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Published on: 21 Oct 2012 - Key Engineering Materials (Trans Tech Publications)

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Antoine Boyer, David Marchat, Didier Bernache-Assollant. Synthesis and characterization of Cx-Siy-HA for bone tissue engineering application. Key Engineering Materials, Trans Tech Publications, 2013, 529-530, pp.100-104. 10.4028/www.scientific.net/KEM.529-530.100 . hal-00734811

HAL Id: hal-00734811

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Synthesis And Characterization Of C_x-Si_y -HA For Bone Tissue Engineering Application

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Keywords: co-substituted hydroxyapatite, carbonate, silicate, bioceramics, pellets, aqueous precipitation.

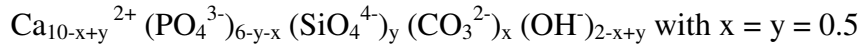
Abstract. The main goal of this work is to prepare carbon and silicon co-substituted calcium hydroxyapatite (C_x-Si_y -HA) for bone tissue engineering application. This study includes the synthesis of pure powders with a controlled amount of carbonate (x) and silicate (y) ions within the apatite structure, their characterization with the establishment of database for different compositions, and the manufacture of dense bioceramics. Carbon-silicon co-substituted hydroxyapatite ($C_{0.5}-Si_{0.5}$ -HA) powders are synthesized by aqueous precipitation. According to structural, spectroscopic and elemental characterizations, silicate and carbonate are included in the apatite lattice and their stoichiometries are controlled. The heat treatments under CO_2 atmosphere allow the sintering of pellets without decomposition of the apatite structure.

Introduction

Hydroxyapatite (HA), $Ca_{10}(PO_4)_6(OH)_2$ is the most commonly used bioceramics for dental and orthopedic implants due to its chemical and crystallographic similarities with bone's mineral part. Natural bone is comparable to a badly crystallized apatitic calcium phosphate with substitutions in its lattice. The apatite lattice is very tolerant with substitutions and vacancies [1]. All biological apatites contain carbonate ions in erratic quantities (between 3-8wt %) [2], filling preferentially in phosphate (B-type) compared to hydroxyl (A-type) positions in the apatite structure [3,4]. The composition depends on bone location, age, sex, *etc.* [3]. Carbonate substitution in apatite decrease the lattice crystallinity, as well as the temperature of the maximum rate of sintering [5,6], and increase its solubility in biological conditions (*in vitro* and *in vivo*) [4]. To improve bioactivity of synthetic grafts as resorption or osteogenesis, the literature talks over incorporation of carbonate or silicate groups in HA. Among trace elements existing in natural bone, silicon show essential characters for bone growth and development [7]. Even if the benefits of soluble silicon have been proved, the bioactivity of silicon substituted hydroxyapatite ceramics was not well defined. Results heterogeneity could be due to uncontrolled composition of Si-HA bioceramics [8]. Adequate solubility of carbonated hydroxyapatite and benefits of soluble silicon could be associated to improve bioactivity of apatite bioceramics. Some studies about carbon-silicon substituted hydroxyapatite were recently published [9–13]. Unfortunately, phase purity was not evidently proved. The main idea of this work was the preparation of a SiO_4 and CO_3 co-substituted hydroxyapatite bioceramics with controlled stoichiometry for bone tissue engineering. The study was first dedicated to the synthesis of C_x-Si_y -HA powders by aqueous precipitation, and second to the preparation of dense pellets.

Materials and methods

Synthesis of apatite powders. C_x-Si_y -HA powders were prepared by aqueous precipitation based on the method of Marchat *et al* [8]. This synthesis method allowed preparing monophasic silicon substituted hydroxyapatite ceramics with controlled composition (e.g. $Si_{0.5}$ -HA). The amount of reagents was calculated assuming silicate and carbonate substitution for phosphate ions according to the following theoretical formula:



A reagent molar ratio Ca/(P+Si+C) of 10/6 was used. Besides, materials of reference were prepared like a pure HA, C_x-HA, Si_y-HA or different C_x-Si_y-HA to establish databases.

A diammonium hydrogen phosphate aqueous solution ((NH₄)₂HPO₄, 99%, Merck, Germany), an ammonium hydrogen carbonate ((NH₄)HCO₃, 99%, Merck, Germany) and an alkaline silicate solution [8] were added to a calcium nitrate solution (Ca(NO₃)₂·4H₂O, 99%, Merck, Germany) by means of peristaltic pumps. Deionized water was used to dissolve the reagents. An argon flow (Air Liquide) was maintained on the reactor to prevent any atmospheric uncontrolled carbonation. The pH was kept at 10.8 by addition of a 28% ammonia solution (Merck, Germany) using a pH stat (Hanna instrument) and the temperature was regulated automatically at 50°C with an external cryothermostat. The suspension was continuously stirred. After complete introduction of reagents, suspension was matured for 48h, and then filtered on Whatmann filter (Ø > 6 µm). Finally, precipitates were dried at 100°C overnight.

Pellets preparation. Protocol consists to precalcine the powders to get a specific surface area around 30 m²/g. 160 mg of powder were first uniaxially pressed at 48 MPa in 8 mm diameter stainless steel die, then isostatically pressed at 300 MPa. Lafon *et al.* studied the synthesis of carbonated hydroxyapatite and their thermal stability under different atmospheres (P_{CO₂}, P_{H₂O}, P_{N₂}) [14,15]. According to their work and a complementary thermal study, it happens a decarbonation of carbonated and silico-carbonated hydroxyapatite powders beyond 800°C under air atmosphere [5,6,14]. To avoid this decomposition, heat-treatments and sintering of C_{0.5}-Si_{0.5}-HA samples were made under CO₂ atmosphere (P_{CO₂} = 1 atm). Heating rates were fixed at ±5°C/min.

Characterizations.

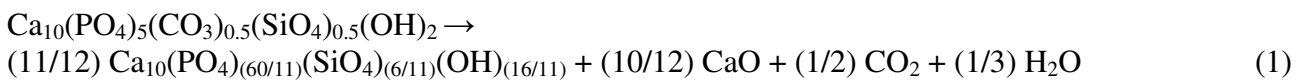
The crystalline phases of powders were studied using X-ray diffraction (XRD). The device used is a θ/2θ X-ray diffractometer, Siemens D5000, using CuKα radiation. The crystalline phases were evaluated on calcined powders over the range 27-38° with a step size of 0.02° and a count time of 9s by step. Phase identification was performed by comparison to standard patterns from International Center for Diffraction Data – Powder Diffraction Files (ICDD-PDF).

Fourier Transformed Infrared (FT-IR) absorption spectra of powders were performed using a Spectrometer MIR TF VERTEX 70. They were recorded over the range 400-4000 cm⁻¹ with a resolution of 2 cm⁻¹ and with an iteration of 64 successive scans. All spectra were normalized from the ν₄ band of the phosphate group at 600 cm⁻¹ as usual [4].

Calcium, phosphorous and silicon contents of powders were determined by elemental analysis via Inductively Coupled Plasma Atomic Emission Spectrometry (ICP/AES) (Spectrometer HORIBA, Jobin-Yvon, with Activa model). Powders samples were dissolved in nitric acid solution (pH < 2). Likewise, carbon content of powders was determined by an elemental analyzer using an infrared detector (LECO CS-444 carbon and sulfur analyzer).

Results and discussions

Structural analysis. The XRD patterns of heat-treated (under CO₂) C_{0.5}-Si_{0.5}-HA powders and C_{0.5}-Si_{0.5}-HA pellets (Fig. 1) show an apatitic well crystallized structure (PDF 9-432). No secondary crystallized phase is detected, as the main characteristic line of CaO at 2θ = 37.347° (PDF 37-1497). Thus, the decarbonation of B-site, according to the following reaction (Eq. 1) is not observed [14]:



The CO₂ atmosphere keeps carbonate in apatite lattice and avoids phase decomposition. Thereby, the results indicate that silicon and carbon can be incorporated in the apatitic structure, or in an amorphous phosphate, or both.

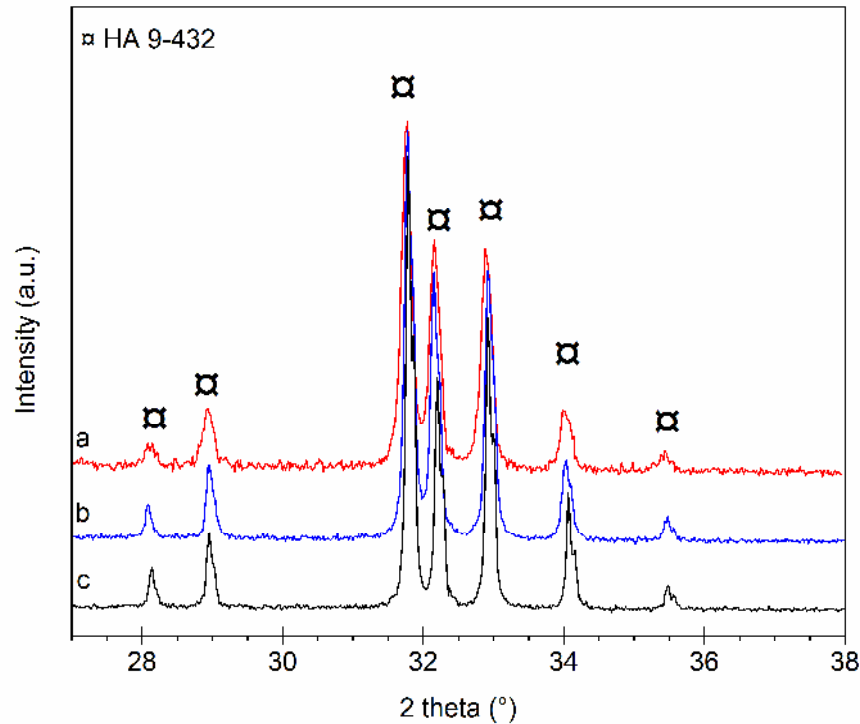


Fig. 1: XRD patterns: (a) C_{0.5}-Si_{0.5}-HA pellets after sintering under CO₂ ; (b) C_{0.5}-Si_{0.5}-HA powders after calcination under CO₂ ; (c) HA powders in reference after calcination 1000°C-15h

Spectrometric analysis. The FTIR patterns of C_{0.5}-Si_{0.5}-HA powders and C_{0.5}-Si_{0.5}-HA pellets after heat-treatment under CO₂ (Fig. 2) reveal the same main phosphate bands as spectra of heat treated HA under air atmosphere : 472 cm⁻¹ (ν₂), 562 and 600 cm⁻¹ (ν₄), 959 cm⁻¹ (ν₁), 1015 and 1080 cm⁻¹ (ν₃). Characteristics bands of carbonate groups are shown on spectra of C-Si-HA samples: from B site (872 cm⁻¹, 1409 cm⁻¹, 1450 cm⁻¹, 1466 cm⁻¹), and A site (754 cm⁻¹, 878 cm⁻¹, 1545 cm⁻¹). The CO₂ atmosphere seems enriched the A sites along the channels in CO₃. Indeed, the ν_S and ν_L modes of hydroxyl group (629 cm⁻¹ and 3572 cm⁻¹) are not observed anymore. More, spectra of C_{0.5}-Si_{0.5}-HA display the specific bands attributed to SiO₄ substitution in hydroxyapatite structure: 503 cm⁻¹ (ν₂ : SiO₄), 521 cm⁻¹ (ν₄ : SiO₄), 750 cm⁻¹ (ν₁ : SiO₄), 846 cm⁻¹ (ν₃ : SiO₄), and 922 cm⁻¹ (Si-OH) [8]. Besides, an extra band appeared at 945 cm⁻¹. This large band results from the degeneration of ν₁ PO₄³⁻ domain at 960 cm⁻¹ due to the presence of OH⁻ vacancies and CO₃²⁻ ions along the channel in the environment of the phosphate ions. Finally, no evidence of secondary compounds as CaO or TCP is detected. Likewise, the SiO₂ specific bands (e.g. ~680 cm⁻¹, ~792 cm⁻¹ and ~870 cm⁻¹ [8]) are not observed. To conclude, results prove that C_{0.5}-Si_{0.5}-HA samples (powders and pellets) are monophasics with carbonate and silicate ions localized in the apatitic lattice.

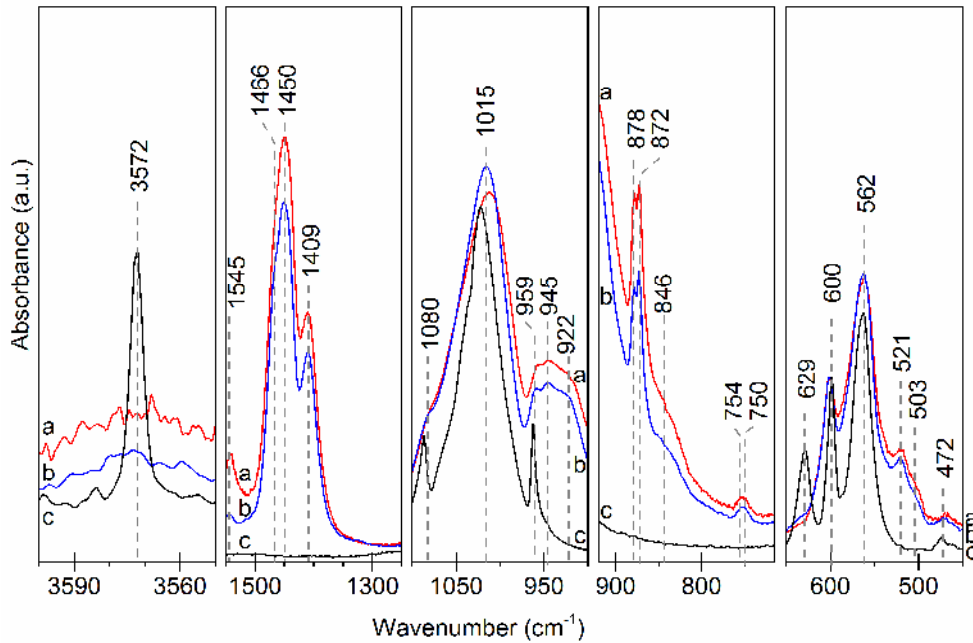


Fig. 2: FTIR patterns: (a) $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ pellets after sintering under CO_2 ; (b) $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ powders after calcination under CO_2 ; (c) HA powders in reference after calcination 1000°C-15h

Elemental analysis. $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ powder and $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ pellet elemental compositions were measured after heat treatment at 1000°C for 1h under CO_2 atmosphere (Table 1). The Ca/P molar ratios are equal to 1.961 and 2.042 for $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ powder and pellet respectively. These results are in coherence with the theoretical ratio Ca/P of reagent introduced equal to 2. For the powders and pellets heat treated under CO_2 atmosphere, samples present a Ca/(P+Si+C) molar ratios inferior than estimated, respectively 1.546 and 1.567. In the ideal case of total co-substitution of carbonate and silicate in B anionic site, the Ca/(P+Si+C) molar ratios for a $C_{0.5}\text{-Si}_{0.5}\text{-HA}$ composition should be equal to $10/6$ ($=1.667$). Difference can be explained by an increase of carbon amount in compounds, *i.e.*, in A site, due to calcination under CO_2 atmosphere. The substitution of OH^- by CO_3^{2-} along the channels throughout this heat treatment is confirmed FTIR spectrometry.

Table 1 : Results from elemental analysis

Compound	Heat treatment	Sintering	Ca/P measured	Ca/(P+Si+C) measured
$C_{0.5}\text{-Si}_{0.5}\text{-HA}$ powder	$1000^\circ\text{C-1h-CO}_2$	/	1.961	1.546
$C_{0.5}\text{-Si}_{0.5}\text{-HA}$ pellet	$835^\circ\text{C-1h-CO}_2$	$1000^\circ\text{C-1h-CO}_2$	2.042	1.567

Conclusion

Carbon and silicon co-substituted calcium hydroxyapatite (C-Si-HA) powders were synthesis by aqueous precipitation. Silicate and carbonate were incorporated in the apatite lattice and their stoichiometry was controlled. The heat treatments under CO_2 atmosphere allowed the sintering of pellets without decomposition of the apatite structure. Powders and pellets obtained were monophasic apatite polysubstituted with silicate and carbonate ions and free of secondary phase. Studies about the improvement of the sintering conditions are in progress, as well as *in vitro* and *in vivo* biological evaluation of dense pellets and scaffolds.

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