

# Lawrence Berkeley National Laboratory

## Recent Work

### Title

No influence of CO<sub>2</sub> on stable isotope analyses of soil waters with off-axis integrated cavity output spectroscopy (OA-ICOS).

### Permalink

<https://escholarship.org/uc/item/88t9v11j>

### Journal

Rapid communications in mass spectrometry : RCM, 31(5)

### ISSN

0951-4198

### Authors

Sprenger, Matthias  
Tetzlaff, Doerthe  
Soulsby, Chris

### Publication Date

2017-03-01

### DOI

10.1002/rcm.7815

Peer reviewed

*Rapid Commun. Mass Spectrom.* **2017**, *31*, 430–436  
(wileyonlinelibrary.com) DOI: 10.1002/rcm.7815

# No influence of CO<sub>2</sub> on stable isotope analyses of soil waters with off-axis integrated cavity output spectroscopy (OA-ICOS)

Matthias Sprenger\* , Doerthe Tetzlaff  and Chris Soulsby

Northern Rivers Institute, School of Geosciences, University of Aberdeen, Elphinstone Road, Aberdeen AB24 3UF, UK

**RATIONALE:** It was recently shown that the presence of CO<sub>2</sub> affects the stable isotope ( $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  values) analysis of water vapor via Wavelength-Scanned Cavity Ring-Down Spectroscopy. Here, we test how much CO<sub>2</sub> is emitted from soil samples and if the CO<sub>2</sub> in the headspace influences the isotope analysis with the direct equilibration method by Off-Axis Integrated Cavity Output Spectroscopy (OA-ICOS).

**METHODS:** The headspace above different amounts of sparkling water was sampled, and its stable isotopic composition ( $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  values) and CO<sub>2</sub> concentration were measured by direct equilibration and by gas chromatography, respectively. In addition, the headspace above soil samples was analyzed in the same way. Furthermore, the gravimetric water content and the loss on ignition were measured for the soil samples.

**RESULTS:** The experiment with the sparkling water showed that CO<sub>2</sub> does not influence the stable isotope analysis by OA-ICOS. CO<sub>2</sub> was emitted from the soil samples and correlated with the isotopic fractionation signal, but no causal relationship between the two was determined. Instead, the fractionation signal in pore water isotopes can be explained by soil evaporation and the CO<sub>2</sub> can be related to soil moisture and organic matter which both enhance microbial activity.

**CONCLUSIONS:** We found, despite the high CO<sub>2</sub> emissions from soil samples, no need for a post-correction of the pore water stable isotope analysis results, since there is no relation between CO<sub>2</sub> concentrations and the stable isotope results of vapor samples obtained with OA-ICOS. © 2016 The Authors. *Rapid Communications in Mass Spectrometry* Published by John Wiley & Sons Ltd.

Stable isotope analysis of soil pore waters ( $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  values) is increasingly being applied in various studies dealing with, for example, flow paths on hillslopes,<sup>[1,2]</sup> travel times in the vadose zone,<sup>[3,4]</sup> and root water uptake.<sup>[5,6]</sup> Several methods for the analysis of the stable isotopes of soil pore waters exist,<sup>[7]</sup> but differences in their results are not yet fully understood.<sup>[7,8]</sup> Of the available methods, the application of the direct equilibration for stable isotope analysis of pore waters<sup>[9]</sup> is increasingly used.<sup>[2–4,10–12]</sup> However, because this direct equilibration method uses laser spectrometry to determine the isotopic composition of the headspace in equilibrium with the soil pore water, volatile compounds can potentially alter the analysis by spectral interferences, when they absorb the laser in similar wavelengths to the isotopologues of water.<sup>[13,14]</sup>

Such a spectral interference was recently found for CO<sub>2</sub> during isotope analyses with the laser-based technique of wavelength-scanned cavity ring-down spectrometry

(WS-CRDS):<sup>[15]</sup> there is an apparent linear relationship between the offset of the isotopic analysis of a standard water vapor and the amount of CO<sub>2</sub> in the vapor. While the  $\delta^{18}\text{O}$  values were overestimated with higher CO<sub>2</sub> concentration, the  $\delta^2\text{H}$  values were underestimated.<sup>[15]</sup> This offset induced by CO<sub>2</sub>, resulting in increased  $\delta^2\text{H}$  values and reduced  $\delta^{18}\text{O}$  values, would directly affect estimates of the so-called deuterium excess (d-excess).<sup>[16]</sup>

The d-excess was introduced by Dansgaard<sup>[16]</sup> as an index for non-equilibrium conditions and is defined as:

$$\text{d-excess} = \delta^2\text{H} - 8 \times \delta^{18}\text{O} \quad (1)$$

The isotopic signal of precipitation has a global average d-excess of about 10 ‰, which results from equilibrium Rayleigh condensation from vapor that was evaporated in a non-equilibrium process from seawater.<sup>[16]</sup> When soil water that originated from precipitation (d-excess = 10 ‰) evaporates, the residual water in the soil will have a d-excess <10 ‰ due to non-equilibrium processes (kinetic fractionation) and the  $\delta^{18}\text{O}$ - $\delta^2\text{H}$  relationship of the water will have a slope <8 in the dual isotope space (the regression line is then called the evaporation line).<sup>[17]</sup> Thus, the offset from the real isotope value induced by CO<sub>2</sub> could be misinterpreted as a signal for evaporation of soil water. Since the  $\delta^{18}\text{O}$ - $\delta^2\text{H}$  relationship in precipitation varies locally, depending on various factors (e.g., temperature and humidity during cloud generation and condensation, altitude effects, continental

\* Correspondence to: M. Sprenger, School of Geosciences, University of Aberdeen, Elphinstone Road, Aberdeen AB24 3UF, UK.  
E-mail: matthias.sprenger@abdn.ac.uk

This is an open access article under the terms of the Creative Commons Attribution License, which permits use, distribution and reproduction in any medium, provided the original work is properly cited.

effects, latitude effects)<sup>[18]</sup> we use the line-conditioned excess (lc-excess) defined as:<sup>[19]</sup>

$$\text{lc-excess} = \delta^2\text{H} - a \times \delta^{18}\text{O} - b \quad (2)$$

where *a* and *b* represent the slope and intercept of the local meteoric water line (LMWL). For precipitation, the lc-excess has an average value of 0 ‰ and soil water plotting below the LMWL will be of lc-excess < 0.

So far, the effect of CO<sub>2</sub> on stable isotope analysis has only been tested for WS-CRDS<sup>[15]</sup> and there is a need to assess if or how CO<sub>2</sub> also affects the analysis with Off-Axis Integrated Cavity Output Spectroscopy (OA-ICOS). Both analyzer systems use the absorption of a near-infrared laser beam by molecules (i.e. isotopologues) in a gaseous sample in a high-finesse optical cavity.<sup>[20,21]</sup> However, directing the laser beam off-axis allows spatial separation of the multiple reflections within the cavity,<sup>[22]</sup> which results in fully resolved OA-ICOS absorption spectra.<sup>[23]</sup> Therefore, unlike WS-CRDS, OA-ICOS does not derive the isotope ratios in the cavity from discrete wavelength sampling measurements. Instead, the isotope ratios are derived from the integrated areas under fully resolved absorption spectra (D. S. Baer, Los Gatos Research, personal communication). How this different instrumental set up for OA-ICOS may affect the influence of CO<sub>2</sub> on the measurements of stable isotopes of water vapor has not yet been studied.

As microbial activity is known to respond to temperature changes within a few hours<sup>[24]</sup> and the aeration of the soil resulting from taking disturbed soil samples enhances soil respiration,<sup>[25]</sup> it is likely that CO<sub>2</sub> is emitted from the soil samples taken into the laboratory. This will probably be of special relevance if soils contain high levels of organic matter, as aeration and temperature increase have been shown to stimulate CO<sub>2</sub> production in peaty soils.<sup>[26]</sup> To date, it has not been tested whether CO<sub>2</sub> is emitted while soil samples are stored prior to isotope analysis.

We address the two following research questions in this study. (1) Does CO<sub>2</sub> alter stable isotope analysis by the direct equilibration method? (2) Do soil samples emit CO<sub>2</sub> during storage in the laboratory?

## EXPERIMENTAL

### Set up

The experimental set up follows two tracks: First, we determined the isotopic composition of different volumes of sparkling water that were de-gassing different amounts of CO<sub>2</sub> to assess the influence of CO<sub>2</sub> on the isotope analysis (Sparkling water experiment). Secondly, we conducted analysis of the pore water stable isotopes of field-moist soil samples and measured the CO<sub>2</sub> emissions from the samples in the lab (Soil water experiment).

For the sparkling water experiment, five samples of sparkling water with known isotopic composition (measured in liquid mode:  $\delta^{18}\text{O} = -7.45 \pm 0.1 \text{ ‰}$  and  $\delta^2\text{H} = -50.8 \pm 0.4 \text{ ‰}$ , *n* = 3), but different volumes of sparkling water, were prepared in airtight bags (Weber Packaging, Güglingen, Germany). Another sample was prepared with non-sparkling (still)

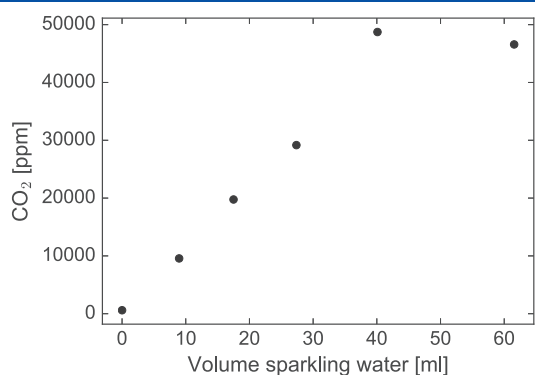
water from the same spring having the same isotopic composition (measured in liquid mode:  $\delta^{18}\text{O} = -7.43 \pm 0.1 \text{ ‰}$  and  $\delta^2\text{H} = -50.7 \pm 0.4 \text{ ‰}$ , *n* = 3). The isotopic composition of the water was determined by Off-Axis Integrated Cavity Output Spectroscopy (OA-ICOS) (triple water-vapor isotope analyzer TWIA-45-EP, Model#: 912-0032-0000, Serial#: 14-0038, Manufactured: 03/2014, Los Gatos Research, Inc., San Jose, CA, USA) running in liquid mode with a precision of  $\pm 0.4 \text{ ‰}$  for  $\delta^2\text{H}$  values and  $\pm 0.1 \text{ ‰}$  for  $\delta^{18}\text{O}$  values, as given by the manufacturer. The bags with the sparkling water were inflated with pressurized dry air before immediately then heat-sealing them. The pressurized dry air had a CO<sub>2</sub> concentration of about 600 ppm. The bags were stored for 2 days under constant temperature to allow for an equilibration between the water and the headspace in the bag until the stable isotope analysis were conducted as described below. The different water volumes of sparkling water exhausted different amounts of CO<sub>2</sub> during the 2 days of storage before isotope and CO<sub>2</sub> analyses (Table 1, Fig. 1).

To compare the isotope data from the non-sparkling water ( $\delta_{\text{nsw}}$  values) with those for the sparkling water ( $\delta_{\text{sw}}$  values), we calculated the differences as  $\Delta\delta = \delta_{\text{sw}} - \delta_{\text{nsw}}$ .

For the soil water experiment, 24 soil samples were taken from peaty podzols of the Bruntland Burn experimental catchment in the Scottish Highlands. A detailed description of the hydrometric and isotopic dynamics of the study site was published previously.<sup>[27]</sup> The soil sampling took place within the upper 20 cm of the profile, which is characterized by high organic matter content (17–80 %)<sup>[28]</sup> and low bulk density ( $0.76 \pm 0.21 \text{ g cm}^{-3}$ ).<sup>[28]</sup> We sampled depth profiles at eight different locations with each four samples taken in 5 cm increments (Table 2). Each sample consisted of 100–300 g of field-wet soil that was stored in airtight bags (Weber Packaging) until performance of the isotope analysis in the lab. When sampling the soil in the field, it was ensured that as little air as possible was present in the bags with the soil samples by manually furling them (no vacuum applied). In this way, the isotopic equilibration and CO<sub>2</sub> exhaust from the soil were limited to the 2 days of controlled storage after adding dry air in the laboratory. Thus, the CO<sub>2</sub> measured after the isotope analysis would stem from the 2 days of storage prior to the analysis. The isotope analyses described below were performed within 1 week after soil sampling.

**Table 1.** Sample characteristics of the sparkling water experiment with the volume of sample, the isotopic composition ( $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  values) of the equilibrated headspace above the sparkling water samples and the CO<sub>2</sub> concentration measured in the headspace

Sample ID	Non-sparkling water [mL]	Sparkling water [mL]	$\delta^{18}\text{O}$ [‰]	$\delta^2\text{H}$ [‰]	CO <sub>2</sub> [ppm]
$\delta_{\text{nsw}}$	22.7	0	-7.5	-50.8	600
$\delta_{\text{sw1}}$	0	9.0	-7.5	-50.2	9,563
$\delta_{\text{sw2}}$	0	17.5	-7.2	-49.4	19,756
$\delta_{\text{sw3}}$	0	27.4	-7.3	-50.1	29,163
$\delta_{\text{sw4}}$	0	40.1	-7.1	-50.2	48,726
$\delta_{\text{sw5}}$	0	61.6	-7.4	-50.7	46,578



**Figure 1.** CO<sub>2</sub> concentration in the headspace above different volumes of sparkling water after 2 days of storage.

### Isotope analysis

The stable isotopes of water in the pore space of the soil samples and the sparkling water samples were analyzed by the direct equilibration method.<sup>[9]</sup> Dry air was added to all samples in the airtight bags after which the bags were heat-sealed and stored for 2 days under constant temperature to allow for an equilibration between the (soil) water sample and the headspace in the bag. Along with the soil and sparkling water samples, standard waters of known isotopic composition were prepared in the same way to derive, via calibration from water vapor isotope data, the liquid-phase isotopic composition relative to the Vienna Standard Mean Ocean Water. The standard waters used for the calibration were seawater ( $\delta^{18}\text{O} = -0.85 \pm 0.1 \text{ ‰}$  and  $\delta^2\text{H} = -5.1 \pm 0.4$

$\text{‰}$ ,  $n = 4$ ), Aberdeen tap water ( $\delta^{18}\text{O} = -8.59 \pm 0.1 \text{ ‰}$  and  $\delta^2\text{H} = -57.7 \pm 0.4 \text{ ‰}$ ,  $n = 4$ ), and condensate of distilled tap water ( $\delta^{18}\text{O} = -11.28 \pm 0.1 \text{ ‰}$  and  $\delta^2\text{H} = -71.8 \pm 0.4 \text{ ‰}$ ,  $n = 4$ ). The liquid water isotopic composition of these standard waters was determined with the above-mentioned OA-ICOS analyzer running in liquid mode. The seawater was freed from salt by distillation with a rotary evaporator to exclude the effect of salt on the direct equilibration method. The prepared standard waters were analyzed at the beginning, in the middle, and at the end of each day of isotope analysis with the direct equilibration method. We assessed the precision of the analyses from the variation of the three measurements for each of the three standard waters over 29 days of analyses during the last year (261 standard water analyses in total). The average standard deviation for the isotope analysis with the direct equilibration method was found to be 0.54 ‰ for  $\delta^{18}\text{O}$  values and 1.39 ‰ for  $\delta^2\text{H}$  values.

After the 2 days of equilibration at  $23 \pm 1^\circ\text{C}$ , all the samples of the soil water experiment and the sparkling water experiment were analyzed subsequently by sampling the headspace in the bags with a needle, while a silicone seal outside the bags served as a septum to prevent laboratory air from entering the bag. A tube was connected to the needle to directly route the vapor to the OA-ICOS instrument running in vapor mode. No carrier gas is needed for this instrument running in vapor mode, since the headspace from the sample bags is directly sucked into the cavity and not diluted. The continuous measurements of  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  values at 1 Hz were performed for 6 min per sample, and a plateau of stable values for the water vapor concentration in the cavity [ppm], the  $\delta^{18}\text{O}$  values [‰], and the  $\delta^2\text{H}$  values [‰] was reached within 3 min.

**Table 2.** Characterization of the soil samples, sampling depth, stable isotopic composition of the pore waters ( $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  values), CO<sub>2</sub> concentration in the headspace during the isotope analysis, gravimetric water content (GWC) loss on ignition (LOI) of the soil samples, and sampling date

Sample no.	Depth [cm]	$\delta^{18}\text{O}$ [‰]	$\delta^2\text{H}$ [‰]	CO <sub>2</sub> [ppm]	GWC [-]	LOI [%]	Sampling date
1	0–5	-6.0	-49.3	45,295	0.68	94	04.08.2016
2	5–10	-5.9	-48.4	49,917	0.70	81	04.08.2016
3	10–15	-6.1	-48.0	31,811	0.57	67	04.08.2016
4	15–20	-7.8	-54.1	10,790	0.33	29	04.08.2016
5	0–5	-6.5	-52.1	32,962	0.65	6	04.08.2016
6	5–10	-6.8	-49.7	33,597	0.28	9	04.08.2016
7	10–15	-7.5	-54.1	18,159	0.26	10	04.08.2016
8	15–20	-6.9	-52.4	10,434	0.32	24	04.08.2016
9	0–5	-6.2	-50.8	44,044	0.49	16	04.08.2016
10	5–10	-6.3	-48.5	19,077	0.31	9	04.08.2016
11	10–15	-7.3	-52.7	23,404	0.29	11	04.08.2016
12	15–20	-7.9	-55.5	15,769	0.25	10	04.08.2016
13	0–5	-6.0	-48.6	42,262	0.59	51	04.08.2016
14	5–10	-5.4	-42.9	28,135	0.32	12	04.08.2016
15	10–15	-6.8	-49.4	15,511	0.21	5	04.08.2016
16	15–20	-7.5	-53.8	5,717	0.28	10	04.08.2016
17	0–5	-5.2	-42.7	41,366	0.67	80	26.07.2016
18	5–10	-5.5	-45.0	26,940	0.52	49	26.07.2016
19	10–15	-6.6	-49.9	11,915	0.36	24	26.07.2016
20	15–20	-7.1	-50.4	2,085	0.16	5	26.07.2016
21	0–5	-5.9	-43.7	18,398	0.16	11	26.07.2016
22	5–10	-6.3	-49.3	15,386	0.12	10	26.07.2016
23	10–15	-7.9	-56.0	29,312	0.35	51	26.07.2016
24	15–20	-8.4	-56.1	21,288	0.26	20	26.07.2016

We averaged the values for  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  over the last 2 min of the 6 min of the continuous measurements to represent the isotopic composition of the sample. The standard deviation over this last 2 min of integration time after allowing for 4 min to stabilize was always  $<0.25\text{‰}$  for  $\delta^{18}\text{O}$  values and  $<0.55\text{‰}$  for  $\delta^2\text{H}$  values. The water vapor concentration at the plateau of the isotope analysis was about 32,000 ppm and its standard deviation over the last 2 min that were used to derive the isotope values was always  $<90$  ppm. To prevent carry-over effects between different samples, the water vapor concentration in the cavity was reduced to  $<200$  ppm with dry air before each individual analysis.

The TWIA-45-EP triple water-vapor isotope analyzer saves time series of various parameters in addition to the above introduced water vapor concentration and the stable isotopes of water. One dimensionless parameter is called 'H2Ob\_10\_PT\_B' and relates to the width of the absorption peaks (R. Provencal, Los Gatos Research, personal communication).

### CO<sub>2</sub> analysis, GWC, and LOI

Directly after the stable isotope analysis – and, thus, after 2 days of equilibration in the sealed bag with dry headspace – the headspaces of all the soil and sparkling water samples were analyzed for their CO<sub>2</sub> concentration by gas chromatography (GC) with a flame ionization detector (CP-9001, Chrompack, Raritan, NJ, USA). A Porapak QS column was used for the GC and the packing was Hayesep Q (60–80 mesh, 2.0 m  $\times$  1.8"  $\times$  2 mm SS). For the analysis of the headspace from the sample bags, 0.5 mL of the sample vapor was manually directly injected straight onto the column with a syringe. Prior to the CO<sub>2</sub> analyses of the headspace of the sparkling water and soil sample bags, a calibration was conducted to relate the area under a GC peak to known CO<sub>2</sub> concentrations. To do so, 0.5 mL of 350, 1000, 3000, 5000, and 10000 ppm concentrations was each injected three times and an average of the resulting area under the GC peak was calculated for each concentration. The resulting linear relationship ( $r = 0.99$ ,  $p < 0.1$ ) was then used to infer the CO<sub>2</sub> concentration from the area under the GC peak via linear regression. The precision of the CO<sub>2</sub> analyses was estimated from the average standard deviation of the area below the GC peak for the three injections of each of the five different standards of known CO<sub>2</sub> concentration used for the calibration. The precision for the analyses of the sparkling water samples (Table 1) and the soil samples 1–16 (Table 2) was 94 ppm and for the soil samples 17–24 (measured on a different day) it was 49 ppm. Since the CO<sub>2</sub> concentrations in the soil and sparkling water samples exceeded the maximum measurement range of the instrument (10,000 ppm), all samples were diluted 1:12 with helium.

The gravimetric water content (GWC) was estimated by oven drying at 105°C overnight and relating the water loss to the field-moist soil mass. The loss on ignition (LOI) was determined by igniting about 10 g of the previously oven-dried soil in a muffle-furnace at 550°C over 2 h and relating the weight loss to the initial dried soil sample.

To statistically test our data for linear relationships, we applied the Pearson correlation (`scipy.stats.pearsonr` in Python) for the sparkling water experiment, since the data was normally distributed according to the Shapiro–Wilk test

(`scipy.stats.shapiro` in Python) ( $p$ -values for the Shapiro–Wilk test for CO<sub>2</sub> ( $p = 0.61$ ),  $\delta^{18}\text{O}$  ( $p = 0.41$ ),  $\delta^2\text{H}$  ( $p = 0.41$ ), and the width of the absorption peak ( $p = 0.75$ ). We applied the Spearman rank-order correlation (`scipy.stats.spearmanr` in Python) for the soil water experiment, since not all data was normally distributed according to the Shapiro–Wilk test ( $p$ -values for GWC ( $p = 0.02$ ), LOI ( $p < 0.01$ ), width of the absorption peak ( $p = 0.11$ ), and CO<sub>2</sub> ( $p = 0.52$ )).

## RESULTS AND DISCUSSION

### Does CO<sub>2</sub> alter the stable isotope analysis with the direct equilibration method?

The experimental set up of applying the direct equilibration method to different volumes of sparkling waters of known isotopic composition allowed us to directly assess the influence of different CO<sub>2</sub> concentrations in the headspace on the stable isotope analysis with OA-ICOS, since higher volumes of sparkling water emitted more CO<sub>2</sub> during the equilibration period (Fig. 1).

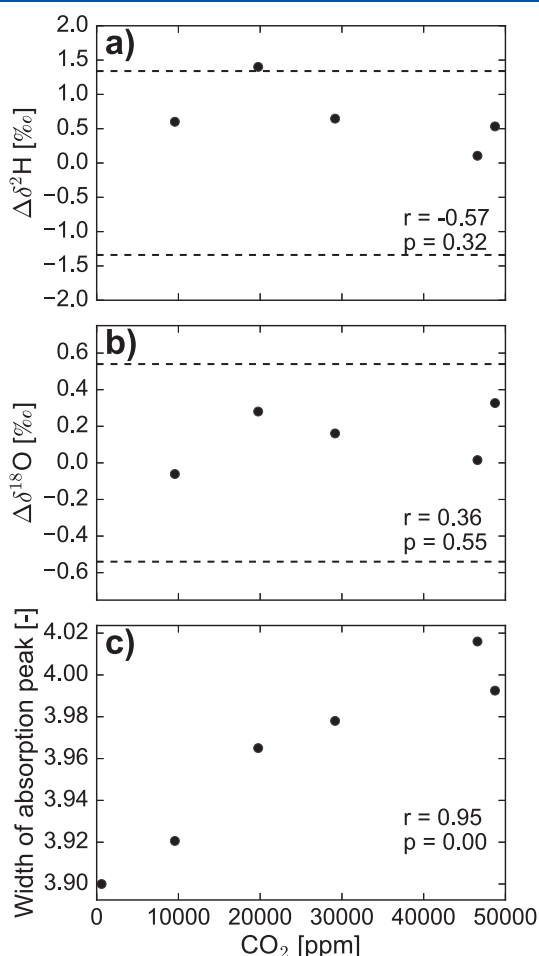
The differences in the analyzed stable isotope composition of the non-sparkling water and the five sparkling water samples of increasing volume leading to increasing CO<sub>2</sub> concentrations did not show a statistically significant relationship (Figs. 2(a) and 2(b)).

The differences between the measured  $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  values in the headspace of the non-sparkling water and those of the sparkling waters of various amounts ( $\Delta\delta$ ) was within the given accuracy range of the instrument for the direct equilibration method over the entire range between 10,000 and 50,000 ppm of CO<sub>2</sub> (dotted lines in Figs. 2(a) and 2(b)). Consequently, there was also no relationship between the CO<sub>2</sub> concentration and the  $l_c$ -excess.

However, there was a significant positive correlation between the parameter representing the width of the absorption peaks ('H2Ob\_10\_PT\_B') and the measured CO<sub>2</sub> in the headspace (Fig. 2(c)). This relationship shows that CO<sub>2</sub> in the cavity affects the measurements of the absorption spectra by widening the absorption peaks. However, the integrated area below the spectra is independent of that widening as long as the water vapor concentration in the cavity is constant (D. S. Baer, Los Gatos Research, personal communication). Since the water vapor concentration during the integration period of the isotope measurements did not vary between the samples, there was no effect of CO<sub>2</sub> on the measured isotope ratios.

As no relationship between the CO<sub>2</sub> concentration of vapor in the headspace and the stable isotopic composition of the headspace was observed, we have to reject the hypothesis that CO<sub>2</sub> in the headspace alters the stable isotope analysis results obtained with the direct equilibration method when conducted with an OA-ICOS instrument. Consequently, in contrast to CRDS, no post-correction of the sampled isotope data is required for the analysis by OA-ICOS. However, this finding will be limited to this specific OA-ICOS laser technology, where the spectral interference does not influence the measurements of isotope ratios. Whether the presence of CO<sub>2</sub> influences other laser spectroscopy techniques such as Fourier transform infrared (FTIR) or quantum cascade laser (QCL) has not yet been tested.



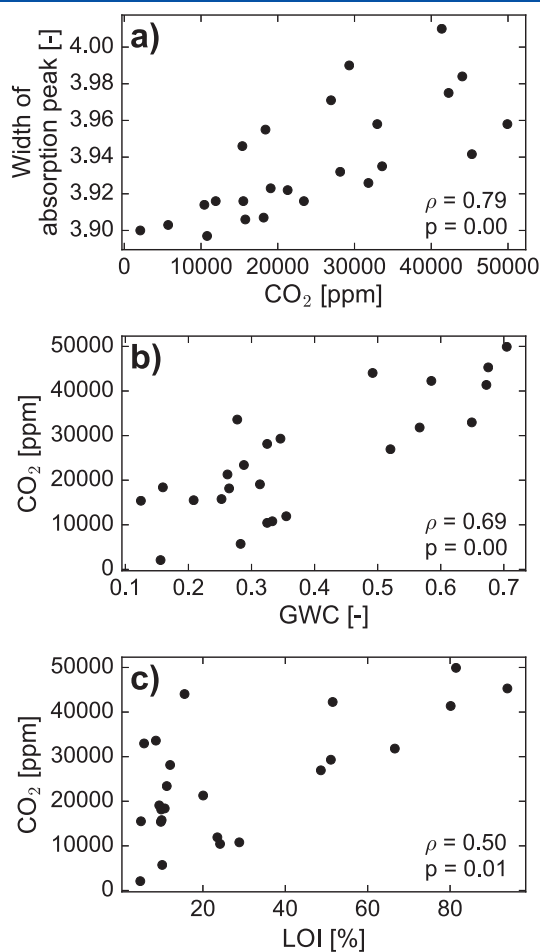


**Figure 2.** Difference between the measured values for (a)  $\delta^2\text{H}$  and (b)  $\delta^{18}\text{O}$  for non-sparkling and sparkling waters as a function of the  $\text{CO}_2$  concentration in the sampled headspace. The dashed lines indicate the measurement precision for the applied direct-equilibration method: 0.54 ‰ for  $\delta^{18}\text{O}$  values and 1.39 ‰ for  $\delta^2\text{H}$  values. (c) Relationship between width of the absorption peak (given as 'H2Ob\_10\_PT\_B' from the isotope analyzer) and the  $\text{CO}_2$  concentration in the sampled headspace. The Pearson correlation coefficient is given as  $r$  and the significance level is given as  $p$ .

### Do soil samples emit $\text{CO}_2$ during storage in the lab?

Similar to the sparkling water experiment, the soil water experiment also revealed that there is a relationship between the width of the absorption peak measured with OA-ICOS and the  $\text{CO}_2$  concentration in the sampled headspace above the soil sample (Fig. 3(a)). Thus,  $\text{CO}_2$  was emitted from the soil samples during the 2 days of equilibration prior to the isotope analysis. The  $\text{CO}_2$  concentrations varied between 2000 and almost 50,000 ppm (Table 2) and are, therefore, within the range of  $\text{CO}_2$  concentrations covered by the sparkling water experiment.

The  $\text{CO}_2$  concentrations correlate with the soil sampling depth ( $\rho = 0.74$ ,  $p < 0.01$ ) with a decrease in  $\text{CO}_2$  over soil depth. The  $\text{CO}_2$  concentration also correlates strongly with the GWC (Fig. 3(c)) and less strongly – but still significantly – with the LOI (Fig. 3(d)). A positive relationship between  $\text{CO}_2$  emissions and soil moisture has also been found in studies in the tropics,<sup>[29]</sup> in semiarid temperate steppes,<sup>[30]</sup>



**Figure 3.** Relationship between (a) the width of the absorption peak (given as 'H2Ob\_10\_PT\_B' by the isotope analyzer) and the  $\text{CO}_2$  concentration in the headspace during the stable isotope analysis, (b)  $\text{CO}_2$  concentration and the gravimetric water content (GWC), and (c)  $\text{CO}_2$  concentration and the loss on ignition (LOI). The Spearman rank-order correlation coefficient is given as  $\rho$  and the significance level is given as  $p$ .

savanna,<sup>[24]</sup> and lab experiments.<sup>[24]</sup> The LOI is probably a too broad measure of the labile soil organic matter available for microbial activity, since studies on peaty soils have shown that the quality of the organic matter plays an important role for decomposition rates in peaty soils.<sup>[26]</sup>

We cannot reject the hypothesis that  $\text{CO}_2$  is being generated in bags with soil samples during the 2 days of equilibration time prior to the isotope analysis. Instead, we see high concentrations of  $\text{CO}_2$  and can relate their variability within the soil profile to soil moisture and to some degree also to the amount of organic material in the soil.

### Implications for applications

For the presented soil samples, the lc-excess values correlate with soil depth ( $\rho = -0.79$ ,  $p < 0.01$ ) with lower lc-excess values in the shallow soil and lc-excess approaching zero at 15–20 cm soil depth. As shown with the sparkling water experiment, for the OA-ICOS instrument used, the  $\text{CO}_2$  concentration does not affect the isotope analysis. Therefore, we interpret the lc-excess pattern as kinetic fractionation of the soil water due to

soil evaporation under non-equilibrium processes. This fractionation is more pronounced in the shallow soil that has a more intense interaction with the atmosphere than at 15–20 cm soil depth, where humidity is higher and evaporative fluxes will be lower under the given climatic conditions of the study site. Consequently, the soil water isotopes plot along an evaporation line of slope 4.17. The decrease in the evaporation fractionation signal with depth is a general pattern shown in numerous previous studies.<sup>[31]</sup> For soils in temperate forests and temperate grasslands, the evaporation fractionation is usually limited to the upper 30 cm.<sup>[31]</sup>

Given the high CO<sub>2</sub> concentrations in the headspace above the soil samples, isotope analysis by WS-CRDS would require a post-analysis correction in order to prevent a misinterpretation of a more pronounced soil evaporation signal.

## CONCLUSIONS

We conclude that CO<sub>2</sub> in vapor samples does not affect the measurement of the stable isotopic composition of the water vapor by off-axis integrated cavity output spectroscopy (OA-ICOS). Therefore, no post-correction is needed when applying the direct equilibration method with OA-ICOS to determine the isotopic composition of soil water. However, as soil samples were shown to emit CO<sub>2</sub> into the headspace during the 2 days of equilibration prior to the isotope analysis, issues could arise when other laser spectrometry techniques, that are more sensitive to spectral interferences, are applied.

Furthermore, we conclude for the pore water isotopic composition of the soil samples from organic rich peaty podzols that the deviation from the LMWL does not result from spectral interferences between CO<sub>2</sub> and the isotopologues of water during the analysis. Instead, the location of the soil water isotope samples in the dual isotope plot indicates kinetic fractionation processes that result from evaporation.<sup>[31–33]</sup> The fact that the evaporation signal, given as the *lc*-excess in this study, gets lower with increasing soil depth strongly supports the above interpretation and is in agreement with numerous stable isotope studies that applied various analysis methods.<sup>[31]</sup>

## Acknowledgements

We are thankful for the support by Audrey Innes with the stable isotope, LOI, and GWC analysis. We thank Jonathan Dick for suggesting that we use sparkling water to generate different CO<sub>2</sub> concentrations in the headspace and Claire Tunaley for proof reading. We further highly appreciate the help of David Galloway and Michael McGibbon from the School of Biological Sciences, University of Aberdeen, with the CO<sub>2</sub> analysis. We are also thankful for the support by Robert Provencal and Doug S. Baer regarding the technical aspects of the isotope analyzer. We would also like to thank the European Research Council (ERC, Project No. GA 335910 VeWa) and the Natural Environment Research Council (NERC, Project No. NE/K000268/1) for funding. We thank three anonymous reviewers for their valuable feedback that helped to improve the manuscript.

## REFERENCES

- [1] J. Garvelmann, C. Külls, M. Weiler. A porewater-based stable isotope approach for the investigation of subsurface hydrological processes. *Hydrol. Earth Syst. Sci.* **2012**, *16*, 631.
- [2] M. H. Mueller, A. Alaoui, C. Kuells, H. Leistert, K. Meusburger, C. Stumpp, M. Weiler, C. Alewell. Tracking water pathways in steep hillslopes by  $\delta^{18}\text{O}$  depth profiles of soil water. *J. Hydrol.* **2014**, *519*, 340.
- [3] M. Sprenger, S. Seeger, T. Blume, M. Weiler. Travel times in the vadose zone: variability in space and time. *Water Resour. Res.* **2016**, *52*, 5727.
- [4] M. Sprenger, M. Erhardt, M. Riedel, M. Weiler. Historical tracking of nitrate in contrasting vineyards using water isotopes and nitrate depth profiles. *Agric. Ecosyst. Environ.* **2016**, *222*, 185.
- [5] M. Beyer, P. Koeniger, M. Gaj, J. T. Hamutoko, H. Wanke, T. Himmelsbach. A deuterium-based labeling technique for the investigation of rooting depths, water uptake dynamics and unsaturated zone water transport in semiarid environments. *J. Hydrol.* **2016**, *533*, 627.
- [6] T. H. M. Volkmann, K. Haberger, A. Gessler, M. Weiler. High-resolution isotope measurements resolve rapid ecohydrological dynamics at the soil-plant interface. *New Phytol.* **2016**, *210*, 839.
- [7] M. Sprenger, B. Herbstritt, M. Weiler. Established methods and new opportunities for pore water stable isotope analysis. *Hydrol. Processes* **2015**, *29*, 5174.
- [8] N. Orłowski, D. L. Pratt, J. J. McDonnell. Intercomparison of soil pore water extraction methods for stable isotope analysis. *Hydrol. Processes* **2016**, *30*, 3434.
- [9] L. Wassenaar, M. Hendry, V. Chostner, G. Lis. High resolution pore water  $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  measurements by  $\text{H}_2\text{O}_{(\text{liquid})} - \text{H}_2\text{O}_{(\text{vapor})}$  equilibration laser spectroscopy. *Environ. Sci. Technol.* **2008**, *42*, 9262.
- [10] M. J. Hendry, E. Schmeling, L. I. Wassenaar, S. L. Barbour, D. Pratt. Determining the stable isotope composition of pore water from saturated and unsaturated zone core: improvements to the direct vapour equilibration laser spectrometry method. *Hydrol. Earth Syst. Sci.* **2015**, *19*, 4427.
- [11] J. Klaus, E. Zehe, M. Elsner, C. Külls, J. J. McDonnell. Macropore flow of old water revisited: experimental insights from a tile-drained hillslope. *Hydrol. Earth Syst. Sci.* **2013**, *17*, 103.
- [12] G. Bertrand, J. Masini, N. Goldscheider, J. Meeks, V. Lavastre, H. Celle-Jeanton, J.-M. Gobat, D. Hunkeler. Determination of spatiotemporal variability of tree water uptake using stable isotopes ( $\delta^{18}\text{O}$ ,  $\delta^2\text{H}$ ) in an alluvial system supplied by a high-altitude watershed, Pfyn forest, Switzerland. *Ecohydrol.* **2012**, *7*, 319.
- [13] A. G. West, G. R. Goldsmith, P. D. Brooks, T. E. Dawson. Discrepancies between isotope ratio infrared spectroscopy and isotope ratio mass spectrometry for the stable isotope analysis of plant and soil waters. *Rapid Commun. Mass Spectrom.* **2010**, *24*, 1948.
- [14] W. A. Brand, H. Geilmann, E. R. Crosson, C. W. Rella. Cavity ring-down spectroscopy versus high-temperature conversion isotope ratio mass spectrometry; a case study on  $\delta^2\text{H}$  and  $\delta^{18}\text{O}$  of pure water samples and alcohol/water mixtures. *Rapid Commun. Mass Spectrom.* **2009**, *23*, 1879.
- [15] B. Gralher, B. Herbstritt, M. Weiler, L. I. Wassenaar, C. Stumpp. Correcting laser-based water stable isotope readings biased by carrier gas changes. *Environ. Sci. Technol.* **2016**, *50*, 7074.
- [16] W. Dansgaard. Stable isotopes in precipitation. *Tellus* **1964**, *16*, 436.

- [17] G. Allison. The relationship between  $^{18}\text{O}$  and deuterium in water in sand columns undergoing evaporation. *J. Hydrol.* **1982**, *55*, 163.
- [18] J. R. Gat. Oxygen and hydrogen isotopes in the hydrologic cycle. *Annu. Rev. Earth Planet. Sci.* **1996**, *24*, 225.
- [19] J. M. Landwehr, T. B. Coplen. Line-conditioned excess: a new method for characterizing stable hydrogen and oxygen isotope ratios in hydrologic systems, in *International Conference on Isotopes in Environmental Studies*. IAEA, Vienna, **2006**, p. 132.
- [20] E. R. Crosson. A cavity ring-down analyzer for measuring atmospheric levels of methane, carbon dioxide, and water vapor. *Appl. Phys. B* **2008**, *92*, 403.
- [21] D. S. Baer, J. B. Paul, M. Gupta, A. O'Keefe. Sensitive absorption measurements in the near-infrared region using off-axis integrated-cavity-output spectroscopy. *Appl. Phys. B* **2002**, *75*, 261.
- [22] J. B. Paul, L. Lapson, J. G. Anderson. Ultrasensitive absorption spectroscopy with a high-finesse optical cavity and off-axis alignment. *Appl. Opt.* **2001**, *40*, 4904.
- [23] Los Gatos Research. Off-Axis Integrated Cavity Output Spectroscopy (OA-ICOS). Available: [https://www.et.co.uk/docs/The\\_LGR\\_Advantage\\_ET.pdf](https://www.et.co.uk/docs/The_LGR_Advantage_ET.pdf).
- [24] J. Curiel Yuste, D. D. Baldocchi, A. Gershenson, A. Goldstein, L. Misson, S. Wong. Microbial soil respiration and its dependency on carbon inputs, soil temperature and moisture. *Global Change Biol.* **2007**, *13*, 2018.
- [25] J. Fang, K. Zhao, S. Liu. Factors affecting soil respiration in reference with temperature's role in the global scale. *Chin. Geogr. Sci.* **1998**, *8*, 246.
- [26] M. D. Preston, K. A. Smemo, J. W. McLaughlin, N. Basiliko. Peatland microbial communities and decomposition processes in the James Bay Lowlands, Canada. *Front. Microbiol.* **2012**, *3*, 70.
- [27] D. Tetzlaff, C. Birkel, J. Dick, J. Geris, C. Soulsby. Storage dynamics in hydropedological units control hillslope connectivity, runoff generation and the evolution of catchment transit time distributions. *Water Resour. Res.* **2014**, *50*, 969.
- [28] J. Geris, D. Tetzlaff, J. McDonnell, C. Soulsby. The relative role of soil type and tree cover on water storage and transmission in northern headwater catchments. *Hydrol. Processes* **2015**, *29*, 1844.
- [29] M. S. Johnson, E. G. Couto, O. B. Pinto Jr, J. Milesi, R. S. Santos Amorim, I. A. M. Messias, M. S. Biudes. Soil  $\text{CO}_2$  dynamics in a tree island soil of the Pantanal: the role of soil water potential. *PloS One* **2013**, *8*, e64874.
- [30] C. Zhao, Y. Miao, C. Yu, L. Zhu, F. Wang, L. Jiang, D. Hui, S. Wan. Soil microbial community composition and respiration along an experimental precipitation gradient in a semiarid steppe. *Sci. Rep.* **2016**, *6*, 24317.
- [31] M. Sprenger, H. Leistert, K. Gimbel, M. Weiler. Illuminating hydrological processes at the soil-vegetation-atmosphere interface with water stable isotopes. *Rev. Geophys.* **2016**, *54*, 674.
- [32] Y. Rothfuss, S. Merz, J. Vanderborght, N. Hermes, A. Weuthen, A. Pohlmeier, H. Vereecken, N. Brüggemann. Long-term and high-frequency non-destructive monitoring of water stable isotope profiles in an evaporating soil column. *Hydrol. Earth Syst. Sci.* **2015**, *19*, 4067.
- [33] G. B. Allison, C. Barnes, C. M. Hughes, F. Leaney. Effect of climate and vegetation on oxygen-18 and deuterium profiles in soils, in *Isotope Hydrology 1983*. IAEA, Vienna, **1984**, p. 105.