

# Synthesis and characterization of chitosan–silver nanocomposite

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**Abstract** Chitosan–silver (CS–Ag) nanocomposite materials were synthesized by a simple chemical method. The synthesized CS–Ag nanocomposite contains 20 wt% silver. Silver nanoparticles were synthesized by chemical reduction method as well. The CS–Ag nanocomposite was characterized using Field emission scanning electronic microscope (FESEM), X-ray diffraction (XRD) and Fourier transform infrared spectroscopy (FTIR). The XRD pattern indicated the presence of both silver and chitosan in the nanocomposite. It is observed from the XRD pattern of silver that it is of cubic structure. The spherical morphology of silver nanoparticles was confirmed from the FESEM image. FTIR spectroscopy was used for the structural elucidation. CS–Ag nanocomposite exhibits good antimicrobial and antitumor properties.

**Keywords** Nanocomposite · Biocompatible · Chitosan · Silver

## Introduction

There has been a tremendous increase in the number of patients affected by diseases like diabetes and cancer during the last decade. So, it has become the major focus of researchers to synthesize a compound that will be useful

for the monitoring of certain elements that are present in excess quantities in the blood while used in the form of biosensors, at the same time being useful for the treatment of diseases like cancer.

Chitosan (CS) is a polysaccharide composed of glucosamine and *N*-acetyl glucosamine linked with a  $\beta$ -1-4-glycosidic linkage (Crini et al. 2008). CS is a biopolymer which is biocompatible and can be degraded by enzymes in human body, the degradation products are nontoxic. Commercial CS's are semi-crystalline polymers and crystallinity plays an important role in adsorption efficiency (Trung et al. 2006). CS is a polymer which exhibits a broad spectrum of antimicrobial activity by binding to the negatively charged bacterial cell wall followed by attachment to the DNA, inhibiting its replication. CS is a biopolymer having antibacterial properties (Kurita et al. 1993; Ting and Shen 2005). For the improvement of bioactivity on chitosan, it is often combined to other bioactive materials, such as drugs.

Silver (Ag) nanoparticles have high therapeutic potential and exhibit good antimicrobial activity. Ag nanoparticles have a wide range of antimicrobial activities and exhibit high performance even at a very low concentration. Ag nanoparticles have been identified to possess good potential for the treatment of cancer (Sriram et al. 2010). But the major disadvantage of using silver alone is that it is not specific at targeting the cancer cells and also it is toxic to the normal cells when exposed for a longer time when the size of silver used is  $>20$  nm (Vaidyanathan et al. 2009)

CS–Ag nanocomposite is one of the rare composite materials that is seen to possess a capability of being used as a biosensor as well as in the treatment of cancer as the chitosan present in the nanocomposite is very specific to the cancer cells. It prolongs the action of silver on the affected cells while preventing the normal cell from the

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effect of silver. One more advantage of this nanocomposite is that it is biodegradable, i.e., it can be degraded by the enzymes present in the body making it suitable for the treatment of cancer. Apart from the treatment of cancer, the nanocomposite also possesses good antimicrobial (Sanpui et al. 2008) and biosensing activity.

In this work, CS–Ag nanocomposite was synthesized by the chemical method with 20 wt% Ag. CS–Ag was successfully synthesized and characterized with Fourier transform infrared spectroscopy (FTIR), X-ray diffraction analysis (XRD) and field emission scanning electron microscopy (FESEM).

## Experimental

### Materials

Chitosan (medium molecular weight), silver nitrate were purchased from Sigma-Aldrich. Sodium borohydride, sodium citrate sodium sulphate and acetic acid were of SRL grade.

### Synthesis

The chemical method was used for the synthesis of chitosan, silver and chitosan–Silver nanocomposite. Chitosan (0.25 g) was dissolved in 2 % acetic acid, while stirring followed by sonication for 15 min. This was then filtered to obtain a clear solution. 20 % Na<sub>2</sub>SO<sub>4</sub> was then added drop wise to the aq. chitosan solution, which was simultaneously stirred and sonicated. The solution was left undisturbed, centrifuged at 8,000 rpm to collect the CS nanoparticles (Baek et al. 2008). Silver nanoparticles were obtained by dissolving 0.4225 g of AgNO<sub>3</sub> in water and stirring for about 20 min. To this sodium citrate solution (73.5 g) was added. NaBH<sub>4</sub> solution (1.318 g) was added drop wise to the above solution and stirred. Sodium citrate was used as a stabilizing or capping agent to prevent the aggregation of nanoparticles and to control the size of the nanoparticles, while sodium borohydride was used as a reducing agent (Tolaymat et al. 2010). The settled particles were collected by filtration.

For the synthesis of CS–Ag nanocomposite, 0.5 g of chitosan was dissolved in 2 % acetic acid, to which 1 % PVA solution was added. 20 % Na<sub>2</sub>SO<sub>4</sub> solution was added drop wise to the above solution under simultaneous stirring and sonication. 0.1 g of synthesized silver accounting for 20 wt% was added to the solution which was stirred under sonication for ample time. 20 wt% indicates that the amount of silver taken for the preparation of the nanocomposite amounts to 20 % of the weight of chitosan used. The calculation is as shown below.

**Table 1** Examples for the calculation of weight percentage of silver

| S. no. | Amount of chitosan taken (g) | Amount of silver taken (g) | wt% |
|--------|------------------------------|----------------------------|-----|
| 1      | 0.5                          | 0.025                      | 5   |
| 2      | 0.5                          | 0.05                       | 10  |
| 3      | 0.5                          | 0.075                      | 15  |
| 4      | 0.5                          | 0.1                        | 20  |

$$\left(\frac{0.1}{0.5}\right) \times 100 \% = 20 \%$$

A table (Table 1) showing the calculation of weight percentage of Silver is shown below. The solution was then centrifuged at 8,000 rpm for 20 min after which particles were collected and dried at room temperature.

CS–Ag nanocomposites containing various amounts of silver like 5, 10, 15 and 20 wt% were synthesized. Of these the composite containing 20 wt% of silver is chosen here for discussions as the properties of silver will be more pronounced when the amount of silver present in the composite is more. So it would be easier for us to study the properties of silver as well chitosan.

### Particle characterization

The structure analysis of the prepared samples was done using a Rich Siefert 3000 XRD meter using Cu K $\alpha$  radiation ( $\lambda = 1.5406 \text{ \AA}$ ), surface morphology of CS–Ag nanocomposite was observed using a HITACHI-SU 6600 FESEM. Functional group confirmations were assessed by SHIMADZU IR-affinity FTIR spectrometer.

## Results and discussion

The structural properties of the synthesized CS–Ag nanocomposite were analyzed using the XRD technique. The obtained XRD pattern for chitosan is shown in Fig. 1. The peaks appeared at  $2\theta$  values of 11.7° and 20.2° which match well with the literature values (Harish Prashanth et al. 2002; Wang et al. 2004; Kong et al. 2005). The broadening of the peaks is due to the amorphous nature of the polymer (Rhim et al. 2006). There is no impurity peaks observed in the XRD pattern.

The XRD pattern of silver is shown in Fig. 2 which clearly indicates the formation of silver in a single phase. Peaks were obtained at  $2\theta$  values of 37.9°, 44.0° and 63.9° which are in perfect agreement to that of the JCPDS card no. 89-3722 (Raffi et al. 2008). Peak fitting was done for the obtained pattern, the crystal structure was found to be simple cubic and the lattice parameters was determined to

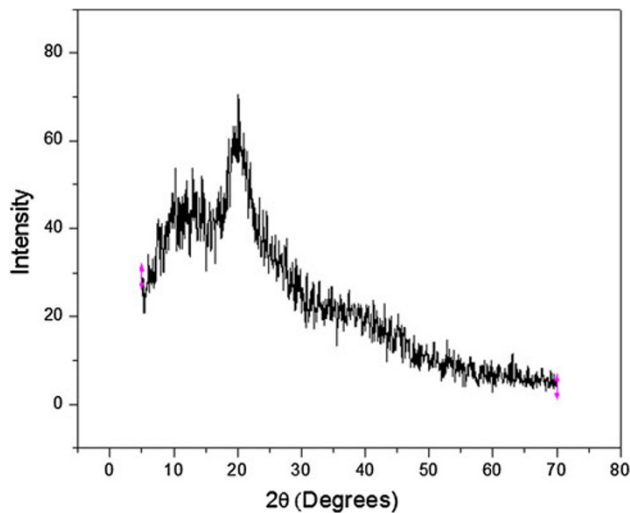


Fig. 1 XRD pattern of chitosan

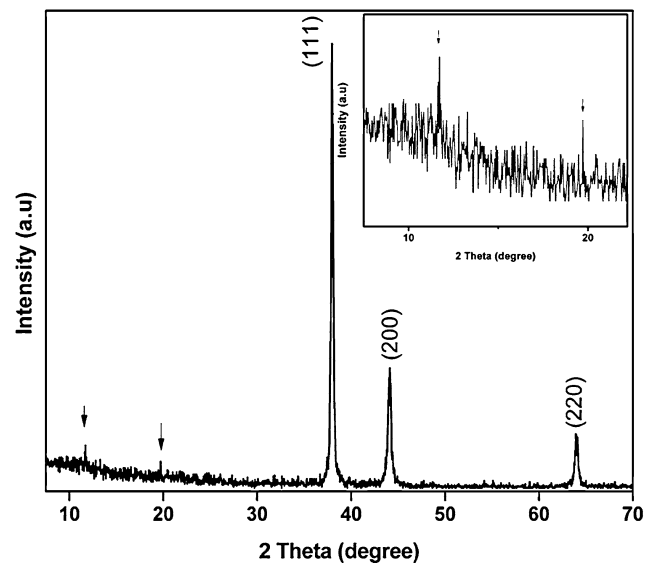


Fig. 3 XRD pattern of chitosan–silver nanocomposite

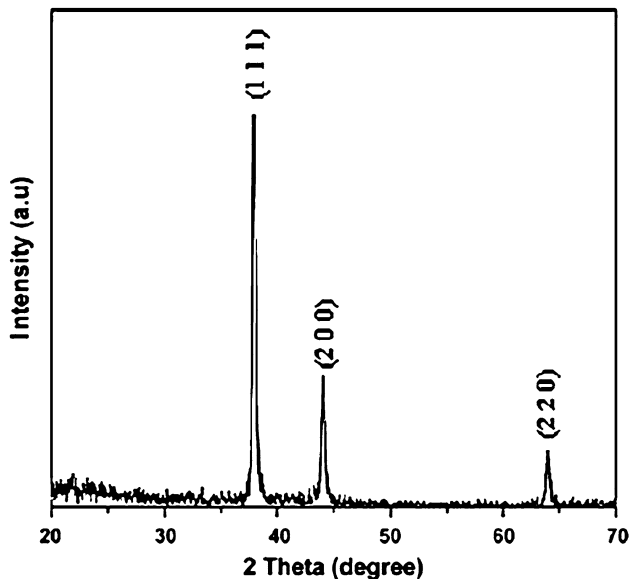


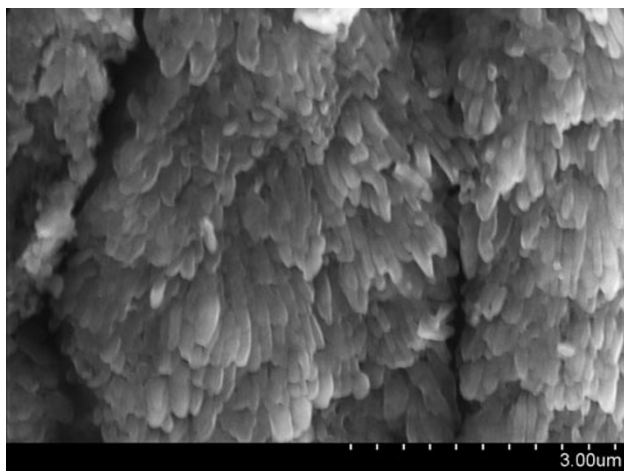
Fig. 2 XRD pattern of silver

be  $a' = 4.111 \text{ \AA}$ . The  $(hkl)$  values are indicated in the Fig. 2. The average crystallite size of the silver nanoparticles calculated using the Scherrer formula was 23 nm.

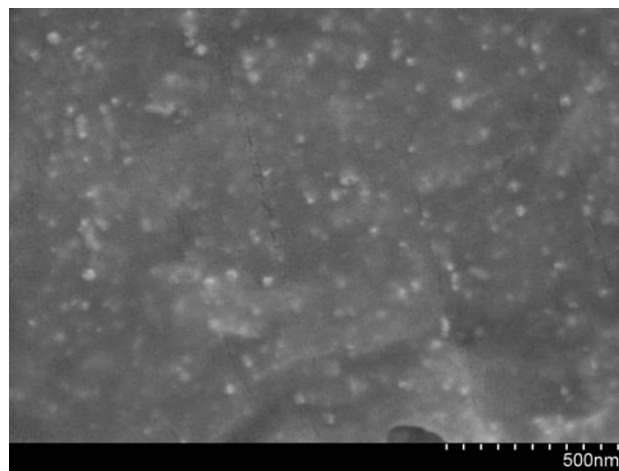
The presence of chitosan as well as silver peaks is observed from the XRD pattern of CS–Ag nanocomposite which is shown in Fig. 3. This seems to be in agreement with the values reported in the literature (Rhim et al. 2006). The chitosan peaks are represented using arrows, while the silver peaks are designated with their  $(hkl)$  values. The regions of the graph at which the chitosan peaks appear are separately represented in an inset. The CS peaks were obtained at  $2\theta$  values of  $11.7^\circ$ ,  $19.8^\circ$  and silver peaks were obtained at  $37.9^\circ$ ,  $44.0^\circ$  and  $63.9^\circ$ , respectively.

The surface morphology of synthesized CS–Ag nanocomposite was analyzed using the FESEM technique. The FESEM image of CS shows that the particles are in the form of bundles with a leaf morphology (Fig. 4). The FESEM image of silver nanoparticles shows spherical shaped particles (Fig. 5) (Cynthia Jemima Swarnavalli et al. 2011). The size of the particles is seen to be almost uniform. The FESEM image (Fig. 6) of CS–Ag nanocomposite shows a mixture of CS and Ag wherein the silver nanoparticles are seen to be enveloped by the chitosan polymer (Yoshizuka et al. 2000). The synthesized particles are in the form of aggregates. Prepared nanoparticles show low dispersibility and high stability. The reduction of agglomeration is seen to occur when the chitosan is allowed to dissolve for a longer duration of time, followed by the dispersion of silver nanoparticles in the chitosan solution for about an hour before the process of reduction. The FESEM image of such a well dispersed and stabilized synthesized composite is shown in Fig. 7. It is observed that the silver nanoparticles are embedded in a matrix of chitosan and also that the size of the composite is confined to the nano range.

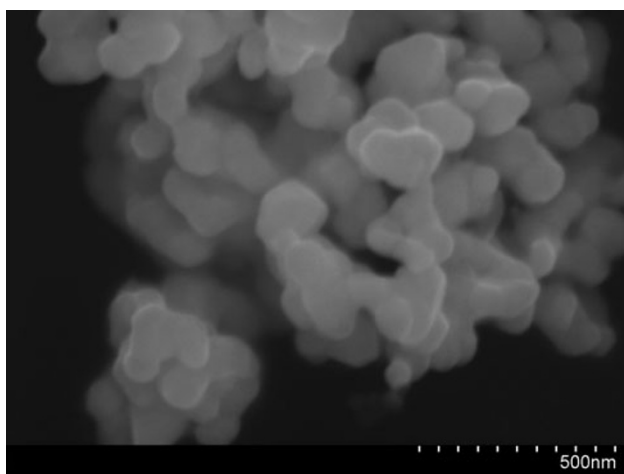
The FTIR spectrum (Fig. 8) of the silver nanoparticles shows O–H stretching at  $3,441$  and  $2,929 \text{ cm}^{-1}$  and N–H bending at  $1,639 \text{ cm}^{-1}$ , this is in agreement with the FTIR results of Mallikarjuna et al. (2011). The FTIR spectrum of chitosan shows O–H stretching at  $3,433 \text{ cm}^{-1}$ , C–H and C–N stretching at  $2,920 \text{ cm}^{-1}$ , N–H bending at  $1,647 \text{ cm}^{-1}$ , N–H angular deformation in CO NH plane at  $1,536 \text{ cm}^{-1}$  and C–O–C band stretching at  $1,109 \text{ cm}^{-1}$  which matches well with the report by Saraswathy et al. (2001) and Wazed Ali et al. (2001). In the FTIR spectrum of CS–Ag nanocomposite, the shifting of the CS peaks is observed which



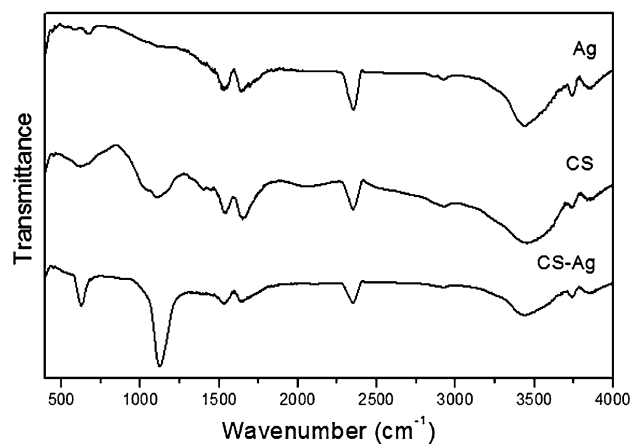
**Fig. 4** FESEM image of chitosan



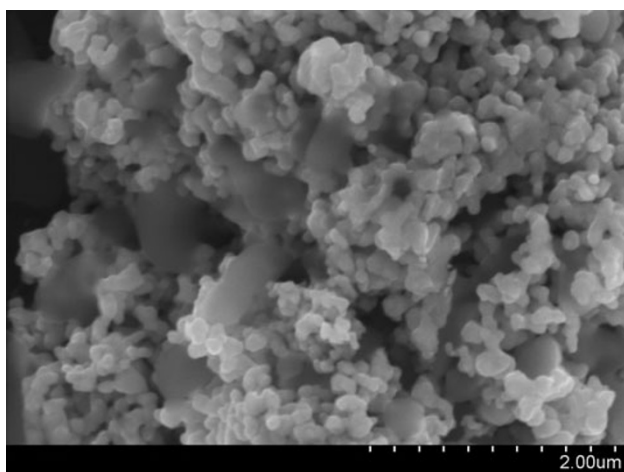
**Fig. 7** FESEM image of synthesized chitosan-silver nanocomposite exhibiting high dispersibility and stability



**Fig. 5** FESEM image of silver nanoparticles



**Fig. 8** FTIR spectrum of silver (Ag) nanoparticles, chitosan (CS), chitosan-silver (CS-Ag) nanocomposite



**Fig. 6** FESEM image chitosan-silver nanocomposite

may be due to the interaction of Ag with CS in the nanocomposite. The other changes that are noticeable are the reduction in the intensity of the hydroxyl peak and the

increase in the intensity of the C–O stretching which are in agreement with the FTIR results of Gupta et al. (2008) and An et al. (2011).

## Conclusions

Chitosan, silver nanoparticles and chitosan-silver nanocomposite was synthesized using simple chemical method. The XRD patterns of CS, Ag and CS-Ag are seen to be free from impurity peaks and the respective XRD patterns indicate the formation of Cs, Ag and CS-Ag nanocomposite clearly. The FESEM image of Ag indicates the formation of spherical nanoparticles which are of uniform size. The FESEM image of CS shows the formation of CS in the form of bundles. IN the FESEM image of the CS-Ag nanocomposite the silver nanoparticles are seen to be enveloped by CS, and agglomeration is also observed. It is

evident that agglomeration can be prevented on dispersing Ag in CS for a longer duration of time. The FESEM pattern obtained for the samples after sufficient dispersion shows the silver nanoparticles to be embedded in a matrix of CS. The shift in the peaks of CS observed in the FTIR spectrum indicates the formation of CS–Ag composite. It is evident from all the Characterizations done that the CS–Ag nanocomposites have been prepared successfully.

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