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Mechanical Property and Structure of Covalent Functionalised Graphene/Epoxy Nanocomposites

SUBJECT AREAS:
MECHANICAL AND
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NANOPARTICLES

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Thermally reduced graphene nanoplatelets were covalently functionalised via Bingel reaction to improve their dispersion and interfacial bonding with an epoxy resin. Functionalised graphene were characterized by microscopic, thermal and spectroscopic techniques. Thermal analysis of functionalised graphene revealed a significantly higher thermal stability compared to graphene oxide. Inclusion of only 0.1 wt% of functionalised graphene in an epoxy resin showed 22% increase in flexural strength and 18% improvement in storage modulus. The improved mechanical properties of nanocomposites is due to the uniform dispersion of functionalised graphene and strong interfacial bonding between modified graphene and epoxy resin as confirmed by microscopy observations.

Most of all graphitic materials, graphene is a single layer of sp^2 carbon atoms and is the building block for all carbon allotropes such as fullerene and carbon nanotubes. Since the advent of graphene in 2004¹, the scientific community interest in graphene and graphene based materials has grown exponentially and the number of publications increased dramatically². This interest mainly stems from a combination of incredible mechanical, electrical and thermal properties of graphene. Before graphene, another famous carbon family member- carbon nanotubes had attracted numerous attention however recent studies have shown that graphene could surpass the properties of carbon nanotubes². This is attributed to the incredibly high specific surface area, unique graphitised plane structure and extremely high charge mobility of graphene³.

In order to improve the properties of polymers, graphene based sheet has been incorporated both into thermosetting⁴⁻⁸ and thermoplastic polymers⁹⁻¹². Similar to inclusion of other nanofillers such as clay and CNTs and regardless of polymer matrix used, the challenge in processing of graphene reinforced composites is to achieve a strong interface and a uniform dispersion of graphene nanoplatelets within polymer matrix¹³. The true realization of remarkable properties of graphene could only happen if graphene nanoplatelets are well dispersed in matrix and that there is a strong interfacial adhesion between graphene nanoplatelets and polymer matrix¹⁴.

To this end, efforts have been made to develop approaches to tackle processing issues associated with graphene based materials. Chemical functionalization of graphene could facilitate uniform dispersion and improve compatibility with polymer matrices. In addition, chemical functionalization allows tuning of the chemical and physical properties of resulting materials which may give rise to novel properties^{15,16}. Chemical functionalization of graphene has been achieved through covalent and non-covalent bonding of various organic functional groups such as isocyanate¹⁷, polystyrene¹⁸, polyglycidyl methacrylate¹⁹, triphenylene⁴, dopamine²⁰. While in covalent functionalization approach functional groups form chemical bonds with the group presenting on graphene surface, non-covalent functionalization method is mainly based on weak van der Waals force or π - π interaction of aromatic molecules on the graphene basal plane and non-covalent functional groups. Although non-covalent attachment avoids the interruption of graphitic planes in graphene giving rise to less defect sites, covalent functionalization ensures a strong and irreversible bond between the polymer matrix and graphene¹⁴.

Among various approaches to covalent functionalization of graphene surface, cycloaddition reaction has been successfully applied for the organic functionalization of carbon based nanostructures including chemically converted graphene^{21,22}. The significance of this chemistry method lies in its versatility in introducing organic derivatives which enables application in biotechnology, energy and polymer composites^{23,24}. Herein we report



organic functionalization of thermally reduced graphene via Bingel reaction which is an example of 2,1-cycloaddition reaction. The Bingel cyclopropanation has been commonly employed for the functionalization of fullerenes²⁵ and single walled carbon nanotubes (SWNTs)^{26,27}.

More recently, solution exfoliated graphite was chemically modified using Bingel reaction with the aid of sonication and microwave irradiation which led to highly functionalised graphene with good dispersability in organic solvents²⁸. Despite the ease of Bingel reaction conditions, to date very few studies if any applied cyclopropanation reaction on thermally reduced graphene and investigated the influence of resulting functionalised graphene on composite properties.

In the present work, thermally reduced graphene was functionalised using Bingel reaction and the Bingel modified graphene was incorporated into epoxy resin. The morphology and mechanical properties of Bingel functionalised graphene/epoxy composites were investigated to evaluate the effect of chemical functionalization on the dispersion status and interface in the resulting composites.

Experimental

Materials. Natural flake graphite used in this study was kindly provided by Asbury Carbons (grade 3243). All reagents were purchased from Sigma-Aldrich and used as received. Both Epon 862 epoxy resin (Bisphenol F) and Epikure W aromatic amine curing agent were purchased from Hexion Specialty Chemicals.

Preparation and chemical functionalization of graphene. Natural flake graphite was oxidised in a solution of nitric acid, sulphuric acid and potassium chlorate for 96 h²⁹. Graphene oxide (GO) was dried in a vacuum oven at 80°C for 24 h. Vacuum dried GO was placed in a quartz tube and flushed with Ar for 10 min. The quartz tube was then quickly inserted into a Lindberg tube furnace preheated to 1050°C and held in the furnace for 30 s³⁰ to obtain thermally reduced graphene (TRG). Functionalisation of TRG was conducted according to Bingel reaction reported by Coleman and co-workers²⁶. A suspension of TRG in dry dichlorobenzene (DCB) was prepared by sonication. A mixture of 1,8-diazabicyclo[5.4.0]undecene (DBU) and diethylbromomalonate were added and stirred under inert atmosphere for 15 h. The mixture was filtered and washed with DCB followed by ethanol. The resulting product was dispersed and stirred in a mixture of NaOH, methanol and THF and stirred at room temperature and was then centrifuged³¹. The supernatant solution was discarded and the sample was dispersed in HCl solution at room temperature. The mixture was centrifuged and washed a few times by distilled water. The functionalised graphene hereafter

denoted as FG, was then dried at 60°C for 5 h in vacuum oven. The synthesis method is illustrated in Figure 1.

Composite preparation. Graphite nanoplatelets (0.1 wt%) were dispersed in acetone (~1 mg/ml) using a probe sonicator (1000 W, 20 KHz) for 30 min. The epoxy resin was then added and the suspension was ultrasonicated for one more hour. During the ultrasonication, the mixture was kept cool by an ice bath. To remove acetone, the above mixture was heated at 70°C while constantly stirring before heating in a vacuum oven at 70°C for 8 hours. Epicure W curing agent was then added and the mixture was degassed with stirring under vacuum at room temperature for 30 min. The mixing ratio of the epoxy and curing agent was 100 : 26.4 as recommended by the manufacturer. To fabricate nanocomposite samples, this mixture was poured into a preheated Teflon mould and cured in an oven at 120°C for 2 hours followed by 2 hours post curing at 177°C. Neat epoxy samples were prepared using the same procedure.

Measurements and characterisations. Fourier transform infrared (FTIR) spectra were recorded by a FTIR spectrophotometer (Bruker Optics) using the KBr method. Scanning electron microscopy (SEM) images were obtained by a Zeiss supra 55VP4 scanning electron microscope. Thermogravimetric analysis (TGA) was measured under a nitrogen atmosphere using a Perkin Elmer at a heating rate of 5°C/min. Transmission electron microscopy (TEM) images were obtained by a JEOL 2100 microscope. Wide angle X-ray diffractometry (XRD) data were collected using a Panalytical XRD instrument with Cu radiation 1.5406 Å. Raman measurements were performed at room temperature at 514 nm. Three-point bending flexural tests were conducted according to ASTM D790 using a Lloyd tensile tester. Dimensions of specimen were 70 mm (length) × 10 mm (width) × 3.3 mm (thickness). The tests were performed with a 10 kN load cell at a cross-head speed of 2 mm/min. At least five specimens from each sample were tested. Dynamic mechanic analysis (DMA) was performed on a TA Instruments Q800 in the cantilever bending mode at a frequency of 1 Hz. Typical dimension for DMA sample were 56 mm (length) × 13 mm (width) × 3 mm (thickness). Measurements were carried out from 30°C to 180°C at a heating rate of 2°C/min and a strain of 0.05%.

Results and Discussion

The morphologies of starting materials as well as thermally reduced graphene (TRG) and functionalised graphene (FG) were characterised using TEM and SEM as shown in Figure 2. A representative TEM image of TRG is shown in Figure 2a. TRG exhibits transparent

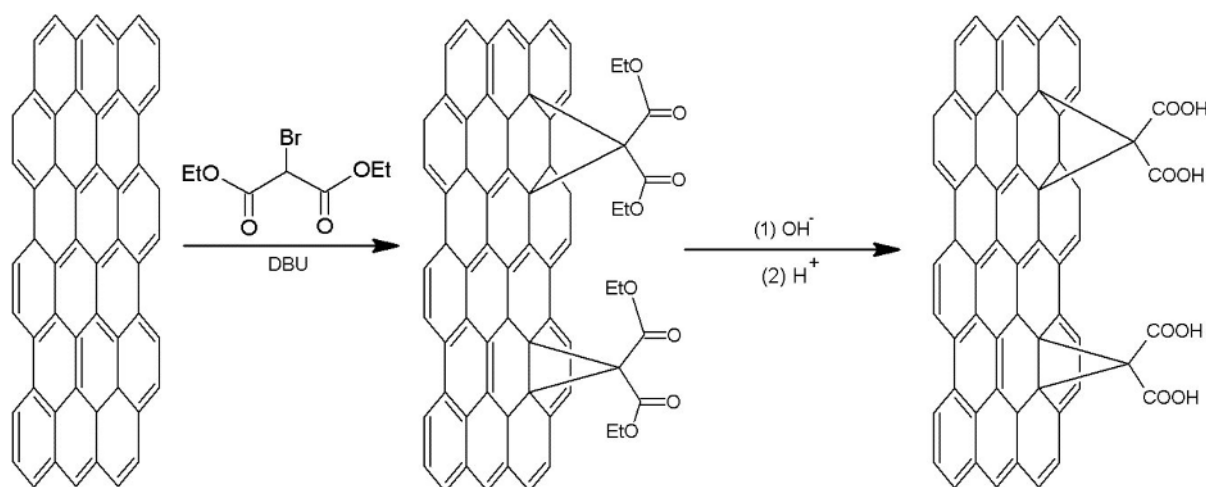


Figure 1 | Synthesis of functionalised graphene via bingel reaction.

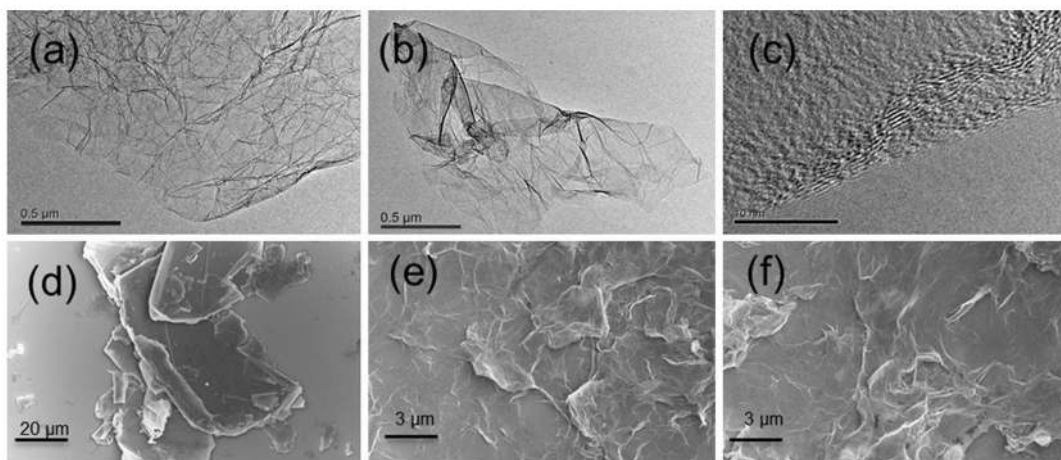


Figure 2 | Micrographs of TEM images of (a & c) TRG and (b) FG; SEM images of (d) natural flake graphite, (e) TRG and (f) FG.

and wrinkled sheet structure which is due to the epoxy groups forming chains across the graphene surface³². No obvious morphological changes were observed for FG (Figure 2b) indicating that chemical functionalization process did not result in restacking and agglomeration of graphene layers. High resolution TEM (Figure 2c) demonstrates that the TRG composed of ~ 4 – 5 individual graphene sheets within each platelet. Figure 2d presents the SEM image of natural flake graphite. As shown graphite flakes demonstrate a layered structure and graphite sheets are stacked densely. As shown in Figure 2e graphene oxide has been successfully reduced to graphene nanoplatelets and presents a typical rippled structure that has been preserved through chemical functionalization (Figure 2f).

Figure 3 shows XRD patterns of the graphite, GO, TRG and FG. The graphite exhibits a sharp diffraction peak centered at $2\theta = 26.5^\circ$ corresponding to the (0 0 2) graphite plane composed of well-ordered graphenes with an interlayer spacing of 3.35 Å. This peak disappears for GO and a relatively low peak at $2\theta = 10.5^\circ$ appears, corresponding to the diffraction of the (0 0 2) graphite oxide plane. The interlayer spacing of GO is calculated according to the Bragg's law and the calculated value is 8.41 Å. This implies that as graphite transforms to GO most oxygen atoms are bonded to the graphite

surface and that graphite is expanded when oxidized^{19,33}. After high temperature heat treatment, (002) peak disappears indicating that the thermally reduced graphene sheets are disordered. Removal of oxidised functional groups results in exfoliation to graphene nanosheets, which was further confirmed by TEM. The randomly disordered graphene XRD patterns remains unchanged after functionalization indicating that chemical functionalization process did not give rise to restacking of graphene layers.

In the FT-IR spectrum of the graphene oxide sample (Figure 4), the peaks at ~ 3432 and 1711 cm^{-1} are attributed to $-\text{OH}$ and $\text{C}=\text{O}$ bands, respectively. Upon reduction of graphene oxide to graphene, the $\text{C}=\text{O}$ band disappears and new bands at 2928 and 2865 cm^{-1} arise representing the C-H stretch vibrations of the methylene group. Functionalised graphene displays a peak at 1731 cm^{-1} characteristic band for $\text{C}=\text{O}$ stretch of the COOH group. The presence of carboxylic functional group is further confirmed by the strong and broad band at 3412 cm^{-1} .

Raman spectroscopy is employed to study the transformation of graphite to graphene and functionalised graphene. Raman spectra of graphite, graphene and functionalised graphene are shown in Figure 5. In the graphite spectra there is a Raman active bond which

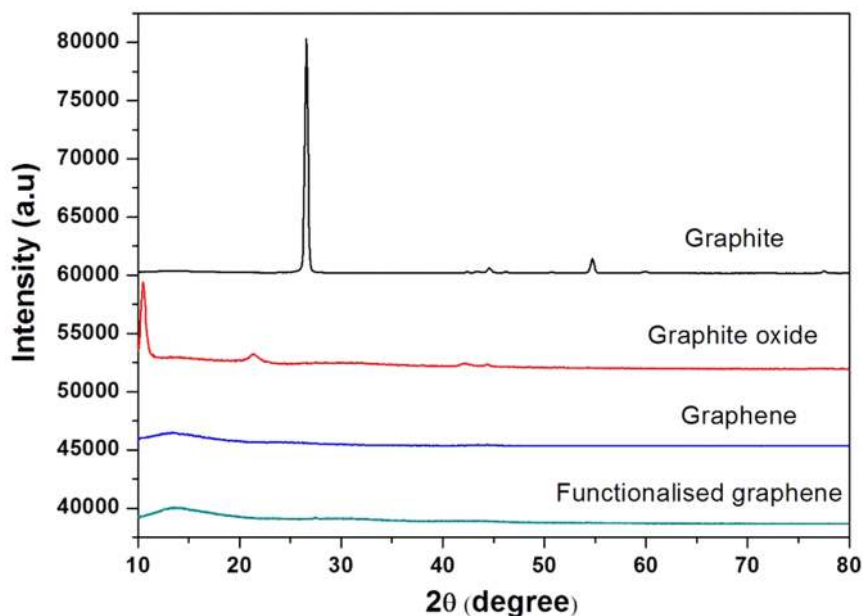


Figure 3 | XRD plots for graphite, GO, TRG and FG.

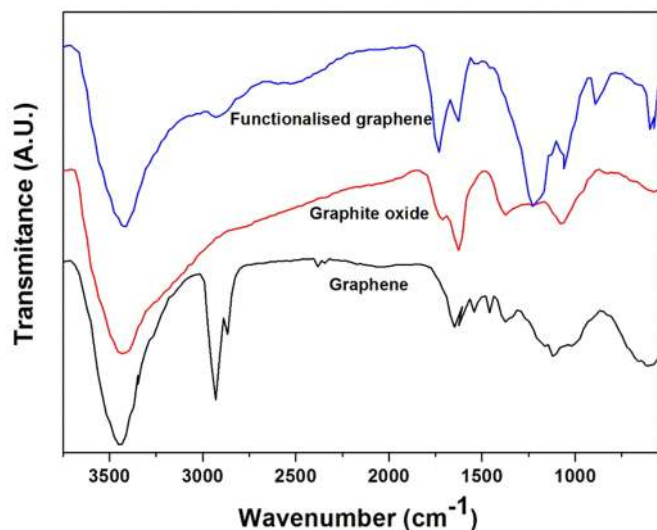


Figure 4 | FTIR spectra of graphene, thermally reduced graphene oxide and functionalised graphene.

is the vibrational G band at 1578 cm^{-1} representing in-plane bond stretching vibration of sp^2 bonded carbon atoms³⁴. There are significant structure changes in thermally reduced graphene leading to appearance of D band at 1342 cm^{-1} indicating the presence of defects due to oxidation and thermal reduction of graphite. The G band in thermally reduced graphene slightly shifts to lower frequencies and shows higher intensity and broadened, and the D band appears which demonstrates an increased D/G intensity ratio of 0.70 compared to 0.05 in graphite. In functionalised graphene D band becomes broad and there is a further increase in D/G intensity ratio, from 0.05 of graphite to 0.90 which confirms the successful attachment of functional groups on graphene nanosheets. Also, the G band shifts back to the position of the G band in graphite that could be attributed to the graphitic self-healing as previously reported in literature³⁵.

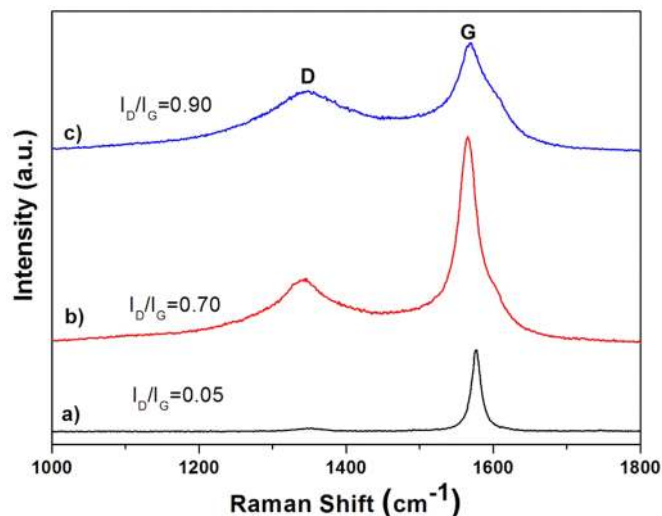


Figure 6 | TGA plots of graphite, GO, TRG and FG under nitrogen atmosphere and heating rate of $10^\circ\text{C}/\text{min}$.

TGA is a valuable and fairly simple method in investigating the thermal stability as well as further analysing the presence of functional group in carbonaceous materials. As seen in Figure 6, graphite is a thermally stable material with no weight loss over the entire testing temperature range. GO, however, has a slight initial mass loss at $\sim 100^\circ\text{C}$ which could be attributed to the evaporation of water molecules^{36,37}. The main weight loss ($\sim 40\%$) for GO takes place at $\sim 220^\circ\text{C}$ and is attributed to the pyrolysis of the reactive oxygen containing functional groups, yielding CO_2 , CO and H_2O vapors. This is in agreement with findings of other researchers^{17,38}. Functionalised graphene on the other hand has a significantly higher thermal stability than GO with the main weight loss occurring at about 500°C . Higher onset degradation temperature for functionalised graphene compared to GO could be due to the presence of more stable oxygen containing functional groups. TRG is thermally stable

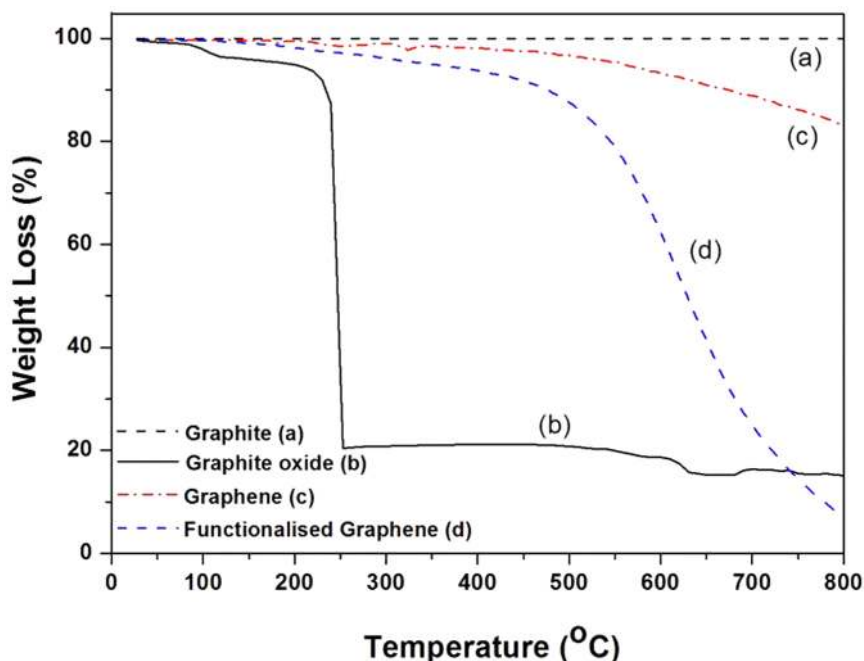


Figure 5 | Raman spectra of (a) graphite, (b) thermally reduced graphene and (c) functionalised graphene.

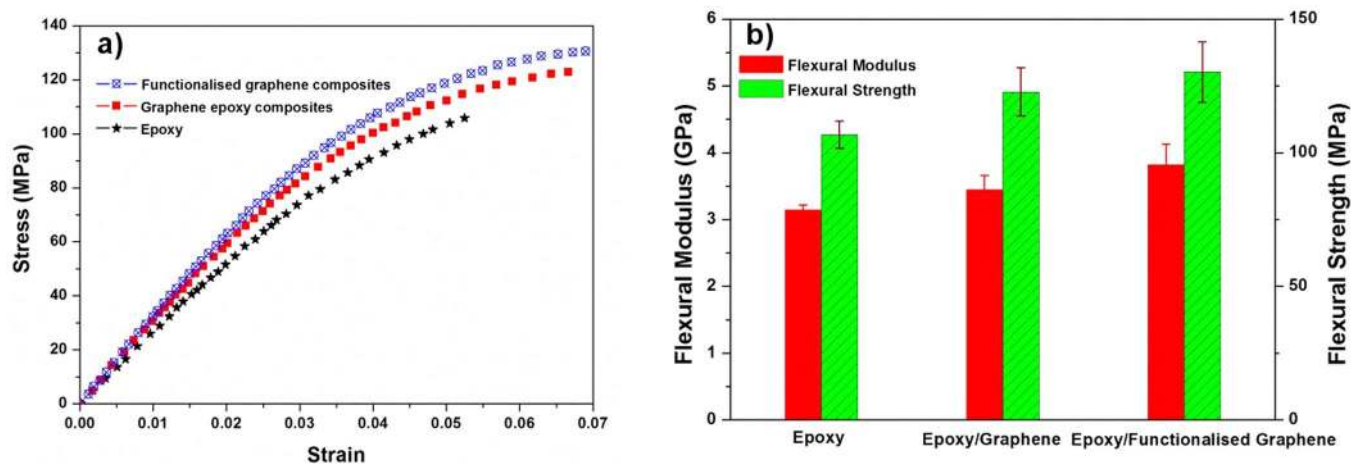


Figure 7 | (a) Stress–strain curves; (b) flexural modulus and strength of the composite samples.

up to $\sim 600^{\circ}\text{C}$, with the overall weight loss of $\sim 16\%$ at 800°C . It's been suggested in previous studies that thermal reduction of graphite oxide restores the graphitic structure resulting in a more thermally stable material³⁹. Due to excellent thermal properties, graphene is used for thermal management of various materials⁴⁰.

The three point bending or flexural test is a useful and commonly used test due to the mixture between tensile and compressive forces that are likely to be encountered in the real world application of composites. This test was done to investigate how the addition of FG influences the mechanical properties of epoxy matrix. Samples with 0.1 wt% TRG and FG as well as neat epoxy were prepared and tested for flexural properties. Figure 7 shows strain-stress plots for neat epoxy and composites containing TRG and FG. The addition of TRG and FG shows the increase in flexural strength of epoxy matrix by $\sim 15\%$ and 22% , respectively. As shown in TEM images, TRG exhibit a wrinkled topology due to extremely small thickness. Nanoscale surface roughness and wrinkled nature of TRG enhances mechanical interlocking leading to better adhesion⁴¹. The presence of carboxylic functional group on FG which may form covalent bonding with epoxy matrix further improves the interfacial bonding with epoxy leading to further improvement in mechanical properties.

In order to obtain more information regarding the interfacial adhesion between FG and epoxy matrix, fractured surface of epoxy samples containing TRG and FG was further studied using SEM imaging and is shown in Figure 8. As can be seen from Figure 8a, the epoxy composite containing TRG exhibits a relatively smooth fracture surface with river-like patterns in the direction of fracturing force, a brittle failure as a result of weak bonding between TRG and epoxy. In examining SEM images (not shown here) we also found that the fracture surface of epoxy/TRG is not uniform meaning that in some regions the surface is mirror-like. This can suggest irregular distribution and dispersion of TRG within epoxy matrix. The incorporation of FG into epoxy matrix exhibits different feature and fracture surface morphology (Figure 8b). The fracture surface of FG/epoxy composites is rough with many small faceted features. This further suggests that the strong interfacial adhesion between functionalised graphene and epoxy matrix discourages fracture, leading to a rough surface. As shown in red circles in Figure 8c, functionalised graphene platelets are embedded in epoxy matrix. No obvious pull out can be observed and the epoxy matrix appears to sufficiently wet the graphene surface (Figure 8d) which suggests that FG present strong filler-matrix interfacial bonding and encourages dispersion

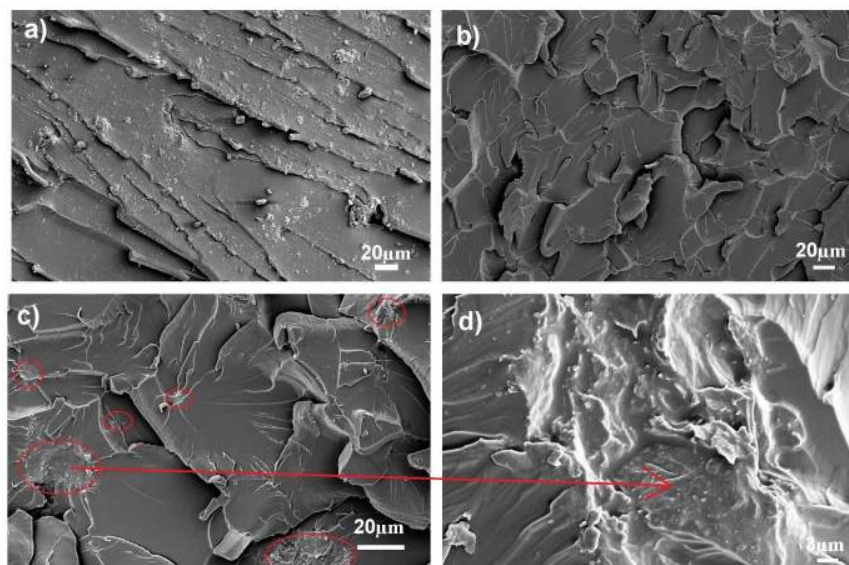


Figure 8 | Scanning electron microscopy (SEM) images of the fractured surface of (a) epoxy/TRG and (b–c) epoxy/FG. Arrow in Figure 8d shows the magnified image of circled area in Figure 8c.

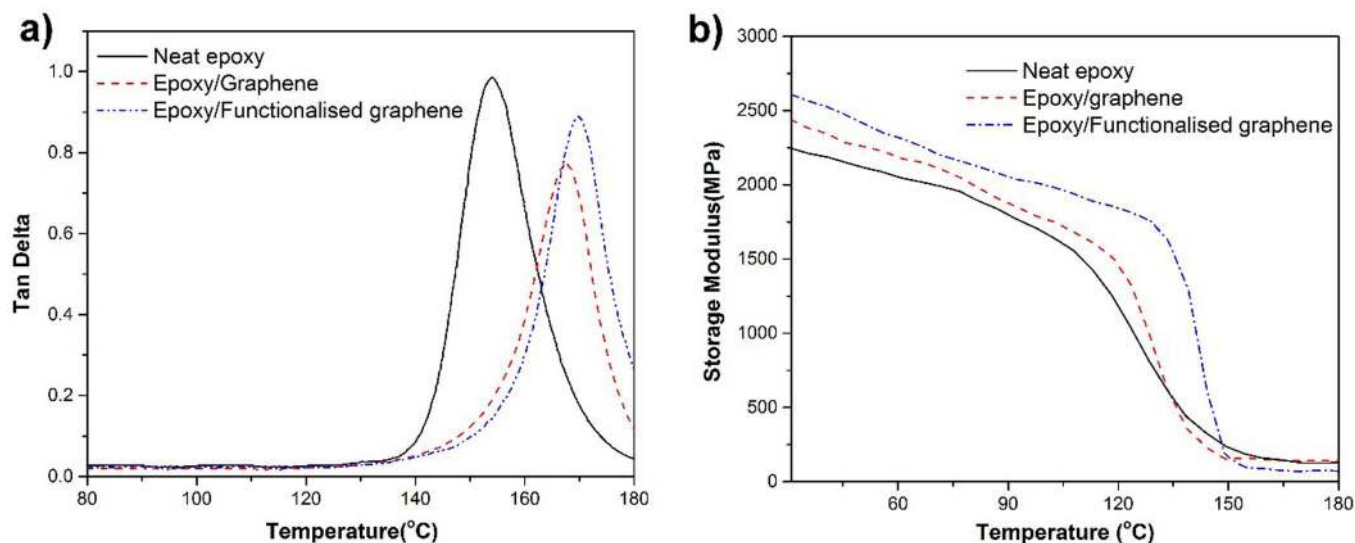


Figure 9 | (a) Tan delta and (b) Storage modulus curves for neat epoxy and composite samples containing TRG and FG.

as a result of chemical functional group present on the surface of graphene. Strong interface and improved dispersion have a significant influence on the mechanical properties of composites⁴².

Figure 9a shows the tan delta plot for neat epoxy and epoxy composites containing TRG and FG. The peak of tan delta curve is considered as glass transition temperature (T_g) of polymers. Neat epoxy resin shows a T_g of 153.9°C. With the addition of TRG and FG, T_g increases to 167.6 and 169.8°C, respectively. The rise in T_g upon inclusion of nanofillers is often associated with a restriction in molecular motion and higher degree of crosslinking indicating significant changes in polymer chains dynamics^{43,44}. In addition to this, molecular dynamic studies have confirmed that surface geometry of nanofillers could influence polymer mobility⁴⁵. Therefore, it's plausible that the nanoscale surface roughness and wrinkled nature of graphene surface induce interfacial interactions through mechanical interlocking with polymer chains and hinder molecular mobility leading to increase in T_g upon addition of TRG and FG. On the other hand, the presence of carboxylic functional group on the surface of graphene enhances uniform dispersion within resin matrix which in turn leads to creating an interphase zone surrounding FG nanoplatelets. This increased surface area in contact with resin together with formation of covalent bonds between FG and epoxy matrix, constrains the resin matrix chain mobility even further and results in a larger T_g shift. Higher T_g in epoxy/FG composites reflects the more effective chain restriction ability of FG. The higher peak height of tan delta for epoxy/FG composites compared to epoxy/TRG could also indirectly indicates a uniform dispersion of the FG in the epoxy resin⁴⁶.

The storage modulus plots for neat epoxy matrix and composites containing TRG and FG are presented in Figure 9b. The storage modulus in composite containing nanofillers is highly influenced by the interfacial bonding strength between reinforcement filler and resin matrix⁴⁷. Inclusion of TRG and FG into neat epoxy demonstrates an increase in storage modulus in the glassy region. The storage modulus of epoxy at 30°C is 2.25 GPa and that of nanocomposite containing TRG and FG is 2.44 and 2.61 GPa, respectively. This indicates about 10% and 18% increase, respectively, compared to the neat epoxy. The two dimensional sheet geometry of TRG and its high aspect ratio enables effective load transfer. FG forms a strong interface interaction with the epoxy matrix as a result of functional groups present in its chemical structure. Covalent bonding between carboxylic group of FG and epoxy could improve the load transfer and consequently enhance the reinforcing effect. As the temperature increases, the storage modulus falls, indicating energy dissipation

which occurs during the transition of the glassy state to a rubber state. Investigating the storage modulus in rubbery region i.e higher temperature, it is noted that the storage modulus for the composite samples containing FG is the highest of the three. At 135°C the storage modulus of the FG/epoxy composite is about 2.7 times higher than both the neat epoxy and TRG reinforced samples.

Conclusions

In this study, we reported covalent functionalization of thermally reduced graphene through Bingel reaction. The reaction conditions provided a platform for functional groups to covalently attach onto the graphene skeleton. Bingel modified graphene showed a randomly disordered structure with enhanced thermal stability. Functionalised graphene were homogeneously dispersed in epoxy resin to produce nanocomposites. Nanocomposite containing a small functionalised graphene loading demonstrated improved flexural strength and modulus. The incorporation of functionalised graphene enhanced the glass temperature of epoxy resin. This shift along with improved mechanical properties suggests stronger interfacial interactions between graphene and epoxy as a result of chemical modification of thermally reduced graphene. Further investigations into the properties of epoxy composites containing various concentrations of Bingel modified graphene is currently underway and will be reported in future.

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Author contributions

Conceived and designed the experiments: M.N. Performed the experiments: M.N., J.W., A.A., H.K., N.H. and L.L. Data analysis: M.N. and J.W. Contributed reagents/materials/analysis tools: M.N., Y.C. and B.F. Wrote the manuscript: M.N. All authors contributed to the technical review of the manuscript.

Additional information

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