

# Magnesium Borohydride: From Hydrogen Storage to Magnesium Battery\*\*

Rana Mohtadi,\* Masaki Matsui, Timothy S. Arthur, and Son-Jong Hwang

Since Bogdanović and Schwickardi illustrated the possibility of reversibly storing hydrogen in sodium alanate,<sup>[1]</sup> extensive research efforts have been dedicated to investigating the hydrogen storage potential of complex metal hydrides. In particular, borohydrides have attracted great interest because of their superior gravimetric hydrogen content.<sup>[2]</sup> Of these, magnesium borohydride  $\text{Mg}(\text{BH}_4)_2$ , first reported in 1950<sup>[3]</sup> and more recently studied for hydrogen storage, has attracted attention because of its relatively low hydrogen-release temperature and reversibility.<sup>[2a,4]</sup> Furthermore, borohydrides are strong reducing agents that are widely used in organic and inorganic syntheses. This reducing power translates to high stability against electrochemical reduction; this stability could be exploited in highly reductive environments, such as battery anodes. Therefore, for the first time, we have conducted research towards harnessing this property of borohydrides for their use in rechargeable batteries. In particular, we have been focusing on utilizing a  $\text{Mg}(\text{BH}_4)_2$  based electrolyte in a rechargeable magnesium battery.

Recently, magnesium batteries have received increased attention as alternatives to the lithium-based battery because of the high volumetric capacity ( $3832 \text{ mA h cm}^{-3}$ ), improved safety (nondendritic), and abundance of Mg metal.<sup>[5]</sup> Despite the potential of Mg batteries, several key challenges need to be overcome for this technology to become viable. For instance, current state-of-the-art electrolytes use organomagnesium salts and complexes as they are the only ones known to be compatible with the Mg anode that allow for reversible electrochemical Mg deposition and stripping.<sup>[5b,6]</sup> Although some of these electrolytes have shown impressive stability against electrochemical oxidation, they were also found to be corrosive.<sup>[6]</sup> This property was attributed to the presence of chlorides in either/both their cations and anions.<sup>[6]</sup> Conventional inorganic and ionic salts such as  $\text{Mg}(\text{ClO}_4)_2$  were found to be incompatible with the Mg anode as a result of the

formation of an ion-blocking layer formed by their electrochemical reduction.<sup>[6]</sup> Hence, the discovery of halide-free electrolytes with high reductive stabilities is crucial for realizing a practical rechargeable Mg battery system.

Herein, we propose a new class of electrolytes based on  $\text{Mg}(\text{BH}_4)_2$  for a Mg battery. We show the first example of electrochemical reversible Mg deposition/stripping in a halide-free inorganic salt in both tetrahydrofuran (THF) and dimethoxyethane (DME) solvents. An increase of several orders of magnitude in the current densities, and high coulombic efficiencies of up to 94% are observed in DME when  $\text{LiBH}_4$  is used as an additive. Furthermore, we use this electrolyte in a rechargeable Mg battery, thus giving the first example of a borohydride electrolyte in a battery system. This work also illustrates the unique properties of borohydrides and opens the door for designing a whole new class of electrolytes for Mg batteries.

Mg deposition/stripping was studied for  $\text{Mg}(\text{BH}_4)_2$  in ether solvents. Figure 1a shows the cyclic voltammogram obtained for 0.5 M  $\text{Mg}(\text{BH}_4)_2/\text{THF}$  where a reversible reduction–oxidation process took place with onsets at  $-0.6 \text{ V}/0.2 \text{ V}$  and a 40% coulombic efficiency (Figure 1a, inset), thus indicating reversible Mg deposition and stripping. X-ray diffraction (XRD) confirmed that the deposited product from the galvanostatic reduction of the above solution (Figure 1b) was hexagonal Mg, hereby establishing the compatibility of  $\text{Mg}(\text{BH}_4)_2$  with Mg metal. The electrochemical oxidative stabilities measured on platinum, stainless steel, and glassy carbon electrodes were 1.7, 2.2, and 2.3 V, respectively (Figure S7). These results showed that for the first time: 1)  $\text{Mg}(\text{BH}_4)_2$  is electrochemically active in THF, that is, ionic conduction is possible, and 2) reversible magnesium deposition/stripping from an inorganic, relatively ionic (Mg Bader charge is  $+1.67$ )<sup>[7]</sup> and halide-free salt is feasible. Although these results are promising, to make this electrolyte more practical for use in batteries the electrochemical performance needs to be improved by lowering the overpotentials, and achieving higher current density and coulombic efficiency. In addition, the demonstration of this performance in less-volatile solvents would make  $\text{Mg}(\text{BH}_4)_2$  based electrolytes even more practical. Therefore, DME was selected (its boiling temperature is  $19^\circ\text{C}$  higher than that of THF) for further investigations. The cyclic voltammogram obtained for 0.1 M  $\text{Mg}(\text{BH}_4)_2/\text{DME}$  is shown in Figure 1c where a substantial improvement in the electrochemical performance compared to  $\text{Mg}(\text{BH}_4)_2/\text{THF}$  was evident from: 1) a 10-fold increase in the current density, 2) a reduction in the overpotentials (deposition/stripping onsets at  $-0.34 \text{ V}/0.03 \text{ V}$  versus  $-0.6 \text{ V}/0.2 \text{ V}$  in THF), and 3) a higher coulombic efficiency of 67% (40% in THF). These findings suggested that the Mg

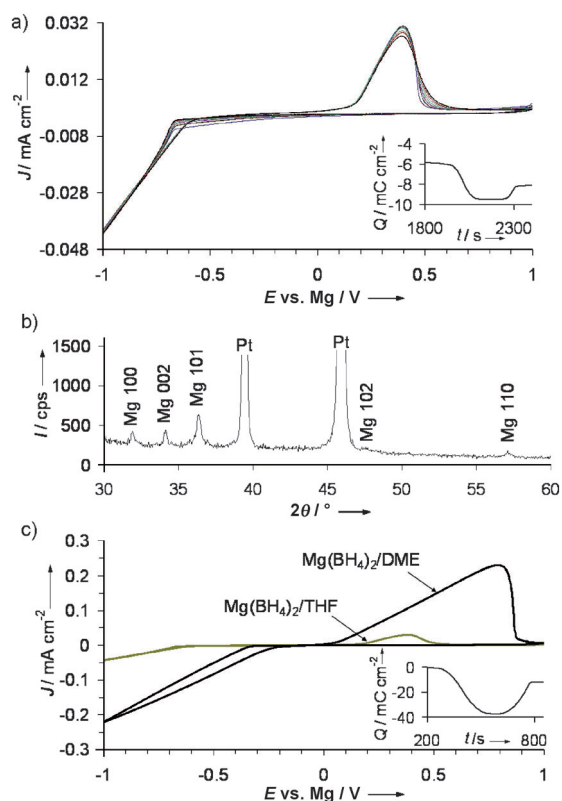
[\*] Dr. R. Mohtadi, Dr. M. Matsui, Dr. T. S. Arthur  
Materials Research Department  
Toyota Research Institute of North America  
Ann Arbor, MI 48105 (USA)  
E-mail: rana.mohtadi@tema.toyota.com

Dr. S.-J. Hwang  
Division of Chemistry and Chemical Engineering  
California Institute of Technology (USA)

[\*\*] The authors would like to thank Dr. T. Matsunaga and Dr. C. Bucur for the discussions. Funding was provided by Toyota Motor Engineering and Manufacturing North America Inc.

Supporting information for this article is available on the WWW under <http://dx.doi.org/10.1002/anie.201204913>.

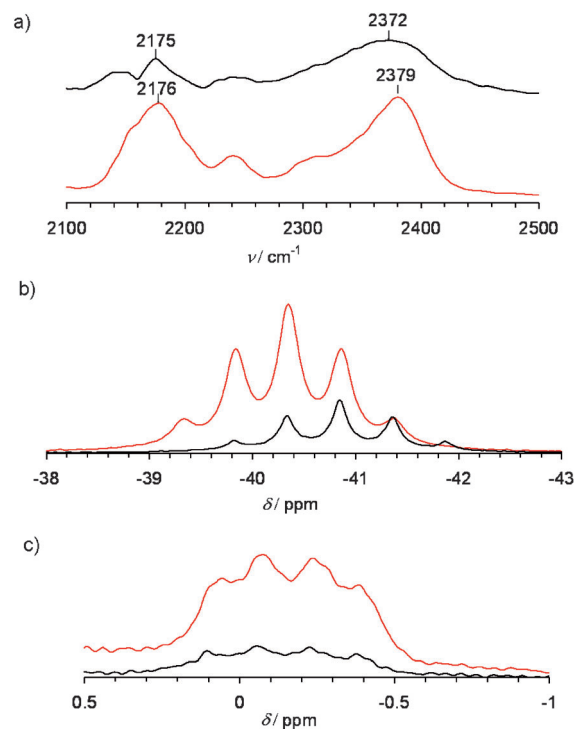
Re-use of this article is permitted in accordance with the Terms and Conditions set out at <http://angewandte.org/open>.



**Figure 1.** For 0.5 M  $\text{Mg}(\text{BH}_4)_2/\text{THF}$ : a) Cyclic voltammogram (8 cycles), inset shows deposition/stripping charge balance (third cycle), and b) XRD results following galvanostatic deposition of Mg on a Pt working electrode. c) Cyclic voltammogram for 0.1 M  $\text{Mg}(\text{BH}_4)_2/\text{DME}$  compared to 0.5 M  $\text{Mg}(\text{BH}_4)_2/\text{THF}$ . Inset shows deposition/stripping charge balance for  $\text{Mg}(\text{BH}_4)_2/\text{DME}$ . All experiments used Pt working electrode and Mg reference/counter electrodes.

electroactive species was present in higher concentration and had increased mobility in DME despite the lower solubility of  $\text{Mg}(\text{BH}_4)_2$  in DME versus THF.

These results demonstrated that for the  $\text{Mg}(\text{BH}_4)_2$  electrolyte, the electrochemical performance in DME is higher than that in THF. In contrast, organomagnesium electrolytes show an optimum electrochemical performance in THF.<sup>[5b]</sup> To further improve the electrochemical performance, it was pertinent to characterize the electroactive species in  $\text{Mg}(\text{BH}_4)_2$  solutions. Therefore, IR and NMR spectroscopic analyses (Figure 2) were conducted for 0.5 M  $\text{Mg}(\text{BH}_4)_2/\text{THF}$  and 0.1 M  $\text{Mg}(\text{BH}_4)_2/\text{DME}$ . The IR B–H stretching region (2000–2500  $\text{cm}^{-1}$ ) showed two strong widely separated bands ( $\text{Mg}(\text{BH}_4)_2/\text{THF}$ : 2379  $\text{cm}^{-1}$ , 2176  $\text{cm}^{-1}$  and  $\text{Mg}(\text{BH}_4)_2/\text{DME}$ : 2372  $\text{cm}^{-1}$ , 2175  $\text{cm}^{-1}$ ); note that the spectra for 0.1 and 0.5 M of  $\text{Mg}(\text{BH}_4)_2$  in THF are similar (Figure S2). These IR spectra are similar to those of covalent borohydrides<sup>[8]</sup> and those of  $\text{Mg}(\text{BH}_4)_2$  solvates from THF and diethyl ether<sup>[9]</sup> where two hydrogen atoms in  $\text{BH}_4^-$  are forming a bridge to one metal atom ( $\mu$  bonding). Therefore, we assigned the bands at the higher and lower B–H frequencies to terminal and bridging B–H vibrations ( $\text{B}-\text{H}_t$  and  $\text{B}-\text{H}_b$ ), respectively. The band and shoulder at 2304 and 2240  $\text{cm}^{-1}$  were assigned to asymmetric  $\text{B}-\text{H}_t$  and  $\text{B}-\text{H}_b$  vibrations, respectively. As complete disso-



**Figure 2.** For  $\text{Mg}(\text{BH}_4)_2$  in THF (red line) and in DME (black line): a) IR spectra, b)  $^{11}\text{B}$  NMR spectra, and c)  $^1\text{H}$  NMR spectra.

ciation of  $\text{Mg}(\text{BH}_4)_2$  into discrete ions is unlikely (as other borohydrides are in ethers),<sup>[10]</sup> we propose that  $\text{Mg}(\text{BH}_4)_2$  is present as the contact ion pair  $\text{Mg}\{(\mu\text{-H})_2\text{BH}_2\}_2$ , which partially dissociates into  $[\text{Mg}\{(\mu\text{-H})_2\text{BH}_2\}]^+$  and  $\text{BH}_4^-$  as in [Eq. (1)]; since the different B–H bands most likely overlap, it is not possible to distinguish all the species.



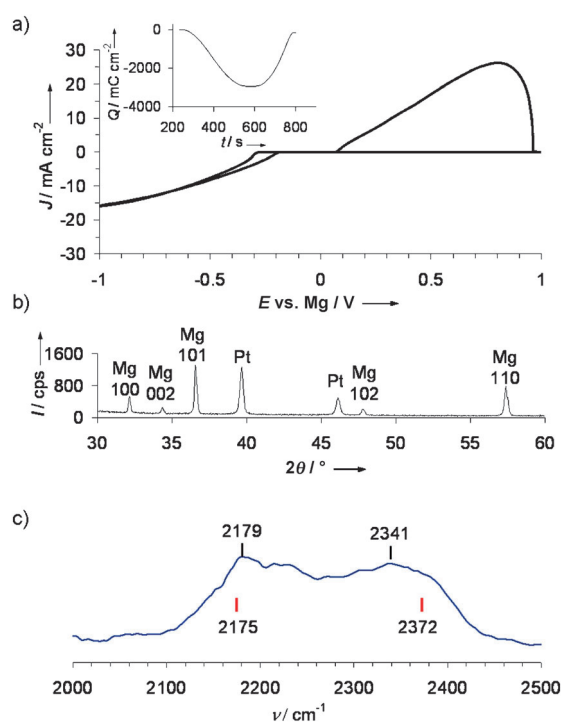
Where  $[\text{Mg}\{(\mu\text{-H})_2\text{BH}_2\}]^+$  may further dissociate:



For the spectrum of  $\text{Mg}(\text{BH}_4)_2/\text{DME}$ , although the main features present in the spectrum of  $\text{Mg}(\text{BH}_4)_2/\text{THF}$  were retained, the  $\nu\text{B}-\text{H}_t$  band is broader and shifted to a lower value and the  $\nu\text{B}-\text{H}_b$  intensity is relatively weaker. Although  $\nu\text{B}-\text{H}_t$  band broadening suggests a pronounced presence of a species similar to that found in THF, the shift in the band maximum indicates a more-ionic B–H bond (the  $\nu\text{B}-\text{H}_t$  shift is similar to those resulting from  $\text{BH}_4^-$  ions that have enhanced ionic character, such as in stabilized covalent borohydrides).<sup>[8]</sup> In addition, the relative weakening in  $\nu\text{B}-\text{H}_b$  intensity suggests that there is more free  $\text{BH}_4^-$ . The NMR spectrum of  $\text{BH}_4^-$  in DME (Figure 2b and c) indicates that there is increased boron shielding as the associated signal is shifted by about 0.5 ppm (quintet in  $^{11}\text{B}$  NMR spectrum), and slightly reduced proton shielding (0.01 ppm, quartet in  $^1\text{H}$  NMR spectrum); these results are consistent with B–H bonds that have a higher ionic character than those in  $\text{BH}_4^-$  in THF (distinguishing  $\text{B}-\text{H}_t$  from  $\text{B}-\text{H}_b$  is not possible likely

because of rapid hydrogen exchange). These findings are evidence of weaker interactions between  $\text{Mg}^{2+}$  and  $\text{BH}_4^-$  within the ion pair and an enhanced dissociation in DME [Eq. (1) and (2)]. So despite the fact that DME has a slightly lower dielectric constant (7.2) compared to THF (7.4), its chelation properties (owing to the presence of two oxygen sites per molecule)<sup>[11]</sup> resulted in an enhanced dissociation and thus an improved electrochemical performance.

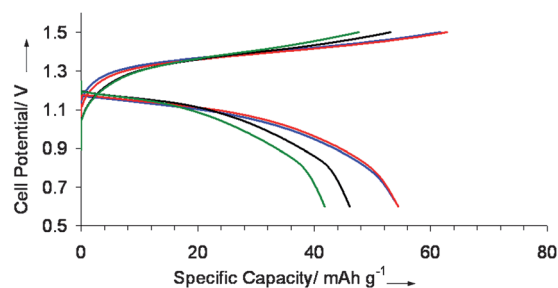
Based on the understanding gained of the nature of  $\text{Mg}(\text{BH}_4)_2$  in solution, we hypothesized that electrochemical performance would be enhanced when the association within the ion pair is weakened. To achieve this, an additive that has an acidic cation with the following characteristics is desirable: 1) reductive stability comparable to  $\text{Mg}(\text{BH}_4)_2$ , 2) nonreactive, 3) halide free, and 4) soluble in DME. Hence,  $\text{LiBH}_4$  was selected as it fulfils all of the above criteria. Mg deposition and stripping was studied in DME using different molar ratios of  $\text{LiBH}_4$  to  $\text{Mg}(\text{BH}_4)_2$ ; Figure 3 a shows the cyclic voltammo-



**Figure 3.** For  $\text{LiBH}_4$  (0.6 M)/ $\text{Mg}(\text{BH}_4)_2$  (0.18 M) in DME: a) Cyclic voltammogram (inset shows deposition/stripping charge balance). b) XRD results following galvanostatic deposition of Mg on a Pt disk. c) IR spectra (red | indicates band maxima for  $\text{Mg}(\text{BH}_4)_2/\text{DME}$ ).

gram obtained for 3.3:1 molar  $\text{LiBH}_4$  to  $\text{Mg}(\text{BH}_4)_2$  (Figure S8a and S8b show the cyclic voltammograms for different concentrations). The use of  $\text{LiBH}_4$  resulted in an increase of two orders of magnitude in the current density (i.e. oxidation peak current  $J_p = 26 \text{ mA cm}^{-2}$ ), and in a higher coulombic efficiency of up to 94%. We attribute the deposition/stripping currents solely to Mg because of the absence of Li after galvanostatic deposition (Figure 3b), and also the lack of electrochemical activity in a  $\text{LiBH}_4/\text{DME}$  solution (Figure S8a). The ionic character of  $\text{BH}_4^-$  was enhanced, as

evidenced by lower  $\nu\text{B-H}_i$  and higher  $\nu\text{B-H}_b$  bands in the IR spectrum (Figure 3c), thus implying that  $\text{LiBH}_4$  has a role in increasing  $\text{Mg}(\text{BH}_4)_2$  dissociation (the B-H bands for  $\text{LiBH}_4/\text{DME}$  occur at lower values, Figure S9). The coulombic efficiency was proportional to the molar ratios of  $\text{LiBH}_4/\text{Mg}(\text{BH}_4)_2$  (Figure S10). A rechargeable Mg battery with a Chevrel phase  $\text{Mo}_6\text{S}_8$  cathode, an Mg metal anode, and this optimized electrolyte (Figure 4) demonstrated reversible cycling capabilities at a  $128.8 \text{ mA g}^{-1}$  rate (capacity retention and cathode magnesiation are shown in Figure S11 and Figure S12). We are currently investigating the sources of the overcharge and capacity fade.



**Figure 4.** Charge/discharge profiles with Mg anode/Chevrel phase cathode for 3.3 molar  $\text{LiBH}_4/\text{Mg}(\text{BH}_4)_2$  in DME. Cycle 1 (blue), cycle 2 (red), cycle 20 (black), cycle 40 (green).

In summary, unprecedented reversible Mg deposition and stripping from an inorganic and relatively ionic salt was obtained in THF and DME. Higher current density and lower overpotentials were achieved in DME compared to those in THF. Substantial enhancement in the coulombic efficiency and the current density was accomplished by the addition of  $\text{LiBH}_4$ . Battery performance was demonstrated using a Chevrel phase cathode. Although the oxidative stability (1.7 V vs. Mg on platinum) currently limits  $\text{Mg}(\text{BH}_4)_2$  utilization with high voltage cathodes, reversibility in the absence of halides and THF makes this salt extremely unique and these findings very important for designing a whole new class of  $\text{Mg}(\text{BH}_4)_2$  based electrolytes. Currently, we are investigating improving the oxidative stability of  $\text{Mg}(\text{BH}_4)_2$ . In addition, the exact nature of the electroactive species in the presence and the absence of the additive is being studied to guide the design of  $\text{Mg}(\text{BH}_4)_2$  based electrolytes. This work provides a stepping stone for extending the applications of  $\text{Mg}(\text{BH}_4)_2$  and underscores the beauty and versatility of the chemistry of borohydrides.

### Experimental Section

Magnesium borohydride ( $\text{Mg}(\text{BH}_4)_2$ , 95%) lithium borohydride ( $\text{LiBH}_4$ , 90%), anhydrous tetrahydrofuran (THF), and dimethoxyethane (DME) were purchased from Sigma-Aldrich. Cyclic voltammetry was conducted in a three-electrode cell with Mg wire/ribbon as reference/counter electrodes. The electrochemical testing was conducted in an argon filled glovebox with  $\text{O}_2$  and  $\text{H}_2\text{O}$  amounts kept

below 0.1 ppm. Details of the analyses and battery testing conducted are described in the Supporting Information.

Received: June 23, 2012

Published online: August 21, 2012

**Keywords:** borohydrides · electrochemistry · hydrogen · magnesium

- 
- [1] B. Bogdanović, M. Schwickardi, *J. Alloys Compd.* **1997**, 253–254, 1.
- [2] a) H.-W. Li, Y. Yan, S. Orimo, A. Züttel, C. M. Jensen, *Energies* **2011**, 4, 185; b) R. Mohtadi, P. Sivasubramanian, S.-J. Hwang, A. Stowe, J. Gray, T. Matsunaga, R. Zidan, *Int. J. Hydrogen Energy* **2012**, 37, 2388.
- [3] E. Wiberg, R. Bauer, *Z. Naturforsch. B* **1950**, 5, 397.
- [4] K. Chłopek, C. Frommen, A. Leòn, O. Zabara, M. Fichtner, *J. Mater. Chem.* **2007**, 17, 3496.
- [5] a) D. Aurbach, Z. Lu, A. Schechter, Y. Gofer, H. Gizbar, R. Turgeman, Y. Cohen, M. Moshkovich, E. Levi, *Nature* **2000**, 407, 724; b) N. Amir, Y. Vestfrid, O. Chusid, Y. Gofer, D. Aurbach, *J. Power Sources* **2007**, 174, 1234; c) M. Armand, J. Tarascon, *Nature* **2008**, 451, 652; d) H. S. Kim, T. S. Arthur, G. D. Allred, J. Zajicek, J. G. Newman, A. E. Rodnyansky, A. G. Oliver, W. C. Boggess, J. Muldoon, *Nat. Commun.* **2011**, 2, 427.
- [6] J. Muldoon, C. B. Bucur, A. G. Oliver, T. Sugimoto, M. Matsui, H. S. Kim, *Energy Environ. Sci.* **2012**, 5, 5941.
- [7] Z. Łodziana, M. J. Van Setten, *Phys. Rev. B* **2010**, 81, 024117.
- [8] a) N. Davies, P. H. Bird, M. G. H. Wallbridge, *J. Chem. Soc. A* **1968**, 2269; b) P. Sivasubramanian, R. Mohtadi, R. Zidan, K. Pariyadat, C. L. Levertte, M. L. Fetterolf, *Appl. Spectrosc.* **2012**, 66, 591.
- [9] a) H. Nöth, *Z. Naturforsch. B* **1982**, 37, 1499; b) M. Bremer, H. Nöth, M. Warchhold, *Eur. J. Inorg. Chem.* **2003**, 111.
- [10] A. E. Shirk, D. F. Shriver, *J. Am. Chem. Soc.* **1973**, 95, 5901.
- [11] N. S. Poonia, A. V. Bajaj, *Chem. Rev.* **1979**, 79, 389.
-