
Nitric Oxide Reduction with Hydrogen, Carbon Monoxide, and Hydrocarbons over Gold Catalysts

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The reactivity of nitric oxide with reducing gases, using supported gold catalysts is reviewed and assessed in comparison with the use of other precious metal catalysts. Applications for this technology in the purification of exhaust gases are envisaged.

Gold was once thought to be poorly active as a heterogeneous catalyst. However, since we reported that gold shows extraordinarily high catalytic activity for the oxidation of carbon monoxide when it is deposited as nanoparticles on selected metal oxides (1, 2), gold has attracted a growing interest from scientists and engineers (3, 4).

The catalytic reduction of nitrogen oxides has been one of the most important scientific challenges during the last few decades. The most practical and convenient method for removing nitric oxide (NO) seems to be catalytic reduction using unburnt compounds such as carbon monoxide (CO), hydrogen, and hydrocarbons contained in the exhaust gas. There are two different practical conditions for catalysts to work; *ie* stoichiometric and lean-burn. Under stoichiometric conditions, where the amount of air supplied is controlled to a level just sufficient for the complete combustion of the fuel, three-way catalysts are used. Only noble metals such as rhodium, palladium, and platinum can be used as catalytic components, due to their durability under both the oxidizing and reducing parts of the cycles. These noble metals are usually supported on oxides with oxygen-storage capacity, such as CeO₂ deposited on Al₂O₃. On the other hand, diesel engines and lean-burn gasoline engines are operated under lean-burn conditions, where the oxygen is fed in excess, 10 to 20% more than is required to meet the stoichiometry for combustion of the fuel.

Catalytic reduction of nitrogen oxides (NO_x) by hydrocarbons in the presence of excess oxygen is a reasonable target, because a certain amount (100 - 1000 ppm) of propene and ethene is usually contained in the exhaust gases of lean-burn gasoline engines and diesel engines (5-7). It has been reported that a variety of metal ion-exchanged zeolites such as Pt-MFI (8, 9) and Cu-MFI (10, 11) exhibit good catalytic performances for the reduction of nitric oxide by light alkenes. When natural gas is used as the fuel for gas turbines and lean-burn gas engines, light alkanes rather than alkenes remain as major reducing components in the exhaust gases (12). The reduction of nitric oxide with methane has been attempted over catalysts containing cobalt (13), iridium (14), palladium (15), *etc.* Ethane and propane have also been studied as reducing agents over Co-β-zeolite catalyst (16). Carbon monoxide and hydrogen are also contained in exhaust gases, however, but their contribution to the reduction of NO_x is not well understood.

Supported gold catalysts are surprisingly active for the oxidation of carbon monoxide with oxygen and moderately active for the oxidation of hydrogen by oxygen (2). In this present paper we provide an overview of our recent work on the catalysis of gold for the reduction of nitrogen oxides by carbon monoxide, hydrogen and hydrocarbons, and attempt to compare gold with other transition metals.

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EXPERIMENTAL

Preparation and Characterization of Gold Catalysts

The catalytic properties of gold are very dependent on the methods employed for preparation of the catalysts. The method used determines the size and the contact structure of gold particles deposited on the metal oxide supports. Co-precipitation (2) and deposition-precipitation techniques (17) are effective for preparing gold catalysts such as gold supported on oxides containing iron, Au/Al₂O₃, Au/ZnO, and Au/TiO₂. One important point in the preparation procedure is that gold tetrachloride anion is transformed into gold hydroxide precursor which is mixed with the hydroxides of the other metals (co-precipitation) or deposited on metal oxide supports (deposition-precipitation).

In this work, gold was deposited on Al₂O₃, TiO₂, and MgO using a deposition-precipitation technique (17). After the pH of an aqueous solution of chloroauric acid (HAuCl₄) was adjusted to 7.0 by adding the appropriate quantity of an aqueous 0.1M sodium hydroxide (NaOH) solution, metal oxide powder was dispersed in the solution under agitation for 1 h to deposit gold hydroxide (Au(OH)₃) exclusively on the surface of metal oxide particles. The precursor obtained was washed several times to remove sodium and chloride ions and dried under vacuum for 18 h, and finally calcined in air at 773 K for 5 h. In the case of the MgO support, no sodium hydroxide was added, and Au(OH)₃ was deposited at pH 9-10 in an aqueous solution of HAuCl₄ containing magnesium citrate. Gold supported on ZnO was prepared by co-precipitation (2) from a mixed aqueous solution of HAuCl₄ and Zn(NO₃)₂ followed by similar procedures to those for deposition-precipitation.

Specific surface areas of the catalysts prepared were determined by the BET single point method using a surface area analyzer (Quantasorb Jr). The amount of gold deposited on metal oxides was determined using an inductively coupled plasma spectrometer ICP-AES (Seiko SPS 1200 VR). The X-ray diffraction pattern of each sample before and after the reaction was measured using a Rigaku X-ray powder diffractometer (RINT 2000). The crystallite size of gold was calculated using Scherrer's equation.

Catalytic Activity Measurements

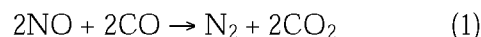
Catalytic activity was measured using a fixed-bed flow reactor (18). A sample (300 mg) was placed in a quartz tube, heated to 573 K for 30 min in a stream of helium at a flow rate of 100 cm³ min⁻¹, and then kept at 573K for 30 min. For nitric oxide reduction with

hydrocarbons, the reactant gas mixture containing NO 1000 ppm, C₃H₆ 1000 ppm, O₂ 5.0 vol%, and H₂O 10.0 vol% in a helium background was passed through the catalyst bed at a flow rate of 100 cm³ min⁻¹ unless otherwise stated. The catalyst temperature was raised stepwise and maintained for 2 h at each temperature. Reactants and products were analysed with a gas chromatograph (Shimadzu GC-14B), and nitric oxide (NO) and nitrogen dioxide (NO₂) were analysed with a chemiluminescent NO_x meter (Yanaco CLS-88US). Calibration was carried out with a standard gas containing known concentrations of the components.

RESULTS AND DISCUSSION

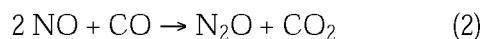
Nitric Oxide Reduction by Carbon Monoxide over Supported Gold Catalysts

Reduction of nitric oxide by carbon monoxide is one of the reactions occurring in catalytic converters for the purification of the engine exhaust gases (Equation 1):



It is well known that rhodium is the most active amongst all the precious metals for the catalytic conversion of NO by CO (19, 20). Although there are some differences in the literature concerning the relative orders of activity reported for the platinum group metals (21-24), a well-established order is Rh > Pd > Pt.

Table 1 shows the experimental results for the NO + CO reaction in the absence of oxygen over gold supported on several metal oxides. Over Au/Al₂O₃, the reaction of NO with CO took place even at a temperature as low as 323 K producing nitrous oxide (N₂O):



Over rhodium supported on Al₂O₃ the reaction takes place at temperatures above 400 K (21, 22), showing that, in principle, gold catalysts are more active at low temperatures than rhodium catalysts. Harrison *et al* (25) have proposed that NO decomposition can make a major kinetic contribution to the NO + CO reaction over supported rhodium catalysts. At low temperatures, N₂O is the principal product and it is only after the reaction temperature is raised to 573K that nitrogen formation takes place (26). In contrast, the formation of nitrogen from NO took place at 373 K in the presence of Au/Al₂O₃ catalyst.

Among precious metals, the catalytic activity order

Table 1 Reduction of Nitric Oxide with Carbon Monoxide on Supported Gold Catalysts^a

Catalyst	Conversion of NO to N ₂ and N ₂ O (%)				
	300 K N ₂ , N ₂ O	323 K N ₂ , N ₂ O	373 K N ₂ , N ₂ O	423 K N ₂ , N ₂ O	473 K N ₂ , N ₂ O
Au/MgO		0.0, 5.4	8.2, 15.2	42.7, 0.2	81.5, 0.0
Au/Al ₂ O ₃		0.0, 8.9	13.4, 24.7	57.2, 0.4	92.3, 0.0
Au/TiO ₂		0.0, 7.2	10.9, 18.4	47.6, 0.0	85.9, 0.0
Au/ α -Fe ₂ O ₃	0.0, 32.5	0.0, 49.5	22.6, 51.4	78.0, 0.0	100, 0.0
Au/ZnO		0.0, 12.7	18.9, 32.7	62.4, 0.0	97.4, 0.0
Au/MnFe ₂ O ₄	0.0, 98.9	0.2, 98.5	81.1, 4.9	96.5, 0.0	100, 0.0
Au/CoFe ₂ O ₄	0.0, 64.5	0.0, 96.1	45.6, 32.7	94.3, 0.0	100, 0.0
Au/NiFe ₂ O ₄	0.0, 97.4	0.4, 98.2	86.4, 12.0	94.0, 0.0	100, 0.0
Au/ZnFe ₂ O ₄	0.0, 84.1	0.0, 97.8	83.8, 15.4	100, 0.0	100, 0.0
Au/NiFe ₂ O ₄ ^b	0.0, 82.4	0.0, 89.4	92.1, 8.2	95.2, 0.0	100, 0.0
Au/NiFe ₂ O ₄ /Al ₂ O ₃ ^c	0.0, 24.7	0.0, 39.4	15.5, 42.8	62.3, 0.2	95.3, 0.0

^a Reaction gas: 0.1 vol% CO and 0.1 vol% NO in helium at a space velocity of 20,000 h⁻¹ cm³ g catalyst⁻¹. The gold loading on metal oxides was 1 wt%.

^b Wet conditions, reaction gas: 0.1 vol% CO, 0.1 vol% NO and 1.8 vol% H₂O in a helium background at a space velocity of 20,000 h⁻¹ cm³ g catalyst⁻¹.

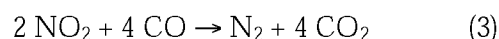
^c A bead-shaped catalyst, the loading of gold and NiFe₂O₄ on Al₂O₃-bead was 0.2 and 2 wt%, respectively.

for the NO + CO reaction is consistent with the order of CO inhibition strength (23). The infrared measurements for CO adsorption on the surface of gold particles on metal oxides show that CO inhibition is very weak (27), which could account for the highest catalytic activity for the NO + CO reaction with gold catalysts.

The catalytic activity for NO reduction by CO over supported gold catalysts changes drastically depending on the kind of metal oxide supports employed. It is especially noteworthy that oxides containing iron provide the highest catalytic activity for the NO + CO reaction. Over Au/NiFe₂O₄, the reduction of NO by CO took place even at 300K to form N₂O, and at 373K most of the NO was reduced to nitrogen. This extraordinarily high activity of gold catalysts can be maintained in the presence of moisture (Table 1). These results show that gold catalysts are superior in their low-temperature catalytic activity to Rh-modified (28), Pt-modified (29) and Pd-modified catalysts (30, 31).

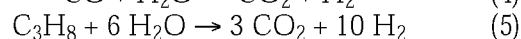
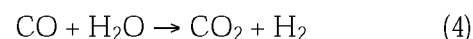
For the application to lean-burn exhaust gases, the NO + CO reaction in the presence of oxygen is of more practical importance, especially at lower temperatures. Figure 1 compares NO reduction by CO (1 vol%) over Au/Al₂O₃ with that for a

mechanical mixture of Mn₂O₃ with Au/Al₂O₃ in the presence of a high concentration oxygen (5 vol%). Over Au/Al₂O₃, the NO + CO reaction was strongly inhibited by the presence of oxygen, giving a maximum conversion of NO to N₂, of 5 %. The O₂ + CO reaction takes place in preference to the NO + CO reaction. On the other hand, the mechanical mixture of Mn₂O₃ with Au/Al₂O₃ gave a higher NO conversion to N₂ (21 %) at 623 K. It is likely that Mn₂O₃ catalyses the reaction of NO with O₂ to form nitrogen dioxide (NO₂), which then reacts with CO adsorbed on the surface of gold particles supported on Al₂O₃ (Equation 3):



Reduction of Nitric Oxide by Hydrogen over Supported Gold Catalysts

Hydrogen is also present in the exhaust gases from combustion of hydrocarbons because it can be generated either by the water gas shift reaction (Equation 4) or the steam reforming reaction (eg Equation 5):



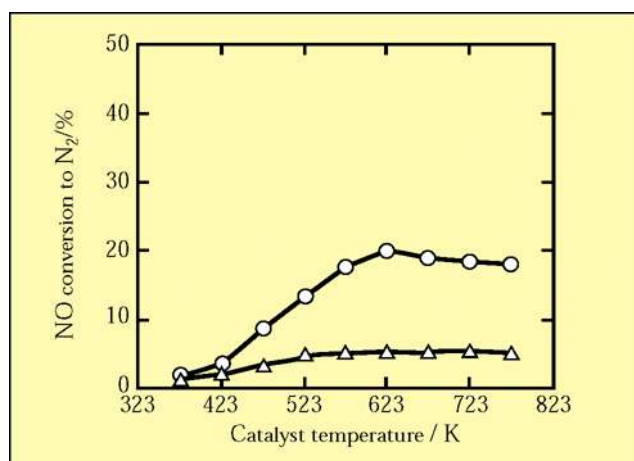
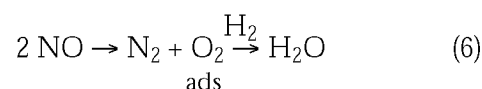


Figure 1 NO conversion to N_2 as a function of catalyst temperature over $Au/Al_2O_3 + Mn_2O_3$ and Au/Al_2O_3 in the $NO + CO + O_2$ reaction. (○), $Au(0.17\text{wt}\%)/Al_2O_3 + Mn_2O_3$ (19:1 in weight); (△), $Au(0.17\text{wt}\%)/Al_2O_3$. Reaction gas: 0.1 vol% NO, 0.1 vol% CO, and 5.0 vol% O_2 in a helium background at a space velocity of $20,000\text{ h}^{-1}\text{ cm}^3\text{ g-catalyst}^{-1}$.

Table 2 lists the results for the reduction of NO by H_2 over Au, Pd, Pt, Rh, Ru, and Ir catalysts in the presence of oxygen. The relative catalytic activity sequence is $Pt > Rh > Pd > Au$, $Ru > Ir$. Taylor (21) proposed that the reaction paths for NO reduction by hydrogen over Pd and Pt supported on Al_2O_3 are NO

decomposition over the metals followed by the removal of adsorbed oxygen species by H_2 (Equation 6):



The above order for NO reduction by H_2 can be explained by the fact that gold gives the middle activity amongst the precious metals for both NO decomposition and the $O_2 + H_2$ reaction. Over Pd/TiO_2 , the two conversion maxima for NO reduction by hydrogen appear at temperatures of 400 K and 600 K, and this is closely related to NO oxidation to NO_2 (32). The combination of NO oxidation by oxygen and NO_2 reduction by hydrogen may also give further improvement of NOx conversion to N_2 as was the case with the $NO + CO + O_2$ reaction.

Reduction of Nitric Oxide by Hydrocarbons over Supported Gold Catalysts

Support Effect of Metal Oxides

Table 3 shows that the temperature range at which the reduction of NO to N_2 by propene takes place varies with the type of metal oxide support used for gold. Gold supported on ZnO , $\alpha\text{-Fe}_2O_3$ and ZrO_2 is active at lower temperatures. This is especially notable at 523K, where the Au/ZnO catalyst gives a maximum of 49% in the conversion to N_2 and, in addition, 16% conversion to N_2O (Equation 7):

Table 2 Reduction of Nitric Oxide with Hydrogen over Supported Precious Metals Catalysts^a

Catalyst	Conversion of NO to N_2 and N_2O (%)				
	323 K N_2, N_2O	373 K N_2, N_2O	473 K N_2, N_2O	573 K N_2, N_2O	623 K N_2, N_2O
Ru/Al_2O_3	0.0, 0.0	5.2, 3.2	27.3, 5.8	7.9, 0.0	1.2, 0.0
Rh/Al_2O_3	0.2, 1.5	5.4, 34.5	0.5, 0.0	0.5, 0.0	0.2, 0.0
Pd/Al_2O_3	0.0, 0.8	2.7, 0.0	1.2, 0.0	2.1, 0.0	1.2, 0.0
Pd/TiO_2	2.8, 6.8	21.7, 25.8	2.8, 2.6	27.4, 17.5	24.4, 15.7
Ir/Al_2O_3	0.0, 0.0	6.4, 2.8	32.4, 0.2	12.4, 0.0	7.9, 0.0
Pt/Al_2O_3	0.4, 4.4	6.6, 55.6	0.2, 0.0	2.8, 0.0	0.2, 0.0
Pt/TiO_2	1.5, 5.8	10.5, 40.1	0.4, 0.2	0.6, 0.0	0.2, 0.0
Au/Al_2O_3	0.0, 0.0	6.4, 2.8	32.4, 0.2	12.4, 0.0	7.9, 0.0
Au/TiO_2	0.0, 0.0	7.8, 1.5	34.3, 0.0	10.7, 0.0	4.2, 0.0

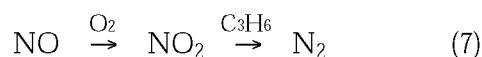
^a Reaction gas: 0.1 vol% NO, 0.3 vol% H_2 , 5.0 vol% O_2 , and 10.0 vol% H_2O in helium, at a space velocity of $20,000\text{ h}^{-1}\text{ cm}^3\text{ g catalyst}^{-1}$. The loading of precious metal on metal oxides was 1 wt%.

Table 3 Metal Loadings, Mean Diameters of Gold Particles, and Specific Surface Areas of the Tested Catalysts, and NO Conversion to N₂ and N₂O over Several Gold Catalysts and Metal Oxides

Catalyst	Preparation method ^a	Au loading ^b (wt.%)	d _{Au} ^c (nm)	Surface area (m ² g ⁻¹)	NO conversion to N ₂ (%) ^d						
					473K	523K	573K	623K	673K	723K	773K
Au/MgO	DP	0.85	5.1	166	8.6	12.6	20.2	42.5	37.7		
Au/Al ₂ O ₃	DP	0.17	-	172	3.8	12.3	30.8	50.9	80.7	80.5	61.6
Au/Al ₂ O ₃	DP	0.28	4.9	173	4.2	6.8	12.2	34.1	50.7	39.4	
Au/Al ₂ O ₃	IMP	1.1	32.0	174			0.1	18.2	15.5	10.3	
Au/TiO ₂	DP	0.76	3.3	39	5.5	8.1	18.4	30.4	18.9		
Au/α-Fe ₂ O ₃	DP	0.79	3.4	23	6.1	12.3	7.8	3.1	1.7		
Au/ZnO	CP	1.2	3.5	36	24.4	49.2	35.1	23.5	14.0		
Au/ZrO ₂	DP	0.84	3.8	24	12.3	32.4	24.5	17.4	8.2		
Al ₂ O ₃	-	-	-	174				0.5	1.2	1.8	2.6

^a DP; deposition-precipitation, IMP; impregnation, CP; co-precipitation. ^b Gold loading was determined by inductively coupled plasma spectrometry (ICP-AES).

^c Crystallite size of gold was calculated by using Scherrer's equation from XRD data. ^d NO reduction by propene, reaction condition; 0.1 vol% NO, 0.1 vol% C₃H₆, 5.0 vol% O₂, and 10.0 vol% H₂O in He, Space Velocity = 20,000 h⁻¹cm³g-catalyst⁻¹.



Gold supported on MgO and TiO₂ gives the middle-temperature conversion of NO. Gold (1wt%) on Al₂O₃ can provide appreciably higher conversion to N₂ with a maximum of 51 % at 673K and at this temperature the formation of N₂O was no longer observed. Because the minimum detectable concentration of N₂O was 10 ppm using our analytical instrumentation, the above result showed that NO conversion to N₂O was less than 2%. In contrast, Au/Al₂O₃ prepared by impregnation, which contains gold particles much larger than those obtained for deposition-precipitation, is much less active for NO reduction by propene. As is the case for CO oxidation (33), the preparation method used for gold catalysts is also very important in NO reduction and the size of the gold particles is one of the controlling factors for catalytic activity (34). The lower content of gold (0.2 wt%) on Al₂O₃ gives much higher NO conversion to N₂ (80 %), although it requires the relatively higher reaction temperature of 723 K.

In comparison with the above gold catalysts, platinum and palladium supported on metal oxides exhibit activities at lower temperatures but the propene and oxygen compete to a greater extent, resulting in lower maximum conversions to N₂ (35, 36).

Au/Al₂O₃ Catalysts

Since gold supported on alumina gives the highest

maximum conversion to nitrogen, its catalytic behaviour for NO reduction has been studied in greater detail.

Effect of Moisture in the Reactants

Comparison of the effect of moisture on NO reduction with propene is remarkably different for Au/Al₂O₃ and Al₂O₃ (Figure 2). Although the conversion to N₂ over Al₂O₃ markedly decreased with increasing concentration of water, over Au/Al₂O₃ an increase in water concentration gave rise to small increases in the conversion to nitrogen. The enhancing effect of moisture has also been observed in CO oxidation over other supported gold catalysts (37) and provides gold catalysts with a significant potential advantage in applications for combustion of exhaust gases, since these usually contain water at a concentration higher than 10 vol%. The catalysts which contain the oxides of base transition metals such as Cu, Ag, and Co, as catalytically active species were appreciably deactivated by moisture (12, 38, 39). In the case of Pt supported on metal oxides, NO conversion is maintained even in the presence of high concentrations of water, but this is accompanied by enhanced formation of N₂O (36, 40, 41).

Effect of Oxygen in the Reactants

Figure 3 indicates that NO conversion is very low (8.4%) in the absence of oxygen, and that the addition of oxygen in the reactant stream appreciably enhances the reduction of NO up to a conversion of 69% (at an oxygen concentration of 4 vol%). The conversion to

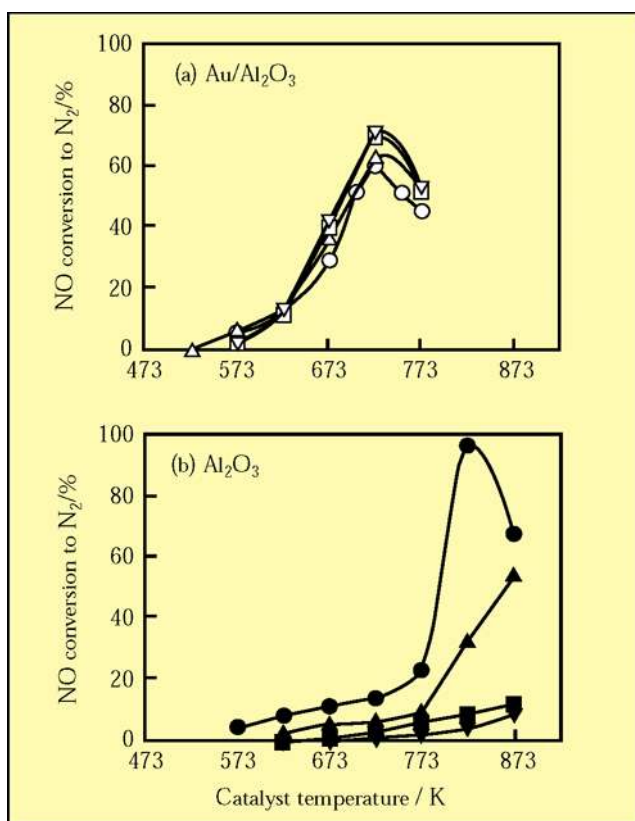


Figure 2 NO reduction to N_2 as a function of catalyst temperature over (a) Au/Al_2O_3 and (b) Al_2O_3 under different concentrations of H_2O . ($\circ, \triangle, \square, \nabla$), $Au(0.17 \text{ wt. \%})/Al_2O_3$; ($\bullet, \blacktriangle, \blacksquare, \blacktriangledown$), Al_2O_3 ; (\circ, \bullet), dry; ($\triangle, \blacktriangle$), $H_2O = 1.8 \text{ vol\%}$; (\square, \blacksquare), $H_2O = 5.1 \text{ vol\%}$; ($\nabla, \blacktriangledown$), $H_2O = 9.8 \text{ vol\%}$. Reaction gas: 0.1 vol\% NO , $0.1 \text{ vol\% } C_3H_6$, $5.0 \text{ vol\% } O_2$, and H_2O in a helium background at a space velocity of $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

N_2 remained almost constant in the oxygen concentration range of 6 to 20 vol%. This shows that oxygen is indispensable for the reduction of NO by propene. The steady conversion of NO on Au/Al_2O_3 catalysts at high oxygen concentration range is advantageous for applications with exhaust gases under lean burn conditions, which usually contain oxygen at a concentration from 1 to 15 vol%.

Figure 4 shows the effects of CO addition on NO reduction by propene over Au/Al_2O_3 . The CO addition gave a positive effect of higher yields of N_2 ($60 \rightarrow 75\%$, at 673 K).

Reaction Pathways in the Reduction of Nitric Oxide by Hydrocarbons

Since oxygen appreciably enhances the reduction of

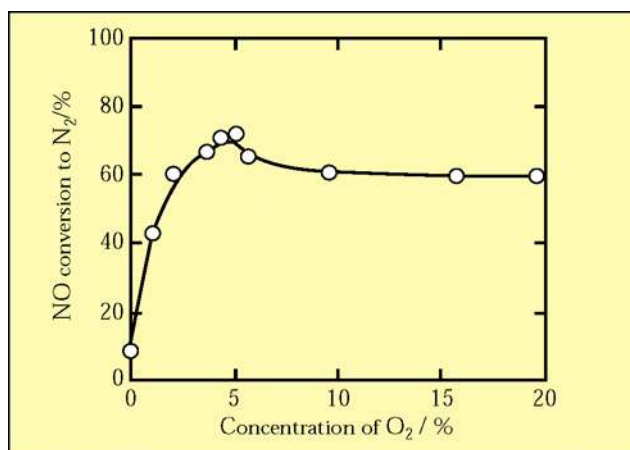


Figure 3 NO conversion to N_2 as a function of the concentration of oxygen over $Au(0.17 \text{ wt\%})/Al_2O_3$ at 723 K . Reaction gas: 0.1 vol\% NO , $0.1 \text{ vol\% } C_3H_6$, $5.1 \text{ vol\% } H_2O$, and O_2 in a helium background at a space velocity of $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

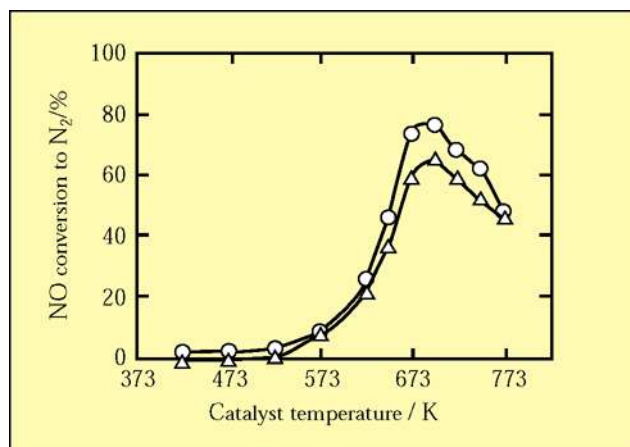


Figure 4 Effect of the CO addition into NO reduction by propene over $Au(1 \text{ wt\%})/Al_2O_3$. (\circ), 987 ppm NO , $750 \text{ ppm } C_3H_6$, 987 ppm CO , $4.9 \text{ vol\% } O_2$, and $5.1 \text{ vol\% } H_2O$ in a He background; (\triangle) 987 ppm NO , $750 \text{ ppm } C_3H_6$, $4.9 \text{ vol\% } O_2$, and $5.1 \text{ vol\% } H_2O$ in a helium background. Space velocity was $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

NO by hydrocarbons, it is likely that the oxidation of NO by oxygen ($2NO + O_2 \rightarrow 2NO_2$) takes place first, and then NO_2 is reduced by hydrocarbons to N_2 (34, 42, 43). As shown in Figure 5, the reaction of NO with molecular oxygen to form NO_2 takes place over Au/Al_2O_3 at temperatures above 573 K, just where the reduction of NO by propene starts (Figure 2), and the conversion to NO_2 reaches almost the equilibrium value at 723 K. In addition, moisture in the reactant

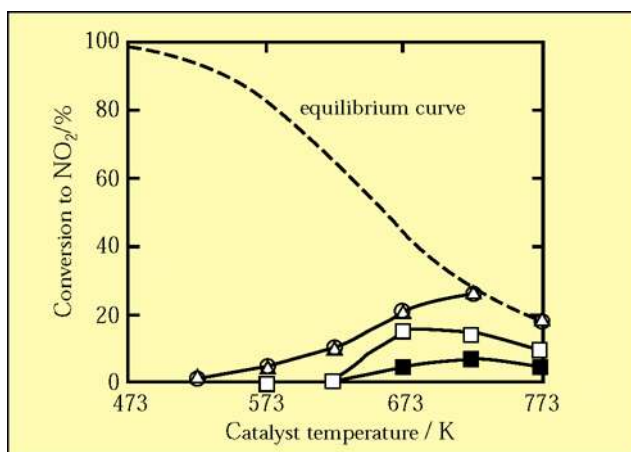


Figure 5 Conversion in NO oxidation as a function of catalyst temperature. (○, △), Au(1wt%)/Al₂O₃; (□, ■), Al₂O₃; (○, □), 0.1 vol% NO and 5.0 vol% O₂ in a helium background; (△, ■), 0.1 vol% NO, 5.0 vol% O₂, and 5.1 vol% H₂O in a helium background. Space velocity was 20,000 h⁻¹ cm³ g⁻¹ catalyst⁻¹

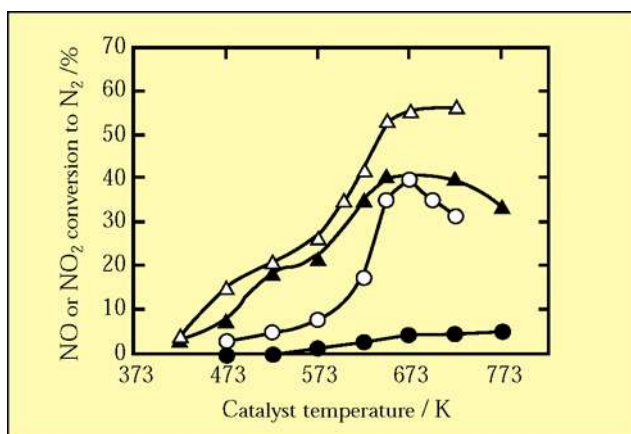


Figure 6 Effect of oxygen on the reduction of NO and NO₂ with propene over Au(1wt%)/Al₂O₃. (○), 0.1 vol% NO, 0.05 vol% C₃H₆, 5.0 vol% O₂, and 5.1 vol% H₂O in a helium background; (●), 0.1 vol% NO, 0.05 vol% C₃H₆, and 5.1 vol% H₂O in a helium background; (▲), 0.1 vol% NO₂, 0.05 vol% C₃H₆, 5.0 vol% O₂, and 5.1 vol% H₂O in a helium background; (△), 0.1 vol% NO₂, 0.05 vol% C₃H₆, and 5.1 vol% H₂O in a helium background. Space velocity = 20,000 h⁻¹ cm³ g⁻¹ catalyst⁻¹

stream does not depress the oxidation of NO with O₂. The reduction of NO₂ to N₂ by propene always takes place at lower temperatures than NO reduction, yielding higher conversions to N₂ (Figure 6), and

proceeds even in the absence of oxygen. These results indicate that the reactivity of NO₂ with propene is much larger than that of NO. Accordingly, the reduction of NO to N₂ over Au/Al₂O₃ catalysts can be assumed to take place *via* the formation of NO₂ by the oxidation of NO with oxygen as the first and slowest step and the subsequent reduction of NO₂ by propene. Since nitric oxide reacts with propene in a molar ratio of almost unity and the conversion of NO reaches about 70% on Au/Al₂O₃, propene may be oxidized first by NO₂ and then the subsequent oxidation of propene oxygenates taking place mainly with oxygen (34). This feature is one of the greatest constraints of supported gold catalysts, because it results in relatively low selectivity of hydrocarbons to NO reduction, compared with that of Cu-MFI catalysts (36).

Burch *et al* (44, 45) have proposed that NO decomposition, which produces adsorbed oxygen and nitrogen atoms, plays a major role in the reduction of NO by hydrocarbons over Pt supported on metal oxides. The reaction between the adsorbed oxygen atoms and hydrocarbons from the gas phase takes place on the surface of platinum particles supported on metal oxide. The relatively higher selectivity to N₂O over Pt on metal oxides can be explained by the formation of N₂O from NO and dissociated nitrogen atoms adsorbed on platinum particles.

Effect of Addition of Mn₂O₃ to Au/Al₂O₃ Catalysts

Since the oxidation of NO to NO₂ appears to be a rate-determining step in NO reduction, acceleration of this reaction would mean that NO reduction would take place at lower temperatures. Based on this hypothesis, Mn₂O₃ which is the most active catalyst for the oxidation of NO to NO₂ and less active for hydrocarbon oxidation with oxygen has been combined with Au/Al₂O₃ through mechanical mixing (18). As shown in Figure 7, this combination can maintain high conversion from lower to higher temperatures and offers one of the best catalytic performances in NO conversion among the catalysts so far reported. This bifunctional mechanism, composed of NO oxidation and NO₂ reduction by hydrocarbons has been proposed and reviewed by Misono (15, 48), Kikuchi (46), and Hamada (42). The catalytic performance of the mechanical mixture of Mn₂O₃ with Au/Al₂O₃ is at its best in the composition of Mn₂O₃ : Au(0.2 wt%)/Al₂O₃ = 1 : 19 in weight and is competitive with that of Pt-MFI [36] in its low-temperature activity. This catalyst gives a negligibly small production of N₂O and is superior to most of the catalysts reported to be active at high temperatures. It should also be noted that this

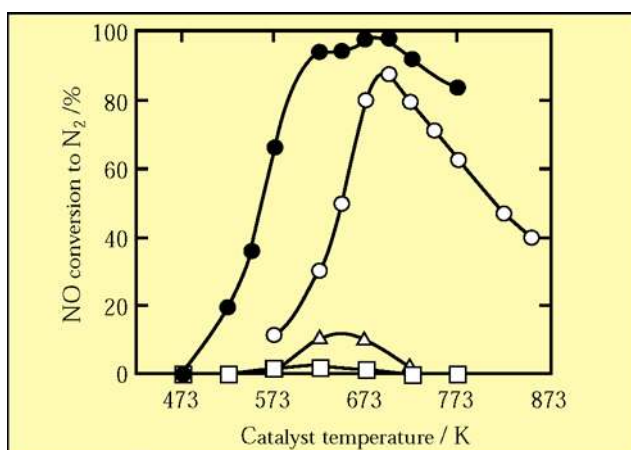


Figure 7 Effect of addition of Mn_2O_3 to Au/Al_2O_3 in the reduction of NO with propene. (○), $Au(0.17 \text{ wt.}\%)/Al_2O_3$; (△), $Au(0.16 \text{ wt.}\%)/Mn_2O_3$; (□), Mn_2O_3 ; (●), $Au(0.17 \text{ wt.}\%)/Al_2O_3 + Mn_2O_3$ (19:1 in weight). Reaction gas: 0.1 vol% NO, 0.1 vol% C_3H_6 , 5.0 vol% O_2 , and 10.0 vol% H_2O in a helium background. Space velocity = $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

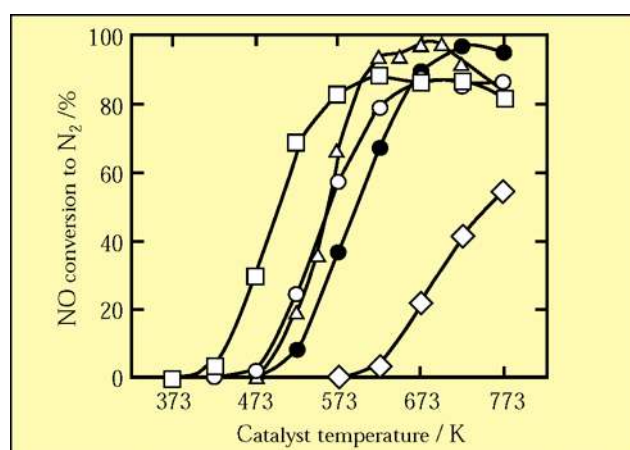


Figure 9 Effect of mechanical mixing Mn_2O_3 to Au/Al_2O_3 on the reduction of NO with C_1 - C_3 hydrocarbons. (◇), CH_4 ; (□), C_2H_4 ; (○), C_2H_6 ; (●), C_3H_8 ; (△), C_3H_6 ; $Au(0.17 \text{ wt.}\%)/Al_2O_3 : Mn_2O_3 = 19:1$ (in weight). Reaction gas: 0.1 vol% NO, 0.1 vol% HC, 5.0 vol% O_2 , and 10 vol% H_2O in a helium background. Space velocity = $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

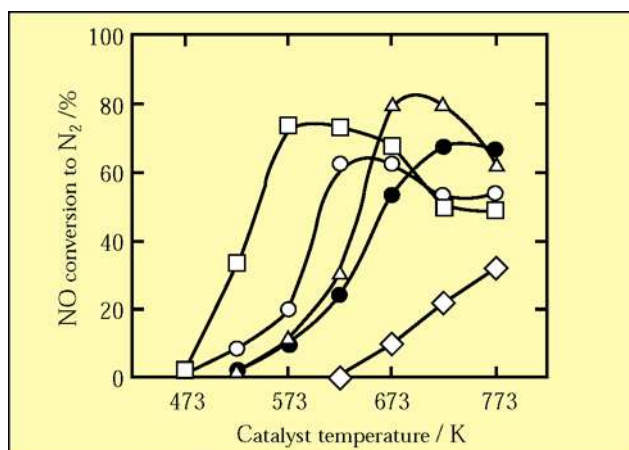


Figure 8 Temperature dependence of NO conversion to N_2 over Au/Al_2O_3 using C_1 - C_3 hydrocarbons. (◇), CH_4 ; (□), C_2H_4 ; (○), C_2H_6 ; (●), C_3H_8 ; (△), C_3H_6 . Gold loading was 0.2 wt%. Reaction gas: 0.1 vol% NO, 0.1 vol% HC, 5.0 vol% O_2 , and 10 vol% H_2O in a helium background. Space velocity = $20,000 \text{ h}^{-1} \text{ cm}^3 \text{ g-catalyst}^{-1}$

modified gold catalyst can maintain its catalytic performance under such severe conditions as 20 vol% oxygen and 10 vol% water vapour.

Other Hydrocarbons as Reductants

Not only propene, but also ethene, ethane and propane

can be used as effective reductants for NO on Au/Al_2O_3 in the presence of oxygen and moisture (47). Figure 8 shows the conversion of NO to N_2 on the Au/Al_2O_3 catalyst as a function of temperature when C_1 - C_3 hydrocarbons are used as reductants. Ethene shows the highest reactivity for NO reduction, but methane requires reaction temperatures higher than 623 K. The light-off temperature for the reduction of NO became higher in the following order of hydrocarbons; ethene < ethane < propene, propane < methane. In the case of ethene, as an exception, the direct reaction pathway between NO and ethene may also take place at temperatures below 523K in addition to the route via NO_2 as intermediate, because NO oxidation with O_2 over Au/Al_2O_3 is very slow at lower temperatures.

As shown in Figure 9, the mechanical mixing of Mn_2O_3 with Au/Al_2O_3 enhances the NO reduction performance even when other hydrocarbons are used instead of propene.

CONCLUSIONS

Gold supported on a variety of metal oxides shows good catalytic performance for the reduction of nitric oxide by hydrocarbons such as propene, propane, ethene and ethane. The addition of carbon monoxide

and hydrogen to the reactant gas can improve the conversion of nitric oxide (NO) to nitrogen. Reaction temperature regions and nitric oxide conversion are widely affected by the selection of metal oxide supports. Zinc oxide support makes gold the most active at low temperatures. Amongst the gold catalysts evaluated, the alumina support showed the highest conversion to nitrogen, but at higher temperatures such as 673K. Reduction of nitric oxide on gold catalysts is enhanced by the presence of both oxygen (15 vol%) and water (10 vol%).

The reduction of nitric oxide to nitrogen is assumed to take place *via* the formation of nitrogen dioxide (NO₂), formed by the oxidation of nitric oxide with oxygen as the first and rate-determining step, and the nitrogen dioxide is then reduced by the hydrocarbon. The mechanical mixture of Au/Al₂O₃ with Mn₂O₃, which is the most active for nitric oxide oxidation, showed appreciably enhanced conversions to nitrogen at temperatures between 523K to 773K. It is not only competitive with Pt-MFI in its low-temperature activity but is also superior to most of the catalyst materials reported so far at higher temperatures. Although selectivity of propene to nitric oxide reduction against oxygen consumption (hydrocarbon efficiency) is inferior to other metal catalysts, supported gold presents one of the promising catalyst candidates for practical applications for emission control of stationary and road vehicle engines.

Over gold catalysts, the relative order of reactivity is O₂ + CO > NO₂ + CO > NO + CO > O₂ + H₂ > NO₂ + H₂ > NO + O₂ > NO₂ + C₃H₆ > O₂ + C₂H₆ > NO + C₃H₆. The combination of NO oxidation to NO₂ and NO₂ reduction by reducing gases such as carbon monoxide, hydrogen, and hydrocarbons should be taken into account when seeking further improvements in the performance of gold-based catalysts.

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