
This is an electronic reprint of the original article.
This reprint may differ from the original in pagination and typographic detail.

Tavakkoli, Mohammad; Flahaut, Emmanuel; Peljo, Pekka; Sainio, Jani; Davodi, Fatemeh; Lobiak, Egor V.; Mustonen, Kimmo; Kauppinen, Esko I.

Mesoporous Single-Atom-Doped Graphene-Carbon Nanotube Hybrid

Published in:
ACS Catalysis

DOI:
[10.1021/acscatal.0c00352](https://doi.org/10.1021/acscatal.0c00352)

Published: 17/04/2020

Document Version
Publisher's PDF, also known as Version of record

Published under the following license:
CC BY

Please cite the original version:
Tavakkoli, M., Flahaut, E., Peljo, P., Sainio, J., Davodi, F., Lobiak, E. V., Mustonen, K., & Kauppinen, E. I. (2020). Mesoporous Single-Atom-Doped Graphene-Carbon Nanotube Hybrid: Synthesis and Tunable Electrocatalytic Activity for Oxygen Evolution and Reduction Reactions. *ACS Catalysis*, 10(8), 4647-4658. <https://doi.org/10.1021/acscatal.0c00352>

This material is protected by copyright and other intellectual property rights, and duplication or sale of all or part of any of the repository collections is not permitted, except that material may be duplicated by you for your research use or educational purposes in electronic or print form. You must obtain permission for any other use. Electronic or print copies may not be offered, whether for sale or otherwise to anyone who is not an authorised user.

Mesoporous Single-Atom-Doped Graphene–Carbon Nanotube Hybrid: Synthesis and Tunable Electrocatalytic Activity for Oxygen Evolution and Reduction Reactions

Mohammad Tavakkoli,* Emmanuel Flahaut, Pekka Peljo, Jani Sainio, Fatemeh Davodi, Egor V. Lobiak, Kimmo Mustonen, and Esko I. Kauppinen



Cite This: *ACS Catal.* 2020, 10, 4647–4658



Read Online

ACCESS |



Metrics & More



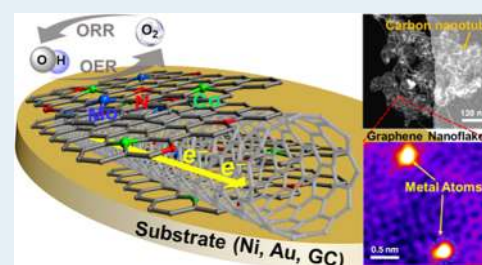
Article Recommendations



Supporting Information

ABSTRACT: Mesoporous heteroatom-doped carbon-based nanomaterials are very promising as catalysts for electrochemical energy conversion and storage. We have developed a one-step catalytic chemical vapor deposition method to grow a highly graphitized graphene nanoflake (GF)–carbon nanotube (CNT) hybrid material doped simultaneously with single atoms of N, Co, and Mo (N–Co–Mo–GF/CNT). This high-surface-area material has a mesoporous structure, which facilitates oxygen mass transfer within the catalyst film, and exhibits a high electrocatalytic activity and stability in oxygen reduction and evolution reactions (ORR and OER) in alkaline media. We have shown that in this metal (M)–N–C catalyst, M (Co, Mo)–C centers are the main sites responsible for OER, while, for ORR, both M and N–C centers synergistically serve as the active sites. We systematically investigated tuning of the ORR and OER activity of the porous catalyst depending on the choice of the underlying substrate. The ORR kinetic current and OER activity for N–Co–Mo–GF/CNT were significantly enhanced when the catalyst was deposited onto a Ni substrate, resulting in an advanced electrocatalytic performance compared to the best bifunctional ORR/OER catalysts reported so far. Using a developed scanning electrochemical microscopy analysis method, we demonstrated that the higher OER reactivity on Ni was attributable to the formation of underlying catalyst/Ni interfacial sites, which are accessible due to the porous, electrolyte-permeable structure of the catalyst.

KEYWORDS: single-atom electrocatalysis, heteroatom doping, graphene–carbon nanotube hybrid, mesoporous electrocatalyst, oxygen reduction/evolution reaction, substrate effect, scanning electrochemical microscopy



1. INTRODUCTION

The need for low-cost electrocatalysts based on earth-abundant materials, rather than critical noble metals, such as Pt, Ru, and Ir, for catalyzing oxygen reduction reaction (ORR) and oxygen evolution reaction (OER) has motivated a large body of research. ORR and OER are the most important electrochemical reactions that limit the efficiencies of fuel cells, water electrolyzers, and metal–air batteries. An ORR or OER electrode comprises three main components: (i) an electrocatalyst film composed of active sites for catalysis, (ii) a catalyst support, which is typically made of high-surface-area and porous carbon-based materials that can enhance the number of active sites on the surface and improve the conductivity of the catalyst, and (iii) a conductive substrate on which the electrocatalyst (or catalyst/support) is deposited. All three components require a rational design for the production of advanced electrodes for electrochemical energy storage and conversion devices.

Among various nonprecious electrocatalysts, nitrogen (N)-doped carbon materials (N–C) have recently evolved into metal-free ORR and OER catalysts.^{1,2} The catalytic activity of

N–C electrocatalysts for ORR^{3–6} and OER⁷ can be further improved by codoping with transition metals (M–N–C catalysts). M–N–C catalysts are generally synthesized by pyrolysis of a mixture of nitrogen-containing chemicals, transition metal inorganic salts, and carbon-based supporting materials.⁸ Carbon-supporting materials for fuel cell electrodes should possess high porosity, high conductivity, a large surface area, and high electrochemical stability.^{4,9,10} Porous materials facilitate electrolytic mass transport through the electrode and increase accessibility to the active sites.^{2,3} For oxygen mass transfer, mesopores (pore size 2–50 nm) are superior to micropores (pore size <2 nm).^{11,12} Mesoporous structures enhance the portion of electrochemically available active sites,³ and they have been reported to be promising porous

Received: January 21, 2020

Revised: March 9, 2020

Published: March 20, 2020



electrodes,^{12–14} which can activate extra active sites within the pores and enhance the electrocatalytic activity for ORR. Recently, N-doped mesoporous carbons containing transition metals have been reported to be promising bifunctional electrocatalysts for ORR/OER¹⁵ and water electrolysis.¹⁶ Carbon black has been widely utilized as the catalyst support due to its high availability and low cost; it nevertheless contains a high number of micropores, which hinder oxygen transport in the electrocatalyst and reduce the accessibility of active sites.^{4,10} Carbon black also contains an abundance of dangling bonds and defects, which can easily form surface oxides, resulting in corrosion under electrochemical oxidation.¹⁷ In contrast, catalyst supports that consist of a highly graphitized carbon structure, such as carbon nanotubes (CNTs) and graphene, are reported to be more stable.^{10,17,18} CNTs and graphene have high conductivity and can be doped by heteroatoms, which differ from carbon in their electronegativity. The heteroatom doping of graphitic carbon nanomaterials, regardless of whether the dopant has a lower (as B, P, S, and transition metals) or higher (as N) electronegativity than that of carbon, can polarize adjacent carbon atoms, change the electronic properties of carbon skeletons,^{2,19} and improve the catalytic activity of carbon-based electrocatalysts for various electrochemical reactions, including ORR^{20–22} and OER.^{23,24} Therefore, designing high-surface-area mesoporous M–N–C electrocatalysts using CNT and graphene supports is a promising route to producing highly active and robust materials for catalyzing ORR and OER.

Graphene–CNT hybrid nanomaterials have demonstrated promising performance in various applications.^{25–31} The simultaneous growth of graphene and CNT has, however, rarely been reported,^{26,27,32–34} but typically resulted in graphitic structures with a relatively high number of defects. In graphene–CNT hybrid films, the CNTs form a conductive framework to which the graphene flakes adhere, significantly improving overall conductivity.^{26,32} Due to an increase in tunneling efficiency, the conductivity of a graphene–CNT hybrid can surpass that of a CNT film.³⁵ Here, we report a synthesis method that produces a high-quality graphene–CNT hybrid in which the graphene flakes are doped with N, Co, and Mo heteroatoms for enhanced electrocatalytic performance.

Substrates over which electrocatalysts are deposited play an important role in defining electrocatalytic activity. The substrate can modify the surface morphology, porosity, electronic structure, and conductivity of electrocatalysts.^{36,37} For ORR, the role of the substrate has not been well studied in alkaline media. However, experimental investigations have shown that the OER electrocatalytic activities of manganese,^{38,39} cobalt,⁴⁰ and Ni-based⁴¹ oxides, as well as Fe (oxy)hydroxide,⁴² NiCeO_x,⁴³ and amorphous Co(OH)₂⁴⁴ in alkaline media are dramatically improved using an AuO_x/Au substrate instead of commonly used glassy carbon (GC). The Au substrate can potentially diffuse into the electrocatalyst film.⁴³ An electrocatalyst possessing a porous, electrolyte-permeable structure furthermore provides access to underlying active sites,^{42,45} allowing the formation of metal oxide–gold interfacial sites that catalyze the OER at lower overpotentials than metal oxide sites.⁴³ Ni foam has also been widely used in recent years for the enhancement of the OER activity of various electrocatalysts.⁴⁶ However, Ni foam offers a large active surface area and a highly continuous porous three-dimensional (3D) network enhancing accessibility to the active sites and thus electrocatalytic activity. This makes a systematic

comparison with other substrates difficult, as it is unclear how much activity is enhanced by the formation of catalyst/substrate interfacial active sites, relative to the influence of the structure and morphology of the substrate.

Here, we introduce a facile and scalable one-step synthesis method for the production of a novel electrocatalyst composed of few-layer graphene nanoflakes (GFs) and CNTs doped with N, Co, and Mo heteroatoms (N–Co–Mo–GF/CNT). In catalysis science, the new field of single-atom catalysts (SACs) with isolated metal atoms dispersed on solid supports has attracted wide research attention because of the maximum atom-utilization efficiency and unique properties of SACs.⁴⁷ Among SACs, those with carbon-based⁴⁸ and particularly with graphene or graphene-like⁴⁹ supports are widely investigated catalysts because of the extraordinary physicochemical properties of such supports. However, owing to the high surface energy of single atoms, immobilization of atomic metal centers on the support for the fabrication of SACs has been challenging.⁴⁷ Here, in contrast to other synthetic strategies for SACs,^{47–49} the SAC is produced during the synthesis of the carbon support (GF/CNT), providing a facile one-step synthesis process for the fabrication of high-performance SACs without any extra cost arising from the immobilization process of the single-atom metals on the support. This high-surface-area mesoporous catalyst shows high ORR and OER activities in alkaline media. Furthermore, we have systematically investigated the role of Ni, Au, and GC flat disk substrates on the ORR and OER catalytic activities of a high-performance mesoporous catalyst (N–Co–Mo–GF/CNT) in comparison to well-established ORR and OER catalysts such as Pt/C and RuO₂, respectively. Our study of substrate effects on the OER and ORR catalytic activity of porous materials establishes a basis for the rational design of electrodes with optimized activity and provides guidelines for future studies.

2. RESULTS AND DISCUSSION

2.1. Synthesis of Catalyst Materials. N–Co–Mo–GF/CNT was synthesized by modifying a catalytic chemical vapor deposition (CCVD) method⁵⁰ that was initially developed for the scalable synthesis of CNTs. The CCVD synthesis of the CNTs is a chemical vapor deposition (CVD) process in which the carbon precursor molecules are catalytically decomposed at high temperatures on the surface of a metallic support that acts as the catalyst for the growth of CNTs. The mesoporous N–Co–Mo–GF/CNT catalyst was prepared using a CCVD synthesis reactor as depicted in Figure 1. The synthesis process has been optimized to grow a GF/CNT hybrid via a one-step synthesis process. Briefly, Mg_{0.99}(Co_{1–x}Mo_x)_{0.01}O ($x \sim 0.25$) was used as the catalyst to synthesize graphitic carbon nanomaterials. This catalyst was developed previously for the synthesis of double-walled CNTs.^{50,51} Using this catalyst, the growth of crumpled graphene structures has been also observed during the CCVD synthesis of the CNTs,⁵¹ probably by the direct catalytic decomposition of CH₄ on MgO oxide.^{51,52} In this work, the growth conditions were optimized to grow a graphene–CNT hybrid material rather than CNTs or graphene layers. The catalyst was prepared by combustion synthesis of a stoichiometric mixture of (NH₄)₆Mo₇O₂₄·4H₂O, magnesium and cobalt nitrates, and citric acid fuel as explained in refs 50 and 53. The catalyst powder was then placed in a furnace under a flow of 205 sccm H₂ and 45 sccm CH₄. The furnace temperature was increased from ambient temperature to 1000 °C, kept there for 6 min, and then slowly decreased to

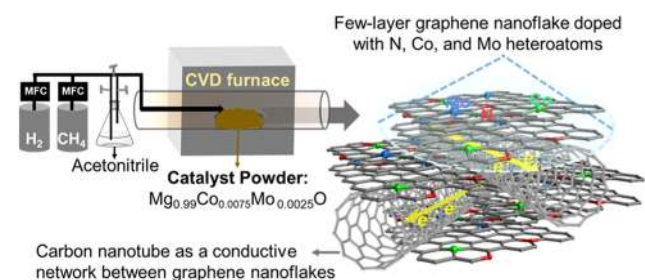


Figure 1. Scheme of the synthesis reactor used for the growth of the N–Co–Mo–GC/CNT and Co–Mo–GC/CNT materials (MFC is an abbreviation for mass flow controller). The schematic illustration of the N–Co–Mo–GC/CNT sample is based on high-resolution scanning transmission electron microscopy (STEM) images and electron energy loss spectroscopy (EELS) analysis.

room temperature. The heating and cooling were performed at a rate of $5\text{ }^{\circ}\text{C min}^{-1}$. For the nitrogen-doped material, acetonitrile was introduced into the reactor by bubbling the feedstock of hydrogen and methane (total flow rate = 250 sccm) through an impinger containing anhydrous acetonitrile for 6 min with the furnace temperature at $1000\text{ }^{\circ}\text{C}$ so that $\sim 1\text{ mL}$ acetonitrile was utilized in this process. The time during which the N source was introduced into the synthesis reactor was extremely short in comparison to the overall growth time of the material. Hence, the sample synthesized without N heteroatoms (denoted the Co–Mo–GF/CNT sample) would likely have a similar structure to that synthesized with N dopants, except for the N heteroatoms in part of the graphitic carbon network.

After cooling to room temperature, the sample was processed with an aqueous solution of HCl to dissolve the remaining catalyst powder and unprotected catalytic nanoparticles. The sample was washed with deionized water until a neutral pH was obtained. After washing, the wet sample was frozen and then dried by lyophilization. During this process, only those metallic atoms/nanoparticles that were firmly embedded in the graphitic carbon network were protected against dissolution in the acid. After this step, inductively coupled plasma atomic emission spectroscopy (ICP-AES) revealed 4.96 and 1.65 wt % for Co and Mo, respectively, in

the N–Co–Mo–GF/CNT hybrid material (see details of the elemental analysis in the [Supporting Information](#)).

The resulting material was dispersed in ethanol (5 mg mL^{-1}) using ultrasound sonication for 1 h, followed by magnetic stirring for 3 days, so that a homogeneous ink was formed. This ink was used for the material characterizations and the electrochemical measurements.

2.2. Characterization of the Synthesized Material.

2.2.1. Atomic Resolution Electron Microscopy Characterization.

High-resolution transmission electron microscopy (HR-TEM) images of the N–Co–Mo–GF/CNT hybrid material are shown in [Figure 2a,b](#), where the presence of 50–70 nm GFs is visible. Scanning TEM (STEM) was also employed ([Figure 2c–h](#)), in which an angstrom-sized 60 keV electron probe was raster-scanned over the sample, and images were recorded with high- or medium-angle annular dark-field (HAADF and MAADF) detectors. Electron energy loss spectroscopy (EELS) with single-atom precision was also performed. The signals from the detectors and EELS can be obtained simultaneously, allowing direct correlation of images and spectroscopic data. In MAADF and especially HAADF imaging, the contrast of the image is highly sensitive to the atomic number or the mass of nuclei (Z -contrast images). Z -contrast allows individual heavy atoms and small metallic nanoparticles to be distinguished from the low- Z carbon support based on their brightness.^{54–58} Here, both annular detectors were used, and elemental identification at the atomic resolution was conducted by EELS as has been previously done for heteroatom-doped graphene and CNTs.^{54,55,59,60} [Figure 2c,d](#) demonstrates how the CNTs bridge the graphene nanoflakes, leading to the expected improvement in the material conductivity. Metallic Co and Mo atoms were not detected in the monolayer graphene ([Figure 2e,f](#)), whereas N was identified directly from the STEM image, for which the corresponding EELS point spectrum is shown in [Figure S1](#). [Figure 1f](#) demonstrates the presence of pentagon and heptagon rings in the hexagonal crystalline structure leading to strain-related corrugation in the structure of GFs. In contrast to monolayer GFs, individual metal atoms were identified in multilayers, as is evident in [Figures S2](#) and [2g,h](#). [Figure S2](#) shows example STEM images of Co and Mo heteroatoms with corresponding EELS maps. In a binary collision with a 60 keV

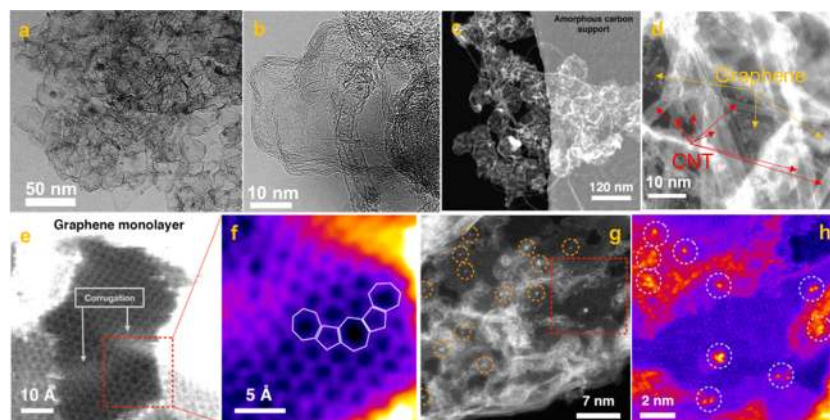


Figure 2. (a, b) HR-TEM and (c–h) STEM images of N–Co–Mo–GF/CNT. (a) TEM image of 30–50 nm GFs and (b) few-layer GF, (c, d) STEM image of the GFs interconnected by CNTs, with (d) showing CNTs bridging the GFs, (e, f) high-resolution STEM images taken from a graphene monolayer, showing the honeycomb structure, and (g, h) STEM images of double- or few-layer GFs. Individual metal atoms are observed as bright spots on the carbon support.

electron, >1 eV in-plane kinetic energy can be transferred to a heavy nucleus such as that of Co,⁶¹ which is enough to activate adatom diffusion in few-layer graphene encapsulation. The fact that the metal atoms remained completely stable during our STEM experiments serves as a strong indication of their covalent nature. STEM images taken from the CNTs in the N-Co-Mo-GF/CNT material are shown in Figure S3. No metallic dopants were observed in the sidewall of CNTs, hinting that the CNTs are mainly a conductive network bridging the GFs for efficient electron transport, rather than providing additional active M-C sites for electrocatalysis. Figure S4 shows single-, double-, and few-layer graphene nanoflakes in the synthesized material. In some parts of the sample, Co and Mo nanoparticles were also detected. The frequency at which the particles were observed was nevertheless much lower than that of the individual metallic atoms, as is also visible in Figure S5.

2.2.2. Raman Spectroscopy. Raman spectroscopy is a powerful method that allows the detection of graphitization in carbon nanomaterials. Generally, the Raman spectra of sp²-hybridized carbon exhibit major spectral components emerging from defects (D-band at 1300–1360 cm⁻¹) and in-plane G-band vibrations at 1570–1600 cm⁻¹.^{62–64} Earlier studies have shown that the disorder-induced D-band is increased when the graphitic carbon lattice is doped with foreign atoms.⁶⁴ Figure 3

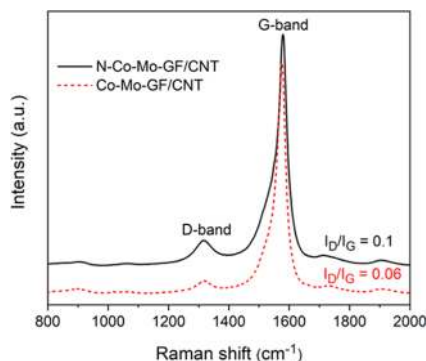


Figure 3. Raman spectra obtained from N-Co-Mo-GF/CNT (black line) and Co-Mo-GF/CNT (red dashed line) samples. The spectra show a low I_D/I_G ratio for the one-step synthesized graphene/CNT hybrids, reflecting the high level of graphitization of the samples. The data are normalized with respect to G-band intensities.

shows the Raman D- and G-bands of Co-Mo-GF/CNT with and without N heteroatom doping. The highly crystalline sp² carbon structure of the Co-Mo-GF/CNT sample is reflected in its relatively low I_D/I_G ratio of 0.06. Nitrogen doping increased the I_D/I_G ratio to 0.1, which can be attributed to the substitution of N heteroatoms in the graphitic carbon. The observed ratio is significantly lower than that reported for other synthesized graphene-CNT hybrids in the literature.^{25–27,31–33} This demonstrates that the synthesis method reported here allows the production of conductive heteroatom-doped graphene-CNT structures with higher quality than has previously been possible.

2.2.3. Pore-Size Analysis. Figure S6 shows the nitrogen adsorption-desorption isotherm of N-Co-Mo-GF/CNT, which exhibits a typical IV-type isotherm curve with a distinct hysteric loop associated with capillary condensation taking place in the mesopores.^{65,66} The shape of the hysteric loop can be attributed to slit-shaped pores, which are observed for

nonrigid aggregates of platelike particles.⁶⁶ Figure S6 shows the cumulative pore volume and pore-size distributions of the N-Co-Mo-GF/CNT calculated using the Barrett-Joyner-Halenda (BJH) method, revealing an average pore width of 11.5 nm. Moreover, a high specific surface area of 911 m² g⁻¹ was measured using the Brunauer-Emmett-Teller (BET) method.

2.3. Oxygen Reduction Activity and Discussion. To evaluate the activity of the N-Co-Mo-GF/CNT catalyst for ORR on various Ni, Au, and GC substrates, rotating disc electrode (RDE) voltammetry measurements were carried out at different rotation rates in 0.1 M KOH (see details of the electrochemical procedures in the Supporting Information). Figure 4a demonstrates the ORR linear sweep voltammetry

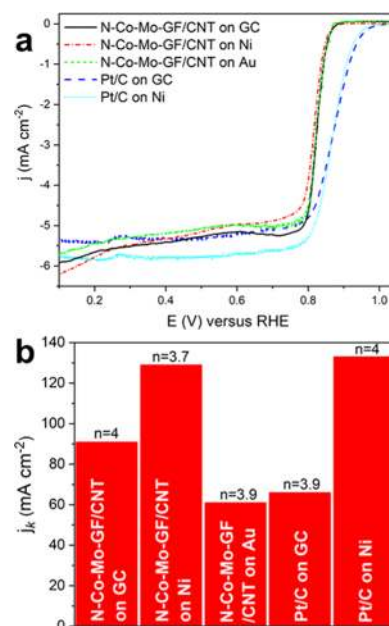


Figure 4. (a) ORR polarization curves in an O₂-saturated 0.1 M KOH solution at 1600 rpm and (b) ORR kinetic currents and number of transferred electrons obtained at 0.7 V vs RHE for N-Co-Mo-GF/CNT on GC, Ni, and Au substrates, as well as for Pt/C on GC and Ni substrates.

(LSV) comparison of N-Co-Mo-GF/CNT and conventional Pt/C 20 wt % on different substrates at a rotation of 1600 rpm. The corresponding ORR Tafel plots are depicted in Figure S7. The prepared catalyst exhibited a remarkable ORR catalytic activity on all of the substrates with a half-potential of ~0.83 V, and an onset potential of ~0.9 V relative to the reversible hydrogen electrode (RHE). The ORR Tafel slopes for the N-Co-Mo-GF/CNT catalyst were 27, 37, and 44 mV dec⁻¹ on Ni, GC, and Au substrates, while, for Pt/C, they were 54 mV dec⁻¹ on both GC and Ni substrates (Figure S7).

The kinetic current and the number of electrons involved per O₂ in the ORR on N-Co-Mo-GF/CNT were calculated based on the Koutecky-Levich (K-L) equation, as explained in the Supporting Information. The ORR LSV polarization curves at various rotation rates (400, 700, 900, 1200, 1600, 2000, and 2500 rpm) for N-Co-Mo-GF/CNT on GC, Ni, and Au substrates and for Pt/C on GC and Ni substrates, together with the corresponding linear K-L plots through the inverse current density (j^{-1}) as a function of the inverse of the square root of the rotation speed ($\omega^{-1/2}$) at 0.7 V, are shown in

Figure S8. The kinetic current density (j_k), measured from the y -intercept of linear K–L plots at a potential of 0.7 V vs RHE, is depicted in Figure 4b. For the N–Co–Mo–GF/CNT on various substrates, j_k increased from 91 mA cm⁻² on GC to 129 mA cm⁻² on Ni, but decreased to 61 mA cm⁻² on the Au substrate. Similarly, for Pt/C, the improvement in j_k on GC vs the Ni substrate was significant, so that j_k increased from 66 to 133 mA cm⁻². These results show that the Ni substrate can considerably improve the ORR kinetic current for both Pt/C and N–Co–Mo–GF/CNT catalysts, rendering Ni a promising substrate for alkaline ORR. The corresponding number of transferred electrons on all tested substrates was close to 4 (Figure 4b), indicating that the ORR was dominated by a four-electron process.

Retention of the ORR current was approximately 97% under 0.70 V for 21 h using the N–Co–Mo–GF/CNT electrocatalyst, showing higher stability than Pt/C (Figure S9). The N–Co–Mo–GF/CNT electrocatalyst also exhibited an improved ORR catalytic activity and stability in comparison to the previously reported multi-heteroatom-doped porous carbon catalyst.⁶⁷ The activity of this catalyst for ORR is comparable to that of the best heteroatom-doped carbon nanomaterials,² nonprecious M–N–C,^{3,4,68–71} and single- (and few-) atom⁷² ORR electrocatalysts reported so far in terms of the ORR half-wave potential, kinetic current density, and stability.

2.4. Oxygen Evolution Activity and Discussion. The electrocatalytic activity of N–Co–Mo–GF/CNT for OER was also investigated by RDE measurements in 0.1 M KOH. The OER polarization curves of the prepared catalyst on GC, Ni, and Au substrates are compared with a well-established water oxidation catalyst, RuO₂, on GC and Ni substrates in Figure 5a. On the GC substrate, RuO₂ exhibited a slightly higher activity at low current densities (<20 mA cm⁻²) than N–Co–Mo–GF/CNT. However, at high current densities (>20 mA cm⁻²), the synthesized catalyst suppressed RuO₂ tested in this work so that it required a potential of 1.68 V, rather than 1.72 V for RuO₂, to reach a current of 50 mA cm⁻². In contrast to manganese- and cobalt-based oxides^{38–41,44} and NiCeO_x⁴³ which demonstrate drastic improvements in activity for OER when the surface or substrate is enriched with Au, we did not observe a significant enhancement of activity with an Au substrate. Nevertheless, the OER activity of the N–Co–Mo–GF/CNT catalyst was remarkably enhanced when it was coated on a Ni substrate. The onset overpotential and the required overpotential to reach 10 mA cm⁻² ($\eta_{\text{OER},10}$) decreased by ~50 and ~77 mV, respectively, on Ni in comparison to the GC substrate. For RuO₂, the onset overpotential over Ni was similar to that over GC, but the ORR current improved and the $\eta_{\text{OER},10}$ was ca. 26 mV lower. Nonetheless, for the N–Co–Mo–GF/CNT catalyst, which has a high-surface-area mesoporous structure, the electrolyte can more readily permeate the underlying layers and the Ni/catalyst interface, where a synergistic effect occurs, forming new active sites with a higher reactivity and a lower onset overpotential. Figure S10 illustrates the synergistic effect between the Ni substrate and the N–Co–Mo–GF/CNT electrocatalyst for enhancing activity for OER. The OER Tafel plots of the prepared catalyst and RuO₂ are shown in Figure 5b. The Tafel slope values for the N–Co–Mo–GF/CNT catalyst on Ni, Au, and GC substrates follow a trend of Ni < GC < Au, where a lower slope indicates a more active electrode. RuO₂ also had a lower Tafel slope on Ni than on

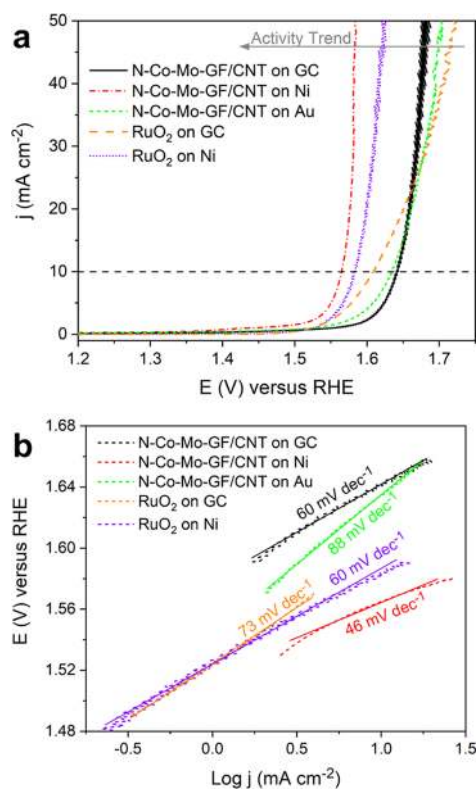


Figure 5. (a) OER polarization curves and (b) corresponding OER Tafel plots of the N–Co–Mo–GF/CNT catalyst on GC, Ni, and Au compared to Pt/C on GC and Ni substrates in 0.1 M KOH.

GC. Since the Tafel slope is related to the reaction mechanism, the difference perhaps arises from a change in the rate-determining step of the OER.

The N–Co–Mo–GF/CNT catalyst on Ni required overpotentials of just 330 and 350 mV to reach 10 and 50 mA cm⁻² (denoted $\eta_{\text{OER},10}$ and $\eta_{\text{OER},50}$), respectively. The catalyst's performance upon a 24 h continuous operation at 20 mA cm⁻² was also highly stable (Figure S11). The OER electrocatalytic performance of N–Co–Mo–GF/CNT on Ni was close to that reported previously for highly active OER electrocatalysts,^{2,73–78} especially at high currents.

Substrate effects enhancing the OER activity of RuO₂ have recently been reported.⁷⁹ Here, we demonstrated that for a mesoporous electrolyte-permeable catalyst, such as N–Co–Mo–GF/CNT, the role of the substrate could be more significant. We believe that this represents a route to further developing the activity of mesoporous carbon-based catalysts through the addition of suitable active metals to the substrate, or to the surface of the catalyst, where new active sites are formed at the metal/catalyst interface.

Hence, the high-surface-area mesoporous N–Co–Mo–GF/CNT hybrid material represents a stable ORR/OER electrocatalyst with a high activity that can be further improved using a Ni substrate. This catalyst exhibited an advanced electrocatalytic performance in terms of both ORR and OER activities compared to the best bifunctional ORR/OER catalysts reported in the literature.^{2,80–85}

2.5. Active Sites for Oxygen Evolution and Reduction Reactions. **2.5.1. Active Sites Over the Glassy Carbon Substrate.** The C–C sites in pristine graphene and CNTs are almost catalytically inactive. Carbon nanomaterials doped by heteroatoms, however, form new active sites that can activate

them for ORR^{20–22} and OER.^{23,24} In the case of the N–Co–Mo–GF/CNT catalyst as an M–N–C catalyst, the N–C and/or the M (Co, and Mo)–C sites act as the active sites for ORR and OER. To find out whether the N–C and/or M–C sites are responsible for the observed activity, an electrocatalyst was synthesized without the temporary introduction of acetonitrile, thus lacking a nitrogen source (see Section 2.1 above). This material is denoted Co–Mo–GF/CNT.

The comparison of the ORR activity of the Co–Mo–GF/CNT and N–Co–Mo–GF/CNT samples (Figures 6a and

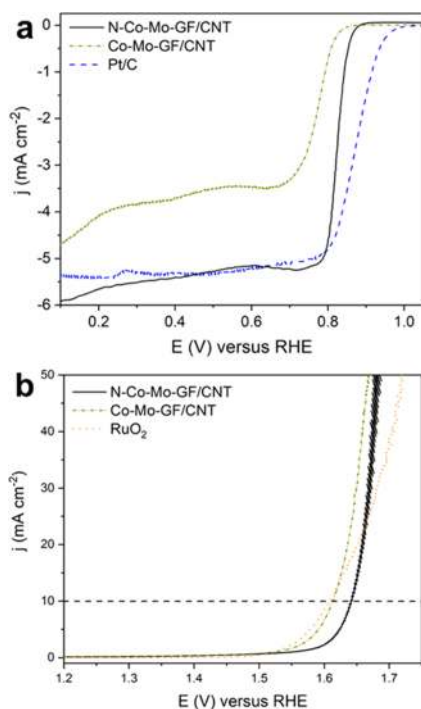


Figure 6. Effect of N heteroatoms on the ORR and OER activities of the N–Co–Mo–GF/CNT electrocatalyst. (a) ORR and (b) OER polarization curves of Co–Mo–GF/CNT compared to that of N–Co–Mo–GF/CNT. The polarization curves were obtained in a 0.1 M KOH solution at a scan rate of 5 mV s^{−1} and a 1600 rpm rotation speed. The electrocatalysts were deposited on a GC substrate.

S12) revealed that the material without N sites remained active for ORR, but that its activity was considerably lower than that in the presence of N heteroatoms. In comparison, doping with N sites resulted in an ORR half-wave potential increase of ~50 mV and a 7-fold enhancement in kinetic current density (from 13 to 91 mA cm^{−2} at 0.7 V). Meanwhile, the number of

transferred electrons also increased from 3.2 to 4 (Figure S12). This change in ORR activity indicates that both the N–C and M–C sites are synergistically responsible for ORR activity.

The comparison of the OER activity of the Co–Mo–GF/CNT and N–Co–Mo–GF/CNT samples is shown in Figure 6b. Doping with N sites increased the $\eta_{\text{OER},10}$ and $\eta_{\text{OER},50}$ by 27 and 13 mV, respectively, indicating that the sample without N has slightly higher OER activity than the N-doped sample. The Tafel slope of the N-doped sample is, however, slightly lower than that without N (Figure S13). We can, therefore, conclude that the main OER active sites in the N–Co–Mo–GF/CNT are likely to be the M–C sites. The slight decrease in OER activity after doping with N heteroatoms can be attributed to the increase in the number of defect sites in the graphitic network, which is also corroborated by Raman spectroscopy (Figure 3). The stability of Co–Mo–GF/CNT for OER was also tested by a chronopotentiometry measurement at 20 mA cm^{−2}, where a stable performance upon continuous operation for 12 h was observed (Figure S14).

2.5.2. Active Sites over the Ni Substrate. **2.5.2.1. X-ray Photoelectron Spectroscopy (XPS): Investigating the Possibility of Ni Diffusion into the Catalyst Film, and Changes in Catalyst Surface Composition during Electrocatalysis.** XPS was used to study the chemical composition of the N–Co–Mo–GF/CNT sample. Figure 7 shows the cobalt 2p, molybdenum 3d, and nitrogen 1s regions of the sample both before and after electrochemical ORR and OER measurements. The related carbon 1s, oxygen 1s, and survey spectra, as well as atomic concentrations, are provided in the Supporting Information (Figure S15 and Table S1).

Figure 7a displays the Co 2p spectra of the N–Co–Mo–GF/CNT sample. The Co 2p_{3/2} peak has been deconvoluted assuming the presence of two chemical states: Co(0) and Co(II).⁸⁶ In the pristine material, mainly Co(0) is found at a 2p_{3/2} binding energy of 778.1 eV, but a ca. 25% contribution of Co(II) is also required to explain the spectra. The binding energies and spectral shape correspond to Co(0). After OER, a typical Co 2p spectrum of Co(II) is observed, with a 2p_{3/2} peak at roughly 780.24 eV,⁸⁶ and a small contribution from Co(0) is still observed at ~778 eV.

The Mo 3d spectra are shown in Figure 7b. The spectra have been deconvoluted, assuming the presence of Mo(0) and Mo(VI).⁸⁷ In the pristine material, the spectrum shows a clear 3d_{5/2} peak at 228.2 eV corresponding to Mo(0) and another 3d_{5/2} peak at roughly 232.3 eV corresponding to Mo(VI).⁸⁷ The 3d_{3/2} peak of Mo(0) at roughly 231.3 eV overlaps with the Mo(VI) 3d_{5/2} peak. The 3d_{3/2} peak of Mo(VI) is found at a binding energy of 235.5 eV. After ORR/OER, the Mo 3d

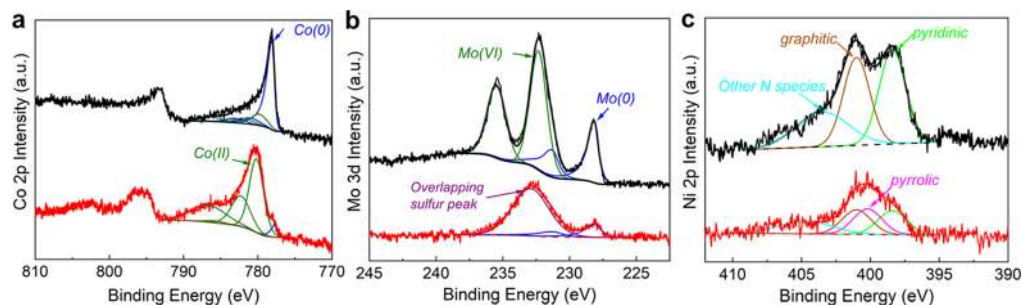


Figure 7. X-ray photoelectron spectra of the N–Co–Mo–GF/CNT catalyst before (black lines) and after (red lines) the ORR and OER measurements on the Ni substrate: (a) Co 2p region, (b) Mo 3d region, and (c) N 1s region including the deconvolution.

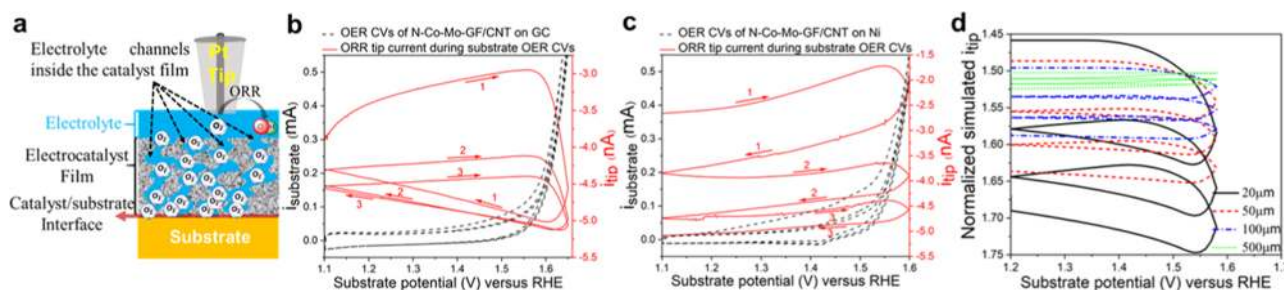


Figure 8. (a) Schematic representation of the Pt SECM tip for approaching the electrocatalyst surface to observe the feedback from the electrocatalyst and substrate, as well as O_2 diffusion from the substrate/catalyst interface to the tip. (b, c) Results of substrate generation–tip collection experiments, where the substrate potential was scanned to the OER region at a scan rate of 5 mV s^{-1} , while the tip potential was kept at a constant value of 0.3 V to drive ORR. The tip was placed close to the surface of N–Co–Mo–GF/CNT on (b) GC and (c) Ni substrates. (d) Simulated transient tip ORR currents that obtained for the experimental OER CVs in (c) for the several oxygen diffusion distances, showing that a similar ORR tip current response to that experimentally measured in (c) is reproduced if the OER takes place significantly further from the catalyst top surface, at approximately $5 \mu\text{m}$ from the tip. The tip current in (d) is normalized by the ORR current in the bulk.

spectrum is dominated by an overlapping sulfur 2s peak at roughly 232.9 eV , which is attributed to sulfate impurities from the electrolyte. Some Mo(0) is still observed, but its contribution has been clearly reduced (see Table S1 in the Supporting Information). Also, the lack of a $3d_{3/2}$ peak at 235.54 eV related to Mo(VI) indicates that the amount of oxidized molybdenum has significantly decreased.

The oxidation of Co and Mo after OER measurements can be attributed to the Co and Mo nanoparticles that were initially encapsulated in carbon layers but oxidized during the OER. During the OER, the graphene layers that protect the metal nanoparticles can be stripped off (Figure S16) due to the presence of highly reactive pentagonal rings and significant strain on the C–C bonds. This causes permeation of the electrolyte into the metal core and subsequent oxidation.⁷⁷ The decrease in the amount of Mo after the electrochemical measurements can be attributed to the anodic dissolution of molybdenum in the alkaline solution.⁸⁸

Figure 7c shows high-resolution N 1s spectra of the N–Co–Mo–GF/CNT sample before and after the electrochemical measurements. The spectra have been deconvoluted assuming a generally used model with the following N-functional groups: pyridinic-N (ca. 398.3 eV), pyrrolic-N (ca. 400.1 eV), graphitic-N (ca. 400.9 eV), and other N species (ca. $403\text{--}405 \text{ eV}$).^{89,90} The peak in binding energies between 403 and 405 eV can be attributed to various oxidized nitrogen configurations, clustered N substitutions, hydrogenation of the nitrogen dopants, and N_2 molecules trapped inside carbon nanotubes or between graphene layers.⁸⁹ The relative concentrations of the different nitrogen species in the pristine sample were 38% pyridinic-N, 35% graphitic-N, and 27% other N species. After OER, the pyrrolic-N component was also required for a satisfactory fit, yielding relative concentrations of 25% pyridinic-N, 26% pyrrolic-N, 25% graphitic-N, and 24% other N species. The XPS analysis of the N–Co–Mo–GF/CNT sample revealed the presence of 1 atom % N before and after ORR/OER measurements (Table S1).

According to earlier experimental work with graphene nanoribbon networks doped with nitrogen, the electron-donating graphitic-N moieties serve as active sites for ORR, while the electron-withdrawing pyridinic-N sites are responsible for OER.¹ Here, however, we could not verify that the N moieties served as the active sites for OER, but, instead, we observed that single metal atoms in the graphene flakes were the main sites responsible for OER.

To detect possible Ni diffusion from the substrate into the N–Co–Mo–GF/CNT film during electrochemical measurements, the catalyst material over the Ni substrate was transferred onto an Au substrate after the ORR/OER measurements. This prevented substrate Ni detection in the XPS measurements. In these measurements, Ni was not detected, meaning that Ni did not diffuse into the catalyst during the electrochemical measurements. This is in clear contrast to what has been observed for $NiCeO_x$ deposited onto an Au electrode, in which XPS depth profiles demonstrated that Au from the substrate could diffuse into the catalyst film.⁴³

2.5.2.2. Scanning Electrochemical Microscopy (SECM) Analysis: Unraveling Where the OER Reaction Takes Place on the Ni Substrate. The OER takes place on the catalyst surface and/or in the underlying layers where new sites are formed at the substrate/catalyst interface. To study where the OER reaction actually takes place, scanning electrochemical microscopy (SECM) was used. SECM is a scanning probe technique capable of imaging local reactivity at a high spatial resolution. Here, using SECM, we established that the enhanced OER activity of N–Co–Mo–GF/CNT over the Ni (rather than the GC) substrate was due to newly formed interfacial sites, thus yielding a synergistic effect.

Figure S17 shows the theoretical approach curves of the SECM tip for the positive (enhancement in the tip current) and negative (reduction in the tip current) feedbacks, as well as the overlay of the experimental approach curves for the tip on Ni and on the N–Co–Mo–GF/CNT mesoporous film deposited on Ni. A similar response was also observed on the GC substrate (data not shown here). A negative feedback to the tip ORR current was observed over both GC and Ni when approaching the surface (Figure S17) because the substrate can block the flux of the oxygen to the Pt tip. In contrast, for the N–Co–Mo–GF/CNT mesoporous film, a positive feedback on the Pt tip was observed, as illustrated in Figure S17. This indicates that some oxygen was trapped within the catalyst layer due to hydrophobicity or the porous surface structure of the GFs and CNTs. A similar observation has been reported when approaching a porous poly(vinylidene fluoride) membrane.⁹¹ As the tip depletes oxygen from the solution, oxygen from the porous film partitions into the solution, resulting in similar behavior to that observed when approaching a water–air interface.^{92–94} The simulations described in the Supporting Information and Figures S17 and S18 were conducted to reproduce the positive feedback for

the tip approach curve over a porous catalyst, which traps oxygen molecules within the porous layer. Mass transport in porous electrodes has been studied previously⁹⁵ by SECM, but not with gaseous species partitioning into the porous electrode. The simulated approach curve reproduced well the experimentally observed positive feedback to the ORR tip current that was obtained on the N–Co–Mo–GF/CNT electrocatalyst surface as shown in Figure S17. The oxygen from the film could be depleted by applying a sufficiently low potential (~ 0 – 0.3 V vs RHE) to the substrate for reducing oxygen. When this was done with the Pt tip (kept at 0.3 V vs RHE) near the substrate, the ORR tip current approached zero. Interestingly, when the substrate potential was removed, after a sufficient time, a positive feedback was again observed when approaching the mesoporous film. This indicates that the dissolved oxygen from the solution can partition into the porous electrode, thus restoring the initial situation.

In the substrate generation–tip collection (SG–TC) mode of SECM, as schematically shown in Figure 8a, oxygen that evolves on the substrate during the OER is subsequently reduced and detected on the probe tip placed close to the substrate surface.⁹⁶ When OER cyclic voltammetry (CV) is carried out on the substrate, the ORR current response on the SECM tip will vary depending on the distance between the tip and the sites at which OER takes place on the substrate.⁹⁶ Here, depending on how the SECM tip current responds to the formation of O_2 on the substrate, we can interpret whether the OER takes place within the catalyst layer or at the catalyst/substrate interface. This approach can hence confirm whether the enhanced activity of N–Co–Mo–GF/CNT over the Ni (rather than the GC) substrate is indeed due to newly formed interfacial sites with a synergistic effect.

The SG–TC SECM experiments were carried out with the N–Co–Mo–GF/CNT catalyst deposited on GC (Figure 8b) and Ni (Figure 8c) substrates. The substrate potential was scanned to the OER region, while the tip potential was kept at a constant value of 0.3 V to drive the ORR sufficiently. The experimental results in Figure 8c show that it took much longer for oxygen to reach the tip when Ni was used as the substrate, resulting in a stepwise increase in the tip current. When GC was used as the substrate, the response was more immediate but also decayed much faster. This indicates that when Ni was used as the substrate, the evolved oxygen remained in the catalyst layer, while, on GC, the oxygen was released faster. To simulate this effect qualitatively, the oxygen diffusion distance to the tip was varied between 20 and 500 μm by increasing the thickness of the porous layer, as shown in Figure 8d. Simulations showed that when the diffusion distance was small, the tip response was fast, and the decay in the current was also more rapid. When the distance increased, however, the time for oxygen to reach the tip was greater, leading to a similar response to that observed experimentally on the Ni substrate (Figure 8c). This indicates that on GC, oxygen evolution takes place close to the catalyst surface and/or the oxygen is released rapidly from the catalyst layer. Meanwhile, on the Ni substrate, OER takes place deeper in the catalyst layer (see Scheme S1). These results indicate that OER, in fact, takes place close to the electrocatalyst/Ni interface, where new active sites can be formed with a higher OER electrocatalytic activity than that of sites on the catalyst surface. These results are in good agreement with an earlier report by Snook et al.⁹⁶ They demonstrated that by increasing the distance between the tip and the substrate, the ORR current on the tip does not

immediately decrease when the OER current on the substrate is reduced during the backward scan of OER CV, due to the time required for the oxygen generated at the substrate to diffuse to the probe tip. As a result, when the O_2 diffusion length from the substrate is increased, the ORR on the tip also takes place later than the OER onset potential.⁹⁶ This also explains why the ORR tip current in Figure 8b increases at almost the same time at which the substrate starts to produce O_2 (at OER onset potential), while, in Figure 8c, the enhancement in the tip current takes place later than O_2 production on the substrate. These SG–TC SECM experiments thus provide a novel technique to observe where the electrocatalytic reaction takes place in porous multilayer electrocatalysts.

3. CONCLUSIONS

In summary, a high-surface-area mesoporous hybrid of graphene nanoflakes and CNTs doped with single atoms of N, Co, and Mo has been synthesized using a fast and scalable, low-cost one-step chemical vapor deposition process. The material shows high activity and stability for catalyzing ORR and OER in alkaline media. In the prepared catalyst, we demonstrate that N–C sites do not serve as the main active sites for OER activity, although they do have a synergistic effect with M–C sites, thus enhancing activity for ORR. We have further demonstrated how the use of Ni, Au, and GC substrates affects the electrocatalytic activity of the synthesized catalyst for ORR and OER. The use of a Ni substrate, in comparison to GC and Au, increases the ORR kinetic current and significantly improves the OER activity of the mesoporous catalyst. Based on XPS analysis, no Ni diffusion from the substrate to the porous catalyst was detected during the electrochemical measurements. Instead, enhanced electrocatalytic activity is attributed to the catalyst/Ni interfacial sites that are accessible through the pores of the electrocatalyst film. By applying the developed SECM analysis method, we further confirm that instead of the catalyst surface, a significant amount of the evolved O_2 at the low overpotentials during OER originated from the underlying catalyst/Ni interface. We have thus established a new catalyst for alkaline OER and ORR with tunable reactivity depending on the choice of the underlying substrate, which opens new avenues for the growth of heteroatom-doped graphene–CNT hybrids for various applications.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acscatal.0c00352>.

STEM and TEM images; N_2 adsorption/desorption isotherms; additional electrochemical ORR and OER measurements; additional details on electrochemical procedures, physical, and chemical characterizations; XPS; SECM analysis; and related references (PDF)

■ AUTHOR INFORMATION

Corresponding Author

Mohammad Tavakkoli – Department of Applied Physics, Aalto University School of Science, FI-00076 Aalto, Finland;

orcid.org/0000-0003-4859-1922;

Phone: +358504140950; Email: mohammad.tavakkoli@aalto.fi

Authors

Emmanuel Flahaut – CIRIMAT, Université de Toulouse, CNRS, INPT, UPS, UMR CNRS-UPS-INP No 5085, Université Toulouse 3 Paul Sabatier, 31062 Toulouse, France

Pekka Peljo – Research Group of Physical Electrochemistry and Electrochemical Physics, Department of Chemistry and Material Sciences, Aalto University School of Chemical Engineering, FI-00076 Aalto, Finland; orcid.org/0000-0002-1229-2261

Jani Sainio – Department of Applied Physics, Aalto University School of Science, FI-00076 Aalto, Finland

Fatemeh Davodi – Department of Chemistry and Material Sciences, Aalto University School of Chemical Engineering, FI-00076 Aalto, Finland

Egor V. Lobiak – Nikolaev Institute of Inorganic Chemistry, SB RAS, 630090 Novosibirsk, Russia; orcid.org/0000-0001-5085-5531

Kimmo Mustonen – Faculty of Physics, University of Vienna, 1090 Vienna, Austria; orcid.org/0000-0002-0953-7299

Esko I. Kauppinen – Department of Applied Physics, Aalto University School of Science, FI-00076 Aalto, Finland; orcid.org/0000-0003-1727-8810

Complete contact information is available at: <https://pubs.acs.org/10.1021/acscatal.0c00352>

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

This work has been supported by the Academy of Finland (ANCED project) and the Magnus Ehrnrooth Foundation. The authors also thank Prof. Kari Laasonen for fruitful discussions and Dr. Hua Jiang for taking the TEM images.

REFERENCES

- (1) Yang, H. B.; Miao, J.; Hung, S.-F.; Chen, J.; Tao, H. B.; Wang, X.; Zhang, L.; Chen, R.; Gao, J.; Chen, H. M.; Dai, L.; Liu, B. Identification of catalytic sites for oxygen reduction and oxygen evolution in N-doped graphene materials: Development of highly efficient metal-free bifunctional electrocatalyst. *Sci. Adv.* **2016**, *2*, No. e1501122.
- (2) Paul, R.; Du, F.; Dai, L.; Ding, Y.; Wang, Z. L.; Wei, F.; Roy, A. 3D Heteroatom-Doped Carbon Nanomaterials as Multifunctional Metal-Free Catalysts for Integrated Energy Devices. *Adv. Mater.* **2019**, *31*, No. 1805598.
- (3) Lee, S. H.; Kim, J.; Chung, D. Y.; Yoo, J. M.; Lee, H. S.; Kim, M. J.; Mun, B. S.; Kwon, S. G.; Sung, Y.-E.; Hyeon, T. Design Principle of Fe–N–C Electrocatalysts: How to Optimize Multimodal Porous Structures? *J. Am. Chem. Soc.* **2019**, *141*, 2035–2045.
- (4) Yan, X.-H.; Xu, B.-Q. Mesoporous carbon material co-doped with nitrogen and iron (Fe–N–C): high-performance cathode catalyst for oxygen reduction reaction in alkaline electrolyte. *J. Mater. Chem. A* **2014**, *2*, 8617–8622.
- (5) Wu, G.; More, K. L.; Johnston, C. M.; Zelenay, P. High-Performance Electrocatalysts for Oxygen Reduction Derived from Polyaniline, Iron, and Cobalt. *Science* **2011**, *332*, 443–447.
- (6) Parvez, K.; Yang, S.; Hernandez, Y.; Winter, A.; Turchanin, A.; Feng, X.; Müllen, K. Nitrogen-Doped Graphene and Its Iron-Based Composite As Efficient Electrocatalysts for Oxygen Reduction Reaction. *ACS Nano* **2012**, *6*, 9541–9550.
- (7) Davodi, F.; Mühlhausen, E.; Tavakkoli, M.; Sainio, J.; Jiang, H.; Gökce, B.; Marzun, G.; Kallio, T. Catalyst Support Effect on the Activity and Durability of Magnetic Nanoparticles: toward Design of Advanced Electrocatalyst for Full Water Splitting. *ACS Appl. Mater. Interfaces* **2018**, *10*, 31300–31311.

- (8) Wu, G.; Johnston, C. M.; Mack, N. H.; Artyushkova, K.; Ferrandon, M.; Nelson, M.; Lezama-Pacheco, J. S.; Conradson, S. D.; More, K. L.; Myers, D. J.; Zelenay, P. Synthesis–structure–performance correlation for polyaniline–Me–C non-precious metal cathode catalysts for oxygen reduction in fuel cells. *J. Mater. Chem.* **2011**, *21*, 11392–11405.

- (9) Dicks, A. L. The role of carbon in fuel cells. *J. Power Sources* **2006**, *156*, 128–141.

- (10) Sharma, S.; Pollet, B. G. Support materials for PEMFC and DMFC electrocatalysts—A review. *J. Power Sources* **2012**, *208*, 96–119.

- (11) Yarlagadda, V.; Carpenter, M. K.; Moylan, T. E.; Kukreja, R. S.; Koestner, R.; Gu, W.; Thompson, L.; Kongkanand, A. Boosting Fuel Cell Performance with Accessible Carbon Mesopores. *ACS Energy Lett.* **2018**, *3*, 618–621.

- (12) Park, J.; Nabae, Y.; Hayakawa, T.; Kakimoto, M.-a. Highly Selective Two-Electron Oxygen Reduction Catalyzed by Mesoporous Nitrogen-Doped Carbon. *ACS Catal.* **2014**, *4*, 3749–3754.

- (13) Proietti, E.; Jaouen, F.; Lefèvre, M.; Larouche, N.; Tian, J.; Herranz, J.; Dodelet, J.-P. Iron-based cathode catalyst with enhanced power density in polymer electrolyte membrane fuel cells. *Nat. Commun.* **2011**, *2*, No. 416.

- (14) Liang, H.-W.; Wei, W.; Wu, Z.-S.; Feng, X.; Müllen, K. Mesoporous Metal–Nitrogen-Doped Carbon Electrocatalysts for Highly Efficient Oxygen Reduction Reaction. *J. Am. Chem. Soc.* **2013**, *135*, 16002–16005.

- (15) Ding, J.; Wang, P.; Ji, S.; Wang, H.; Linkov, V.; Wang, R. N-doped mesoporous FeN_x/carbon as ORR and OER bifunctional electrocatalyst for rechargeable zinc-air batteries. *Electrochim. Acta* **2019**, *296*, 653–661.

- (16) Ding, J.; Ji, S.; Wang, H.; Linkov, V.; Gai, H.; Liu, F.; Liu, Q.; Wang, R. N-Doped 3D Porous Ni/C Bifunctional Electrocatalysts for Alkaline Water Electrolysis. *ACS Sustainable Chem. Eng.* **2019**, *7*, 3974–3981.

- (17) Shao, Y.; Yin, G.; Zhang, J.; Gao, Y. Comparative investigation of the resistance to electrochemical oxidation of carbon black and carbon nanotubes in aqueous sulfuric acid solution. *Electrochim. Acta* **2006**, *51*, 5853–5857.

- (18) Antolini, E.; Gonzalez, E. R. Ceramic materials as supports for low-temperature fuel cell catalysts. *Solid State Ionics* **2009**, *180*, 746–763.

- (19) Hu, C.; Liu, D.; Xiao, Y.; Dai, L. Functionalization of graphene materials by heteroatom-doping for energy conversion and storage. *Prog. Nat. Sci.: Mater. Int.* **2018**, *28*, 121–132.

- (20) Yang, L.; Jiang, S.; Zhao, Y.; Zhu, L.; Chen, S.; Wang, X.; Wu, Q.; Ma, J.; Ma, Y.; Hu, Z. Boron-Doped Carbon Nanotubes as Metal-Free Electrocatalysts for the Oxygen Reduction Reaction. *Angew. Chem.* **2011**, *123*, 7270–7273.

- (21) Yang, Z.; Yao, Z.; Li, G.; Fang, G.; Nie, H.; Liu, Z.; Zhou, X.; Chen, X.; Huang, S. Sulfur-Doped Graphene as an Efficient Metal-free Cathode Catalyst for Oxygen Reduction. *ACS Nano* **2012**, *6*, 205–211.

- (22) Jin, J.; Pan, F.; Jiang, L.; Fu, X.; Liang, A.; Wei, Z.; Zhang, J.; Sun, G. Catalyst-Free Synthesis of Crumpled Boron and Nitrogen Co-Doped Graphite Layers with Tunable Bond Structure for Oxygen Reduction Reaction. *ACS Nano* **2014**, *8*, 3313–3321.

- (23) Davodi, F.; Tavakkoli, M.; Lahtinen, J.; Kallio, T. Straightforward synthesis of nitrogen-doped carbon nanotubes as highly active bifunctional electrocatalysts for full water splitting. *J. Catal.* **2017**, *353*, 19–27.

- (24) Ma, T. Y.; Dai, S.; Jaroniec, M.; Qiao, S. Z. Graphitic Carbon Nitride Nanosheet–Carbon Nanotube Three-Dimensional Porous Composites as High-Performance Oxygen Evolution Electrocatalysts. *Angew. Chem., Int. Ed.* **2014**, *53*, 7281–7285.

- (25) Yan, Z.; Ma, L.; Zhu, Y.; Lahiri, I.; Hahm, M. G.; Liu, Z.; Yang, S.; Xiang, C.; Lu, W.; Peng, Z.; Sun, Z.; Kittrell, C.; Lou, J.; Choi, W.; Ajayan, P. M.; Tour, J. M. Three-Dimensional Metal–Graphene–Nanotube Multifunctional Hybrid Materials. *ACS Nano* **2013**, *7*, 58–64.

- (26) Chen, Z.; Chen, X.-D.; Wang, H.; Li, X.; Lin, L.; Chen, K.; Ci, H.; Wu, X.; Zhang, Y.; Zhang, Y.; Liu, Z. One-Step Growth of Graphene/Carbon Nanotube Hybrid Films on Soda-Lime Glass for Transparent Conducting Applications. *Adv. Electron. Mater.* **2017**, *3*, No. 1700212.
- (27) Lin, C.-C.; Lin, Y.-W. Synthesis of Carbon Nanotube/Graphene Composites by One-Step Chemical Vapor Deposition for Electrodes of Electrochemical Capacitors. *J. Nanomater.* **2015**, *2015*, No. 741928.
- (28) Maarouf, A. A.; Kasry, A.; Chandra, B.; Martyna, G. J. A graphene-carbon nanotube hybrid material for photovoltaic applications. *Carbon* **2016**, *102*, 74–80.
- (29) Cheng, Q.; Tang, J.; Ma, J.; Zhang, H.; Shinya, N.; Qin, L.-C. Graphene and carbon nanotube composite electrodes for supercapacitors with ultra-high energy density. *Phys. Chem. Chem. Phys.* **2011**, *13*, 17615–17624.
- (30) Yu, D.; Dai, L. Self-Assembled Graphene/Carbon Nanotube Hybrid Films for Supercapacitors. *J. Phys. Chem. Lett.* **2010**, *1*, 467–470.
- (31) You, B.; Wang, L.; Yao, L.; Yang, J. Three dimensional N-doped graphene-CNT networks for supercapacitor. *Chem. Commun.* **2013**, *49*, 5016–5018.
- (32) Dong, X.; Li, B.; Wei, A.; Cao, X.; Chan-Park, M. B.; Zhang, H.; Li, L.-J.; Huang, W.; Chen, P. One-step growth of graphene-carbon nanotube hybrid materials by chemical vapor deposition. *Carbon* **2011**, *49*, 2944–2949.
- (33) Li, H.-F.; Wu, F.; Wang, C.; Zhang, P.-X.; Hu, H.-Y.; Xie, N.; Pan, M.; Zeng, Z.; Deng, S.; Wu, M. H.; Vinodgopal, K.; Dai, G.-P. One-Step Chemical Vapor Deposition Synthesis of 3D N-doped Carbon Nanotube/N-doped Graphene Hybrid Material on Nickel Foam. *Nanomaterials* **2018**, *8*, No. 700.
- (34) Lobiak, E. V.; Bulusheva, L. G.; Fedorovskaya, E. O.; Shubin, Y. V.; Plyusnin, P. E.; Lonchambon, P.; Senkovskiy, B. V.; Ismagilov, Z. R.; Flahaut, E.; Okotrub, A. V. One-step chemical vapor deposition synthesis and supercapacitor performance of nitrogen-doped porous carbon-carbon nanotube hybrids. *Beilstein J. Nanotechnol.* **2017**, *8*, 2669–2679.
- (35) Liao, Y.; Mustonen, K.; Tulić, S.; Skákalová, V.; Khan, S. A.; Laiho, P.; Zhang, Q.; Li, C.; Monazam, M. R. A.; Kotakoski, J.; Lipsanen, H.; Kauppinen, E. I. Enhanced Tunneling in a Hybrid of Single-Walled Carbon Nanotubes and Graphene. *ACS Nano* **2019**, *13*, 11522–11529.
- (36) Kishor, K.; Saha, S.; Sivakumar, S.; Pala, R. G. S. Enhanced Water Oxidation Activity of the Cobalt(II,III) Oxide Electrocatalyst on an Earth-Abundant-Metal-Interlayered Hybrid Porous Carbon Support. *ChemElectroChem* **2016**, *3*, 1899–1907.
- (37) Koza, J. A.; He, Z.; Miller, A. S.; Switzer, J. A. Electrodeposition of Crystalline Co_3O_4 —A Catalyst for the Oxygen Evolution Reaction. *Chem. Mater.* **2012**, *24*, 3567–3573.
- (38) Gorlin, Y.; Chung, C.-J.; Benck, J. D.; Nordlund, D.; Seitz, L.; Weng, T.-C.; Sokaras, D.; Clemens, B. M.; Jaramillo, T. F. Understanding Interactions between Manganese Oxide and Gold That Lead to Enhanced Activity for Electrocatalytic Water Oxidation. *J. Am. Chem. Soc.* **2014**, *136*, 4920–4926.
- (39) Seitz, L. C.; Hersbach, T. J. P.; Nordlund, D.; Jaramillo, T. F. Enhancement Effect of Noble Metals on Manganese Oxide for the Oxygen Evolution Reaction. *J. Phys. Chem. Lett.* **2015**, *6*, 4178–4183.
- (40) Yeo, B. S.; Bell, A. T. Enhanced Activity of Gold-Supported Cobalt Oxide for the Electrochemical Evolution of Oxygen. *J. Am. Chem. Soc.* **2011**, *133*, 5587–5593.
- (41) Yeo, B. S.; Bell, A. T. In Situ Raman Study of Nickel Oxide and Gold-Supported Nickel Oxide Catalysts for the Electrochemical Evolution of Oxygen. *J. Phys. Chem. C* **2012**, *116*, 8394–8400.
- (42) Zou, S.; Burke, M. S.; Kast, M. G.; Fan, J.; Danilovic, N.; Boettcher, S. W. Fe (Oxy)hydroxide Oxygen Evolution Reaction Electrocatalysis: Intrinsic Activity and the Roles of Electrical Conductivity, Substrate, and Dissolution. *Chem. Mater.* **2015**, *27*, 8011–8020.
- (43) Ng, J. W. D.; García-Melchor, M.; Bajdich, M.; Chakhranont, P.; Kirk, C.; Vojvodic, A.; Jaramillo, T. F. Gold-supported cerium-doped NiOx catalysts for water oxidation. *Nat. Energy* **2016**, *1*, No. 16053.
- (44) Sayeed, M. A.; Herd, T.; O'Mullane, A. P. Direct electrochemical formation of nanostructured amorphous $\text{Co}(\text{OH})_2$ on gold electrodes with enhanced activity for the oxygen evolution reaction. *J. Mater. Chem. A* **2016**, *4*, 991–999.
- (45) Burke, M. S.; Kast, M. G.; Trotochaud, L.; Smith, A. M.; Boettcher, S. W. Cobalt-Iron (Oxy)hydroxide Oxygen Evolution Electrocatalysts: The Role of Structure and Composition on Activity, Stability, and Mechanism. *J. Am. Chem. Soc.* **2015**, *137*, 3638–3648.
- (46) Chaudhari, N. K.; Jin, H.; Kim, B.; Lee, K. Nanostructured materials on 3D nickel foam as electrocatalysts for water splitting. *Nanoscale* **2017**, *9*, 12231–12247.
- (47) Chen, Y.; Ji, S.; Chen, C.; Peng, Q.; Wang, D.; Li, Y. Single-Atom Catalysts: Synthetic Strategies and Electrochemical Applications. *Joule* **2018**, *2*, 1242–1264.
- (48) Gawande, M. B.; Fornasiero, P.; Zboril, R. Carbon-Based Single-Atom Catalysts for Advanced Applications. *ACS Catal.* **2020**, *10*, 2231–2259.
- (49) Fei, H.; Dong, J.; Chen, D.; Hu, T.; Duan, X.; Shakir, I.; Huang, Y.; Duan, X. Single atom electrocatalysts supported on graphene or graphene-like carbons. *Chem. Soc. Rev.* **2019**, *48*, 5207–5241.
- (50) Flahaut, E.; Bacsa, R.; Peigney, A.; Laurent, C. Gram-scale CCVD synthesis of double-walled carbon nanotubes. *Chem. Commun.* **2003**, 1442–1443.
- (51) Bortolamiol, T.; Lukanov, P.; Galibert, A.-M.; Soula, B.; Lonchambon, P.; Datas, L.; Flahaut, E. Double-walled carbon nanotubes: Quantitative purification assessment, balance between purification and degradation and solution filling as an evidence of opening. *Carbon* **2014**, *78*, 79–90.
- (52) Weibel, A.; Mesguich, D.; Chevallerier, G.; Flahaut, E.; Laurent, C. Fast and easy preparation of few-layered-graphene/magnesia powders for strong, hard and electrically conducting composites. *Carbon* **2018**, *136*, 270–279.
- (53) Flahaut, E.; Peigney, A.; Bacsa, W. S.; Bacsa, R. R.; Laurent, C. CCVD synthesis of carbon nanotubes from (Mg,Co,Mo)O catalysts: influence of the proportions of cobalt and molybdenum. *J. Mater. Chem.* **2004**, *14*, 646–653.
- (54) Inani, H.; Mustonen, K.; Markevich, A.; Ding, E.-X.; Tripathi, M.; Hussain, A.; Mangler, C.; Kauppinen, E. I.; Susi, T.; Kotakoski, J. Silicon Substitution in Nanotubes and Graphene via Intermittent Vacancies. *J. Phys. Chem. C* **2019**, *123*, 13136–13140.
- (55) Mustonen, K.; Markevich, A.; Tripathi, M.; Inani, H.; Ding, E.-X.; Hussain, A.; Mangler, C.; Kauppinen, E. I.; Kotakoski, J.; Susi, T. Electron-Beam Manipulation of Silicon Impurities in Single-Walled Carbon Nanotubes. *Adv. Funct. Mater.* **2019**, *29*, No. 1901327.
- (56) Ortalan, V.; Uzun, A.; Gates, B. C.; Browning, N. D. Direct imaging of single metal atoms and clusters in the pores of dealuminated HY zeolite. *Nat. Nanotechnol.* **2010**, *5*, 506–510.
- (57) Nellist, P. D.; Pennycook, S. J. Direct Imaging of the Atomic Configuration of Ultradispersed Catalysts. *Science* **1996**, *274*, 413–415.
- (58) Jiang, H.; Ruokolainen, J.; Young, N.; Oikawa, T.; Nasibulin, A. G.; Kirkland, A.; Kauppinen, E. I. Performance and early applications of a versatile double aberration-corrected JEOL-2200FS FEG TEM/STEM at Aalto University. *Micron* **2012**, *43*, 545–550.
- (59) Arenal, R.; March, K.; Ewels, C. P.; Rocquefelte, X.; Kociak, M.; Loiseau, A.; Stéphan, O. Atomic Configuration of Nitrogen-Doped Single-Walled Carbon Nanotubes. *Nano Lett.* **2014**, *14*, 5509–5516.
- (60) Willke, P.; Amani, J. A.; Sinterhauf, A.; Thakur, S.; Kotzot, T.; Druga, T.; Weikert, S.; Maiti, K.; Hofsäss, H.; Wenderoth, M. Doping of Graphene by Low-Energy Ion Beam Implantation: Structural, Electronic, and Transport Properties. *Nano Lett.* **2015**, *15*, 5110–5115.
- (61) Egerton, R. F. Beam-Induced Motion of Adatoms in the Transmission Electron Microscope. *Microsc. Microanal.* **2013**, *19*, 479–486.

- (62) Dresselhaus, M. S.; Dresselhaus, G.; Saito, R.; Jorio, A. Raman spectroscopy of carbon nanotubes. *Phys. Rep.* **2005**, *409*, 47–99.
- (63) Dresselhaus, M. S.; Jorio, A.; Hofmann, M.; Dresselhaus, G.; Saito, R. Perspectives on Carbon Nanotubes and Graphene Raman Spectroscopy. *Nano Lett.* **2010**, *10*, 751–758.
- (64) Dresselhaus, M. S.; Jorio, A.; Filho, A. G. S.; Saito, R. Defect characterization in graphene and carbon nanotubes using Raman spectroscopy. *Philos. Trans. R. Soc., A* **2010**, *368*, 5355–5377.
- (65) Donohue, M. D.; Aranovich, G. L. Classification of Gibbs adsorption isotherms. *Adv. Colloid Interface Sci.* **1998**, *76–77*, 137–152.
- (66) Alothman, Z. A. A Review: Fundamental Aspects of Silicate Mesoporous Materials. *Materials* **2012**, *5*, 2874–2902.
- (67) Huang, P.; Li, H.; Huang, X.; Chen, D. Multiheteroatom-Doped Porous Carbon Catalyst for Oxygen Reduction Reaction Prepared using 3D Network of ZIF-8/Polymeric Nanofiber as a Facile-Doping Template. *ACS Appl. Mater. Interfaces* **2017**, *9*, 21083–21088.
- (68) Sharma, M.; Jang, J.-H.; Shin, D. Y.; Kwon, J. A.; Lim, D.-H.; Choi, D.; Sung, H.; Jang, J.; Lee, S.-Y.; Lee, K. Y.; Park, H.-Y.; Jung, N.; Yoo, S. J. Work function-tailored graphene via transition metal encapsulation as a highly active and durable catalyst for the oxygen reduction reaction. *Energy Environ. Sci.* **2019**, *12*, 2200–2211.
- (69) Qiu, K.; Chai, G.; Jiang, C.; Ling, M.; Tang, J.; Guo, Z. Highly Efficient Oxygen Reduction Catalysts by Rational Synthesis of Nanoconfined Maghemite in a Nitrogen-Doped Graphene Framework. *ACS Catal.* **2016**, *6*, 3558–3568.
- (70) Chen, Y.; Ji, S.; Zhao, S.; Chen, W.; Dong, J.; Cheong, W.-C.; Shen, R.; Wen, X.; Zheng, L.; Rykov, A. I.; Cai, S.; Tang, H.; Zhuang, Z.; Chen, C.; Peng, Q.; Wang, D.; Li, Y. Enhanced oxygen reduction with single-atomic-site iron catalysts for a zinc-air battery and hydrogen-air fuel cell. *Nat. Commun.* **2018**, *9*, No. 5422.
- (71) Yan, X.; Yao, Y.; Chen, Y. Highly Active and Stable Fe-N-C Oxygen Reduction Electrocatalysts Derived from Electrospinning and In Situ Pyrolysis. *Nanoscale Res. Lett.* **2018**, *13*, No. 218.
- (72) Sharif, T.; Gracia-Espino, E.; Chen, A.; Hu, G.; Wågberg, T. Oxygen Reduction Reactions on Single- or Few-Atom Discrete Active Sites for Heterogeneous Catalysis. *Adv. Energy Mater.* **2020**, *10*, 1902084.
- (73) Zhou, W.; Huang, D.-D.; Wu, Y.-P.; Zhao, J.; Wu, T.; Zhang, J.; Li, D.-S.; Sun, C.; Feng, P.; Bu, X. Stable Hierarchical Bimetal–Organic Nanostructures as High-Performance Electrocatalysts for the Oxygen Evolution Reaction. *Angew. Chem., Int. Ed.* **2019**, *58*, 4227–4231.
- (74) Tavakkoli, M.; Nosek, M.; Sainio, J.; Davodi, F.; Kallio, T.; Joensuu, P. M.; Laasonen, K. Functionalized Carbon Nanotubes with Ni(II) Bipyridine Complexes as Efficient Catalysts for the Alkaline Oxygen Evolution Reaction. *ACS Catal.* **2017**, *7*, 8033–8041.
- (75) Kauffman, D. R.; Alfonso, D.; Tafeni, D. N.; Lekse, J.; Wang, C.; Deng, X.; Lee, J.; Jang, H.; Lee, J.; Kumar, S.; Matrangola, C. Electrocatalytic Oxygen Evolution with an Atomically Precise Nickel Catalyst. *ACS Catal.* **2016**, *6*, 1225–1234.
- (76) Gao, M.; Sheng, W.; Zhuang, Z.; Fang, Q.; Gu, S.; Jiang, J.; Yan, Y. Efficient Water Oxidation Using Nanostructured α -Nickel-Hydroxide as an Electrocatalyst. *J. Am. Chem. Soc.* **2014**, *136*, 7077–7084.
- (77) Tavakkoli, M.; Kallio, T.; Reynaud, O.; Nasibulin, A. G.; Sainio, J.; Jiang, H.; Kauppinen, E. I.; Laasonen, K. Maghemite nanoparticles decorated on carbon nanotubes as efficient electrocatalysts for the oxygen evolution reaction. *J. Mater. Chem. A* **2016**, *4*, 5216–5222.
- (78) Lei, C.; Lyu, S.; Si, J.; Yang, B.; Li, Z.; Lei, L.; Wen, Z.; Wu, G.; Hou, Y. Nanostructured Carbon Based Heterogeneous Electrocatalysts for Oxygen Evolution Reaction in Alkaline Media. *ChemCatChem* **2019**, *11*, 5855–5874.
- (79) Browne, M. P.; Mills, A. Determining the importance of the electrode support and fabrication method during the initial screening process of an active catalyst for the oxygen evolution reaction. *J. Mater. Chem. A* **2018**, *6*, 14162–14169.
- (80) Wang, L.; Stoerzinger, K. A.; Chang, L.; Zhao, J.; Li, Y.; Tang, C. S.; Yin, X.; Bowden, M. E.; Yang, Z.; Guo, H.; You, L.; Guo, R.; Wang, J.; Ibrahim, K.; Chen, J.; Rusydi, A.; Wang, J.; Chambers, S. A.; Du, Y. Tuning Bifunctional Oxygen Electrocatalysts by Changing the A-Site Rare-Earth Element in Perovskite Nickelates. *Adv. Funct. Mater.* **2018**, *28*, No. 1803712.
- (81) Kirsanova, M. A.; Okatenko, V. D.; Aksyonov, D. A.; Forslund, R. P.; Mefford, J. T.; Stevenson, K. J.; Abakumov, A. M. Bifunctional OER/ORR catalytic activity in the tetrahedral YBaCo₄O_{7.3} oxide. *J. Mater. Chem. A* **2019**, *7*, 330–341.
- (82) Hu, X.; Huang, T.; Tang, Y.; Fu, G.; Lee, J.-M. Three-Dimensional Graphene-Supported Ni₃Fe/Co₉S₈ Composites: Rational Design and Active for Oxygen Reversible Electrocatalysis. *ACS Appl. Mater. Interfaces* **2019**, *11*, 4028–4036.
- (83) Yang, H. B.; Miao, J.; Hung, S.-F.; Chen, J.; Tao, H. B.; Wang, X.; Zhang, L.; Chen, R.; Gao, J.; Chen, H. M.; Dai, L.; Liu, B. Identification of catalytic sites for oxygen reduction and oxygen evolution in N-doped graphene materials: Development of highly efficient metal-free bifunctional electrocatalyst. *Sci. Adv.* **2016**, *2*, No. e1501122.
- (84) Xu, N.; Zhang, Y.; Zhang, T.; Liu, Y.; Qiao, J. Efficient quantum dots anchored nanocomposite for highly active ORR/OER electrocatalyst of advanced metal-air batteries. *Nano Energy* **2019**, *57*, 176–185.
- (85) Bu, Y.; Nam, G.; Kim, S.; Choi, K.; Zhong, Q.; Lee, J.; Qin, Y.; Cho, J.; Kim, G. A Tailored Bifunctional Electrocatalyst: Boosting Oxygen Reduction/Evolution Catalysis via Electron Transfer Between N-Doped Graphene and Perovskite Oxides. *Small* **2018**, *14*, No. 1802767.
- (86) Biesinger, M. C.; Payne, B. P.; Grosvenor, A. P.; Lau, L. W. M.; Gerson, A. R.; Smart, R. S. C. Resolving surface chemical states in XPS analysis of first row transition metals, oxides and hydroxides: Cr, Mn, Fe, Co and Ni. *Appl. Surf. Sci.* **2011**, *257*, 2717–2730.
- (87) Baltrusaitis, J.; Mendoza-Sanchez, B.; Fernandez, V.; Veenstra, R.; Dukstiene, N.; Roberts, A.; Fairley, N. Generalized molybdenum oxide surface chemical state XPS determination via informed amorphous sample model. *Appl. Surf. Sci.* **2015**, *326*, 151–161.
- (88) Povey, A. F.; Metcalfe, A. A. The anodic dissolution of molybdenum in alkaline solutions — X-ray photoelectron spectroscopic studies. *J. Electroanal. Chem. Interfacial Electrochem.* **1977**, *84*, 73–81.
- (89) Susi, T.; Pichler, T.; Ayala, P. X-ray photoelectron spectroscopy of graphitic carbon nanomaterials doped with heteroatoms. *Beilstein J. Nanotechnol.* **2015**, *6*, 177–192.
- (90) Chen, S.; Bi, J.; Zhao, Y.; Yang, L.; Zhang, C.; Ma, Y.; Wu, Q.; Wang, X.; Hu, Z. Nitrogen-Doped Carbon Nanocages as Efficient Metal-Free Electrocatalysts for Oxygen Reduction Reaction. *Adv. Mater.* **2012**, *24*, 5593–5597.
- (91) Lin, T.-E.; Cortés-Salazar, F.; Lesch, A.; Qiao, L.; Bondarenko, A.; Girault, H. H. Multiple scanning electrochemical microscopy mapping of tyrosinase in micro-contact printed fruit samples on polyvinylidene fluoride membrane. *Electrochim. Acta* **2015**, *179*, 57–64.
- (92) Toikkanen, O.; Lähteenmäki, M.; Moisio, T.; Forssell, P.; Partanen, R.; Murtomäki, L. Study of Oxygen Transfer across Milk Proteins at an Air–Water Interface with Scanning Electrochemical Microscopy. *J. Agric. Food. Chem.* **2014**, *62*, 2284–2288.
- (93) Slevin, C. J.; Ryley, S.; Walton, D. J.; Unwin, P. R. A New Approach for Measuring the Effect of a Monolayer on Molecular Transfer across an Air/Water Interface Using Scanning Electrochemical Microscopy. *Langmuir* **1998**, *14*, 5331–5334.
- (94) Cortés-Salazar, F.; Deng, H.; Peljo, P.; Pereira, C. M.; Kontturi, K.; Girault, H. H. Polyethylene C coated microelectrodes for scanning electrochemical microscopy. *Electrochim. Acta* **2013**, *110*, 22–29.
- (95) Haensch, M.; Balboa, L.; Graf, M.; Silva Olaya, A. R.; Weissmüller, J.; Wittstock, G. Mass Transport in Porous Electrodes Studied by Scanning Electrochemical Microscopy: Example of Nanoporous Gold. *ChemElectroChem* **2019**, *6*, 3160–3166.

(96) Snook, G. A.; Duffy, N. W.; Pandolfo, A. G. Detection of Oxygen Evolution from Nickel Hydroxide Electrodes Using Scanning Electrochemical Microscopy. *J. Electrochem. Soc.* **2008**, *155*, A262–A267.