

# Synthesis, characterization, spectroscopic and crystallographic investigation of Cobalt(III) schiff base complex with two perpendicular diamine coumarin ligands

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## ABSTRACT

New transition metal complex of Cobalt(III) of the ligand (*E*)-3-(1-(2-aminoethylimino)ethyl)-4-hydroxy-2H-chromen-2-one, derived from condensation of ethylene diamine with 3-acetyl-4-hydroxy-chromene-2-one have been synthesized by reaction of cobalt(III) salt and the ligand, in amounts equal to metal-ligand molar ratio of 1:2. Both the Schiff base and the complex of Co(III) were characterized by IR, UV-vis, <sup>1</sup>H NMR- and <sup>13</sup>C NMR-spectroscopy techniques. Single crystal X-ray diffraction investigation, at low temperature T = 120 K, shows that the cobalt complex is triclinic P-1, a = 10.426(5) Å, b = 11.3234(2) Å, c = 15.729(5) Å, α(°) = 70.102(4), β(°) = 86.049(4), γ(°) = 82.497(4), Z = 2, and its structure consists of isolated [Co(III)(C<sub>13</sub>H<sub>13</sub>N<sub>2</sub>O<sub>3</sub>)<sub>2</sub>]<sup>+</sup> complex cations with distorted octahedral geometry, ClO<sub>4</sub><sup>-</sup> counter anions, acetone solvent and water molecules. The crystal cohesion is stabilized by hydrogen bonds between ligands and water molecules, and ionic interactions between complex cations and counter anions.

**Keywords:** Hydroxycoumarin Derivative; Co<sup>III</sup> Complex; <sup>1</sup>H NMR; UV-Visible; Single Crystal X-Ray Diffraction

## 1. INTRODUCTION

Coumarins and their derivatives [1,2] were found to exhibit a variety of biological and pharmacological activities and have attracted considerable interest because of their potential beneficial effects on human health [3,4].

A series of coumarin derivatives were recently synthesized and studied for their structure and anticoagulant activity and cytotoxic [5-10]. A number of coumarins have been investigated for complexing ability [11,12]. The mechanisms of cytotoxic action of a large number of different coordination compounds have been discussed in relation to the development of new antitumor agents.

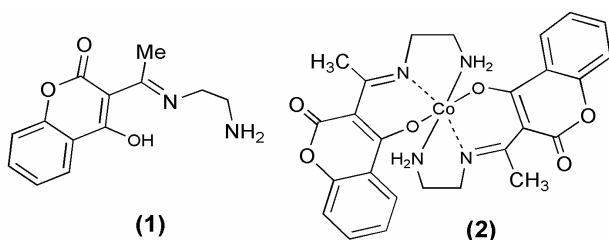
A wide range of medicinal applications of metal complexes of coumarins was investigated. It was found that in some cases, the metal complexes obtained revealed a higher activity than their biological ligands [9-11].

In the past decade, Schiff bases have received much attention, mainly because of their extensive applications in the synthesis of coordination compounds and catalysis [13-17]. This is due to the fact that Schiff bases offer opportunities for inducing substrate chirality, tuning factors of metal centered electronic and improve the solubility and stability of homogeneous or heterogeneous catalysts, *i.e.* [18,19]. They are also increasingly important as biochemical, analytical and antimicrobial reagents [20]. Further, significant advances have also been achieved in the field of straightforward preparation of new chiral ligands, such as chiral diamines and complexes, used for asymmetric catalysis [21]. For example, complexes based on chiral diamines have been employed for chiral hydrogenations and transfer hydrogenations [22]. Both homogeneous and heterogeneous diamine complexes have proved highly effective in enantioselective catalysis [23,24].

In continuation of our previous research [25-27] on synthesis and characterization of new asymmetric Schiff base complex, we describe here the synthesis and characterization of the tridentate Schiff base (*E*)-3-(1-(2-aminoethylimino)ethyl)-4-hydroxy-2H-chromen-2-one (1),

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then its complexation reaction with cobalt(III) to obtain the complex numbered as (2).



## 2. EXPERIMENTAL

### 2.1. Measurements

Proton nuclear magnetic resonance ( $^1\text{H NMR}$ ) spectra were run on a Bruker AV-300 NMR spectrometer and chemical shifts expressed as  $\delta$  (parts per million) values with TMS as internal standard. Multiplicities of proton resonance are designated as singlet (s), doublet (d), triplet (t), quartet (q), and multiplet (m). IR spectra were recorded in KBr on a Bruker FT-IR spectrophotometer. The electronic spectra in the 200 - 900 nm range were obtained on a Varian spectrophotometer 50 scan, using solutions with the concentration of  $10^{-5}$  M in DMSO. Single crystal was characterized by an Oxford Diffraction Gemini R Ultra CCD diffractometer. All the chemicals are commercially available and they were used without further purification. All the solvents were dried using standard methods before use.

### 2.2. Synthesis of the Schiff Base (1)

The reaction of 3-acetyl-4-hydroxy-2H-chromen-2-one with the ethylene diamine in the presence of methanol provided a high yield, as detailed below.

A solution of (0.6 mL, 5 mmol) ethylene diamine in methanol was added drop wise stirring to a methanol solution of (2 g, 5 mmol) 3-acetyl-4-hydroxy-chromene-2-one, already prepared by the method cited in [28], and then the mixture was refluxed for 1 h. The precipitated Schiff base was filtered and recrystallized from DMSO. Yield: 1.7 g (85%).

IR (selected,  $\text{cm}^{-1}$ ): 3097 [ $\nu(\text{C-H arom})$ ], 3207 [ $\nu(\text{NH}_2)$ ], 1637 [ $\nu(\text{C=O})$ ], 1618 [ $\nu(\text{C=N})$ ], 1585 [ $\nu(\text{C=C})$ ], 1279 [ $\nu(\text{C-O})$ ].

UV-vis  $\lambda_{\text{max}}/\text{nm}$  (DMSO): 235, 250, 320.

$^1\text{H NMR}$  (300 MHz, DMSO):  $\delta$  13.76 (s, H, OH), 7.23 - 8.27 (m, 4H, H arom), 4.00 (s, 2H,  $\text{NH}_2$ ), 2.69 (d, 2H,  $\text{CH}_2$ ), 2.48 (t, 2H,  $\text{CH}_2$ ), 2.07 (s, 3H,  $\text{CH}_3$ ).

$^{13}\text{C NMR}$  (75 MHz, DMSO):  $\delta$  179.63 C(21), 177.20 C(10), 161.80 C(8), 152.99 C(6), 134.12 C(2), 125.63 C(4), 123.68 C(3), 120.04 C(1), 116.20 C(9), 96.47 C(5), 42.81 C(25), 34.40 C(24), 18.31 C(52).

### 2.3. Synthesis of the Complex

#### $[\text{Co}(\text{C}_{13}\text{N}_2\text{O}_3\text{H}_{13})_2]\text{ClO}_4, 2(\text{C}_3\text{H}_6\text{O}), \text{H}_2\text{O} (2)$

In the synthesis of (2), one equivalent of ethylene diamine condenses with 4-hydroxy-2H-chromen-2-one to give 1:1 Schiff base (1) that is sequestered as,  $\text{Co}^{\text{III}}$  complex (2) (Figure 1).

A solution of (0.34 g, 1 mmol) of  $\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 4\text{H}_2\text{O}$  dissolved in 50 mL methanol was added to the freshly prepared Schiff base (0.41 g, 2 mmol). Upon addition of the metal, the solution immediately turned brown. An aqueous solution of sodium perchlorate to this solution was added and then the color turned red. The mixture was stirred for 3 h and the solid was collected by gravity filtration and dried under vacuum. Yield: 70%.

IR (selected,  $\text{cm}^{-1}$ ): 3414 [ $\nu(\text{OH})$ ], 1602 [ $\nu(\text{C=N})$ ], 1418 [ $\nu(\text{C-O})$ ], 1084.19 [ $\nu(\text{Cl-O})$  ( $\text{ClO}_4^-$ )], 523 [ $\nu(\text{M-N})$ ], 418 [ $\nu(\text{M-O})$ ].

UV-vis  $\lambda_{\text{max}}/\text{nm}$  (DMSO): 552, 486, 346, 261.

$^1\text{H NMR}$  (300 MHz, DMSO):  $\delta$  7.19 - 7.8 (m, 4H, H-arom.), 3.99 (t, 2H,  $\text{CH}_2$ ), 2.90 (s, 3H,  $\text{CH}_3$ ), 2.49 (d, H, NH), 2.48 (d, 2H,  $\text{CH}_2$ ).

$^{13}\text{C NMR}$  (75 MHz, DMSO):  $\delta$  172.56 (C10), 161.85 (C21), 152.8 (C8), 133.42 (C4), 125.92 (C2), 123.63 (C4), 119.38 (C1), 115.56 (C5), 100.83 (C9), 55.54 (C25), 42.94 (C24), 22.35 (C52), 119.38 (C3).

### 2.4. X-Ray Data Collection and Structure Solution and Refinement

Single crystals, suitable for X-ray data collection, were obtained by slow evaporation of the solvent from a solution of (2) in  $\text{CH}_3\text{COCH}_3$ . The crystals were filtered and washed with acetone. The diffraction data were collected with an Oxford Diffraction Gemini R Ultra CCD. Semi-empirical absorption corrections from equivalents (multi-scan) were applied, using graphite-monochromated  $\text{Cu K}\alpha$

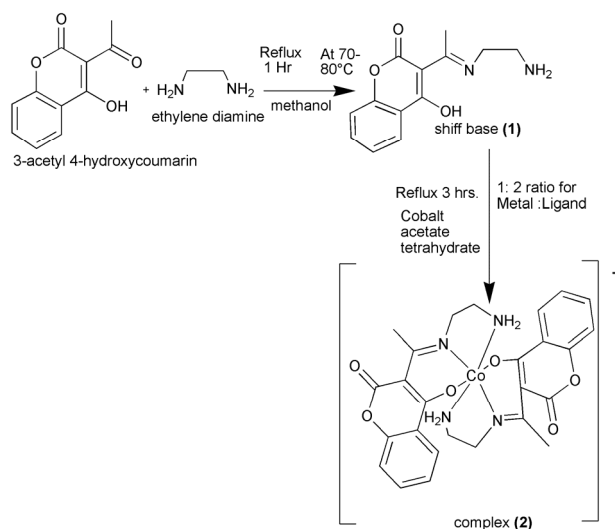


Figure 1. The synthesis steps of the title compound.

radiation at 120 K. The structure was solved with direct methods using SIR 2004 [29]. An anisotropic refinement, by full-matrix least-squares was performed with JANA2006 [30], and the residual finally converged to  $R_1 = 0.08$ . Non-hydrogen atoms were anisotropic and the hydrogen atom positions were included in the riding mode, except for those of water molecule, where they are introduced and refined with distance and angle restrictions. Acetone solvent molecules are not coordinated to cobalt. They just, fill voids in the crystal lattice. One of the two acetone molecules is quite involved in disorder and so, we let it to possess two positions in the same void. For this purpose, a split model in which two slightly different positions for the atoms C45, C44, O46 and C47 was used. Both refined positions had different occupancy and their sum was significantly different from 1. The disorder of this solvent molecule has negative influences on the structure because of the relatively high number of electrons that are disordered with the solvent. This, together with the very weak diffractions, meant that the R values are higher than usual, but the molecular structure is clear. Further details are given in **Table 1**. Selected bond lengths and angles are listed in **Table 2**. The structure exhibits intermolecular hydrogen bonds (**Table 3**).

**Table 1.** Crystal data and structure refinement for cobalt complex.

Empirical formula	C <sub>30</sub> H <sub>46</sub> ClI Co <sub>1</sub> N <sub>4</sub> O <sub>13</sub>
Formula weight	783.06 g.mol <sup>-1</sup>
Temperature	120 (K)
Wavelength (Å)	1.5418 Å
Space group	'P - 1'
<i>a</i> (Å)	10.4261(6)
<i>b</i> (Å)	11.3234(5)
<i>c</i> (Å)	15.7297(7)
$\alpha$ (°)	70.102(4)
$\beta$ (°)	86.049(4)
$\gamma$ (°)	82.497(4)
Volume (Å <sup>3</sup> )	1730.62(15)
<i>Z</i>	2
Dcalc (Mg/m <sup>3</sup> )	1.503
F(000)	848
Crystal size (mm <sup>3</sup> )	0.4 * 0.15 * 0.02
Absorption coefficient (mm <sup>-1</sup> )	5.222
Theta range for data collection	3.01° - 62.69°
Range of <i>h, k, l</i>	-12/12, -11/11, -17/17
Reflections collected number	23897
Total Reflections number	5397
[ <i>R</i> <sub>int</sub> = 0.06]	[ <i>R</i> $\sigma$ ( <i>I</i> ) = 0.13]
Refinement method	full-matrix least-squares on F <sup>2</sup>
Refined parameters/ Restraints/constraints	476/2/152
Final R indices [ <i>I</i> > 3 $\sigma$ ( <i>I</i> )]	R1 = 0.08, wR1 = 0.09

**Table 2.** Selected bond distances (Å) and angles (°).

Bond Lengths (Å)			
Co-O(5)	1.882(4)	C(26)-C(27)	1.522(10)
Co-O(9)	1.900(5)	C(24)-C(25)	1.503(8)
Co-N(4)	1.944(5)	C(23)-C(51)	1.514(8)
CO-N(1)	1.911(5)	C(21)-C(52)	1.503(9)
CO-N(2)	1.915(5)	C(20)-O(4)	1.212(9)
CO-N(3)	1.932(6)		
N(4)-C(24)	1.483(8)		
N(1)-C(21)	1.310(7)		
Selected Angles (°)			
O(5)-CO-O(9) O1 C1 C2	92.3(2)	CO-N(1)-C(21)	127.7(4)
O(5)-CO-N(4)	177.9(2)	CO-N(1)-C(25)	110.4(4)
O(5)-CO-N(1)	93.21(19)	CO-N(2)-C(25)	126.7(5)
O(9)-CO-N(1)	86.4(2)	C(4)-O(7)-C(8)	123(5)
O(9)-CO-N(2)	92.9(2)	N(4)-CO-N(3)	91.8(2)
O(9)-CO-N(3)	177.34(19)		
N(1)-CO-N(3) O1 C1 C2	94.1(2)		
N(2)-CO-N(3)	86.6(2)		
O(1)-O(5)-C(10)	124.9(4)		
O(1)-N(1)-C(21)	127.7(4)		

**Table 3.** Hydrogen bonds for the title compound (Å and °).

<i>D</i> - <i>H</i> ... <i>A</i>	<i>d</i> ( <i>D</i> ... <i>H</i> )	<i>d</i> ( <i>H</i> ... <i>A</i> )	<i>d</i> ( <i>D</i> ... <i>A</i> )	$\angle$ ( <i>DHA</i> )
N3-H2n3...O8 <sup>b</sup>	0.87	2.14	2.963(7)	156.78
N3-H1n3...Ow <sup>a</sup>	0.87	2.19	2.999(6)	154.19
N4-H2n4...O'1	0.87	2.23	3.05(7)	157.63
N4-H1n4...Ow <sup>a</sup>	0.87	2.16	2.904(8)	143.35
Ow-Hw1...O4	0.96(8)	1.89(8)	2.752(7)	148(11)
Ow-Hw2...O'2 <sup>c</sup>	0.96(5)	2.31(4)	3.118(8)	142(5)
Ow-Hw2...O'4 <sup>c</sup>	0.96(5)	2.28(5)	3.018(9)	133(4)

<sup>a</sup>x - 1, y, z; <sup>b</sup>-x, 1 - y, -z; <sup>c</sup>1 - x, -y, 1 - z.

### 3. RESULTS AND DISCUSSIONS

The Schiff base was isolated in 85% yield and was obtained cleanly, melting point 170°C, yellow in color, insoluble in common organic solvents, partially soluble in chloroform but soluble in dimethyl sulfoxide (DMSO). The reaction of the ligand and metal salt gave the corresponding metal complex in good yield, stable and non-hygroscopic in nature. The complex was prepared red in color and was found to be soluble in DMSO, CH<sub>3</sub>COCH<sub>3</sub> but was only partially soluble in other organic solvents such as CHCl<sub>3</sub> and CH<sub>3</sub>OH. It decomposes at about 322°C.

### 3.1. Spectral Characterization of the Ligand and Complex

The IR,  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra of the synthesized materials confirm the formation of ligand and complex with the proposed structure of the type  $[\text{Co}(\text{C}_{13}\text{N}_2\text{O}_3\text{H}_{13})_2]\text{ClO}_4 \cdot 2(\text{C}_3\text{H}_6\text{O}), \text{H}_2\text{O}$ .

### 3.2. IR Spectrum

The IR spectrum of the Schiff base exhibits characteristic intensity band at  $1618\text{ cm}^{-1}$  which is attributed to the  $\nu(\text{C}=\text{N})$  vibration. Another band at  $1637\text{ cm}^{-1}$  is assigned to  $\nu(\text{C}=\text{O})$ , lactonyl carbon of the coumarin moiety. The Schiff base has a strong absorption band at  $3207\text{ cm}^{-1}$ , which corresponds to the  $\text{NH}_2$  stretching frequency [31]. A medium intensity band at  $1585\text{ cm}^{-1}$  is regarded as an aromatic  $\text{C}=\text{C}$  stretching vibrations.

In comparison with the spectrum of the Schiff base, the complex exhibits the band of  $\nu(\text{C}=\text{N})$  in the region  $1602\text{ cm}^{-1}$ , showing the shift of band to lower wave numbers indicating that, the azomethine nitrogen is coordinated to the metal ion. The high intensity band due to enolic  $\text{C}=\text{O}$  appeared at  $1325\text{ cm}^{-1}$  in the Schiff base, appeared as a medium to high intensity band in the  $1418\text{ cm}^{-1}$  region in the complex.

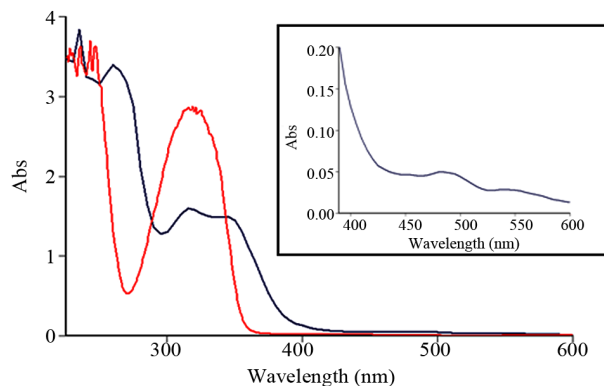
These observations support the formation of  $\text{M}-\text{O}$  bonds via deprotonation. The broad band of medium intensity occurring in the  $3414\text{ cm}^{-1}$  region, corresponds to of water molecules.

The bands in the region of  $410 - 375$  and  $524 - 480\text{ cm}^{-1}$  in the complex are assigned to stretching frequencies of ( $\text{M}-\text{O}$ ) and ( $\text{M}-\text{N}$ ) bonds respectively. The absorption bands near  $523\text{ cm}^{-1}$  and  $418\text{ cm}^{-1}$  are related to the coordination of Schiff-base to cobalt ion through  $\text{Co}-\text{O}$  and  $\text{Co}-\text{N}$  bands [32].

### 3.3. Electronic Spectrum

The electronic spectrum (**Figure 2**) of the free Schiff base ligand shows absorption bands in the ranges  $215 - 260$  and  $270 - 367\text{ nm}$ . The absorption maxima at  $235\text{ nm}$  and  $250\text{ nm}$  are due to  $\pi-\pi^*$  transitions for the benzene ring and enol ring respectively. That observed at  $320\text{ nm}$  is reasonably assigned to  $\pi-\pi^*$  transition for azomethine moiety. The shoulder at  $338\text{ nm}$  corresponds to  $n-\pi^*$  transitions for azomethine and  $\text{NH}_2$  groups [33]. In the spectrum of the complex (**Figure 2**), the absorption bands at ca.  $261, 320, 346\text{ nm}$  of the enol, azomethine and  $\text{NH}_2$  chromophores  $\pi-\pi^*$  and  $n-\pi^*$  transitions are shifted compared to the ligand indicating that the enolic oxygen, the imine and amine nitrogen are involved in the coordination with metal ion.

The spectrum of the metal complex shows (inset of **Figure 2**) that the absorption around  $400 - 600\text{ nm}$  is due to ligand to metal charge transfer and  $d-d^*$  transition band of the metal in the complex.



**Figure 2.** UV-vis absorption spectra of ligand (red spectrum) and complex (blue spectrum). The inset shows the vis. Spectrum of the complex in the region  $400 - 600\text{ nm}$ .

### 3.4. NMR Spectra

The Schiff base and the complex has been characterized by  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra.

In  $^1\text{H}$  NMR spectrum of the Schiff base, signal at  $13.76$  is ascribed to enolic-OH and the methyl protons from the acetyl fragment are identified at  $2.07\text{ ppm}$ . The  $^1\text{H}$ -NMR spectrum of the Schiff base showed the characteristic coumarin aromatic proton signals in the  $7.23 - 8.27\text{ ppm}$ . Two additional signals at  $2.48\text{ ppm}$  and  $2.69\text{ ppm}$  are ascribed to  $\text{CH}_2-\text{CH}_2$ . The signal around  $4.00\text{ ppm}$  is ascribed to  $\text{NH}_2$ .

The comparison of the  $^1\text{H}$ -NMR spectra of the complex and ligand leads to: The resonance due to enolic-OH, around  $13.76\text{ ppm}$  in the Schiff base disappears in the spectrum of complex; this confirms that, the hydroxyl group reacted with metal ion via deprotonation.

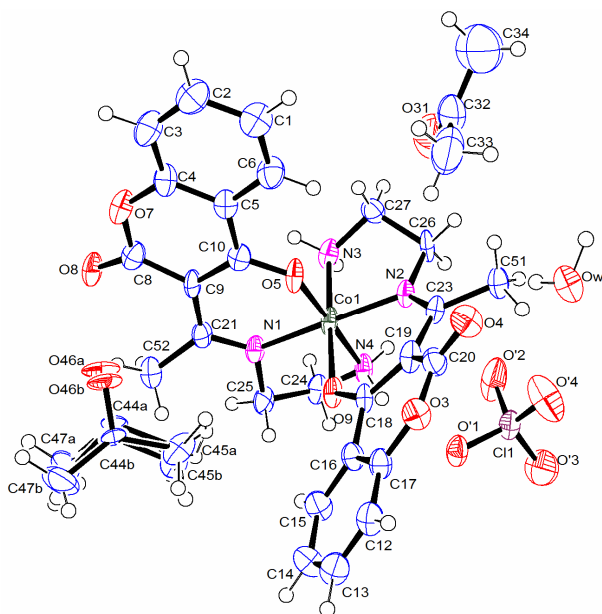
In  $^{13}\text{C}$  NMR spectra of the Schiff base, the signals appeared in the region  $116.20 - 152.99\text{ ppm}$  are assigned to aromatic carbons. The signals at  $179.63, 161.80, 18.31\text{ ppm}$  are due to  $\text{C}=\text{N}, \text{C}=\text{O}$  and  $\text{CH}_3$  respectively. Additional signals at  $34.40\text{ ppm}$  and  $42.81\text{ ppm}$  are ascribed to  $\text{CH}_2-\text{CH}_2$ .

The comparison of the  $^{13}\text{C}$  NMR spectra of the complex and ligand leads to: The azomethine ( $\text{C}=\text{N}$ ) carbon signal occurring, as single at  $179.63\text{ ppm}$  in the ligand, moves upfield to  $161.85$  upon complexation. This provides further support for the coordination of the ligand through the azomethine nitrogens to metal ions.

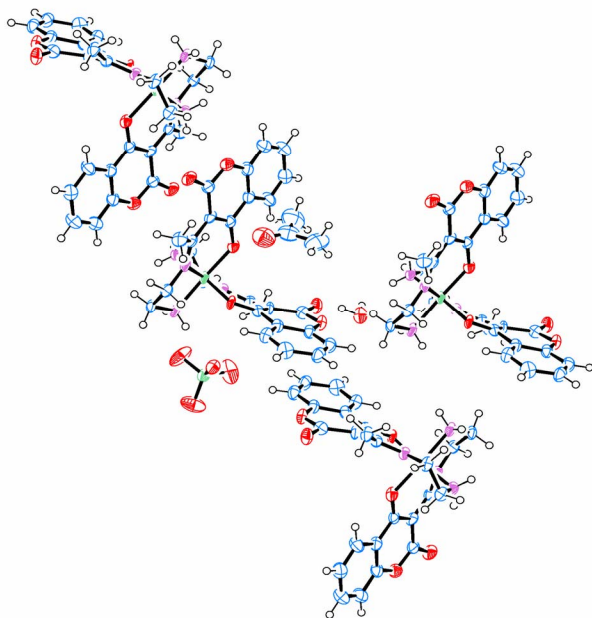
### 3.5. X-Ray Structure Determination of Complex Cobalt

A relatively high coordination number of 6 is achieved by Co, involving two tridentate Schiff base ligands, as shown in **Figure 3** The molecular structure of the complex consists of discrete  $[\text{Co}(\text{C}_{13}\text{H}_{13}\text{N}_2\text{O}_3)_2]^+$  cation,  $\text{ClO}_4^-$  counter anion, acetone solvent and water molecules.

**Figure 4** Shows the packing of the structure. The cobalt(III) center has a distorted octahedral environment,



**Figure 3.** ORTEP stereoscopic projection of the molecular structure of  $[\text{Co}(\text{L})_2]\text{ClO}_4 \cdot 2(\text{C}_3\text{H}_6\text{O}) \cdot \text{H}_2\text{O}$ . Thermal ellipsoids are given at 50% probability.



**Figure 4.** The packing of the structure.

with four nitrogen atoms and two oxygen atoms of the two perpendicular tridentate ligands.

The complex coordinates through imido nitrogen and enolic oxygen with both the nitrogen donors and the enolic oxygen donors being *cis*. The Co-O [1.882(4) Å and 1.944(5) Å] bond lengths are in the range of those expected for  $\text{Co}^{\text{III}}$  in an octahedral environment [34]. The distortion, from regular octahedral geometry is evident

from the N(1)-Co-N(2) bond angle of  $179.2(2)^\circ$ , from O(5)-Co-N(3) and O(9)-Co-N(4) where the angles are  $90.3(2)^\circ$  and  $85.6(2)^\circ$  respectively. The calculated average values of the distortion indices [34,35], corresponding to the different angles and distances in the octahedron, when assuming that the Co-O and Co-N bonds are identical,  $[\text{DI}((\text{O}/\text{N})\text{Co}(\text{O}/\text{N})) = 0.003, \text{DI}(\text{Co}(\text{O}/\text{N})) = 0.011, \text{DI}((\text{O}/\text{N})(\text{O}/\text{N})) = 0.055]$  exhibit a pronounced distortion of the Co(O/N) distances, if compared to (O/N)Co(O/N) angles and (O/N)(O/N) distances.

The diamino fragment is twisted at a torsion angle of  $13.3(1)^\circ$  from the fused ring plane. The dimensions in the shift base correspond closely to standard values [36] except for the bonds C9-C8, C9-C10 and C10-C5, where C9-C10, a double bond, is longer (at 1.414(9) Å) than the expected 1.340 Å, C10-C5 is longer (at 1.447(10) Å) than the expected 1.440 Å, and C9-C8 is shorter (at 1.456(9) Å) than the expected 1.465 Å.

The bond angles C3-C4-O7 and C6-C5-C10 at the junction of the two rings are  $117.6(6)^\circ$  and  $123.6(6)^\circ$ , respectively. Here, as well as in other coumarin compounds [37], an important asymmetry in the O-C-O bond angles was detected;  $\text{O7-C8-O8} = 114.8(6)^\circ$ .

The crystal structure cohesion is stabilized by N---H...O hydrogen bonds between neighboring ligands, and O---H...O links between water molecules and the ligands on the one hand and counter-anions on the other, and ionic interactions between complex cations and counter-anions, these hydrogen bonds link the molecules into a three-dimensional frame. In the O---H...O hydrogen bonds, water molecules were proton donors. Strong H-bonding was observed when the lactone group was a potential hydrogen acceptor (see **Table 3** for the hydrogen bond [Ow-Hw1...O4]).

#### 4. CONCLUSION

The newly synthesized Schiff base which acts as a tridentate ligand was derived in good yield from condensation of ethylene diamine with 3-acetyl-4-hydroxychromene-2-one. It coordinated with the metal ion with two azomethine nitrogen and enolic oxygen atom by deprotonation. The bonding of ligand to metal ion is confirmed by spectroscopic investigation. The structure of  $[\text{Co}(\text{L})_2]^+$  was determined by single crystal X-ray diffraction and revealed that the NNO-donor sets of two ligand fragments form a distorted octahedral configuration at the Cobalt ion. A perpendicularity in the arrangement of the two Schiff base ligands is observed. The structure of the complex revealed the existence of network in one-dimensional zigzag chain propagating along the *c* axis in which the perchlorate ions, the acetone solvent and water molecules are inserted between the *ab* planes.

## 5. SUPPLEMENTARY MATERIAL

CCDC 840711 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif).

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