

INVITED REVIEW

The slow relaxation mode: from solutions to gel networks

Junfang Li¹, To Ngai² and Chi Wu^{2,3}

In the past, many laser light-scattering experimental results revealed that besides the fast relaxation mode, there existed an additional slow mode in semidilute solutions. This slow mode has been assigned to a variety of origins, but there has been no clear and well-accepted explanation. As the polymer concentration increases, the slow relaxation mode persists in the concentrated region, in melts and in gels in which polymer chains are crosslinked instead of entangled. The slow relaxation mode has also been reported for charged macromolecules in aqueous and nonaqueous solutions. However, it is generally thought to be different in nature from that observed in semidilute neutral polymer solution. In recent years, armed with novel solution preparation methods and some specially designed polymers, we have reexamined the dynamics of polymer chains, especially the slow mode, in semidilute neutral polymer solutions, dilute polyelectrolyte solutions and gels, which are reviewed here. Our results suggest that the slow mode can be qualitatively considered as hindered motions of interacting chains even though the nature of interaction can be very different; namely, from the weak segment–segment interaction in a less good solvent to strong electrostatic interaction among polyelectrolyte chains, and even to chemical crosslinking inside gel networks.

Polymer Journal (2010) 42, 609–625; doi:10.1038/pj.2010.59

Keywords: anionic polymerization; laser light scattering; polyelectrolyte; polymer interaction; semidilute; slow relaxation mode; sol–gel transition

INTRODUCTION

The dynamics of chains in solutions and gels is an old subject of polymer science. Although the behaviors of individual isolated polymer chains in dilute solution under the θ - and good solvent condition are fairly understood,^{1–27} the slow dynamics of polymer chains in semidilute, concentrated solutions as well as in gel networks still remains a challenging problem since the 1980s. Because of the existence of overlapping and entanglement of polymer chains, new dynamics processes such as interchains interactions and disentanglement come into the problem. Many theoretical and experimental works have been carried out to clarify the properties of these systems.^{28–75} Among these, dynamics and relaxation processes investigated by dynamic laser light scattering (LLS) are of particular interest. In the past, many dynamic LLS experimental results revealed that besides the fast relaxation mode, there existed an additional slow relaxation mode in semidilute solution. It is a widely held view that the fast mode is related to the cooperative diffusion of chain segments between two neighboring entanglement points ('blob'), whereas the slow mode has been assigned to a variety of origins, such as reptation of a clustering of polymer chains through the entangled coils, the scattering vector (q)-independent relaxation of a transient network,

the q^2 -dependent translational diffusion of large aggregates or even dust particles. Therefore, a clear understanding with regard to the origins of the slow relaxation mode is still lacking. Note that as the polymer concentration increases, the slow relaxation generally persists in concentrated solutions, melts as well as gels in which polymer chains are crosslinked instead of entangled.

The slow relaxation mode has also been commonly reported for nearly all charged macromolecules, including both synthetic and biological polyelectrolytes in aqueous and nonaqueous solutions. The interpretation of the slow mode, especially for those very slow relaxation modes observed in salt-free or low-salt polyelectrolyte solutions, is very controversial.^{76–95} It has been attributed to large multichain domains formed because of electrostatic interaction or some insoluble clusters or even a trace amount of dust particles introduced during the imperfect preparation of polymer solution. Actually, not everyone accepts or recognized such a slow mode, even though it has been repeatedly observed in many dynamic LLS experiments for more than three decades, because of some problems or questions related to previous light-scattering experiments, such as some earlier premature data analysis methods and the preparation of dust-free viscous solutions. As a general phenomenon, which still

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Received 28 April 2010; revised and accepted 4 June 2010

remains an important controversial problem in polymer science, it certainly deserves a careful reexamination so that we can find its physical nature. Moreover, advancements of LLS instrumentation, including the full digital time correlator and computer, have made the dynamic study of slow relaxation mode much easier and reliable.

In recent years, by designing, synthesizing and using some novel polymers or sample preparation methods, we have reexamined the dynamics of slow relaxation mode in polyelectrolyte, semidilute solutions and gelling systems by a combination of static and dynamic LLS.^{75,96–107} In this article, we mainly review what has been carried out in our laboratory during the past 10 years. We will outline the details of each system in the following sections by starting with a brief introduction of LLS.

LLS

In a broad definition, LLS could be grouped as inelastic (for example, Raman, fluorescence and phosphorescence) and elastic (no adsorption) light scattering. However, in polymer and colloid science, light scattering is normally referred to in terms of static (elastic) or dynamic (quasi-elastic) measurements, or both.¹⁰⁸ Static LLS as a classic and absolute analytical method measures the time-average intensity and it has been long and widely used to characterize both synthetic and natural macromolecules.¹⁰⁹ On the other hand, dynamic LLS measures the intensity fluctuation. This is where the word dynamic comes from. The visibility of the scattering objects (macromolecules or colloidal particles) in LLS depends on the refractive index difference between the scattering object and the background ($dn=n-n_0$).

In the last two decades, owing to the advance of stable laser, ultra-fast electronics and personal computers, LLS, especially dynamic LLS, has evolved from a very special instrument for physicists and physical chemists to a routine analytical tool in polymer laboratories or even to a daily quality-control device in production lines. Commercially available research-grade LLS instruments nowadays are normally capable of making static and dynamic measurements simultaneously for studies of colloidal particles in suspensions or macromolecules in solutions as well as in gels and viscous media. In our laboratory, we have used a slightly modified commercial LLS spectrometer (ALV/DLS/SLS-5022f, ALV GmbH, Langen, Germany) equipped with a cylindrical 22-mW UNIPHASE He-Ne laser ($\lambda_0=632$ nm, JDS Uniphase Corp., Milpitas, CA, USA) as the light source and a sensitive Avalanche Photo Diode (APD) detector. The incident light beam was vertically polarized with respect to the scattering plane and the intensity was regulated with a beam attenuator (Newport Corp., Mountain View, CA, USA, M-925B) so as to avoid localized heating in the light-scattering cuvette. It is worth mentioning that many reviews, books, proceedings and chapters have been published on the details of LLS instrumentation and theory. Herein, we will concentrate only on the relevant details.

When a light beam is incident on a solution, and the solvent refractive index is different from that of the solute (macromolecules or colloidal particles), the incident light is scattered by each illuminated macromolecule or colloidal particle in all directions.^{9–11} If all scattering species are stationary, the scattered light intensity at each direction would be a constant, that is, independent of time. However, in reality, all the solute in solution are undergoing constant Brownian motions, which leads to both fluctuations of the scattered intensity pattern on the detection plane and the fluctuations of net scattered light intensity, $I(t)$, with time in sufficiently small detector area. The fluctuation rates can be related to different relaxation processes such as translational and rotational diffusions as well as internal motions of the macromolecules. The faster the relaxation process, the faster the intensity fluctuations will be. In the homodyne or self-beating system, the

intensity fluctuations can relate to the dynamic pair correlating of the scattering species:

$$\langle \delta I_q(0) \delta I_q(t) \rangle = |S(q, t)|^2 = \langle S(q) \rangle^2 \left[1 + \beta_{\text{app}} |g_{\text{solute}}^{(1)}(q, t)|^2 \right] \quad (1)$$

where $q = |\vec{q}| = \frac{4\pi n}{\lambda_0} \sin(\frac{\theta}{2})$ with n , λ_0 and θ are the solvent refractive index, the wavelength of the light in vacuum and scattering angle, respectively, β is the coherent factor depending on the detection optics, $S(q, t)$ is the dynamic scattering factor, $S(q)$ is the structure factor, $g^{(1)}(q, t)$ is the normalized field–field time correlation functions, defined as:

$$\langle E(0, q) E^*(t, q) \rangle / \langle E(0) E^*(0) \rangle, \text{ If } t \rightarrow \infty, \quad g^{(1)}(q, t) = 0, \text{ and}$$

$$S(q, t = \infty) = S(q) = \int \exp[i\vec{q} \cdot (\vec{r}_{\text{mi}} - \vec{r}_{\text{nj}})] dN d\vec{r}$$

For polymer in dilute solution without long-range pair interaction, the structure factor can relate to molar mass and form factor of the scattering particle by the follow equation:

$$S(q) \propto N_p N P(q) \propto C M (1 - 2C M A_2 + \dots) (1 - \frac{q^2}{3} R_g^2 + \dots) \quad (3)$$

where N_p , N , C , M , A_2 and R_g are the number of polymer chain, the number of monomer in each polymer, the concentration, molar mass of polymer, the second virial coefficient and the radius of gyration, respectively. In static LLS, the excess absolute time-averaged scattered light intensity, known as the excess Rayleigh ratio $R_{\text{v}}(q)$, of a dilute polymer solution at concentration C (g ml^{-1}) is related to the weight average molar mass M_w , the root-mean square z -average radius of gyration $\langle R_g^2 \rangle_z^{1/2}$ (or written as $\langle R_g \rangle$), and the scattering vector q as:

$$\frac{KC}{R_{\text{v}}(q)} \approx \frac{1}{M_w} (1 + \frac{1}{3} \langle R_g^2 \rangle q^2 - \dots) + 2A_2 C - \dots \quad (4)$$

where $K=4\pi^2 n^2 (dn/dC)^2 / (N_A \lambda_0^4)$ with N_A and dn/dC being the Avogadro number and the specific refractive index increment, respectively. This is the most basic equation in static LLS. With $R_{\text{v}}(q)$ measured over a series of C and q , one can obtain $\langle R_g \rangle$ and A_2 , respectively, from the slopes of $[KC/R_{\text{v}}(q)]_{C \rightarrow 0}$ vs q^2 and $[KC/R_{\text{v}}(q)]_{q \rightarrow 0}$ vs C ; and M_w from $[KC/R_{\text{v}}(q)]_{C \rightarrow 0, q \rightarrow 0}$. The Zimm plot, that is, $KC/R_{\text{v}}(q)$ vs $(q^2 + kC)$ with k being an adjustable constant, allows both q and C extrapolations to be made on a single grid.

In dynamic LLS, the intensity–intensity time auto-correlation function $G^{(2)}(q, t)$, defined as $\langle I(q, 0) I(q, t) \rangle / \langle I(q) \rangle^2$, in the homodyne mode or self-beating mode was measured, where t is the delay time and $\langle I(q) \rangle$ is the time-average scattering intensity, that is, the measured baseline. $G^{(2)}(q, t)$ is related to $g^{(1)}(q, t)$ by the Siegert relation as:

$$G^{(2)}(q, t) = A \left[1 + \beta |g^{(1)}(q, t)|^2 \right] \quad (5)$$

where A is the measured baseline. For monodispersed spherical scatters, $|g^{(1)}(q, t)|$ is theoretically represented as an exponential decay function:

$$|g^{(1)}(q, t)| = G \exp(-\Gamma t) \quad (6)$$

where G and Γ are the factor of proportionality and the line width, respectively, and $\Gamma=1/\tau$, the characteristic decay time representing the rate of dynamic relaxation. For a polydispersed dynamic relaxation, equation (6) may be generalized as:

$$|g^{(1)}(q, t)| = \int_0^\infty G(\Gamma) e^{-\Gamma t} d\Gamma \quad (7)$$

where $G(\Gamma)$ is called the line-width distribution and $G(\Gamma)d\Gamma$ is the statistic weight of the dynamic relaxation, which possess line width or related to the polymer distribution in dilute polymer solution.

The Laplace inversion of each measured $G^{(2)}(q,t)$ can lead to one $G(\tau)$ on the basis of equations 5 and 7. For a diffusive relaxation, $\Gamma(1/\tau)$ is related to the translational diffusion coefficient D by $(\Gamma/q^2)_{c \rightarrow 0, q \rightarrow 0} \rightarrow D$, so that $G(\Gamma)$ can be converted into a transitional diffusion coefficient distribution $G(D)$ or further a hydrodynamic radius distribution $f(R_h)$ applying the Stokes–Einstein equation, $R_h = (k_B T / 6\pi\eta) / D$, where k_B , T and η are the Boltzmann constant, the absolute temperature and the solvent viscosity, respectively. Note that η is local viscosity. In general, we cannot conclude that the slower mode should be caused by larger scattering object if the local viscosity is not uniform.

We should note that the Laplace inversion is an ill-conditioned problem because of the bandwidth limitation of photon correlation instruments, some unavoidable noises in the measured time correlation function, and a limited number of data points. Sometimes we used a combination of two stretch exponential functions to fit the correlation function:

$$\left[G^{(2)}(q, t) - A \right] / A = \beta_{\text{apparent}} \left\{ A_f \exp[-(t/\tau_f)^{\delta_f}] + A_s \exp[-(t/\tau_s)^{\delta_s}] \right\}^2 \quad (8)$$

where $\beta_{\text{apparent}} \leq \beta$, an apparent coherent factor, A_f and A_s are intensity contributions of the fast and slow modes, respectively, and $A_f + A_s = 1$. The details of the data analysis could be referred to the articles.^{97–107,110,111}

The slow mode in semidilute neutral polymer solutions

When the concentration of polymer is higher than the overlap concentration (C^*), polymer chains start to ‘touch’ each other and the solution enters the semidilute regime. C^* has been differently defined as $3M/(4\pi N_A R_g^3)$, or $M/(2^{3/2} N_A R_g^3)$, and $[\eta]^{-1}$, where M , R_g , N_A and $[\eta]$ are the molar mass, the radius of gyration of polymer chains, the Avogadro constant and the intrinsic viscosity, respectively. The difference of C^* calculated based on these definitions could be as large as five times.^{112,113} On the other hand, at the entanglement concentration (C_e), polymer chains start to interwind with each other, where C_e can be estimated from the ratio of the chain length (L) and the entanglement length (L_e).¹¹⁴ C_e is higher than C^* . Note that for chains shorter than L_e , they are not able entangle with each other even at higher concentrations.

The mean-field theory can effectively describe some properties of polymer solutions over the whole concentration range, but fails to explain other static and dynamic properties because of strong intra- and interchain interaction among covalently bonded monomers. Using the concept of a ‘blob’, the reptation and tube model, de Gennes and colleagues developed some scaling laws to predict various properties of semidilute solutions.^{114–117}

In a transient gel model, there exists only one characteristic length or one dynamic process in semidilute solutions if the solvent is thermodynamically good or athermal. In other words, the thermal energy ($k_B T$) can only agitate one short segment of the chain so that its gravity center undergoes a random Brownian motion inside a confined volume. The static and dynamic correlation lengths (ξ_S and ξ_D) can be measured from the angular dependence of the average scattering intensity and the average cooperative diffusion coefficient (D_C), respectively. Previous studies are well represented by two master curves: $\xi_S \sim C^{-0.72 \pm 0.01}$ and $\xi_D \sim C^{-0.70 \pm 0.01}$, slightly deviated from $\xi \sim C^{-0.75}$ predicted in theories.^{28,112} For a distance longer than ξ , the

segment–segment interaction is ‘completely’ screened out in an athermal solvent.^{116,118,119}

Besides the fast relaxation mode, an additional slow relaxation mode has also been observed in the intensity–intensity time correlation functions of various semidilute polymer solutions.^{28–39} This slow relaxation mode was assigned to different physical origins, such as the reptation of the entire chain inside a ‘tube’ made of its surrounding chains,^{114,115,117} the scattering vector (q)-independent relaxation of a transient network,^{29,34,35} the q^2 -dependent translational diffusion of large aggregates or even dust particles,^{32,33,37,38,112} and internal motions of large transient chain clusters.⁹² Later, it was realized that the reptation is not observable in dynamic LLS because of its nature for homopolymer system.

Theoretically, Brochard and de Gennes⁴⁰ attributed D_C to the osmotic modulus (M_π), the elastic modulus (M_0) and the friction coefficient (ζ) as $D_C = (M_\pi + M_0) / \zeta C$. Adam and Delsanti^{34,35} allowed the transient gel network to relax and assigned the q -independent slow mode with a viscoelastic nature and a characteristic time (τ_r). On the other hand, Wang *et al.*^{41–46} related the slow relaxation to the viscoelasticity under the Θ condition from the osmotic pressure fluctuation without using the concept of a transient gel network. The cooperative diffusion and the viscoelastic effect are generally mixed, depending on the frequency and a coupling parameter (β). $\beta = (C/\rho)(V_S - V_P)/V_S$, proportional to the difference between partial specific volumes of polymer and solvent ($V_S - V_P$) and ρ is the solution density. The influence of longitudinal modulus on the concentration fluctuation leads to a broad distribution.^{44–46} If $\beta = 1$, there exists only one stress-relaxation mode. D_C becomes identical to that predicted in the transient gel network model.^{40,114,115}

Experiments around the Θ condition indicate that the solvent quality has a profound effect on the slow relaxation.^{48–54} Actually, not everyone accepts or recognizes such a slow mode, even though it has been repeatedly observed for more than three decades. In the last two decades, advancements of LLS instrumentation, including the full digital time correlator and computer, have made the dynamic study of semidilute solutions much easier and more reliable, but the preparation of dust-free viscous semidilute solutions still remains an experimental challenge.

Traditionally, viscous dust-free semidilute solutions are prepared by slow evaporation of solvent from a dust-free dilute solution (the increase of polymer concentration). Figure 1 shows how the polymer concentration (C) of four polystyrene (PS) standards increases with the evaporation time (t). The inset shows that the time-average excess scattering intensity ($\langle I \rangle_{q \rightarrow 0} = \langle I \rangle_{\text{solution}} - \langle I \rangle_{\text{solvent}}$) decreases as C increases, following a scaling of $\langle I \rangle_{q \rightarrow 0} \propto C^{-0.25 \pm 0.05}$. For a polymer solution, $\langle I \rangle_{q \rightarrow 0} \propto (k_B T / V) C / (\partial \pi / \partial C)$, where V is the volume and $\partial \pi / \partial C$ is the osmotic compressibility. In dilute solution, $\partial \pi / \partial C = N_A k_B T / M$ and then $\langle I \rangle_{q \rightarrow 0} \propto C$. In semidilute solutions, $\partial \pi / \partial C \propto C^{-5/4}$ so that $\langle I \rangle_{q \rightarrow 0} \propto C^{-1/4}$. The scaling in the inset agrees well with the prediction and shows that solvent evaporation gradually induces the dilute polymer solution to the semidilute regime. As expected, only one fast relaxation mode is observed over the entire concentration range, consistent with some of the previous studies.^{112,116} However, it is worth mentioning that the whole process lasted for about 1 year. Even using such a painful process, one still faces a concentration gradient problem, that is, the upper layer is likely more concentrated than the bottom layer.

On the basis of the definition of C^* , one can also change R_g or M to switch a solution from dilute to semidilute for a fixed polymer concentration. One interesting advantage of the *in situ* change of R_g or M or both inside an LLS cell is that the polymer solution can

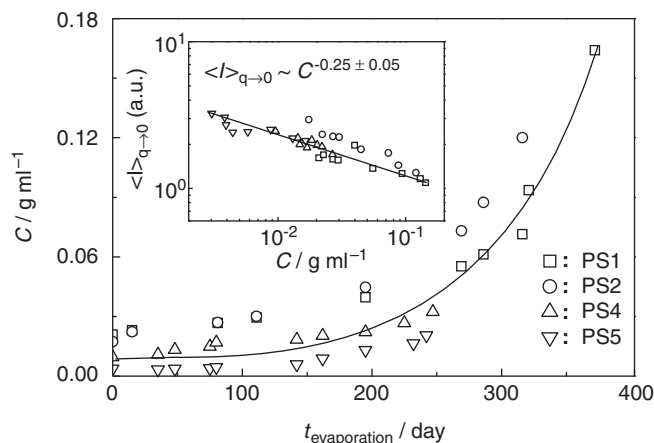


Figure 1 Evaporation time dependence of concentration (C) of four PS standards (PS1, PS2, PS4 and PS5 in Li *et al.*⁷⁵) in toluene. The inset shows concentration dependence of average scattering intensity ($\langle I \rangle_{q \rightarrow 0}$) as scattering vector (q) approaches zero. Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

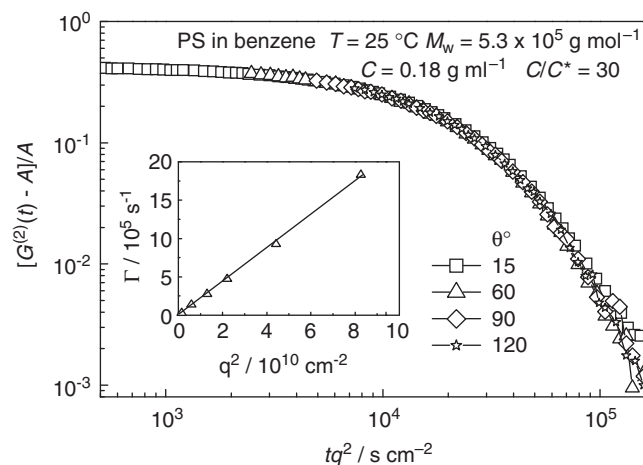


Figure 2 Scattering vector dependence of intensity–intensity time correlation functions of PS in benzene, where delay time (t) is scaled with q^2 . The inset shows scattering vector dependence of average line width. Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

completely avoid possible problems of dust, solubility and inhomogeneity. To achieve that, we used anionic living polymerization⁷⁵ to prepare dust-free semidilute/concentrated solutions (up to 30%) of narrowly distributed PS chains directly inside the LLS cell with a high vacuum stopcock, starting from a dust-free monomer solution. Subsequently, we comparatively reexamined the solution dynamics of PS chains in both good solvent (benzene) and a less good/ θ solvent (cyclohexane).

For the solution studied in Figure 2, we terminated living PS chains and measured its weight average molar mass (M_w) after the end of polymerization. We found that M_w is $\sim 5.3 \times 10^5 \text{ g mol}^{-1}$ and the polymer concentration is 0.18 g ml^{-1} , that is, $C/C^* \sim 30$. Importantly, Figure 2 shows that even with such a high concentration, there is still only one fast relaxation mode, agreeing well with the prediction in an athermal solvent. The scaling of the delay time (t) with q^2 makes the measured time correlation functions at different scattering angles collapse into one curve. The inset shows that the average line width

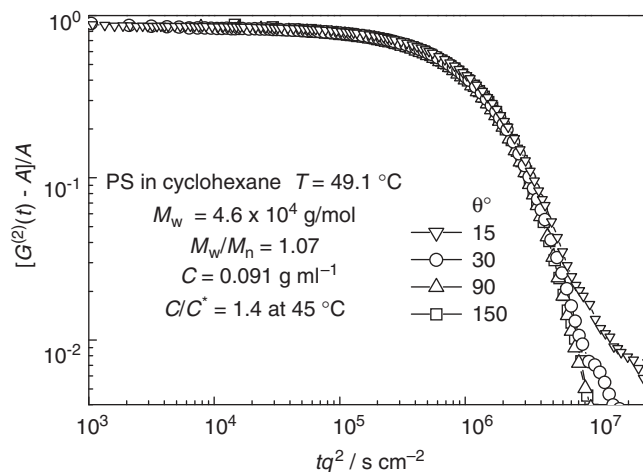


Figure 3 Scattering vector dependence of intensity–intensity time correlation functions of PS in cyclohexane, where delay time (t) is scaled with q^2 . Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

($\langle \Gamma \rangle_f = 1/\langle \tau \rangle_f$) is a linear function of q^2 , passing through the origin, a characteristic of diffusive relaxation. This fast relaxation mode could be modeled as the cooperative motions of the ‘blobs’ confined inside a ‘tube’ made of its surrounding chains. Such a fast relaxation mode is also independent of the solution chains temperature (T) after the scaling of the delay time (t) with T and the solvent viscosity, which is expected because benzene is an athermal solvent for PS in the temperature range studied.

With the same prepared method, we further studied the solution dynamics of PS in a less good/ θ solvent, cyclohexane. As expected, there is only one fast relaxation mode in cyclohexane when the solution is dilute, independent of the scattering angle. As soon as the solution becomes semidilute, that is, $C/C^* > 1$, an additional slow relaxation mode appears in the measured intensity–intensity time correlation function. The analysis of such a time correlation function on the basis of equation (8) leads to two characteristic relaxation times ($\langle \tau \rangle_f$ and $\langle \tau \rangle_s$). Figure 3 shows that the slow mode becomes more obvious at smaller scattering angles. Note that in LLS, the observation length is directly proportional to $1/q$, that is, large objects are more obvious at smaller scattering angles. Also note that here C/C^* is only ~ 1.4 at 45°C and $C/C_c \sim 0.1$ so that most of the chains are not entangled with each other. In a less good solvent, those entangled chains behave like large transient clusters within the delay time window in dynamic LLS due to relatively strong segment–segment interaction in comparison with in an athermal solvent. Moreover, we found that the characteristic line width ($\langle \Gamma_s \rangle$) is scaled to q as $\langle \Gamma_s \rangle \sim q^3$, indicating that the slow mode reflects some internal motions of large scattering objects. The concentration dependence of the scaling exponent (α_s) between $\langle \Gamma_s \rangle$ and q will be discussed more in the follow section.

As the solvent quality also affects the solution dynamics, we studied the temperature dependence of the slow relaxation mode in cyclohexane. Figure 4 shows typical plots of the temperature dependence of intensity–intensity correlation function for the PS in cyclohexane, and Figure 5 shows the corresponding plots of characteristic relaxation time distribution function analyzed by the CONTIN method. It clearly reveals that at a low scattering angle, the decreases of the solution temperature makes the slow mode less obvious and results in a decrease of its intensity contribution (the peak area). At 37°C , the

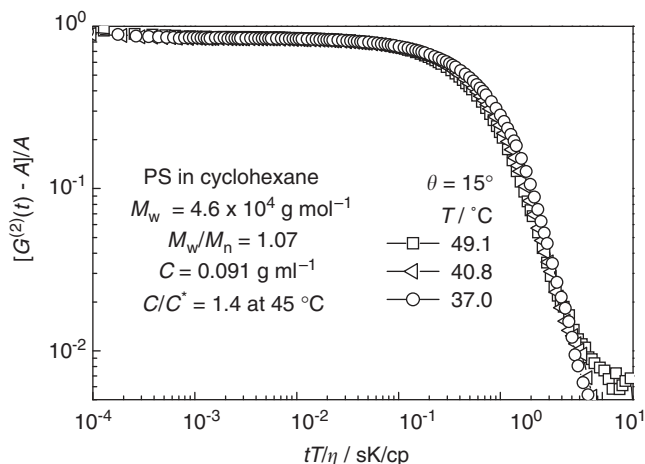


Figure 4 Temperature dependence of intensity–intensity time correlation functions of PS in cyclohexane, where delay time (t) is scaled with solution temperature (T) and solvent viscosity (η). Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

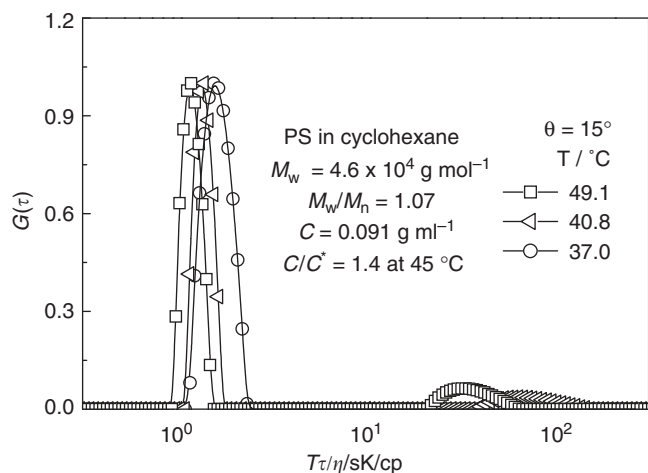


Figure 5 Temperature dependence of characteristic relaxation time distributions $G(\tau)$ of PS in cyclohexane, where characteristic relaxation time (τ) is scaled with solution temperature (T) and solvent viscosity (η). Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

slow mode completely disappears in the whole angular range and only one fast mode is observable. As discussed before, the decrease of the solution temperature makes the solvent quality poorer and the segment–segment interaction stronger so that PS chains contract as the temperature decreases, resulting in an increase of C^* , which can switch a solution from semidilute to dilute for a given C , that is, from $C/C^* > 1$ to $C/C^* < 1$ when C/C^* is not higher at higher temperature.

Figure 6 shows that for a more concentrated solution, that is, $C/C^* = 12$ or more precisely $C/C_e > 1$, the fast mode is collapsed together when the decay time (t) is scaled by q^2 , revealing its diffusive nature. In contrast with Figure 3, here the slow mode becomes more obvious at larger angles, which is different from those observed in the above lower concentration region. It implies that the slow mode is not related to some large scattering objects. Further increase of the polymer concentration makes the slow mode even visible at low scattering angles, as shown in Figure 7. Here, $C/C^* > 50$ and the solution appears like a gel without any visible flow within a short time, but the solution is homogeneous and its scattering intensity is

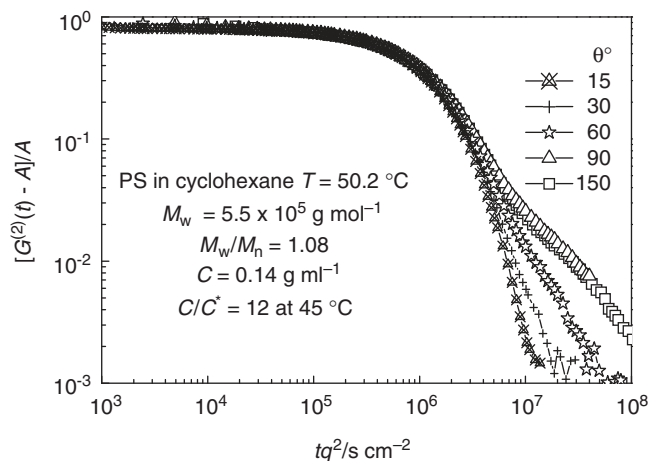


Figure 6 Scattering vector dependence of intensity–intensity time correlation functions of PS in cyclohexane, where delay time (t) is scaled with q^2 . Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

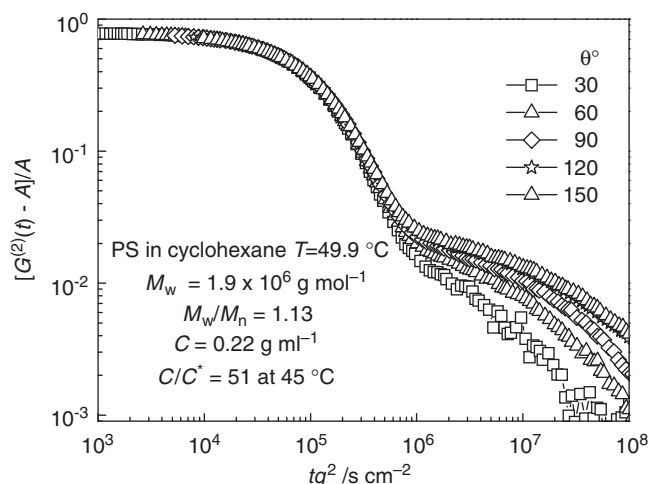


Figure 7 Scattering vector dependence of intensity–intensity time correlation functions of PS in cyclohexane (Solution 10 in Li *et al.*⁷⁵), where delay time (t) is scaled with q^2 . Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

independent of the position of the scattering volume inside the solution. Figure 7, once again shows that the fast mode collapses together after delay time (t) is scaled by q^2 , indicating that it remains diffusive. At the same time, Figure 8 shows that for a given scattering angle, the slow mode become slower and its intensity contribution (A_s) increases as the solution temperature decreases. Note that the decrease of the solution temperature has two opposite effects on the slow mode. On the one hand, it increases the segment–segment interaction and enhances the slow mode; on the other hand, the chain contraction disentangles the transient clusters. When the latter is dominant, the slow mode becomes weak or even disappears at lower solution temperatures. However, when C is much greater than C_e , the chain contraction cannot switch the solution from semidilute to dilute so that decreasing the solution temperature only increases the segment–segment interaction. This presumably is the reason why the slow mode becomes slower and more visible at lower temperatures as shown in Figure 8.

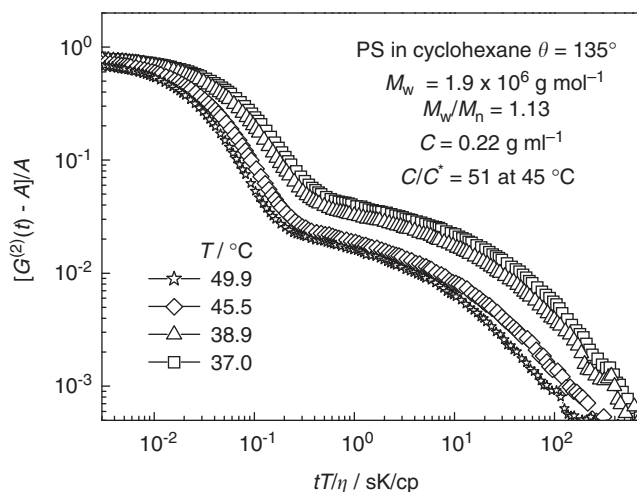


Figure 8 Temperature dependence of intensity-intensity time correlation functions of PS in cyclohexane, where delay time (t) is scaled with solution temperature (T) and solvent viscosity (η). Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

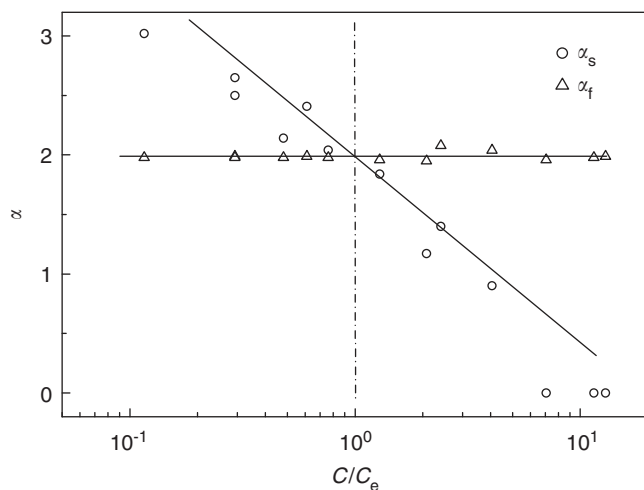
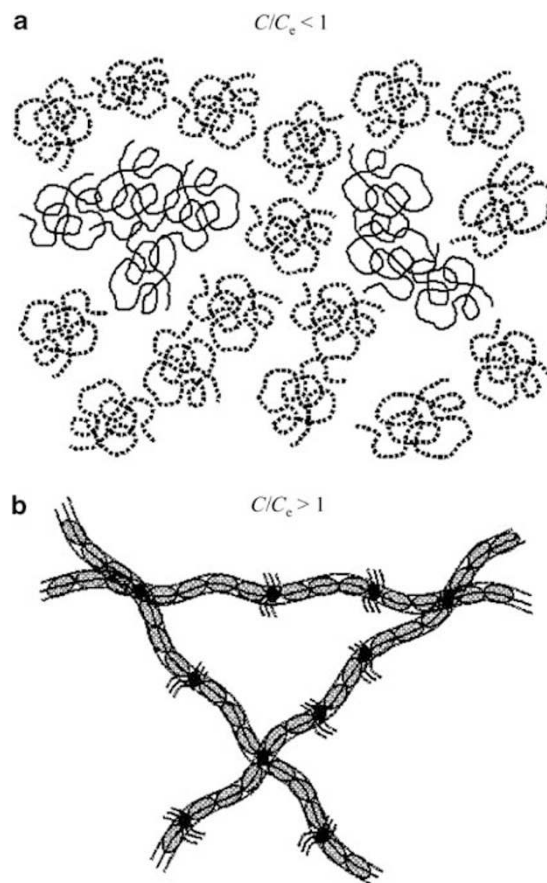


Figure 9 Polymer concentration dependence of two scaling exponents (α_f and α_s) in $\langle \Gamma \rangle_f \propto q^{2\alpha_f}$ and $\langle \Gamma \rangle_s \propto q^{2\alpha_s}$ for PS in cyclohexane (a good solvent at 45 °C). Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

Figure 9 summarizes the concentration dependence of the scaling exponents (α_f and α_s) in $\langle \Gamma \rangle_f \propto q^{2\alpha_f}$ and $\langle \Gamma \rangle_s \propto q^{2\alpha_s}$ for PS in cyclohexane solution. As expected, for the fast mode, $\langle \Gamma \rangle_f \propto q^2$ over the entire concentration range studied, reflecting the diffusive Brownian motion of the gravity center of each segment (blob). In contrast, for the slow mode, α_s decreases as C/C_e increases. We can roughly divide Figure 9 into two concentration regions (left and right) using $\alpha_s = 2$ as a reference (the dashed line). On the left ($C/C_e \leq 1$), the slow mode is more visible at smaller scattering angles, but in the right region ($C/C_e > 1$), its intensity contribution to $G(\tau)$ increases with the scattering angle. As mentioned before, the LLS theory tells us that larger objects scatter more light at a lower scattering angle. Therefore, in the left region, the slow mode should be related to scattering objects larger or comparable with $1/q$; whereas in the right region, it is related to scattering objects smaller than $1/q$. The LLS theory also tells us that



Scheme 1 Semidilute polymer solutions in a less good solvent in different concentration regions. (a) $C/C_e < 1$, where individual chains and ‘clusters’ have the same microenvironment, but different sizes; and (b) $C/C_e > 1$, where all the chains are overlapped and entangled with each other. Light and dark spheres represent ‘blobs’ between two entanglement points and around entanglement points, respectively. Reproduced with permission from Li *et al.*⁷⁵ Copyright (2008) American Chemical Society.

when the scaling exponent (α) between the characteristic line width ($\langle \Gamma \rangle$) and the scattering vector (q) is > 2 , we are observing a mixture of diffusion and internal motions of larger scattering objects. However, it is not clear why α is < 2 in a three-dimensional space.

In short, the comparative studies of different solvents, concentrations and temperatures convincingly demonstrated that the slow mode is real in semidilute solution. Previously, we speculated that the slow mode might be related to some long-range density fluctuation. The recent results indicate that for $C/C_e \leq 1$, the slow mode is related to some large structures (scattering objects or long-range density fluctuation), as it is more visible at lower scattering angles and the scaling exponent (α_s) is higher than 2. Such large structures could be attributed to a few transient clusters made of lightly entangled chains due to relatively stronger segment–segment interaction in a less good solvent. In this case, most of the chains are not entangled with each other and they coexist with those transient ‘clusters’, as schematically shown in Scheme 1a. Strictly speaking, these solutions are not semidilute but a transitional region between dilute and semidilute solutions. Our results reveal that one should use C_e instead of C^* to mark the real boundary of semidilute solutions. In the transitional region ($C^* \leq C \leq C_e$), it is these transient chain entanglements that make the solution microscopically inhomogeneous over a short time scale ($\sim 10^{-3}$ s).

When $C > C_e$, all the chains inside the solution entangle with each other to form one huge 'cluster'. Each chain could be visualized to arrest inside a 'tube' made of its surrounding chains within a short time and 'divided' into a number of segments (blobs) with a dimension limited by the tube diameter (the correlation length). Within each blob, the segment is excited by thermal energy ($k_B T$) so that its center of gravity undergoes a Brownian motion. These thermal blobs have a similar size. In an athermal solvent, there is no interaction between them and the tube so that they experience a similar microenvironment, resulting in only one (fast) relaxation mode. In a less good or θ solvent, the segment–segment interaction near the entanglement points is expected to be stronger. Therefore, the blobs near each entanglement point should move slower than those in the middle between two entanglement points; namely, the blobs experience two different microenvironments, as schematically shown in Scheme 1b. The increase of polymer concentration makes the tube thinner. Up to one point, the motions of the segments perpendicular to the tube near the entanglement point is so limited that the segments can only randomly oscillate in the tube. This might explain why its characteristic relaxation time becomes less dependent on or even independent of the scattering vector.

The temperature dependence studies also support our speculation about the nature of the slow mode. The decrease of the solution temperature makes the segment–segment interaction stronger so that each chain contracts. When $C/C_e < 1$, the chain contraction disentangles the transient clusters so that the slow mode becomes weak or even disappears at lower solution temperatures. However, when $C > C_e$, the chain contraction cannot switch the solution from semidilute to dilute so that decreasing the solution temperature only increases the segment–segment interaction. This is why the slow mode becomes slower and more visible at lower temperatures.

The slow mode in PNIPAM/H₂O system

The solution dynamics in poly(*N*-isopropylacrylamide) (PNIPAM)/H₂O system was selected and examined by dynamic LLS because of its thermal-responsive properties, which exhibits a lower critical solution temperature around 32 °C. In such a way, similar to the PS in cyclohexane, we are able to use temperature dependence of the solvent–polymer interaction to investigate effects of the shrinkage of polymer chains with increasing temperature, and consequently disentanglement of transient network, which in turn can prove the coexistence of individual chains and large transient network in semidilute solutions when $C/C_e < 1$.

For PNIPAM in semidilute aqueous solutions, Hirotsu and co-workers¹²⁰ found that the correlation function measured by dynamic LLS could be fitted to single exponential quite well through the whole temperature range investigated (25–40 °C). They claimed that the translational motion of polymer molecules is profoundly inhibited due to the relative immobility of densely packed coils and the intermolecular entanglement effect; the correlation function represents the relaxation of the internal or pseudogel mode. However, our recent result found that the single relaxation mode measured from dilute PNIPAM aqueous solution by dynamic LLS turns into one fast and one slow relaxation mode when the concentration enters into semidilute regime. Figure 10a shows typical plots of the intensity–intensity correlation function for the PNIPAM/H₂O system in both the dilute and semidilute regimes, and Figure 10b shows the corresponding plots of characteristic relaxation time distribution function analyzed by the CONTIN method. When $C < C^*$ ($C^* = 1.9\text{--}2.8 \times 10^{-3} \text{ g ml}^{-1}$ at 25 °C), a single relaxation mode is observed, which is related to the mutual diffusion of individual polymer chains. As the concentration increases

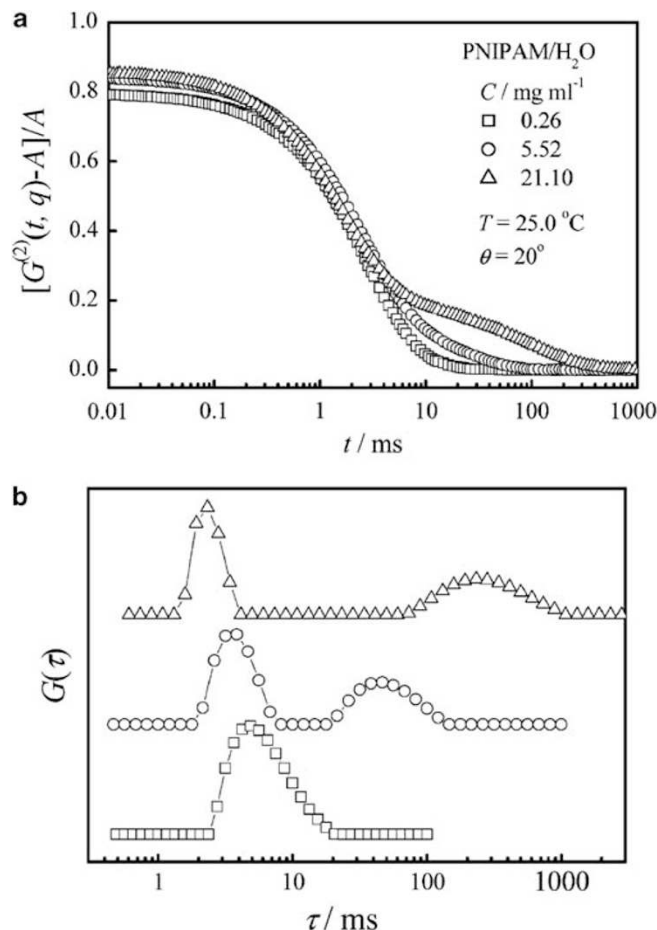


Figure 10 (a) Concentration dependence of intensity–intensity time correlation function $[G^{(2)}(t, q) - A]/A$ and (b) their corresponding characteristic relaxation time distribution function $G(\tau)$ of PNIPAM/H₂O in both dilute and semidilute regimes. Reproduced with permission from Yuen *et al.*⁹⁶ Copyright (2006) American Chemical Society.

and enters the semidilute regime, the chains begin to overlap and entangle, and new dynamical processes involving interchain interaction and disentanglement begin to occur besides the fast relaxation mode. The slow mode becomes more and more evident, and its characteristic relaxation time shifts toward slow direction with increasing concentrations.

Figure 11 shows the temperature dependence of characteristic relaxation time distribution function obtained from a semidilute solution. The data on the x axis is shifted by (T/η) so that the peak positions at different temperatures can be directly compared, where η is the solvent viscosity at the absolute temperature. It is clear that the fast mode shifts to the slower direction, whereas the slow mode becomes faster with increasing temperatures when $T < 30.0$ °C. The effects are just contrary to that of increasing concentrations. The solvent quality tends to be poor with increasing temperature, which will result in fewer entanglement points and, consequently, a loose network with more defects or larger structural fluctuation. However, at $T = 30.0$ °C, phase separation happens. After keeping the solution at this temperature for 32 h, the characteristic relaxation time distribution function turns out to be single narrow mode, which is q^2 dependent and related to collective diffusion of aggregates. The temperature dependence studied for PNIPAM/H₂O further supports

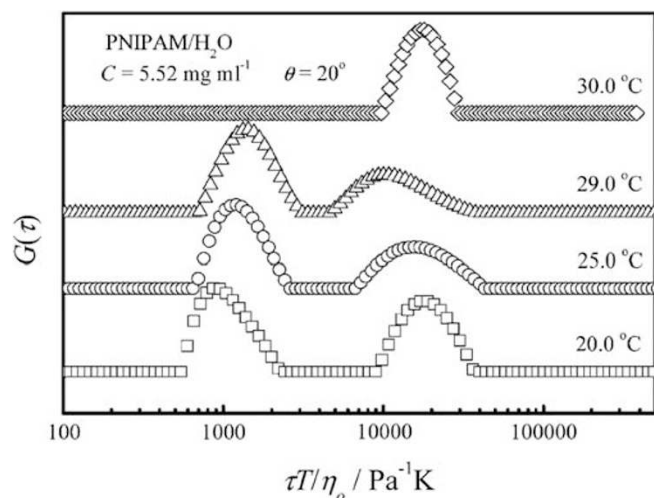


Figure 11 Temperature dependence of characteristic relaxation time distribution function $G(\tau)$ for a semidilute aqueous solution of PNIPAM. Reproduced with permission from Yuen *et al.*⁹⁶ Copyright (2006) American Chemical Society.

our speculation about the nature of the slow mode that for $C/C_c < 1$, it is related some large structures.

The slow mode in PEO-PPO-PEO/H₂O system

Block copolymers normally consist of a few chemically distinct short chains joined by covalent bonds. In a solvent selectively good only for one block, they can display more complicated structures and morphology, such as polymeric micelles, than homopolymers even in dilute solution.^{121,122} As for triblock copolymers in a selective solvent, the results become even more complicated because triblock chains can form flower-like micelles or interconnected clusters or even a gel network, depending on whether the solvent is selectively good for the middle or the two end blocks. In semidilute solution, because of the existence of overlapping and entanglement, new dynamical processes come into the additional problem. Therefore, in comparison with dynamics in homopolymers, the results of block copolymer chains in semidilute solution were more controversial and diversified.^{123–141}

The system of poly(ethylene oxide)-poly(propylene oxide)-poly(ethylene oxide) (PEO-PPO-PEO) in H₂O was also selected to investigate the solution dynamics because of the well-known fact that it can exhibit temperature-induced micellization phenomenon at a certain concentration range. Figure 12a shows typical plots of intensity–intensity correlation function for the PEO-PPO-PEO/H₂O system in both the dilute and semidilute regimes, and Figure 12b shows the corresponding plots of characteristic relaxation time distribution analyzed by the CONTIN method. The apparent coherence is lower for dilute solution in Figure 12a due to its weak scattering intensity. The same tendency with that in the PNIPAM/H₂O system can be seen as increasing concentrations. Here, $T=20\text{ °C}$ is below the critical micellization temperature of all the mentioned concentration according to Hatton *et al.*¹⁴² The slow mode is two orders of magnitude slower than the fast mode for PEO-PPO-PEO/H₂O at the concentration which is twice C^* .

For nonionic polymeric surfactants PEO-PPO-PEO triblock copolymer in aqueous solutions, the dehydration gradually becomes significant with increasing temperature and consequently results in a decrease of the size of copolymer chains. When the temperature is above the critical micellization temperature of a certain concentration, micelles will form. Figure 13 shows the temperature dependence of

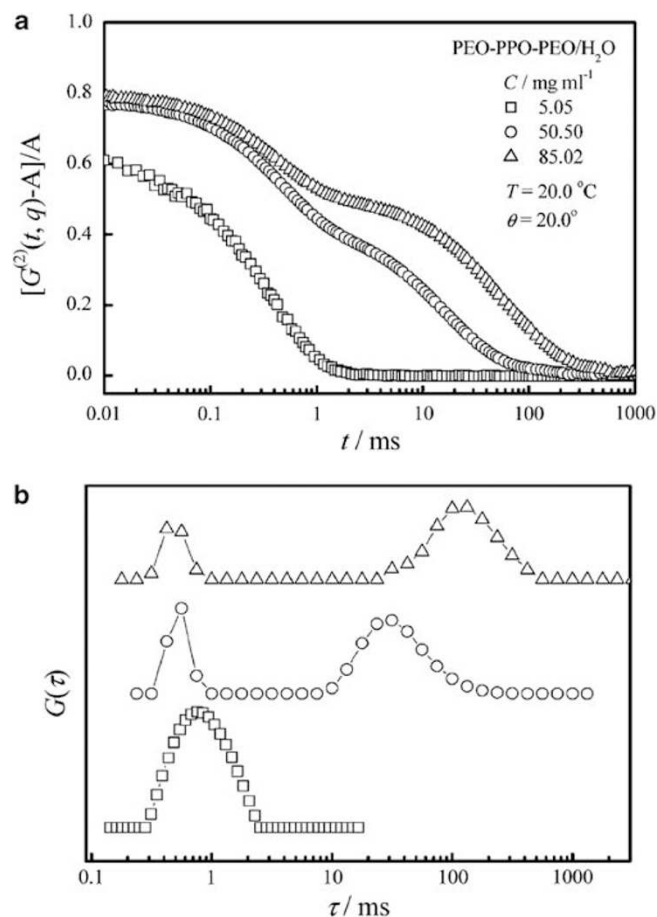


Figure 12 (a) Concentration dependence of intensity–intensity time correlation function $[G^{(2)}(t, q) - A]/A$ and (b) their corresponding characteristic relaxation time distribution function $G(\tau)$ of PEO-PPO-PEO/H₂O in both dilute and semidilute regimes. Reproduced with permission from Yuen *et al.*⁹⁶ Copyright (2006) American Chemical Society.

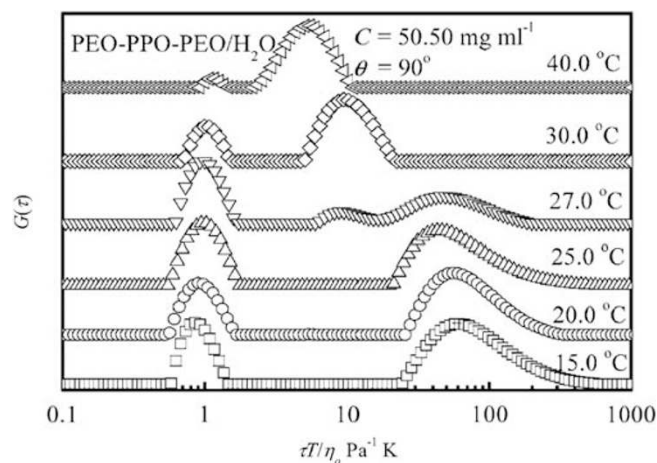


Figure 13 Temperature dependence of characteristic relaxation time distribution function $G(\tau)$ for a semidilute aqueous solution of PEO-PPO-PEO. Reproduced with permission from Yuen *et al.*⁹⁶ Copyright (2006) American Chemical Society.

characteristic relaxation time distribution function obtained from a semidilute solution. Below 25 °C, the same tendency of fast and slow relaxation mode changing with temperature was observed, compared

with that in PNIPAM system. With increasing temperature, it is clear that an intermediate mode appears accompanied by the disappearance of the slow mode. The intermediate mode is evident at high temperature, and its relaxation rate is q^2 dependent; therefore it can be interpreted as collective diffusion of micelles.

In summary, the special thermal-sensitive properties of PNIPAM/H₂O and PEO-PPO-PEO/H₂O systems enable us to follow the temperature-induced shrinkage of polymer chains and, consequently, disentanglement of transient network in semidilute solutions formed by overlapped or entangled polymer chains, without changing the overall concentrations. The difference of the nature of slow mode between these two systems is that aggregates are formed in the PNIPAM/H₂O system but micelles are formed in the PEO-PPO-PEO/H₂O systems when the solvent becomes poor.

Effect of crosslinking on dynamics during the sol–gel transition

Up to this point, our results clearly show that the slow relaxation observed in the semidilute solutions is real, not because of previously suggested artifacts in the solution preparation. The concentration and temperature dependence of the characteristic relaxation time of slow mode in different systems also suggest that such a slow mode has two different natures. When C is lower than the entanglement concentration (C_e), it can be related either to the large temporal clusters formed by some entangled chains. However, at higher entanglement concentration, the restricted interaction of polymer chains near the entanglement points likely results in the slow mode. If this is the case, one may wonder that *in situ* crosslinking of polymer chains in a semidilute solution can artificially increase the correlation between different entanglement points, consequently increasing the relative contribution of the slow mode as crosslinking proceeded.^{101–103}

To prove that, we synthesized a linear poly(methyl methacrylate-*co*-7-acryloyloxy-4-methylcoumarin) (P(MMA-*co*-AMC)) copolymer and studied the dynamics of such a copolymer in semidilute solutions by dynamic LLS. The advantage of using such a system is relied on the fact that two coumarin groups can undergo a (2+2) photodimerization when irradiated with an ultraviolet light (310 nm) as shown in Scheme 2. Such a dimerization reaction can be used to bind the copolymer chains together in chloroform in a completely controllable

manner so that we are able to investigate the effect of crosslinking on the fast and slow relaxation modes of semidilute solutions at each photo-reaction stage of the sol–gel transition.

Figure 14 shows that in the initial semidilute solution before the photo-crosslinking, $g^{(2)}(q, \tau)$ has two distinct relaxation modes. The fast relaxation corresponds to the cooperative diffusion of the subchains between two entangled points, which is well described by the ‘blob’ or scaling theory,^{116,117} whereas the controversial slow relaxation can be attributed to the relatively stronger segment–segment interaction near the entanglement points makes the congested polymer chain nonuniform.^{99,102,103} Previously, we also speculated that such slow mode might be related to some long-range density fluctuation.¹⁰² As the crosslinking proceeds, the slow relaxation becomes even slower, and at the same time, its contribution to the total scattering intensity increases, reflecting in the raise of its amplitude. It clearly shows that the crosslinking enhances the restricted movement of the subchains near the entanglement points, resulting in the slow mode. The flowing

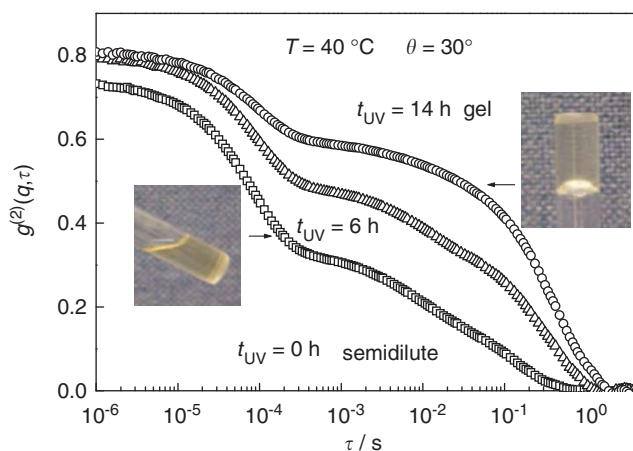
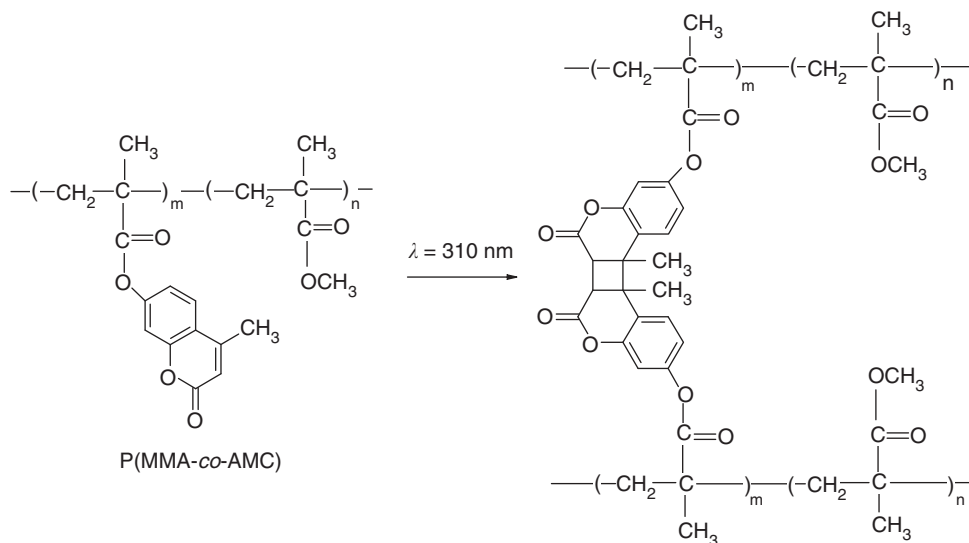


Figure 14 Normalized intensity–intensity time correlation functions of crosslinkable P(MMA-*co*-AMC) chains in chloroform during the sol–gel transition. Reproduced with permission from Ngai *et al.*¹⁰² Copyright (2004) American Chemical Society.



Scheme 2 Photodimerization-induced crosslinking of P(MMA-*co*-AMC) copolymer chains in chloroform. Reproduced with permission from Ngai *et al.*¹⁰³ Copyright (2003) American Chemical Society.

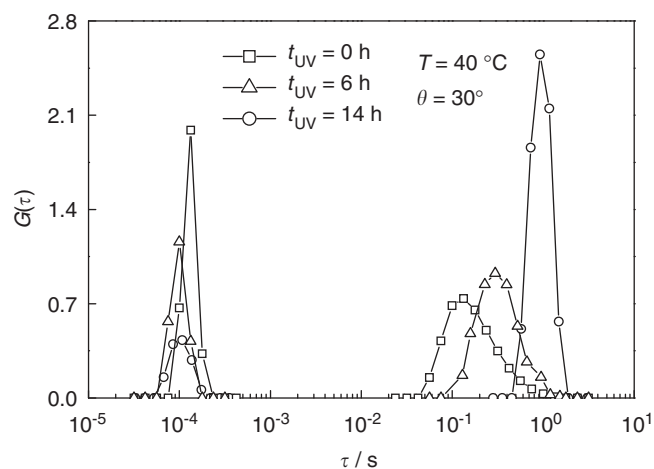


Figure 15 Peak area-normalized intensity distribution of characteristic decay time $f(\tau)$ of crosslinkable P(MMA-co-AMC) chains in chloroform during the sol-gel transition, where $f(\tau)$ is τ -weighted $G(\tau)$, that is, $f(\tau) \equiv \tau G(\tau)$, because the x axis is logarithmic. Reproduced with permission from Ngai *et al.*¹⁰² Copyright (2004) American Chemical Society.

semidilute solution finally becomes a macroscopically immobile chemical gel after ~ 9 h ultraviolet irradiation, as shown in the insets. Note that even in the gel state, $g^{(2)}(q, \tau)$ can still completely relax to zero even when the relaxation is fairly slow, indicating that there is no static frozen-in component inside. It is important to note that the complete relaxation of $g^{(2)}(q, \tau)$ is not due to an improper normalization.^{143,144} This is because here the intercept of $g^{(2)}(q, \tau)$ at $\tau \rightarrow 0$ increases during the sol-gel transition. The lack of static frozen-in components is supported by the sample-position-independent scattering intensity; for example, no speckle was detected when we randomly turned the LLS cell to let the laser hit different parts of the gel.

Figure 15 shows the variation of $G(\tau)$ from a semidilute solution to a gel network, where $\tau = 1/\Gamma$. During the sol-gel transition, the bimodal distribution is clearly preserved, although the relative contribution of the fast relaxation decreases. Shibayama and colleagues¹⁴⁵ also reported that the fast mode remained in the formation of silica gels. Another interesting point is that the position of the peak related to the fast relaxation nearly remains during the sol-gel transition, different from other gelling systems.^{127,146} The peak related to the slow relaxation progressively shifts to a longer delay time, that is, the crosslinking hinders motions of interacting polymer chains. In other words, the motion of the segments near the entanglement point (or crosslinked point) is so limited that the segments may only randomly oscillate. It should be addressed that the crosslinking mostly occurs at the entangled points because only two very nearby coumarin groups can undergo the (2+2) photocycloaddition. Therefore, the crosslinking should have a limited effect on the length of the subchains ('blobs'). This explains why the peak related to the fast relaxation remains its position.

Figure 16 shows the effect of temperature on the intensity weighting (relative contribution) of fast relaxation (A_f) during the sol-gel transition. It indicates that the intensity contribution of the A_f increases with the temperature, but the trend is quite different with and without the crosslinking. Note that $A_f + A_s = 1$ so that we only plot A_f vs T . The fast relaxation of the crosslinked chains contributes very little at lower temperatures. It should be noted that the overall time-averaged scattering intensity ($\langle I \rangle_T$) measured in static LLS is the sum of $\langle I \rangle_{\text{fast}}$ and $\langle I \rangle_{\text{slow}}$ and $\langle I \rangle_T$ increases ~ 180 times as the

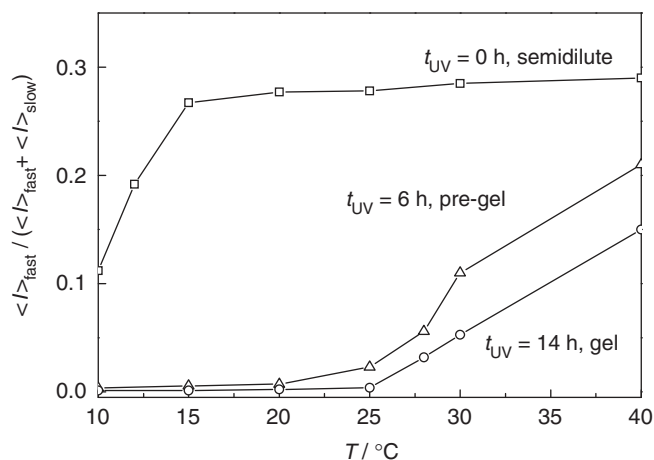


Figure 16 Temperature dependence of intensity weighting (relative contribution) of fast relaxation (A_f) of crosslinkable P(MMA-co-AMC) chains in chloroform during the sol-gel transition. Note that $A_f + A_s = 1$. Reproduced with permission from Ngai *et al.*¹⁰² Copyright (2004) American Chemical Society.

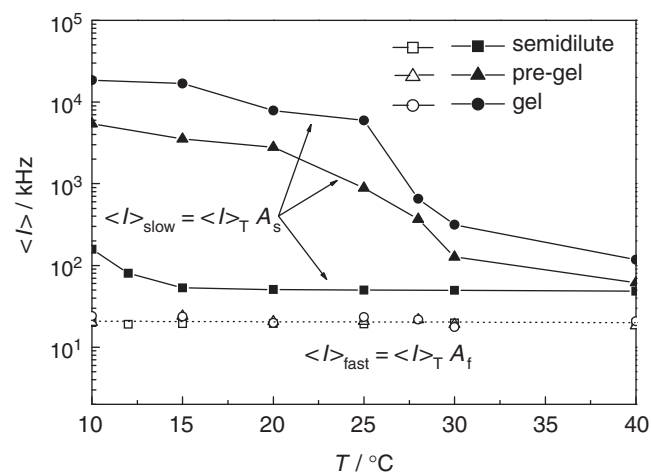


Figure 17 Temperature dependence of time-averaged scattering intensities ($\langle I \rangle_{\text{fast}}$ and $\langle I \rangle_{\text{slow}}$) related to fast and slow relaxation of crosslinkable P(MMA-co-AMC) chains in chloroform during the sol-gel transition, where $\langle I \rangle_T$ is overall time-averaged scattering intensity measured in static LLS. Reproduced with permission from Ngai *et al.*¹⁰² Copyright (2004) American Chemical Society.

temperature decreases. Therefore, it is necessary and right to calculate the absolute average scattering intensities ($\langle I \rangle_{\text{fast}}$ and $\langle I \rangle_{\text{slow}}$), respectively, for the fast and slow modes, instead of A_f and A_s . The results are shown in Figure 17. It is helpful to note that static and dynamic LLS measurements were carried out in a back-to-back manner on the same sample position under identical conditions and at the same q . On the basis of Figures 16 and 17, we know that the decrease of A_f is attributed to the increase of $\langle I \rangle_{\text{slow}}$. This is understandable because the fast relaxation is related to the motions of the subchains (blobs) between two entangled (crosslinked) points and the temperature should have no effect on the average length of the subchains and the refractive index contrast between the blobs and solvent, as the temperature dependence of the refractive index (dn/dT) is only in the order of 10^{-4} .

As for the increase of $\langle I \rangle_{\text{slow}}$, we should note that in semidilute solution polymer chains are highly entangled with each other. Inside the volume of one chain, there also exist many other chains, especially when the chain is long. Assuming that each chain with a radius of gyration (R_g) has n_s segments and there are n_p chains inside its occupied space ($\sim R_g^3$), the scattering intensity ($\langle I \rangle$) from the $n_s n_p$ segments on the basis of the LLS theory is related to the position of each segment (r) as follows,¹⁰⁸

$$\langle I \rangle_T \propto \left\langle \sum_{j=1}^{n_s n_p} \sum_{k=1}^{n_s n_p} \exp[i\vec{q} \cdot (\vec{r}_j - \vec{r}_k)] \right\rangle \quad (9)$$

If the motions of $n_s n_p$ segments are fully correlated, $\langle I \rangle_T \propto (n_s n_p)^2$ when $qR_g \ll 1$; although if the motions of $n_s n_p$ segments are completely independent, $\langle I \rangle_T \propto n_s n_p$. In reality, $\langle I \rangle_T$ is between $n_s n_p$ and $(n_s n_p)^2$ in semidilute solution because the segments between different chains are only partially correlated. The crosslinking makes the segments between different chains more correlated so that the scattering intensity increases. On the other hand, the increase of the solution viscosity at lower temperatures enhances the hydrodynamic interaction of the segments between different chains. This is why $\langle I \rangle_{\text{slow}}$ increases with a decreasing temperature.

The slow mode in salt-free or low-salt polyelectrolytes solution

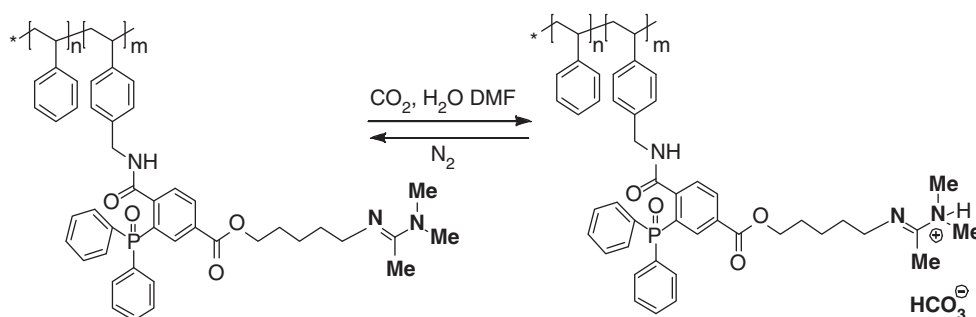
Dynamics of polyelectrolytes in salt-free or low-salt solutions has attracted much attention after Lin *et al.*¹¹¹ reported their dynamic LLS observation of an extremely slow diffusive mode in the low-salt poly(L-lysine) solutions. Later, it was shown that there are actually a fast and a slow relaxation mode in dynamic LLS measurements of salt-free or low-salt dilute polyelectrolyte solution in comparison with the translational diffusive relaxation mode of individual neutral chains with similar lengths. Some reported slow relaxation modes were really slow with a hydrodynamic size of 10^2 – 10^3 nm or larger. Gradually, such two modes have been reported for nearly all charged macromolecules, including both synthetic and biological polyelectrolytes in aqueous and nonaqueous solutions.^{76–84} Therefore, these two modes are general features of salt-free or low-salt polyelectrolyte solutions.

In the past, the fast mode in polymer dilute and semidilute solutions has been attributed to different origins, such as propagation of excitations in a polyelectrolyte pseudolattice or the free diffusion of the noncaged chains/particles. When polymer chains are charged in a salt-free or low-salt solution, their translational diffusive relaxation modes split into a fast and a slow diffusive relaxation mode. Muthukumar^{86–88} and Lee and colleagues¹¹¹ suggested that the fast mode, independent of the chain length, would be related to the coupled diffusion (D_f) of polyelectrolyte chains and their small counterions, namely, in a salt-free

or low-salt solution; the counterions must be condensed nearby the chain backbone, and the fluctuation of condensed counterions toward and away from each chain backbone induces an electric field that leads to an electrophoresis-mobility-related diffusive relaxation.

On the other hand, the interpretation of the slow mode, especially for those very slow relaxation modes observed in salt-free or low-salt polyelectrolyte solutions, is even more controversial. It has been attributed to dynamics of large multichain domains (transient aggregates or clusters) formed due to electrostatic interaction^{76–80} or some insoluble chain clusters, or even a trace amount of large particles introduced during the imperfect preparation of solution.^{89–93} The transient cluster interpretation was supported by studies of poly(styrene-sulfonate) with sodium counterions in aqueous solutions under dialysis. For a long time, this transient cluster has been attributed to the effective interaction between similarly charged segments, namely, the overlapping of the ion clouds of neighboring polyions, so that the more loosely associated small ions are 'shared' by two or more polyions chains, resulting in a fluctuating dipole field that tends to retard the relative motions of those participating polyions.

Experimentally, a polymer solution is normally prepared by allowing a macroscopic piece of polymer to dissolve in a solvent. The so-called permanent chain clusters must be from the incomplete dissolution in the solution preparation, not the chain association, because we do not place individual chains into a solvent. Cong *et al.*⁹⁵ addressed those possible problems and recently prepared a 'virgin' poly(styrenesulfonate) sample directly from 4-styrene-sulfonic sodium salt in an aqueous medium. They then used dialysis to *in situ* change the salt concentration. Even in that case, the solution cleanliness still deteriorates after the dialysis. To avoid some possible problems in sample preparation, ideally, one would like to have a polymer solution in which polymer chains can be *in situ* switched from a completely neutral to a fully charged state by some chemical reactions under a mild condition. We have searched such a system for a long time. Recently, using the Staudinger ligation to attach 4-(5'-(1-(dimethylamino)-ethylidene-amino)pentyl)-1-methyl-2-(diphenylphosphino)terephthalate to poly(*p*-azidomethylstyrene)-*co*-polystyrene, we successfully prepared a novel polymer, poly(4-(4-((5-(1-(dime-thylamino)ethylidene-amino)pentyl-oxy)carbonyl)-2-(diphenylphosphoryl)benzoyloxy-methyl-styrene)-*co*-poly(styrene) (P('amidine'MS)-*co*-PS), which undergo a reversible neutral-charged-neutral transition when CO₂ and N₂ are alternatively bubbled through its *N,N*-dimethylformamide (DMF) solution as illustrated in Scheme 3. Armed with this novel polymer, we are able to rule out any interference of impurities or permanent chains clusters on solutions, and *in situ* study the variation of solution dynamics of polymer chains during the neutral-charged-neutral transition.



Scheme 3 Reversible change between P('amidine')MS-*co*-PS and its charged state in DMF with 0.5% H₂O after alternating bubbling of CO₂ and N₂. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

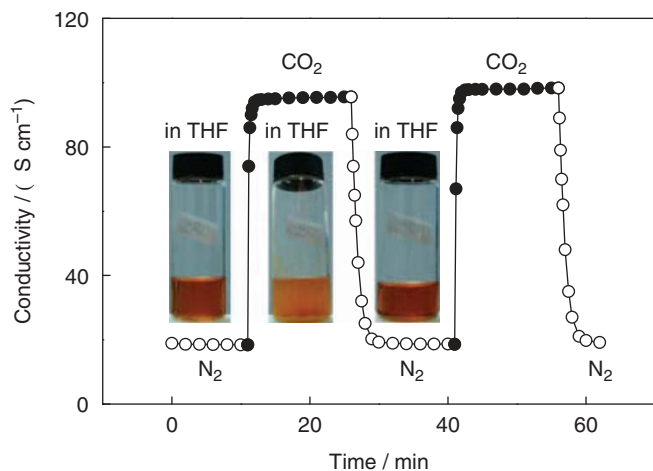


Figure 18 Conductivity of P('amidine'MS)-*co*-PS in DMF with 0.5% H₂O at 25 °C after alternative bubbling of CO₂ and N₂, where polymer concentration is 7.3 mg ml⁻¹. Insets are photos of P('amidine'MS)-*co*-PS in THF with 0.5% H₂O in different states before and after CO₂ and N₂ bubbling, where polymer concentration is 25.0 mg ml⁻¹. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

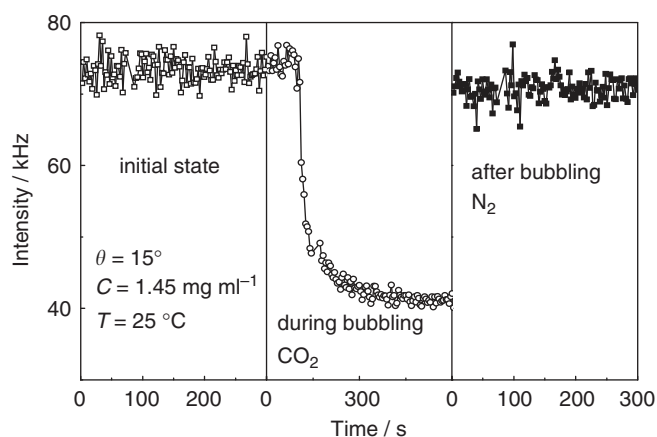


Figure 19 Time dependence of time-average scattering intensity of P('amidine'MS)-*co*-PS in DMF with 0.5% H₂O at different states before and after CO₂ and N₂ bubbling. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

Figure 18 shows that after CO₂ and N₂ gases were alternatively bubbled 30 min through each solution of P('amidine'MS)-*co*-PS in DMF at 25 °C, the solution conductivity dramatically changes five times and switches between two constant values. The switching is completely reversible. The inset photos in Figure 18 are the appearance of the solution of P('amidine'MS)-*co*-PS in a different solvent (tetrahydrofuran, THF) with 0.5% H₂O, in which the charged P('amidine'MS)-*co*-PS chains have an extremely low solubility. The change solution appearance from clear to cloudy and from cloudy to clear during the alternative bubbling of CO₂ and N₂ clearly shows the expected neutral-charged-neutral transition of P('amidine'MS)-*co*-PS chains in THF.

Dynamic LLS was also used to follow such neutral-charged-neutral transition of P('amidine'MS)-*co*-PS chains. Figure 19 shows that as individual P('amidine'MS)-*co*-PS chains are gradually charged during the CO₂ bubbling, the time-average scattering intensity ($\langle I \rangle$) decreases ~45% over a period of ~500 s. The alternative bubbling of CO₂ or N₂ can reversibly switch $\langle I \rangle$ between two average values.

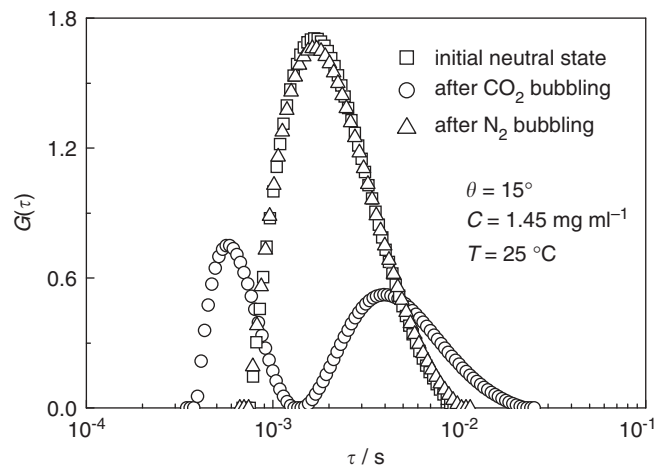


Figure 20 Characteristic relaxation time distributions in one cycle of CO₂ and N₂ bubbling of P('amidine'MS)-*co*-PS solution. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

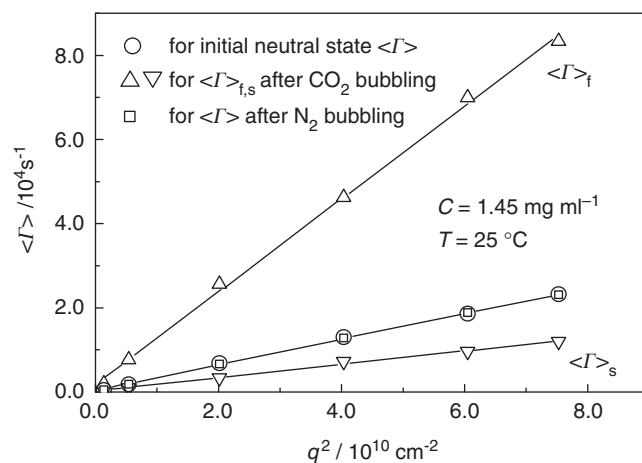


Figure 21 The scattering vector (q) dependence of average characteristic line widths ($\langle I \rangle$) in one cycle of CO₂ and N₂ bubbling of P('amidine'MS)-*co*-PS solution. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

It has been well known in LLS that $\langle I \rangle$ is proportional to $(\partial C / \partial \pi)_T$, where C and π are polymer concentration and solution osmotic pressure, respectively. When polymer chains are charged, it is more difficult to induce the concentration fluctuation for a given osmotic pressure change because of electrostatic repulsion. Therefore, such a decrease of $\langle I \rangle$ after the chains are charged is understood.

Figure 20 shows variation of the characteristic relaxation time distribution $G(\tau)$ in one cycle of CO₂ and N₂ bubbling of the P('amidine'MS)-*co*-PS solution. When the polymer chains are their initial neutral state before the CO₂ bubbling, there is only one relaxation mode. When each P('amidine'MS)-*co*-PS chain is charged after the CO₂ bubbling, $G(\tau)$ splits into two peaks. The N₂ bubbling can return $G(\tau)$ to its initial state, further indicating that the neutral-charged-neutral transition is completely reversible. The scattering vector (q) dependence of average characteristic line widths ($\langle I \rangle$) in the corresponding one cycle of CO₂ and N₂ bubbling of the P('amidine'MS)-*co*-PS solution is shown in Figure 21. No matter whether the chains are charged, the average characteristic line width

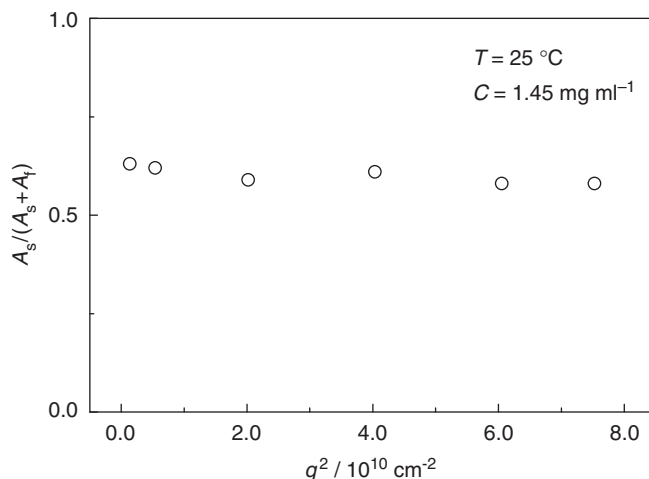


Figure 22 Scattering vector (q) dependence of intensity contribution of slow relaxation mode (A_s) after N_2 bubbling of P('amidine'MS)-*co*-PS solution. Note that $A_s + A_f = 1$. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

$\langle \Gamma \rangle = 1/\langle \tau \rangle$ is a linear function of q^2 , passing through the origin, indicating that these relaxation modes are diffusive. Further, Figure 22 shows that the intensity contribution of the slow relaxation mode (A_s) nearly remains a constant at different scattering angles, where $A_f + A_s = 1$. The q independence of the scattering intensity reveals that the slow mode is not related to scattering objects larger than $1/q$, or ~ 35 – 190 nm.

These results are similar to those recently observed by Russo and colleagues,⁹⁵ but they still attributed the slow mode to some temporal chain aggregates. We found that our results agree well with the interpretation of dynamics of individual polyelectrolyte chains in salt-free dilute solutions proposed by Muthukumar.^{86–88} Pusey and Tough¹⁴⁷ attributed the fast and slow modes to the mutual and self-diffusion, respectively. For example, in a short time, each chain moves over a short distance and the diffusion of counterions makes the charged particles to diffuse faster. After the relaxation of such 'collective diffusion', the 'self-diffusion' of individual chains in a long time or over a long length must be retarded by those surrounding interacting particles and becomes detectable in LLS. Our fast mode agrees well with the interpretation of the coupled diffusion of individual polyelectrolyte chains and their counterions in salt-free dilute solutions proposed by Lin *et al.*¹¹¹ and Muthukumar,^{86–88} whereas the slow mode is related to the self-diffusion of the center-of-mass of individual chains under long-range electrostatic interaction-induced constrains of other surrounding chains, just like a cage, because our results reveal that the slow mode is related to small but slowly moving subjects, in other words, the long-range electrostatic interaction may slow down the self-diffusion of the center-of-mass of individual chain.

Figure 23 shows how $G(\tau)$ changes in the process of backing to the neutral state when N_2 is slowly bubbled through the solution. It is clear that during the charge-to-neutral transition, $\langle \tau \rangle_{\text{fast}}$ increases, that is, D_{fast} decreases, so that the coupled fast relaxation slows down. At the same time, the total scattering intensity $\langle I \rangle$ and the contribution from the slow relaxation (A_s) increase because the electrostatic interaction constrained the motion of chains weakly and the intensity of the fast mode, which is contributed to the coupled diffusive relaxation of individual polyelectrolyte chains and their counterions, diminishes as each chain becomes neutral.

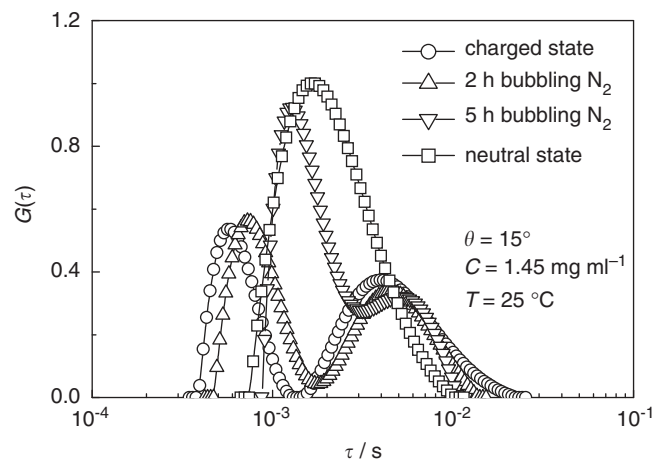


Figure 23 Characteristic relaxation time distributions $G(\tau)$ during the charge-to-neutral transition induced by slowly bubbling N_2 through P('amidine'MS)-*co*-PS solution. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

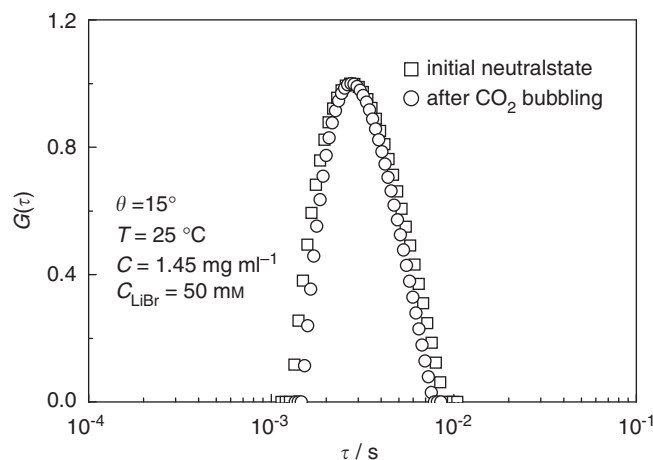
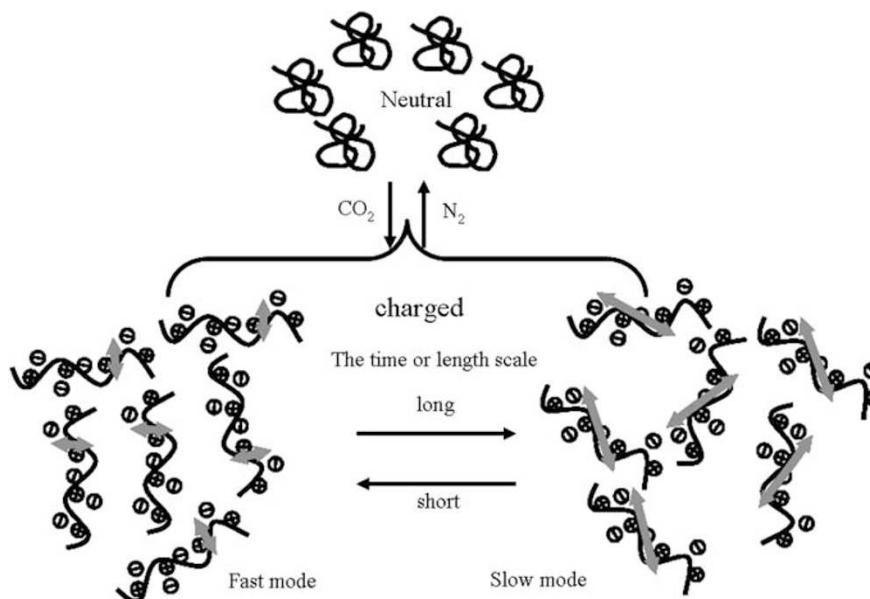


Figure 24 Characteristic relaxation time distributions $G(\tau)$ of P('amidine'MS)-*co*-PS solution in the presence of 50 mM LiBr before and after CO_2 bubbling. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

When the high amount of salt LiBr was added into the solution, as expected, we can find that there is only one relaxation mode after the CO_2 bubbling. Figure 24 shows that the CO_2 bubbling has no effect on $G(\tau)$ of charged P('amidine'MS)-*co*-PS chains in dilute solutions; namely, the addition of 50 mM LiBr sufficiently suppress electrostatic interaction among charges on the chain backbone so the charged chains behave like in the initial neutral state. The slightly larger $\langle R_h \rangle$ is attributed to the chain swelling because the addition of LiBr makes DMF a better solvent.

In summary, by using a novel P('amidine'MS)-*co*-PS that can be *in situ* and reversibly switched between a neutral state to a charged state by alternative bubbling of CO_2 and N_2 , our LLS results reconfirm that when individual polymer chains are charged in salt-free dilute solutions, their translational diffusive relaxation in the neutral state splits into a fast and a slow diffusive mode due to long-range electrostatic interaction. As expected, adding high amount of salt can suppress the electrostatic interaction and the charged chains behave like the neutral ones with only one translational diffusive relaxation mode. The fast



Scheme 4 Schematic of fast and slow relaxation modes after each P('amidine' MS)-co-PS chain becomes charged after the CO₂ bubbling; namely, the fast mode is because of the coupling of the diffusion of the chain and the wiggling of chain segments induced by fluctuation of counterions condensed nearby the charged chain backbone in a salt-free or low-salt solution and the slow mode to the self-diffusion of the charged chain with an extended conformation. Reproduced with permission from Zhou *et al.*¹⁰⁷ Copyright (2009) American Chemical Society.

mode can be attributed to the coupled diffusion originated from a convective current generated by an induced electric field arising from fluctuation of all charged species (charges on the chain and counterions) in the solution in a short time or length scale, whereas the slow mode is related to slowly moving scattering objects but it is not large, supporting the assumption of the retarded self-diffusion as Scheme 4.

Conclusion

Using *in situ* anionic polymerization to directly prepare macroscopically homogeneous dust-free semidilute/concentrated PS solutions inside a LLS cell, we comparatively reexamined solution dynamics of PS in an athermal solvent (toluene) and in a less good/ θ solvent (cyclohexane in the range 34–50 °C) by a combination of static and dynamic LLS. We confirm that in toluene there is only one fast diffusive relaxation mode in the measured intensity–intensity time correlation function even for the solution with a concentration (C) 30 times higher than the overlapping concentration (C^*). In cyclohexane, an additional slow mode appears as soon as $C/C^* > 1$. Our results show that the slow mode is real, not because of previously suggested artifacts in the solution preparation. The concentration, and time dependence of the characteristic relaxation time of the slow mode ($\langle \tau \rangle_s$), the scaling exponent (α_s) between $\langle \tau \rangle_s$ and the scattering vector (q), and the relative intensity contribution (A_s) of the slow mode suggests that such a slow mode has two different natures, depending on the polymer concentration.

When C is higher than the overlapping concentration (C^*) but lower than the entanglement concentration (C_e), only some of the chains are entangled together to form large transient clusters, which coexist with individual (nonentangled) chains in the solution. In such a transitional region between dilute and semidilute solutions, the relatively stronger segment–segment interaction in a less good/ θ solvent makes the movements of the chains inside each cluster more correlated within a short delay time window (< 1 s) and leads to the slow mode. The temperature dependence of the slow relaxation mode in two different systems, PNIPAM/H₂O and PEO-PPO-PEO/H₂O

further confirms the coexistence of individual chains and large temporal chain clusters in the semidilute regimes.

When $C > C_e$, all the chains are entangled with each other to form one huge 'cluster'. Each chain is confined inside a 'tube' made of its surrounding chains and the gravity center of each chain segment (blob) fluctuates inside the tube within a short time scale ($\sim 10^{-3}$ s). In an athermal solvent, there is no interaction among the chains so that all the blobs are in the same microenvironment, leading to only one fast relaxation mode. In a less good/ θ solvent, the relatively stronger segment–segment interaction near the entanglement points makes the tube nonuniform. It can be visualized that each chain is inside a nonuniform tube with a band-like structure. Therefore, the blobs experience two different microenvironments. The restricted movement of the blobs near the entanglement points results in the slow mode. The use of a special (2+2) photoaddition to induce *in situ* crosslinking of polymer chains rather than the entanglement in a semidilute solution further indicates an expected increase of the relative contribution of the slow mode as the crosslinking proceeds because the crosslinking enhances the restricted movement of the polymer chains near the crosslinked (entangled) points.

By using a novel P('amidine' MS)-co-PS that can be *in situ* and reversibly switched between a neutral to a charged state by alternative bubbling of CO₂ and N₂, we reconfirm that when individual polymer chains are charged in salt-free dilute solutions, their translational diffusive relaxation in the neutral state splits into a fast and a slow diffusive mode (D_f and D_s), respectively, due to the long-range electrostatic interaction. The fast mode can be attributed to the coupled diffusion originated from a convective current generated by an induced electric field arising from fluctuation of all charged species (charges on the chain and counterions) in the solution in a short time or length scale. However, there is still no decisive evidence to differentiate whether the slow relaxation is related to large temporal aggregates formed because of the overlapping of ion clouds among different chains or to the self-diffusion of individual chains retarded by surrounding chains (interchain friction) in a long time or length

scale. Our results indicate that the slow mode is not related to large but slowly moving scattering objects, supporting the assumption of the retarded self-diffusion. Consequently, we conclude that the slow mode observed in the semidilute, polyelectrolyte, or even gelling system is real and can be generalized as hindered motions of interacting polymer chains even though the nature of interaction can be different, including, the weak segment–segment interaction in a less good solvent to strong electrostatic interaction among polyelectrolyte chains, and even to chemical crosslinking inside gel networks.

ACKNOWLEDGEMENTS

The financial support of the National Natural Scientific Foundation of China (NNSFC) Project (20904049, 20934005 and 50773077), and the financial support of the Hong Kong Special Administration Region General Research Fund (CUHK402506, 2160291; and CUHK403908, 2160361) is gratefully acknowledged.

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