

Synthesis, Structure, and Biological Assays of Novel Trifluoromethyldiazepine–Metal Complexes

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A new series of Cu^{II}, Ni^{II}, Co^{II}, and Mn^{III} complexes have been synthesised from the (6Z)-6-(7-trifluoromethyl-1,2,3,4-tetrahydro-5H-1,4-diazepin-5-ylidene)cyclohexa-2,4-dien-1-one (HDZP) ligand. These complexes were characterised by elemental, spectroscopic (IR and UV-vis), and thermal analysis. The crystal structure of Cu-DZP was solved by X-ray diffraction methods. The complex crystallises in the monoclinic $P2_1/c$ space group, with two molecules per unit cell. The crystal lattice is stabilised by different intra and intermolecular interactions. Hirshfeld surface analysis was employed to obtain additional information about interactions that are responsible for the crystal packing. Quantitative examination of the fingerprint plots indicated the dominant contribution of H···H and H···X (X = O, F) interactions in the crystal packing. In addition, C–H···chelate ring (CR) and C–H··· π interactions are described in detail and evaluated using DFT calculations. The antibacterial properties and the mechanism of inhibition of the main bacterial resistant mechanism, the biofilm, of the metal complexes and free ligand were investigated. [Mn(DZP)₃] \cdot 2H₂O was the most active complex against the *Pseudomonas aeruginosa* biofilm formation with an inhibition of 40%. However, none of the complexes inhibit more than 25% of the Gram negative bacteria microbial development. The most meaningful result was the bactericidal effect of [Co(DZP)₂(H₂O)₂] \cdot 2H₂O against the Gram positive bacteria, *Staphylococcus aureus*, which inhibits the bacterial development and significantly reduces the biofilm formation at low concentration.

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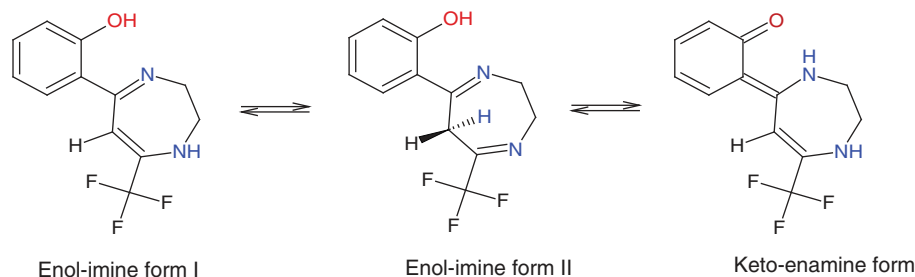
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Introduction

Diazepines (DZPs) and their derivatives are nitrogen-containing heterocyclic compounds which show relevant biological properties and numerous applications in medicinal chemistry. It is well known that 1,4-diazepine heterocyclic compounds such as Zometapine, Etizolam, Brotizolam, Clozapine, Dibenzepine, and Diazepam^[1–3] show a wide range of biological activities being active in the treatment of insomnia,^[4] epilepsy, and

alcoholism.^[5] In recent decades, numerous structural modifications on the diazepine core have been performed in order to expand their biological activities including antibacterial,^[6,7] antioxidant,^[8] antiviral,^[9] anticonvulsant,^[10] and antitumour^[11] properties. In addition, herbicide and anti-HIV activities have been also reported.^[12,13] Generally, DZP derivatives with seven-membered rings are synthesised by a cyclocondensation process with a high regioselectivity. Moreover, DZPs are widely



Scheme 1. Different tautomeric forms of HDZP.

used as synthons for the synthesis of triazole and oxadiazole derivatives.^[14–16] In the last years, the dependence on the stereochemistry of the seven-membered DZP ring derivatives with biological activity have attracted remarkable interest.^[17] The different tautomeric forms of 1,4-DZPs were studied by different experimental techniques such as X-ray diffraction (XRD), nuclear magnetic resonance (NMR), UV-visible spectroscopy, and theoretical calculations. From these studies, it can be concluded that the chair form of the DZP ring is the most stable conformation, although the rapid inter-conversion involving the boat-shape form was also studied by NMR spectroscopy for the DZP ring in 1,4-DZP derivatives and related benzodiazepines.^[18] The introduction of the $-CF_3$ group allowed improvement of both the physicochemical and biological properties of the host compounds, as described in related trifluoromethyl chromones.^[19]

Metal complexes have been widely used in pharmacology and medicinal chemistry for the treatment of many diseases. The interaction between DZP derivatives and metal ions to form metal-based drugs has been investigated by some groups.^[20] Generally, 1,4-DZP can act as a monodentate ligand coordinating through its N atom,^[20,21] but in some cases as a bridging ligand.^[22] Romba and co-workers have reported the structural features of a new tridentate ligand 1,4-diazepan-6-amine and a complete study of the structural and spectroscopic characterisation of Cu^{II} , Ni^{II} , and Zn^{II} metal complexes has been performed.^[23] The inclusion of metal ions into DZP structures can induce changes in the pharmacological properties of the ligand. The use of metal ions in the synthesis of new antibacterial compounds is related to the different mechanism of action that they can present, which is different from the organic ligand, and which generally possess a single molecular target.^[24] For example, Cu^{II} ions can damage cellular membranes in bacteria in order to interact with DNA, causing bacterial death.^[24] In addition, some 1,4-DZP derivatives were found to be highly selective and sensitive chemosensors for Cu^{II} distinguishing it from other cations.^[25]

Non-covalent interactions involving π -systems are relevant in the field of supramolecular chemistry, crystal engineering, and in the study of crystal packing of molecules. In general, organic molecules with planar geometry participate in $C-H \cdots \pi$ and π -stacking interactions. In addition, other planar moieties can also participate in ‘un-conventional’ or ‘non-classical’ $C-H \cdots \pi$ and $\pi \cdots \pi$ contacts. It is well documented that planar chelate rings (CR) with delocalised π -bonds participate in π -stacking interactions analogous to those of aromatic molecules in transition metal complexes. Recent investigations have suggested that these types of interactions are responsible in the stabilisation of inorganic and hybrid metal–organic compounds.^[26–29] In this work we have demonstrated the relevance

of $C-H \cdots \pi(CR)$ in the crystal packing of the $[Cu(DZP)_2] \cdot 2H_2O$ metal complex.

Biofilms are communities of microorganisms that inhabit a self-produced biopolymer. In this environment, antimicrobial resistance is superior to free-living bacteria.^[30] The biofilm adheres to several surfaces and causes persistent contamination in various industries including food.^[31] To solve this problem, new cleaning and disinfection agents with anti-biofilm activity are required. For this reason, the possible antibacterial and anti-biofilm activity of metal complexes and the free ligand was investigated.

Herein, the synthesis, structural, and spectroscopic characterisation of Cu^{II} , Ni^{II} , Co^{II} , and Mn^{III} complexes with a 1,4-DZP derivative, HDZP = (6Z)-6-(7-trifluoromethyl-1,2,3,4-tetrahydro-5H-1,4-diazepin-5-ylidene)cyclohexa-2,4-dien-1-one (Scheme 1) is reported. Recently, this compound was fully characterised and its tautomerism was studied in solid, solution, and gas phase.^[32] The synthesised compounds were characterised by FT-IR, UV-vis, and diffuse reflectance spectroscopic studies. The crystal structure of the Cu^{II} complex has been solved by X-ray diffraction methods and a detailed analysis of the main intermolecular interactions has been evaluated by Hirshfeld surface (HS) analysis. Thus, the compound surface was mapped by using d_{norm} , shape index, and curvedness properties, for the exploration of the packing modes and visualisation of the intermolecular interactions. The role of non-covalent interactions in the stability of the crystal packing was evaluated by using density functional theory (DFT) calculations and topological analysis of the electron density by Bader’s atoms in molecules (AIM) approach.

Experimental

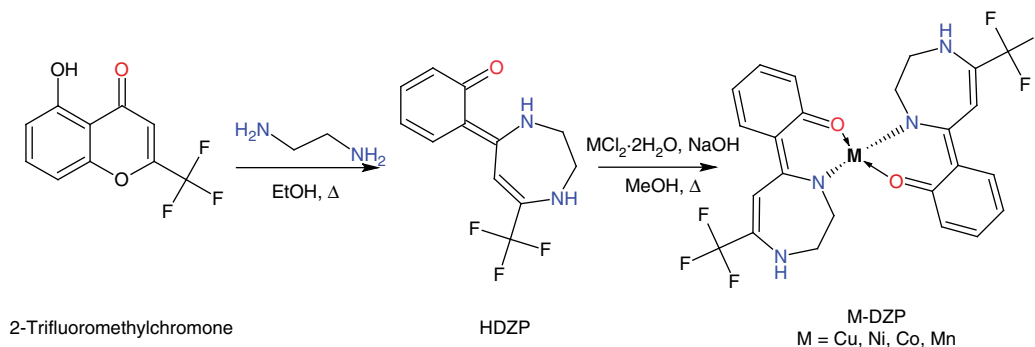
Synthetic Procedures

The ligand HDZP was synthesised as previously reported^[32,33] by mixing ethanolic solutions of 2-trifluoromethylchromone (4 mmol) and ethylenediamine (4 mmol) at room temperature as shown in Scheme 2. The orange solid thus obtained was recrystallised from hot ethanol.

Anal. Calc. for $C_{12}H_{11}F_3N_2O$: C 56.25, H 4.29, N 10.94%. Found: C 56.40, H 4.32, N 10.75.

Synthesis of $[Cu(DZP)_2] \cdot 2H_2O$ and $[Ni(DZP)_2]$ Complexes

$[Cu(DZP)_2] \cdot 2H_2O$ and $[Ni(DZP)_2]$ complexes were prepared as follows: NaOH (0.08 g, 2 mmol) was added to a solution of HDZP (0.512 g, 2 mmol) in 5 mL of methanol. The dispersion was mixed with $CuCl_2 \cdot 2H_2O$ (0.170 g, 1 mmol) or $NiCl_2 \cdot 6H_2O$ (0.238 g, 1 mmol) dissolved in 5 mL of methanol. The mixtures were refluxed at 60°C for 4 h with continuous stirring



Scheme 2. Synthetic route for the ligand HDZP and its complexes.

(see Scheme 2). After cooling, the solids obtained were filtered off, washed with cold methanol, and dried in a desiccator under CaCl_2 . Only single crystals of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ adequate for structural study were obtained from slow evaporation of the solvent (ethanol) at room temperature.

$[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$. Yield: 75%. Red crystals. Anal. Calc. for $\text{CuC}_{24}\text{H}_{24}\text{F}_6\text{N}_4\text{O}_4$: C 47.25, H 3.94, N 9.19%. Found: C 47.32, H 4.01, N 9.25.

$[\text{Ni}(\text{DZP})_2]$. Yield: 80%. Orange powder. Anal. Calc. for $\text{NiC}_{24}\text{H}_{20}\text{F}_6\text{N}_4\text{O}_2$: C 50.64, H 3.52, N 9.85%. Found: C 50.15, H 3.62, N 9.83.

Synthesis of $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$ and $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ Complexes

The synthetic method for both complexes was similar to $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ and $[\text{Ni}(\text{DZP})_2]$, except that in these cases $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (0.238 g, 1 mmol) and $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ (0.198 g, 1 mmol) were used. The resulting solutions were left to evaporate slowly at room temperature and the obtained powders were dissolved in ethyl acetate and the precipitated NaCl was separated by filtration. The solvent was evaporated to dryness and the solids were re-crystallised from hot methanol. Due to the amorphous nature of both complexes, single crystals could not be isolated.

$[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$. Yield: 85%. Green powder. Anal. Calc. for $\text{CoC}_{24}\text{H}_{28}\text{F}_6\text{N}_4\text{O}_6$: C 44.93, H 4.37, N 8.73%. Found: C 44.85, H 4.30, N 8.65.

$[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$. Yield: 82%. Brown powder. Anal. Calc. for $\text{MnC}_{36}\text{H}_{34}\text{F}_9\text{N}_6\text{O}_5$: C 50.47, H 3.97, N 9.81%. Found: C 50.72, H 3.98, N 9.77.

Materials and Physical Measurements

Reagent grade compounds were used for synthesis without further purification. The IR absorption spectra were recorded in the solid state using KBr pellets on a FTIR Perkin Elmer GX1 in the $4000\text{--}400\text{ cm}^{-1}$ frequency range with 4 cm^{-1} spectroscopic resolution. Diffuse reflectance UV-Vis spectra of HDZP and its metal complexes were measured using a Shimadzu UV-2600 spectrophotometer equipped with the integrating sphere (BaSO_4 was used as reference). Elemental analysis of carbon, hydrogen, and nitrogen were performed using a Carlo Erba EA1108 analyser. Thermogravimetric measurements were performed on a Shimadzu TG-50 thermo-balance between 25 and 800°C at a heating rate of 5°C min^{-1} under air flow. The magnetic susceptibility measurement for $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ was performed with a magnetic susceptibility balance (MSB-Auto) at room temperature and diamagnetic corrections were performed using Pascal's constants.

X-Ray Diffraction Data and Structure Refinement of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$

X-Ray diffraction intensities were collected on an Oxford Xcalibur Gemini, Eos CCD diffractometer employing ω scans with θ and κ -offsets and processed with *CrysAlisPro*.^[34] The unit cell constants were derived from least-squares fitting to the angular setting of all gathered reflections having intensities above seven times the measurement errors. Absorption distortion of diffraction data were corrected employing the empirical multi-scan method. The crystal structure was resolved with the dual method implemented in *SHELXT*^[35] and refined with *SHELXL*.^[36] The electron density map showed two types of disorder in the lattice:

- (i) Disordered solvent, as confirmed by FT-IR absorption showing the signature of a loosely bonded crystallisation water and by thermogravimetric analysis (TGA) data indicating the loss of two solvent molecules per complex at a relatively low temperature (around 85°C); the solid can thus be described as $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$. This disorder could not be modelled satisfactorily in terms of a discrete solvent distribution and therefore it was resorted to a procedure described by Van der Sluis and Spek^[37] and implemented in the program *SQUEEZE* included in the *PLATON* suit of programs^[38] to remove from the diffraction data the contribution due to this electron density;
- (ii) Rotational disorder of the $-\text{CF}_3$ group. This was modelled in terms of three angular conformations whose occupancies summed up to one. The $-\text{CF}_3$ conformers were refined restraining the C–F and F...F distances to be respectively equal to one another. The H-atoms were positioned at the expected geometrical locations and refined with the riding model.

Table 1 shows a summary of crystal data and structure refinement results. Crystallographic data have been deposited at the Cambridge Crystallographic Data Centre (CCDC) with the reference number CCDC 1848051.

Hirshfeld Surface Calculations

HS analysis and the decomposed two-dimensional fingerprint (FP) plots^[39–42] were used to understand the nature of intermolecular interactions which stabilise the crystal packing and to quantify the contribution of different contacts. The HS and their associated FP plots were generated using the *CrystalExplorer17* program,^[43] using the crystallographic information files (cif) obtained from the X-ray single crystal measurements. The d_{norm} HS were mapped over a fixed colour scale of -0.025 au (red) to

Table 1. Crystallographic data and structure refinement results for [Cu(DZP)₂]·2H₂O

Parameter	[Cu(DZP) ₂]·2H ₂ O
Empirical formula	C ₂₄ H ₂₀ CuF ₆ N ₄ O ₂
Formula weight	573.98
Temperature [K]	294(2)
Wavelength [Å]	0.71073
Crystal system	Monoclinic
Space group	<i>P</i> 2 ₁ / <i>c</i>
<i>a</i> [Å]	8.5718(4)
<i>b</i> [Å]	12.6039(5)
<i>c</i> [Å]	12.6422(4)
β [deg.]	98.803(4)
Volume [Å ³]	1349.75(9)
Z, density [calc., mg m ⁻³]	2, 1.412
Absorpt. coeff. [mm ⁻¹]	0.877
F(000)	582
Crystal size [mm ³]	0.296 × 0.221 × 0.207
θ -range for data collect [deg.]	3.233 to 28.927
Index ranges	−11 ≤ <i>h</i> ≤ 10, −15 ≤ <i>k</i> ≤ 16, −17 ≤ <i>l</i> ≤ 15
Reflections collected	6168
Independent reflections	2930 [<i>R</i> (int) 0.0279]
Observed reflections [<i>I</i> > 2 σ (<i>I</i>)]	2271
Completeness [%]	99.8 (to $\theta = 25.242^\circ$)
Refinement method	Full-matrix least-squares on <i>F</i> ²
Data / restraints / parameters	2930/72/223
Goodness-of-fit on <i>F</i> ²	1.078
Final <i>R</i> indices ^A [<i>I</i> > 2 σ (<i>I</i>)]	<i>R</i> ₁ 0.0428, <i>wR</i> ₂ 0.1038
<i>R</i> indices (all data)	<i>R</i> ₁ 0.0595, <i>wR</i> ₂ 0.1132
Larg. diff. peak and hole [e Å ⁻³]	0.266 and −0.290

$$^A R_1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|}, wR_2 = \frac{[\sum w(|F_o|^2 - |F_c|^2)^2 / \sum w(|F_o|^2)]^{1/2}}{}$$

0.75 au (blue), shape index mapped in the colour range of −1.0 au (concave) to 1.0 au (convex) and curvedness in the range of −4.0 au (flat) to 4.0 au (singular). The 2D FP plots were displayed using the translated 0.6–2.4 Å range including reciprocal contacts.

Computational Details

DFT calculations for [Cu(DZP)₂]·2H₂O were performed using the *Gaussian 09*^[44] program at the B3LYP/cc-pVDZ approximation. In order to obtain the calculated energies involved in the intermolecular interactions that stabilise the crystal packing, the crystallographic coordinates were used instead of those from the optimised structure of the assemblies.

The molecular electrostatic potential (MEP) surface was computed at the same level of theory, using the 0.001 a.u. isosurface. To analyse the bonding nature of the compounds, topological analysis of the electron density was performed using the atoms in molecules (AIM) approach.^[45]

Biological Assays

Growth Test

A 96-well polystyrene microplate assay (flat base) was performed by determining the absorbance. In each well of the microplate (four repetitions), 5 μ L of compound solutions (4000 and 400 μ g mL⁻¹) were added. Bacterial suspensions were carried out in Luria–Bertani (LB) and Müller–Hinton (MH) broth from overnight cultures in a MH agar plate of *Pseudomonas aeruginosa* PAO1 and *Staphylococcus aureus* ATCC 6538, respectively. The bacterial suspension was adjusted to a final

concentration of 10⁷ colony forming units (CFU) mL⁻¹ and 195 μ L of that culture was placed in each microplate well, assaying concentrations of 100 and 10 μ g mL⁻¹ of the compounds. Solvent control (four repetitions) with 5 μ L of DMSO and 195 μ L of the culture were used. The 96-well plates were statically incubated at 37°C for 24 h and growth was detected as turbidity (560 nm) using a microplate reader (Thermo Fisher Scientific Oy Ratastie 2, FI-01620, Vantaa, Finland).

Biofilm Formation Assay

A micro-method reported previously^[46] was used for the biofilm quantification, with few modifications^[47] and the influence of the complexes were evaluated after 24 h incubation, using crystal violet. The results obtained were processed in *Excel* and analysed with the statistical program *Minitab 1.6*. An asterisk was used to indicate significant differences (*P* < 0.05).

Minimal Inhibitory Concentration (MIC) and Minimal Bactericidal Concentration (MBC) Determination

This test was performed in micro plates. The MIC of the most active compound found in the growth test described in the section *Growth Test* was determined by the broth micro dilution technique according to the Clinical and Laboratory Standards Institute (CLSI).^[48] The concentrations tested were 8, 16, 32, 64, 128, 256, 512, and 1024 μ g mL⁻¹. The bacterial inoculum (100 μ L) containing 5 × 10⁵ CFU mL⁻¹ was individually added to each well. All controls were performed for sterility, viability, solvent effect (DMSO 2%), and an antibiotic (ciprofloxacin) used as reference. After incubation at 37°C during 24 h, for controls and samples under study, it was determined in which well the formation of a microbial precipitate was not observed. The MIC was the concentration of each sample that inhibits microbial development. From these samples, a 10 μ L aliquot was extracted and plated in fresh medium for 24 h at the same temperature to evaluate which samples were able to kill bacteria (MBC).

Results and Discussion

Synthesis of Metal Complexes

The synthesis of the complexes was performed by mixing methanolic solutions of CuCl₂·2H₂O, NiCl₂·6H₂O, CoCl₂·6H₂O, and MnCl₂·4H₂O for [Cu(DZP)₂]·2H₂O, [Ni(DZP)₂], [Co(DZP)₂(H₂O)₂]·2H₂O, and [Mn(DZP)₃]·2H₂O, respectively and the corresponding HDZP ligand (see *Scheme 2*). The ligand-to-metal molar ratio was 2 : 1 for all complexes. The resulting products were coloured powders and the crystal structure of [Cu(DZP)₂]·2H₂O was solved by single crystal X-ray diffraction measurements. Elemental analysis and TGA data were satisfactory with the proposed empirical formula. These results also revealed that the initial ligand-to-metal ratio of 2.1 was maintained for Cu^{II}, Ni^{II}, and Co^{II} complexes, but the ratio of 3 : 1 was observed for the Mn^{III} complex. In [Mn(DZP)₃]·2H₂O, Mn^{II} was oxidised by atmospheric oxygen to Mn^{III} under the reaction conditions.^[49] The value of the magnetic moment at room temperature (3.00 μ _B) is close to the spin-only value of 2.88 μ _B, as expected for a low-spin d⁴ manganese(III) ion.

Spectroscopic Studies

FT-IR Spectroscopy

The FT-IR spectra of the complexes were compared with free HDZP, as shown in *Fig. 1*. The HDZP spectrum has numerous bands and the most relevant ones related to the coordination sites

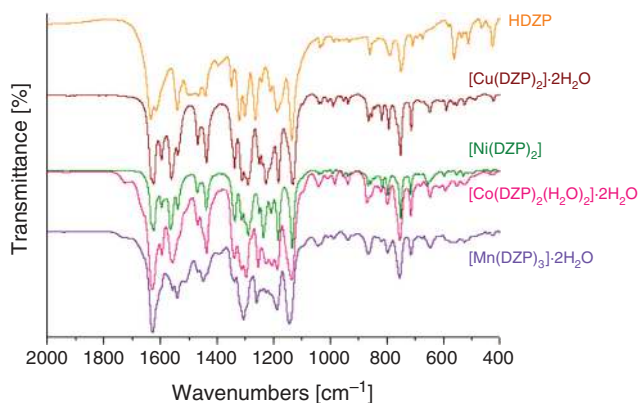


Fig. 1. FT-IR spectra of HDZP ligand and its Cu^{II}, Ni^{II}, Co^{II}, and Mn^{III} complexes.

could be assigned (See Table S1, Supplementary Material). The HDZP ligand could exist in three tautomeric forms, as shown in Scheme 1. The crystal structure of the HDZP ligand indicates that the most stable form in the solid state is the *keto-enamine* tautomer. This is in accordance with the intense absorption band observed at 1635 cm^{-1} in the IR spectrum of the solid, assigned to the C=O stretching mode. Several bands between 3080 and 2865 cm^{-1} in the free ligand and its metal complexes are attributed to C–H stretching vibrations of aromatic and aliphatic hydrogens. The bands corresponding to the $\nu(\text{O–H})$ modes of non-coordinated water molecules in $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ are not clearly visible in the IR spectrum because they are overlapped with the bands related to the N–H stretching mode (see Fig. S1, Supplementary Material). The spectra of $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$ and $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ show a broad band in the $3550\text{--}3375\text{ cm}^{-1}$ frequency range, according to the existence of water molecules in the complexes (see Figs S3 and S4, Supplementary Material). The other corresponding bands of water molecules could not be assigned, due to the fact that they overlapped with some other modes. The frequency of the N1–H stretching of the diazepine, observed at 2768 cm^{-1} in the free ligand, was not detected in the spectra of the complexes, confirming the deprotonation of the N1–H moiety. The C7–N1 stretching vibration located at around 1540 cm^{-1} in HDZP is shifted towards higher frequencies upon complexation with metal ions. The C=O stretching mode of the *keto-enamine* tautomer of HDZP in the solid state appears as an intense band at 1635 cm^{-1} whereas, for all complexes, the C=O stretching vibration is shifted to lower frequencies (between $1624\text{--}1628\text{ cm}^{-1}$) because of the increased length of the C=O bond coordinated to the metal ions. The band at 1615 cm^{-1} associated with C–C stretching modes of the phenyl ring changed slightly upon complexation, indicating that the complex formation alters the ring structure. The bands corresponding to the vibrations of the diazepine ring were, in some extent, affected by the interaction of HDZP with the metal ions (see Table S1). The evident negative shift of the C–C=O bending mode in the complexes suggests that the ligand coordinates the metal ions by the oxygen atoms, which is confirmed by the crystal structure determination of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ (see below). The metal complexes exhibit bands in the range $552\text{--}593\text{ cm}^{-1}$ and $451\text{--}488\text{ cm}^{-1}$ which are assigned to M–N and M–O stretching modes, indicating that the coordination of the metal ions with the HDZP ligand took place probably via carbonyl oxygen and N1 atoms of the DZP ring after deprotonation.

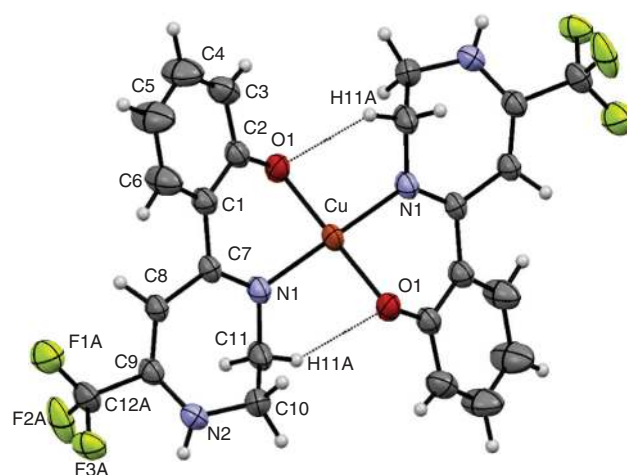


Fig. 2. Plot of the molecular structure of $[\text{Cu}(\text{DZP})_2]$, showing the atomic labelling and the displacement ellipsoids of non-H atoms at the 50% probability level. Intramolecular hydrogen bonds are shown as dashed lines. For clarity, from the three rotationally split positions that model the rotational disorder of the $-\text{CF}_3$ group, only the conformation having the largest occupancy is shown.

Electronic Spectroscopy

The electronic absorption spectra of the ligand and its metal complexes (Fig. S5, Supplementary Material) were recorded in the solid state at room temperature. The electronic data of the HDZP ligand and the $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$, $[\text{Ni}(\text{DZP})_2]$, $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$, and $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ complexes are summarised in Table S2 (Supplementary Material). The ligand shows absorption bands at 254 and 279 nm attributed to $\pi \rightarrow \pi^*$ transitions and a band at 352 nm which could be assigned to $n \rightarrow \pi^*$ transitions. The electronic spectra of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ and $[\text{Ni}(\text{DZP})_2]$ exhibit three bands at 231 , 290 , and 337 nm and at 234 , 268 , and 334 nm , respectively. These absorptions correspond to intra-ligand transitions. The absorption bands at 492 and 674 nm in the electronic spectra of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ are assigned to ${}^2\text{B}_{1g} \rightarrow {}^2\text{B}_{2g}$ and ${}^2\text{B}_{1g} \rightarrow {}^2\text{A}_{1g}$ transitions, respectively. The broad band at 563 nm in the $[\text{Ni}(\text{DZP})_2]$ spectra is attributed to the ${}^1\text{A}_{1g} \rightarrow {}^1\text{B}_{1g}$ electronic transition. These results are consistent with the square-planar geometry of Cu^{II} and Ni^{II} in the complexes.^[50] The $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$ and $[\text{Ni}(\text{DZP})_2]$ spectra exhibit absorption bands at 387 and 389 nm , respectively which can be attributed to ligand to metal charge transfer (LMCT) transitions. The $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$ spectrum shows bands at 418 and 629 nm attributed to ${}^4\text{T}_{1g}(\text{F}) \rightarrow {}^4\text{T}_{1g}(\text{P})$ and ${}^4\text{T}_{1g}(\text{F}) \rightarrow {}^4\text{A}_{2g}(\text{F})$ transitions, respectively. These bands are in accordance with an octahedral geometry of the Co^{II} ion.^[51] The electronic spectrum of the $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ complex exhibits an absorption at 408 nm attributed to LMCT, typically of Mn^{III} complexes.^[52]

Description of Crystal and Molecular Structure of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$

Fig. 2 shows an ORTEP^[53] drawing of the $[\text{Cu}(\text{DZP})_2]$ complex and selected bond lengths and angles are listed in Table 2. The experimental geometrical parameters are compared in Table S3 (Supplementary Material) with the corresponding calculated ones at the B3LYP/cc-pVDZ level of theory. The metal is sited on a crystallographic inversion centre, in a square planar environment, coordinated to two deprotonated DZP anions, acting as bidentate ligands through its carbonyl oxygen

($d(\text{Cu}-\text{O}) = 1.925(2) \text{ \AA}$) and amine nitrogen ($d(\text{Cu}-\text{N}) = 1.979(2) \text{ \AA}$) atoms nearly along their electron lone-pair lobes maximising their overlap with the copper $d(x^2-y^2)$ orbital.

The bond lengths and angles are like those observed in related complexes.^[54–56] In a previous work, the crystal structure of the free HDZP ligand was fully elucidated by DRX, resulting in the *keto-enamine* being the preferred tautomer in the solid state ($\text{O}=\text{C}(\text{R})-\text{C}=\text{C}-\text{N}-\text{H}(\text{R})$).^[33] Most bond lengths of the DZP ligand in the copper complex are remarkably similar to the corresponding ones of the HDZP ligand. Upon its coordination to the Cu^{II} ion, the most significant changes are the lengthening of the carbonyl $\text{C}=\text{O}$ bond (in $+0.016 \text{ \AA}$ or $\sim 5\sigma$) and to a lesser extent to a shortening of the amino $\text{C}-\text{N}$ bond (in $-0.01 \text{ \AA} = -2.8 \sigma$). In addition, a lengthening of the bond connecting the aromatic and the DZP rings ($\text{C}1-\text{C}7$: $1.475(4) \text{ \AA}$) and a shortening of the $\text{C}1-\text{C}2$ bond ($1.421(4) \text{ \AA}$, aromatic ring) with respect to the free ligand (1.459 and 1.304 \AA , respectively) is observed. This can be understood assuming that the interatomic distance changes observed between the free and the coordinate HDZP ligand corresponds to a charge distribution oriented to the resonance form ($\text{O}^--\text{C}(\text{R})=\text{C}=\text{N}(\text{R})$). Since the ligand is deprotonated before the complex formation, the interatomic distances reflect the weight of each resonance forms

Table 2. Experimental and calculated bond lengths and angles for $[\text{Cu}(\text{DZP})_2]$

Bonds	Experimental ^A	Calculated ^B
Lengths [\AA]		
Cu–N	1.979(2)	2.011
Cu–O	1.925(2)	1.927
N1–C7	1.305(3)	1.319
N1–C11	1.458(4)	1.459
C11–C10	1.533(4)	1.537
C10–N2	1.442(4)	1.455
N2–C9	1.334(3)	1.355
C9–C12A	1.519(4)	1.523
C9–C8	1.354(4)	1.369
C8–C7	1.456(3)	1.464
C7–C1	1.475(4)	1.473
C2–O	1.320(3)	1.309
C2–C1	1.421(3)	1.438
C2–C3	1.401(4)	1.425
Angles [deg.]		
O#1–Cu–N1	92.37(9)	91.22
O#1–Cu–O	180.0	179.9
N1–Cu–N#1	180.0	179.9
C7–N1–Cu	123.94(18)	123.4
C11–N1–Cu	117.38(17)	115.8
C2–O–Cu	116.52(16)	123.4

^AExperimental parameters obtained in this work.

^BCalculated at B3LYP/cc-pVDZ level.

depicted in Scheme 3. Considering the observed results, it can be concluded that II is the predominant form.

The $[\text{Cu}(\text{DZP})_2]$ crystal structure is further stabilised by an extensive network of intermolecular $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds between neighbouring complexes ($d(\text{N}2\cdots\text{O}') = 2.871(3) \text{ \AA}$, $d(\text{H}\cdots\text{O}') = 2.17 \text{ \AA}$, $\angle(\text{N}2-\text{H}\cdots\text{O}') = 139^\circ$), giving rise to a layered structure parallel to the crystal (100) plane (see Fig. 3). More details of the hydrogen bonding interactions are listed in Table S3.

The weak non-covalent interactions such as $\text{C}-\text{H}\cdots\text{X}$ ($\text{X} = \text{N}, \text{O}, \text{S}$, or halogen) and π -stacking are weaker than classical hydrogen bonds and they have an important role in structural biology and supramolecular chemistry.^[57,58] In addition, these interactions also have the ability to regulate antigen–antibody recognition.^[59] In the crystal structure of $[\text{Cu}(\text{DZP})_2]$, two identical intramolecular $\text{C}-\text{H}\cdots\text{O}$ contacts are observed (see Fig. 2). Intramolecular hydrogen bonds occur between $\text{C}11-\text{H}11\text{A}\cdots\text{O}$ ($2.47(3) \text{ \AA}$) atoms with the formation of five-membered rings ($\text{Cu}-\text{N}1-\text{C}11-\text{H}11\text{A}-\text{O}$). The crystal structure of $[\text{Cu}(\text{DZP})_2]$ shows chelate ring (CR) $\cdots\text{H}$ interactions, with a $\text{Cg}1\cdots\text{H}$ distance of 3.047 \AA ($\text{Cg}1 = \text{Cu}/\text{O}/\text{C}2/\text{C}1/\text{C}7/\text{N}1$), as shown in Fig. 4. Recently, numerous investigations reported evidence of $\text{C}-\text{H}\cdots\pi$ and π -stacking interactions with the π -cloud of a six-membered chelate ring due to the planarity and delocalisation of the π -system.^[27–29] Also, there are intermolecular $\text{C}-\text{H}\cdots\pi$ interactions ($\text{C}11-\text{H}11\text{A}\cdots\text{Cg}2 = 3.875 \text{ \AA}$), where $\text{Cg}2$ is the ring centroid defined by atoms $\text{C}1-\text{C}6$ (see Fig. S6, Supplementary Material). The nature of the $\text{CR}\cdots\text{H}$ interaction is discussed in the theoretical study (section *Theoretical Results*).

Hirshfeld Surface Analysis

HS analysis is an important tool in the study of crystal structures of numerous inorganic and organic compounds, since this method proposes a facile way of obtaining information about intermolecular interactions that stabilise the crystal packing.^[41]

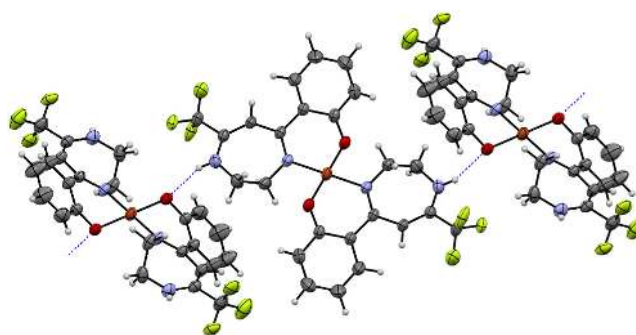
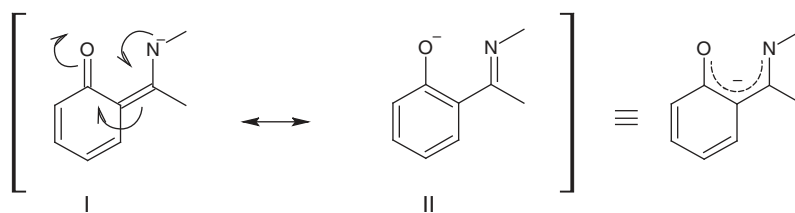


Fig. 3. Partial view of the crystal structure of $[\text{Cu}(\text{DZP})_2]$ showing $\text{N}-\text{H}\cdots\text{O}$ intermolecular hydrogen bonds as dashed lines.



Scheme 3. Resonance forms of the HDZP ligand.

The HSs mapped over d_{norm} properties are displayed in Fig. 5a. The HSs, mapped over shape index and curvedness, are shown in Figs S6 and S7 (Supplementary Material). The deep red visible spots on the HS is indicative of O...H and F...H contacts. The dominant interactions are observed in d_{norm} surface plots as bright areas in Fig. 5a. The full 2D FP plot of [Cu(DZP)₂] illustrates the significant differences between the intermolecular interaction patterns. In the FP plot (Fig. 5b), O...H, F...H, and C...H interactions are represented by spikes in the bottom right area, in the same way, the H...O, H...F, and H...C contacts are visualised in the top left region.

According to the HS analysis, the non-directional H...H contacts, comprising 25.3% of the total HS, are one of major contributors to the crystal packing. As shown in Fig. 5, the shortest H...H contacts in the FP plot are at ($d_c + d_i$) 2.45 Å. The H...H contacts are generally characterised by broader or sharper spikes in the FP plots. Furthermore, a subtle feature, the splitting of the H...H contact in the FP plot is evident for [Cu(DZP)₂]. This splitting occurs when the shorter contact is between three atoms, rather than a direct two-atom contact.

The proportions of H...O/O...H interactions comprise 6.3% of the total HS and are characterised by spikes at $d_c + d_i$

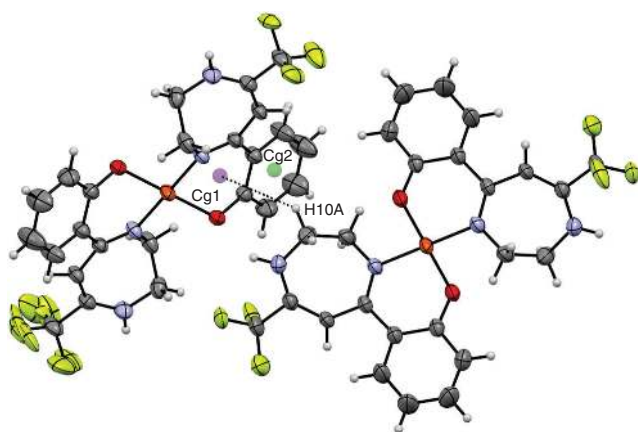


Fig. 4. Crystal packing of [Cu(DZP)₂] showing C–H... π (CR) contacts. Cg1 and Cg2 are centroids defined by Cu/O/C2/C1/C7/N1 and C1–C6 atoms, respectively.

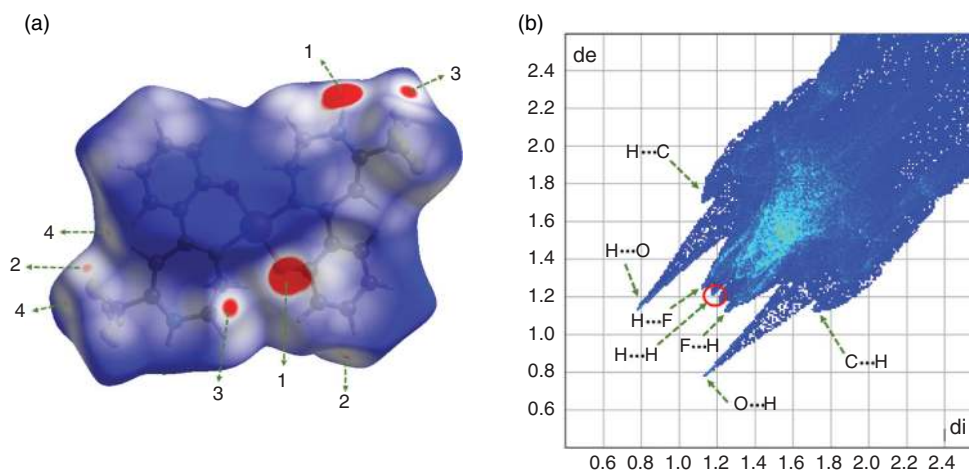


Fig. 5. (a) Hirshfeld surface mapped over d_{norm} and (b) full 2D fingerprint plot of the [Cu(DZP)₂] complex showing the spikes of the main intermolecular contacts in the crystal packing.

~1.95 Å, associated to N1–H...O1 intermolecular interactions. These interactions seem to be relevant, as indicated by the deep red spots in the d_{norm} map labelled 1.

The most important contacts are H...F/F...H, with a magnitude of 39.9% of the total HS area. The red spots labelled as 2 and 4 on the d_{norm} surface are associated to C4–H4...F1A and C6–H6...F1A interactions, respectively (see Fig. S9, Supplementary Material) with H...F1A distances of 2.590 and 2.658 Å. The red spots labelled 3 in the d_{norm} map are associated to C–H...F intermolecular interactions. The F3A atom is involved with the H10B of the methylene group embedded in the DZP ring through a non-classical intermolecular hydrogen bond C10–H10B...F3A, with a H...A distance of 2.660 Å (see Fig. S10, Supplementary Material).

The FP plots also show C...H/H...C intermolecular interactions that involve 14.2% of the total HS.

The shape index and curvedness maps (Figs S7 and S8) are significant indicators for π -stacking interactions within the crystal structure. For [Cu(DZP)₂], no complementary red and blue triangles have been observed in the shape index map indicating the absence of π ... π interactions.

Thermal Analysis

TGA was applied to investigate the thermal stability and behaviour of [Cu(DZP)₂].2H₂O, [Ni(DZP)₂], [Co(DZP)₂-(H₂O)₂].2H₂O, and [Mn(DZP)₃].2H₂O complexes. The results are shown in Figs S11–S14 (Supplementary Material).

The [Cu(DZP)₂].2H₂O complex decomposes in a three-stage process. The first mass loss step is observed at ~85°C and corresponds to the removal of two water molecules (exp. mass loss: 6.54%, theoretical: 6.00%). The dehydration process was observed at low temperature because the water molecules are weakly bonded (uncoordinated water molecules). The second and third mass losses are associated to the degradation of DZP ligands to form CuO as a final product. The observed mass loss (87.0%) is in very good agreement with the calculated value (87.5%).

The curve for the thermal decomposition of [Ni(DZP)₂] exhibits two pronounced mass loss steps. It starts at temperatures above 300°C and the two consecutive steps are associated with the removal of two DZP molecules to form NiO as the final product (exp. mass loss: 88.0%, theoretical mass loss: 87.0%).

The first step for the thermal decomposition of $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$ finishes at $\sim 130^\circ\text{C}$ with an experimental mass loss of 5.54 % (theoretical value: 5.62 %), and can be attributed to the loss of two non-coordinated water molecules. The second decomposition process finishes at 200°C and corresponds to the removal of two coordinated water molecules with a mass loss of 5.66 % (calculated: 5.58 %). The next degradation steps occur within the range of $200\text{--}450^\circ\text{C}$ and are attributed to the removal of two DZP ligands to form CoO as the final product. The experimental mass loss (89.1 %) agrees with that calculated (88.3 %) for the complete exclusion of all DZP ligands. In accordance with elemental analysis and thermal data we can conclude that the Co^{II} ion is in an octahedral environment coordinated to two DZP ligands and two water molecules. A proposed molecular structure for $[\text{Co}(\text{DZP})_2(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$ is shown in Fig. S15 (Supplementary Material).

The thermal decomposition of $[\text{Mn}(\text{DZP})_3] \cdot 2\text{H}_2\text{O}$ is quite different. The first step finishes at 150°C and it is due to the loss of two weakly bonded water molecules (exp. mass loss: 4.26 %, theoretical: 4.20%). The subsequent decomposition processes are associated with the combustion of the three DZP ligands, with a total mass loss of 93.0%, in accordance with the calculated value (92.0%) expected for the formation of Mn_3O_4 as the final product. These results indicate that the Mn^{III} is octahedrally coordinated to three DZP ligands and crystallised with two non-coordinated water molecules (see Fig. S15).

Theoretical Results

The theoretical study was performed to compare the energetic features of the different types of intermolecular interactions

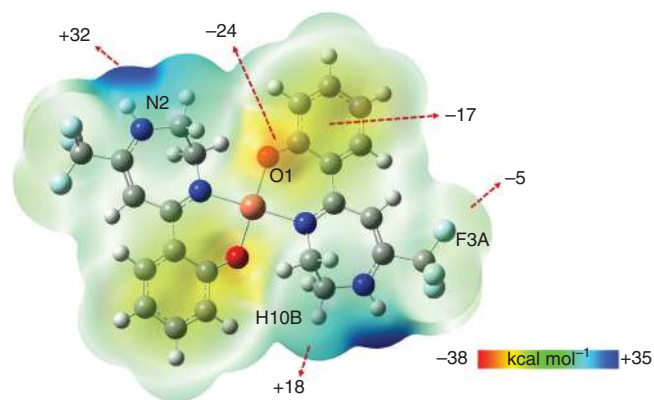


Fig. 6. MEP map of $[\text{Cu}(\text{DZP})_2]$ showing the energetic values at selected points in kcal/mol.

observed in the crystal packing of the $[\text{Cu}(\text{DZP})_2]$ complex. In particular, $\text{N-H}\cdots\text{O}$, $\text{C-H}\cdots\text{F}$, and $\text{C-H}\cdots\pi$ interactions have been analysed, since they are decisive to understand the crystal packing of the complex. We have computed the molecular electrostatic potential (MEP) to understand the electronic nature of the different donor–acceptor sites and the aromatic and chelate rings (Fig. 6). As shown in Fig. 6, the most negative electrostatic potential corresponds to the region of the O atoms from the C=O group ($-24 \text{ kcal mol}^{-1}$) and the most positive part is located in the region of the N2-H bond of the diazepine ring ($+32 \text{ kcal mol}^{-1}$). These results are indicative that the hydrogen bonds between these groups should be energetically favoured. In addition, the MEP values are negative over the phenyl rings ($-17 \text{ kcal mol}^{-1}$) indicating that the $\text{C-H}\cdots\pi$ interactions should be preferred. The MEP value at H10B ($+18 \text{ kcal mol}^{-1}$) is less positive than the H bonded to N2, indicating that the $\text{C10-H10B}\cdots\text{F3A}$ interactions are less favoured in the crystal structure.

The dimer shown in Fig. 4 was optimised and the interaction energies were also computed. Moreover, the distribution of bond critical points (BCPs) and bond paths by using the AIM approach were also obtained. A bond path combined with a BCP connecting two atoms is a clear indication of an interaction.^[46] In accordance with the crystal structure of $[\text{Cu}(\text{DZP})_2] \cdot 2\text{H}_2\text{O}$, the strongest interactions $\text{N2-H}\cdots\text{O1}$ are responsible for the crystal packing of molecules. AIM analysis shows a complex set of molecular interactions that are outlined in Table 3 and the molecular graphs showing the intermolecular interactions are shown in Fig. S16 (Supplementary Material). The results indicate that the values of electron density and the Laplacian of electron density at the BCP for all intermolecular contacts (criteria range: 0.002–0.035 a.u. and 0.024–0.139 a.u. for ρ and $\nabla^2(\rho)$, respectively) confirm the existence of every intermolecular interaction shown in Fig. S16 (Supplementary Material). The energies of the intermolecular interactions at the BCP have been calculated according to the formula proposed by Espinosa et al.: $E(\text{H}\cdots\text{A}) = 0.5V$, where V is the electron potential energy density at the BCP.^[60] It is well known that the values of electron density and its Laplacian can be used as a measure of the strength of the interaction.^[61] The strongest interactions are related to the highest values of electron density at the BCP. In accordance with the results reported in Table 3, the $\text{N3-H}\cdots\text{O1}$ hydrogen bond is the strongest, with an interaction energy value of $-4.895 \text{ kcal mol}^{-1}$ and seems to be responsible for the dimer formation. The self-assembled dimer in $[\text{Cu}(\text{DZP})_2]$ is also stabilised by a combination of weak hydrogen bonds ($\text{C-H}\cdots\text{F}$) and $\text{C-H}\cdots\pi$ interactions. Based on results listed in Table 3, the $\text{C10-H10B}\cdots\text{F3A}$ interaction has

Table 3. Characteristics of the bond critical points (BCPs) of the intermolecular interactions in Cu-DZP

ρ , electron density; $\nabla^2(\rho)$, Laplacian of the electron density; G , electron kinetic energy density; H , electron energy density; V , electron potential energy density; $E(\text{H}\cdots\text{A})$, interaction energy

Interaction	$\text{H}\cdots\text{A}$ [Å]	ρ	$\nabla^2(\rho)$	G [a.u.]	H [a.u.]	V [a.u.]	$E(\text{H}\cdots\text{A})$ [kcal mol ⁻¹]
$\text{N2-H}\cdots\text{O1}$	2.111	0.01880	0.06356	0.01527	-0.00062	-0.01589	-4.895
$\text{C4-H4}\cdots\text{F1A}$	2.590	0.00536	0.02652	0.00538	-0.00125	-0.00663	-2.081
$\text{C10-H10B}\cdots\text{F3A}$	2.660	0.00358	0.01955	0.00360	-0.00129	-0.00489	-1.533
$\text{C10-H10B}\cdots\text{F2A}$	2.844	0.00491	0.02180	0.00444	-0.00097	-0.00540	-1.696
$\text{C10-H10A}\cdots\text{Cg1}$	3.047	0.00558	0.01800	0.00360	-0.00090	-0.00451	-1.414
$\text{C11-H11A}\cdots\text{Cg2}$	3.875	0.00305	0.00896	0.00176	-0.00048	-0.00224	-0.703
$\text{H3}\cdots\text{H8}$	2.863	0.00160	0.00628	0.00110	-0.00047	-0.00157	-0.492

a lower electron density and interaction energy than the C4–H4...F1A.

AIM analysis was used to provide additional evidence for the existence of C–H... chelate ring (CR) interactions. Fig. S16 (Supplementary Material) shows the AIM plot of the [Cu-(DZP)₂] complex wherein it can be observed that each C–H... CR interaction between the chelate ring and the H10A from the DZP ring methylene group is characterised by a bond critical point that connects the H10A atom to the chelate ring carbon atom formed by N1/C7/C1/C2/O1/Cu atoms, thus confirming the interaction. The high value of electron density and the interaction energy (−1.414 kcal mol^{−1}) evidence that these interactions are stronger than conventional C–H...π interactions. These results are similar to those reported by Maity and co-workers.^[29] The supramolecular assembly of the complex is also governed by C–H...π interactions involving the centroid (Cg2) and the H11A from the methylene group of the diazepine ring and its existence was confirmed by AIM analysis with a C11–H11A...Cg2 interaction energy of −0.703 kcal mol^{−1}. In addition, there is a H3...H8 intermolecular interaction observed in the molecular pair ($E_{H...A} = -0.492$ kcal mol^{−1}) which also provides additional stabilisation of the crystal structure.

Biological Activities

Inhibition of Growth and Biofilm Formation

It is known that the formation of metal complexes can modify the activity of the ligand. In this work, the four complexes were tested against a model of Gram negative and Gram positive bacteria. The results showed that the antibacterial effect (growth inhibition) of the compounds against *Pseudomonas aeruginosa* PAO1 was less than 25 % with respect to the growth control (Fig. 7). In general, Gram-negative bacteria are less sensitive to antimicrobials due to their external lipopolysaccharide membrane, which restricts the diffusion of hydrophobic compounds.^[62] In addition, *P. aeruginosa* has various active expulsion systems to eliminate antimicrobials that reach inside the microorganism.^[63]

The antibiofilm properties were higher in [Co(DZP)₂(-H₂O)₂].2H₂O, [Cu(DZP)]₂.2H₂O, or [Mn(DZP)₃].2H₂O complexes compared with those observed for the free ligand (HDZP). In these cases, the antibiofilm activity was higher than the antimicrobial activity. The [Mn(DZP)₃].2H₂O complex

produced the higher inhibition of the biofilm (40 and 31 %) at both concentrations of 100 and 10 μg mL^{−1}, respectively (see Fig. 7).

With respect to the Gram positive bacteria, *Staphylococcus aureus* ATCC 6538, the presence of Co-DZP at 100 μg mL^{−1} inhibited 90 % of the bacterial growth (Fig. 8). In order to determine the involved action mechanism, the MIC and MBC were determined. The MIC was 128 μg mL^{−1} and the MBC was 512 μg mL^{−1}, indicating a good potentiality of the complex as an antimicrobial. These results are in concordance with recent studies that demonstrated the bactericidal property of metallo-surfactants derived from cobalt oxide against several strains of *Staphylococcus aureus*.^[64] The author reported an important activity of several complexes present in nanoparticles at low concentration (near to 0.1 mM).

With respect to biofilm inhibition, it was higher than the growth inhibition in almost all complexes, with the exception of Cu-DZP. In presence of 100 μg mL^{−1} of Ni-DZP, Mn-DZP, and Co-DZP, the *S. aureus* biofilm development decreased 30, 55, and 99 % respectively. In addition, at 10 μg mL^{−1}, Co-DZP inhibits 45 % of the biofilm formation (Fig. 8).

In previous studies it was determined that the Cu²⁺ strongly influences the biofilm composition of *S. aureus*, in different ways, mainly by altering the structure of the exopolysaccharides that form the biofilm.^[65] The stress response to the complexes containing Cu^{II} could be explained by the higher production of biofilm.

The increase in the biofilm production observed with Mn-DZP at lower concentrations and with Cu-DZP at lower and higher concentrations can be correlated with previous studies with several compounds against *S. aureus*.^[66,67] The increase in biofilm production might be part of a defence mechanism against the complexes. However, this mechanism is usually inefficient when the concentration is increased.^[66,68] In particular, copper interferes with exopolysaccharides altering the absorption bands and hence the composition, possibly increasing the biomass by stress.

It is well documented that the presence of free bivalent cations such as Cu^{II}, Ni^{II}, Co^{II}, or Zn^{II} improves adhesion between biofilm cells.^[69] For this reason, a complex with a partial lipophilic character allows its insertion with the respective incorporated cations and interferes with the development and formation of the biofilm. Chohan and co-workers have

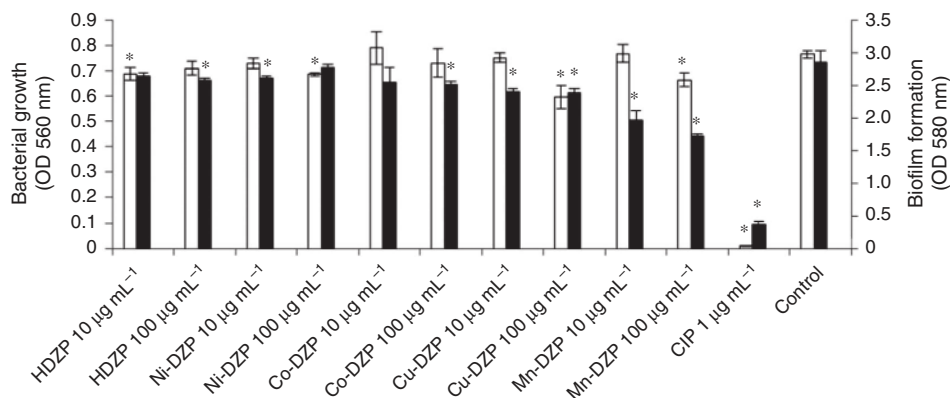


Fig. 7. Growth (□) and biofilm formation (■) by *P. aeruginosa* PAO1 cultures grown in LB broth in absence and presence of 10 and 100 μg/mL of HDZP, [Ni(DZP)₂] (Ni-DZP), [Co(DZP)₂(H₂O)₂].2H₂O (Co-DZP), [Cu(DZP)₂].2H₂O (Cu-DZP), and [Mn(DZP)₃].2H₂O (Mn-DZP). The inhibitor control was CIP (Ciprofloxacin) 1.0 μg mL^{−1}. The error bars indicate the standard deviation ($n = 4$). An asterisk was used to indicate significant differences ($P < 0.05$).

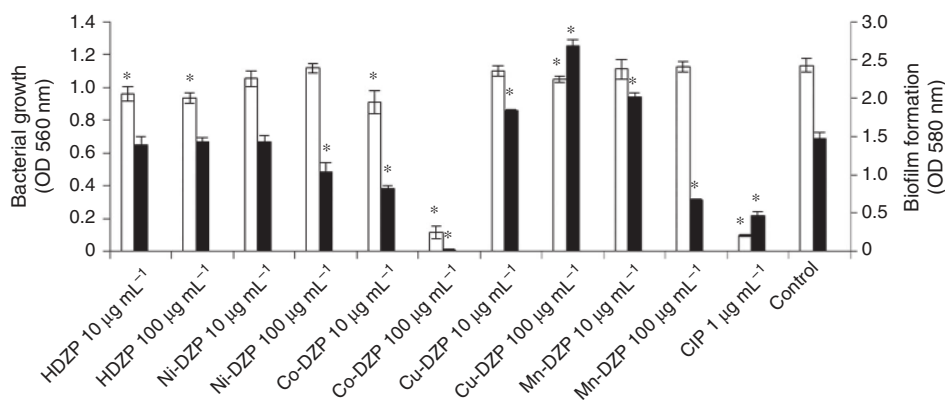


Fig. 8. Growth (□) and biofilm formation (■) by *S. aureus* ATCC 6538 cultures grown in MH broth in absence and presence of 10 and 100 µg mL⁻¹ of HDZP, [Ni(DZP)₂] (Ni-DZP), [Co(DZP)₂(H₂O)₂·2H₂O (Co-DZP), [Cu(DZP)₂·2H₂O (Cu-DZP), and [Mn(DZP)₃·2H₂O (Mn-DZP). The inhibitor control was CIP (Ciprofloxacin) 1.0 µg mL⁻¹. The error bars indicate the standard deviation ($n = 4$). An asterisk was used to indicate significant differences ($P < 0.05$).

reported that the polarity of the metal complexes is generally lower than those of the ligands. This effect is related to the greater electron delocalisation which improves the lipophilic character with respect to the ligand. This electronic effect allows a better diffusion of complexes in the lipid membrane and therefore increases its biological activities.^[70] Previous studies also suggest that numerous metal chelates exhibit slightly higher antibacterial activity than free ligands. This effect is clearly explained by Overton's concept and the Tweedy chelation theory.^[71] In addition, it is well known that some transition metals can act as active centres in bioactive molecules present in living systems.

According to the present results, [Co(DZP)₂(H₂O)₂·2H₂O is a good candidate to use against *S. aureus*, because this compound inhibits the microorganism growth and decreases significantly the main bacterial resistant mechanism, biofilm.

Conclusions

In this work, the synthesis, spectroscopic, and structural characterisation of metal complexes with (6Z)-6-(7-trifluoromethyl-1,2,3,4-tetrahydro-5H-1,4-diazepin-5-ylidene)cyclohexa-2,4-dien-1-one (HDZP) as ligand is reported. The crystal structure of the [Cu(DZP)₂] complex was determined by X-ray diffraction methods. The crystal packing of the complex is stabilised by classical N–H···O hydrogen bonds and the solid state structure exhibits interesting chelate ring (CR)–H and C–H···π interactions. The nature of these intermolecular interactions has been studied through DFT calculations and Hirshfeld surface analysis. The results described herein might be useful to understand the solid state architecture of different organic–inorganic hybrid materials containing Cu^{II}-chelate rings and aromatic molecules as ligands. The growth inhibitory activity and anti-biofilm activity against *S. aureus* and *P. aeruginosa* were evaluated for the ligand HDZP and its metal complexes. The [Co(DZP)₂(H₂O)₂·2H₂O complex is a good candidate to use against *Staphylococcus aureus* since in its presence both bacterial growth and biofilm formation decreases significantly.

Supplementary Material

IR spectra of the ligand HDZP and its Cu^{II}, Ni^{II}, Co^{II} and Mn^{III} complexes (Figs S1–S4); electronic spectra of the complexes (Fig. S5); Hirshfeld surfaces mapped over shape index and curvedness properties (Figs S7 and S8); crystal packing of the

Cu^{II} complex (Figs S6, S9 and S10); TGA curves for the thermal decomposition of the complexes (Figs S11–S14); a proposed molecular structure of Co^{II} and Mn^{III} complexes (Fig. S15); distribution of bond critical points in a dimer of the Cu^{II} complex (Fig. S16); vibrational data for the ligand and complexes (Table S1); absorption bands with the corresponding assignment (Table S2); and hydrogen bond geometrical parameters of the Cu^{II} complex (Table S3) are available on the Journal's website.

Conflicts of Interest

The authors declare no conflicts of interest.

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