



Research Article

Photo catalytic degradation of methylene blue and methyl orange from aqueous solution using solar light onto chitosan bi-metal oxide composite

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Abstract

Chitosan bi-metal oxide (CBMO) is prepared by co-precipitation method to determine the degradation ability of methylene blue (MB) and methyl orange (MO) dyes using solar light irradiation. The characteristics of CBMO are studied by FTIR, UV, SEM, EDX and XRD studies. The photocatalytic degradation of studies MB and MO were carried out by varying the parameters such as pH, irradiation time, dye concentration, composite dosage. The photo catalytic study shows the maximum percentage of degradation at pH 8 for MB and 3 for MO at irradiation time of 120 min and 0.15 g of CBMO and is confirmed by UV spectral studies. Based on the equilibrium data value Langmuir isotherm is well fitted, which indicates the formation of strong monolayer and it obeys pseudo second order kinetic model. Among methylene blue and methyl orange dyes, methyl orange was highly degraded (74.05%) by the chitosan-bimetal oxide adsorbent. And also 62.05% of dyes can be removed from industrial waste water by CBMO. According to the experimental results it is clearly evident that the chitosan-bimetal oxide composite is expected to be an effective product for the remediation of dyes from waste water.

Keywords Chitosan bi-metal oxide · Degradation · Methylene blue · Methyl orange · Solar light · Langmuir isotherm · Second order kinetic model

1 Introduction

In the present scenario dyes, pigments and colour products play a vital role in various activities like food production, cosmetics, pharmaceuticals, leather, beverage and also in paper industries [1]. It contains complex structure, non-biodegradability and harmfulness towards animals, plants and environment. Around 700,000 varieties of dyes are produced per year in the worldwide market, in which most of the dyes are discharged from textile industries and contaminated the water resources [2, 3]. MB and MO are most commonly used cationic and anionic dyes in dyeing, printing, cosmetics and leather industries [4]. Releases of these dyes into aquatic steams restrict the penetration of

sunlight that leads to destroy the life of aquatic organisms. It causes several health hazards to human beings like vomiting, breathing problem, diarrhoea and skin irritation [5]. Now a day's degradation and removal of these dyes before discharging into the water bodies is the challenging task. Several methods are available to treat dye effluents such as coagulation, electro precipitation, reverse osmosis, adsorption, evaporation, flocculation, advanced oxidation process and photo catalytic degradation by using UV and solar light irradiation [6, 7]. Among all these methods degradation by solar light irradiation is the effective and easy method to remove dyes from dye effluent, because the solar energy is naturally available and does not produce sludge during degradation and quietly low

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cost when compared to other methods [8, 9]. According to previous literature number of attempts is made by many researchers for the degradation of MB and MO dye by photo catalytic degradation (UV/Solar irradiation) using various materials such as Cu₂O/LDH [10], Chitosan Zinc Oxide hybrid composite [11], BiFeO₃ [12], ZnO [13], CuCl₂ + STS, Chitosan-g-Poly (acrylamide)/ZnS [14], Chitosan/LaFe_{0.8}Cu_{0.2}O₃ and Xylan/LaFe_{0.8}Cu_{0.2}O₃ [15].

Chitosan is the second abundant natural polysaccharide derivative obtained from the deacetylation of chitin next to cellulose which has the capacity to degrade organic pollutants due to its numerous amide groups in its inter linked structure [16]. Modification of Chitosan by grafting process increases the degradation ability with its superior properties. Due to its inter and intra molecular hydrogen bonding between H⁺ and -NH₂, Chitosan dissolves only in acid medium [17]. Modified Chitosan composites are more suitable for anionic dyes than cationic dyes.

Generally the metal oxides possess good degradation ability towards organic pollutants due to its extensive properties like high surface area, low production and regeneration cost [18] and it is superior to other traditional adsorbents due to the absence of internal diffusion resistance. Based on previous literature studies many researchers involved in the preparation of bi-metal oxides in wastewater treatment, the reason is that it produces high degradation efficiency than mono metallic oxides; monometallic oxides are deactivated easily by pH alterations and it produces low degradation efficiency [19–22].

The objective of this present work is to prepare the modified chitosan by Zn–Mg bimetal oxide using grafting process (insertion of functional groups) and to examine the degradation ability towards MB and MO and to investigate the degradation ability of CBMO onto industrial waste water. Optimization of photocatalytic degradation parameters and characterisation of CBMO composites were examined.

2 Experimental

2.1 Materials

MB is a classical cationic dye with molecular formula C₁₆H₁₈N₃ClS, and λ_{max} 663 nm and MO dye is a typical anionic dye with molecular formula C₁₄H₁₄N₃NaO₃S with λ_{max} 464 nm; it is used as adsorbate to identify the degradation ability of CBMO. Analytical grade chemicals and reagents of Zinc Sulphate, Magnesium Sulphate, Chitosan, Acetic acid, HCl, NaOH are procured from Universal Scientific Company, Coimbatore, India. The entire study was carried out by using double distilled water. A stock solution of 1000 mg/L of MB and MO was prepared by dissolving 1 g

of dye in the double distilled water. Industrial wastewater is collected from Mangalam, Tirupur district.

2.2 Preparation of CBMO

Double distilled water is used to prepare equimolar solutions of Zinc Sulphate and Magnesium Sulphate and the solutions are mixed well and the mixture is stirred in magnetic stirrer for 3 h to obtain Zn-Mg Bimetal Oxide. The resultant product is washed well with double distilled water to reach neutral pH. The prepared product is filtered and dried at room temperature. The Zn-Mg Bimetal Oxide and the chitosan in 5% acetic acid solution is mixed thoroughly and the whole content is agitated in a magnetic stirrer for 2 h at room temperature [23]. After agitation the product is filtered, dried and calcinated in a hot air oven. The calcinated CBMO is crushed and packed in an air tight container for further studies.

2.3 Characterization of CBMO

The physiochemical characteristics such as moisture content, yield, pH, composite surface acidity and basicity, iodine number and the presence of carboxyl, carboxylic and phenolic groups are studied and reported. The surface characteristics such as SEM, EDX, XRD and FTIR also studied and reported for the prepared CBMO.

2.4 Photo catalytic studies

The photocatalytic degradation efficiency of CBMO is studied by varying irradiation time, pH, CBMO dose and initial dye concentration of MB and MO using batch mode experiments. It is carried out by using 250 ml borosil conical flasks. 50 ml of dye is taken in a series of 250 ml conical flasks with various dye concentration and known amount of CBMO is stirred well. The stirred mixture is agitated in a bench shaker for 10–15 min for the attainment of equilibrium between composite and adsorbate solution. 0.1 N HCl and 0.1 N NaOH is used to adjust the pH ranges from 2 to 11. After equilibrium attainment the whole content is exposed to sunlight irradiation. The absorbance is measured at regular interval using UV –VIS spectrophotometer 119 with regular interval ranges from 200–800 nm to determine the degradation percentage of MB and MO onto CBMO. The reaction is duplicated and the means values are measured and reported. The percentage of degradation efficiency is calculated by the following equation

$$\eta = \frac{A_0 - A_t}{A_0} \quad (1)$$

where A_0 is the initial absorbance of the dye and A_t is the absorbance of dye after degradation [24].

3 Results and discussion

3.1 Physiochemical characterization of CBMO

The Physiochemical characteristic of CBMO is studied and summarized in the Table 1. Moisture content is measure of the action of composite. If the moisture content is high it dilutes the action of composite and it needs extra load of composite dosage to degrade dye. From the results it is clear that the moisture content of CBMO is low (2.98%) compared to raw chitosan (7.84%) shows the better removal ability of dyes from wastewater than raw chitosan [25]. The surface acidity and basicity of CBMO is 4.05 mmoles/gm and 3.02 mmoles/gm respectively. The surface acidity and basicity is confirmed by Boehm titration, which quantifies the presence of oxygenated acidic groups and basic groups on modified composite. The Boehm titration results revealed the presence of phenyl carboxyl groups on its surface is measured in mill equivalent/gram [26]. The degradation ability of the composite is directly proportional to Iodine Number (1384.31 mg/g) and Yield (84.7%). If the iodine number and yield is high the degradation ability is also high due to large surface area. The surface area of raw chitosan is 1.46 m²/g; the result showed that the surface area of CBMO is rapidly increases by modification process. The surface area of CBMO is 41.9 m²/g enriches the uptake ability of dyes and heavy metals than raw chitosan [27]. The electrical neutrality of the composites is determined by zero point charge (pH_{zpc}). The result indicated that the pH_{zpc} of CBMO depend upon the raw material and activated agent. The pH_{zpc} of CBMO is 7.2, if the solution pH is below than pH_{zpc} (positively charged surface) enhances anionic dye

degradation and if the pH is above than pH_{zpc} (negatively charged surface) enhances cationic dye degradation [28].

3.2 Surface characterization of CBMO

The SEM image of the prepared CBMO at 1000× magnification is shown in Fig. 1. It represents the surface morphology of the prepared composite. CBMO surface has single phase primary particles and it has porous structure [29].

Qualitative confirmation of CBMO is studied by EDX technique as shown in the Fig. 2. The mass % and atomic mass % of CBMO is summarized in Table 2.

X-ray diffraction pattern of CBMO is shown in Fig. 3. The 2θ values lies between 20 and 30°, 45° and 65° indicates the composite is partially crystalline and amorphous in nature [30].

Characterization of functional groups is confirmed by FTIR spectral peaks. The sharp peak at 3254.67 cm⁻¹ indicates the presence of free -OH group. The wide bond lies between 2000 and 2021 cm⁻¹ represents the presence of overtone bands. The band at 1651.74 cm⁻¹ show the presence of N-H bending of primary amines, similar result was observed in the previous literature [31]. The peak

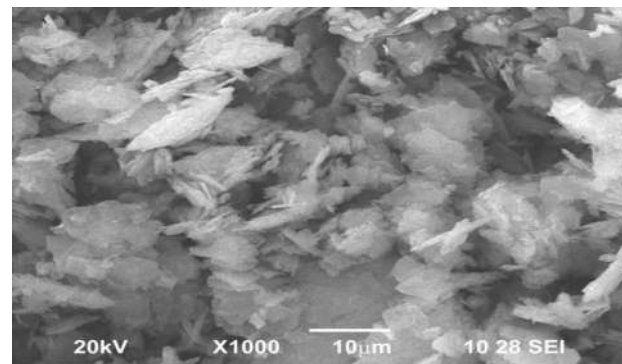


Fig. 1 SEM image of CBMO

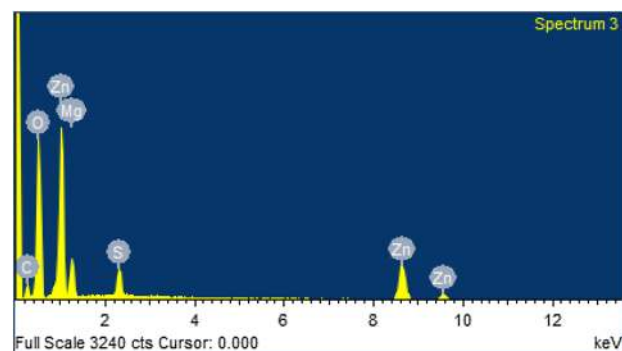


Fig. 2 EDX image of CBMO

Table 1 Physiochemical characteristics of CBMO

Parameters	CBMO
Yield (%)	84.7
Moisture content (%)	2.98
pH	6.8
Surface acidity (mmoles/g)	4.05
Surface basicity (mmoles/g)	3.02
pH_{zpc}	7.2
Boehm titration (m eq/gm) Basic sites	2.25
Phenolic and Carboxylic groups	1.28
Carboxyl groups	0.52
Iodine number (mg/g)	1384.31
Surface area (m ² /g)	41.9

Table 2 EDX elemental analysis of CBMO

Elements (CBMO)	Mass %	Atomic %
C	15.08	24.00
O	53.14	63.50
Mg	4.94	3.89
S	2.52	1.50
Zn	24.33	7.12

at 1377.47 cm^{-1} represents the stretching of C–N bond between chitosan and metal oxide [32]. The band ranges from 1070.71 to 1310.31 cm^{-1} indicating C=O stretching in alcohols and phenols [33] is shown in Fig. 4.

3.3 Photocatalytic studies for the degradation of MB and MO dye onto CBMO using solar light irradiation

3.3.1 Effect of irradiation time

The effect of irradiation time is the predominant factor to determine the degradation capacity of CBMO towards MB and MO. It is studied by the varied irradiation time from 20 to 140 min at optimum dose of 0.15 g of CBMO. From the results of photo catalytic studies, the initial dye concentration is fixed as 10 mg/L. The degradation percentage of MB is 67.43% and MO is 74.05% respectively. Degradation of the MB and MO increases with increase in irradiation time up to 120 min which attributed to higher number of active sites. After that there is no optimal change observed in the degradation efficiency since lesser number of active sites are available in the surface of CBMO [34]. The relationship between irradiation time

Fig. 3 XRD image of CBMO

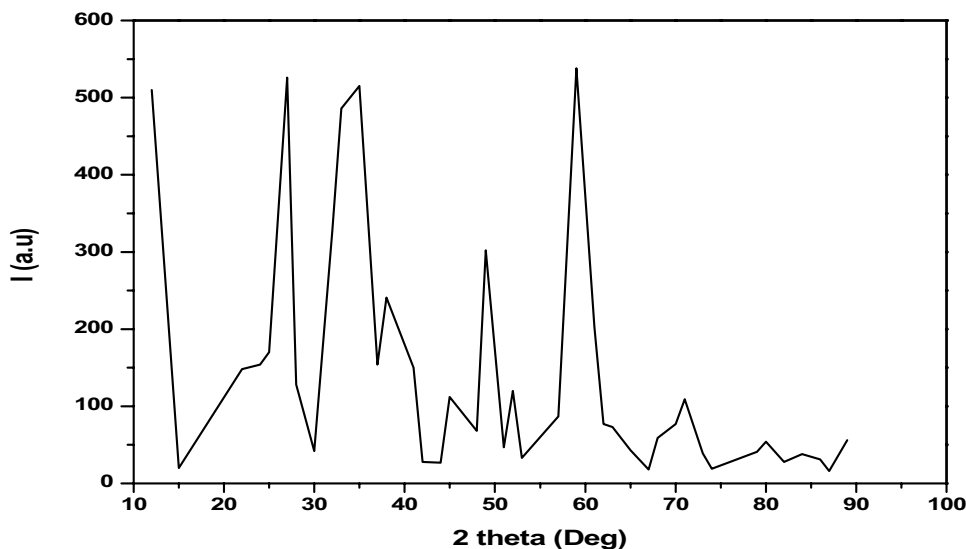
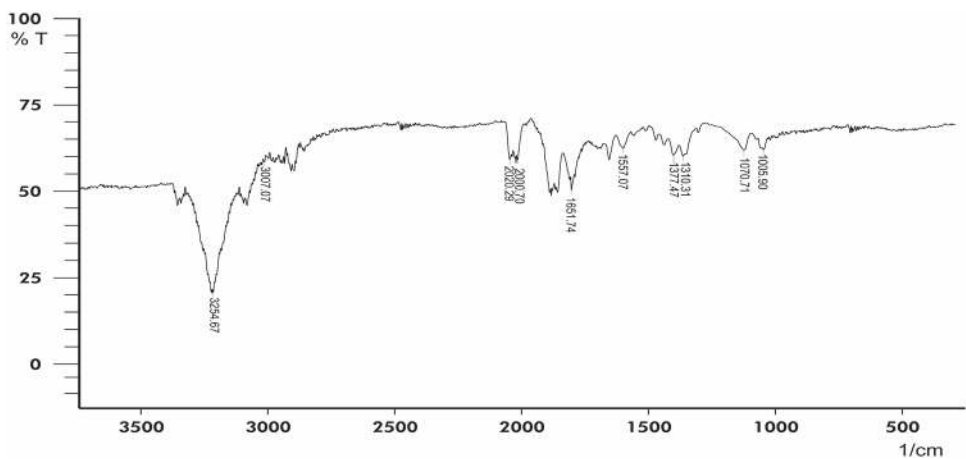


Fig. 4 FTIR image of CBMO



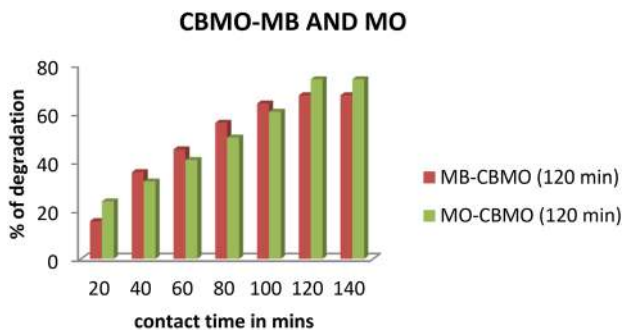


Fig. 5 Effect of irradiation time and % of degradation (CBMO dose=0.15 g, pH=8 for MB and 3 for MO and initial dye concentration 10 mg/L)

and % of degradation is shown in Fig. 5. The results indicate that the degradation of MO is higher than MB onto CBMO, therefore the composite favours anionic dye than cationic dye.

3.3.2 Effect of pH

The degradation of MB and MO onto CBMO as a function of hydrogen ion concentration is investigated from pH 2 to 11 at optimum dose of 0.15 g of composite and the irradiation time of 120 min. The solution pH is the primary tool to find the degradation efficiency of CBMO. The relationship between pH and % of degradation is shown in Fig. 6. Maximum percentage of degradation is 67.43% at pH 8 for MB and 74.05% at pH 3 for MO respectively. If zero point charge of CBMO (7.2) is lower than the solution pH, it enhances the methylene blue dye degradation because it possesses negative charge on its surface. Based on the observation of effect of pH, the zero point charge of the CBMO is higher than the solution pH it enriches the degradation of anionic dyes

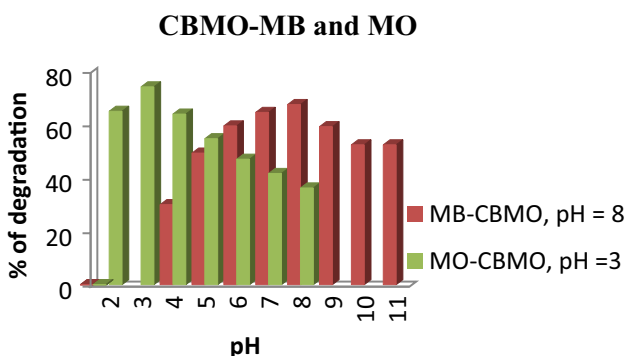


Fig. 6 Effect of pH and % of degradation (CBMO dose=0.15 g, irradiation time=120 min and initial dye concentration 10 mg/L)

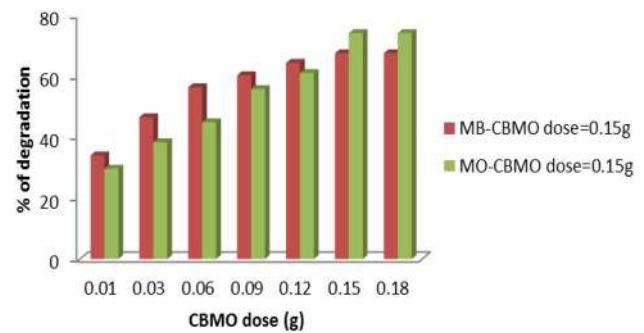


Fig. 7 Effect of CBMO dose and % of degradation (pH=8 for MB and 3 for MO, irradiation time=120 min and initial dye concentration 10 mg/L)

since the surface of the composite is positively charged [35]. Similar results are observed by several researchers using modified chitosan composites as a photo catalyst in dye removal process. From the result it may come to the conclusion that CBMO can be used as an effective photo catalyst to degrade MO than MB.

3.3.3 Effect of CBMO dose

The degradation capacity of MB and MO onto CBMO is studied by varying the amount of composite from 0.01 to 0.18 g at optimum conditions. The plot of CBMO dose and the % of degradation of MB and MO are shown in the Fig. 7. From the figure it is clear that the percentage of degradation increases steadily with increase in CBMO dose up to 0.15 g at lower concentration of dye solution. The maximum degradation of MB is 67.43% and 74.05% for MO. After this there is no considerable change in the percentage of degradation in MB and MO, this may be attributed that the formation of cloudy and non-transparent solution is due to the excess addition of CBMO with dye solution, this restrict the penetration of light into the dye solution [36]. Hence 0.15 g of the composite is fixed as optimum dose for entire study.

3.3.4 Effect of Initial dye concentration

Effect of initial dye concentration is analysed by varying the concentration of MB and MO concentration which ranges from 10 mg/L to 70 mg/L and the relationship between initial dye concentration and % of degradation is shown in Fig. 8. The % of degradation of MB dye decreased from 67.43 to 22.8% and MO decreased from 74.05 to 21.9%. When the concentration of dye increases the uptake capacity of CBMO decreases, because the formation of OH^- reduced at this stage, it is the prime and foremost reason for degradation of dye [37].

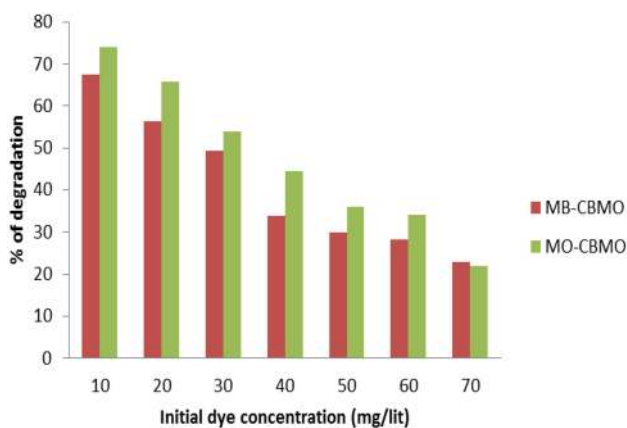


Fig. 8 Effect of Initial dye concentration on the degradation efficiency of CBMO (pH=8 for MB and 3 for MO CBMO dose=0.15 g and irradiation time= 120 min)

3.3.5 Conformation analysis for the degradation of methylene blue and methyl orange

The degradation of MB and MO by solar light irradiation is confirmed by UV-Visible spectroscopic studies as shown in Fig. 9a, b. The quantity of dye degraded is determined by using UV-VIS spectrophotometer 119 at 663 nm for MB and 464 nm for MO. The decrease in absorbance indicates the increase in % of dye degradation at regular interval time of 20 min at optimum conditions in the presence of solar light. The degradation of MB and MO onto CBMO before and after treatment is shown in Fig. 10a, b.

Many researchers have made attempt using various materials for the degradation of MB and MO are reported and summarized in the Table 3. The result shows that the degradation efficiency of CBMO is high with the contact time of 120 min, when compared to other studied materials using UV light and sunlight irradiation.

3.3.6 Comparative study of synthetic and industrial wastewater using CBMO as a photo catalyst

The degradation efficiency of CBMO is also analysed with industrial wastewater, which is collected from Mangalam, Tirupur district. The results of both synthetic and industrial wastewater onto CBMO are compared and summarized in Table 4. The results proved that the percentage degradation of CBMO towards synthetic and industrial wastewater is more or less equal indicates that the CBMO can also be used in industrial level as an effective photo catalyst. Degradation percentage is confirmed by UV- spectral studies for the industrial wastewater onto CBMO and is shown in the Fig. 11. The results of UV spectral analysis showed that there is a decrease in absorbance, indicates the % of dye degraded with respect to time onto CBMO.

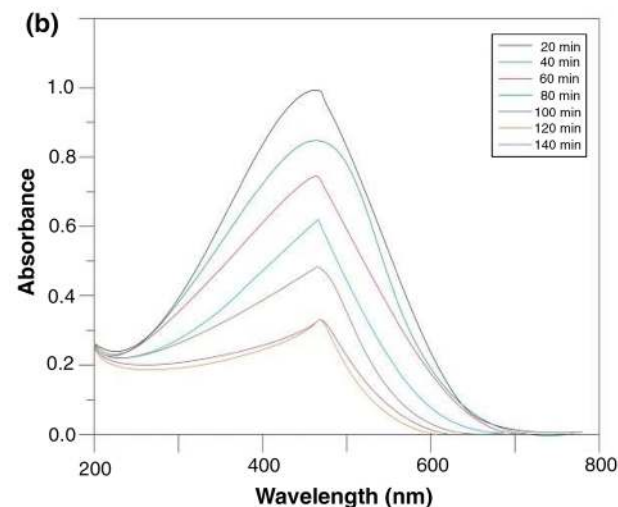
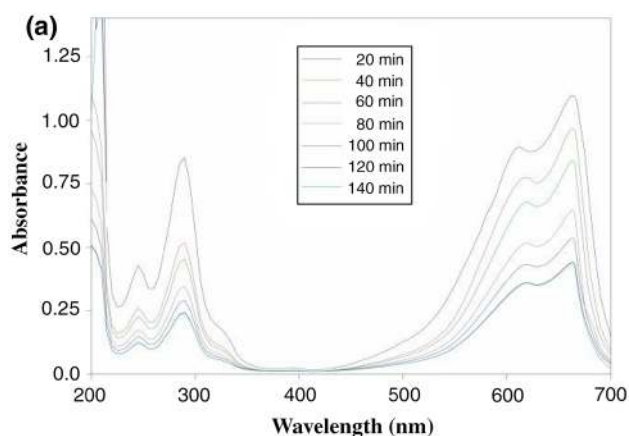


Fig. 9 UV visible sepctrum for the degradation of **a** MB onto CBMO, **b** MO onto CBMO

3.4 Adsorption isotherms

Based on the equilibrium data Freundlich and Langmuir isotherms are used to determine the maximum degradation capacity of MB and MO onto CBMO. The empirical form of Freundlich isotherm and the linear form of Langmuir equation is given below;

$$\text{Log}q_e = \frac{1}{n} \log (C_e) + \log K_f \tag{2}$$

$$\frac{C_e}{q_e} = \frac{1}{q_{max}}b + \frac{C_e}{q_{max}} \tag{3}$$

where q_e is the MB and MO dye concentration in CBMO (mg/g), $1/n$ is the heterogeneity factor, C_e is the concentration of dye solution (mg/L), K_f is the Freundlich constant (mg/g), q_{max} is the maximum degradation potential (mg/g) and b is the Langmuir equilibrium constant



Fig. 10 **a** Degradation of MB onto CBMO before and after treatment. **b** Degradation of MO onto CBMO before and after treatment

[38, 39]. Based on the result the correlation regression coefficient of MB (0.9919) and MO (0.9914) onto CBMO is higher for Langmuir isotherm than Freundlich isotherm model which revealed that the formation of strong mono layer and homogeneous surface during degradation process. The maximum degradation capacity (Q_m) of CBMO for MB is 5.9031 mg/g and for MO is 5.561 mg/g respectively. The Freundlich and Langmuir isotherm models are shown in the Figs. 12 and 13.

The results revealed that the Langmuir Isotherm model is best fit for the equilibrium data than Freundlich Isotherm model. The isotherm parameters are summarized in the Table 5.

3.5 Adsorption kinetics

The rate of the degradation and the mechanism of the adsorption of MB and MO can be analysed by pseudo first and second order kinetics. The linear form of pseudo first order reaction is given as [40, 41].

$$\log (q_e - q_t) = \log q_e - \frac{K_1 t}{2.303} \quad (4)$$

K_1 is the first order rate constant for the degradation of MB and MO onto CBMO is 0.0696 and 0.0712 min^{-1} . The correlation regression coefficient for first order kinetics is 0.853 for MB and 0.8031 for MO degradation. q_e is the uptake of dye at moment t is calculated as 16.59 mg/g and 18.21 mg/g for MB and MO onto CBMO. A graph is plotted

Table 3 Comparative study of methylene blue and methyl orange dye degradation using various materials

Name of the dye	Materials used	% of degradation and irradiation time	Light source
Methylene blue	Cu ₂ O/LDH (layered double hydroxide)	51.8% 6 h	Solar light irradiation
	Chitosan Zinc Oxide hybrid composite	64% 3 h	UV light irradiation
	BiFeO ₃	35% 9 h	Solar light irradiation
	ZnO	50% 4 h	Solar light irradiation
	CuCl ₂ + STS	68.1% 24 h	Solar light irradiation
	CBMO (Present study)	67.43% 2 h	Solar light irradiation
Methyl orange	Chitosan-g-Poly (acrylamide)/ZnS	69% 4 h	Solar light irradiation
	Chitosan/LaFe _{0.8} Cu _{0.2} O ₃	67% 8 h	UV light irradiation
	Xylan/LaFe _{0.8} Cu _{0.2} O ₃	58% 8 h	UV light irradiation
	CBMO (present study)	74.05% 2 h	Solar light irradiation

Table 4 Comparative analysis of synthetic and industrial wastewater onto CBMO

Composite	Degrada-tion time	% of degradation		
		Synthetic wastewater		Industrial wastewater
		Methylene blue	Methyl orange	
CBMO	120 min	67.43	74.05	62.05

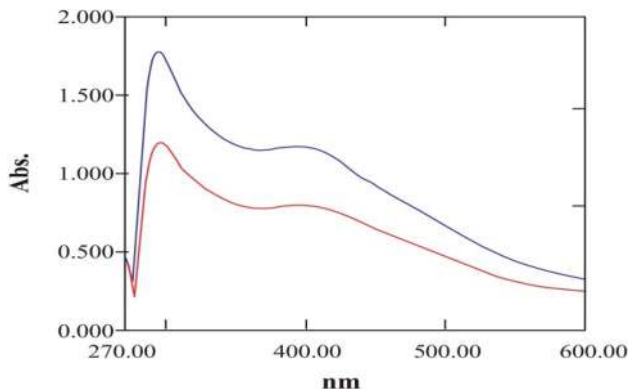


Fig. 11 UV-Spectra of industrial wastewater onto CBMO

between time t and $\log (q_e - q_t)$ which gives a straight line as shown in Fig. 14a, b.

In pseudo second order kinetics a graph is plotted between time (t) versus t/q_t as shown in the Fig. 15. The commonly used form of pseudo second order kinetics is as follows.

$$\frac{t}{q_t} = \frac{1}{K_2} - \frac{t}{q_e} \tag{5}$$

K_2 is the equilibrium rate constant for the adsorption of MB and MO onto CBMO is 0.00072 and 0.00015 g/mg min respectively, q_t is the uptake of dye at moment t and the amount of dye adsorbed at equilibrium time q_e is 24.03 and 39.37 mg/g. The regression coefficient R^2 for the adsorption of MB is 0.9929 and 0.9916 for MO indicates the equilibrium data is well fitted for pseudo second order kinetic model when compared to pseudo first order kinetic model.

The kinetic parameters for the degradation of MB and MO onto CBMO are summarized in the Table 6.

Fig. 12 Freundlich isotherm **a** for MB onto CBMO, **b** for MO onto CBMO

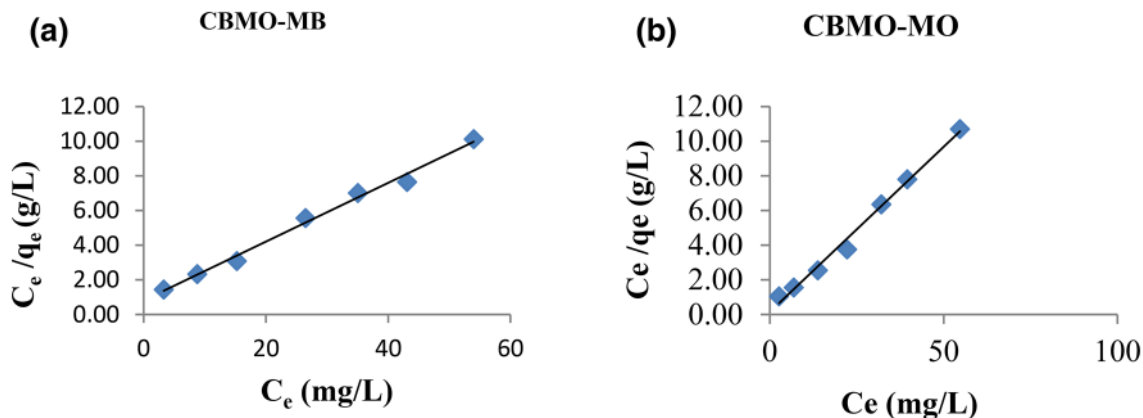
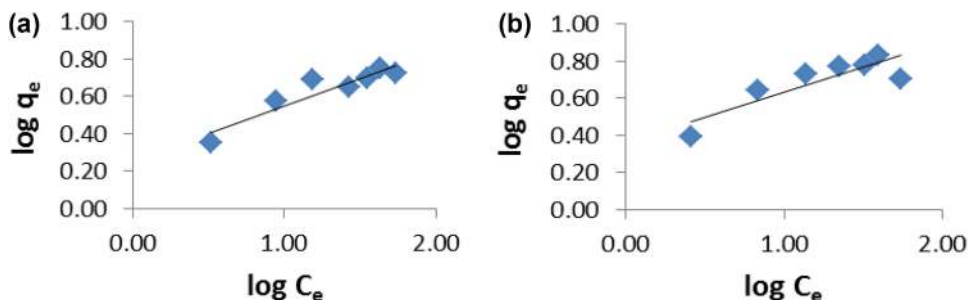


Fig. 13 Langmuir isotherm for **a** MB onto CBMO, **b** MO onto CBMO

Table 5 Isotherm parameters for MB and MO onto CBMO

Isotherm model	Parameters	MB	MO
Freundlich isotherm	1/n	3.385	3.70
	K _F (mg/g)	1.81	2.31
	R ²	0.8702	0.7471
Langmuir isotherm	Q _m (mg/g)	5.9031	5.561
	b (L/mg)	1.151	2.407
	R ²	0.9919	0.9914

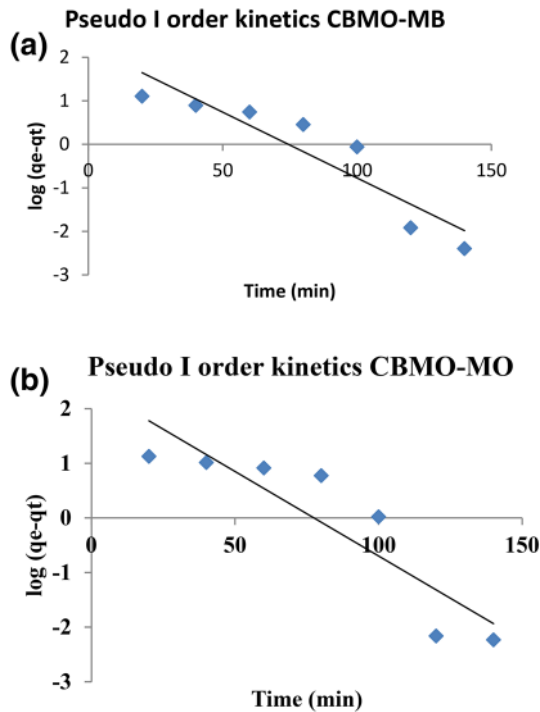


Fig. 14 Pseudo first order kinetic model for the degradation of **a** MB onto CBMO, **b** MO onto CBMO

3.6 Proposed mechanism for the degradation of methylene blue and methyl orange degradation onto CBMO

A possible mechanism has been proposed based on the experimental results of photo catalytic degradation of MB and MO onto CBMO. The overall photo catalytic degradation is depends on the light degradation ability of the composite and separation of excited electron hole pair [42].

According to the above discussion, CBMO produce conduction band via photo catalytic method obviously increases the degradation ability of the composite and generates the separation of photo generated electrons-hole pair; hence a considerable photo catalytic activity

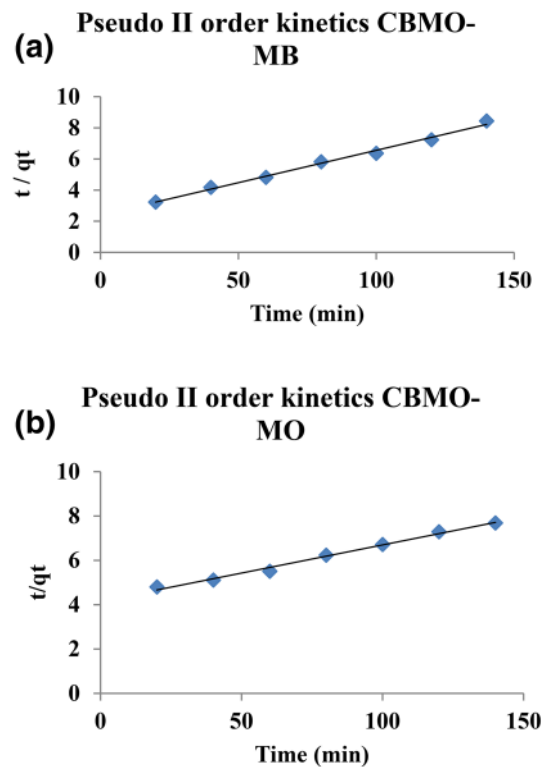
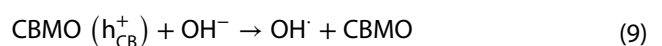
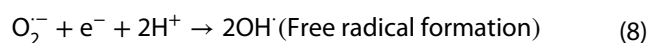
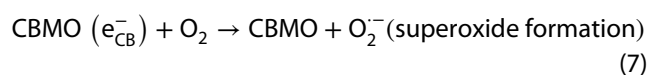
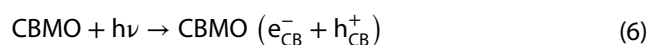


Fig. 15 Pseudo second order kinetics for the degradation of **a** MB onto CBMO, **b** MO onto CBMO

Table 6 Kinetic parameters for the degradation of MB and MO onto CBMO

Kinetic model	Parameters	MB	MO
Pseudo first order kinetics	k ₁ (min ⁻¹)	0.0696	0.0712
	q _e (mg/g)	16.59	18.21
	R ²	0.853	0.8031
Pseudo second order kinetics	k ₂ (g/mg min)	0.00072	0.00015
	q _e (mg/g)	24.03	39.37
	h	0.415	0.24
	R ²	0.9929	0.9916

[43] is achieved. The possible mechanism may be given as follows



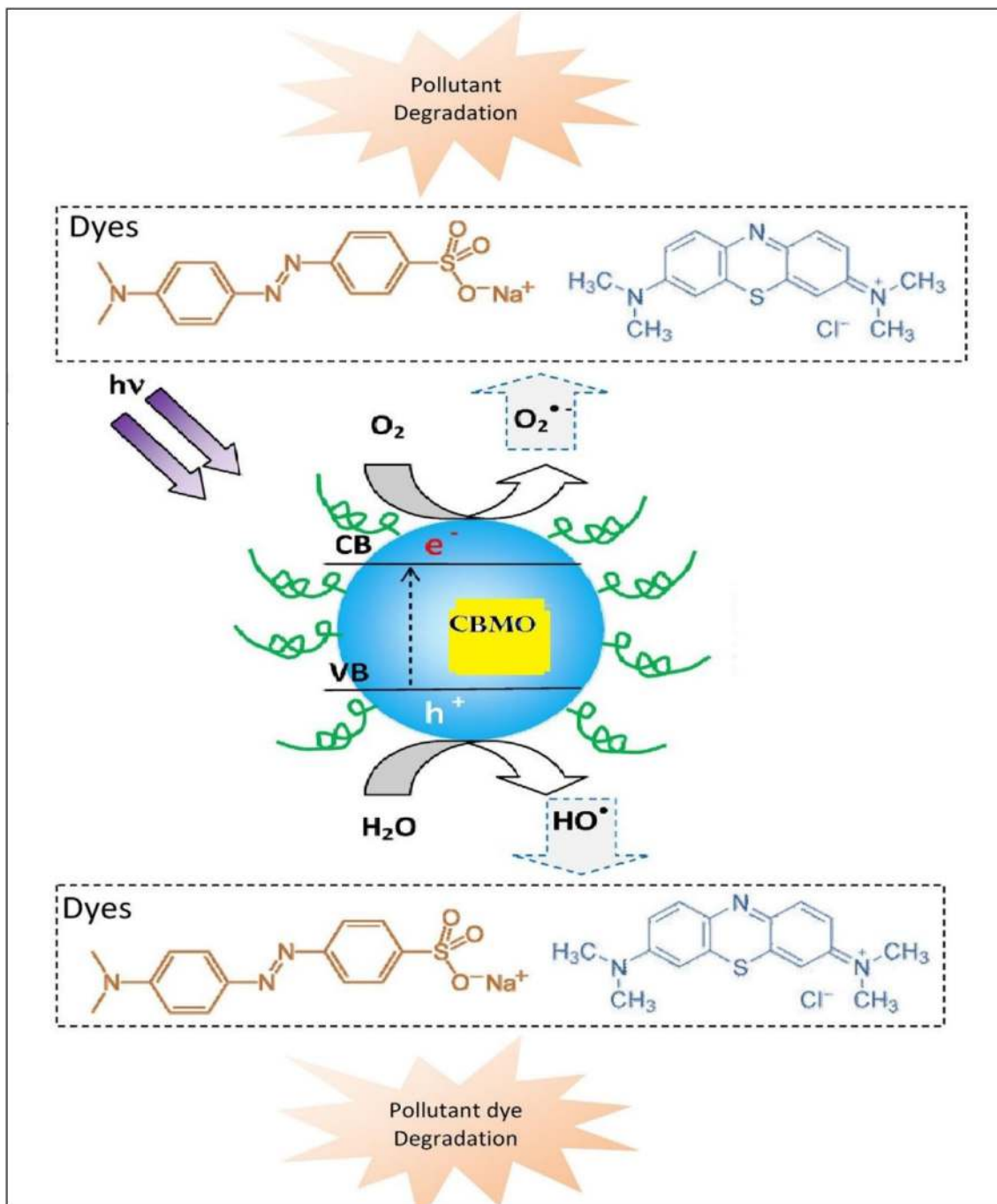
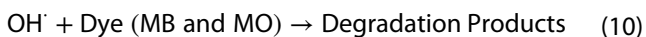


Fig. 16 Proposed mechanisms for the degradation of MB and MO onto CBMO



The predicted mechanism for the degradation of MB and MO onto CBMO is given in the Fig. 16, which gives the clear mechanism of the degradation process.

4 Conclusion

The present research paper focuses on the preparation and utilization of modified chitosan composite (CBMO) to investigate the degradation and adsorption efficiencies of MB and MO dyes by photocatalytic degradation studies using solar light as an illuminant. The physiochemical and surface characterization is studied for the prepared

composites and is reported. By varying irradiation time, pH, CBMO dose and initial dye concentration of MB and MO the optimum conditions are determined and fixed for entire studies. The result of synthetic wastewater is compared with industrial wastewater to determine the maximum % of degradation efficiency of CBMO. Based on results the optimum irradiation time is 120 min, pH 8 for MB and pH 3 for MO and the optimum dose of CBMO is 0.15 g. The maximum degradation of synthetic MB is 67.43% and MO is 74.05% at optimum conditions and for industrial wastewater the maximum degradation is 62.05% indicates the photo catalyst is effectively suitable for both synthetic and industrial wastewater. The maximum adsorption capacity of MB and MO onto CBMO is 5.90 mg/g and 5.56 mg/g respectively. The experimental data is well fitted for pseudo second order kinetics and Langmuir adsorption isotherm shows that the adsorption of MB and MO onto CBMO is composite dependent and monolayer formation occurred in homogeneous surface of the composite. Possible mechanism for the degradation MB and MO onto CBMO is clearly explains the path way of the degradation. Based on this study CBMO can be recommended as a catalyst for the removal of synthetic as well as industrial wastewater.

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Compliance with ethical standards

Conflict of interest The authors declare that they have no conflict of interest.

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