

Structure and Reactivity of Surface Oxides on Pt(110) during Catalytic CO Oxidation

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We present the first structure determination by surface x-ray diffraction during the restructuring of a model catalyst under reaction conditions, i.e., at high pressure and high temperature, and correlate the restructuring with a change in catalytic activity. We have analyzed the Pt(110) surface during CO oxidation at pressures up to 0.5 bar and temperatures up to 625 K. Depending on the O₂/CO pressure ratio, we find three well-defined structures: namely, (i) the bulk-terminated Pt(110) surface, (ii) a thin, commensurate oxide, and (iii) a thin, incommensurate oxide. The commensurate oxide only appears under reaction conditions, i.e., when both O₂ and CO are present and at sufficiently high temperatures. Density functional theory calculations indicate that the commensurate oxide is stabilized by carbonate ions (CO₃²⁻). Both oxides have a substantially higher catalytic activity than the bulk-terminated Pt surface.

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CO oxidation on Pt single crystal surfaces forms a favorite model system for heterogeneous catalysis. Most of our present-day understanding of this system is based on a large number of experiments that have been carried out under laboratory rather than industrial conditions, such as ultrahigh vacuum (UHV) and low temperatures. Presently, a growing number of experimental techniques is being adapted to bridge the so-called *pressure gap* and obtain atomic-scale insight in the structure, composition, and other properties of model catalyst surfaces under (semi) realistic reaction conditions. Examples are scanning tunneling microscopy [1–3], surface x-ray diffraction [4–7], transmission electron microscopy [8], and photoelectron spectroscopy [9].

Previous high-pressure STM observations during CO oxidation on Pt(110) have demonstrated that at a sufficiently high O₂/CO pressure ratio this surface undergoes a structural phase transition, which has a dramatic effect on the reaction mechanism and strongly enhances the reaction rate [2]. The STM movies have been interpreted in terms of the formation of a thin platinum oxide film on the Pt(110) surface, on which the CO molecules oxidize through a Mars–Van Krevelen mechanism, and react with O atoms from the oxide. There is a growing body of evidence from experiments [2,10,11] as well as theory [12] strongly suggesting that similar scenarios are active for CO oxidation on other metal surfaces.

In this Letter we present the first observation by high-pressure surface x-ray diffraction of a correlation between the reactivity and the structure of a catalyst surface. The

experiment performed during CO oxidation on Pt(110) shows that the most active structures are, in fact, two thin oxide film configurations, one of which is stable only when the surface is in contact with *both* reactant gasses, CO and O₂, at elevated temperatures.

The experiments were performed at the ID03 beam line of the European Synchrotron Radiation Facility (ESRF) in a combined UHV-high-pressure surface x-ray diffraction (SXRD) chamber (10⁻¹⁰ mbar – 2 bar), which had a volume of ~2 liter and was equipped with a 360° beryllium window for entrance and exit of the x rays [13]. The sample could be heated up to 1300 K in vacuum. Connected to the chamber were a gas manifold with four high-purity gasses (N47 grade for CO, N55 for all other) and a quadrupole mass spectrometer (QMS) for online gas analysis. The chamber was mounted on a z-axis diffractometer with the crystal surface in a horizontal plane.

A focused beam of monochromatic, 17 keV x-ray photons was impinging on the surface at an angle of 1° (~4 × 10¹² photons/s). The fluorescence radiation in the scattered beam was filtered with a crystal analyzer. We describe the Pt crystal lattice with two perpendicular unit vectors **A**₁ and **A**₂ in the surface plane along [1 $\bar{1}$ 0] and [001], respectively, and the third one, **A**₃, along the [110] surface normal. |**A**₁| = a₀ = 2.774 Å is the Pt nearest-neighbor distance; |**A**₂| = √2a₀ and |**A**₃| = a₀. In reciprocal space **H** and **K** are the unit vectors in the surface plane (parallel to **A**₁ and **A**₂, respectively) and **L** is along the surface normal. Well-ordered, clean Pt(110) surfaces exhibiting the characteristic (1 × 2) missing-row recon-

struction [14] were obtained after several ion bombardment (1 keV Ar⁺) and annealing (1200 K) cycles. The full width at half maximum (FWHM) of a rocking scan around the (0 0.5 0.2) reflection was typically 0.24°, which corresponds to ordered domain sizes of ~1200 Å.

Our procedure to form a well-ordered oxide film was to first lift the (1 × 2) missing-row reconstruction by exposing the surface to 100 mbar of CO at 600 K [15]. Within minutes, the surface reordered to a (1 × 1) bulk-terminated surface with terraces of ~3000 Å. The surface was then cooled to room temperature under CO to preserve the (1 × 1) structure. Subsequently, the CO pressure was reduced to 10 mbar and 500 mbar of O₂ was admitted to the chamber, after which the sample was heated to 650 K. This procedure led to new diffraction peaks at noninteger **H** and **K** values (see Fig. 1) forming a regular, slightly distorted hexagonal pattern, exhibiting angles of 57.4° and 61.3°. As indicated in Fig. 1, the hexagons were aligned with respect to the substrate crystallographic directions. The intensities of the new reflections were similar to those of the surface reflections from the clean substrate, which immediately indicated that they arose from one or more

layers of Pt atoms, rather than from an overlayer of merely O atoms or CO molecules. The unit mesh of the nearly hexagonal Pt layer was very close to that of α-PtO₂ (see Fig. 1), suggesting that the overlayer was a distorted PtO₂ layer, azimuthally aligned with the substrate, and oriented with its *c* axis parallel to the surface normal [16]. Relative to the bulk lattice of α-PtO₂ our surface oxide exhibited compressive strain of 3.4% along the [1 $\bar{1}$ 0] direction and tensile strain of 2.4% along the [001] direction. We have also oxidized the (1 × 2)-reconstructed surface without preexposure to CO. The resulting oxide layers exhibited the same structure, but with very poor order [3], both in and out of plane. Hence, all experiments discussed here were performed with oxide layers grown on the (1 × 1) bulk-terminated surface.

A crystallographic data set consisting of 22 in-plane reflections (up to 4th order, 10 nonequivalent) and 3 crystal truncation rods (2 nonequivalent, Fig. 2) was collected to determine the oxide structure. The oscillations along **L** in the data in Fig. 2 show that several Pt layers were involved in the surface oxide. For a quantitative fit to the intensities [17], we started from a distorted α-PtO₂ film and had as free parameters (i) the thickness of the oxide, (ii) the lattice constant along **A**₃, and (iii) the distance between the oxide film and the substrate. As the scattering of the x rays is largely dominated by the Pt atoms, the stoichiometry and the relative positions of the oxygen atoms were set to those for the bulk oxide. The continuous curves in Fig. 2 correspond to 2.6 ± 0.1 oxide monolayers (ML), i.e., a 40:60 mixture of 2- and 3-ML oxide films, an oxide monolayer being defined as an O-Pt-O triple layer. The fitted Pt-Pt distance is 2.8 Å along **A**₃, which is 36% shorter than the reported distance in bulk PtO₂. The fit to the (00**L**) rod (Fig. 2) confirms this rather high density of the oxide layer compared to reported values for bulk PtO₂. This strong contraction might be due to the very small thickness of the oxide film. The distance between the outermost Pt layer of the substrate and the innermost Pt layer of the oxide was found to be 3.3 ± 0.1 Å, which is very close to the Pt-PtO₂ distance calculated by Helveg *et al.* for a fully oxidized Pt atom row on the missing-row reconstructed Pt(110) surface [18].

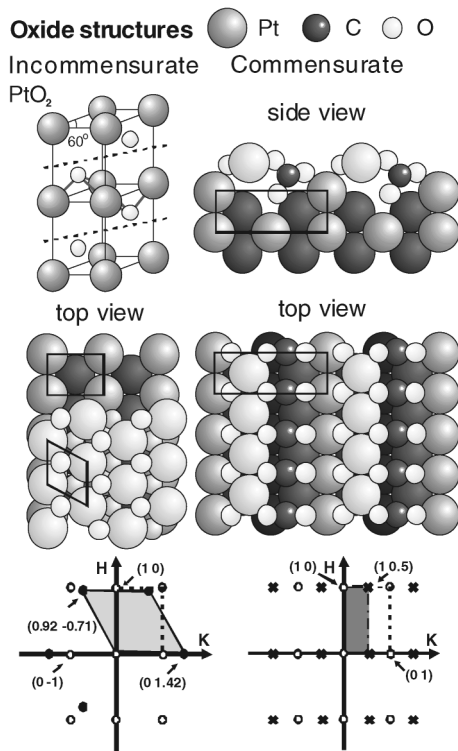


FIG. 1. Bottom left: in-plane reciprocal space map of the Pt(110) substrate (open circles), with the incommensurate, quasi-hexagonal oxide overlayer (solid circles). One diffraction peak could not be measured due to instrumental limitations. Higher orders have been measured as well (not shown). Bottom right: the (1 × 2), commensurate oxide overlayer (crosses). All coordinates are expressed in reciprocal lattice units of Pt(110). The ball models show the real-space structures of the two oxides.

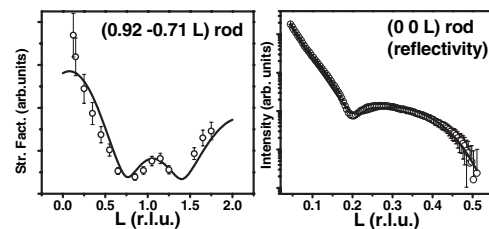


FIG. 2. Structure factor at (**H**, **K**) = (0.92, -0.71) as a function of **L** (left), and reflection intensity curve [(00**L**)-rod] (right), measured at 625 K and 0.5 bar of O₂ on Pt(110). **L** is expressed in reciprocal lattice units of the substrate. The solid curves are a fit for the incommensurate, 2.6 ML, α-PtO₂ oxide film.

Simultaneously with the diffraction measurements we have recorded the CO oxidation reactivity both on the bulk-terminated metal surface and on oxidized surfaces. These experiments were performed in the so-called batch mode, in which the chamber was first filled with 500 mbar of O_2 , with the Pt sample at a temperature in the range 425–625 K. SXRD measurements showed that under these conditions the surface was covered with the α -PtO₂ structure discussed above. Pulses of pure CO were then introduced into the chamber [e.g., at point (a) in Fig. 3]. The CO started reacting to CO₂ as soon as it entered the chamber. From the decrease in the CO and O₂ partial pressures, P_{CO} and P_{O_2} , and the increase in the CO₂ partial pressure, P_{CO_2} , the typical turnover number can be estimated to be approximately 3×10^3 molecules/site/s at 625 K in a mixture of 80 mbar CO and 500 mbar O₂. The comparison of the turnover numbers from experiments at different temperatures suggests that the maximum reaction rate was limited by the diffusion of CO through the O₂ dominated gas mixture to the surface and not by the reactivity of the catalyst surface. Several other groups have found similarly high turnover numbers in comparable conditions [19,20].

In the top panel of Fig. 3 we show the diffracted intensity at (0 1.42 0.5) from the oxide layer. The introduction of CO caused a sudden $\sim 25\%$ drop in the intensity followed by a slow recuperation. During the recovery, the shape of the (0 1.42 L) scan and the in-plane FWHM of the reflection did not change significantly, indicating that the surface remained covered by an equally thick oxide layer throughout the reaction. The intensity drop is ascribed to a roughening of the oxide layer caused by the reaction, which leads to a growing density of pits and protrusions, an effect that has already been seen by STM under similar conditions [2]. Because of the incommensurability of the oxide layer, the in-plane domain size is very small (~ 7 –20 nm depending on growth conditions) and it does not reduce further during the roughening. The fifth CO pulse, at point (b) in Fig. 3,

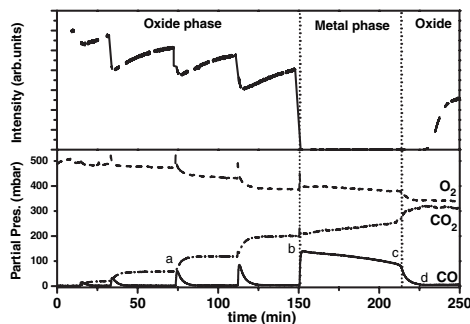


FIG. 3. Simultaneously measured x-ray diffraction intensity at (0, 1.42, 0.5) from the quasihexagonal oxide (top panel) and partial pressures of CO, O₂, and CO₂ (bottom panel). Separate CO pulses were admitted to the reactor, which was initially filled with 500 mbar of O₂ at a temperature of 625 K. The sharp peaks in P_{O₂} are an artefact due to the sudden increase of total pressure at each CO pulse.

was so large that the oxide layer was reduced completely: the diffracted intensity from the oxide layer dropped to zero and the surface reverted to the bulk-terminated (1 × 1) periodicity, found before in pure CO. A simultaneous decrease was observed in the reactivity, by a factor of roughly 25, as calculated from the decay rate of P_{CO} after (b) with respect to the decay rates after the preceding pulses. Also the reaction order was changed, since after (b) P_{CO} varied roughly linearly in time, i.e., with a rate that was independent of the amount of remaining CO, whereas it decayed exponentially, i.e., with a rate proportional to P_{CO}, after the four earlier pulses when the surface was still covered by the oxide.

At point (c) in Fig. 3, at a partial pressure ratio P_{CO}/P_{O₂} of approximately 0.2, the CO₂ formation rate sped up and the variation in all partial pressures switched to exponential again, indicating that the catalyst had reverted to its “high-reactivity” phase. Interestingly, the diffraction intensity associated with the incommensurate oxide overlayer came back only at point (d), 20 minutes after (c).

As Fig. 4 shows, immediately when the reaction sped up, new diffraction peaks appeared at integer **H** values and half-integer **K** values, corresponding to a commensurate (1 × 2) structure (Fig. 1), rather than the incommensurate, quasihexagonal oxide structure, developing later at point (d). The intensities of the new diffraction peaks were high, indicating that also these were due to a structure involving Pt atoms. Figure 5 shows **L** scans of two of the new half-order reflections. They clearly differ from the calculated values expected for the (1 × 2)-reconstructed, clean Pt(110) surface [14], indicating that the structure differed strongly from the known missing-row reconstruction. The solid curves in Fig. 5 show the fit of the **L** scans, based on a simple structural model in which every second close-packed [110] row of surface Pt atoms was strongly dis-

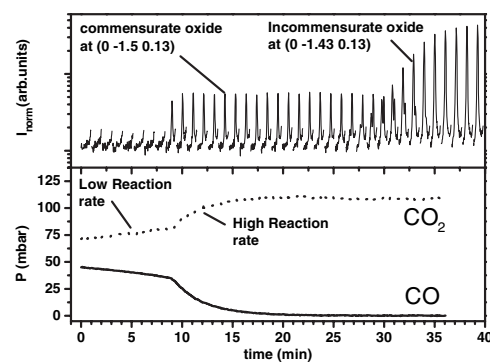


FIG. 4. Diffraction scans in a narrow range from **K** = −1.30 to −1.65 and at (**H** = 0, **L** = 0.13) (top panel), and partial pressures of CO and CO₂, simultaneously measured at P_{O₂} = 500 mbar and $T = 625$ K. Together with the increase in the reaction rate, at $t = 9$ min, a half-order diffraction peak appeared at **K** = −1.50. When nearly all CO had been consumed, at $t \approx 30$ min, this diffraction peak decreased to zero while the peak of α -PtO₂, at **K** = −1.43, appeared.

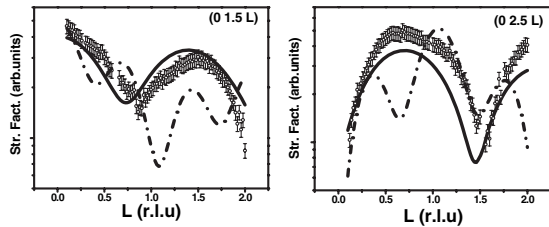


FIG. 5. The (0 1.5 L) and (0 2.5 L) rods of the (1×2) commensurate oxide layer, formed at 625 K and 0.5 bar of O_2 and approximately 35 mbar of CO on Pt(110). The solid curves are fits for the carbonate-stabilized commensurate oxide film, as discussed in the text. The dash-dotted lines are calculated curves for the 1×2 missing-row reconstruction [14].

placed. In addition to this large displacement, small relaxations of all other atoms in the two topmost layers of Pt were considered. Relaxations beyond this second layer did not add any accuracy to the fit, and were hence neglected, as were the C and O atom positions. In the best-fit structure every second Pt surface atom row was displaced by $(0.5 - 0.38 \ 0.67)$, expressed in unit vectors of the (1×1) unit cell, while the relaxations of the other atoms were modest (Fig. 1). With the low-symmetry position of the shifted Pt atom rows, two equivalent mirror versions can be formed, and the best fit was obtained for nearly equal proportions of these.

Although the two surface oxide structures resolved here are different from the structures observed in Ref. [3] after exposure of Pt(110) to atomic oxygen, some of the structural elements might be similar. In particular, local configurations in Figs. 1(b)–(d) of Ref. [3] resemble the geometry of the commensurate surface oxide, discussed above.

The observation that the commensurate oxide structure was only stable under reaction conditions, when both CO and O_2 were present, strongly suggests that the (1×2) oxide was stabilized by the presence of CO or a CO-related species, adsorbed on the oxide surface. In order to further substantiate this notion we have performed density functional theory calculations, following the methods described in Ref. [3]. Structures were investigated with the Pt atom density of Pt(110)- (1×1) and at most one first layer row shifted half a lattice spacing along the A_1 direction, as found in the experiment, and with adjustable displacements along A_2 and A_3 . Only one structure was identified that involved a high-lying Pt row, consistent with the experiment (Fig. 1), and its energy per surface unit cell was within 0.27 eV from the most stable structure found. Since this energy difference is comparable to the accuracy of density functional theory [21], the density functional theory results are supporting the possibility of such a high-lying Pt row. An essential feature of the proposed structure is the presence of one carbonate ion (CO_3^{2-}) per unit cell,

which stabilizes the high Pt row and thus the commensurate oxide monolayer.

In this work we have explicitly demonstrated the role of spontaneously formed surface oxides in a catalytic process under (semi) relevant pressure and temperature conditions. The surface oxides found here on Pt(110) have a significantly higher reaction rate than the original metallic surface. Our observations show that *in situ* measurements under actual reaction conditions are an absolute necessity for a meaningful investigation of the surface structure and chemical behavior of this model catalyst. Clearly, the presence and role of the commensurate oxide layer could not have been found either in experiments under UHV or in so-called pre- and post-reaction experiments.

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