

Synthesis, Characterization and Biocidal Activities of Some Schiff Base Metal Complexes

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Neelakantan, *et al.*: Biological Screening of Schiff Base and Mixed Ligand Complexes

Some new mixed ligand complexes (1-5) of type ML'B (M(II)=Mn(II), Co(II), Ni(II), Cu(II) and Zn(II); HL'= *o*-vanillidene-2-aminobenzothiazole; B= 1,10-phenanthroline) and Schiff base metal complexes of types (ML₂') (6-10) and (M₂L'') (11-15) (HL''= *o*-vanillidene-2-amino-N-(2-pyridyl)-benzene sulfonamide) were synthesized and characterized by elemental analysis and spectral (IR, ¹H NMR and ¹³C NMR) studies. The free ligands and their metal complexes have been screened for their *in vitro* biological activities against bacteria, fungi and yeast. The metal complexes show more potent activities compared with Schiff base ligands.

Key words: Antibacterial activities, antifungal activities, mixed ligand complexes, *o*-vanillin, schiff base, spectral studies

Metal complexes of N and S chelating ligands have attracted considerable attention because of their interesting physicochemical properties and pronounced biological activities. The N and S atoms play a key role in the coordination of metals at the active sites of numerous metallo-bimolecules. *o*-vanillin is a natural aldehyde found in *Andropogon nardus*. It is used to treat bellyaches and also used in spicery^[1]. Schiff bases containing *o*-vanillin possesses antifungal, antibacterial properties^[2] and it acts as a weak inhibitor of tyrosinase, display both antimutagenic and comutagenic properties in *Escherichia coli*^[3]. Heterocycles containing thiazole ring is present in a number of pharmacologically and biologically active compounds. Compounds containing benzothiazole and sulphonamide derivatives were used as antifungal^[4,5], anti-inflammatory^[6], anti-HIV^[7], anticancer^[8], anticarbonic anhydrase^[9], diuretic, hypoglycaemic^[10], antithyroid^[11], antimalarial and in therapeutic fields. In view of the pronounced biological activities of these compounds, we report herein the synthesis and characterization of (i) mixed ligand complexes of Mn(II), Co(II), Ni(II), Cu(II) and Zn(II) derived from 1,10-phenanthroline and *o*-vanillidene-2-aminobenzothiazole (HL') and (ii) Schiff base metal complexes of *o*-vanillidene-2-

amino-N-(2-pyridyl)-benzene sulfonamide (HL''). The biological screening of free ligands and their complexes against different bacteria, fungi and yeast are reported.

MATERIALS AND METHODS

All the chemicals and solvents used were of AR grade (Merck, Mumbai, India) except *o*-vanillin (Fluka, Switzerland) and 2-amino-N-(2-pyridyl)-benzene sulfonamide (Sigma, USA). Melting points of all the compounds were determined in open glass capillaries and are uncorrected. The purity of the Schiff bases was ascertained by TLC on Silica gel-G plates and spots were visualized by using iodine vapours. The elemental analysis was performed using a Thermo Finnigan Flash EA 1112 CHNS analyzer at Central Electro Chemical Research Institute (CECRI), Karaikudi, India. The conductometric measurements of the complexes were carried out in DMSO solution using Systronics 611 conductivity bridge. Vibrational spectra were recorded using KBr pellets on FT-IR Shimadzu 8400S spectrophotometer, in the region 4000 – 400 cm⁻¹ range. The ¹H NMR and ¹³C NMR of Schiff base ligands and their diamagnetic zinc complexes in DMSO-d₆ were recorded on Perkin Elmer R-32 spectrometer using Tetramethylsilane as internal standard at IIT, Chennai, India.

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Biological evaluation:

The newly synthesized ligands and their metal complexes were screened *in vitro* for their antibacterial activity against bacteria: *Escherichia coli*, *Pseudomonas aeruginosa*, *Salmonella Typhi* and *Vibrio parahaemolyticus* by well diffusion method^[12] using agar nutrient. The antifungal activities were tested against fungus: *Aspergillus Niger*, *Penicillium*, *Trichoderma virida* and yeast: *Saccharomyces cerevisiae* by well diffusion method using potato dextrose agar as the medium. Ampicillin and nystatin are used as control for bacteria and fungi, respectively. The suspension of each microorganism was added to a sterile agar medium, then poured into sterile Petri plates and left to solidification. The well was dug in the agar media using sterile metallic borer in each plate. The test solution (3×10^{-3} M) was prepared by dissolving the compounds in DMSO and the well was filled with the test solution using micropipette. The plates were incubated for 24 h in the case of bacteria and 72 h for fungi at 35° . The extracts were subjected to further assay with a series on time basis (24, 48 and 72 h). During this

period, the test solution was diffused and affected the growth of the inoculated microorganisms. Activity was determined by measuring the diameter of the zone showing complete inhibition (mm). Growth of inhibition was compared with the control. The zone of inhibition is given as the average of three independent determinations.

Synthesis of mixed ligand complexes (1-5):

The Schiff base (HL') was synthesized^[13] by the condensation of 20 ml of *o*-vanillin (0.030 g/10 mmol) with 2-amino benzothiazole (0.030 g/10 mmol) in ethanol (1:1 molar ratio). Ten millilitres of Schiff base (0.028 g/10 mmol) and 1,10-phenanthroline (0.020 g/10 mmol) in ethanol was added drop wise to 10 ml of 10 mmol metal salts in hot ethanol (0.020g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$, 0.024g of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$, 0.025 g of $\text{Ni}(\text{CH}_3\text{COO})_2 \cdot 4\text{H}_2\text{O}$, 0.020 g of $\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$ and 0.022 g of $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$). The reaction mixture was refluxed for 3 h on a water bath and the volume of the solution was reduced to half of its original volume. The solid compound (fig. 1) obtained was filtered off, washed with water, diethyl ether and

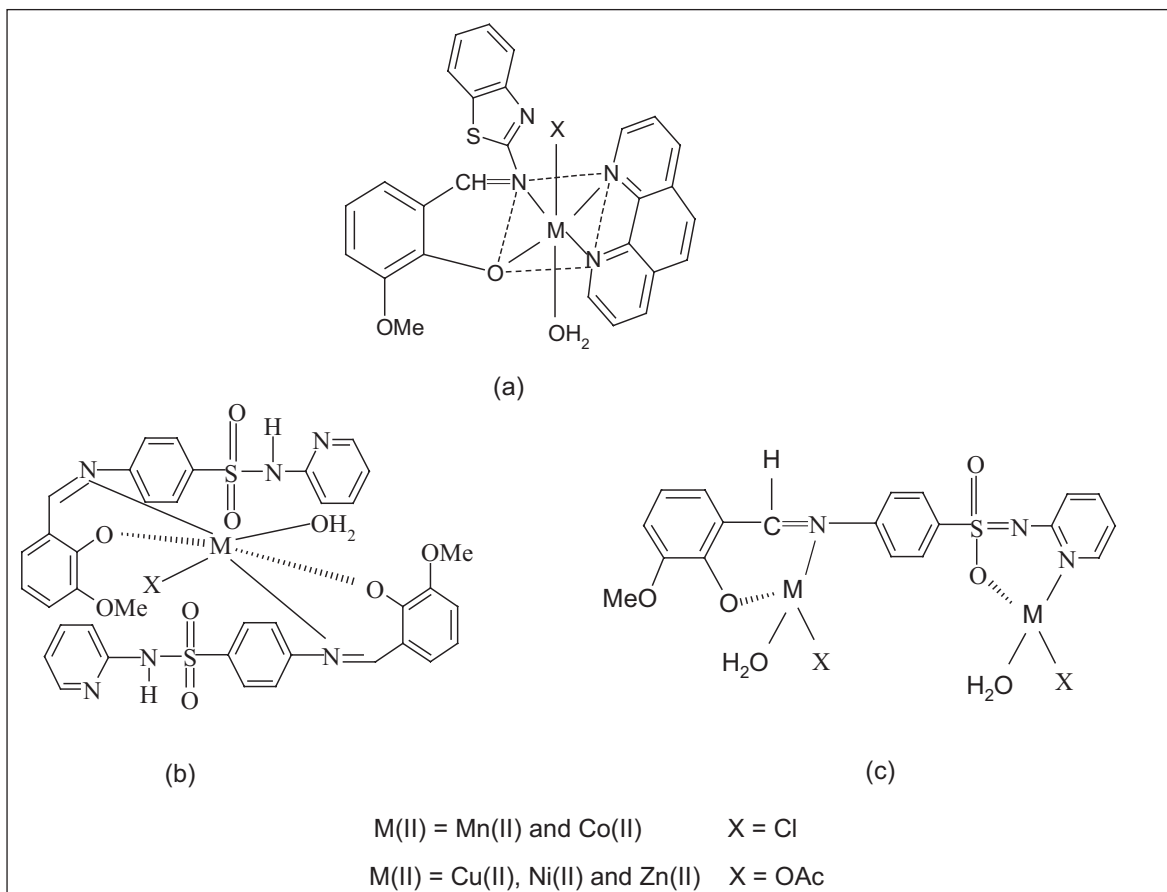


Fig. 1: Structure of the complexes.

(a) Mixed ligand complexes (1-5), (b) Schiff base complexes (6-10) and (c) Schiff base complexes (11-15).

dried in vacuum over CaCl_2 . The color, elemental analysis data, molar conductivity and melting point of Schiff base (HL') and its mixed ligand complexes (1-5) are given in Table 1.

Compound 1, yield: 67%, mp: 205°, IR (KBr) cm^{-1} : 1639 (C=N, azomethine), 1557 (C=N, thiazole ring), 1217 (C-O, phenolic), 746 (C-S-C, thiazole ring), 569 (Mn-N), 424 (Mn-O), 3396 and 856 (H_2O molecule). Compound 2, yield: 62%, mp: 145°, IR (KBr) cm^{-1} : 1618 (C=N, azomethine), 1558 (C=N, thiazole ring), 1215 (C-O, phenolic), 747 (C-S-C, thiazole ring), 572 (Co-N), 430 (Co-O), 3387, 848 (H_2O molecule). Compound 3, yield: 69%, mp: 169°, IR (KBr) cm^{-1} : 1624 (C=N, azomethine), 1558 (C=N, thiazole ring), 1207 (C-O, phenolic), 747 (C-S-C, thiazole ring), 567 (Ni-N), 420 (Ni-O), 3387, 850 (H_2O molecule). Compound 4, yield: 67%, mp: 199°, IR (KBr) cm^{-1} : 1606 (C=N, azomethine), 1560 (C=N, thiazole ring), 1217 (C-O, phenolic), 748 (C-S-C, thiazole ring), 540 (Cu-N), 422 (Cu-O), 3380 and 852 (H_2O molecule). Compound 5, yield: 65%, mp: 130°, IR (KBr) cm^{-1} : 1641 (C=N, azomethine), 1561 (C=N, thiazole ring), 1213 (C-O, phenolic), 748 (C-S-C, thiazole ring), 563 (Zn-N), 437 (Zn-O), 3321, 852 (H_2O molecule), ^1H NMR (TMS, DMSO-d_6) δ ppm: 7.67 (s, 1H, HC=N, azomethine), 6.80–7.40 (d, 4H, Ar-H), 6.90 (s, 1H, HC=N, thiazole ring), 3.70 (s, 3H, OCH_3), 1.80 (s, 3H, OOCCH_3).

Synthesis of Schiff base:

The Schiff base (HL'') was prepared by refluxing a mixture of equimolar quantities of *o*-vanillin (10 ml, 0.300 g/10 mmol) and 2-amino-N-(2-pyridyl)-benzene sulfonamide (10 ml, 0.490 g/10 mmol, Scheme 1) in hot ethanol. After 8 h of refluxing, the reaction mixture was kept at room temperature overnight and the orange colored product was filtered, washed with distilled water, diethyl ether and recrystallized from the same solvent. Yield: 85%, mp: 198°, IR (KBr) cm^{-1} : 1631 (C=N, azomethine), 1269 (C-O, phenolic), 1361, 1138 (SO_2), 3246 (N-H, sulfonamide group), 1025 ($-\text{HC}=\text{N}$, pyridine ring nitrogen), ^1H NMR (TMS, DMSO-d_6) δ ppm: 9.92 (s, 1H, HC=N, azomethine), 12.82 (s, -OH), 7.60–8.00 (d, 4H, Ar - H), 8.40–8.60 (s, pyridine $-\text{HC}=\text{N}$), 11.04 (s, NH), 3.48 (s, 3H, OCH_3), 2.30 (s, 3H, CH_3), 5.72 (s, 2H, SO_2NH group), ^{13}C NMR (TMS, DMSO-d_6) δ ppm: 164.66 (HC=N, azomethine), 155.29 ($-\text{C}-\text{OH}$, phenolic), 151.44 (HC=N, pyridine ring), 142.04–140.36 (C=C, pyridine ring), 148.51 (C-O-C), 121.69–117.96 (Ar, C=C), 56.24 (OCH_3).

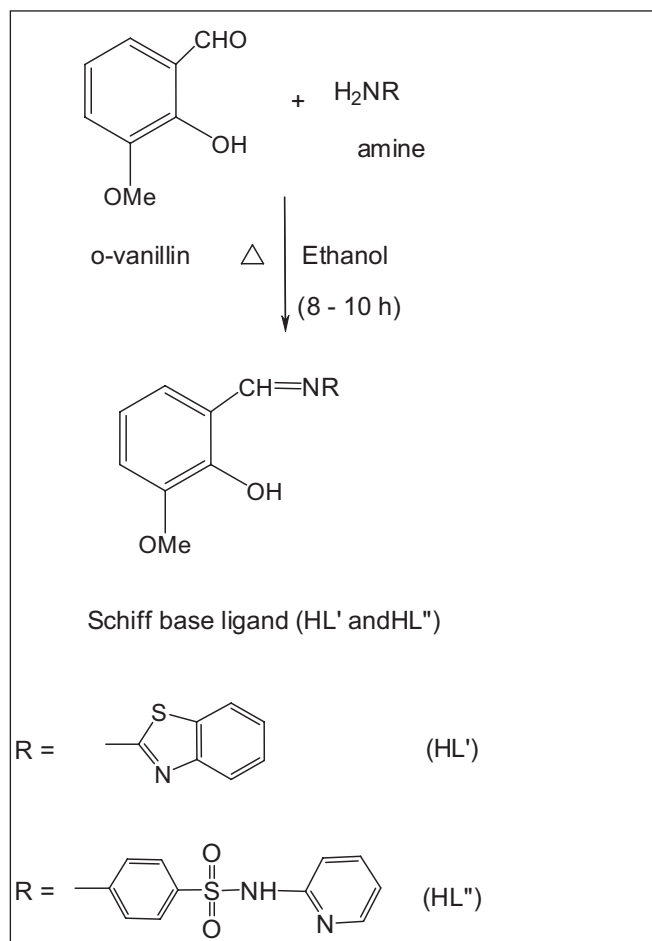
Synthesis of metal complexes (6-15):

Ten millilitres of metal salts (10 mmol, 0.02 g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$, 0.024 g of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$, 0.025 g of $\text{Ni}(\text{CH}_3\text{COO})_2 \cdot 4\text{H}_2\text{O}$, 0.02 g of $\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$ and 0.022 g of $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$) in hot ethanol was added to 10 ml of the Schiff base ligand

TABLE 1: PHYSIO-CHEMICAL PROPERTIES OF SCHIFF BASE AND METAL COMPLEXES

Compound	Color	Mol. formula ^a	Mol. Wt.	Yield (%)	Λ_m ($\Omega^{-1} \text{cm}^2 \text{mol}^{-1}$)	Mp (°)
HL'	Reddish yellow	$\text{C}_{15}\text{H}_{12}\text{N}_2\text{O}_2\text{S}$	284.00	65	6.78	180
HL''	Orange	$\text{C}_{19}\text{H}_{17}\text{N}_3\text{O}_4\text{S}$	383.42	94	4.23	198
1	Pale yellow	$\text{MnC}_{27}\text{H}_{21}\text{O}_3\text{N}_4\text{SCl}$	571.74	67	11.38	205
2	Green	$\text{CoC}_{27}\text{H}_{21}\text{O}_3\text{N}_4\text{SCl}$	575.73	62	17.39	145
3	Light brown	$\text{NiC}_{29}\text{H}_{24}\text{O}_5\text{N}_4\text{S}$	599.04	69	14.22	169
4	Dark brown	$\text{CuC}_{29}\text{H}_{24}\text{O}_5\text{N}_4\text{S}$	603.90	67	9.42	202
5	Yellow	$\text{ZnC}_{29}\text{H}_{24}\text{O}_5\text{N}_4\text{S}$	605.76	65	6.78	130
6	Brown	$\text{MnC}_{38}\text{H}_{34}\text{O}_9\text{N}_6\text{S}_2\text{Cl}$	896.43	54	5.50	210
7	Light Brown	$\text{CoC}_{38}\text{H}_{34}\text{O}_9\text{N}_6\text{S}_2\text{Cl}$	899.88	64	8.10	>300
8	Light Green	$\text{NiC}_{40}\text{H}_{37}\text{O}_{11}\text{N}_6\text{S}_2$	900.22	59	3.00	285
9	Black	$\text{CuC}_{40}\text{H}_{37}\text{O}_{11}\text{N}_6\text{S}_2$	904.93	63	5.13	290
10	Dark yellow	$\text{ZnC}_{40}\text{H}_{37}\text{O}_{11}\text{N}_6\text{S}_2$	906.86	57	3.40	220
11	Light brown	$\text{Mn}_2\text{C}_{19}\text{H}_{19}\text{O}_6\text{N}_3\text{SCl}_2$	718.27	58	9.80	225
12	Brown	$\text{Co}_2\text{C}_{19}\text{H}_{19}\text{O}_6\text{N}_3\text{SCl}_2$	725.17	51	13.90	>300
13	Dark Green	$\text{Ni}_2\text{C}_{23}\text{H}_{25}\text{O}_{10}\text{N}_3\text{S}$	725.85	68	12.90	293
14	Black	$\text{Cu}_2\text{C}_{23}\text{H}_{25}\text{O}_{10}\text{N}_3\text{S}$	663.28	83	10.80	320
15	Light yellow	$\text{Zn}_2\text{C}_{23}\text{H}_{25}\text{O}_{10}\text{N}_3\text{S}$	739.13	56	12.40	254

^aC, H and N are within the limit of $\pm 0.3\%$ and S $\pm 0.4\%$; HL' - *o*-vanillidene-2-aminobenzothiazole; HL'' - *o*-vanillidene-2-amino-N-(2-pyridyl)-benzene sulfonamide



Scheme 1: Synthesis of Schiff base (HL' and HL'') ligands

(0.038 g/10 mmol) in ethanol (1:1). 10 ml of metal salts (20 mmol, 0.04 g of MnCl₂·4H₂O, 0.048 g of CoCl₂·6H₂O, 0.05 g of Ni(CH₃COO)₂·4H₂O, 0.04 g of Cu(CH₃COO)₂·H₂O and 0.044 g of Zn(CH₃COO)₂·2H₂O) in hot ethanol was added to 10 ml of the Schiff base ligand (0.038 g/10 mmol) in ethanol (2:1). The mixture was refluxed for 8 h on water bath. The volume of the solution was reduced to one-third of its original volume and left overnight. The precipitate was filtered off, washed with 50% (v/v) ethanol-water mixture, diethyl ether and dried over fused CaCl₂ in a vacuum desiccator. The elemental analysis, molar conductivity value, color and melting point of Schiff base (HL'') and its metal(II) complexes (6-15) are given in Table 1.

Compound 6, yield: 54%, mp: 210°, IR (KBr) cm⁻¹: 1599 (C=N, azomethine), 1293 (C-O, phenolic), 1362, 1130 (SO₂), 3248 (N-H, sulfonamide group), 1026 (-HC=N, pyridine ring nitrogen), 594 (Mn-N), 521 (Mn-O), 3398, 976, 774 (-OH, H₂O molecule). Compound 7, yield: 64%, mp: >300°, IR (KBr) cm⁻¹:

1602 (C=N, azomethine), 1272 (C-O, phenolic), 1364, 1132 (SO₂), 3244 (N-H, sulfonamide group), 1030 (C=N, pyridine ring nitrogen), 604 (Co-N), 520 (Co-O), 3405, 976, 759 (-OH, H₂O molecule). Compound 8, yield: 59%, mp: 285°, IR (KBr) cm⁻¹: 1610 (C=N, azomethine), 1286 (C-O, phenolic), 1364, 1135 (SO₂), 3246 (N-H, sulfonamide group), 1029 (C=N, pyridine ring nitrogen), 588 (Ni-N), 503 (Ni-O), 3398, 968, 767 (-OH, H₂O molecule), 1616, 1402 (-COO, acetate group). Compound 9, yield: 63%, mp: 290°, IR (KBr) cm⁻¹: 1616 (C=N, azomethine), 1291 (C-O, phenolic), 1362, 1138 (SO₂), 3248 (N-H, sulfonamide group), 1027 (C=N, pyridine ring nitrogen), 569 (Cu-N), 513 (Cu-O), 3415, 980, 775 (-OH, H₂O molecule), 1626, 1412 (-COO, acetate group). Compound 10, yield: 57%, mp: 220°, IR (KBr) cm⁻¹: 1598 (C=N, azomethine), 1278 (C-O, phenolic), 1363, 1136 (SO₂), 3247 (N-H, sulfonamide group), 1026 (C=N, pyridine ring nitrogen), 596 (Zn-N), 507 (Zn-O), 3382, 962, 769 (-OH, H₂O molecule), 1631, 1408 (-COO, acetate group). ¹H NMR (TMS, DMSO-d₆) δ ppm: 10.12 (s, 1H, HC=N, azomethine), 7.65-7.97 (d, 4H, Ar-H), 8.38-8.61 (s, pyridine, -HC=N), 11.00 (s, NH), 3.46 (s, 3H, OCH₃), 2.31 (s, 3H, CH₃), 5.70 (s, 2H, SO₂NH group), 4.73 (s, 2H, H₂O molecule), ¹³C NMR (TMS, DMSO-d₆) δ ppm: 169.42 (HC=N, azomethine), 150.94 (HC=N, pyridine ring), 141.98-140.52 (C=C pyridine ring C), 147.91 (C-O-C), 122.76-118.72 (Ar, C=C), 55.82 (OCH₃).

Compound 11, yield: 58%, mp: 225°, IR (KBr) cm⁻¹: 1615 (C=N, azomethine), 1279 (C-O, phenolic), 1006 (C=N, pyridine ring nitrogen), 589 (Mn-N), 518 (Mn-O), 3403, 987, 779 (-OH, H₂O molecule). Compound 12, yield: 51%, mp: >300°, IR (KBr) cm⁻¹: 1599 (C=N, azomethine), 1286 (C-O, phenolic), 1013 (C=N, pyridine ring nitrogen), 586 (Co-N), 516 (Co-O), 3394, 997, 765 (-OH, H₂O molecule). Compound 13, yield: 68%, mp: 293°, IR (KBr) cm⁻¹: 1605 (C=N, azomethine), 1298 (C-O, phenolic), 1350, 1009, 669 (C=N, pyridine ring nitrogen), 605 (Ni-N), 513 (Ni-O), 3379, 986, 776 (-OH, H₂O molecule), 1631, 1405 (-COO, acetate group). Compound 14, yield: 83%, mp: 320°, IR (KBr) cm⁻¹: 1620 (C=N, azomethine), 1287 (C-O, phenolic), 1012 (C=N, pyridine ring nitrogen), 605 (Cu-N), 517 (Cu-O), 3415, 980, 775 (-OH, H₂O molecule), 1632, 1408 (-COO, acetate group). Compound 15, yield: 56%, mp: 254°, IR (KBr) cm⁻¹: 1612 (C=N, azomethine), 1304 (C-O, phenolic), 1006 (C=N, pyridine ring nitrogen),

579 (Zn–N), 510 (Zn–O), 3410, 989, 787 (–OH, H₂O molecule), 1623, 1408 (–COO, acetate group). ¹H NMR (TMS, DMSO-d₆) δ ppm: 10.03 (s, 1H, HC=N, azomethine), 7.64–7.99 (d, 4H, Ar–H), 8.23–8.54 (s, pyridine, –HC=N), 3.40 (s, 3H, OCH₃), 2.28 (s, 3H,

CH₃), 11.86 (s, 2H, SO₂NH group), 4.67 (s, 2H, H₂O molecule), ¹³C NMR (TMS, DMSO-d₆) δ ppm: 172.26 (–HC=N, azomethine), 164.49 (HC=N, pyridine ring), 140.19 – 138.25 (C=C pyridine ring C), 148.17 (C–O–C), 124.27 – 119.58 (Ar, C=C), 56.15 (OCH₃).

TABLE 2: BACTERIAL ACTIVITIES OF THE SCHIFF BASE AND METAL COMPLEXES

Compound	Diameter of inhibition zone in mm for different microbial species											
	<i>Salmonella typhi</i>			<i>Pseudomonas aeruginosa</i>			<i>Escherichia coli</i>			<i>Vibrio parahaemolyticus</i>		
	24 h	48 h	72 h	24 h	48 h	72 h	24 h	48 h	72 h	24 h	48 h	72 h
HL'	7	8	8	8	9	9	10	11	12	8	9	9
HL''	6	7	7	7	8	8	-	-	-	7	8	8
1	9	10	10	9	11	12	9	11	11	9	10	13
2	-	-	-	14	15	16	-	-	-	8	8	11
3	9	9	11	22	23	23	20	22	22	8	9	9
4	-	9	10	8	10	10	-	-	-	10	10	10
5	9	10	10	-	-	-	12	12	14	10	11	12
6	7	7	8	-	-	-	-	-	-	6	7	9
7	7	7	8	-	-	-	-	-	-	6	7	9
8	-	-	-	-	-	-	-	-	-	5	10	10
9	9	8	7	-	-	-	-	-	-	7	8	9
10	-	-	-	6	7	7	-	-	-	12	13	14
11	6	8	8	-	-	-	6	7	8	8	9	9
12	-	-	-	8	9	9	9	9	10	-	-	-
13	15	15	16	6	7	7	10	12	13	10	10	11
14	-	-	-	7	8	9	-	-	-	8	8	9
15	14	15	17	11	12	12	11	14	14	9	11	12
Control	11	14	16	32	34	34	16	17	18	31	33	34

Each experiment was done in triplicate; [-, less active]; well diffusion method

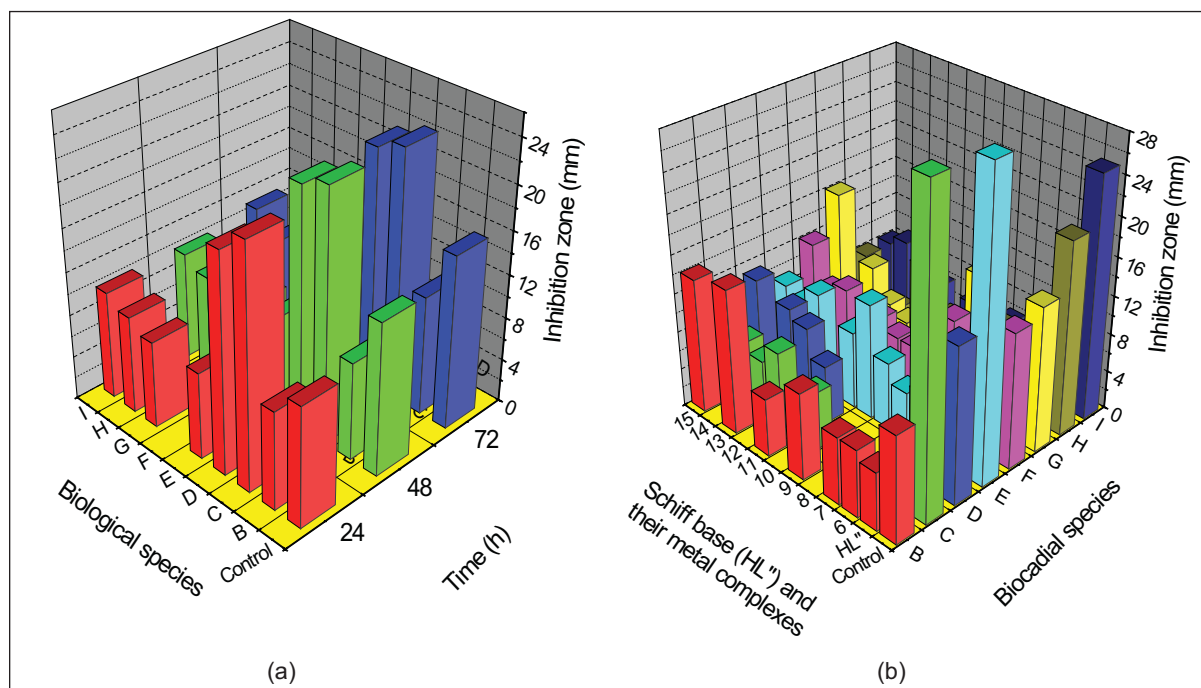


Fig. 2: Biological activities of (a) mixed ligand complex (b) Schiff base complexes.
 (a) Mixed ligand complex (3) at 24, 48 and 72 h; (b) Schiff base complexes (HL'', 6-10 and 11-15) at 24 h by well diffusion method (zone formation in mm); B- *Salmonella typhi*, C- *Pseudomonas aeruginosa*, D- *Escherichia coli*, E- *Vibrio parahaemolyticus*, F- *Aspergillus niger*, G- *Penicillium*, H- *Trichoderma virida* and I- *Saccharomyces species*.

TABLE 3: FUNGAL ACTIVITIES OF THE SCHIFF BASE AND METAL COMPLEXES

Compound	Diameter of inhibition zone in mm for different microbial species											
	<i>Aspergillus niger</i>			<i>Penicillium</i>			<i>Trichoderma virida</i>			<i>Saccharomyces cerevisiae</i>		
	24 h	48 h	72 h	24 h	48 h	72 h	24 h	48 h	72 h	24 h	48 h	72 h
HL'	5	6	6	5	6	6	8	8	9	8	9	9
HL''	6	6	6	-	6	8	6	7	7	4	5	6
1	-	-	-	9	9	10	13	14	14	12	12	13
2	9	9	10	9	10	11	-	-	-	8	8	9
3	-	9	9	8	8	9	9	9	9	10	10	11
4	10	11	11	9	10	10	12	14	14	9	10	10
5	-	10	10	-	9	10	12	12	13	10	11	12
6	10	12	13	10	11	12	8	9	9	8	11	12
7	12	14	16	15	16	16	8	8	8	6	7	8
8	10	12	12	-	-	-	4	4	8	6	8	10
9	7	9	13	-	-	-	4	6	6	-	-	-
10	7	8	16	13	15	16	6	10	10	6	8	8
11	8	8	10	6	8	8	4	4	4	-	-	-
12	-	-	-	7	8	8	-	-	-	6	6	6
13	9	12	12	10	10	11	4	4	4	-	-	-
14	8	8	9	-	-	-	-	-	-	9	9	10
15	12	13	14	16	18	19	8	12	12	8	8	8
Control	14	15	16	15	18	18	20	22	23	25	28	28

Each experiment was done in triplicate; [-, less active]; well diffusion method

RESULTS AND DISCUSSION

The zone of inhibition of the complexes against the growth of bacteria and fungi were given in Table 2 and Table 3, respectively. A representative graph is given in fig 2. From the table it is evident that the complexes (1-5) and (11-15) have higher antibacterial activity than the complexes (6-10). The complexes (1-5) and (6-10) shows higher antifungal activity than the complexes (11-15). In general, the synthesized metal complexes have higher biological activities compared to the free ligands. The increased inhibition activity of the metal complexes can be explained on the basis of Tweedy's chelation theory^[14]. In metal complexes, on chelation the polarity of the metal ion will be reduced to a greater extent due to the overlap of the ligand orbital and partial sharing of the positive charge of the metal ion with donor groups. Further, it increases the delocalization of π - electrons over the whole chelate ring. The large ring size of 1,10-phenanthroline moiety makes the complexes more lipophilic^[15]. This increased lipophilicity enhances the penetration of the metal complexes into lipid membranes and block the metal binding sites in the enzymes^[16]. Metal complexes also disturb the respiration process of the cell and thus block the synthesis of proteins, which restricts further growth of the organisms. This enhancement in inhibiting the growth of bacteria and fungi can also be explained

on the basis of their structure. The azomethine linkage and hetero aromatic moiety in the synthesized complexes exhibit extensive biological activities^[17,18] due to increased liposolubility of the molecules in crossing cell membrane of the microorganism. The presence of electron donor group (-OCH₃) in the complexes also plays a role in enhancing the inhibition activity. The antibacterial activity is found to be in the order; Control>(1-5)>(11-15)>(6-10)>HL'>HL''. The antifungal activity is found to be in the order; Control>(6-10)>(1-5)>(11-15)>HL'>HL''.

ACKNOWLEDGEMENTS

The authors wish to thank the Management, Director and Principal of National Engineering College, Kovilpatti for their keen interest and constant encouragement throughout this investigation.

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Accepted 20 February 2010

Revised 1 December 2009

Received 4 August 2009

Indian J. Pharm. Sci., 2010, 72 (2): 216-222