



Synthesis and Photostability Study of Some Modified Poly(vinyl chloride) Containing Pendant Benzothiazole and Benzimidazole Ring

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Abstract

The influence of introducing benzothiazole and benzimidazole as a pending groups into the repeating unit of PVC has been studied on the bases of photostability measurements. The photostability of modified polymers were studied by monitoring the carbonyl Index (I_{CO}), polyene index (I_{PO}) and hydroxyl index (I_{OH}), all results obtained indicated that PAA, PSS, PNN, PSA and PNA gave lower growth rate of I_{CO} , I_{PO} and I_{OH} with irradiation time with respect to unmodified PVC film (control). The following trends of photostability are obtained



According to the experimental results several reaction mechanisms were proposed depending on the presence of benzothiazole and benzimidazole moieties in the polymer chain. The mode of chain scission was also investigated by measuring the variation of average viscosity molecular weight with irradiation time for modified PVC. The degree of deterioration (α) and the average number of chain scission (S) values were also calculated.

Keywords: PVC, UV-light, Heterocyclic

1. Introduction

Synthesis of polymer-bound chelating ligands and the selective chelation of specific metal ions is a field of active research (Pittmann 1980). A number of ligands including polydentate amines, crown ethers, phosphines, bipyridines, and naphthyridines have been bound with mainly polystyrene divinylbenzene copolymers (Jellinek 1978 & Gardette *et al.* 1993). These studies are mostly concerned with ion-binding and catalytic aspects. Very little information appears to exist as to whether (Andrady *et al.* 1998), and to what extent, such polymer modification can influence the properties of the macromolecule (Aliwi *et al.* 2005). Recently, scientists does able to modified (Yousif *et al.* 2007) PVC by introduction aromatic and heterocyclic moieties through halogen displacement reaction. PVC, thus modified, showed improved overall photochemical stability. The facial chlorine displacement from PVC indicated the possibility on easy anchoring of ligands to PVC matrix. In view of paucity of any information on PVC in this line, we undertook the synthesis of PVC-heterocyclic compounds.. This article will describe the synthetic procedure adopted (Scheme 1) and the photostability of these new polymers towards UV light.

2. Materials and methods

2.1 Preparation of modified polymers

2.1.1 Preparation of modified polymer (PAA)

2-amino benzothiazole (1.67mmole) was added to (0.125g) poly (vinyl chloride) and three drops of pyridine in (25 mL) tetrahydrofuran (THF) was also added. A white precipitate was formed after refluxing the mixture for three hours the modified polymer (PAA) was filtered, washed with redistilled water, ethanol, ether and dried under vacuum.

2.1.2 Preparation of modified polymer (PSS)

2-mercapto benzothiazole (1.49mmole) was added to (0.125g) poly(vinyl chloride) and three drops of pyridine in (25 mL) tetrahydrofuran (THF) was also added. A yellow precipitate was formed after refluxing the mixture for three hours, the modified polymer (PSS) was filtered, washed with redistilled water, ethanol, ether and dried under vacuum.

2.1.3 Preparation of modified polymer (PNN)

2-Amino benzimidazole (1.87mmole) was added to (0.125g) poly(vinyl chloride) and three drops of pyridine in (25 mL) tetrahydrofuran (THF). A brown precipitate was formed after refluxing the mixture for three hours the modified polymer (PNN) was filtered, washed with redistilled water, ethanol, ether and dried under vacuum.

2.1.4 Preparation of modified polymer (PSA)

A solution of mixture of (0.83mmole) 2-amino benzothiazole and (0.75mmole) of 2-mercapto benzothiazole and three drops of pyridine in (25 mL) tetrahydrofuran (THF) was added to (0.125g) poly(vinyl chloride). A brown precipitate was formed after refluxing the mixture for three hours, the modified polymer (PSA) was filtered, washed with redistilled water, ethanol, ether and dried under vacuum.

2.1.5 Preparation of modified polymer (PNA)

A solution of mixture of (0.94mmole) 2-amino benzimidazole and (0.83mmole) 2-amino benzothiazole and three drops of pyridine in (25 mL) tetrahydrofuran (THF) was added to (0.125g) poly(vinyl chloride). A brown precipitate was formed after refluxing the mixture for three hours, the modified polymer (PNA) was filtered, washed with redistilled water, ethanol, ether and dried under vacuum. The physical properties of prepared modified polymers are listed in Table (1)

2.2 Photodegradation measuring methods

2.2.1 Films Preparation

A solution of Poly(vinyl chloride) solution or modified Poly(vinyl chloride) (5 g /100 ml) in tetrahydrofuran (THF) was used to prepare (30 μm) thickness of polymer films, (measured by a micrometer type 2610 A, Germany). The films were prepared by evaporation technique at room temperature for 24 hours. To remove the possible residual tetrahydrofuran solvent, film samples were further dried at room temperature for three hours under reduced pressure.

2.2.2 Accelerated testing technique

Accelerated weatherometer Q.U.V. tester (Q. panel, company, U S A), was used for irradiation of polymers films. The accelerated weathering tester contains stainless steel plate, which has two holes in the front side and a third one behind. Each side contains a lamp (type Fluorescent Ultraviolet Lights) 40 Watt each. These lamps are of the type UV-B 313 giving spectrum range between 290-360 nm with a maximum at wavelength 313nm. The polymer film samples were vertically fixed parallel to the lamps to make sure that UV incident radiation is perpendicular on the samples. The irradiated samples are rotated from time to time to ensure that the intensity of light incident on all samples is the same.

2.2.3 Photodegradation measuring methods

A. Measuring the photodegradation rate of polymer films using infrared spectrophotometry

The degree of photodegradation of polymer film samples was followed by monitoring FTIR spectra in the range 4000-400 cm^{-1} using FTIR 8300 Shimadzu Spectrophotometer. The position of carbonyl absorption is specified at 1722 cm^{-1} , polyene group at (1602 cm^{-1}) and the hydroxyl group at 3500 cm^{-1} . The progress of photodegradation during different irradiation times was followed by observing the changes in carbonyl and polyene peaks. Then carbonyl (I_{co}), polyene (I_{po}) and hydroxyl (I_{OH}) indices were calculated by comparison of the FTIR absorption peak at 1722, 1602 and 3500 cm^{-1} with reference peak at 1328 cm^{-1} , respectively. This method is called band index method which includes (Rabek & Ranby 1975):

$$I_s = \frac{A_s}{A_r} \quad (1)$$

A_s = Absorbance of peak under study.

A_r = Absorbance of reference peak.

I_s = Index of group under study.

Actual absorbance, the difference between the absorbance of top peak and base line (A Top Peak – A Base Line) is calculated using the Base Line method (Rabek & Ranby 1975).

B. Determination of average molecular weight (\overline{M}_v) using viscometry method:

The viscosity property was used to determine the average molecular weight of polymer, using the Mark- Houwink relation (Mark 2007).

$$[\eta] = K\bar{M}_v^\alpha \quad (2)$$

$[\eta]$ = the intrinsic viscosity.

K, α = are constants depend upon the polymer-solvent system at a particular temperature.

The intrinsic viscosity of a polymer solution was measured with an Ostwald U-tube viscometer. Solutions were made by dissolving the polymer in a solvent (gm/100ml) and the flow times of polymer solution and pure solvent are t and t_0 respectively. Specific viscosity (η_{sp}) was calculated as follows:

$$\eta_{re} = \frac{t}{t_0} \quad (3)$$

η_{re} = Relative viscosity.

$$\eta_{sp} = \eta_{re} - 1 \quad (4)$$

The single – point measurements were converted to intrinsic viscosities by the relation 2.

$$[\eta] = (\sqrt{2}/c)(\eta_{sp} - \ln\eta_{re})^{1/2} \quad (5)$$

C = Concentration of polymer solution (g /100 ml).

By applying equation 5, the molecular weight of degraded and un-degraded polymer can be calculated. Molecular weights of PVC with and without additives were calculated from intrinsic viscosities measured in THF solution using the following equation;

$$[\eta] = 1.38 \times 10^{-4} M_v^{0.77} \quad (6)$$

The quantum yield of main chain scission (ϕ_{cs}) (Nakajima *et al.* 1990) was calculated from viscosity measurement using the following relation (7).

$$\phi_{cs} = (CA/\bar{M}_{v,0}) \left[([\eta_0]/[\eta])^{1/\alpha} - 1 \right] / I_0 t \quad (7)$$

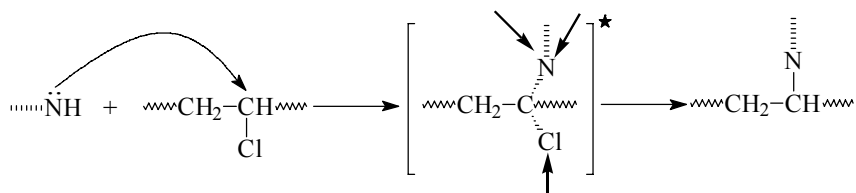
Where: C = concentration; A = Avogadro's number; $(\bar{M}_{v,0})$ = the initial viscosity-average molecular weight; $[\eta_0]$ =

Intrinsic viscosity of PVC polymer before irradiation; I_0 = Incident intensity and t = Irradiation time in second.

3. Results and Discussion

3.1 Synthesis of PVC bound benzothiazole and benzimidazole ring system

Polymers with an aromatic and heterocyclic rings have been drawing the attention of researchers more due to their importance in many aspects of interest, among these polymers are modified poly(vinyl chloride). PVC are modified by introducing aromatic and heterocyclic moieties through halogen displacement reaction mainly by S_{N2} mechanism, modified PVC polymers have developed over the past decade into a fast growing industrial application field. They have also attracted attention because of their importance as catenation ligands. Also, the thermal stabilities of (PVC-aromatic). Polymers are comparable better than unmodified PVC. In addition, most of them showed improved changes in electro physical properties, especially in dielectric properties. In continuation of our earlier studies on the synthesis of modified PVC, we undertook in this research the synthesis of some new modified PVC containing heterocyclic systems viz., benzothiazole and benzimidazole (polymers prepared), with a view to finding out the effect of introducing these moieties on physical and chemical properties of PVC. The reaction was initiated by the nucleophilic attack of (S,N) atoms on the carbon carrying chlorine atom in the polymeric chain followed by departure of chloride anion as a good leaving group.



3.2 Characterization of modified polymers

3.2.1 Infra-red Spectroscopy

The structure of the present polymers was confirmed by Infra-red spectroscopy. The most common bands are tabulated in Table (2), beside the typical IR absorption bands, each of the present polymers showed other band characteristic of its own structure, Polymers showed the following features, a medium to strong band at about $(630) \text{ cm}^{-1}$ which could be attributed to (C-Cl), the appearance of a medium band at about $(1550) \text{ cm}^{-1}$ indicated that the end groups of these polymers are C=C double band, the bands at around $(2900) \text{ cm}^{-1}$ which was assigned to the typical aliphatic (C-H) stretching vibration was also observed, moreover, these polymers exhibited significant bands in the region at (3050) for ν (C-H aromatic), 1635 for ν (C=N) and at 750 for (C-H bending of *O*-disubstituted benzene) that clearly indicated the presence of the benzothiazole and benzimidazole nucleus in the backbone of PVC chain.

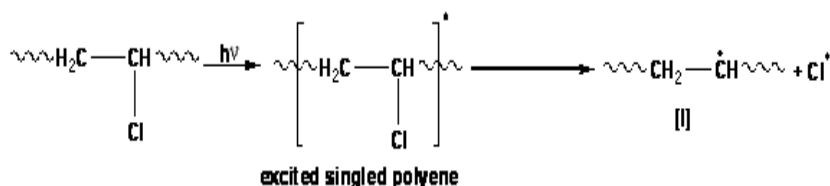
3.2.2 Ultraviolet-visible Spectroscopy

The ultraviolet-visible spectrophotometry technique is used to characterize the polymers in tetrahydrofuran (THF) as a solvent. The ultraviolet visible electronic spectra of the modified polymers, showed an absorption bands could be attributed to the $\pi \rightarrow \pi^*$ electronic transitions, These transitions are assigned in relevance to the structures of polymers. As shown in Table (3).The increase of the absorbance in the uv-vis range for samples containing modified PVC can be explained by the formation of conjugation double bonds (band corresponding to the $\pi \rightarrow \pi^*$ transition) in these modified polymers resulting from introducing an aromatic benzothiazole and benzimidazole rings. Thus, shifting the absorbance to longer wavelengths (i.e the bathochromic effect) is good evidence that modification were formed in our PVC. Moreover, the broadness in the bands caused by the presence of other chromophores such as (C=N, etc) which absorb in the above region and this overlapping bands make the interpretation of the results difficult.

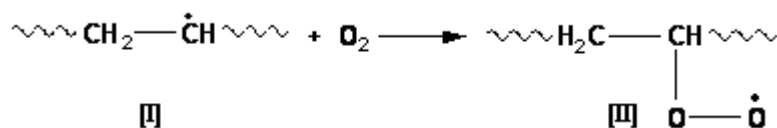
3.3 Photostabilization of modified PVC film

The new modified polymers PAA, PSS, PNN, PSA and PNA. were used as photostabilized modified PVC comparing with PVC films(control).In order to study the photochemical activity of these modified polymers for the photostabilization of films, the carbonyl, polyene and hydroxyl indices were monitored with irradiation time using IR spectroscopy. The formation of carbonyl group (C=O) and polyene (C=C)_n and (O-H) during photolysis of PVC film could be illustrated by the following mechanism.

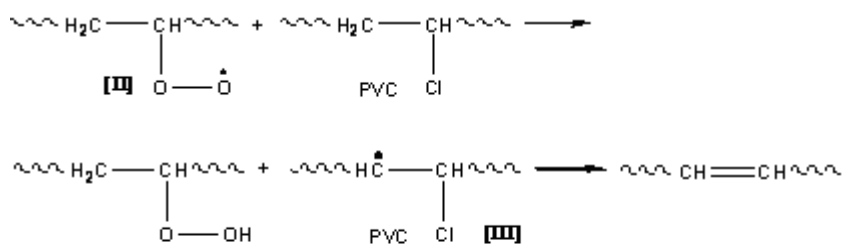
When PVC is photolysed in the presence of (O_2), the primary photochemical processes occur (the excited singlet polyenes).



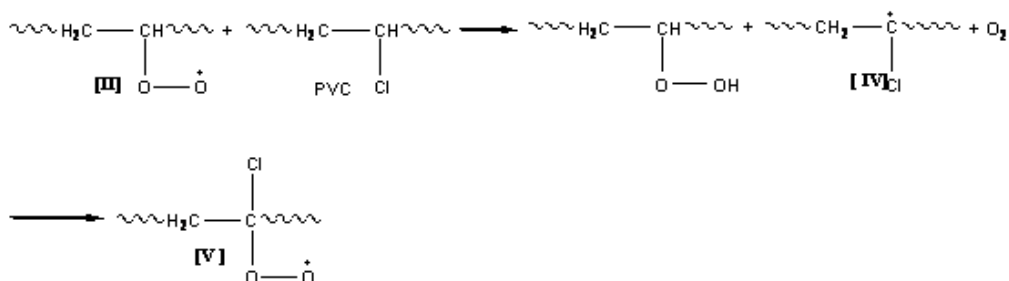
Those excited states will thus disappear by the different routes including the allylic (C-Cl) bond cleavage with formation of (Cl) radical and polyene radical [I]. The polyenyl radical [I] which bears no (Cl) atom in the α -position is very likely to be scavenged by (O_2) to give peroxy radical [II]



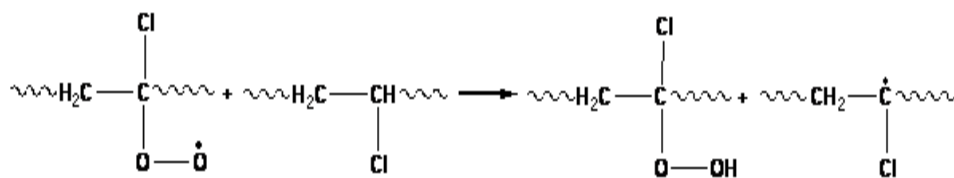
Rate constant measurement on related model compounds suggest that peroxy radical [II] react with the ($-\text{CH}_2-$) and (CH-Cl) groups on PVC at rates that are comparable. The attack on (CH_2) group yields radical [III].



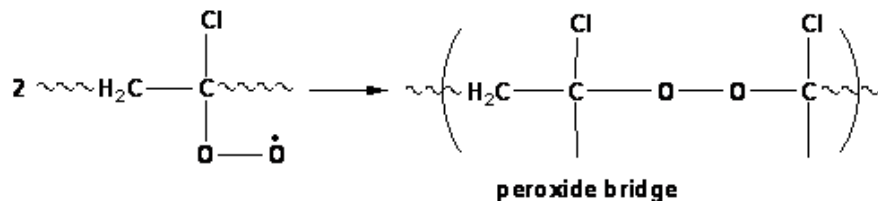
This radical contributes to the chain-dehydrochlorination of PVC. The attack of II on the (-CH-Cl) group yield radical [IV] which possesses no labile β -chlorine and is likely to react with (O_2) to give a γ -chloroalkyl peroxy radical [V].



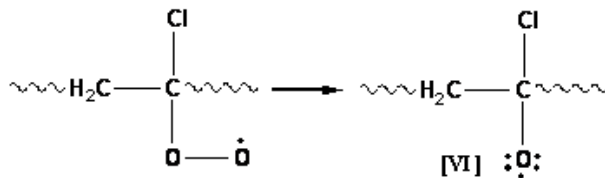
The main oxidation products of PVC are expected to result primary from the various reactions of this radical [V], there are two major routes of the fate (disappearances) of the γ -chloroalkyl peroxy radical. The hydrogen abstraction from PVC with the formation of hydroperoxide groups.



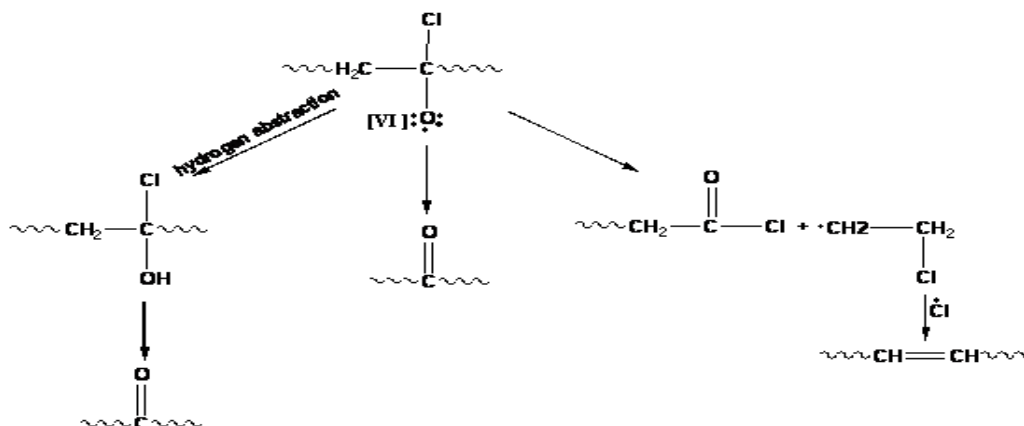
The bimolecular interaction, the later leads to the formation of either a peroxide bridge (termination reaction),



the formation of alkoxy radicals [VI].



The most usual reaction of alkoxy radicals [VI] is the hydrogen abstraction. The unstable α -chloroalcohol formed by reaction of alkoxy radical with PVC will rapidly decompose into the corresponding ketone.



Tertiary alkoxy radicals are also proved to be stabilized by β -scission which may involve either (C-Cl) bond cleavage to form ketone or/and (C-C) bond cleavage to form polyene. The irradiation of PVC films with light of wavelength, $\lambda = 313\text{nm}$ lead to a clear change in their FTIR spectrum, as shown in Figure (1). Appearance of bands at 1772 cm^{-1} and 1722 cm^{-1} respectively,

attributed to the formation of carbonyl group, the first is related to chloro ketone with the second one is due to aliphatic ketone, together with formation of a band at 1604 cm^{-1} related to polyene group. The hydroxyl band appeared at 3500 cm^{-1} is related to alcohol functions (Andrady & Searle 1989). The absorption of the polyene, carbonyl and hydroxyl groups are used to follow the extend of polymer degradation during irradiation. This absorption was calculated as carbonyl index (I_{CO}), polyene index (I_{PO}) and hydroxyl index (I_{OH}). Therefore, one should expect that the growth of carbonyl index is a measure to the extent of degradation. As seen from Figure (2) that the modified polymers PAA, PSS, PNN, PSA and PNA show lower growth rate of carbonyl index with irradiation time with respect to unmodified PVC film (control). these modified polymers might be considered as photostabilized with respect to the PVC polymer. The good photostabilizer show a longer induction period. Therefore, from Figure (3), the PAA is the most active photostabilizer, followed by PSS, PNN, PSA and then PNA which is less active. As it was mentioned before polyene compounds are produced during photodegradation of PVC. Therefore, polyene index (I_{PO}) was monitored with irradiation time. The same trend obtained as carbonyl index. Hydroxyl functions are produced during photodegradation of PVC. Therefore, hydroxyl index (I_{OH}) was monitored with irradiation time for unmodified PVC and modified PVC (PAA, PSS, PNN, PSA and PNA) polymers. From Figure (4) the PAA, PSS, PNN, PSA and PNA show lower growth rate of hydroxyl index with irradiation time with respect to PVC film without modification (control).

3.3.1 Variation of PVC molecular weight during photolysis

An analysis of the relative change in viscosity average molecular weight (\overline{M}_v), Figure (5) shows the plot of (\overline{M}_v) versus irradiation time for modified PVC films and PVC film, with absorbed light intensity of $1.43 \times 10^{-8}\text{ ein. dm}^{-3}\text{ s}^{-1}$. It is worth mentioning that traces of the films with additives are not soluble in THF indicating that cross-linking or branching in the PVC chain does occur during the course of photolysis. The plots indicate a rapid decrease in (\overline{M}_v) initially then it slows down, suggesting that the initial rapid drop in (\overline{M}_v) is due to the main chain scission at various locations that distributed along the polymer chain. The photodegradation becomes slower and bond scission may be random. If the scission of bond is random, the rate of decrease in (\overline{M}_v) should be proportional to the square of the viscosity – average molecular weight at that irradiation time. Therefore, in order to clarify the type of bond scission the rate in the decrease of (\overline{M}_v) with time, the value of $\left(\frac{d\overline{M}_v}{dt}\right)$ was calculated using the relation (8).

$$\frac{d\overline{M}_v}{dt} = \frac{\overline{M}_{v,0} - \overline{M}_{v,t}}{t} \quad (8)$$

Where $\overline{M}_{v,0}$ and $\overline{M}_{v,t}$ are viscosity average molecular weight at (zero) and (t) irradiation time respectively.

The plots Figure (6), shows $\frac{d\overline{M}_v}{dt}$ against $(\overline{M}_v)^2$, which were found to be not linear. The plots show that the weak bond

links is randomly distributed along the polymer chains. For randomly distributed weak bond links, which break rapidly in the initial stages of photodegradation, the degree of deterioration α is given as:

$$\alpha = \frac{m}{\overline{Mv}_0} \quad (9)$$

Where m and Mv_0 are the molecular weights of irradiation time, respectively. The plot of α as a function of irradiation time is shown in Figure (7). The curves indicate that the degradation is taking place by the random breaking of one type of bond. The values of α of the irradiated samples are higher without presence of additives and lower in presence of additives compared to the corresponding values of the additive free PVC. In the initial stages of photodegradation of PVC, the values of α increases rapidly with time, which indicates a random breaking of bonds in the polymer chain. For more support of this view, the number average chain scission (average number cut per single chain) (S) was calculated using the relation (10):

$$S = \frac{\overline{M}_{v,0}}{\overline{M}_{v,t}} - 1 \quad (10)$$

Where $\overline{M}_{v,0}$ and $\overline{M}_{v,t}$ are viscosity average molecular weight at (zero) and (t) irradiation time respectively. The plot of S versus time is shown in Figure (8). The curve indicates an increase in the degree of branching such as might arise from cross-linking occurrence. It is observed that, insoluble material was formed during irradiation, an extra evidence to the idea that cross-linking dose occur. The plot of $1/Pt$ versus irradiation time is adopted to characterize degradation reaction of polymer as shown in Figure (9). The curve indicates that increase in the inverse of number average degree of polymerization with irradiation time

3.4 Suggested mechanisms of photostabilization of PVC by modification

Through the overall results obtained, the efficiency of modified polymers PVC films can be arranged according to the change in the carbonyl, polyene and hydroxyl Indexes as a reference for comparison as shown in Figure (2), (3) and (4).



1-The reaction between PVC and heterocyclic units reduce the number of HCl in this polymer which led to decrease the probability of labrtaion Cl from PVC chain. (Scheme 2).

2-These polymers also function as radical scavengers through energy transfer and by forming unreactive charge transfer complexes between the modified polymers and excited state of the chromophore (POO^*) and stabilize through resonating structures as shown in Scheme (3).

3- The benzothiazole and benzimidazole rings in these modified polymers plays a role in the mechanism of the stabilizer process by acting as UV absorber. The UV light absorption by these polymers containing these heterocyclic units dissipates the UV energy to harmless heat energy Scheme (4). Further more, this ring play a role in resonating structures conjugation of radical in peroxide decomposer Scheme (3), which support these rings as photostabilizer. Another way of degradation reaction characterization is the measurements of the quantum yield of the chain scission (Φ_{cs}). The quantum yield for chain scission was calculated for PVC films of modified polymers. The Φ_{cs} values for these polymers are tabulated in Table (4). The Φ_{cs} values for modified PVC films are less than that of unmdified PVC (control), which increase in the order:



The explanation of low values of Φ_{cs} is due to that, in large molecule of PVC the energy is absorbed at one site, and then the electronic excitation is distributed over many bonds so that the probability of a single bond breaking is small, or the absorbed energy is dissipated by non reactive processes. It is well established that the quantum yield (Φ_{cs}) increases with increasing temperatures around the glass transition temperature, (T_g) of the amorphous polymer, and around the melting temperature of crystalline polymers. In the study presented in this work, the photolysis of PVC film is carried out at a temperature (25 – 40 °C) well blow the glass transition temperature (T_g of PVC = 80 °C). Therefore, the Φ_{cs} dependency on temperature is not expected to be observed.

4. Conclusions

We prepared five modified polymers by the covalent modification of commercial PVC with benzothiazole and benzimidazole binding units. The modified polymers differ in their susceptibility to photooxidative degradation from that unmodified PVC. The photostability effects of these heterocyclic moities was indicated on the basis of carbonyl,

polyene and hydroxyl indexes. The results obtained demonstrates that all the modified polymers are good photostabilizer rather than unmodified PVC.

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Table 1. The Softening point of the prepared polymers

Polymer	Color	Softening point (⁰ C)
PAA	White	92-95
PSS	Yellow	103-106
PNN	Brown	161-167
PSA	Brown	141-146
PNA	Brown	130-138
PVC	Colorless	200-205

Table 2. Characteristic absorption bands of polymers

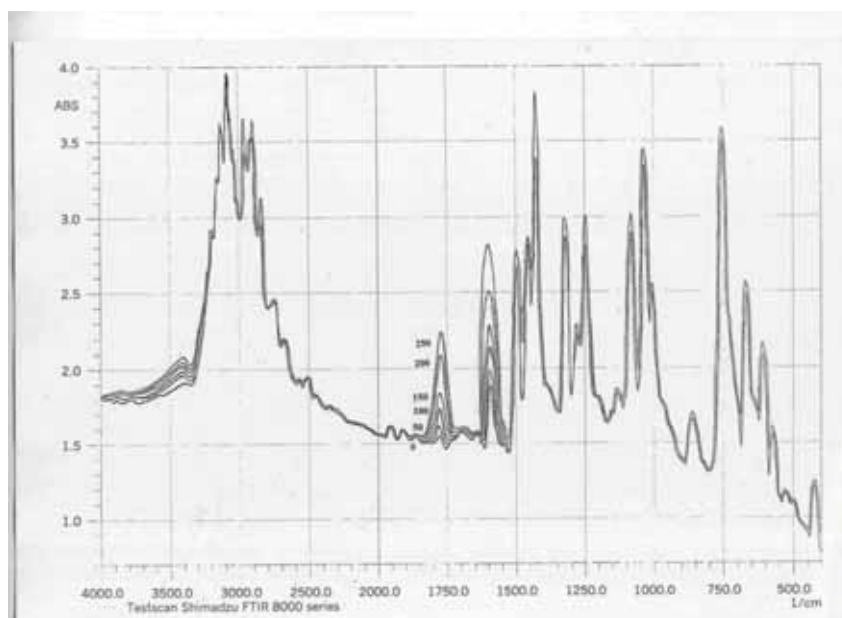
Polymer Symbol	IR Spectral Data (cm ⁻¹)	Assignment
PAA	3274 3058 2923 1635 1531 1438 1303 763 628	N-H stretching C-H stretching of benzene ring C-H stretching of -CH ₂ - C=N stretching of thiazole ring C=C stretching of end of PVC chain C-H bending of -CH ₂ - C-H bending of -CH ₃ C-H bending of ortho disubstituted benzene ring C-Cl stretching
PSS	3025 2900 1630 1589 1434 1350 655	C-H stretching of benzene ring C-H stretching of -CH ₂ - C=N stretching of thiazole ring C=C stretching of end of PVC chain C-H bending of -CH ₂ - C-H bending of -CH ₃ C-Cl stretching
PNN	3375 3070 2931 1646 1569 1458 1319 732 621	N-H stretching C-H stretching of benzene ring C-H stretching of -CH ₂ - C=N stretching (being broad) C=C stretching of end of PVC chain C-H bending of -CH ₂ - C-H bending of -CH ₃ C-H bending of ortho disubstituted benzene ring C-Cl stretching
PSA	3200 3078 2896 1616 1531 748 651	N-H stretching (being broad) C-H stretching of benzene ring C-H stretching of -CH ₂ - C=N stretching (being broad) C=C stretching of end of PVC chain C-H bending of ortho disubstituted benzene ring C-Cl stretching
PNA	3270 2931 1643 1539 1450 1328 750 628	C-H stretching of benzene ring C-H stretching of -CH ₂ - C=N stretching of imidazole ring C=C stretching of end of PVC chain C-H bending of -CH ₂ - C-H bending of -CH ₃ C-H bending of ortho disubstituted benzene ring C-Cl stretching N-H stretching (being broad)
PVC	2970,2910,2860 1429 615.2	C-H stretching of -CH ₂ and -CH- C-H bending of -CH ₂ - C-Cl stretching

Table 3. Electronic spectra of polymers in THF solvent

Polymer	Absorption Bands(nm)
PAA	234, 262, 293
PSS	230, 324
PNN	224, 276, 281
PSA	226, 265
PNA	234, 249, 270
PVC	219, 254

Table 4. Quantum yield (Φ_{cs}) for the chain scission for PVC and modified PVC films (30 μ m) thickness after 250 hrs. irradiation time

Additive (0.5%wt)	quantum yield of main chain scission (Φ_{cs})
PAA	1.11E-07
PSS	1.36E-07
PNN	1.39E-07
PSA	1.59E-07
PNA	2.82E-07
PVC	5.12E-05

Figure 1. Change in IR spectrum of PVC film (30 μ m) in the presence of Cd(L)₂ complex. (Numbers n spectra are irradiation time in hours)

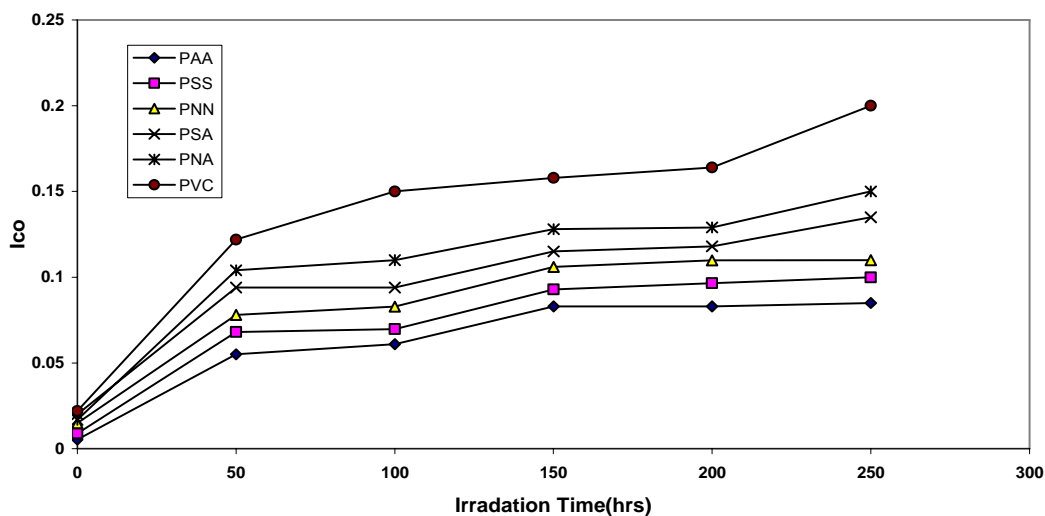


Figure 2. The relationship between the carbonyl index and irradiation time for PVC and modified PVC films (30 μm) thickness

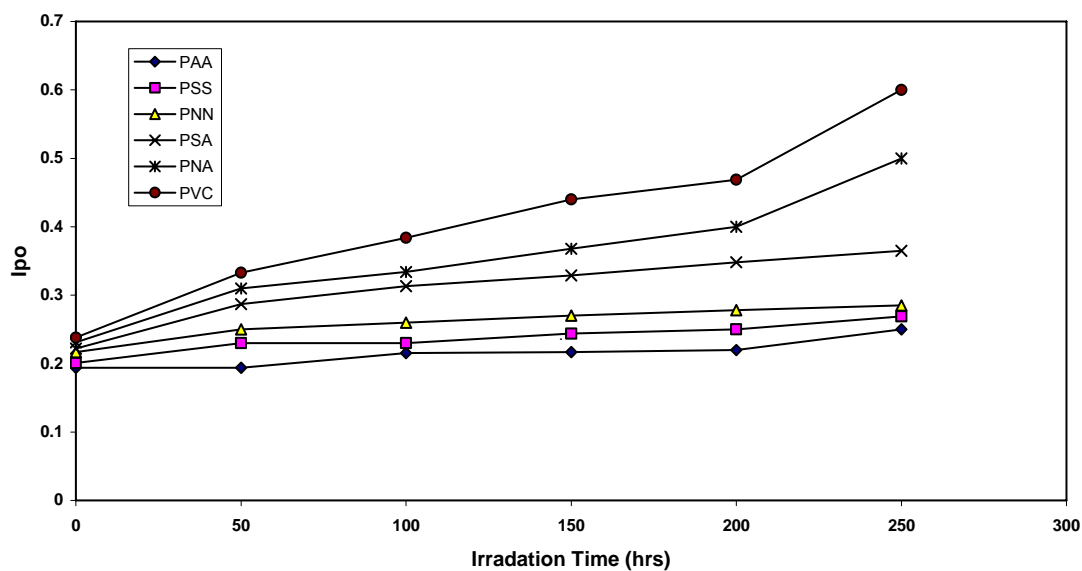


Figure 3. The relationship between the polyene index and irradiation time for PVC and modified PVC films (30 μm) thickness

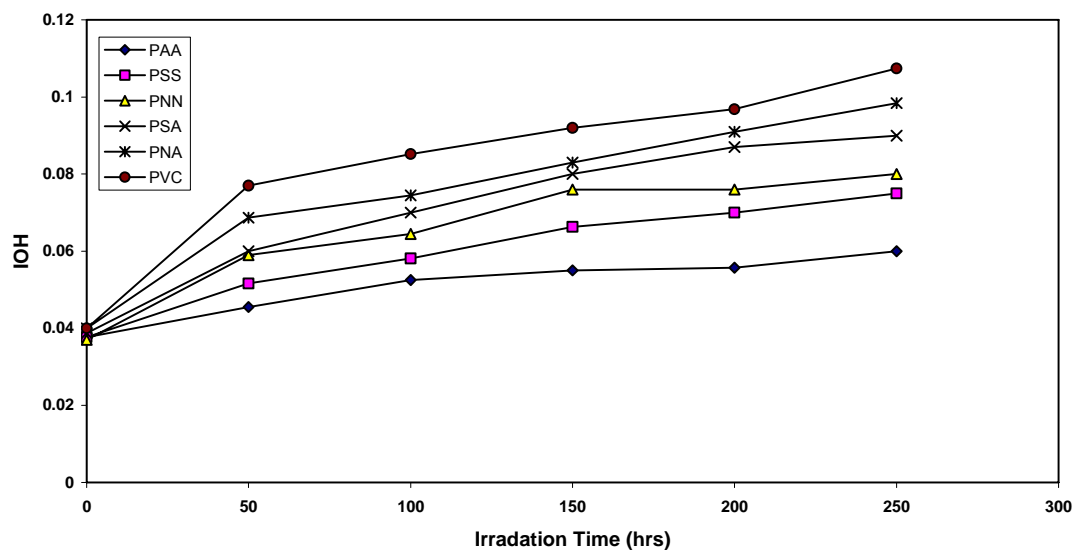


Figure 4. The relationship between the hydroxyl index and irradiation time for PVC and modified PVC films (30 μm) thickness

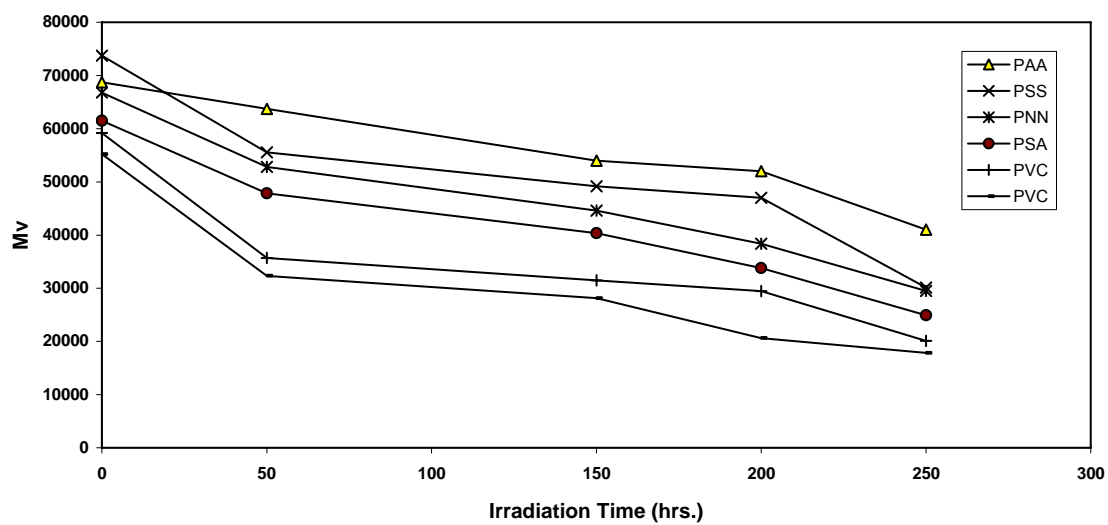


Figure 5. Changes in the viscosity average molecular weight (\overline{M}_v) during irradiation of PVC films (30 μm) (control) and modified PVC

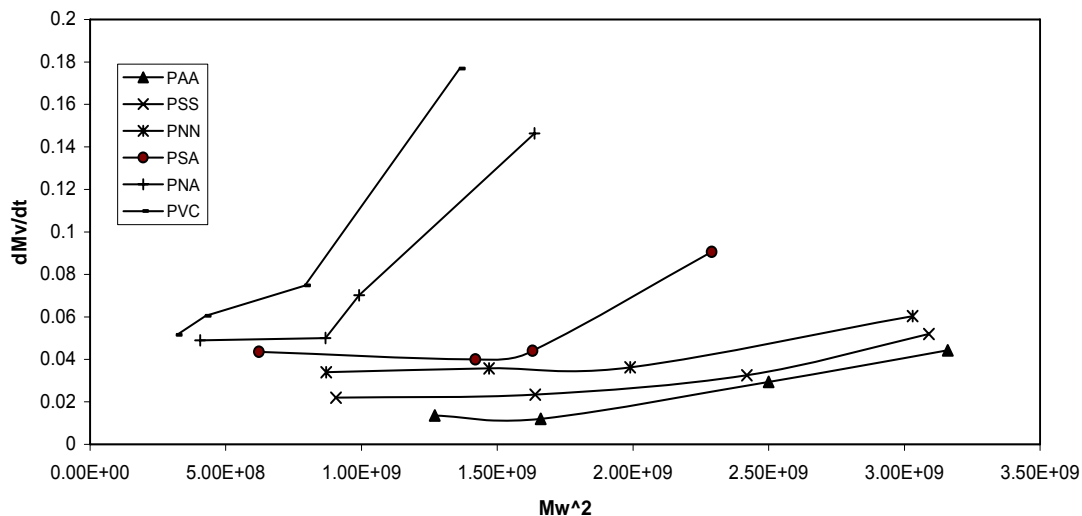


Figure 6. Rate variation of molecular weight change ($d\bar{M}_w/dt$) during irradiation of PVC films (30 μm) (control) and modified PVC films

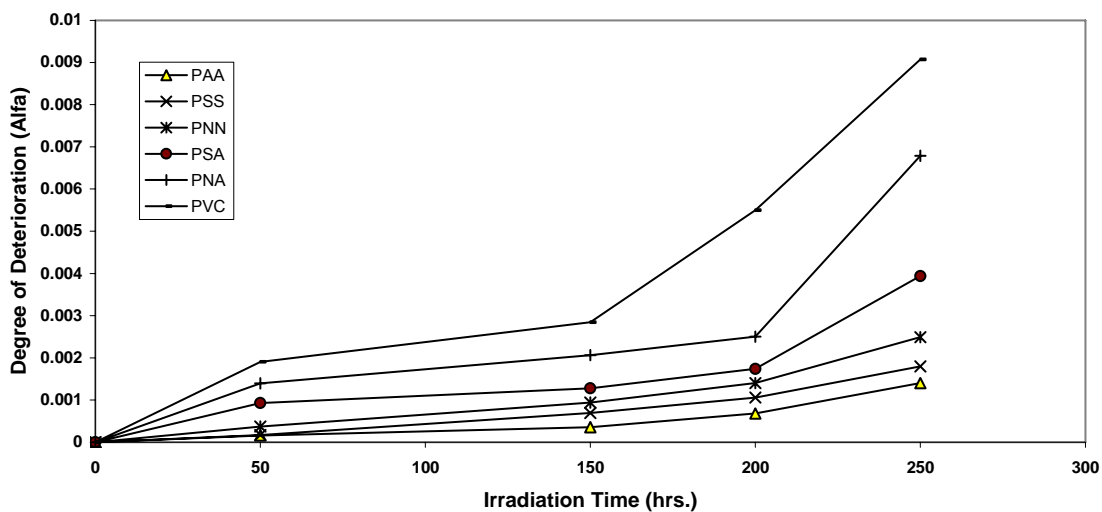


Figure 7. Changes in the degree of deterioration (α) during irradiation of PVC films (30 μm) (control) and modified PVC films

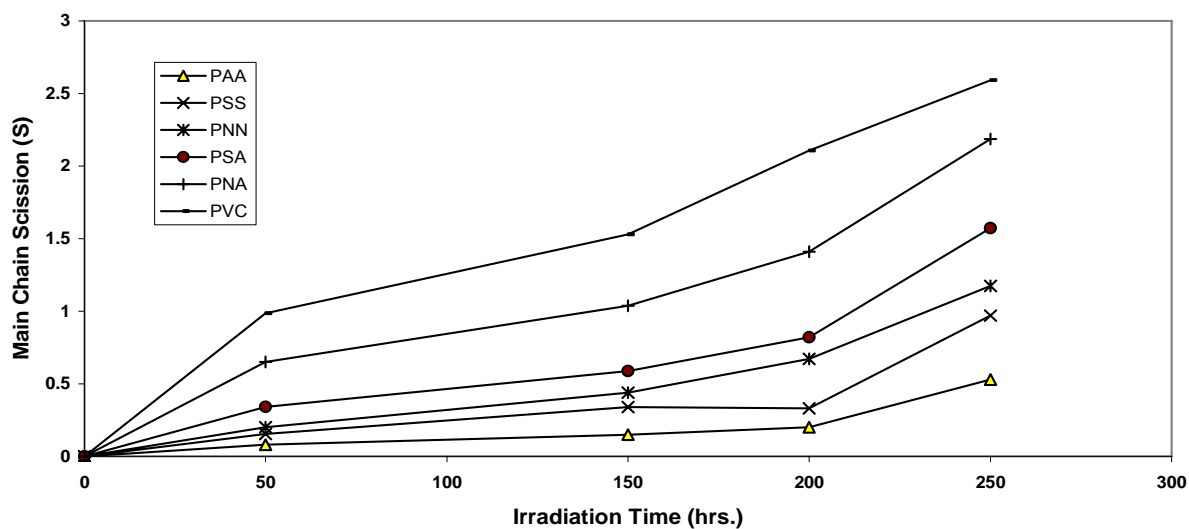


Figure 8. Changes in the number average chains during irradiation of PVC films (30 μ m) (control) and modified PVC films

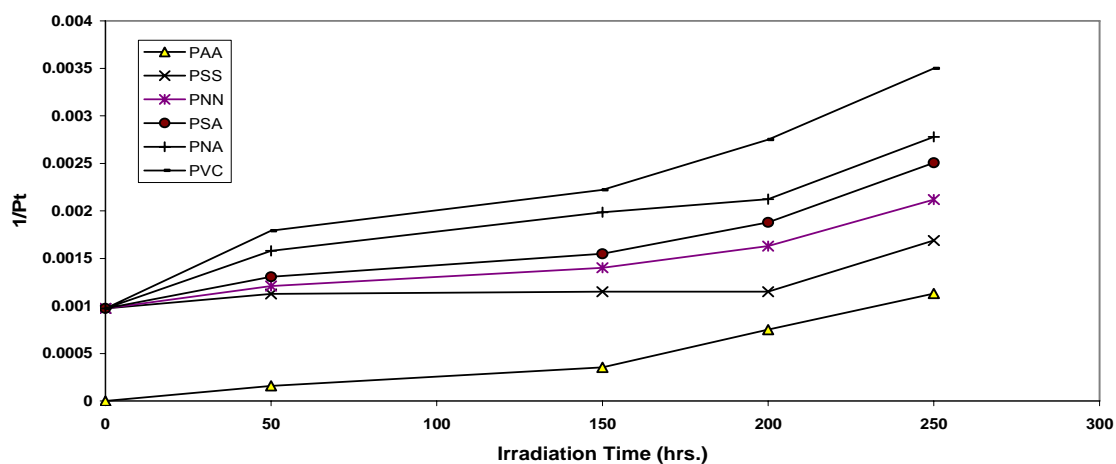
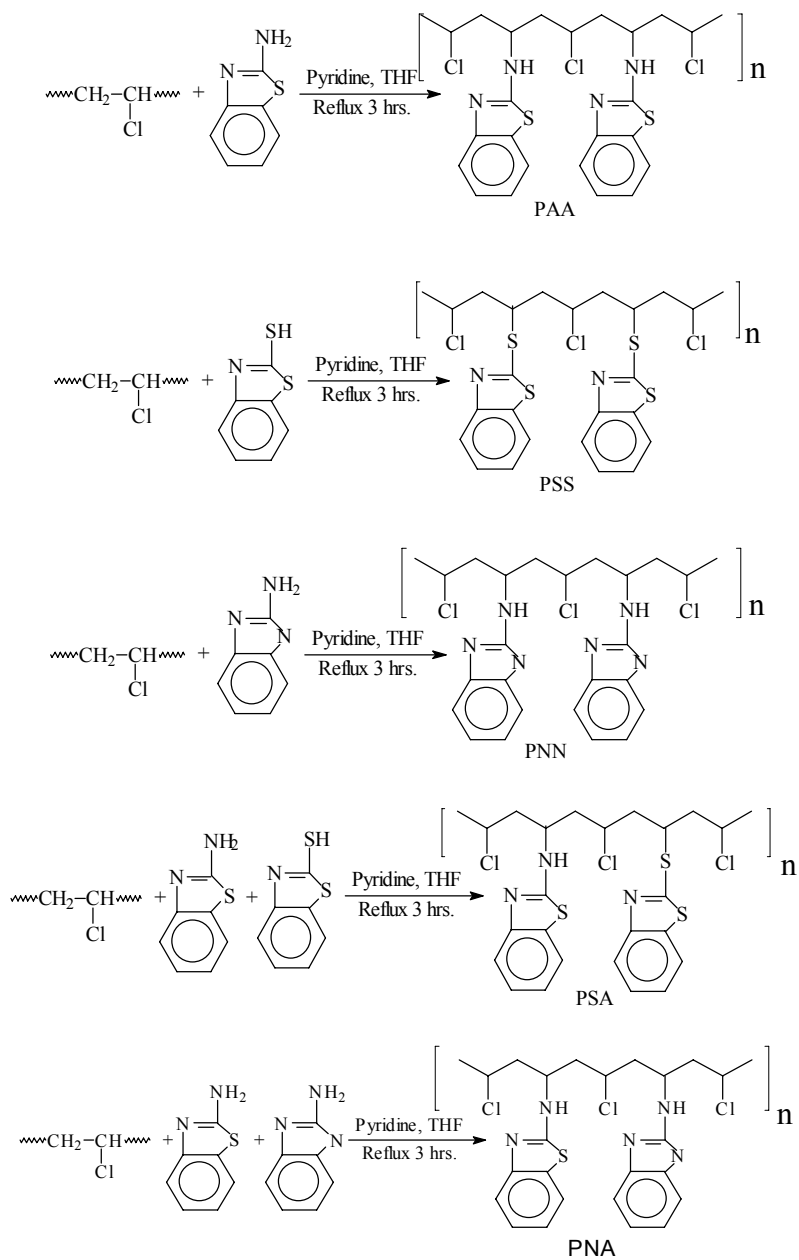
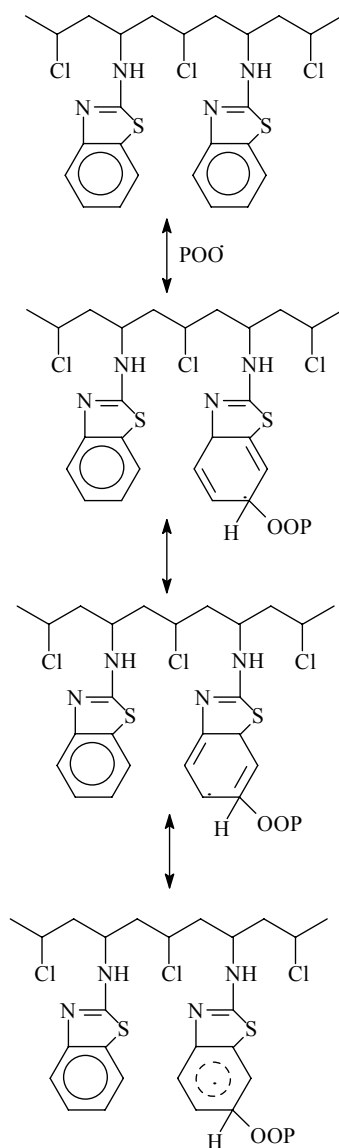


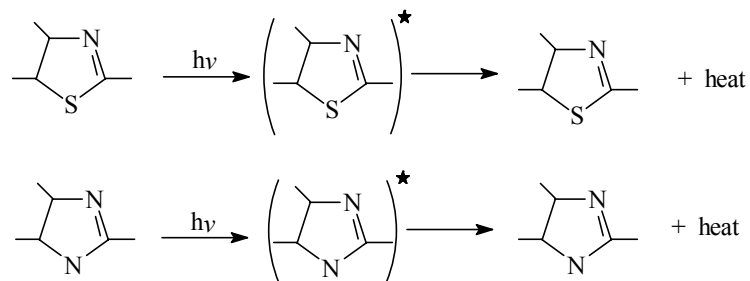
Figure 9. Changes in the reciprocal of number average of polymerization ($1/P_t$) during 313 nm during irradiation of PVC films (30 μ m) (control) and modified PVC films



Scheme 1. Chemical reaction of preparation of polymers



Scheme 2. Suggested mechanism of photostabilization of PAA as radical scavengers through energy transfer and forming unreactive charge transfer and stabilize through resonating structure.



Scheme 3. Suggested mechanism of photostabilization of modified polymers as UV absorber.