}	3.38×10^{-5}	8.96×10^{-6}	1.13×10^{-5}	12.48
10	CALCITE SIHGlow CSH(I) ETTRINGI MONOCARB	7.79×10^{-3}	7.34×10^{-5}	5.78×10^{-5}	9.53×10^{-6}	1.37×10^{-5}	12.11
11	CALCITE SIHGlow SIHGhigh CSH(I) ETTRINGI	7.06×10^{-3}	9.39×10^{-5}	6.24×10^{-5}	9.63×10^{-6}	1.45×10^{-5}	12.07
12	C ₂ ASH ₈ SIHGlow GIBBSITE ETTRINGI MONOCARB	4.90×10^{-3}	3.02×10^{-4}	2.05×10^{-5}	9.91×10^{-6}	1.60×10^{-5}	11.91
13	CALCITE C ₂ ASH ₈ SIHGlow ETTRINGI MONOCARB	4.98×10^{-3}	2.82×10^{-4}	2.28×10^{-5}	1.01×10^{-5}	1.60×10^{-5}	11.92
14	CALCITE C ₂ ASH ₈ SIHGlow SIHGhigh ETTRINGI	5.34×10^{-3}	2.17×10^{-4}	3.50×10^{-5}	9.99×10^{-6}	1.60×10^{-5}	11.95
15	CALCITE C ₂ ASH ₈ GIBBSITE ETTRINGI MONOCARB	4.88×10^{-3}	3.01×10^{-4}	2.07×10^{-5}	1.01×10^{-5}	1.61×10^{-5}	11.91
16	CALCITE C ₂ ASH ₈ SIHGhigh CSH(I) ETTRINGI	5.62×10^{-3}	1.43×10^{-4}	7.62×10^{-5}	9.92×10^{-6}	1.88×10^{-5}	11.97
17	CALCITE GIBBSITE SiO ₂ gel CSH(I) THAUMASI	2.17×10^{-3}	7.16×10^{-6}	5.23×10^{-3}	1.58×10^{-5}	2.49×10^{-5}	10.30
18	CALCITE CSH(I) CSH(II) ETTRINGI THAUMASI	1.97×10^{-2}	9.53×10^{-7}	3.38×10^{-5}	8.96×10^{-6}	3.70×10^{-5}	12.48
19	CALCITE C ₂ ASH ₈ CSH(I) ETTRINGI THAUMASI	4.17×10^{-3}	1.52×10^{-4}	1.04×10^{-4}	1.04×10^{-5}	3.83×10^{-5}	11.84
20	CALCITE C ₂ ASH ₈ GIBBSITE CSH(I) THAUMASI	2.82×10^{-3}	1.65×10^{-4}	1.70×10^{-4}	1.13×10^{-5}	4.15×10^{-5}	11.66
21	CALCITE PORTLAND CSH(II) ETTRINGI THAUMASI	2.20×10^{-2}	5.29×10^{-7}	2.70×10^{-5}	8.93×10^{-6}	4.48×10^{-5}	12.52
22	CALCITE C ₂ ASH ₈ GIBBSITE ETTRINGI THAUMASI	3.28×10^{-3}	1.96×10^{-4}	9.13×10^{-5}	1.09×10^{-5}	6.07×10^{-5}	11.73
23	C ₂ ASH ₈ GIBBSITE CSH(I) ETTRINGI THAUMASI	2.86×10^{-3}	1.64×10^{-4}	1.69×10^{-4}	4.50×10^{-6}	1.04×10^{-4}	11.66
24	GYPSUM PORTLAND CSH(II)ETTRINGI THAUMASI	3.39×10^{-2}	1.01×10^{-10}	2.59×10^{-5}	3.10×10^{-8}	1.24×10^{-2}	12.49
25	GYPSUM CALCITE PORTLAND ETTRINGI THAUMASI	3.39×10^{-2}	1.01×10^{-10}	8.95×10^{-8}	8.95×10^{-6}	1.24×10^{-2}	12.49
26	GYPSUM CSH(I) CSH(II) ETTRINGI THAUMASI	3.17×10^{-2}	1.34×10^{-10}	3.22×10^{-5}	2.53×10^{-8}	1.25×10^{-2}	12.44
27	GYPSUM GIBBSITE SiO ₂ gel CSH(I) THAUMASI	1.60×10^{-2}	4.16×10^{-6}	3.60×10^{-3}	1.03×10^{-8}	1.48×10^{-2}	10.02
28	GYPSUM GIBBSITE CSH(I) ETTRINGI THAUMASI	1.58×10^{-2}	6.93×10^{-6}	1.73×10^{-3}	1.05×10^{-8}	1.50×10^{-2}	10.25
29	GYPSUM CALCITE GIBBSITE SiO ₂ gel THAUMASI	1.53×10^{-2}	1.27×10^{-7}	1.46×10^{-3}	8.68×10^{-5}	1.52×10^{-2}	8.51
30	GYPSUM CALCITE GIBBSITE ETTRINGI THAUMASI	1.54×10^{-2}	7.03×10^{-6}	1.71×10^{-6}	1.08×10^{-5}	1.52×10^{-2}	10.25

Abbreviations: Portland, portlandite; ettringi, ettringite; hemicarb, hemicarboaluminate (see Table 1 for formula); monocarb, monocarboaluminate (see Table 2); C₂ASH₈, stratlingite; see text for definition of the two hydrogarnets (HG) and two C–S–H formulae included in calculations.

Table 3. Stable assemblages with two solids coexisting with an aqueous (aq) phase a

PORTLAND + C ₃ AH ₆ + aq.	CSH(SI) + MONOCARB + aq.	CALCITE + GIBBSITE + aq.
PORTLAND + SIHGlow + aq.	PORTLAND + MONOCARB + aq.	C ₂ ASH ₈ + CSH(SI) + aq.
PORTLAND + ETTRINGITE + aq.	C ₃ AH ₆ + GIBBSITE + aq.	CALCITE + SiO ₂ gel + aq.
PORTLAND + HEMICARB + aq.	SIHGlow + GIBBSITE + aq.	CALCITE + THAUMASITE + aq.
C ₃ AH ₆ + SIHGlow + aq.	GIBBSITE + ETTRINGITE + aq.	GIBBSITE + SiO ₂ gel + aq.
C ₃ AH ₆ + ETTRINGITE + aq.	GIBBSITE + MONOCARB + aq.	GIBBSITE + CSH(SI) + aq.
C ₃ AH ₆ + HEMICARB + aq.	CALCITE + PORTLAND + aq.	GIBBSITE + THAUMASITE + aq.
SIHGlow + ETTRINGITE + aq.	CALCITE + CSH(SII) + aq.	SiO ₂ gel + CSH(SI) + aq.
SIHGlow + HEMICARB + aq.	CALCITE + ETTRINGITE + aq.	SiO ₂ gel + THAUMASITE + aq.
ETTRINGITE + HEMICARB + aq.	CALCITE + MONOCARB + aq.	CSH(SI) + THAUMASITE + aq.
C ₃ AH ₆ + MONOCARB + aq.	CALCITE + CSH(SI) + aq.	CSH(SII) + THAUMASITE + aq.
SIHGlow + MONOCARB + aq.	CALCITE + SIHGlow + aq.	ETTRINGITE + THAUMASITE + aq.
ETTRINGITE + MONOCARB + aq.	CALCITE + SIHGhigh + aq.	C ₂ ASH ₈ + THAUMASITE + aq.
MONOCARB + HEMICARB + aq.	SIHGlow + SIHGhigh + aq.	PORTLAND + THAUMASITE + aq.
PORTLAND + CSH(SII) + aq.	SIHGhigh + CSH(SI) + aq.	GYPSUM + PORTLAND + aq.
SIHGlow + CSH(SII) + aq.	SIHGhigh + ETTRINGITE + aq.	GYPSUM + CSH(SII) + aq.
CSH(SII) + ETTRINGITE + aq.	C ₂ ASH ₈ + SIHGlow + aq.	GYPSUM + ETTRINGITE + aq.
CSH(SII) + HEMICARB + aq.	C ₂ ASH ₈ + GIBBSITE + aq.	GYPSUM + THAUMASITE + aq.
CSH(SII) + MONOCARB + aq.	C ₂ ASH ₈ + ETTRINGITE + aq.	GYPSUM + CALCITE + aq.
SIHGlow + CSH(SI) + aq.	C ₂ ASH ₈ + MONOCARB + aq.	GYPSUM + CSH(SI) + aq.
CSH(SI) + CSH(SII) + aq.	CALCITE + C ₂ ASH ₈ + aq.	GYPSUM + GIBBSITE + aq.
CSH(SI) + ETTRINGITE + aq.	C ₂ ASH ₈ + SIHGhigh + aq.	GYPSUM + SiO ₂ gel + aq.

See text and Tables 1 and 2 for definition of abbreviations.

Table 4. Stable assemblages with three solids coexisting with an aqueous (aq) phase

PORTLAND + C ₃ AH ₆ + SIHGlow + aq.	C ₂ ASH ₈ + GIBBSITE + ETTRINGITE + aq.
PORTLAND + C ₃ AH ₆ + ETTRINGITE + aq.	C ₂ ASH ₈ + GIBBSITE + MONOCARB + aq.
PORTLAND + C ₃ AH ₆ + HEMICARB + aq.	C ₂ ASH ₈ + ETTRINGITE + MONOCARB + aq.
PORTLAND + SIHGlow + ETTRINGITE + aq.	CALCITE + C ₂ ASH ₈ + SIHGlow + aq.
PORTLAND + SIHGlow + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + ETTRINGITE + aq.
PORTLAND + ETTRINGITE + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + MONOCARB + aq.
C ₃ AH ₆ + SIHGlow + ETTRINGITE + aq.	CALCITE + C ₂ ASH ₈ + SIHGhigh + aq.
C ₃ AH ₆ + SIHGlow + HEMICARB + aq.	C ₂ ASH ₈ + SIHGlow + SIHGhigh + aq.
C ₃ AH ₆ + ETTRINGITE + HEMICARB + aq.	C ₂ ASH ₈ + SIHGhigh + ETTRINGITE + aq.
SIHGlow + ETTRINGITE + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + GIBBSITE + aq.
C ₃ AH ₆ + SIHGlow + MONOCARB + aq.	CALCITE + GIBBSITE + ETTRINGITE + aq.
C ₃ AH ₆ + ETTRINGITE + MONOCARB + aq.	CALCITE + GIBBSITE + MONOCARB + aq.
C ₃ AH ₆ + MONOCARB + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + CSH(SI) + aq.
SIHGlow + ETTRINGITE + MONOCARB + aq.	C ₂ ASH ₈ + SIHGhigh + CSH(SI) + aq.
SIHGlow + MONOCARB + HEMICARB + aq.	C ₂ ASH ₈ + CSH(SI) + ETTRINGITE + aq.
ETTRINGITE + MONOCARB + HEMICARB + aq.	CALCITE + GIBBSITE + SiO ₂ gel + aq.
PORTLAND + SIHGlow + CSH(SII) + aq.	CALCITE + GIBBSITE + CSH(SI) + aq.
PORTLAND + CSH(SII) + ETTRINGITE + aq.	CALCITE + GIBBSITE + THAUMASITE + aq.
PORTLAND + CSH(SII) + HEMICARB + aq.	CALCITE + SiO ₂ gel + CSH(SI) + aq.
SIHGlow + CSH(SII) + ETTRINGITE + aq.	CALCITE + SiO ₂ gel + THAUMASITE + aq.
SIHGlow + CSH(SII) + HEMICARB + aq.	CALCITE + CSH(SI) + THAUMASITE + aq.
CSH(SII) + ETTRINGITE + HEMICARB + aq.	GIBBSITE + SiO ₂ gel + CSH(SI) + aq.
SIHGlow + CSH(SII) + MONOCARB + aq.	GIBBSITE + SiO ₂ gel + THAUMASITE + aq.
CSH(SII) + ETTRINGITE + MONOCARB + aq.	GIBBSITE + CSH(SI) + THAUMASITE + aq.
CSH(SII) + MONOCARB + HEMICARB + aq.	SiO ₂ gel + CSH(SI) + THAUMASITE + aq.
SIHGlow + CSH(SI) + CSH(SII) + aq.	CALCITE + CSH(SII) + THAUMASITE + aq.
SIHGlow + CSH(SI) + ETTRINGITE + aq.	CALCITE + ETTRINGITE + THAUMASITE + aq.
SIHGlow + CSH(SI) + MONOCARB + aq.	CSH(SI) + CSH(SII) + THAUMASITE + aq.
CSH(SI) + CSH(SII) + ETTRINGITE + aq.	CSH(SI) + ETTRINGITE + THAUMASITE + aq.
CSH(SI) + CSH(SII) + MONOCARB + aq.	CSH(SII) + ETTRINGITE + THAUMASITE + aq.
CSH(SI) + ETTRINGITE + MONOCARB + aq.	CALCITE + C ₂ ASH ₈ + THAUMASITE + aq.
PORTLAND + CSH(SII) + MONOCARB + aq.	C ₂ ASH ₈ + CSH(SI) + THAUMASITE + aq.
PORTLAND + ETTRINGITE + MONOCARB + aq.	C ₂ ASH ₈ + ETTRINGITE + THAUMASITE + aq.
PORTLAND + MONOCARB + HEMICARB + aq.	C ₂ ASH ₈ + GIBBSITE + CSH(SI) + aq.
C ₃ AH ₆ + SIHGlow + GIBBSITE + aq.	C ₂ ASH ₈ + GIBBSITE + THAUMASITE + aq.
C ₃ AH ₆ + GIBBSITE + ETTRINGITE + aq.	CALCITE + PORTLAND + THAUMASITE + aq.
C ₃ AH ₆ + GIBBSITE + MONOCARB + aq.	PORTLAND + CSH(SII) + THAUMASITE + aq.
SIHGlow + GIBBSITE + ETTRINGITE + aq.	PORTLAND + ETTRINGITE + THAUMASITE + aq.
SIHGlow + GIBBSITE + MONOCARB + aq.	GIBBSITE + ETTRINGITE + THAUMASITE + aq.
GIBBSITE + ETTRINGITE + MONOCARB + aq.	GIBBSITE + CSH(SI) + ETTRINGITE + aq.
CALCITE + PORTLAND + CSH(SII) + aq.	GYP SUM + PORTLAND + CSH(SII) + aq.
CALCITE + PORTLAND + ETTRINGITE + aq.	GYP SUM + PORTLAND + ETTRINGITE + aq.
CALCITE + PORTLAND + MONOCARB + aq.	GYP SUM + PORTLAND + THAUMASITE + aq.
CALCITE + CSH(SII) + ETTRINGITE + aq.	GYP SUM + CSH(SII) + ETTRINGITE + aq.
CALCITE + CSH(SII) + MONOCARB + aq.	GYP SUM + CSH(SII) + THAUMASITE + aq.
CALCITE + ETTRINGITE + MONOCARB + aq.	GYP SUM + ETTRINGITE + THAUMASITE + aq.
CALCITE + CSH(SI) + CSH(SII) + aq.	GYP SUM + CALCITE + PORTLAND + aq.
CALCITE + CSH(SI) + ETTRINGITE + aq.	GYP SUM + CALCITE + ETTRINGITE + aq.
CALCITE + CSH(SI) + MONOCARB + aq.	GYP SUM + CALCITE + THAUMASITE + aq.
CALCITE + SIHGlow + CSH(SI) + aq.	GYP SUM + CSH(SI) + CSH(SII) + aq.
CALCITE + SIHGlow + ETTRINGITE + aq.	GYP SUM + CSH(SI) + ETTRINGITE + aq.
CALCITE + SIHGlow + MONOCARB + aq.	GYP SUM + CSH(SI) + THAUMASITE + aq.
CALCITE + SIHGlow + SIHGhigh + aq.	GYP SUM + GIBBSITE + SiO ₂ gel + aq.
CALCITE + SIHGhigh + CSH(SI) + aq.	GYP SUM + GIBBSITE + CSH(SI) + aq.
CALCITE + SIHGhigh + ETTRINGITE + aq.	GYP SUM + GIBBSITE + THAUMASITE + aq.
SIHGlow + SIHGhigh + CSH(SI) + aq.	GYP SUM + SiO ₂ gel + CSH(SI) + aq.
SIHGlow + SIHGhigh + ETTRINGITE + aq.	GYP SUM + SiO ₂ gel + THAUMASITE + aq.
SIHGhigh + CSH(SI) + ETTRINGITE + aq.	GYP SUM + GIBBSITE + ETTRINGITE + aq.
C ₂ ASH ₈ + SIHGlow + GIBBSITE + aq.	GYP SUM + CALCITE + GIBBSITE + aq.
C ₂ ASH ₈ + SIHGlow + ETTRINGITE + aq.	GYP SUM + CALCITE + SiO ₂ gel + aq.
C ₂ ASH ₈ + SIHGlow + MONOCARB + aq.	

See text and Tables 1 and 2 for definition of abbreviations.

Table 5. Stable assemblages with four solids coexisting with an aqueous (aq) phase

PORTLAND + C ₃ AH ₆ + SIHGlow + ETTRINGITE + aq.	C ₂ ASH ₈ + SIHGlow + SIHGhigh + ETTRINGITE + aq.
PORTLAND + C ₃ AH ₆ + SIHGlow + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + GIBBSITE + ETTRINGITE + aq.
PORTLAND + C ₃ AH ₆ + ETTRINGITE + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + GIBBSITE + MONOCARB + aq.
PORTLAND + SIHGlow + ETTRINGITE + HEMICARB + aq.	CALCITE + GIBBSITE + ETTRINGITE + MONOCARB + aq.
C ₃ AH ₆ + SIHGlow + ETTRINGITE + HEMICARB + aq.	CALCITE + C ₂ ASH ₈ + SIHGhigh + CSH(SI) + aq.
C ₃ AH ₆ + SIHGlow + ETTRINGITE + MONOCARB + aq.	CALCITE + C ₂ ASH ₈ + CSH(SI) + ETTRINGITE + aq.
C ₃ AH ₆ + SIHGlow + MONOCARB + HEMICARB + aq.	C ₂ ASH ₈ + SIHGhigh + CSH(SI) + ETTRINGITE + aq.
C ₃ AH ₆ + ETTRINGITE + MONOCARB + HEMICARB + aq.	CALCITE + GIBBSITE + SiO ₂ gel + CSH(SI) + aq.
SIHGlow + ETTRINGITE + MONOCARB + HEMICARB + aq.	CALCITE + GIBBSITE + SiO ₂ gel + THAUMASITE + aq.
PORTLAND + SIHGlow + CSH(SII) + ETTRINGITE + aq.	CALCITE + GIBBSITE + CSH(SI) + THAUMASITE + aq.
PORTLAND + SIHGlow + CSH(SII) + HEMICARB + aq.	CALCITE + SiO ₂ gel + CSH(SI) + THAUMASITE + aq.
PORTLAND + CSH(SII) + ETTRINGITE + HEMICARB + aq.	GIBBSITE + SiO ₂ gel + CSH(SI) + THAUMASITE + aq.
SIHGlow + CSH(SII) + ETTRINGITE + HEMICARB + aq.	CALCITE + CSH(SI) + CSH(SII) + THAUMASITE + aq.
SIHGlow + CSH(SII) + ETTRINGITE + MONOCARB + aq.	CALCITE + CSH(SI) + ETTRINGITE + THAUMASITE + aq.
SIHGlow + CSH(SII) + MONOCARB + HEMICARB + aq.	CALCITE + CSH(SII) + ETTRINGITE + THAUMASITE + aq.
CSH(SII) + ETTRINGITE + MONOCARB + HEMICARB + aq.	CSH(SI) + CSH(SII) + ETTRINGITE + THAUMASITE + aq.
SIHGlow + CSH(SI) + CSH(SII) + ETTRINGITE + aq.	CALCITE + C ₂ ASH ₈ + CSH(SI) + THAUMASITE + aq.
SIHGlow + CSH(SI) + CSH(SII) + MONOCARB + aq.	CALCITE + C ₂ ASH ₈ + ETTRINGITE + THAUMASITE + aq.
SIHGlow + CSH(SI) + ETTRINGITE + MONOCARB + aq.	C ₂ ASH ₈ + CSH(SI) + ETTRINGITE + THAUMASITE + aq.
CSH(SI) + CSH(SII) + ETTRINGITE + MONOCARB + aq.	CALCITE + C ₂ ASH ₈ + GIBBSITE + CSH(SI) + aq.
PORTLAND + CSH(SII) + ETTRINGITE + MONOCARB + aq.	CALCITE + C ₂ ASH ₈ + GIBBSITE + THAUMASITE + aq.
PORTLAND + CSH(SII) + MONOCARB + HEMICARB + aq.	C ₂ ASH ₈ + GIBBSITE + CSH(SI) + THAUMASITE + aq.
PORTLAND + ETTRINGITE + MONOCARB + HEMICARB + aq.	CALCITE + PORTLAND + CSH(SII) + THAUMASITE + aq.
C ₃ AH ₆ + SIHGlow + GIBBSITE + ETTRINGITE + aq.	CALCITE + PORTLAND + ETTRINGITE + THAUMASITE + aq.
C ₃ AH ₆ + SIHGlow + GIBBSITE + MONOCARB + aq.	PORTLAND + CSH(SII) + ETTRINGITE + THAUMASITE + aq.
C ₃ AH ₆ + GIBBSITE + ETTRINGITE + MONOCARB + aq.	CALCITE + GIBBSITE + ETTRINGITE + THAUMASITE + aq.
SIHGlow + GIBBSITE + ETTRINGITE + MONOCARB + aq.	C ₂ ASH ₈ + GIBBSITE + ETTRINGITE + THAUMASITE + aq.
CALCITE + PORTLAND + CSH(SII) + ETTRINGITE + aq.	C ₂ ASH ₈ + GIBBSITE + CSH(SI) + ETTRINGITE + aq.
CALCITE + PORTLAND + CSH(SII) + MONOCARB + aq.	GIBBSITE + CSH(SI) + ETTRINGITE + THAUMASITE + aq.
CALCITE + PORTLAND + ETTRINGITE + MONOCARB + aq.	GYPNUM + PORTLAND + CSH(SII) + ETTRINGITE + aq.
CALCITE + CSH(SII) + ETTRINGITE + MONOCARB + aq.	GYPNUM + PORTLAND + CSH(SII) + THAUMASITE + aq.
CALCITE + CSH(SI) + CSH(SII) + ETTRINGITE + aq.	GYPNUM + PORTLAND + ETTRINGITE + THAUMASITE + aq.
CALCITE + CSH(SI) + CSH(SII) + MONOCARB + aq.	GYPNUM + CSH(SII) + ETTRINGITE + THAUMASITE + aq.
CALCITE + CSH(SI) + ETTRINGITE + MONOCARB + aq.	GYPNUM + CALCITE + PORTLAND + ETTRINGITE + aq.
CALCITE + SIHGlow + CSH(SI) + ETTRINGITE + aq.	GYPNUM + CALCITE + PORTLAND + THAUMASITE + aq.
CALCITE + SIHGlow + CSH(SI) + MONOCARB + aq.	GYPNUM + CALCITE + ETTRINGITE + THAUMASITE + aq.
CALCITE + SIHGlow + ETTRINGITE + MONOCARB + aq.	GYPNUM + CSH(SI) + CSH(SII) + ETTRINGITE + aq.
CALCITE + SIHGlow + SIHGhigh + CSH(SI) + aq.	GYPNUM + CSH(SI) + CSH(SII) + THAUMASITE + aq.
CALCITE + SIHGlow + SIHGhigh + ETTRINGITE + aq.	GYPNUM + CSH(SI) + ETTRINGITE + THAUMASITE + aq.
CALCITE + SIHGhigh + CSH(SI) + ETTRINGITE + aq.	GYPNUM + GIBBSITE + SiO ₂ gel + CSH(SI) + aq.
C ₂ ASH ₈ + SIHGlow + GIBBSITE + ETTRINGITE + aq.	GYPNUM + GIBBSITE + SiO ₂ gel + THAUMASITE + aq.
C ₂ ASH ₈ + SIHGlow + GIBBSITE + MONOCARB + aq.	GYPNUM + GIBBSITE + CSH(SI) + THAUMASITE + aq.
C ₂ ASH ₈ + SIHGlow + ETTRINGITE + MONOCARB + aq.	GYPNUM + SiO ₂ gel + CSH(SI) + THAUMASITE + aq.
C ₂ ASH ₈ + GIBBSITE + ETTRINGITE + MONOCARB + aq.	GYPNUM + GIBBSITE + CSH(SI) + ETTRINGITE + aq.
CALCITE + C ₂ ASH ₈ + SIHGlow + ETTRINGITE + aq.	GYPNUM + GIBBSITE + ETTRINGITE + THAUMASITE + aq.
CALCITE + C ₂ ASH ₈ + SIHGlow + MONOCARB + aq.	GYPNUM + CALCITE + GIBBSITE + SiO ₂ gel + aq.
CALCITE + C ₂ ASH ₈ + ETTRINGITE + MONOCARB + aq.	GYPNUM + CALCITE + GIBBSITE + THAUMASITE + aq.
CALCITE + C ₂ ASH ₈ + SIHGlow + SIHGhigh + aq.	GYPNUM + CALCITE + SiO ₂ gel + THAUMASITE + aq.
CALCITE + C ₂ ASH ₈ + SIHGhigh + ETTRINGITE + aq.	GYPNUM + CALCITE + GIBBSITE + ETTRINGITE + aq.

See text and Tables 1 and 2 for definition of abbreviations.

more than two solids, but which include only two (or more) of the incompatible solids shown in Table 6 will also not be stable.

The calculations reveal many more phase assemblages than have been found experimentally or are recorded from structures undergoing TSA. However, we can immediately compare results with the limited data recorded in the literature. For example, and contrary to the results of Juel *et al.*,⁷ the phase assemblage thaumasite + calcite + gypsum + ettringite is stable; indeed, this phase assemblage is often noted to occur in cement and concrete undergoing TSA.¹

Discussion

In calculating equilibria in cement systems, certain phases known to be thermodynamically unstable are nevertheless very persistent. For example, C–S–H, the principal-binding phase of cements, is undoubtedly thermodynamically unstable with respect to crystalline CaO–SiO₂–H₂O phases. However, we know that the C–S–H gel is important to the constitution of real cement and concretes. Unaltered historic Portland cements still contain C–S–H and it has been reported to occur in Roman concretes as well as in nature, in

Table 6. Solid phase pairs in the $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$ system at 25°C which, from calculation, are not compatible

GYPSUM + C_2ASH_8 + aq.	C_2ASH_8 + C_3AH_6 + aq.	SIHGhigh + CSH(SII) + aq.
GYPSUM + C_3AH_6 + aq.	C_2ASH_8 + SiO_2gel + aq.	SIHGhigh + MONOCARB + aq.
GYPSUM + SIHGlow + aq.	C_2ASH_8 + CSH(SII) + aq.	SIHGhigh + HEMICARB + aq.
GYPSUM + SIHGhigh + aq.	C_2ASH_8 + HEMICARB + aq.	SIHGhigh + THAUMASITE + aq.
GYPSUM + MONOCARB + aq.	C_3AH_6 + SIHGhigh + aq.	GIBBSITE + CSH(SII) + aq.
GYPSUM + HEMICARB + aq.	C_3AH_6 + SiO_2gel + aq.	GIBBSITE + HEMICARB + aq.
CALCITE + C_3AH_6 + aq.	C_3AH_6 + CSH(SI) + aq.	SiO_2gel + CSH(SII) + aq.
CALCITE + HEMICARB + aq.	C_3AH_6 + CSH(SII) + aq.	SiO_2gel + ETTRINGITE + aq.
PORTLAND + C_2ASH_8 + aq.	C_3AH_6 + THAUMASITE + aq.	SiO_2gel + MONOCARB + aq.
PORTLAND + SIHGhigh + aq.	SIHGlow + SiO_2gel + aq.	SiO_2gel + HEMICARB + aq.
PORTLAND + GIBBSITE + aq.	SIHGlow + THAUMASITE + aq.	CSH(SI) + HEMICARB + aq.
PORTLAND + SiO_2gel + aq.	SIHGhigh + GIBBSITE + aq.	MONOCARB + THAUMASITE + aq.
PORTLAND + CSH(SI) + aq.	SIHGhigh + SiO_2gel + aq.	HEMICARB + THAUMASITE + aq.

hydrothermally-altered impure limestones. If calculations are to be relevant to real cements, it is essential to include these phases. Thus with the conspicuous exception of C-S-H, which is allowed to replace the stable crystalline C-S-H minerals, and silica gel, allowed to replace quartz, the calculations are otherwise allowed to proceed on the basis of true thermodynamic equilibrium. Since thaumasite mainly appears in altered concretes, and since the driving force for chemical and mineralogical reconstitution of cement paste mineralogy is the approach to thermodynamic equilibrium, the authors see no inconsistency in the approach taken, namely: of using equilibrium but admitting selected metastable but persistent phases into the scheme of calculation.

At the temperature used for the calculation (25°C), several AFm minerals observed in cement are not thermodynamically stable. Examples include hydroxy- and sulphate-containing AFm phases. For this reason, they are not included in the calculations whereas other stable AFm phases are included, for example C_2ASH_8 , mono- and hemicarboaluminate. The mineralogy of the system is very sensitive to temperature. At some temperature, probably not much exceeding 25°C , thaumasite will no longer be stable. On the other hand, at temperatures exceeding $\sim 40^\circ\text{C}$, sulphate-AFm becomes thermodynamically stable and thus the range of stable AFm chemistries will broaden. The authors do not at present have the data necessary to repeat the calculations undertaken here at other temperatures, especially at $> 25^\circ\text{C}$, but as part of a wider study of the relationships between cement paste mineralogy, temperature, environmentally-conditioned reactions and durability of the resulting concrete, the quickest way to establish quantitative links with paste mineralogy may be to identify and subsequently obtain and apply the missing data.

Some ranges of solid solution, for example between ettringite and thaumasite, have also not been included in the calculation. The impact of solid solution will be, in general, to enhance the importance of phase assem-

blages containing fewer than the maximum permitted number of phases. Thus sulphate, carbonate and silicate concentrations, which are predicted in the model to form assemblages containing both ettringite and thaumasite may, in practice, give only one phase, typically an ettringite solid solution. These data deficiencies do not invalidate the general conclusions reached here but simply shift the boundaries between phase regions mainly by extending the stability domain of ettringite.

Results of the calculation disclose that thaumasite is stable at 25°C . The question of the kinetics of its formation is not handled by thermodynamics, but once formed, typically but not necessarily exclusively at $< 25^\circ\text{C}$, it can persist stably over a broad temperature. This is supported by field experiments² and by natural occurrences of thaumasite. The thaumasite stability domain is wide and, as a consequence, thaumasite coexists stably with all the phases of the $\text{CaO}-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{CaSO}_4-\text{CaCO}_3-\text{H}_2\text{O}$ system with the exception of hydrogarnets (C_3AH_6 , HG Si low and HG Si high) and with calcium hemicarboaluminate and monocarboaluminate. The incompatibility of assemblages containing carboaluminates is in agreement with the calculation of Juel *et al.*⁷ Hydrogarnet and Si-substituted hydrogarnet are not stable with thaumasite; they react, with transformation to ettringite, while the sulphate concentration in the aqueous phase is still insufficient to stabilise thaumasite. Indeed, if the invariant points are sorted in order of increasing sulphate concentration of the aqueous phase necessary to stabilise the various solid phase assemblages, as has been done in Table 2, the first 16 invariant points do not contain thaumasite but instead contain ettringite. This arises because thaumasite characteristically requires higher sulfate concentrations than ettringite to form. Ettringite is however frequently found to occur with thaumasite. The four invariant points, numbered 17, 20, 27 and 29 in Table 2, are exceptions: gibbsite is always present and the pH is below 12. These four invariant points represent two different chemical environments: invariant points 17 and 20 correspond to low

sulphate concentrations whereas 27 and 29 correspond to characteristically high sulphate concentrations conditioned by the solubility of gypsum. In these conditions thaumasite can form directly without the need for concomitant or anterior formation of ettringite. This is considered to resolve the controversy concerning the need to have ettringite as a precursor for thaumasite formation; while this may occur under some conditions it is not a prerequisite in all ranges of composition. The invariant point 29, of gypsum + calcite + gibbsite + SiO₂ gel + thaumasite, is also very interesting: first because it is supposed to be the final stage of TSA, and second, because it corresponds to a quite low pH aqueous (8.51). Thus while ettringite cannot be in equilibrium with pH lower than about 9.5 (point 30), thaumasite can exist stably to lower pH. This is in disagreement with Gaze and Crammond¹ who report that thaumasite does not form at pH less than 10.5: the current authors fix the limit as approximately one pH unit lower. Also, the conclusions reported by Jallad *et al.*¹⁵ seem to be more related to the addition of other ions used to control the pH rather than to the stability of thaumasite itself.

It should also be noted that high aqueous carbonate contents are not a necessary prerequisite for formation of thaumasite. For example, at invariant point 28, Table 2, the composition of the aqueous solution is calculated to contain 2.53×10^{-8} mol/kg of carbonate. Only when calcite occurs together with thaumasite, as at invariant points 17, 18, 19, 20, 21, 22, 25, 29 and 30, do aqueous carbonate contents increase to $\sim 10^{-5}$ mol/kg, sustained by the relatively high solubility of calcite. While some of the assemblages listed above are only likely to occur in highly altered cements, thaumasite can coexist with C-S-H as for example in assemblages 18, 19, 21 and 24, Table 2. These assemblages might appear in plain or blended cements relatively early in their history of alteration.

Broad consideration of the results shown in the tables, coupled with experience of cements and concrete undergoing chemical alteration, either as a result of internal reaction or reaction with their service environment, suggests that with few exceptions the driving force for chemical-mineralogical alteration results from an approach to a local thermodynamic equilibrium. The exceptions include the persistence of a few well-defined but metastable phases such as C-S-H. While the data given here define all energetically feasible assemblages at 25°C, the reaction pathways of cement undergoing reaction with its service environment are not intuitively obvious. Reaction pathways are determined in practice by a combination of dissolution and reaction: in general, the dissolution component remains unknown in field studies. We need patiently to relate microscopic and diffusion studies to the phase equilibria in order to determine the response of paste mineralogy to diffusional composition gradients in order to infer to what extent dissolution has occurred.

We need additionally to improve the accuracy of thermodynamic predictions, as for example by determining more precisely the extent of solid solutions and of element fractionation between coexisting phases. We also need increasingly to apply thermodynamics, coupled with geochemical limitations on, for example, natural ground water compositions, to devise standard test methods which more accurately reflect the performance of cement concretes in adverse environments. Above all, we need to quantify deteriorations, so as to be able to produce tailored cement and concrete formulations which will give enhanced performance in adverse service environments. Decades of largely empirical research have produced considerable enhancement in concrete durability but the authors submit that a coupled chemical thermodynamic-property model represents the way forward.

Conclusion

The phase diagram of the CaO–Al₂O₃–SiO₂–CaSO₄–CaCO₃–H₂O system has been resolved by calculation of the composition of the stable invariant points. The results demonstrate the usefulness of this approach for the investigation complex of multicomponent systems.

Thaumasite is calculated to be stable over a wide range of aqueous phase compositions and is one of the major phases of the CaO–Al₂O₃–SiO₂–CaSO₄–CaCO₃–H₂O system at 25°C. With respect to ettringite, thaumasite formation generally demands more sulphate ions in the aqueous phase but thaumasite is also stable at pH as low as 8.5 compared to 9.5 for ettringite. In these low pH conditions, thaumasite is expected to form directly without need for the formation of precursor ettringite. These conditions are believed to be responsible for much of the damage occurring during TSA. The phase development model presented here provides a flexible tool with which to predict chemical fluxes in cements undergoing alteration and relate chemical-mineralogical change to engineering properties.

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References

1. GAZE M. and CRAMMOND N. J. The formation of thaumasite in a cement:lime:sand mortar exposed to cold magnesium and potassium sulfate solutions. *Cement and Concrete Research*, 2000, **30**, 209–222.

2. SAHU S., BADGER S. and THAULOW N. Evidence of thaumasite formation in southern California concrete. *Cement and Concrete Composites*, 2002, **24**, 379–384.
3. DAMIDOT D. and GLASSER F. P. Thermodynamic investigation of the CaO–Al₂O₃–CaSO₄–H₂O system at 25°C and influence of Na₂O. *Cement and Concrete Research*, 1993, **23**, 221–238.
4. DAMIDOT D., STRONACH S., KINDNESS A., ATKINS M. and GLASSER F. P. Thermodynamic investigation of the CaO–Al₂O₃–CaCO₃–H₂O system at 25°C and the influence of Na₂O. *Cement and Concrete Research*, 1994, **24**, 563–572.
5. STRONACH S. and GLASSER F. P. Modelling the impact of abundant geochemical components on phase stability and solubility of the CaO–SiO₂–H₂O system at 25°C: Na⁺, K⁺, SO₄²⁻, Cl⁻ and CO₃²⁻. *Advances in Cement Research*, 1997, **9**, No. 36, 167–182.
6. DAMIDOT D. and GLASSER F. P. Investigation of the CaO–Al₂O₃–SiO₂–H₂O system at 25°C by thermodynamic calculations. *Cement and Concrete Research*, 1995, **25**, 22–28.
7. JUEL I., HERFORT D., GOLLOP R., KONNERUP-MADSEN J., JAKOBSEN H. J. and SKIBSTED J. A thermodynamic model for predicting the stability of thaumasite. *Proceedings of the 1st International Conference on Thaumasite in Cementitious Materials, Watford, UK*, 19–21 June 2002.
8. EDGE R. A. and TAYLOR H. F. W. Crystal structure of thaumasite [Ca₃Si(OH)₆·12H₂O] SO₄·CO₃. *Acta Crystallographa*, 1971, **B27**, 594–601.
9. BARNETT S. J., MACPHEE D. E., LACHOWSKI E. E. and CRAMMOND N. J. XRD, EDX and IR analysis of solid solutions between thaumasite and ettringite. *Cement and Concrete Research*, 2002, **32**, 719–730.
10. PLUMMER L. N., PARKHURST D. L., FLEMING G. W. and DUNKLE S. A. *A Computer Program Incorporating Pitzer's Equations for Calculation of Geochemical Reactions in Brines*. US Geological, Survey Water-Resources Investigations Report 88-4153, 1988.
11. PITZER K. S. Thermodynamics of electrolytes I: theoretical basis and general equations. *Journal of Physical Chemistry*, 1973, **77**, 268–277.
12. NONAT A., COURAULT A.-C. and DAMIDOT D. A surface reaction model describing the structure and stoichiometry variations of calcium silicate hydrate (C–S–H) with calcium hydroxide concentration in solution. *Langmuir*, forthcoming.
13. DAMIDOT D. and NONAT A. C₃S hydration in diluted and stirred suspension: (II) Properties of C–S–H precipitated during the two kinetic steps. *Advances in Cement Research*, 1994, **22**, 83–91.
14. JAPPY T. G. and GLASSER F. P. Synthesis and stability of silica substituted hydrogarnet, Ca₃Al₂Si_{3-x}O_{12-4x}(OH)_{4x}. *Advances in Cement Research*, 1991, **4**, No. 1, 1–8.
15. JALLAD K. N., SANTHANAM M. and COHEN M. D. Stability and reactivity of thaumasite at different pH levels. *Cement and Concrete Research*, 2000, **33**, 433–437.