



Review

# Light to Hydrogen: Photocatalytic Hydrogen Generation from Water with Molecularly-Defined Iron Complexes

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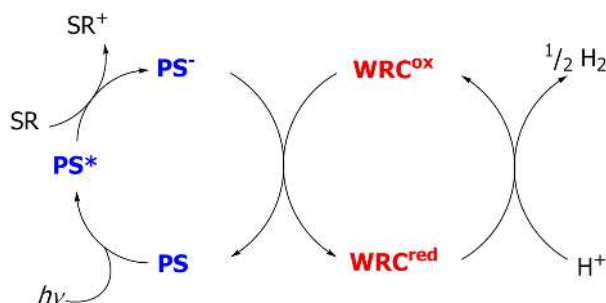
**Abstract:** Photocatalytic hydrogen generation is considered to be attractive due to its combination of solar energy conversion and storage. Currently-used systems are either based on homogeneous or on heterogeneous materials, which possess a light harvesting and a catalytic subunit. The subject of this review is a brief summary of homogeneous proton reduction systems using sacrificial agents with special emphasis on non-noble metal systems applying convenient iron(0) sources. Iridium photosensitizers, which were proven to have high quantum yields of up to 48% (415 nm), have been employed, as well as copper photosensitizers. In both cases, the addition or presence of a phosphine led to the transformation of the iron precursor with subsequently increased activities. Reaction pathways were investigated by photoluminescence, electron paramagnetic resonance (EPR), Raman, FTIR and mass spectroscopy, as well as time-dependent DFT-calculations. In the future, this knowledge will set the basis to design photo(electro)chemical devices with tailored electron transfer cascades and without the need for sacrificial agents.

**Keywords:** hydrogen; iron; copper; hydrogenases; photosensitizer; photocatalysis; water splitting; spectroscopy

## 1. Introduction

An increasing number of pilot plants combining wind mills or photovoltaic devices with classical water electrolysis is installed all over the world. This reflects the progressive contribution of the renewable energies wind and sunlight to a sufficient and sustainable energy supply [1–4] and thereby raises the necessity to store the harvested electric energy due to its fluctuating occurrence. In this respect, hydrogen is of particular interest as the primary product of the interconversion of electric and chemical energy, as well as a secondary energy carrier [1–4]. Besides the hydrogen generation from water, also its storage and back conversion to electrical energy are current topics of intensive research [1–7]. A more efficient way for the conversion of the Sun's energy to hydrogen compared to the combination of photovoltaic devices with water electrolysis [8–17], which possesses an overall efficiency between 10% and 14% [18,19], may constitute the direct photocatalytic water splitting into hydrogen and oxygen. So far, the two half reactions water oxidation and water reduction are mainly studied separately. Disadvantageously, this requires the application of sacrificial reagents (SR) as

electron donors or electron acceptors (Scheme 1) [11–17,20–29]. On the other hand, it allows for a simplification of the complex system of overall water splitting and hence a more detailed understanding of the basic processes as a prerequisite for improvements. Within the development of more efficient and stable catalyst systems for photocatalytic hydrogen generation from water, an additional trend towards the usage of 3d metals is obvious. This is motivated by several advantages of such catalysts due to lower costs and higher abundance compared to noble metals. Exemplarily, significant efforts have been reported applying cobalt and nickel complexes as proton reduction catalysts (WRC) (Although the term “water reduction catalyst” as well as the related abbreviation WRC is commonly used in relevant literature it should be named proton reduction catalyst in the strict sense) [30–40]. For example, various groups investigated cobaloxime-based catalysts [25,31]. Recently, these systems have been outperformed by pentapyridyl cobalt complexes achieving a turnover number (TON) with respect to Co of up to 11,000 with a Re-photosensitizer (PS) and ascorbic acid (SR) [36]. Prominent examples of nickel catalysts constitute the DuBois catalyst [37], as well as tris(2-pyridylthiolate) [38], 2-aminobenzenethiolate, 2-mercaptophenolate and 2-mercaptopyridyl-*N*-oxide [39] nickel complexes by Eisenberg. The DuBois system showed excellent activity with a TOF up to 100,000 s<sup>-1</sup> in the electrocatalytic hydrogen evolution reaction (HER) [41,42]. However, in photocatalytic hydrogen generation, this system achieved only a TON<sub>Ni</sub> of 2700 over 150 h with Ru photosensitizer A and ascorbic acid (SR) [37], while applying thiolate nickel complexes improved the TON<sub>Ni</sub> up to 7300 after 30 h with fluorescein, a xanthen-type organic dye, as the photosensitizer and TEA as SR [40]. A stability of more than 100 h was obtained using this system with TEOA as the electron donor [39]. More than 280,000 turnovers were obtained by replacing the PS and SR by water-soluble CdSe quantum dots and ascorbic acid, respectively [39]. Very recently, also stabilized nickel [43], iron [44], zinc [45] and copper nanoparticles [46] were reported to constitute active WRC in the photocatalytic hydrogen generation from water.



**Scheme 1.** The basic concept of hydrogen generation from water (hydrogen half reaction) via a reductive quenching pathway applying a photosensitizer (PS) and a proton reduction catalyst (WRC) in the presence of the sacrificial reductant (SR).

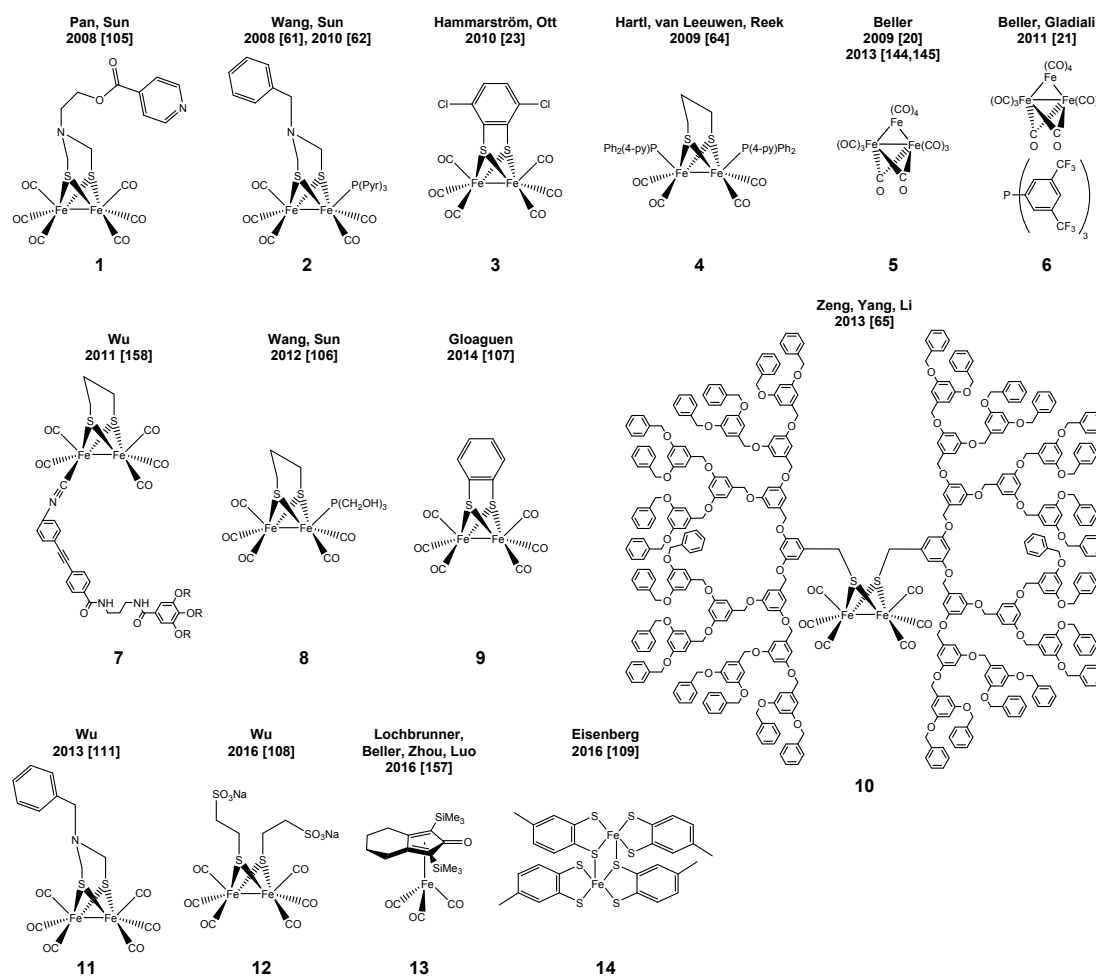
In this review, we present a brief overview about iron-based WRC, including noble metal-free PS systems. Exemplarily, mechanistic investigations and the working mode of a selected system by employing a tool of combined analytical methods are summarized.

## 2. Overview of Molecularly-Defined Iron WRCs and the Respective Photosensitizers

Iron is a very important and abundant base metal constituting, e.g., the active centers in hydrogenases [47–49] and showing impressive activities for hydrogen generation and cleavage with up to 9000 molecules H<sub>2</sub> per second and site [50]. Because of this outstanding activity, especially [FeFe] hydrogenases have attracted much attention for more than a decade, although the real structure regarding the central atom of the dithiolate bridge was identified to be nitrogen only in 2009 [51,52]. Thus, besides CO and CN ligands, the dinuclear iron complex contains an internal base, which is of fundamental importance since it allows for the heterolytic splitting/formation of H<sub>2</sub> by metal ligand

cooperative catalysis. This principle has also been adopted for other (de)hydrogenation reactions applying non-innocent ligand iron complexes [53–60].

Inspired by the above-mentioned lead structure several groups, e.g., Wang and Sun [61,62], as well as Hammarström [23], Ott [63] and Reek [64] successfully developed a variety of Fe-based WRC, which resulted in catalyst turnover numbers of up to 466 regarding Fe (Scheme 2). An overview of the applied PS is provided in Scheme 3; selected conditions and TONs are listed in Table 1. It should be noted here that a direct comparison of turnover numbers is sometimes intricate as it depends on the concentration of the involved components, i.e., WRC, PS and SR.

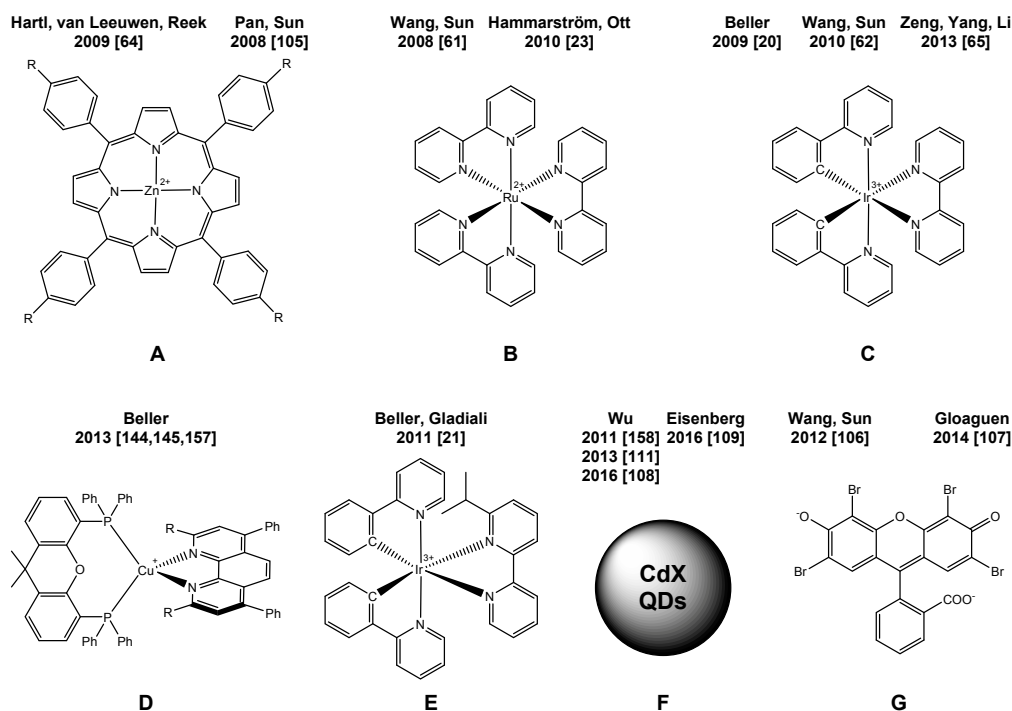


**Scheme 2.** Iron-based WRCs inspired by hydrogenases. Turnover numbers (TONs) with respect to the WRC are reported in Table 1. The used photosensitizers (PS) are shown in Scheme 3.

In 2013, Yu et al. reported a TON of 22,200 with a dendrimer-based diiron hydrogenase mimic applying only 10 nmol of the catalyst [65]. The same year, other mimics were applied for visible light-driven hydrogen production embedded in a MOF [66] or later in a protein matrix [67]. However, in these cases, productivities (TON) of only ca. 5 and 130, respectively, were achieved in the presence of Ru-PS **B** (Scheme 3).

Besides the WRC, the photosensitizer (PS) is a second important component in water reduction. For this purpose, especially ruthenium complexes have played a key role since the 1970s [68–74], later followed by various iridium [75,76], platinum [77–79] and rhenium [80–85] complexes. In contrast, more abundant metals or even metal-free photocatalytic systems were reported: examples include, e.g., iron [86], zinc [24,25,87–89] and magnesium-based [24,90–94] photosensitizers, CdTe [95], CdSe [39] or carbon [96] quantum dots or organic dyes [38–40,97–104] together with either cobalt or nickel catalysts.

Mostly, the reported activities and stabilities were still low. However, Eisenberg recently reported a TON for a nickel WRC as high as 280,000 together with water-soluble CdSe quantum dots with tripodal S-donor capping agents as PS and ascorbic acid as the sacrificial electron donor at pH 4.5 [39].



**Scheme 3.** Photosensitizers (PS) applied in iron-based water reduction systems with the WRCs depicted in Scheme 2. TONs with respect to the photosensitizer are reported in Table 1.

A few of the above-mentioned iron-based WRC were also used as a part of fully noble metal-free water reduction systems applying either Zn porphyrins [64,105] or organic dyes [106,107] instead of Ru-PS [23,61] and Ir-PS [62]. Applying the latter (i.e.,  $EY^{2-}$ ) in a micellar solution resulted in a TON of 117 (Scheme 2: 9, Scheme 3: G) [107].

Very recently, CdSe quantum dots were independently successfully applied as light-harvesting units by Wu [108] and Eisenberg [109] in a system containing different [FeFe]-Hydrogenase mimics (Scheme 2, 12 and 14). In both cases, outstanding TONs of 26,500 and 29,400, respectively, were obtained. Applying  $[Ru(bpy)_3]^{2+}$  as PS resulted in a comparably low TON of only 178 [108]. In addition, co-embedded CdSe quantum dots and the [FeFe]-hydrogenase mimic in lipid membranes were used, however resulting in a much lower TON of 651 [110]. As an alternative, also CdTe quantum dots were applied resulting in the hitherto highest observed  $TON_{Fe}$  as high as 52,800 [111] for 11 (Scheme 2).

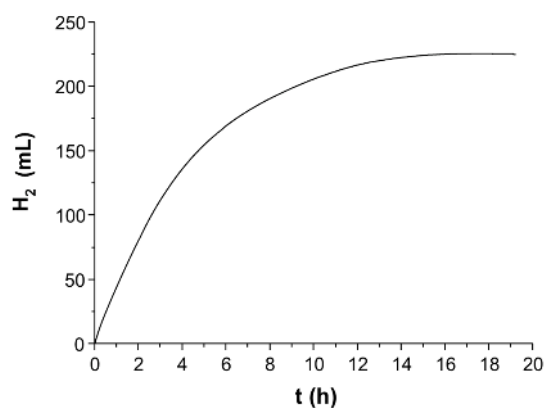
All of these results might be compared, e.g., with the use of an isolated [NiFe] hydrogenase in visible light-driven hydrogen production on  $C_3N_4/TiO_2$  with an excellent productivity (TON) of 580,000 [112].

Surprisingly, the promising application of copper (I) complexes with polypyridine ligands as photosensitizers has been neglected for a long time with only one early communication from the group of Sauvage [113]. As their photo- and electro-chemical properties have been known for decades, they are already applied as active components in organic light-emitting diodes (OLEDs), light-emitting electrochemical cells (LECs) or luminescence-based sensors and dye-sensitized solar cells (DSSCs) [114–125]. Besides, molecularly-defined copper compounds showed a significant ability to act as catalysts for electrochemical water reduction [126] and oxidation [127–130] and  $CO_2$  reduction [131]. In addition, heterogeneous cupric and cuprous oxides have been used as co-catalysts, mainly supported on titania [132–136], or as light-harvesting semiconductors [137–139] for water reduction. Although a

variety of homo- and hetero-leptic Cu complexes have been developed [121,125,140–146], only recently, two different examples of its successful application as photosensitizers in efficient noble metal-free photocatalytic systems for proton reduction have been reported (Scheme 3) [144,145,147].

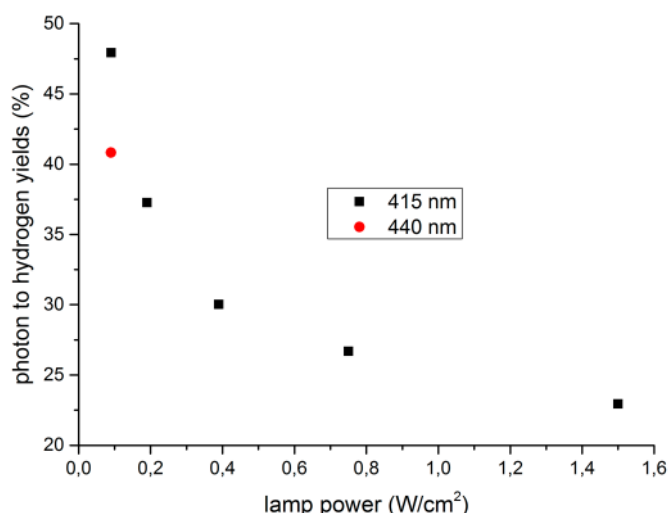
### 3. Light to Hydrogen: Development and Improvement of an Iron Carbonyl-Based Catalytic System

Within the frame of our project “Light to Hydrogen”, an initial focus laid on the development of efficient catalysts and photosensitizers based on abundant metals in order to substitute rare and expensive noble metals. The realization of a noble metal-free system for photocatalytic hydrogen generation from water was achieved in several steps. In the first step, Beller and coworkers identified simple, inexpensive, readily-available and abundant iron(0) carbonyl complexes to be appropriate WRC applying  $[(ppy)_2Ir(bpy)]PF_6$  as PS and triethylamine as SR in 10 mL of a solution of THF/TEA/H<sub>2</sub>O (4/1/1) at 25 °C under Xe-light irradiation for 3 h. In these experiments, turnover numbers with respect to Fe of 114, 141 and 132, respectively, were achieved with either Fe(CO)<sub>5</sub>, Fe<sub>2</sub>(CO)<sub>9</sub> or Fe<sub>3</sub>(CO)<sub>12</sub> as WRC precursors (Scheme 2 and Table 1, Entry 6). Applying the latter under optimized conditions (exclusion of UV light by 420 nm cut-off filter, ratio Ir-PS:WRC = 1.61:1) resulted in a maximum TON<sub>Fe</sub> of 400 and deactivation after three hours (Scheme 2) [20]. A significant improvement of the WRC was achieved by the addition of 1.5 equivalents of tris(4-trifluoromethylphenyl)phosphine P(C<sub>6</sub>H<sub>4</sub>(CF<sub>3</sub>))<sub>3</sub> or tris[3,5-bis(trifluoromethyl)phenyl]phosphine (P(C<sub>6</sub>H<sub>3</sub>(CF<sub>3</sub>)<sub>2</sub>))<sub>3</sub>. Applying the latter together with Fe<sub>3</sub>(CO)<sub>12</sub>, the *in situ*-generated WRC achieved a TON of up to 1610 after 24 h under slightly changed conditions (THF/TEA/H<sub>2</sub>O = 3/2/1, without filter) [76]. Thus, besides the productivity, also the stability of the catalyst system could be increased. Noteworthy, the incident photon to hydrogen efficiency reached 13.4% at 440 nm. Further improvement was achieved by application of novel monocationic Iridium(III)-photosensitizers with the general formula  $[Ir^{III}(C^{\cdot}N)_2(N^{\cdot}N)]^+$  (C<sup>·</sup>N: mono deprotonated cyclometallating phenyl pyridine, N<sup>·</sup>N: neutral bidentate ligand) and  $[Ir(phenyloxazole)_2(bpy)]PF_6$  (bpy: 2,2'-bipyridine) [21,22]. In this series,  $[(2\text{-phenylbenz-oxazole})_2Ir(bpy)]PF_6$  and  $[(ppy)_2Ir(6\text{-}^iPr\text{-}bpy)]PF_6$  (Scheme 3, Table 1, Entry 9) turned out to be the most efficient photosensitizer with  $[HNEt_3][HFe_3(CO)_{11}]$ /tris[3,5-bis(trifluoromethyl)phenyl]-phosphine as the water reduction catalyst and triethylamine as the sacrificial reductant. Accordingly, TONs were improved up to 2770 for the Fe-WRC and 4550 for the Ir-PS applying 15 μmol and 0.5 μmol of the photosensitizer  $[(ppy)_2Ir(6\text{-}^iPr\text{-}bpy)]PF_6$ , respectively. These experiments have been performed under 440-nm irradiation by a Lumatec mercury lamp (Figure 1).



**Figure 1.** Typical hydrogen evolution curve. Reaction conditions: 3.3 μmol  $[HNEt_3][HFe_3(CO)_{11}]$ , 15.0 μmol Ir-PS, 5.0 μmol P(C<sub>6</sub>H<sub>3</sub>-3,5-(CF<sub>3</sub>)<sub>2</sub>)<sub>3</sub>, 440-nm irradiation 1.5 W, 20 mL THF/TEA/H<sub>2</sub>O (3/2/1), 25 °C.

The incident photon to hydrogen yield was also increased to 16.4% applying this system [21]. In order to render light the limiting factor, the lamp performance was decreased to a value of  $0.09 \text{ W}\cdot\text{cm}^{-2}$ . The number of photons was then determined by chemical actinometer experiments [148]. Obviously, only 46%–65% of the provided photons were available inside the reactor depending on the wavelength and the power. Taking this into consideration, significantly improved quantum yields of up to 48% at a wavelength of 415 nm and 41% at 440 nm, respectively, were achieved (Figure 2).



**Figure 2.** Photon to hydrogen yield determination with dependence on lamp performance and wavelength. Reaction conditions:  $3.3 \mu\text{mol Fe}_3(\text{CO})_{12}$ ;  $15.0 \mu\text{mol} [(\text{ppy})_2\text{Ir}(\text{bpy})]\text{PF}_6$ ;  $5.0 \mu\text{mol}$  ligand (tris[3,5-bis(trifluoromethyl)phenyl]-phosphine); ( $\text{Fe}_3(\text{CO})_{12}/\text{PR}_3 = 1:1.5$ ); 10 mL THF/TEA/ $\text{H}_2\text{O}$  (3:2:1);  $25^\circ\text{C}$ , 20 h.

Noteworthy, a noble metal-free water reduction system could be realized applying a copper-based PS instead of the Ir complex. In this respect, we focused on heteroleptic copper complexes with the general formula  $[\text{Cu}(\text{N}^{\wedge}\text{N})(\text{P}^{\wedge}\text{P})]^+$  (D in Scheme 3, Table 1, Entries 7, 8 and 16). While the absorption and fluorescence spectra of these complexes are similar to those of ruthenium [73,74,149–153] or iridium [21] complexes with polypyridine ligands, the excited state lifetime reaches up to  $54 \mu\text{s}$  and, thus, demonstrates that the non-radiative decay is indeed slow and does not interfere with photocatalytic electron transfer processes. A key issue is the improvement of steric and electronic factors in order to get long-lived metal to ligand charge transfer (MLCT) excited states. This was achieved in heteroleptic complexes, especially those bearing the bathocuproine ligand. The methyl groups located at the 2- and 9-position are important to avoid the expansion of the coordination sphere of the metal center [154] and at the same time favor the tetrahedral ( $T_d$ ) geometry. Thus, in the MLCT excited state, the flattening to the preferred square planar geometry is significantly reduced [155,156]. A number of molecularly-defined Cu-PS has been synthesized based on various combinations of bidentate phosphines and amines. Subsequent tests as PS in the photocatalytic proton reduction in the presence of  $[\text{Fe}_3(\text{CO})_{12}]$  as WRC and TEA as SR (THF:TEA: $\text{H}_2\text{O} = 4:3:1$ ) revealed a TON of up to 1330 for the Cu-PS (Scheme 3) [144,145]. Thus, these noble metal-free systems already achieved productivities in the same order of magnitude as those containing Ru- or Ir-PS.

Finally, the activity of the catalytic proton reduction was improved applying Knölker's iron complexes instead of iron carbonyl complexes. In combination with the mentioned Cu-PS, the activity was increased up to 15 times (Scheme 2 and Table 1, Entry 16) [157]. Noteworthy, this class of complexes allows for a metal ligand cooperative catalysis, as the oxygen atom at the cyclopentadienyl ring provides a basic center, which can be protonated like the azadithiolate moiety in the  $[\text{FeFe}]$  hydrogenase, enabling facile heterolytic hydrogen generation [157].

**Table 1.** Selected reported iron-based WRCs (Scheme 2) in photocatalytic hydrogen generation in the presence of photosensitizers (Scheme 3) and sacrificial agents.

Entry	WRC	PS	Conditions <sup>1</sup>	Light ( $t_{irr}/h$ )	TON <sub>WRC</sub>	TON <sub>PS</sub>	Ref.
1	1	A R = H	10 mL DCM, 25 $\mu$ M 1; 50 $\mu$ M A; 5 mM thiosalicylic acid, 1 mM TFA	Xe, $\lambda \geq 400$ nm (2.5 h)	0.16	- <sup>2</sup>	[105]
2	2	B	5 mL ACN, 5 mL H <sub>2</sub> O, 10 $\mu$ mol 2, 1 $\mu$ mol B, 1 mmol ascorbic acid	Xe, $\lambda \geq 400$ nm (3 h)	4.3	86	[61]
3	2	C	acetone/water (9:1 v/v), 0.05 M 2, 0.05 M C, 0.14 M TEA	Xe, $\lambda \geq 400$ nm (8 h)	466	- <sup>2</sup>	[62]
4	3	B	1.5 mL DMF/H <sub>2</sub> O (1:1), 14 $\mu$ M 3, 140 $\mu$ M B, 100 mM ascorbic acid (adjusted to pH 5.5 with 1 M NaOH)	455–850 nm (2.5 h)	200	20	[23]
5	4	A R = H + OMe	5 mL toluene, 5 $\mu$ mol 4, 20 $\mu$ mol A, 50 $\mu$ mol [N <sup>3</sup> Pr <sub>2</sub> EtH] [OAc]	Xe, $\lambda \geq 390$ nm (80 min)	5	4	[64]
6	5	C	10 mL THF/TEA/H <sub>2</sub> O (4:1:1 v/v/v), 18.5 $\mu$ mol 5, 7.5 $\mu$ mol C	Xe, $\lambda \geq 420$ nm (output 1.5 W) (6 h)	400	3035	[20]
7	5	D R = Me	10 mL THF/TEA/H <sub>2</sub> O (4:1:1 v/v/v), 5 $\mu$ mol 5, 3.5 $\mu$ mol D	Xe, $\lambda \geq 200$ nm (output 1.5 W) (27 h)	547	781	[144]
8	5	D R = <sup>s</sup> Bu	10 mL THF/TEA/H <sub>2</sub> O (4:1:1 v/v/v), 5 $\mu$ mol 5, 3.5 $\mu$ mol D	Xe, $\lambda \geq 200$ nm (output 1.5 W) (60 h)	931	1330	[145]
9	6	E	20 mL THF/TEA/H <sub>2</sub> O (3:2:1, v/v/v), 6: 3.3 $\mu$ mol [Fe] + 5.0 $\mu$ mol L, 0.5 or 15.0 $\mu$ mol E	Hg, $\lambda = 440$ nm (output 1.5 W) (20 h)	2770	4550	[21]
10	7	F X = Te	10 mL H <sub>2</sub> O, 156 $\mu$ M 7, 50 $\mu$ M F, 85.2 mM ascorbic acid	Hg, $\lambda \geq 400$ nm (18 h)	505	79	[158]
11	8	G	10 mL EtOH/H <sub>2</sub> O (1:1, v/v), 0.1 mM 8, 1.0 mM G, 10% TEA	Xe, $\lambda > 450$ nm (15 h)	226	59	[106]
12	9	G	10 mL H <sub>2</sub> O, 0.1 mM 9, 20 mM sodium dodecyl sulfate, 0.2 mM G, 10 vol % TEA	LED, $\lambda = 455$ nm (output 0.3 W) (4.5 h)	117	58	[107]
13	10	C	10 mL acetone/H <sub>2</sub> O (9:1, v/v), 0.25–1.0 $\mu$ M 10, 0.5 mM C, 0.6 M TEA	Xe, 400 < $\lambda$ < 800 nm (8 h)	22,200	44	[65]
14	11	F X = Te	10 mL H <sub>2</sub> O/MeOH (3:1, v/v), 1.0 $\mu$ M 11, 0.684 $\mu$ M F + 10 mg chitosan, 0.2 M ascorbic acid	LED, $\lambda = 410$ nm (60 h)	52,800	- <sup>2</sup>	[111]
15	12	F X = Se	10 mL H <sub>2</sub> O, 1.0 $\mu$ M 12, 5.2 $\mu$ M F, 0.2 M ascorbic acid	LED, $\lambda = 520$ nm (12 h)	26,500	- <sup>2</sup>	[108]
16	13	D R = Me	10 mL THF/TEA/H <sub>2</sub> O (4:3:1 v/v/v), 12 $\mu$ mol 13, 12 $\mu$ mol Et <sub>4</sub> NOH, 3.5 $\mu$ mol D	Xe, $\lambda \geq 200$ nm (output 1.5 W) (7 h)	131	449	[157]
17	14	F X = Se	5 mL EtOH/H <sub>2</sub> O (1:1, v/v), 0.5 $\mu$ M 14, 0.2 $\mu$ M F (with 80 $\mu$ M capping agent), 0.2 M ascorbic acid	LED, $\lambda = 520$ nm (output 0.15 W) (80 h)	29,400	- <sup>2</sup>	[109]

<sup>1</sup> Abbreviations: irr, irradiation; ref., reference; DCM, dichloromethane; TFA, trifluoroacetic acid; ACN, acetonitrile; TEA, triethylamine; DMF, dimethylformamide; THF, tetrahydrofuran. <sup>2</sup> Not provided.

#### 4. Improving Mechanistic Understanding by an Approach of Combined Analytical Methods

The project “Light to Hydrogen” brought together research groups with catalytic, analytical, opto-physical, as well as theoretical expertise, which provided an excellent opportunity to gain deeper mechanistic insights. As a consequence, the Ir–Fe water (proton) reduction system belongs to the most investigated and best understood ones reported so far. Therefore, it will be demonstrated in the following section how a combination of various analytical and theoretical methods clarifies the intricate work mode of the different components including the detection of intermediates. These comprehensive mechanistic studies included Raman, NMR, EPR, *in situ*, as well as operando FTIR spectroscopy and DFT calculations [159]. In particular, the trimeric complex [HFe<sub>3</sub>(CO)<sub>11</sub>]<sup>−</sup> was identified as a key intermediate during water reduction by NMR, Raman and *in situ* FTIR spectroscopy. Comparison with a synthesized [HNEt<sub>3</sub>][HFe<sub>3</sub>(CO)<sub>11</sub>] proved the assignment of the Raman bands at 223 and 165 cm<sup>−1</sup> to this intermediate, as well as the characteristic band pattern in the IR spectrum (bands at 2064, 1999,

1993, 1975, 1941 and 1748  $\text{cm}^{-1}$ ). In addition, the monomeric  $[\text{HFe}(\text{CO})_4]^-$  species was detected via *in situ* IR spectroscopy (bands at 1998, 1908 and 1878  $\text{cm}^{-1}$ ), while EPR spectroscopy enabled the identification of the dimeric  $[\text{H}_2\text{Fe}_2(\text{CO})_7]^-$  ( $g = 2.0432$ ) and  $[\text{Fe}_2(\text{CO})_8]^{\bullet-}$  ( $g = 2.0385$ ) intermediates. The latter two dimeric species are products of a side reaction and constitute intermediates toward  $[\text{HFe}(\text{CO})_4]^-$  (M11), which acts as a resting state [160,161]. Deactivation pathways during catalysis are both CO release from the WRC mediated by light irradiation and the decomposition of the Ir-PS especially at high PS/WRC ratios. Then, the released ligand is transferred towards the iron center of the WRC to form  $[\text{Fe}(\text{CO})_3(\text{bpy})]$ , which was detected by *in situ* IR spectroscopy. Further, these assumptions are supported by DFT calculations [161,162]. Based on the obtained data, a proposal for the mechanism has been developed (Scheme 4).

In the left cycle, electrons of  $[(\text{ppy})_2\text{Ir}(\text{bpy})]^+$  (M1) are promoted via light irradiation to form the excited  $[(\text{ppy})_2\text{Ir}(\text{bpy})]^{+\bullet}$  (M2). Time-resolved photoluminescence experiments showed reductive quenching of M2 by TEA, which is the rate-determining step of the overall system. The resulting  $[(\text{ppy})_2\text{Ir}(\text{bpy})]$  (M3) possesses a reduction potential sufficiently high to reduce  $[\text{Fe}_3(\text{CO})_{11}]^{\bullet-}$  (M6), as well as  $[\text{HFe}_3(\text{CO})_{11}]^-$  (M7) in the right cycle [163]. The sacrificial reductant TEA is converted to oxidized TEA radicals and finally to acetaldehyde after hydrolysis.

The precursor  $\text{Fe}_3(\text{CO})_{12}$  (M5) is quickly reduced already in the dark by electron transfer from TEA forming the electron-rich 19-electron complex  $[\text{Fe}_3(\text{CO})_{12}]^{\bullet-}$ . This unstable complex undergoes decarbonylation to the electron-deficient 17-electron complex  $[\text{Fe}_3(\text{CO})_{11}]^{\bullet-}$ , as well as  $[\text{Fe}_2(\text{CO})_8]^{\bullet-}$  and  $\text{Fe}(\text{CO})_5$  (M6). Under light irradiation, reduction of M6 by Ir-PS $^-$  and subsequent proton transfer result in the formation of the active catalyst  $[\text{HFe}_3(\text{CO})_{11}]^-$  (M7). Subsequently, a second electron transfer is supposed to form  $[\text{HFe}_3(\text{CO})_{11}]^{2-}$  (M8) followed by a proton transfer to form  $[\text{H}_2\text{Fe}_3(\text{CO})_{11}]^-$  (M9).  $\text{H}_2$  is released from M9, and  $[\text{Fe}_3(\text{CO})_{11}]^{2-}$  (M10) is restored after further reduction by the Ir-PS $^-$  in order to restart the cycle. In parallel, M7 can also be converted into  $[\text{Fe}_2(\text{CO})_8]^{\bullet-}$  and the resting state  $[\text{HFe}(\text{CO})_4]^-$  (M11) by incorporation of CO and reduction (*vide supra*). The reverse reaction from M11 to M7 and reactivation of the WRC are mediated by irradiation [76,160,161,163].

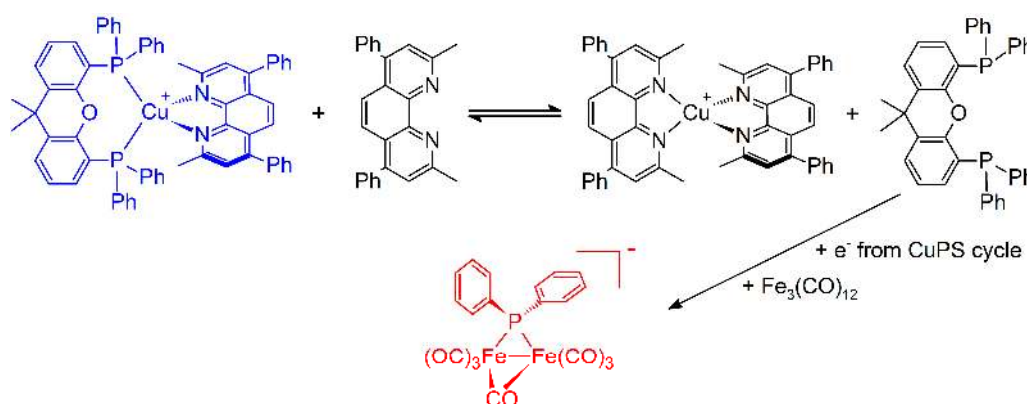
It should be noted that this broad range of methods is unparalleled for a single proton reduction system. The rich chemistry of  $\text{Fe}_3(\text{CO})_{12}$  in reaction solution in the dark is extended to several more species under illumination in the presence of an Ir-PS with sufficiently high reduction potential in its excited state.

Further, the exact role of added phosphine ligand and its enhancing effect on the TON was of special interest. Recent *in situ* FTIR experiments reveal the formation of  $[\text{Fe}_2(\text{CO})_7\text{P}(\text{C}_6\text{H}_4(\text{CF}_3))_2]^-$  (bands at 2021, 1971, 1941 and 1922  $\text{cm}^{-1}$ ), when  $\text{P}(\text{C}_6\text{H}_4(\text{CF}_3))_3$  was added to the reaction mixture (Figure 3, top). This phosphidodiferrate is more stable than the phosphorus-deficient compounds  $[\text{HFe}_3(\text{CO})_{11}]^-$  (M7) and  $[\text{HFe}(\text{CO})_4]^-$  (M11), and thus, it constitutes the only iron carbonyl complex present in solution after 15 h of light irradiation (Figure 3, right). At this time, hydrogen generation still continues (Figure 3, left, violet curve), while it stops in the absence of a phosphine ligand (Figure 3, left, grey curve). The spectral assignment of  $[\text{Fe}_2(\text{CO})_7\text{P}(\text{C}_6\text{H}_4(\text{CF}_3))_2]^-$  is confirmed by DFT calculations and comparison with  $[\text{Fe}_2(\text{CO})_7\text{PPh}_2]^-$  [164–166]. The latter shows the same IR contributions shifted by 5  $\text{cm}^{-1}$  to lower wavenumbers due to the less electron withdrawing effect of the phenyl group compared to  $\text{C}_6\text{H}_4(\text{CF}_3)$ .





Figure 4 [141,165]. Thus, a part of the released xantphos is available as the precursor for the PR<sub>2</sub> unit. Besides, the heteroleptic [Cu(bathocuproine)(xantphos)]PF<sub>6</sub> was established to present the active PS [166].



**Figure 4.** Equilibrium between hetero- and homoleptic Cu-PS and the formation of the phosphidodiferrate initiated by reductively-induced xantphos fragmentation. The active Cu-PS and WRC are highlighted in blue and red, respectively.

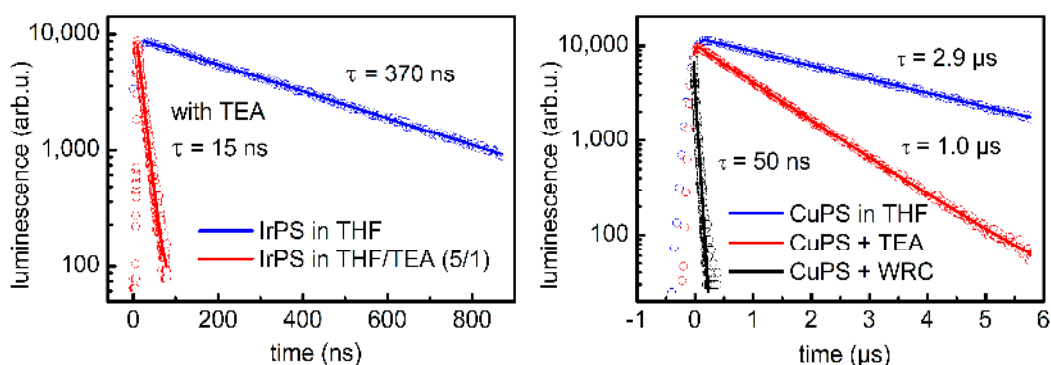
Taking the equilibrium in Figure 4 into consideration, we successfully developed an *in situ* method for the synthesis of the organometallic copper complex. Instead of the time-consuming pre-synthesis of molecularly-defined complexes, the application of this method allowed for further optimization of the quantity of ligands leading to improved system efficiency [166].

To investigate the primary steps initiated by the absorption of light for both systems, the Ir-PS/Fe-WRC, as well as the Cu-PS/Fe-WRC, time-resolved UV/Vis absorption and photoluminescence measurements applying a femtosecond pump-probe setup and a streak camera were performed. The UV/Vis absorption spectrum of the Ir-PS [(ppy)<sub>2</sub>Ir(bpy)]<sup>+</sup> C results from electronic transitions to ligand centered states and to metal to ligand charge transfer (MLCT) states involving the ppy- and the bpy-ligands at longer wavelengths. After optical excitation ultrafast intersystem crossing (ISC) takes place within the time resolution of the pump-probe experiment of 150 fs [167]. Then, vibrational redistribution and relaxation within the triplet manifold results in the population of the lowest triplet MLCT state on the picosecond timescale. In this state, the excited electron is located on the bpy-ligand, and if the ppy-ligand is originally excited, an ultrafast interligand charge transfer is observed [167].

In the case of the Cu-PS D, ISC takes 7 ps and is preceded by a flattening of the complex structure within the first picosecond [156]. As in the case of the Ir-PS C, the long lifetime of the resulting <sup>3</sup>MLCT is a crucial factor for the performance as a sensitizer in photocatalysis.

The electron transfer steps between the photosensitizer and the SR and the WRC were studied by quenching of the photoluminescence from the <sup>3</sup>MLCT state (Figure 5). In pure THF, the Ir-PS [(ppy)<sub>2</sub>Ir(bpy)]<sup>+</sup> C exhibited a lifetime of 0.37 μs, which reduces drastically to 15 ns in the presence of 17 vol % TEA (Figure 5, left). This is caused by reductive quenching, i.e., an electron transfer from the SR to the metal center of the Ir-PS. The quenching rate can be compared with the calculated diffusion rate. It turns out that the transfer happens only for a small fraction of collisions between TEA and Ir-PS molecules since a specific collision geometry is required for this process [163]. However, the high TEA concentration compensates for the low collision yield and guarantees an effective transfer. At the applied WRC concentrations, the WRC does not significantly contribute to the quenching. Rather, it accepts the electron from the already reduced Ir-PS, and the transfer chain follows a reductive pathway. In the case of the Cu-PS, the situation is different (Figure 5, right). The lifetime of the <sup>3</sup>MLCT in THF is 2.9 μs at a Cu-PS concentration of 0.35 mM. This concentration is also applied in the photocatalytic experiments, but results already in some self-quenching of the PS. If it is reduced to

0.02 mM, the lifetime doubles to 6.4  $\mu\text{s}$  [145]. Adding TEA (17 vol %) causes a moderate reduction of the luminescence lifetime to 1  $\mu\text{s}$ , while it strongly decreases down to 50 ns if the TEA solution also contains 0.5 mM of the WRC. Obviously, the first electron transfer step is here from the Cu-PS to the WRC and is hence associated with an oxidation of the Cu-PS. The reduction of the sensitizer by the SR back to its original configuration occurs as the second step on longer time scales. Accordingly, the electron transfer sequence proceeds via an oxidative route. However, the dominant pathway depends on the specific photocatalytic system, as was shown by the application of Knölker's iron complexes **13** with Cu-PS **D**, where the reductive pathway is dominant again [157].



**Figure 5.** Time resolved photoluminescence and fitted exponential decays of the Ir-PS [(ppy)<sub>2</sub>Ir(bpy)]<sup>+</sup> **C** (left) and the Cu-PS **D** (right) in THF. Adding 17 volume percent of TEA (red) results in quenching of the luminescence with respect to pure solvent (blue). This effect is particular strong for the Ir-PS, while in the case of the Cu-PS, the addition of 0.5 mM WRC [HFe<sub>3</sub>(CO)<sub>11</sub>]<sup>−</sup> (black) leads to a much faster luminescence decay.

## 5. Summary and Outlook

The productivities of non-noble metal-based WRCs were significantly increased during the last 6–8 years: Thus, the TONs for Ni and Co were improved by four and three orders of magnitude, respectively. The same trend is obvious for iron. While the first successful experiments containing an iron-based WRC gave a TON of 90 in 2006, already five years later, a TON of 2770 was achieved. This trend is continuing, as is demonstrated by the latest records of >58,000 turnovers in 2013. Thus, these WRCs already show a comparable or even better performance compared to previously applied Pt, Rh- or Ru-based WRCs. An additional trend focuses on the development of fully noble metal-free photocatalytic systems in order to substitute expensive Ir and Ru photosensitizers by semiconductors, quantum dots or organic dyes.

In this article, we present the development of a water reduction system containing only base metals. In the first step, a readily-available, simple and inexpensive iron carbonyl compound has been shown to catalyze the proton reduction together with an iridium photosensitizer as the light-harvesting component. The performance of this Fe-WRC was stepwise improved up to a TON of 2770 by the addition of an appropriate phosphine and the development of the more efficient PS [(ppy)<sub>2</sub>Ir(6-*i*Pr-bpy)]PF<sub>6</sub>. The active Fe-WRC was identified to be a phosphidodiferrate [Fe<sub>2</sub>(CO)<sub>7</sub>P(C<sub>6</sub>H<sub>4</sub>(CF<sub>3</sub>)<sub>2</sub>)<sub>2</sub>]<sup>−</sup>. Quantum efficiencies of up to 48% were achieved with these systems applying blue light (415 nm). Finally, we successfully substituted the Ir-PS by molecularly-defined heteroleptic copper photosensitizers of the general formula [Cu(N<sup>∗</sup>N)(P<sup>∗</sup>P)]PF<sub>6</sub>. A combination of Fe<sub>3</sub>(CO)<sub>12</sub> and Cu-PS in the light-driven water reduction using TEA as the sacrificial reagent resulted in turnover numbers of up to 1330 and a stability of more than two days. Noteworthy, the Cu-PS can even be generated in an *in situ* method from the precursors during light irradiation, thus avoiding time-consuming synthesis procedures. Besides the catalytic tests, a careful investigation of the reaction mechanism allowed for a deep insight into the reaction mode, including loss channels.

Spectroscopic studies showed that after optical excitation of the photosensitizer, intramolecular processes result in a fast population of the lowest <sup>3</sup>MLCT state, which itself has an intrinsic lifetime of some tenth to several ten microseconds depending on the sensitizer. This long lifetime allows for efficient electron transfer processes from the SR and to the WRC. Whether the reductive or the oxidative pathway prevails depends again on the specific photocatalytic system.

So far, homogeneous catalyst materials have been tested mainly in the water reduction half reaction applying sacrificial reagents as electron donors. This allows for a simplification of the complex system of the overall water splitting and, hence, a more detailed understanding of the basic processes as a prerequisite for improvements. However, a challenge for the next time has to be the application of the developed catalyst systems in photocatalytic or photoelectrocatalytic overall water splitting. This will help avoid the undesired influences of sacrificial reagents onto the catalytic systems, like side reactions, deactivation and changing driving forces.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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