

CORRECTION

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Cite this: *Chem. Soc. Rev.*, 2021,
50, 10700

DOI: 10.1039/d1cs90075g

rsc.li/chem-soc-rev

Correction: Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone

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Correction for 'Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone' by Igor V. Alabugin et al., *Chem. Soc. Rev.*, 2021, DOI: 10.1039/d1cs00386k.

The authors regret that Fig. 4 and 116 contained errors in the labeling of the acceptor orbitals of the carboxylic acids. The corrected versions of these figures are presented here.



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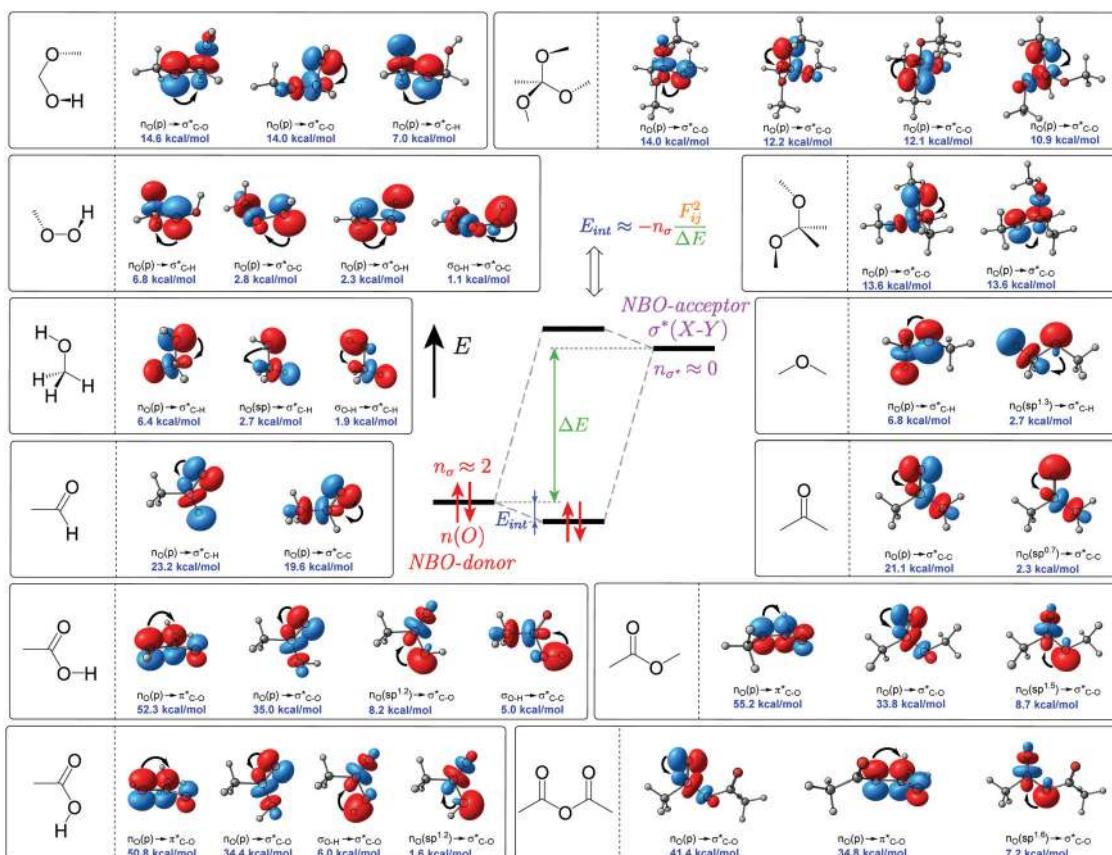


Fig. 4 A selection of donor–acceptor anomeric interactions in organic O-functionalities. E_{int} values are computed for the Natural Bond Orbital method (NBO) using the second-order perturbation approximation for the PBE0-D3BJ/aug-cc-pVTZ/CPCM(H₂O) level of theory. The orbital interaction diagram at the center of the scheme shows the equation used to calculate the interaction energy— E_{int} —and its contributors: F_{ij} is the Fock matrix element between natural bond orbitals i and j , ε_i and ε_j are the energies of the σ and σ^* NBO's, and n_i is the population of the donor orbital.²⁷

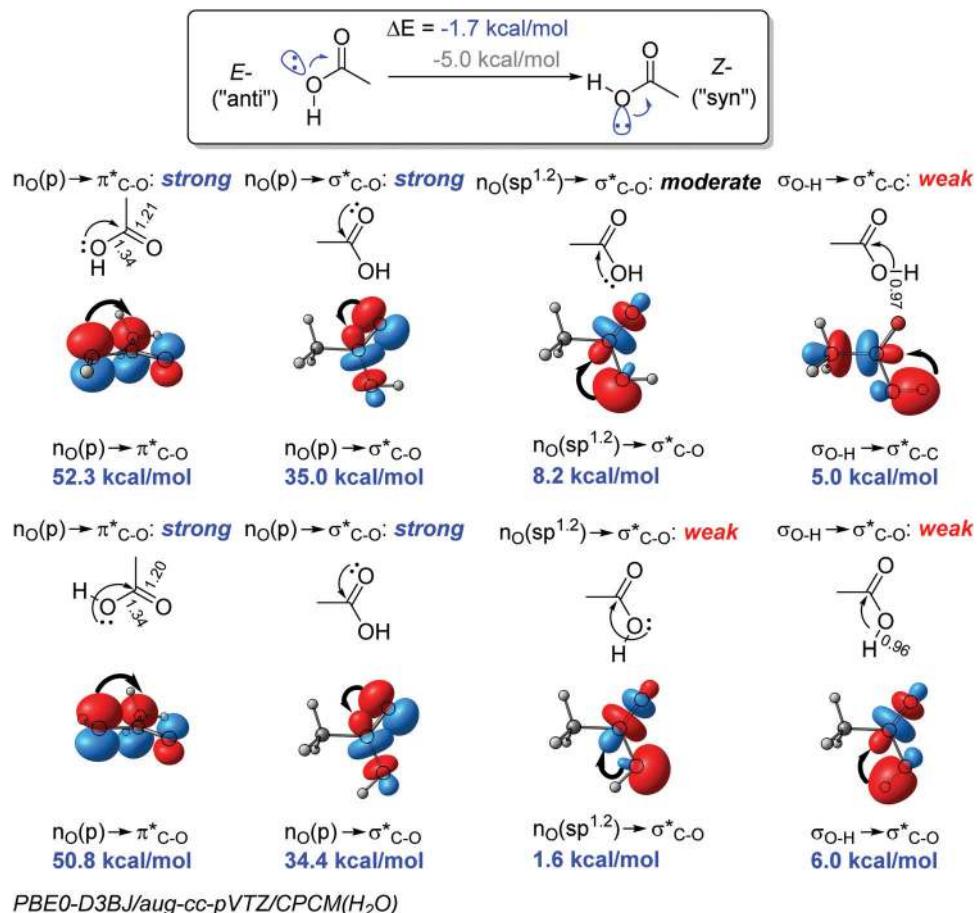


Fig. 116 Stereoelectronic “portrait” and conformational preferences of acetic acid.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.