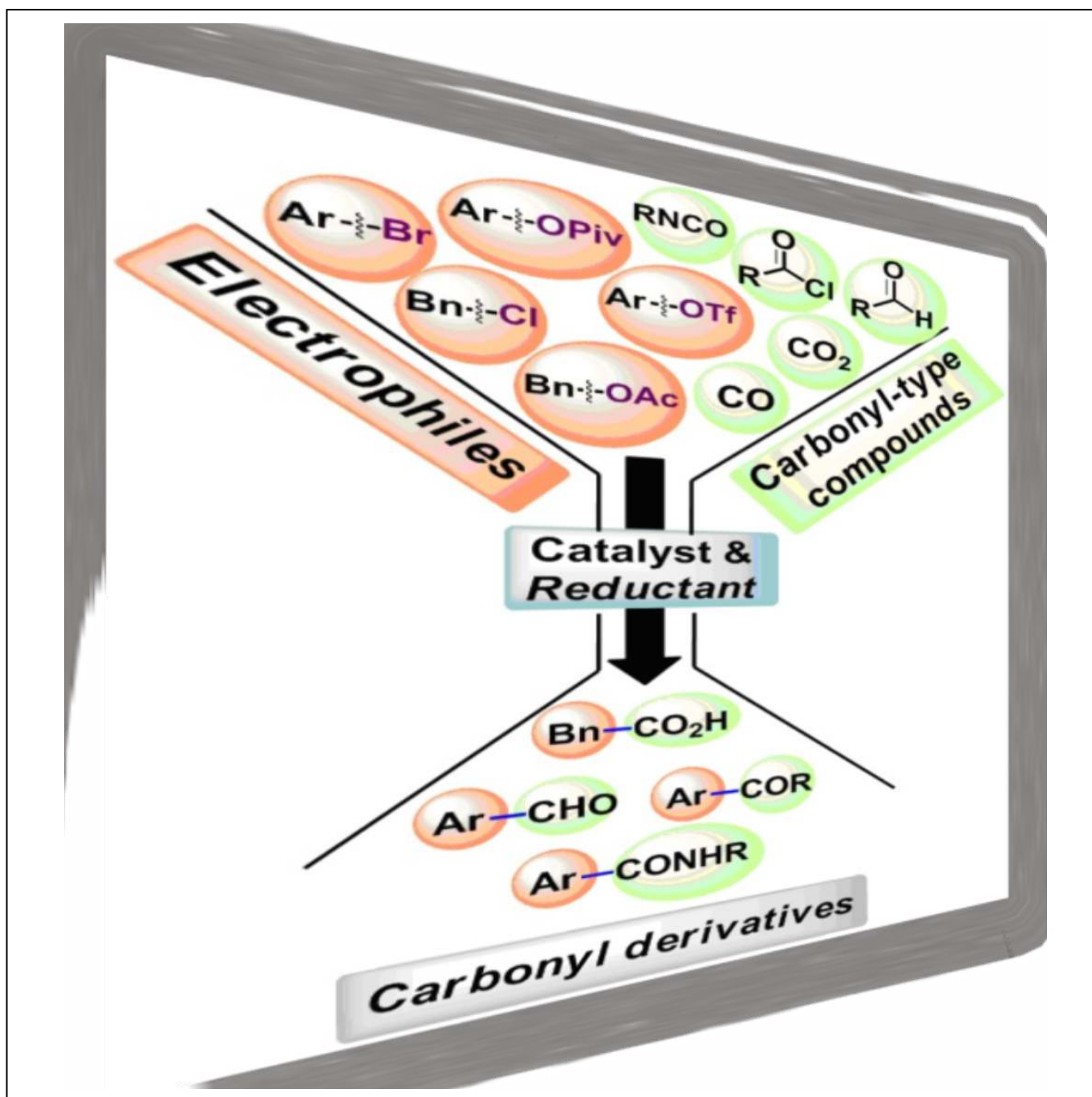


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Metal-catalyzed Reductive Coupling Reactions of Organic Halides with Carbonyl-type Compounds

Toni Moragas^[a], Arkaitz Correa^{*[a]} and Ruben Martin^{*[a,b]}

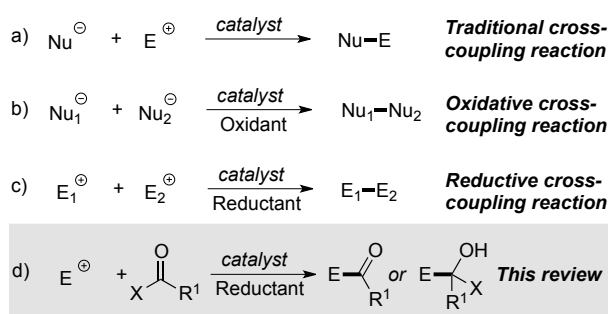


Introduction

Abstract: Metal-catalyzed reductive couplings of aryl halides and (pseudo)halides with carbonyl-type compounds have undergone impressive development along the last years. These methodologies have shown to be a powerful alternate strategy, practicality aside, to the use of stoichiometric, well-

defined and, in some cases, air-sensitive organometallic species. In this review, the recent findings in this field are summarized, with particular emphasis on the mechanistic interpretation of the results and the future aspects of this area of expertise.

Metal-catalyzed cross-coupling reactions have undoubtedly changed the landscape of organic synthesis, allowing for the design of innovative, yet practical, tactics for the formation of carbon-carbon and carbon-heteroatom bonds.^[1] Importantly, these technologies have found immediate application when preparing agrochemicals, polymers and pharmaceuticals, among others. Indeed, it is rather difficult to find any advanced total synthesis of target molecules of considerable complexity that do not use a metal-catalyzed cross-coupling event. The underlying mechanisms of such "classical" transformations tacitly suggest that nucleophiles react with electrophilic counterparts assisted by a transition metal (Scheme 1, path a). Thus, a myriad of different cross-coupling methodologies have been reported under such premise.^[1] The success of these "classical" cross-coupling reactions has contributed to the perception that we have all the methods that we need to build up molecular complexity. Fortunately, scepticism, uncertainty and, in particular, curiosity has always been the main driving force behind scientific activity. Indeed, the number of possible combinations of two or more substances in the presence of a certain catalyst is virtually astronomical, indicating that cross-coupling methodologies cannot be simply limited to a simple combination of a nucleophile with an electrophilic partner. Such a vision has inspired chemists to design new techniques within the cross-coupling arena by using unconventional substrate combinations, thus increasing our ever-growing organic chemistry portfolio. Among them, the metal-catalyzed oxidative cross-coupling reaction of two different carbon nucleophiles has shown to be a powerful alternative when designing new carbon-carbon bond-forming reactions (Scheme 1, path b).^[2] Unfortunately, however, the air sensitivity and low chemoselectivity profile of Grignard reagents, organozinc or organoboron species, together with their low accessibility limit the application profile of these elegant processes.



Scheme 1. Classical cross-coupling reactions vs unconventional oxidative or reductive coupling events.

Commonly employed organometallic species in cross-coupling reactions are typically prepared from the corresponding organic halides.^[3] Therefore, the ability to promote a catalytic cross-coupling event of organic halides with an additional electrophilic partner would be highly appreciated (Scheme 1, path c). The advantages for developing such a route are primarily associated to the readily availability of organic halides and the avoidance of stoichiometric and highly reactive organometallic species. However, the inherent low reactivity of electrophiles and the proclivity towards undesired pathways constitute serious drawbacks to be overcome. The recent years have witnessed the development of new metal-catalyzed reductive coupling reactions of organic halides, achieving a high degree of molecular complexity while increasing the applicability of these synthetically attractive transformations.^[4,5] Although different electrophilic counterparts can be utilized with remarkable efficiency,^[5] the use of carbonyl-type compounds is particularly attractive since these motifs rank amongst the most versatile synthons in organic synthesis.^[3] Therefore, the utilization of such motifs would allow the rapid and straightforward synthesis of advanced intermediates possessing the carbonyl unit, increasing the flexibility in synthetic design (Scheme 1, path d).^[6] While realizing the preparative potential of such reductive coupling events, we identified a need to review the most recent advances in this vibrant field of expertise when utilizing organic halides and carbonyl-type compounds, including the use of CO or even CO₂ as coupling partners. The latter is particularly interesting as the use of CO₂ in reductive cleavage events could potentially revolutionize the means to access carboxylic acids in a direct fashion from simple renewable carbon feedstock. Although formidable advances have been reported within the field of electrochemistry,^[7] the purpose of this review is to primarily focus on the most recent and important developments in metal-catalyzed reductive coupling reactions with carbonyl-type compounds over the last years not promoted by

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electrochemical methods, including synthetic applications and mechanistic aspects, when applicable.

Toni Moragas Solà was born in 1984 in Barcelona (Spain). He conducted his undergraduate studies at the University of Barcelona, obtaining his M.Sc. in 2009. He then moved to the United Kingdom to carry out his PhD studies under the guidance of Dr. Robert Stockman at the University of Nottingham. In 2013, he moved to his current position as a postdoctoral fellow in the laboratories of Prof. Ruben Martin at ICIQ (Tarragona). His research interests are primarily focused on the field of metal-catalyzed carbon dioxide fixation and C–O bond cleavage processes.



Arkaitz Correa studied chemistry at the University of the Basque Country (UPV-EHU). He completed his PhD studies in 2006 under the guidance of Prof. Esther Domínguez. Along that time he did a short stay with Prof. Ben L. Feringa at University of Groningen. In 2007 he undertook his first postdoctoral studies with Prof. Carsten Bolm at RWTH-Aachen University (2007-2008). Subsequently, he joined the group of Prof. Ruben Martin at ICIQ (Tarragona) as a postdoctoral fellow (2008-2010). In 2010 he started a collaborative project with Bayer CropScience at the CSIC under the supervision of Prof. J. M. Lassaletta. In April 2011, he returned to the group of Prof. Ruben Martin at ICIQ as Juan de la Cierva researcher. Recently, he has been awarded with a Ramón y Cajal research contract to continue his research career at UPV-EHU in the Basque Country.

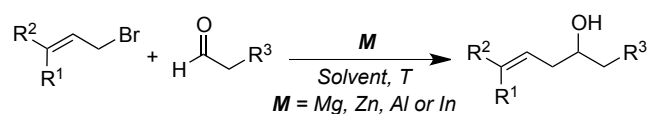


Ruben Martin conducted his undergraduate studies at the University of Barcelona. He then received his PhD in 2003 from the University of Barcelona under the guidance of Prof. Antoni Riera. In 2004, he moved to the Max-Planck Institut für Kohlenforschung as a Humboldt postdoctoral fellow with Prof. Alois Fürstner. In 2005 he then undertook further postdoctoral studies at MIT with Prof. Stephen L. Buchwald as a MEC-Fulbright fellow. In September 2008 he initiated his independent career as an assistant professor at the ICIQ (Tarragona). In July 2013 he was promoted to associate professor and subsequently to ICREA Research Professor. His research is focused on the development of new metal-catalyzed activation of inert bonds. His work has been rewarded, among others, by the RSEQ Sigma-Aldrich prize for Outstanding Emerging Investigators (2010), Thieme Chemistry Award (2011), Eli Lilly Young Investigator Award (2011), and he currently holds a European Research Council (ERC) Starting Grant (2011).



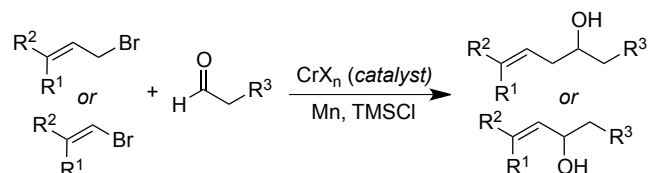
Reductive Coupling Reactions of Unactivated Organic Halides with Carbonyl-type Compounds

The addition of organolithium or Grignard reagents, among others, to carbonyl compounds has evolved to a mature tool in organic synthesis.^[8] While no doubt robust synthetic tools, these methodologies are still limited to the use of stoichiometric amounts of well-defined, and in some cases air- and moisture-sensitive organometallic species that are ultimately prepared from the corresponding organic halides. From an ideal perspective, the crucial organometallic species should be generated *in situ* without the need for using sensitive reagents. Therefore, the means to effect a reductive coupling reaction of organic halides with carbonyl compounds in the presence of a reducing agent will present a considerable amount of attractive features, both from an experimental ease and step-economical standpoint. The successful development of the Barbier reaction could be visualized as the first step towards achieving such goal in which an activated alkyl halide reacts with a carbonyl compound in the presence of simple, inexpensive and water insensitive Mg, Zn, Al or In, among others.^[9]



Scheme 2. General scheme for the Barbier reaction of allyl electrophiles.

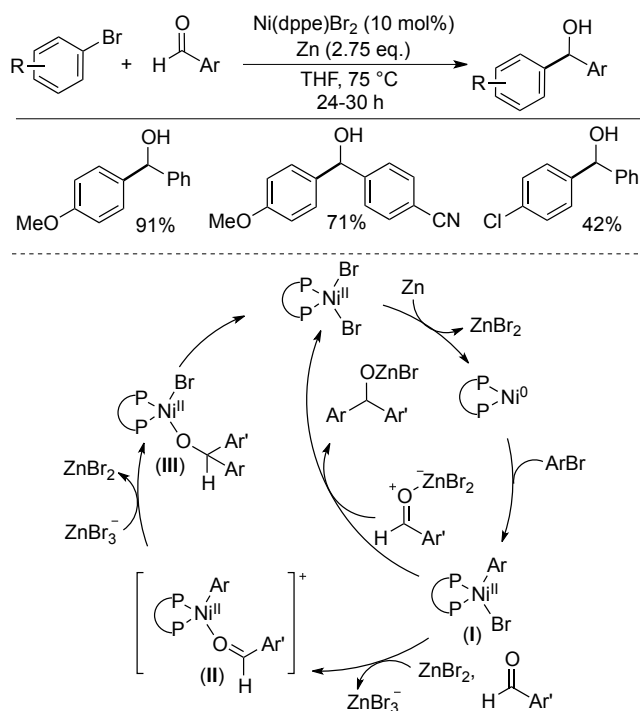
A considerable step forward in Barbier-type reactions was the development of the mild and chemoselective *Nozaki-Hiyama-Kishi* (NHK) reaction catalytic in Cr,^[10] in which activated vinyl or allyl (pseudo)halides react with carbonyl compounds, typically aldehydes, with catalytic amounts of CrX_n, Mn and in the presence of an activating reagent, typically TMSCl. The importance of this reaction is manifested by its prolific use in total synthesis, such as in the impressive total synthesis of palytoxin by Kishi,^[11] as well as in pharmaceutical industry. Interestingly, this reaction can be even performed in an enantioselective fashion,^[12] providing access to valuable synthetic intermediates from simple precursors.



Scheme 3. General scheme for the Cr-catalyzed NHK reaction.

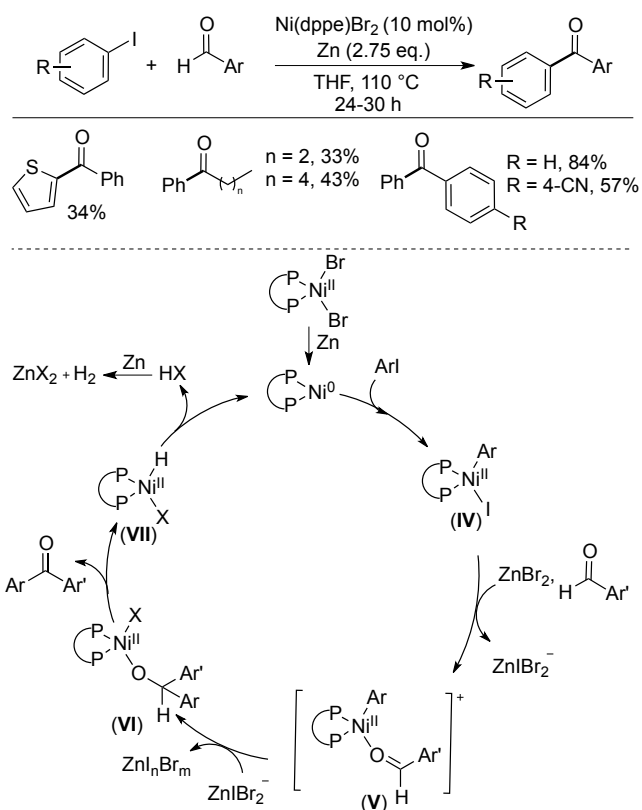
While Barbier-type reactions, including the NHK reaction, have undoubtedly impacted the field of organic chemistry, these reactions are primarily limited to counterparts with a weak C–X bond and particularly reactive carbonyl compounds. Therefore, the extension to less activated counterparts have been considered a quest for organic chemists. In line with this notion, Majumdar and Cheng developed a Ni-catalyzed reaction between aryl bromides and aldehydes under reductive conditions to afford secondary alcohols (Scheme 4).^[13] Generally, aromatic carbinols can be obtained in high yields and with good chemoselectivity profile employing the NiBr₂(dppe)/Zn couple. Unfortunately, however, the reaction was limited to the use of particularly activated aromatic aldehydes. While the temperature proved to be important for improving the selectivity of the process, the nature of the ligand played a critical role for success; indeed, the reaction only took place at some extent when bidentate phosphines or

bisoxazolines were employed as ligands. Unfortunately, no stereoselectivity was observed with the use of chiral bisoxazolines as ligands. The authors suggested a mechanism consisting of an initial Ni(II) to Ni(0) reduction mediated by Zn followed by an oxidative addition into the C–Br bond (I). Subsequently, it was postulated that the ZnBr₂ generated *in situ* in the initial reduction step acted as a halide scavenger, thus opening up a open coordination site and setting up the stage for binding the aldehyde motif into the metal center (II).^[14]



Scheme 4. Ni-catalyzed reductive coupling of aryl bromides with benzaldehydes.

Shortly after, the Cheng group reported a similar procedure for the synthesis of aromatic ketones.^[15] Intriguingly, the exact catalyst system previously employed for aryl bromides^[13] affords aromatic ketones when aryl iodides are employed as substrates at higher temperatures (Scheme 5). In this case, the generality of the reaction was expanded to the use of heterocyclic and aliphatic aldehydes as well. While a similar sequence of events was postulated when employing aryl iodides, the nickel alkoxide **VI** was believed to undergo β -hydride elimination to give the aryl ketone and a nickel hydride which, upon reductive elimination recovered the catalytically active Ni(0) species. Although Zn was suggested to act as an acid scavenger affording ZnX₂ and H₂, no H₂ evolution was detected in the reaction mixture. Interestingly, gas release could be observed in the presence of a Zn/ZnI₂ couple.

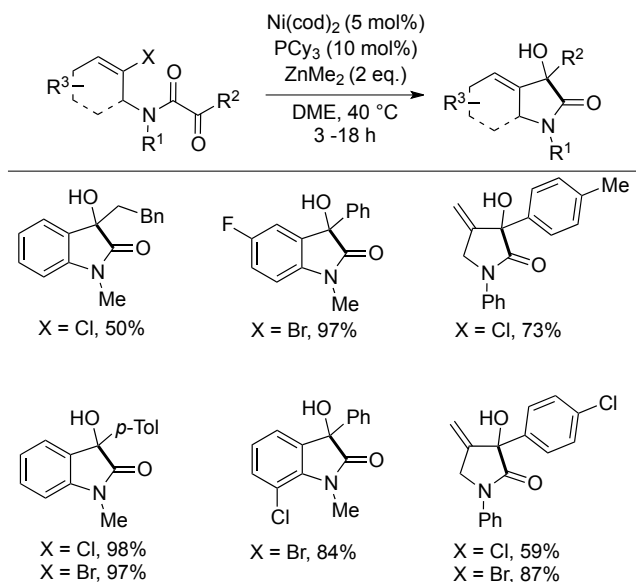


Scheme 5. Ni-catalyzed reductive coupling of aryl iodides with benzaldehydes.

In 2007, Durandetti and co-workers expanded the scope of metal-catalyzed reductive coupling of aryl halides beyond the use of aldehydic counterparts for the synthesis of β -hydroxyesters in a Reformatsky-type reaction.^[16] The authors demonstrated that ketones could also be employed as coupling partners by using nickel catalysis and manganese as reducing agent. More recently, Jia and Gao discovered a Ni-catalyzed intramolecular addition of aryl or vinyl halides to ketamides for preparing 3-hydroxyoxindoles (Scheme 6).^[17] In sharp contrast with Cheng's observations,^[13,15] monodentate and rather bulky PCy₃ provided the best results, with ZnMe₂ being used as the reducing agent. Unlike Chen's procedure,^[13,15] the method allowed for the coupling of rather challenging aryl chlorides, an observation that might be tentatively attributed to the use of bulky and electron-rich PCy₃.^[14] As for many other reductive catalytic events,^[5] the need for Zn-reagents might suggest the intermediacy of organozinc species that are responsible for the observed reactivity. Control experiments, however, evidenced that arylzinc species were not competent as reaction intermediates. This observation was fostered by the non-negligible enantioselectivity observed (up to 37% e.e.) when a chiral phosphoramidite ligand was used.

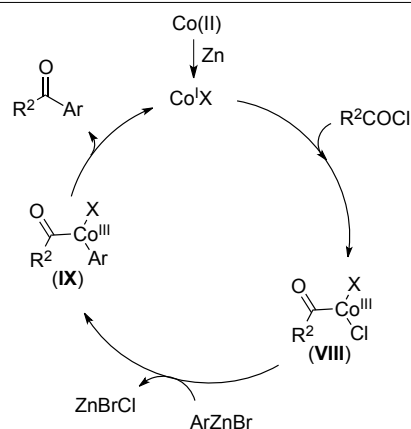
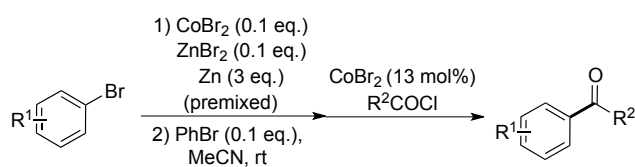
While aldehydes and ketones can be successfully coupled in catalytic reductive coupling reactions with great efficiency, other types of carbonyl-type compounds have also proven to be suitable substrates. The considerably higher reactivity of acid chlorides as compared with other carbonyl compounds makes them ideal coupling partners in reductive coupling events. Gosmini *et al.* discovered that acid chlorides could be utilized, for the first time, in reductive coupling events of aryl bromides mediated by Zn and in the presence of catalytic amounts of ZnBr₂ and CoBr₂ for preparing aromatic ketones (Scheme 7).^[18] The suggested mechanism is based on the ability of Co(I) species to trigger an initial oxidative addition to the acyl–Cl

bond in order to produce a rather elusive Co(III) intermediate (**VIII**). This proposed mechanism is somewhat reminiscent of a number of elegant electrochemical methods reported in the literature for similar means.^[19] A final transmetalation (**IX**) and reductive elimination was postulated to afford the desired product while recovering the catalytically active Co(I) species.^[20]

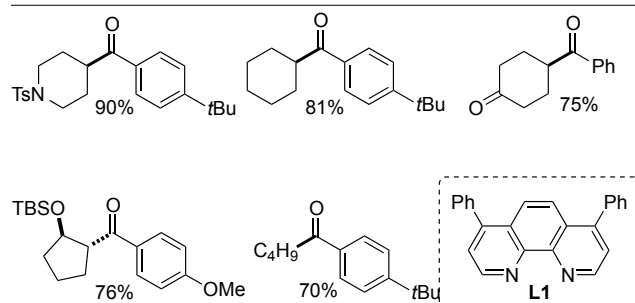
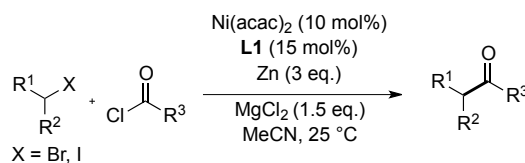


Scheme 6. Ni-catalyzed intramolecular reductive coupling of aryl bromides and chlorides with ketamides.

Although no doubt innovative, this methodology requires the use of stoichiometric amounts of organometallic, and pyrophoric compounds, hence significantly limiting its synthetic utility. A more direct reductive carbonylation using acid chlorides has been recently developed by Gong and co-workers in which *unactivated* alkyl halides could be used as coupling partners (Scheme 8).^[21] Primary and sterically-encumbered secondary alkyl iodides were efficiently coupled with Ni precatalysts in combination with phenantroline-type ligands (**L1**). Not surprisingly, much less reactive alkyl bromides were not effective, obtaining the expected product in moderate yields. The utility of this methodology was expanded with the development of a *one-pot* reaction from benzoic acids to afford differently substituted aromatic ketones, thus constituting a powerful alternative to the use of organolithium derivatives for similar purposes. Furthermore, the reaction could easily be scaled up without compromising the yield.^[22] Control experiments determined that formation of organozinc intermediates was kinetically disfavoured and that the reaction takes place *via* a distinctive mechanistic scenario. Given the known ability of Ni catalysts to promote Negishi-type couplings,^[23] the authors suggested that the mechanism was based on a Ni(I)/Ni(III) couple.^[24] In this manner, it was believed that Ni(I) triggers the formation of alkyl radicals with concomitant formation of Ni(II) species. A radical recombination afforded Ni(III) species that were subsequently reduced to Ni(I) by Mn. Oxidative addition of the latter to the acid chloride motif and reductive elimination yielded the desired ketone while regenerating the initial Ni(I) species.



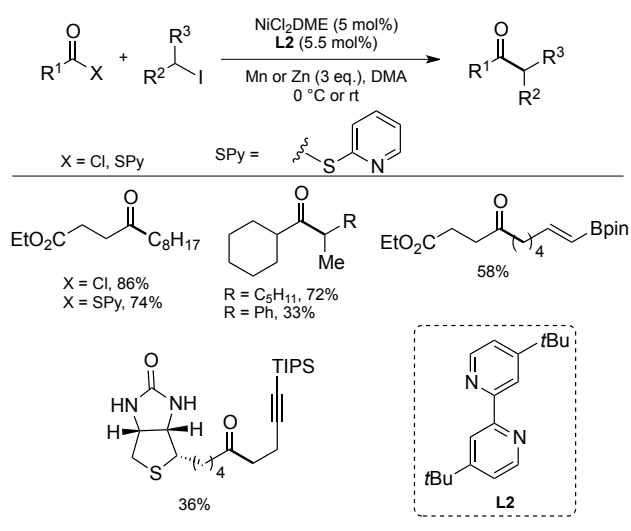
Scheme 7. Mechanism for the Co-catalyzed reductive coupling of aryl bromides with acid chlorides *via* the *in situ* generation of organozinc species.



Scheme 8. Ni-catalyzed reductive coupling of alkyl halides with acyl chlorides.

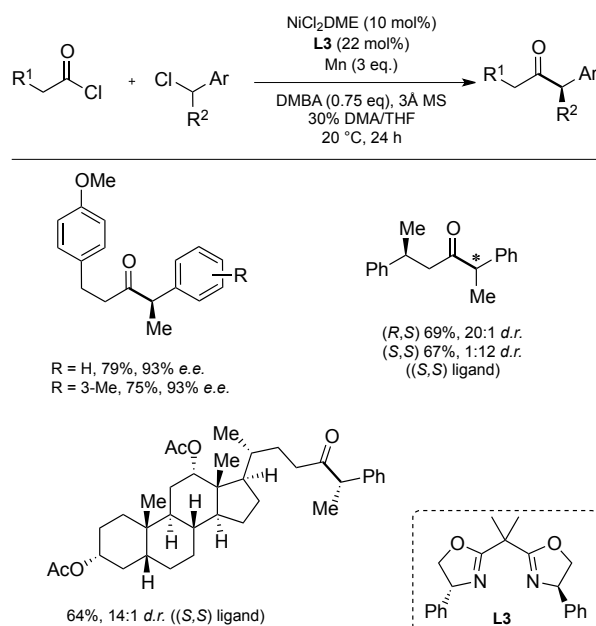
A more suitable methodology for the synthesis of dialkyl ketones was developed by Weix and co-workers using Ni catalysts (Scheme 9).^[25] In this case, 4,4'-di-*tert*-butyl-2,2'-bipyridine (dtbpy, **L2**) allowed for the coupling of primary and secondary alkyl iodides with acid chlorides or thioesters (Scheme 9). The reaction displayed a remarkable functional group tolerance and even highly hindered ketones could be obtained in good yields. However, low yields were usually obtained with aromatic acid chlorides, showing the limitations of this strategy. A slightly different mechanism was proposed for this reaction. Prompted by the work of Yamamoto,^[26] the authors hypothesized that the *in situ* generated Ni(II) oxidative addition species (**X**) could disproportionate to give intermediates **XI** and **XII**. While the reduction of **XII** promoted by Mn would regenerate the initial Ni(0) species, intermediate **XI** could transfer the alkyl residue into the acid chloride motif. It is worth noting that such mechanistic scenario is significantly different from commonly proposed pathways within the realm of Ni-catalyzed reductive cleavage reactions.^[5] Unfortunately, however, an *in depth* mechanistic study was not performed to lend support for disproportionation events within the catalytic

cycle. Although tentative, we speculate that such scenarios might come into play for related reductive cross-coupling events.



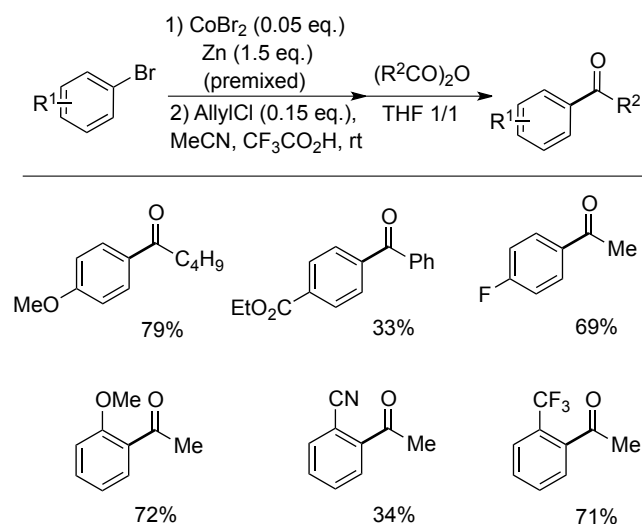
Scheme 9. Ni-catalyzed cross-coupling reaction of acid chlorides with secondary alkyl halides.

Reisman and co-workers recently reported an excellent contribution to the field of metal-catalyzed reductive cleavage of organic halides and carbonyl-type compounds (Scheme 10).^[27] Particularly, the authors designed an enantioconvergent reductive coupling of acid chlorides with activated benzyl halides. While a variety of chiral ligands were tested, the use of a chiral bisoxazoline ligand (**L3**) provided the best results with enantioselectivities up to 94% and with an excellent chemoselectivity profile. Interestingly, it was revealed that the inclusion of additives such as benzoic acids was crucial; in line with the same notion, a delicate balance between reactivity and stereoselectivity can be attained with the use of mixed solvent mixtures. The mechanism of the reaction was believed to proceed *via* a Ni(I)/Ni(III) couple (Scheme 10). After initial oxidative addition into the acid chloride motif, a reduction promoted by Mn would result in **XIV**.^[24] Subsequently, an oxidative addition into the C–Cl bond of the benzylic framework *via* a stereoconvergent radical recombination would afford **XVI**, setting up the stage for a reductive elimination en route to the desired coupling product. A final reduction promoted by Mn would regenerate the active Ni(0) species. Although not contemplated by the authors, one might have envisioned an alternate pathway consisting of the *in situ* formation of ketenes that would react with the corresponding alkyl halide in a reductive cross-coupling event.



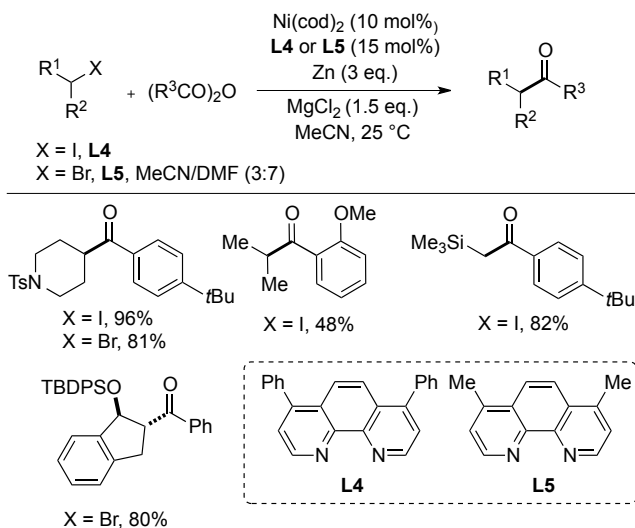
Scheme 10. Ni-catalyzed enantioselective reductive coupling of benzyl chlorides with acid chlorides.

Following their interest in Co-catalyzed reductive cross-coupling reactions,^[18] Gosmini and co-workers discovered that a range of aryl bromides could also be efficiently coupled with acid anhydrides.^[28] In contrast with the use of acid chlorides (Scheme 7),^[18] the diminished reactivity of acid anhydrides allowed the reaction to be performed in one pot (Scheme 11). In this case, allyl chloride was required as additive to enhance the formation of the organozinc species and reduce the formation of undesired byproducts. Control experiments demonstrated that CoBr₂ might serve as a precatalyst en route to the catalytically active Co(I) species.^[24] While the efficiency of the reaction was strongly dependent on the nature of the anhydride (with alkyl substituents giving the best results), the substitution pattern on the aromatic bromide also had an effect on the reaction outcome, an observation that was corroborated by a Hammett study.



Scheme 11. Co-catalyzed reductive coupling reaction of aryl bromides with acid anhydrides via the *in situ* generation of organozinc species.

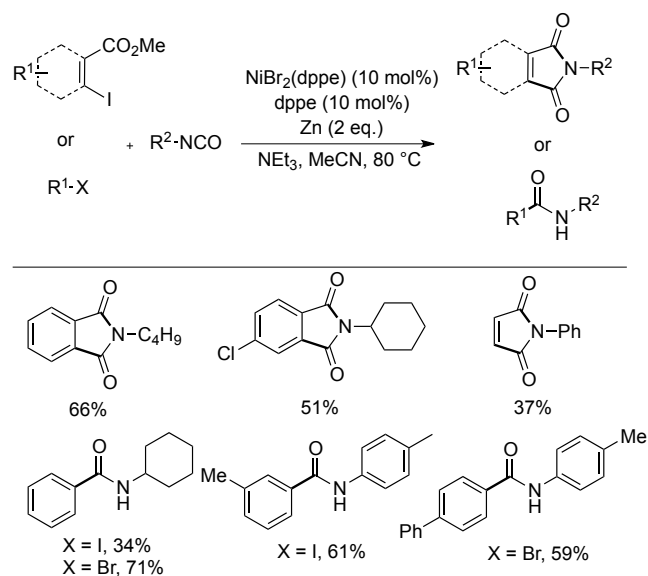
More recently, the Ni-catalyzed reductive coupling reaction of acid anhydrides was extended to the use of *unactivated* alkyl halides as coupling partners.^[29] Lin and Gong showed that alkyl iodides and alkyl bromides could be efficiently coupled with a variety of acid anhydrides using Ni(cod)₂ as the catalyst in combination with phenanthroline-type ligands (**L4** and **L5**), MgCl₂ as an additive and Zn as reducing agent in acetonitrile at room temperature (Scheme 12). The MgCl₂ was believed to play with dual roles, both as a Lewis acid activating the anhydride motif and as a salt scavenger in the Zn surface. Although the reaction was general for both primary and secondary alkyl halides, it was found that cyclic secondary alkyl halides gave the best results. Interestingly, good yields were obtained regardless of the nature of the anhydride employed. Furthermore, the authors demonstrated the practicality of the method by generating the acid anhydride *in situ* from the corresponding carboxylic acid and Boc anhydride, albeit larger quantities of the latter were required. Although initial experiments demonstrated that a Negishi-type pathway *via* the intermediacy of organozinc species may be operative using alkyl iodides, the authors did not support such scenario when alkyl bromides were employed. In this case, it was postulated that the reaction proceeded *via* ArCONi(II) species that subsequently reacted with the corresponding alkyl bromide.



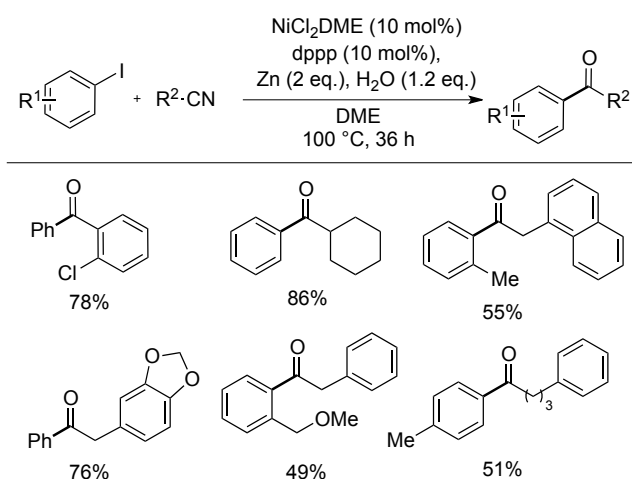
Scheme 12. Ni-catalyzed reductive coupling of unactivated alkyl halides with acid anhydrides.

In 2005, the use of isocyanates as coupling partners under catalytic reductive conditions was developed by Hsieh and Cheng for the formation of imide and amide derivatives.^[30] The catalytic protocol was based on the use of Ni(dppe)Br₂ employing Zn as final reductant (Scheme 13). While the presence of catalytic amounts of NEt₃ was found to accelerate the reaction, the inclusion of dppe significantly decreased the amounts of homocoupling products in the intramolecular approach. While *in depth* mechanistic studies were not performed, the authors favoured a pathway consisting of an oxidative addition of a Ni(0) to the corresponding iodoarene followed by a migratory insertion into the isocyanate motif. Intriguingly, the intermolecular reaction was better conducted with aryl bromides in the absence of NEt₃ since iodobenzene derivatives primarily resulted in homocoupling products.

Prompted by a seminal work of Larock using Pd catalysts,^[31] Hsieh *et al.* described an intermolecular catalytic coupling of aryl halides with aromatic and aliphatic nitrile derivatives using Ni catalysts and Zn as reductant (Scheme 14).^[32] Optimization studies revealed that the bidentate nature of the phosphine ligand utilized and the Ni source were absolutely crucial for success, with dppp and NiCl₂·DME providing the best results. However, the presence of electron-donating or electron-withdrawing groups on the aryl iodide had a negative impact on reactivity. Likewise, the inclusion of *ortho*-substitution significantly lower the efficiency of the process, thus showing the inherent limitations of this method. Although the use of tertiary aliphatic nitriles did not lead to the formation of products, a wide range of secondary nitriles could be successfully employed using this approach. A plausible mechanism involved an oxidative addition of an *in situ* generated Ni(0) into the C–I bond of the iodoarene, followed by coordination and insertion of the cyano group. The resulting cationic species might undergo reductive elimination en route to an iminonickel(II) intermediate. A final hydrolysis delivered the product and recovered back the Ni(II) precatalyst.

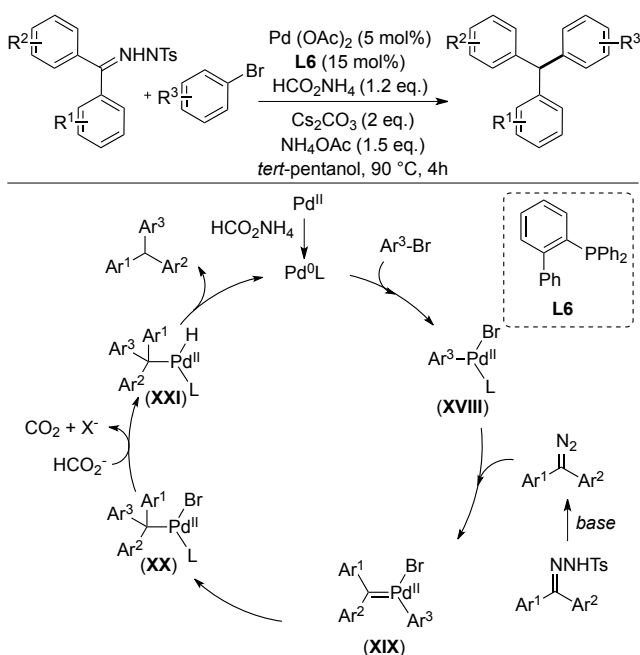


Scheme 13. Ni-catalyzed reductive coupling of haloarenes with isocyanates.



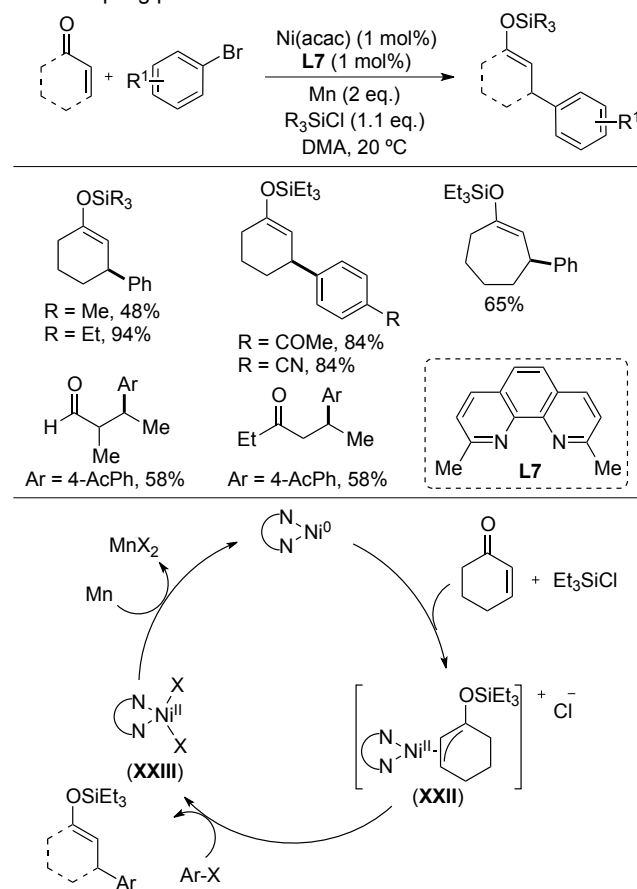
Scheme 14. Ni-catalyzed reductive coupling of iodoarenes with nitriles.

Zhang, Wang and co-workers have recently established an unconventional strategy for the reductive cross-coupling reaction of aryl bromides with *N*-tosylhydrazones to prepare synthetically relevant triarylmethane derivatives (Scheme 15).^[33] This approach required the use of Pd catalysts in combination with ammonium formate as reducing agent. The use of a phosphine ligand (**L6**), *tert*-pentanol as solvent, and ammonium acetate as additive provided the best results while minimizing the formation of reduced arenes. Under these optimized conditions, a wide range of differently substituted *N*-tosylhydrazones and aryl bromides were tolerated, and only the introduction of heteroaromatic substituents or *ortho*-substituted aryl bromides hampered the reaction. The mechanism suggested for this transformation was believed to proceed *via* initial reduction of Pd(OAc)₂ to Pd(0) promoted by ammonium formate followed by oxidative addition to the aryl bromide to afford an arylpalladium(II) intermediate (**XVIII**). The latter reacts with a diazo compound formed *in situ* to deliver a palladium (II) carbene species (**XIX**). A subsequent [1,2]-aryl migration into the palladium carbene afforded an intermediate **XX** that reacted with formate to release CO₂ with concomitant formation of a Pd(II) hydride **XXI** that ultimately resulted in the final product by a final reductive elimination event.

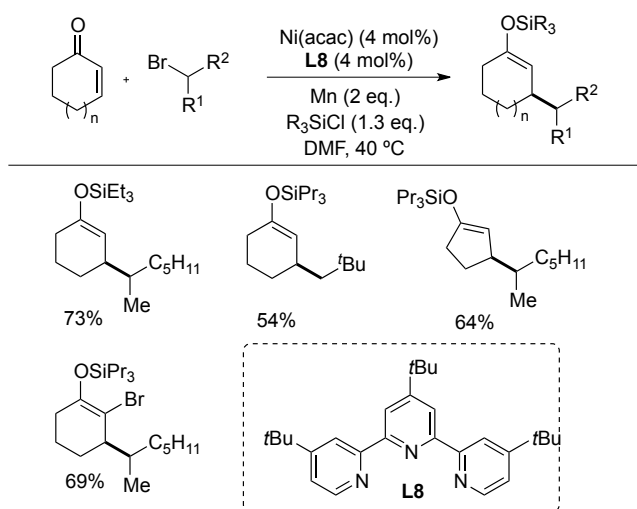


Scheme 15. Mechanism of the Pd-catalyzed reductive synthesis of triarylmethanes.

Recently, Weix and co-workers reported an extensive investigation on the reductive cross-coupling reaction of aryl halides with enones as electrophiles to afford substituted silyl enol ethers (Scheme 16).^[34] A broad screening of ligands revealed that the use of neocuproine (**L7**) or bipyridine allows the use of aromatic halides as electrophilic coupling partners with cyclic and non-cyclic enones. The reaction proceeds with low catalyst loadings (1 mol%), a remarkable finding taking into account the typical high loadings required in Ni-catalyzed reductive cleavage events.^[5] Several chlorosilanes were also tested, with TESCi giving the best results. The authors favored a mechanism consisting of an oxidative addition into the enone backbone (**XXII**) followed by reaction with the corresponding aromatic halide to furnish the desired silyl enol ether and a Ni(II) species (**XXIII**) that is further reduced to the propagating Ni(0) species by Mn. While the previous methodology was robust enough to accommodate a series of aryl halides, the extension of alkyl halides has only been accomplished recently by the Weix's group.^[35] In this case, the use of a bulky terpyridine (**L8**) was crucial for the minimization of dimer adducts. The reaction proceeds with moderate-to-good yields and is general for different chlorosilanes. Likewise, secondary and even tertiary unactivated iodo- and bromoalkanes could be utilized. Furthermore, the authors demonstrated that this reaction could be carried out in gram-scale (up to 10 mmol). Unfortunately, the use of chloroalkanes resulted in the formation of the homocoupling product.



Scheme 16. Ni-catalyzed conjugate addition of aromatic halides to enones.

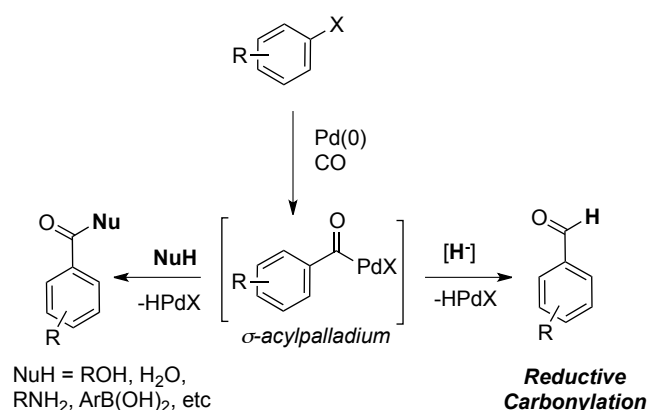


Scheme 17. Ni-catalyzed conjugate addition of secondary halides to enones.

Beyond any reasonable doubt, the recent developments in the area of metal-catalyzed reductive coupling reactions of organic halides with carbonyl-type compounds will set the standards for the discovery of more challenging scenarios within this field of expertise. Despite the advances realized, the means to promote asymmetric transformations *via* metal-catalyzed reductive coupling events has virtually been unexplored with the exception of the elegant work by Reisman.^[27] Therefore, it is tempting to speculate that further improvements in this area will probably be focused within the realm of asymmetric catalysis, thus opening up new horizons to be implemented in this relatively unexplored area of research.

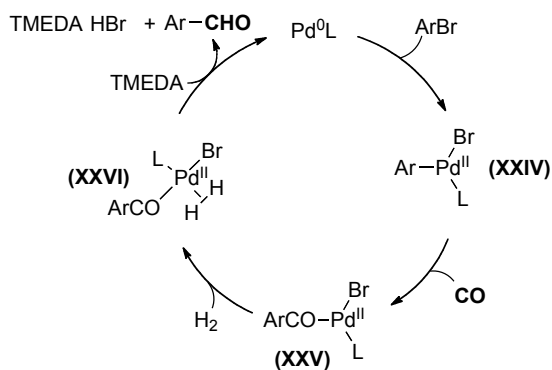
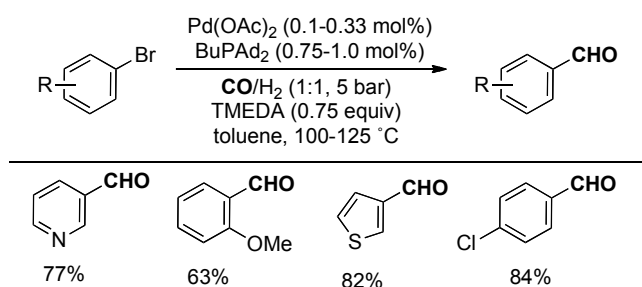
Reductive Carbonylation Reactions

Carbonylation reactions with gaseous carbon monoxide (CO) are key industrial processes for the atom-efficient introduction of C1 units into organic backbones en route to valuable synthetic intermediates such as aldehydes, alcohols and carboxylic acid derivatives, compounds of utmost importance in pharmaceuticals and agrochemicals.^[36] Over the last years, a wide variety of *classical* palladium-catalyzed carbonylation processes have been described using organic halides as substrates en route to aryl ketone or benzamides, allowing the use of CO as a readily available feedstock.^[36] These reactions follow a common mechanistic scenario consisting of the intermediacy of σ -acylpalladium species resulting from a CO insertion into σ -arylpalladium species (Scheme 18). Subsequently, such intermediates are intercepted by an *in situ* treatment with different nucleophiles such as alcohols, amines or even carbon nucleophiles such as boronic acids or alkynes, among others (Scheme 18, bottom left).^[36] While structurally related, a reductive carbonylation event constitutes a more challenging process since the presence of hydride donors might result in hydrogenolysis of the organic halide via reduction of the intermediate oxidative addition species (Scheme 18, bottom right). If successful, such a method might represent a straightforward and practical alternative, particularly in industry, to the classical synthesis of formyl derivatives based on the treatment of stoichiometric amounts of organometallic species with dimethylformamide.



Scheme 18. Pd-catalyzed carbonylation reactions.

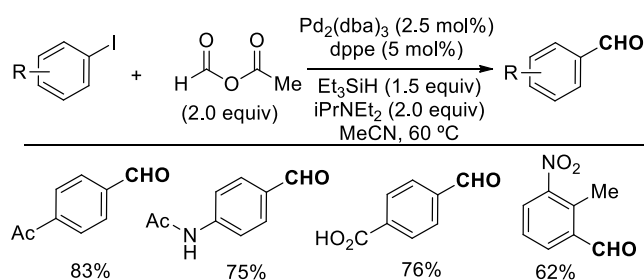
In 1974, Heck and co-workers reported the first palladium-catalyzed reductive carbonylation of aryl and vinyl halides with gaseous CO utilizing hydrogen as reducing agent.^[37] Unfortunately, such seminal discovery was conducted at high pressures (80-100 bar) and temperatures (80-150 °C) and large amounts of catalyst were required, hence limiting its synthetic application profile. Prompted by such precedents, many research groups described milder alternatives utilizing other hydride sources such as tributyltin hydride (Bu_3SnH),^[38] organosilanes^[39] and particularly attractive and readily available formate salts.^[40] In 2006, Beller reported a rather general and highly efficient catalytic reductive carbonylation that used a catalytic system based on $\text{Pd}(\text{OAc})_2$ as the precatalyst and di-1-adamantyl-*n*-butylphosphine ($n\text{BuPAD}_2$, CataCXium A) as the ligand (Scheme 19).^[41] Interestingly, the scope of the reaction included the coupling of (hetero)aryl and vinyl bromides by using simple and advantageous Syngas (CO/H_2 1:1) at lower pressures (5 bar) together with TMEDA as base at 100 °C. Remarkably, these processes were performed on multigram scale for the synthesis of aromatic aldehydes. They further studied in detail the reaction mechanism and unambiguously characterized relevant palladium complexes that suggested a scenario consisting of an oxidative addition, which was found to be rate-determining, followed by CO insertion into the Pd-C bond and subsequent hydrogen transfer aided by the base to deliver the targeted benzaldehyde.^[42]



Scheme 19. Pd-catalyzed reductive carbonylation of aryl halides.

While a significant step forward, the use of rather expensive and air-sensitive *n*BuPAD₂ constituted a serious drawback to be overcome. Preliminary studies from Beller and co-workers revealed that the topology of the phosphine-type ligand had a crucial impact on the reaction outcome,^[41,42] while the replacement of the adamantyl units in *n*BuPAD₂ by *tert*-butyl groups had little influence on reactivity, the use of *Pt*Bu₃ resulted in no conversion to products.^[41,42] In line with these observations, they described a new protocol for the reductive carbonylation of aryl and heteroaryl bromides based on the use of cheaper and easy-to-handle di-*tert*-butylphosphinite ligands with Syngas (5 bar) at 100 °C.^[43] Unlike the former protocol based upon *n*BuPAD₂ that exhibited a wide substrate scope, this new catalyst system was found inefficient when coupling particularly challenging 3-bromopyridine and sterically hindered substrates such as mesityl bromide. An additional bonus of this method came from the observation that (*t*Bu)₂PCl can be used as a surrogate of the corresponding di-*tert*-butylphosphinite by *in situ* treatment with TMEDA in *n*-propanol.

With the aim of designing new cost-efficient catalyst that enhance the synthetic utility of the reductive carbonylation of aryl halides, Nagarkar and co-workers reported an alternate protocol involving a combination of Pd(acac)₂ and commercially available and bench stable 1,1-bis(diphenylphosphino)methane (dppm) as supporting ligand.^[44] Under these conditions, the formylation of certain aryl and heteroaryl iodides occurred smoothly at 100 °C with Syngas (10 bar). In this case, 3-iodopyridine was successfully formylated; unfortunately, 2-iodothiophene and aryl iodides bearing electron-withdrawing substituents exhibited poor reactivity. It is worth noting, however, that this method relied on the use of comparatively more expensive aryl iodides as coupling partners; indeed, the use of aryl bromides was found to be highly inefficient for their purposes.

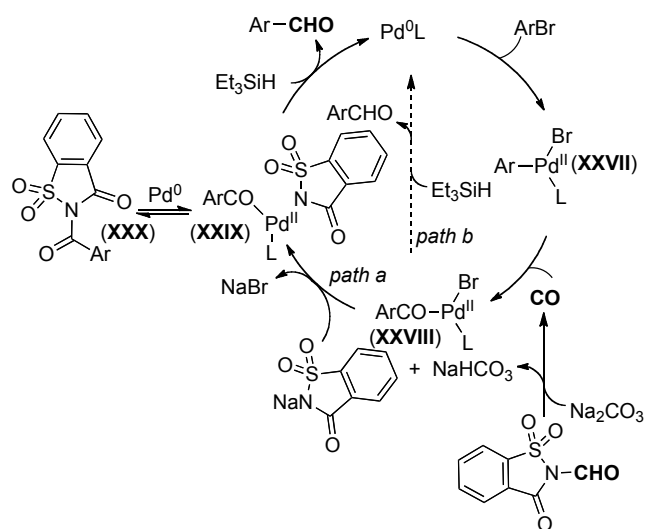
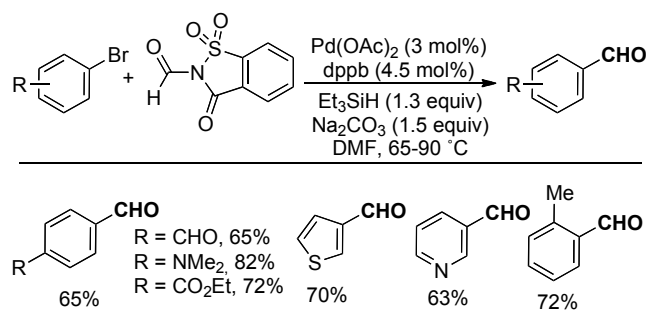


Scheme 20. Use of mixed anhydride as CO source in Pd-catalyzed reductive carbonylation reactions.

Owing to the high toxicity of CO and the requirement for high-pressure equipment in many instances, a number of more practical and advantageous CO-free carbonylation reactions have been developed. In 2004 Cacchi and co-workers devised a Pd-catalyzed reductive carbonylation of aryl iodides utilizing a thermally unstable mixed anhydride derived from acetic and formic acid as carbonylation source and Et₃SiH as reducing agent (Scheme 20).^[45] They found that the use of Pd₂(dba)₃ in combination with 1,1-bis(diphenylphosphino)ethane (dppe), *i*Pr₂EtN as base in CH₃CN at 60 °C led to the conversion of a wide variety of aryl iodides with different electronic environments into the corresponding benzaldehydes. The formylation of electron-poor aryl iodides was conducted in DMF in the presence of 3.0 equivalents of LiCl; the use of such additive, however, is not totally understood, although cation solvation might likely play a role.

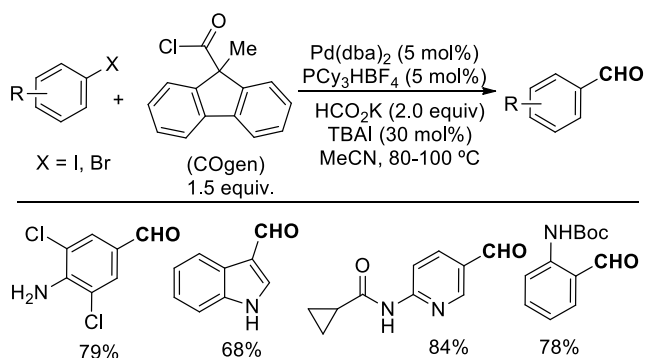
In 2013, Manabe and co-workers introduced the use of *N*-formylsaccharin as a cheap and easy-to-handle CO source in the Pd-catalyzed reductive carbonylation of aryl bromides with Et₃SiH as reducing agent (Scheme 21).^[46] This CO surrogate was first developed by the group of Cossy and applied as a formylating agent of amines.^[47] Control experiments verified that the decarbonylation of *N*-formylsaccharin to *in situ* produce CO and saccharin quantitatively occurred within 30 min in the presence of a mild base such as Et₃N or Na₂CO₃. Interestingly, the process was not strongly ligand-dependent as compared with previous reductive carbonylations for similar means.^[41,44] In this respect, a variety of bidentate phosphines showed excellent activity being 1,1-bis(diphenylphosphino)buthane (dppb) the ligand of choice for evaluating the preparative scope of the process. The method was efficient for the formylation of a wide range of aryl and heteroaryl bromides at 65-90 °C under low catalyst loadings, and in the presence of sensitive functional groups. Although some *ortho* substituted aryl bromides underwent the desired cross-coupling reaction, the extension to more sterically-demanding 2,6-disubstituted aryl bromides was unfruitful. Whereas the use of aryl iodides provided similar results, aryl triflates showed comparatively lower reactivity. Regarding the reaction mechanism, the authors hypothesized that an alternate pathway to the commonly accepted route involving the trapping of acylpalladium species by Et₃SiH may be operative. In order to shed light into the mechanism, several tests were conducted, indicating the following: (a) the use of *N*-benzoylsaccharin under the optimal reaction conditions provided benzaldehyde in 64% HPLC yield; (b) formation of acylsaccharin was confirmed by HRMS analysis when using 2-bromonaphthalene as substrate, and it could be trapped by addition of *n*BuOH to deliver the corresponding *n*-butyl ester. In line with these observations, a plausible catalytic cycle was suggested (Scheme 21, bottom) in which a ligand exchange occurred from **XXVIII** with *in situ* formed saccharin en route to **XXIX** that might be in equilibrium with **XXX** (*path a*). Final

hydride transfer from Et₃SiH would furnish the corresponding aldehyde and recover back the active catalyst; however, the authors could not rule out an alternate pathway consisting of a hydride transfer from **XXVIII** with Et₃SiH (*path b*).



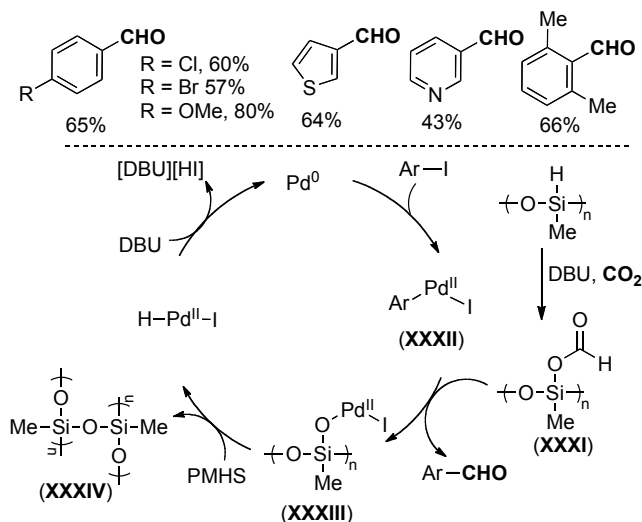
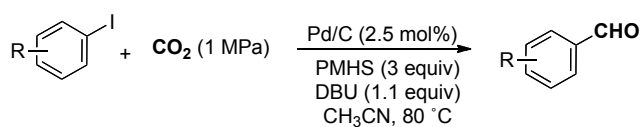
Scheme 21. Use of *N*-formylsaccharin as CO source in Pd-catalyzed reductive carbonylation reactions.

In recent years, Skrydstrup and co-workers elegantly implemented 9-methylfluorene-9-carbonyl chloride (COgen) as an innovative CO source in a sealed two-chamber reactor (COware) for conducting a vast array of Pd-catalyzed carbonylation processes.^[48] Interestingly, the use of simple and cheap potassium formate as reducing agent allowed for the reductive formylation of a wide variety of aryl iodides and bromides (Scheme 22).^[49] High selectivity towards the formylation was achieved when using Pd(dba)₂ along with monodentate phosphine ligands such as PCy₃ (or its HBF₄ salt) or CataXCium A; intriguingly, the use of bulkier P(*t*Bu)₃ led to the preferential formation of carboxylic acids.^[50] Remarkably, TBAI was required as phase-transfer agent for enhancing the catalyst performance. The protocol was characterized by its robustness and compatibility with the presence of numerous functional groups such as free-amines, aryl tosylates and nitrogen-containing heterocycles. The application for rapidly preparing isotopically labelled aldehydes in the presence of ¹³C-*COgen* and/or DCO₂K is particularly noteworthy given the few methods available to produce such compounds.^[49] The synthetic utility of this method was illustrated by the rapid synthesis of ¹³C-labeled florbetaben, a powerful β-amyloid binding compound.^[49]



Scheme 22. Pd-catalyzed reductive carbonylation of aryl halides in a two-chamber reactor.

Recently, the Liu group has implemented the advantageous use of carbon dioxide (CO₂) as environmentally friendly C1 source (*vide infra*) in palladium-catalyzed reductive formylation processes (Scheme 23).^[51] They found that the conversion of aryl iodides into the corresponding benzaldehydes could be accomplished with CO₂ (1 MPa) in the presence of Pd/C as catalyst, 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) as base and poly(methylhydrosiloxane) (PMHS) as hydride source in MeCN at 80 °C. This protocol was particularly efficient for the formylation of numerous aryl iodides; whereas substrates bearing electron-donating groups provided good to excellent yields, the presence of electron-withdrawing groups afforded the corresponding aldehydes in much lower yields. As expected, in these cases dehalogenated arenes were observed as a result of a rapid hydride transfer from the siloxane to the *in situ* formed oxidative addition species prior CO insertion. Remarkably, *ortho* substituents did not hamper the reaction and aryl bromides and chlorides did not undergo the formylation process. Interestingly, they gained some mechanistic insights by performing certain control experiments: (a) the formation of [DBU][HI] as an ionic liquid was detected, thus revealing a dual role of the base by activating the CO₂ and neutralizing the HI released in the process; (b) benzoic acid did not produce benzaldehyde under the standard conditions ruling out an initial carboxylation process; (c) ¹H-NMR analysis confirmed the formation of silylformate (**XXXI**) when mixing PMHS and CO₂ in the presence of DBU. Accordingly, they proposed that silylformate species (**XXXI**) react with the oxidative addition species (**XXXII**) to deliver the targeted aldehyde and silyloxypalladium complex (**XXXIII**). Finally, subsequent hydrogen transfer from PMHS and reduction of HPdI assisted by DBU regenerates the active Pd(0) catalyst.



Scheme 23. Pd-catalyzed carbonylation of aryl iodides with CO₂.

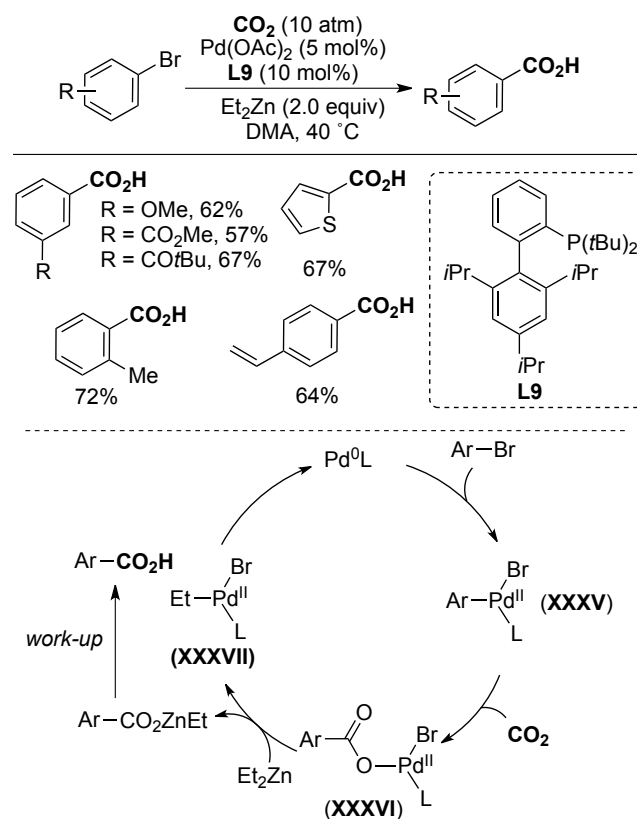
Despite the advances realized, most of the reductive carbonylation reactions reported to date are restricted to the use of aryl iodides, bromides or particularly activated aryl triflates. A major challenge in these endeavours for the years to come would imply the development of mild reductive carbonylations applicable to cheaper aryl chlorides and other environmentally friendly coupling partners such as C–O electrophiles derived from phenols.

Reductive Carboxylations with CO₂

Carbon dioxide (CO₂) is highly abundant, non-toxic, inexpensive and remarkably attractive as an excellent renewable C1 feedstock.^[52] The significant increase of global CO₂ emissions in the atmosphere along with its negative impact on the environment have undoubtedly prompted chemists to face practical means for fixing CO₂ into organic molecules. The successful implementation of such technologies could represent a direct alternative to the use of phosgene or CO as C1 source. Still, however, the high thermodynamic and kinetic stability of CO₂ constitutes a serious drawback when devising CO₂ fixation processes. Among all conceivable scenarios, the use of transition-metal catalysts represents one of the most powerful techniques to achieve CO₂ insertion processes. Despite the latest elegant methodologies for fixing CO₂ into useful organic compounds,^[53] the next section is mainly focused on the recently disclosed metal-catalyzed reductive carboxylations of organic halides and (pseudo)halides with CO₂ in the field of homogeneous catalysis.

Owing to the prevalence of carboxylic acids in a wide range of medically relevant compounds and organic materials, a plethora of novel synthetic routes, particularly in a catalytic fashion, have recently been developed. In this regard, the direct carboxylation of organometallic species and other carbon nucleophiles with CO₂ have emerged as powerful and advantageous alternatives to classical synthetic approaches for preparing carboxylic acids.^[54,55] Despite the advances realized, the need for stoichiometric amounts of well-defined and air-

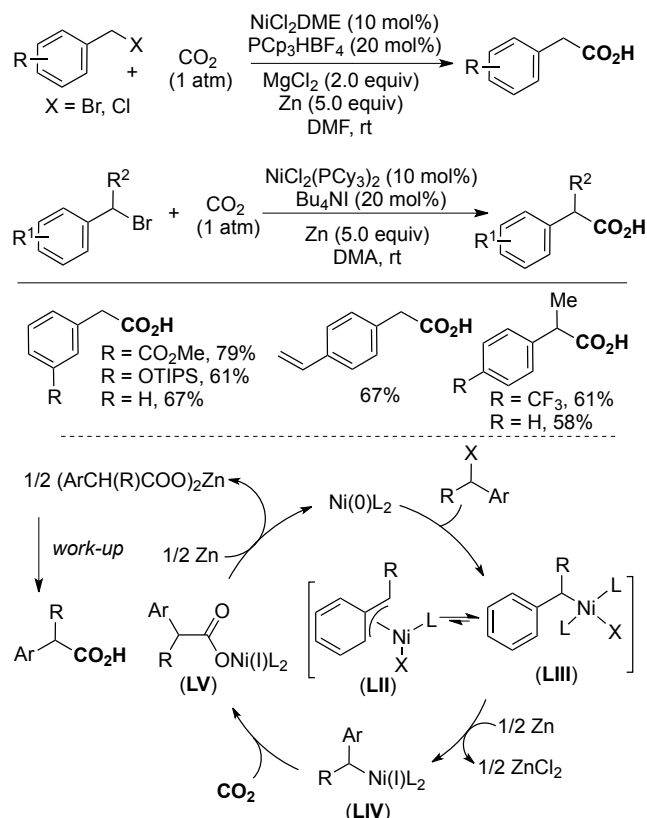
sensitive organometallic species limit the application profile of these rather appealing events. Therefore, the development of alternate user-friendly carboxylative processes operating in the absence of sensitive metal species and using readily available substrates is of utmost synthetic importance.



Scheme 24. Pd-catalyzed reductive carboxylation of aryl bromides.

Prompted by a seminal discovery of Osakada and Yamamoto using stoichiometric amounts of nickel salts,^[56] and by the excellent electrochemical methods reported by Périchon, Jutand, and Amatore,^[57] the Martin group reported in 2009 a Pd-catalyzed reductive carboxylation of aryl bromides with CO₂ utilizing Et₂Zn as reducing agent (Scheme 24).^[58] As for many other cross-coupling methodologies,^[1] the nature of the ligand was critical for success, with *t*BuXPhos showing a unique reactivity. The use of structurally-related phosphines provided variable amounts of either dehalogenated or Negishi-type products, thus revealing the crucial importance of using a particular bulky and electron-rich phosphine to rapidly insert CO₂ into the Pd–C bond (XXXVI) while inhibiting the competitive transmetalation with Et₂Zn. Interestingly, reducing agents other than Et₂Zn were found much less selective, leading to reduced arene or homocoupling products. While one might argue that the use of Et₂Zn could compromise the chemoselectivity profile, this was not the case and numerous functional groups such as amines, (thio)ethers, ketones, alkenes, esters and even oxiranes were perfectly tolerated. Noteworthy, *ortho* substituents did not hamper the reductive carboxylation reaction. However, the method was restricted to the use of aryl bromides as substrates; indeed, it was found that aryl chlorides remained intact under the reaction conditions while the use of aryl iodides provided the corresponding carboxylic acid in comparatively lower yields. The mechanism was believed to involve an oxidative addition of an aryl bromide to an *in situ* generated Pd(0) catalyst, subsequent CO₂ insertion

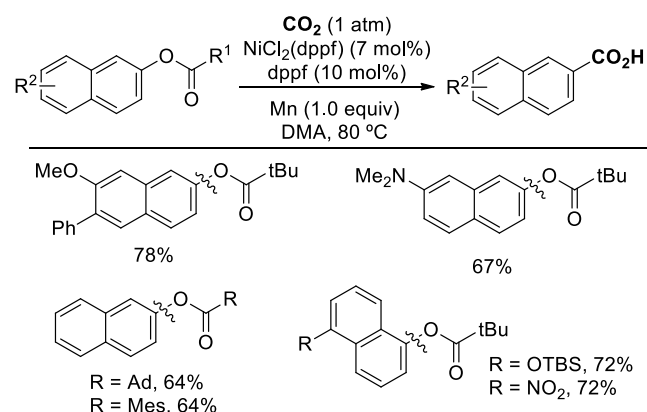
alcohols and ketones. Moreover, sterically hindered aryl iodides smoothly underwent the reductive carboxylation, even at room temperature. The proposed reaction mechanism involved the reduction of CuI to Cu(0) species assisted by Et₂Zn, subsequent oxidative addition of the aryl iodide, CO₂ insertion into the resulting Ar–Cu(I) intermediate (L) and final transmetalation to deliver the corresponding zinc carboxylate with concomitant regeneration of the Cu(0) catalyst (Scheme 26, bottom). Control experiments performed with Hg as additive evidenced the possible involvement of copper clusters. Although a reversible aryl transfer step between copper and zinc species was postulated, the CO₂ insertion event was believed to occur more rapidly into the C–Cu bond than in the C–Zn bond.



Scheme 27. Ni-catalyzed reductive carboxylation of benzyl halides.

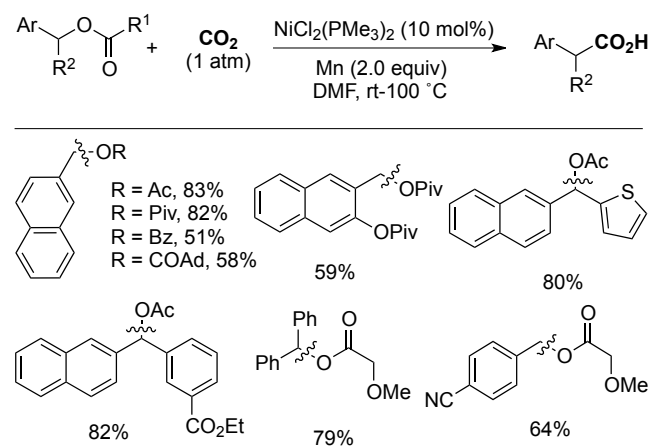
The Martin group has recently extended the scope of the reductive carboxylation processes to the use of activated benzyl bromides and chlorides en route to synthetically valuable phenylacetic acids (Scheme 27).^[62,63] They found that primary benzyl halides were smoothly carboxylated at room temperature and atmospheric pressure of CO₂ in the presence of a catalyst system based upon a cocktail of NiCl₂·glyme, PCp₃·HBF₄ and MgCl₂ as additive. Alternatively, the carboxylation of secondary and tertiary benzyl halides could also be accomplished by using NiCl₂(PCy₃)₂ as catalyst and tetrabutylammonium iodide (TBAI) as additive. In all cases Zn dust was the reducing agent of choice. As for other related carboxylative protocols, the method accommodated a wide number of functional groups and substitution patterns; of particular significance is the tolerance of terminal olefins and aryl halides, moieties that are known to promote related carboxylative events with Ni catalysts, thus showing the remarkable chemoselectivity of the method towards the benzyl halide motif. Following up a similar rationale to previous methods, the authors suggested a mechanism involving an initial

oxidative addition of benzyl halide to *in situ* generated Ni(0) species, delivering a nickel(II) intermediate that likely coexists in a rapid equilibrium of η³- and η¹-bound nickel species (LII and LIII; Scheme 27, bottom). Zn is then believed to assist the generation of a more nucleophilic Ni(I) species that rapidly undergo CO₂ insertion, an observation that is in analogy with related electrochemical methods.^{[24],[57]} An alternative hypothesis consisting of a comproportionation event of LII or LIII with a Ni(0) center could not be entirely excluded.^[64] Subsequently, a final transmetalation with Zn delivers the targeted carboxylic acid while regenerating the active propagating Ni(0) species. Noteworthy, the authors managed to isolate and characterize by X-ray crystallography complex LII. The study of its reactivity, together with the ability of radical traps to inhibit the reaction pointed towards the generation of Ni(I) species via SET processes.



Scheme 28. Ni-catalyzed reductive carboxylation of C(sp²)-O electrophiles.

In recent years unactivated C–O electrophiles have emerged as advantageous and environmentally friendly (pseudo)halide coupling partners in a wide range of cross-coupling processes.^[65] Martin and co-workers recently implemented their use in Ni-catalyzed reductive carboxylations with CO₂, thus avoiding the typical halide waste associated to the use of aryl halides (Scheme 28).^[66] The particular combination of NiCl₂(dppf) and Mn dust in DMA as solvent allowed the carboxylation of several C(sp²)-O bonds to occur at 80 °C and atmospheric pressure of CO₂. The use of co-catalytic amounts of 1,1'-bis(phenylphosphino)ferrocene (dppf) proved essential to enhance the catalyst efficiency. As expected, the nature of the ligand exerted a profound influence on the reaction outcome and other related ferrocene-type ligands entirely inhibited the reductive carboxylation reaction.



Scheme 29. Ni-catalyzed reductive carboxylation of C(sp³)-O electrophiles.

Notably, the reductive carboxylation of more challenging C(sp³)-O motifs was also within reach when using NiCl₂(PMe₃)₂ as catalyst and Mn as reducing agent in DMF at room temperature (Scheme 29).^[66] This catalyst system was found applicable for the carboxylation of a wide variety of benzyl-type substrates. In this respect, not only bulky benzyl pivalates but also simpler acetates and benzoates smoothly underwent the reductive process. Remarkably, both primary and secondary benzyl-type derivatives were perfectly accommodated and, unlike previous carboxylative protocols based on the use of organic halides, the authors found that this new protocol was amenable for the coupling of substrates bearing various heterocycles. Despite the synthetic value of the method, the scope was primarily restricted to the use of π -extended systems, an observation that is in line with recent developments in the C-O bond-cleavage arena.^[67] It was postulated that a partial dearomatization of the arene ring might occur prior oxidative addition; therefore, an extended π -system might bind to the Ni center in a η^2 -fashion via the Dewar-Chat-Chatt-Duncanson model, thus retaining, unlike a regular arene, certain aromaticity that provides extra stabilization.^[68] The authors overcome such limitation by introducing traceless hemilabile directing groups when using C(sp³)-O electrophiles (Scheme 29).^[69] The authors postulated that the presence of such directing group might accelerate the rate of oxidative addition with regular arenes while opening up coordination sites on the Ni center, thus facilitating the binding of CO₂ and its subsequent insertion event. The use of 2-methoxy acetate derivatives was particularly effective, allowing for the successful reductive carboxylation of differently substituted benzyl ester derivatives lacking π -extended systems. In analogy with previous studies, they proposed a reaction pathway involving Ni(I) species as the key intermediates.

While the field of metal-catalyzed reductive carboxylation of organic halides or (pseudo)halides has reached a considerable level of development, this field of expertise is still at its infancy when compared with other CO₂ fixation methods. This is likely due to the fact that the underlying mechanism from which these reactions operate is fundamentally speculative at this point. Therefore, we anticipate that *in depth* mechanistic studies might lead to the foundation of new reductive carboxylative events counterintuitive at first sight. We certainly expect a bright future in the years to come in this particular area of expertise.

Summary and Outlook

This mini-review seeks to provide an up-to-date overview of the latest advances in the field of metal-catalyzed reductive couplings of carbonyl-type compounds with organic halides and/or (pseudo)halide derivatives. The wealth of literature data reported in recent years suggests that these methods might represent a direct and straightforward alternative to the common use of well-defined, stoichiometric, and in many cases, air-sensitive organometallic species. The key aspect of metal-catalyzed reductive coupling reactions relies on the *in situ* generation of catalytically competent organometallic species that rapidly react with a variety of electrophilic counterparts containing carbonyl-type compounds. Despite the advances realized, some of these protocols are still limited to the use of aryl halides or activated (pseudo)halides such as aryl sulfonates; therefore, the implementation of more appealing

coupling partners such as cheap aryl chlorides and environmentally friendlier C-O electrophiles would open up new synthetic opportunities of paramount chemical significance. At present, little knowledge has been gathered regarding the mechanism of these transformations and only indirect evidences have been performed in order to propose a rational mechanistic scenario. The study of reaction mechanisms via the isolation of the putative reaction intermediates has demonstrated its critical role for the discovery of new reactions within the cross-coupling arena. Therefore, we anticipate that efforts along these lines would have a significant impact on the field of metal-catalyzed reductive coupling reactions. Additionally, the implementation of robust and widely-tolerant asymmetric processes within this field of expertise would allow for the design of new synthetic routes to advanced intermediates via the use of unconventional bond-disconnections and from simple precursors.

Acknowledgements

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- [1] For selected reviews: a) D. S. Surry, S. L. Buchwald, *Chem. Sci.* **2011**, *2*, 27-50; b) R. Jana, T. P. Pathak, M. S. Sigman, *Chem. Rev.* **2011**, *111*, 1417-1492; c) G. C. Fu, *Acc. Chem. Res.* **2008**, *41*, 1555-1564; d) R. Martin, S. L. Buchwald, *Acc. Chem. Res.* **2008**, *41*, 1461-1473; e) D. Alberico, M. E. Scott, M. Lautens, *Chem. Rev.* **2007**, *107*, 174-238; f) *Metal-Catalyzed Cross-Coupling Reactions* (Eds: F. Diederich, A. Meijere) Wiley-VCH, Weinheim, **2004**.
- [2] For selected reviews: a) B. -J. Li, Z. -J. Shi, *Chem. Soc. Rev.* **2012**, *41*, 5588-5598; b) C. Liu, H. Zhang, W. Shi, A. Lei, *Chem. Rev.* **2011**, *111*, 1780-1824; c) W. Shi, C. Liu, A. Lei, *Chem. Soc. Rev.* **2011**, *40*, 2761-2776; d) C. -L. Sun, B. -J. Li, Z. -J. Shi, *Chem. Commun.* **2010**, *46*, 677-685.
- [3] Bew, S. P. *In Comprehensive Organic Functional Groups Transformation II*; Katritzky, A. R., Taylor, R. J. K., Eds.; Elsevier: Oxford, 2005 ; pp 19.
- [4] For related reductive coupling processes not using organic halides as coupling partners: a) "Organonickel Chemistry" in *Organometallics in Synthesis: Fourth Manual*, (Eds: B.H. Lipshutz) Wiley, Hoboken, N.J., **2013**; b) M. Jeganmohan, C. H. Cheng, *Chem. Eur. J.* **2008**, *14*, 10876-10886; c) J. Montgomery, G. J. Somunen, *Top. Curr. Chem.* **2007**, *279*, 1-23; d) R. M. Moslin, K. Miller-Moslin, T. F. Jamison, *Chem. Commun.* **2007**, 4441-4449; e) E. Skucas, M. -Y. Ngai, V. Komanduri, M. J. Krische, *Acc. Chem. Res.* **2007**, *40*, 1394-1401; f) J. Montgomery, *Angew. Chem.* **2004**, *116*, 3980-3998; *Angew. Chem. Int. Ed.* **2004**, *49*, 3890-3908; g) S. Ikeda, *Angew. Chem.* **2003**, *115*, 5276-5278; *Angew. Chem. Int. Ed.* **2003**, *42*, 5120-5122.
- [5] For selected references on the use of organic halides in reductive coupling events see: a) L. -Y. Liao, X. -R. Kong, X. -F. Duan, *J. Org. Chem.* **2014**, *79*, 777-782; b) Q. Qian, Z. Zang, S. Wang, Y. Chen, K. Lin, H. Gong, *Synlett* **2013**, 619-624; c) Y. Peng, L. Luo, C. -S. Yan, J. -J. Zhang, Y. -W. Wang, *J. Org. Chem.* **2013**, *78*, 10960-10967; d) D. A. Everson, B. A. Jones, D. J. Weix, *J. Am. Chem. Soc.* **2012**, *134*, 6146-6159; e) S. Wang, Q. Qian, H. Gong, *Org. Lett.* **2012**, *12*, 3352-3355; f) Y. Chang-Song Yan, Y. Peng, X. -B. Xu, Y. -W. Wang, *Chem. Eur. J.* **2012**, *18*, 6039-6048; g) L. L. Anka-Lufford, M. R. Prinsell, D. J. Weix, *J. Org. Chem.* **2012**, *77*, 9989-10000; h) D. A. Everson, R. Shrestha, D. J. Weix, *J. Am. Chem. Soc.* **2010**, *132*, 920-921.
- [6] a) P. Wender, B. L. Miller, *Nature* **2009**, *460*, 197-201; b) B. M. Trost, *Acc. Chem. Res.* **2002**, *35*, 695-705; c) B. M. Trost, *Angew. Chem. Int. Ed. Engl.* **1995**, *34*, 259-281.
- [7] For selected reviews see: a) J. -Y. Nedelec, J. Perichon, M. Troupel, *Top. Curr. Chem.* **1997**, *185*, 141-173; b) J. Chaussard, J. C. Folest, J.

- Y. Nedelec, J. Perichon, S. Sibille, M. Troupel, *Synthesis* **1990**, 5, 369-381.
- [8] P. Knochel, W. Dohle, N. Gommermann, F. F. Kneisel, F. Kopp, T. Korn, I. Sapountzis, V. Anh Vu, *Angew. Chem.* **2003**, *115*, 4438-4456; *Angew. Chem. Int. Ed.* **2003**, *48*, 4302-4320.
- [9] For selected examples: a) H. Tanaka, S. Yamashita, T. Hamatani, Y. Ikemoto, S. Torii, *Chem. Lett.* **1986**, 1611-1614; b) C. Peter, J. Einhorn, J. L. Luche, *Tetrahedron Lett.* **1985**, *26*, 1449-1452; c) C. Perier, J. L. Luche, *J. Org. Chem.* **1985**, *50*, 910-912; d) T. Hiyama, M. Sawahata, Y. Kusano, *Chem. Lett.* **1985**, 611-612; e) M. Wada, K. Akiba, *Tetrahedron Lett.* **1985**, *26*, 4211-4212; f) T. Imamoto, T. Kusumoto, Y. Tawarayama, Y. Sugiura, T. Mita, *J. Org. Chem.* **1984**, *49*, 3904-3912; g) G. P. Boldrini, D. Savoia, E. Tagliavini, C. Trombin, A. Umani-Ronchi, *J. Org. Chem.* **1983**, *48*, 4108-4111; h) J. Nokami, J. Otera, T. Sudo, R. Okawara, *Organometallics*, **1983**, *2*, 191-193; i) T. Hiyama, M. Obayashi, A. Nakamura, *Organometallics*, **1982**, *1*, 1249-1251; j) J. L. Luche, J. C. Damiano, *J. Am. Chem. Soc.* **1980**, *102*, 7927-7300; k) C. Blomberg, F. A. Hartog, *Synthesis* **1977**, 18-30.
- [10] a) J. J. Miller, M. S. Sigman, *J. Am. Chem. Soc.* **2007**, *129*, 2752-2753; b) A. Fürstner, *Chem. Rev.* **1999**, *99*, 991-1045; c) A. Fürstner, N. Shi, *J. Am. Chem. Soc.* **1996**, *118*, 2533-2534.
- [11] R. W. Armstrong, J. M. Beau, S. H. Cheon, W. J. Christ, H. Fujioka, W. H. Ham, L. D. Hawkins, H. Jin, S. H. Kang *J. Am. Chem. Soc.* **1989**, *111*, 7530-7533.
- [12] G. C. Hargarden, P. J. Guiry, *Adv. Synth. Catal.* **2007**, *349*, 2407-2424
- [13] K. K. Majumdar, C. –H. Cheng, *Org. Lett.* **2000**, *2*, 2295-2298.
- [14] For related reductive couplings of aldehydes with allylic carbonates see: Z. Tan, X. Wan, Z. Zang, Q. Qian, W. Deng, H. Gong, *Chem. Commun.* **2014**, 3827-3830.
- [15] Y. –C. Huang, K. K. Majumdar, C. –H. Cheng, *J. Org. Chem.* **2002**, *67*, 1682-1684.
- [16] M. Durandetti, C. Gosmini, J. Perichon, *Tetrahedron* **2007**, *63*, 1146-1153.
- [17] J. –X. Hu, C. –Y. Li, W. –J. Sheng, Y. –X. Jia, J. –R. Gao, *Chem. Eur. J.* **2011**, *17*, 5234-5237.
- [18] H. Fillon, C. Gosmini, J. Perichon, *Tetrahedron* **2003**, *59*, 8199-8202.
- [19] For selected examples on the electrochemical coupling reactions of carbonyl-type compounds, see: a) M. Durandetti, C. Meignen, J. Perichon, *Org. Lett.* **2003**, *5*, 317-320; b) S. Seka, O. Buriez, J. Perichon, *Chem. Eur. J.* **2003**, *9*, 3597-3603; c) S. Seka, O. Buriez, J. –Y. Nedelec, J. Perichon, *Chem. Eur. J.* **2002**, *8*, 2534-2538; d) M. Durandetti, J. –Y. Nedelec, J. Perichon, *Org. Lett.* **2001**, *3*, 2073-2076; e) C. Amatore, A. Jutand, J. Perichon, Y. Rollin, *Monatsh. Chem.* **2000**, *131*, 1293-1304; f) M. Durandetti, J. Perichon, J. –Y. Nedelec *Tetrahedron Lett.* **1999**, *40*, 9009-9013; g) A. Conan, S. Sibille, J. Perichon, *J. Org. Chem.* **1991**, *56*, 2018-2024.
- [20] G. Cahiez, A. Moyeux, *Chem. Rev.* **2010**, *110*, 1435-1462.
- [21] F. Wu, W. Lu, Q. Qian, Q. Ren, H. Gong, *Org. Lett.* **2012**, *14*, 3044-3047.
- [22] W. Lu, Z. Liang, Y. Zhang, F. Wu, Q. Qian, H. Gong, *Synthesis* **2013**, 2234-2240.
- [23] a) S. W. Smith, G. C. Fu *J. Am. Chem. Soc.* **2008**, *130*, 12645-12647; b) C. Fisher, G. C. Fu, *J. Am. Chem. Soc.* **2005**, *127*, 4594-4595; c) J. Zhou, G. C. Fu, *J. Am. Chem. Soc.* **2003**, *125*, 14726-14727.
- [24] For related catalytic electrochemical methods based on the generation of metal species in odd oxidation states, see: ref. 19
- [25] A. C. Wotal, D. J. Weix, *Org. Lett.* **2012**, *14*, 1476-1479.
- [26] T. Yamamoto, T. Kohara, A. Yamamoto, *Bull. Chem. Soc. Jpn.* **1981**, *54*, 2010-2016; b) T. Yamamoto, T. Kohara, K. Osakada, A. Yamamoto, *Bull. Chem. Soc. Jpn.* **1983**, *56*, 2147-2153.
- [27] A. H. Cherney, N. T. Kadunce, S. E. Reisman, *J. Am. Chem. Soc.* **2013**, *135*, 7442-7445.
- [28] I. Kazmierski, M. Bastienne, C. Gosmini, J. –M Paris, J. Pèrichon, *J. Org. Chem.* **2004**, *69*, 936-942.
- [29] H. Yin, C. Zhao, H. You, K. Lin, H. Gong, *Chem. Commun.* **2012**, 7034-7036.
- [30] J. –C. Hsieh, C. –H. Cheng, *Chem. Commun.* **2005**, 4554-4556.
- [31] a) A. A. Pletnev, R. C. Larock, *Tetrahedron Lett.* **2002**, *43*, 2133-2136; b) A. A. Pletnev, R. C. Larock, *J. Org. Chem.* **2002**, *67*, 4554-4556.
- [32] J. –C. Hsieh, Y. –C. Chen, A. –Y. Cheng, H. –C. Tseng, *Org. Lett.* **2012**, *14*, 1282-1285.
- [33] Y. Xia, F. Hu, Z. Liu, P. Qu, R. Ge, C. Ma, Y. Zhang, J. Wang, *Org. Lett.* **2013**, *15*, 1784-1787.
- [34] R. Shrestha, S. C. M. Dorn, D. J. Weix, *J. Am. Chem. Soc.* **2013**, *135*, 751-762.
- [35] R. Shrestha, D. J. Weix, *Org. Lett.* **2011**, *13*, 2766-2769.
- [36] a) X. –F. Wu, H. Neumann, M. Beller, *Chem. Rev.* **2013**, *113*, 1-35; b) X. –F. Wu, H. Neumann, M. Beller, *Chem. Soc. Rev.* **2011**, *40*, 4986-5009; c) Q. Liu, H. Zhang, A. Lei, *Angew. Chem.* **2011**, *119*, 10978-10989; *Angew. Chem. Int. Ed.* **2011**, *50*, 10788-10799; d) A. Brennfürher, H. Neumann, M. Beller, *Angew. Chem.* **2009**, *121*, 4176-4196; *Angew. Chem. Int. Ed.* **2009**, *48*, 4114-4133;
- [37] A. Schoenberg, R. F. Heck, *J. Am. Chem. Soc.* **1974**, *96*, 7761-7764.
- [38] a) A. B. Smith III, L. Kürti, A. H. Davulcu, Y. S. Cho, *Org. Process. Res. Dev.* **2007**, *11*, 19-24; b) V. P. Baillargeon, J. K. Stille, *J. Am. Chem. Soc.* **1986**, *108*, 452-461; c) V. P. Baillargeon, J. K. Stille, *J. Am. Chem. Soc.* **1983**, *105*, 7175-7176.
- [39] a) L. Ashfield, C. F. J. Barnard, *Org. Process Res. Dev.* **2007**, *11*, 39-43; b) D. C. Behenna, J. L. Stockdill, B. M. Stoltz, *Angew. Chem.* **2007**, *119*, 4155-4158; *Angew. Chem. Int. Ed.* **2007**, *46*, 4077-4080; c) H. Kotsuki, P. K. Datta, H. Suenaga, *Synthesis* **1996**, 470-472.
- [40] a) M. –Z. Cai, H. Zhao, J. Zhou, C. –S. Song, *Synth. Commun.* **2002**, *32*, 923-926; b) Y. Ben-David, M. Portnoy, D. Milstein, *J. Chem. Soc. Chem. Commun.* **1989**, 1816-1817; c) I. Pri-Bar, O. Buchman, *J. Org. Chem.* **1984**, *49*, 4009-4011.
- [41] a) A. Brennfürher, H. Neumann, M. Beller, *Synlett* **2007**, 2537-2540; b) A. Brennfürher, H. Neumann, S. Klaus, T. Riermeier, J. Almèna, M. Beller, *Tetrahedron* **2007**, *63*, 6252-6258; c) S. Klaus, H. Neumann, A. Zapf, D. Strübing, S. Hübner, J. Almèna, T. Riermeier, P. Groß, M. Sarich, W. –R. Krahnert, K. Rossen, M. Beller, *Angew. Chem.* **2006**, *118*, 161-165; *Angew. Chem. Int. Ed.* **2006**, *45*, 154-158.
- [42] A. G. Sergeev, A. Spannenberg, M. Beller, *J. Am. Chem. Soc.* **2008**, *130*, 15549-15563.
- [43] H. Neumann, R. Kadyrov, X. –F. Wu, M. Beller, *Chem. Asian J.* **2012**, *7*, 2213-2216.
- [44] A. S. Singh, B. M. Bhanage, J. M. Nagarkar, *Tetrahedron Lett.* **2011**, *52*, 2383-2386.
- [45] S. Cacchi, G. Fabrizi, A. Goggiamani, *J. Comb. Chem.* **2004**, *6*, 692-694.
- [46] T. Ueda, H. Konishi, K. Manabe, *Angew. Chem.* **2013**, *125*, 8773-8777; *Angew. Chem. Int. Ed.* **2013**, *52*, 8611-8615.
- [47] T. Cocheta, V. Bellostaa, A. Greinerb, D. Rocheb, J. Cossy, *Synlett* **2011**, 1920-1922.
- [48] S. Korsager, D. U. Nielsen, R. H. Taaning, T. Skrydstrup, *Angew. Chem.* **2013**, *125*, 9945-9948; *Angew. Chem. Int. Ed.* **2013**, *52*, 9763-9766 and references cited therein.
- [49] S. Korsager, R. H. Taaning, A. T. Lindhardt, T. Skrydstrup, *J. Org. Chem.* **2013**, *78*, 6112-6120.
- [50] See for example: a) S. Korsager, R. H. Taaning, T. Skrydstrup, *J. Am. Chem. Soc.* **2013**, *135*, 2891-2894; b) M. N. Burhardt, R. H. Taaning, T. Skrydstrup, *Org. Lett.* **2013**, *15*, 948-951 and references cited therein; c) P. Hermange, A. T. Lindhardt, R. H. Taaning, K. Bjerglund, D. Lupp, T. Skrydstrup, *J. Am. Chem. Soc.* **2011**, *133*, 6061-6071.
- [51] B. Yu, Y. Zhao, H. Zhang, J. Xu, L. Hao, X. Gao, Z. Liu, *Chem. Commun.* **2014**, *50*, 2330-2333.
- [52] For recent reviews dealing with the use of CO₂: a) Y. Tsuji, T. Fujihara, *Chem. Commun.* **2012**, 9956-9964; b) M. Cokoja, C. Bruckmeier, B.

- Rieger, W. A. Herrmann, F. E. Kühn, *Angew. Chem.* **2011**, *123*, 8662-8690; *Angew. Chem. Int. Ed.* **2011**, *50*, 8510-8537; c) R. Martin, A. W. Kleij, *ChemSusChem* **2011**, *4*, 1259-1263; d) K. Huang, C. -L. Sun, Z. -L. Shi, *Chem. Soc. Rev.* **2011**, *40*, 2435-2452; e) *Carbon Dioxide as Chemical Feedstock* (Eds: M. Aresta), Wiley-VCH, Weinheim, **2010**; f) T. Sakakura, J. -C. Choi, H. Yasuda, *Chem. Rev.* **2007**, *107*, 2365-2387 and references cited therein.
- [53] Recent selected examples: a) O. Jacquet, X. Frogneux, C. D. N. Gomes, T. Cantat, *Chem. Sci.* **2013**, *4*, 2127-2131; b) S. Kikuchi, K. Sekine, T. Ishida, T. Yamada, *Angew. Chem.* **2012**, *124*, 7095-7098; *Angew. Chem. Int. Ed.* **2012**, *51*, 6989-6992; c) L. Zhang, J. Cheng, B. Carry, Z. Hou, *J. Am. Chem. Soc.* **2012**, *134*, 14314-14317; d) S. Li, W. Yuan, Y. Ma, *Angew. Chem.* **2011**, *123*, 2626-2630; *Angew. Chem. Int. Ed.* **2011**, *50*, 2578-2582; e) T. Fujihara, T. Xu, K. Semba, J. Terao, Y. Tsuji, *Angew. Chem.* **2011**, *123*, 543-547; *Angew. Chem. Int. Ed.* **2011**, *50*, 523-527; f) H. Mizuno, J. Takaya, N. Iwasawa, *J. Am. Chem. Soc.* **2011**, *133*, 1251-1253; g) I. I. F. Boogaerts, S. P. Nolan, *J. Am. Chem. Soc.* **2010**, *132*, 8858-8859; h) L. Zhang, J. Cheng, T. Ohishi, Z. Hou, *Angew. Chem.* **2010**, *122*, 8852-8855; *Angew. Chem. Int. Ed.* **2010**, *49*, 8670-8673.
- [54] For selected references: a) K. Kobayashi, Y. Kondo, *Org. Lett.* **2009**, *11*, 2035-2037; b) J. Takaya, S. Tadami, K. Ukai, N. Iwasawa, *Org. Lett.* **2008**, *10*, 2697-2700; c) T. Ohishi, M. Nishiura, Z. Hou, *Angew. Chem.* **2008**, *120*, 5876-5879; *Angew. Chem. Int. Ed.* **2008**, *47*, 5792-5795; d) H. Ochiai, M. Jang, K. Hirano, H. Yorimitsu, K. Oshima, *Org. Lett.* **2008**, *10*, 2681-2683; e) C. S. Yeung, V. M. Dong, *J. Am. Chem. Soc.* **2008**, *130*, 7826-7827; f) K. Ukai, M. Aoki, J. Takaya, N. Iwasawa, *J. Am. Chem. Soc.* **2006**, *128*, 8706-8707.
- [55] A. Correa, R. Martin, *Angew. Chem.* **2009**, *121*, 6317-6320; *Angew. Chem. Int. Ed.* **2009**, *48*, 6201-6204.
- [56] K. Osakada; R. Sato, R.; T. Yamamoto, T. *Organometallics* **1994**, *13*, 4645-4647.
- [57] For selected examples on the metal-catalyzed electrochemical reductive carboxylation of organic halides: a) S. Torii, H. Tanaka, T. Hamatani, K. Morisaki, A. Jutand, F. Pflüger, J. F. Fauvarque, *Chem. Lett.* **1986**, 169-172; b) J. F. Fauvarque, A. Jutand, M. François, M. A. Petit, *J. Appl. Electrochem.* **1988**, *18*, 116-119; c) C. Amatore, A. Jutand, *J. Am. Chem. Soc.* **1991**, *113*, 2819-2825; d) C. Amatore, A. Jutand, F. Khalil, M. F. Nielsen, *J. Am. Chem. Soc.* **1992**, *114*, 7076-7085.
- [58] A. Correa, R. Martin, *J. Am. Chem. Soc.* **2009**, *131*, 15974-15975.
- [59] T. Fujihara, K. Nogi, T. Xu, J. Terao, Y. Tsuji, *J. Am. Chem. Soc.* **2012**, *134*, 9106-9109.
- [60] F. B. Sayyed, Y. Tsuji, S. Sakaki, *Chem. Commun.* **2013**, 10715-10717.
- [61] H. Tran-Vu, O. Daugulis, *ACS Catal.* **2013**, *3*, 2417-2420.
- [62] T. León, A. Correa, R. Martin, *J. Am. Chem. Soc.* **2013**, *135*, 1221-1224.
- [63] For related synthesis of phenylacetic acids not using organic halides via a hydrocarboxylation event with CO₂, see: a) M. D. Greenhalgh, S. P. Thomas, *J. Am. Chem. Soc.* **2012**, *134*, 11900-11903; b) C. M. Williams, J. B. Johnson, T. Rovis, *J. Am. Chem. Soc.* **2008**, *130*, 14936-14937; c) H. Hoberg, S. Gross, A. Milchereit, *Angew. Chem.* **1987**, *99*, 567-569; *Angew. Chem. Int. Ed.* **1987**, *26*, 571-572; d) H. Hoberg, Y. Peres, A. J. Milchereit, *J. Organomet. Chem.* **1986**, *307*, C38-C40.
- [64] For selected comproportionation events in route to Ni(I) species, see: (a) J. Cornella, E. Gómez-Bengoia, R. Martin *J. Am. Chem. Soc.* **2013**, *135*, 1997-2009; (b) A. Velian, S. Lin, A. J. M. Miller, M. W. Day, T. Agapie, *J. Am. Chem. Soc.* **2010**, *132*, 6296-6297; (c) G. D. Jones, J. L. Martin, C. McFarland, O. R. Allen, R. E. Hall, A. D. Haley, R. J. Brandon, T. Kononova, P. J. Desrochers, P. Pulay, D. A. J. Vivic, *J. Am. Chem. Soc.* **2006**, *128*, 13175-13183.
- [65] For recent reviews: a) M. Tehetena, N. K. Garg, *Org. Process Res. Dev.* **2013**, *17*, 29-39; b) J. Yamaguchi, K. Muto, K. Itami, *Eur. J. Org. Chem.* **2013**, 19-30; c) A. Correa, J. Cornella, R. Martin, *Angew. Chem.* **2013**, *125*, 1928-1930; *Angew. Chem. Int. Ed.* **2013**, *52*, 1878-1880; d) B. M. Rosen, K. W. Quasdorf, D. A. Wilkson, N. Zhang, A. -M. Resmerita, N. K. Garg, B. Percec, *Chem. Rev.* **2011**, *111*, 1346-1416;
- e) B. -J. Li, D. -G. Yu, C. -L. Sun, Z. -J. Shi, *Chem. Eur. J.* **2011**, *17*, 1728-1759; f) D. -G. Yu, B. -J. Li, Z. -J. Shi, *Acc. Chem. Res.* **2010**, *43*, 1486-1495.
- [66] A. Correa, T. León, R. Martin, *J. Am. Chem. Soc.* **2014**, *136*, 1062-1069.
- [67] The use of extended π -systems has been a requisite for seminal discoveries in the area of inert C-O bond-cleavage reactions, with most of compounds in the substrate scope being extended π -systems. See for example: a) H. M. Wisniewska, E. C. Swift, E. R. J. Jarvo, *J. Am. Chem. Soc.* **2013**, *135*, 9083-9095; b) B. L. H. Taylor, E. C. Swift, J. D. Waetzig, E. R. J. Jarvo, *J. Am. Chem. Soc.* **2011**, *133*, 389-391; c) B. L. Taylor, M. R. Harris, E. R. J. Jarvo, *Angew. Chem.* **2012**, *124*, 7910-7913; *Angew. Chem. Int. Ed.* **2012**, *51*, 7790-7793; d) D. -G. Yu, Z. -J. Shi, *Angew. Chem.* **2011**, *123*, 7235-7238; *Angew. Chem. Int. Ed.* **2011**, *50*, 7097-7100; e) D. -G. Yu, B. -J. Li, S. -F. Zheng, B. -T. Guan, B. -Q. Wang, Z. -J. Shi, *Angew. Chem.* **2010**, *122*, 4670-4674; *Angew. Chem. Int. Ed.* **2010**, *49*, 4566-4570.
- [68] a) J. Chatt, L. A. Duncanson, L. M. Venanzi, *J. Chem. Soc.* **1955**, 4456-4460; b) J. Chatt, L. A. Duncanson, *J. Chem. Soc.* **1953**, 2939-2947; c) M. J. S. Dewar, *Bull. Soc. Chim. Fr.* **1951**, *18*, C71-79.
- [69] G. Rousseau, B. Breit, *Angew. Chem.* **2011**, *123*, 2498-22543; *Angew. Chem. Int. Ed.* **2011**, *50*, 2450-2494.

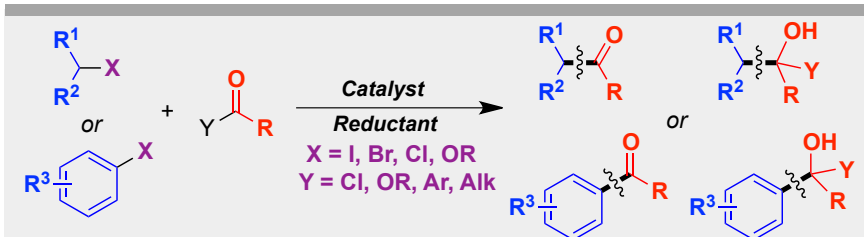
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FULL PAPER



Metal-catalyzed reductive couplings of a variety of electrophilic aryl halides and (pseudo)halides with carbonyl-type compounds have undergone impressive development along the last years. In this review, the recent findings in this field are summarized, with particular emphasis on the mechanistic interpretation of the results and the future aspects of this area of expertise.

■ Catalytic Reductive Coupling Reactions

Toni Moragas, Arkaitz Correa and Ruben Martin**

■■ - ■■

■ Metal-catalyzed Reductive Coupling Reactions of Organic Halides with Carbonyl-type Compounds