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Transition Metal Complexes of 1, 4(2'-Hydroxyphenyl-1-yl) di-imino azine: Synthesis, Characterization and Antimicrobial Studies

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Abstract: The synthesis and characterization of first row transition metal complexes of the 1, 4(2'-hydroxyphenyl-1-yl) di-imino azine {1,4(2'HPDA)} are reported. The complexes have been characterized by elemental analysis, molar conductance, magnetic studies, IR, ¹H NMR and UV-visible studies. They have the stoichiometry of the type [M{1,4(2'HPDA)}₂ 2H₂O] and [M'L₂] where M = Mn(II), Fe(III), Co(II), Ni(II) and Cu(II), and M' = ZrO(II), VO(II), Zn(II), Cd(II), and Hg(II). The antibacterial and antifungal activity of the metal complexes has been investigated. Both ligand and complexes have shown good antibacterial and antifungal activity

Keywords: 1,4(2'-Hydroxyphenyl-1-yl) di-imino azine {1,4(2'HPDA)}, IR, ¹HNMR, Antimicrobial

Introduction

Metal ions play a vital role in a vast number of biological processes. The ions with biologically active ligands are a subject of considerable interest. Some of the biologically active compounds act via chelation¹. But for most of them little is known about how metal binding influences their ability. Some heterocyclic azines are known to inhibit marine tumor growth² and acts as fluorescent brightening agents and are photosensitizers³. Azines were developed for use as ion selective optical sensors⁴. Mixed azines between opioid antagonist and steroidal ketones⁵ have shown various metal complexes for their antimicrobial and antifungal properties.

Cisplatin exerts its biological action by preferential binding of the *cis*-(NH₃)₂Pt²⁺ unit to the N7 sites of adjacent guanine residues of DNA forming intrastrand cross-links⁶⁻⁷. The recent observation that derivatives of 5- and 6-uracilmethylphosphonate (Umpa2-), in combination with *Cisplatin*, prolong the survival time of mice with lymphoid leukemia⁸, prompted a study of their acid-base properties⁹. It was shown that the stability of these M(Umpa) complexes¹⁰ is solely determined by the basicity of the phosphonate group and these two ligands may be considered as simple analogues of uridine 5'-monophosphate (UMP2-). In an effort to combine the antiviral with the anticancer effects, the acid-base and metal ion-binding properties of PMEA (Nucleotide Analogue 9-[2-(Phosphonomethoxy)ethyl]adenine) and of its ternary complexes, (Dien)Pt(PMEA-*N1*) and (Dien)Pt(PMEA-*N7*) (Dien = diethylenetriamine), were studied and compared by Kampf *G et al.*¹¹.

One of the most remarkable results is that the metal nonbinding properties of the phosphonate group are little affected by the nucleobasecoordinated (Dien) Pt²⁺ units. In this context it is important to understand the effect of metal ions upon their binding to nucleobase residues, i.e. how they affect the acid-base properties of nucleobases. This is crucial with regard to their metal nonbinding, hydrogen-bonding¹² and stacking properties. In view of the above discussions it has been planned to synthesize ZrO(II), VO(II), Mn(II), Fe(III), Co(II), Ni(II), Cu(II), Zn(II), Cd(II) and Hg(II) complexes of {1,4(2'HPDA)} to deduce their structure by various techniques such as elemental analysis, molar conductance, UV-Vis, Magnetic susceptibility, IR and ESR data. We have also carried out the antimicrobial activity of the ligand and their metal complexes.

Experimental

All chemicals used were of analytical grade. The benzophenone hydrazine hydrate was prepared by their literature methods. The metal content was determined as already reported. The nitrogen content was estimated micro-analytically.

Preparation of 1, 4(2'-hydroxyphenyl-1-yl) di-imino azine {1,4(2'HPDA)}

Benzophenone is condensed with hydrazine hydrate in presence of catalytic amount of concentrated hydrochloric acid. The resulting product was treated with aromatic aldehyde in the molar ratio of 1: 2. The reaction mixture was refluxed on water bath for about 4 h. The solid azine separated after a few minutes. The product was filtered out and washed successively with alcohol. The compound was re-crystallized from alcohol and dried in vacuum over fused CaCl₂ [Mol. Formula C₁₄H₁₂N₂O M.P. = 187° C, yield = 75 %].

Preparation of complexes

Metal complexes were prepared by adding (0.2 mole, 20mL) aqueous metal(II) chlorides to 0.1 mole in 50mL alcohol. The complexes did not separate immediately, the reaction was refluxed on water bath for about 2 h, during which complexes start separating out. The solution was concentrated and cooled to room temperature. The crystalline light colored complexes were filtered out, washed with water and alcohol to remove the unreacted salts and ligand. In the case of VO(II) and ZrO(II) metal oxy chlorides were used. All the complexes were dried in open air and kept in vacuum desiccator.

Physical measurements

The magnetic susceptibility of the complexes was made on Gouy balance using Hg[Co(NCS)₄] as the calibrant at room temperature. The IR spectra were obtained from Perkin-Elmer (model 783) in KBr pellets. The electronics spectra of VO(II), Mn(II), Fe(III),

Co(II), Ni(II), and Cu(II) complexes were recorded on Elico CL-24 spectrophotometer in DMF (10^{-3} M) solution. The conductance measurements were made on Elico CM-82T conductivity bridge in DMF (10^{-3} M) solution. All the complexes gave satisfactory analysis for metal and nitrogen content. The Electron spin resonance spectra of Copper(II), complex in poly crystalline state were recorded on Varian X – band ESR spectrometer diphenyl picryl hydrazine (DPPH) free radical of “g” marker ($g = 2.0027$) at room temperature.

Results and Discussion

Analytical data for the complexes indicate 1:2 stoichiometries for all the complexes (Table 1). All complexes possess high melting points and are stable in air and partially or insoluble in common organic solvents and soluble in DMF, DMSO and pyridine. The molar conductance of the complexes are in the range 8.19- 25.00 $\text{ohm}^{-1} \text{cm}^2 \text{mole}^{-1}$ in DMF (10^{-3} M) solution. These values are too low to account for any electrolytic behavior of the complexes.

Magnetic and Electronic spectral studies

Electronic spectral studies of VO(II), Mn(II), Fe(III), Co(II), Ni(II) and Cu(II) complexes were carried out in DMSO solution (1×10^{-3} M). The electronic spectrum of oxovanadium (IV) complex exhibited three bands¹² in the region 22000-13000 cm^{-1} . The present oxovanadium complex of the ligand {1,4(2'HPDA)}, have displayed three bands in the region 12820.51-23809.52 cm^{-1} . Bajgur *et al.*¹³, have reported three absorption bands in the region 13000-13650 cm^{-1} , 16300-16800 cm^{-1} and 21000-22000 cm^{-1} for vanadium complexes derived from carbohydrazones. In view of these facts and on the basis of available reports the vanadium complex under the present study has been assigned the square pyramidal geometry. Sakurai *et al.*¹⁴, have reported the magnetic susceptibility values in the range of 1.7 to 1.8 for the complexes, confirming that they all in the mononuclear vanadyl state. In the present investigation, the magnetic moment value observed for VO(II) complex of the ligand {1,4(2'HPDA)} is 1.6 BM. This value is agreeable with the spin only value for oxovanadium(II) complex, indicates that square pyramidal geometry around the VO(II) ion.

Magnetic studies show that the ground term of the manganese(II) ion is the sextet. The only sextet term of the d^5 configuration in octahedral stereochemistry is the ${}^6A_{1g}$. The transitions of the spectrum are assigned as from the ${}^6A_{1g}$ ground term to the quartet excited terms¹⁵. The electronic spectrum of the Mn(II) complex of ligand {1,4(2'HPDA)} indicated by the bands in the region 15151.51-16666.66 cm^{-1} due to ${}^6A_{1g} \rightarrow {}^4T_{1g}({}^4G)$ (ν_1), ${}^6A_{1g} \rightarrow {}^4E_g({}^4G)$ (ν_2), ${}^6A_{1g} \rightarrow {}^4T_{2g}({}^4D)$ (ν_3), ${}^6A_{1g} \rightarrow {}^4T_{1g}({}^4P)$ (ν_4) transitions respectively, indicating octahedral geometry¹⁶ for these complexes. In the present investigation the observed magnetic moment values for Mn(II) complex of ligand {1,4(2'HPDA)} is 5.25 BM, which is well within the range expected for octahedral geometry around the central metal ion¹⁷.

The absorption spectra of high-spin iron(III) complexes have been measured with difficulty¹⁸. The ground state of high spin octahedrally coordinated Fe (III) complexes is ${}^6A_{1g}$. The four lowest energy bands are due to the transition from ${}^6A_{1g}$ to ${}^4T_{1g}$, ${}^4T_{2g}$, 4E_g , and ${}^4A_{1g}$, excited states. A very sharp band near 25000 cm^{-1} is due to the transitions to the accidentally degenerate ${}^4A_{1g}$ and 4E_g levels. These bands possibly arise from an admixture of spin quartet character to the ground state through spin-orbital coupling. The transitions ${}^6A_{1g} \rightarrow {}^4A_{1g}$, ${}^4E_g(G)$ and ${}^4E_g(D)$ are of interest since their energies are independent of the magnitude of D_q and this permits a comparison of the free ion and bound ion term separations. In the present investigation, the Fe(III) complex of ligand {1,4(2'HPDA)} have shown three bands in the region 13513.51-20833.33 cm^{-1} which are characteristic of octahedral geometry. The observed magnetic moment value for Fe(III) complex in the present investigation is 5.67 B.M which is a characteristic of octahedral geometry¹⁹.

Table 1. Elemental analysis, Color, Melting point and Conductance data for {1,4(2'HPDA)} and Metal complexes

S. No	Ligand/complexes	Analysis%				M.P	Color	Magnetic moment BM	Molar Conductance $\Omega^{-1}\text{Cm}^2\text{Mol}^{-1}$
		Found (Calculated)							
		C	H	N	M				
01	{1,4(2'HPDA)}	75.25 (75.00)	5.48 (5.38)	12.75 (12.50)	-	187°C	Yellow	-	-
02	ZrO[{1,4(2'HPDA)}] ₂	59.04 (60.29)	4.05 (3.99)	11.51 (10.04)	17.12 (16.37)	>300°C	Yellow	-	25
03	VO[{1,4(2'HPDA)}] ₂	66.00 (65.50)	4.56 (4.28)	10.32 (10.99)	10.20 (9.99)	>300°C	Yellow Green	1.68	45
04	Mn[{1,4(2'HPDA)}] ₂ .2H ₂ O	61.21 (62.57)	5.27 (4.87)	11.06 (10.42)	11.46 (10.23)	>300°C	Brown	5.25	31
05	Fe[{1,4(2'HPDA)}] ₂ .2H ₂ O	61.70 (62.47)	4.89 (4.83)	10.69 (10.41)	11.96 (10.38)	>300°C	Reddish Brown	5.67	32
06	Co[{1,4(2'HPDA)}] ₂ .2H ₂ O	58.85 (62.11)	5.03 (4.80)	10.60 (10.35)	11.28 (10.89)	>300°C	Orange	4.65	09.23
07	Ni[{1,4(2'HPDA)}] ₂ .2H ₂ O	59.76 (62.14)	5.30 (4.80)	9.60 (10.35)	11.96 (10.89)	> 300°C	Yellow Green	3.23	21.22
08	Cu[{1,4(2'HPDA)}] ₂ .2H ₂ O	60.34 (61.50)	5.28 (4.76)	10.92 (10.26)	10.20 (11.64)	> 300°C	Brown	1.86	8.301
09	Zn[{1,4(2'HPDA)}] ₂	65.12 (65.62)	4.34 (5.07)	11.28 (10.93)	13.03 (12.88)	>300°C	Orange	-	30
10	Cd[{1,4(2'HPDA)}] ₂	60.05 (60.17)	4.50 (3.93)	10.11 (10.02)	21.08 (20.03)	>300°C	Yellow	-	22
11	Hg[{1,4(2'HPDA)}] ₂	51.62 (51.99)	3.66 (3.40)	8.89 (8.66)	31.32 (31.04)	>300°C	Brown	-	23

The electronic spectra of Co(II) complex of ligand {1,4(2'HPDA)} exhibited bands in the region 10101.01, 18181.81 and 23923.44 cm^{-1} due to the ${}^4\text{T}_{1g}(\text{F}) \rightarrow {}^4\text{T}_{2g}(\text{F})$ (ν_1), ${}^4\text{T}_{1g}(\text{F}) \rightarrow {}^4\text{A}_{2g}(\text{F})$ (ν_2) and ${}^4\text{T}_{1g}(\text{F}) \rightarrow {}^4\text{T}_{1g}(\text{P})$ (ν_3) transitions respectively. These transitions suggest octahedral geometry for the Co(II) complexes. These assignments are in good agreement with the reported values^{20,21}. The observed magnetic moment value for Co(II) complexes of {1,4(2'HPDA)} is 4.65 BM, which indicates clearly the octahedral geometry of the Co(II) complex²².

Electronic spectra of Ni(II) complex of the ligand {1,4(2'HPDA)} in the present investigation exhibited three bands in the region 11135.83 – 11428.57 cm^{-1} , in DMSO solution. These bands are assigned due to ${}^3\text{A}_{2g}(\text{F}) \rightarrow {}^3\text{T}_{2g}(\text{F})$ (ν_1), ${}^3\text{A}_{2g}(\text{F}) \rightarrow {}^3\text{T}_{1g}(\text{F})$ (ν_2) and ${}^3\text{A}_{2g}(\text{F}) \rightarrow {}^3\text{T}_{1g}(\text{P})$ (ν_3) transitions respectively. Moreover, the ratio ν_2/ν_1 is in the range 1.21 – 1.38 is indicative of octahedral stereochemistry for these Ni(II) complexes. Koji *et al.*,²³ have reported a band around ~950 nm for d-d component of the Ni(II) complex. In the present study, electronic spectral band in the region 11135.83 – 11428.57 cm^{-1} (~900 nm) is attributable to a d-d component of the Ni(II) complexes²³. The magnetic moment values for Ni(II) complexes of ligands {1,4(2'HPDA)} is 3.23 BM, which is well within the range expected for octahedral geometry around the central metal ion²⁴.

Cu(II) complex of the ligand {1,4(2'HPDA)}, showed broad band at 16129.00 cm^{-1} . Campbell and Grzeskowiak²⁵ have suggested that tetrahedrally distorted octahedral geometry for Cu(II) complexes derived from thiosemicarbazide ligands. The observed broad band in the present Cu(II) complex of ligand {1,4(2'HPDA)} can be assigned to the envelope of ${}^2\text{B}_{1g} \rightarrow {}^2\text{E}_g$, and ${}^2\text{B}_{1g} \rightarrow {}^2\text{E}_{2g}$ transitions in distorted octahedral geometry of Cu(II). The observed magnetic moment value for Cu(II) Complex 1.86 BM. This value is agreeable to spin only value. Hence observed magnetic moment for the Cu(II) complexes under study that indicates it has distorted octahedral²⁶ configurations.

IR Spectra

The chemical structures of the symmetric and asymmetric dimmers are similar to benzylic dimmers. Except for the hydroxyl groups, the hydroxyl stretching regions of FT-IR spectra for the symmetric and asymmetric dimmers will reflect the distribution of hydrogen bonding in the dimmers. In the study of hydrogen bonding for a benzoxazine model dimer using methyl ammine by Dunkers²⁷, it was shown that it is possible to form several kinds of hydrogen bonded species, such as – OH O inter molecular hydrogen bonding, -OH O intra molecular hydrogen bonding and –H ...N intramolecular hydrogen bonding. The band centered at 3400 cm^{-1} shows the existence of – OH O intermolecular hydrogen bonding, and the broad band between 2500 cm^{-1} and 3300 cm^{-1} might be designated for – OH O intra molecular hydrogen bonding and – OH N intramolecular hydrogen bonding. It is well established that the Schiff bases having o-hydroxy group either on aldehyde or aniline residue can form intramolecular hydrogen bonding with azomethine nitrogen²⁸. This has direct impact on the ν_{OH} vibrations and the band due to ν_{OH} shifts to the lower frequency with broadening of the band and decrease in the intensity. The extent of shift depends on the strength of hydrogen bonding²⁹. The stronger the hydrogen bonding, the greater is the shift towards the lower frequency side and this is followed by an immense change in the intensity of the band. In our earlier reports³⁰ we have observed hydrogen bonded -OH with nitrogen atom of azomethine group appeared in the region 2875-2746 cm^{-1} .

In addition to these bands in the ligand {1,4(2'HPDA)}, a high intensity band observed in the range 1265 cm^{-1} is due to the phenolic $\nu_{\text{C-O}}$ vibrations³¹. In phenols, both -OH bending and phenolic $\nu_{\text{C-O}}$ vibrations are coupled and they give rise to absorption at 1340 and 1175 cm^{-1} for free -OH and at 1230 cm^{-1} for bending -OH. Teyssinc and Charette³² have assigned a band at

1280 cm^{-1} to the $-\text{OH}$ bending vibrations and other at 1324 cm^{-1} to the $\nu_{\text{C-O}}$ vibrations in the acetyl acetone Schiff bases. Assignment in the ligand has been made by comparing the spectra with those of a model compound *viz.*, benzylidene aniline and salicylidene aniline. In the present investigation the band due to hydrogen bonded $-\text{OH}$ group was observed at 3040 cm^{-1} in the case of {1,4(2'HPDA)} ligand. The medium bands observed in the region 1619-1575 cm^{-1} are assigned to $\nu_{\text{C=N}}$ group³³. In the present study of the ZrO(II), VO(II), Mn(II), Fe(III), Co(II), Ni(II) and Cu(II), complexes of ligand {1,4(2'HPDA)}, indicates the presence of weak band around 847 – 803 cm^{-1} , which may be assigned due to the presence of coordinated water molecule in the complexes. Whereas in case of Zn(II), Cd(II), and Hg(II) complexes of ligand {1,4(2'HPDA)}, absence of coordinated water molecule.

The assignment of the band to various $\nu_{\text{M-O}}$ and $\nu_{\text{M-N}}$ vibrations in the lower frequency region appears to be complicated as the ligand vibrations interfere in this region. The low frequency skeletal vibrations due to M-O and M-N stretching provide direct evidence for the complexation. This is because a large dipole moment change is involved in the vibration of M-O bond in comparison to that of M-N bond³⁴. Choudhary *et al.*,³⁵ have reported bands in the region 600-450 cm^{-1} for M-O and 400-310 cm^{-1} for M-N vibrations respectively. Bipin *et al.*,³⁶ have also reported that the vibrations in the region 565-552 cm^{-1} and the bands around 414 cm^{-1} for M-N and M-N respectively. Present study of the complexes $\nu_{\text{M-N}}$ vibration occur in the region 540 – 490 cm^{-1} and $\nu_{\text{M-O}}$ occur in the region 450 – 400 cm^{-1} have been assigned.

ESR Spectra

In order to obtain more information about the magnetic environment of the Cu(II) complexes, powder samples were used to record X-band ESR spectra of the complexes at room temperature. ESR spectra of the polycrystalline Cu(II) complex 'g' values obtained from the spectra are presented in Table-2. In the present case, copper (II) complexes, measured in polycrystalline sample at room temperature, give the following values: $g_{\parallel} = 2.06$, $g_{\perp} = 2.13$ for the $[\text{Cu}\{1,4(2'HPDA)\}_2 2\text{H}_2\text{O}]$ and $g_{\parallel} = 2.03$, $g_{\perp} = 2.10$. The trend, $g_{\parallel} < g_{\perp}$ showed that the electron is delocalized in d_{z^2} orbital of the ground state of Cu(II) and the spectra are characteristic of axial (compressed octahedral) symmetry. The parameter G, determined as $G = (g_{\parallel} - 2) / (g_{\perp} - 2)$ is found to be much less than 4 suggesting considerable interaction in the solid state³⁷.

Table 2. Electron spin resonance spectral data

S.No	Complex	g_{\parallel}	g_{\perp}	g_e	G_{av}	G
1	$\text{Cu}\{1,4(2'HPDA)\}_2 2\text{H}_2\text{O}$	2.06	2.16	2.0039	2.11	3.75

¹H NMR Spectrum of the ligand {1, 4(2'HPDA)}

The ¹H NMR spectrum of ligand 1-Phenyl-4-(2'-hydroxyphenyl-1-yl) Di-imino azine ({1, 4(2'HPDA)}) showed a peak at δ 2.35 (s, 3H) due to $-\text{CH}_3$ group on the phenol ring. Two azomethine protons have resonated as a singlet at δ 8.73 (s, 2H). Proton due to phenolic $-\text{OH}$ has appeared at δ 12.29 (s, 1H) as a singlet whereas the eight aromatic protons have resonated in the region δ 7.51-8.30 (m, 8H) as a multiplet.

¹H NMR spectra of Cd (II) complex of {1, 4(2'HPDA)}

¹H NMR spectrum of Cd (II) complex of ligand {1, 4(2'HPDA)} and the sample was scanned in DMSO-*d*₆ showed a peak at δ 2.2 (s, 6, 2xArCH₃) due to six methyl protons at 4-position of phenyl ring. Four methylene protons have resonated at δ 3.6 (s, 4, 2xCH₂) and four aromatic protons of the phenyl ring have resonated as a singlet at δ 7.5 (s, 4, 2xArH₂) respectively. Four azomethine protons have appeared at δ 8.4 (s, 4, $-\text{CH}=\text{N}-$) as a singlet. The

proton of phenolic OH group which was observed at δ 12.0 in case of ligand has disappeared in complex, indicating coordination of the phenoxide ion to the metal ions *via* deprotonation.

Antibacterial Activity

The comparison of inhibition zone values for the metal complexes are given in Table 3 and Table 4, reveals that the antimicrobial activity could be mainly due to the structure of the complexes and also the oxidation state of the metal ions. These results must be directly related to the greater biological activity exhibited by the tetrahedral Zn(II), Cd(II) and Hg(II) complexes compared to the square pyramidal ZrO(II) and VO(II) and octahedral Mn(II), Fe(III), Co(II), Ni(II) and Cu(II) complexes.

Table 3. Antibacterial activities of {1,4(2'HPDA)} and their metal complexes

S. No.	Sample Name	<i>Klebsiella</i> mm	<i>Pseudomonas</i> mm
	{1,4(2'HPDA)}Ligand)	-	-
1	Mn[{1,4(2'HPDA)}] ₂ 2H ₂ O	-	-
2	Co[{1,4(2'HPDA)}] ₂ 2H ₂ O	17.00	12.5
3	Ni[{1,4(2'HPDA)}] ₂ 2H ₂ O	06.25	11.5
4	Cu[{1,4(2'HPDA)}] ₂ 2H ₂ O	13.25	14.66
5	Fe[{1,4(2'HPDA)}] ₂ 2H ₂ O	14.33	08.00
6	ZrO[{1,4(2'HPDA)}] ₂	14.66	14.66
7	VO [{1,4(2'HPDA)}] ₂	12.66	11.33
8	Zn [{1,4(2'HPDA)}] ₂	12.66	17.33
9	Cd [{1,4(2'HPDA)}] ₂	18.33	15.66
10	Hg [{1,4(2'HPDA)}] ₂	18.00	22.33
	Standards Gentamicin	19.00	17.35

5.00 to 8.00 mm are less active, above 10.00 to 13.50 mm are moderately active, 13.50 to 23.33 are significantly active.

Table 4. Antifungal Activities of various synthesized metal {1,4(2'HPDA)} complexes:

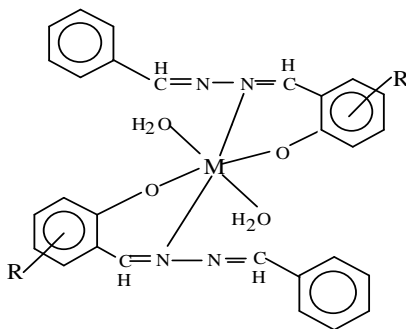
Sl No.	Name of the complexes	<i>A.Niger</i> mm	<i>A.Flavous</i> mm
	{1,4(2'HPDA)}	11.00	05.00
1.	Mn[{1,4(2'HPDA)}] ₂ 2H ₂ O	12.66	07.00
2.	Co[{1,4(2'HPDA)}] ₂ 2H ₂ O	16.44	05.66
4.	Ni[{1,4(2'HPDA)}] ₂ 2H ₂ O	10.66	07.00
4.	Cu[{1,4(2'HPDA)}] ₂ 2H ₂ O	09.66	07.66
5.	Fe[{1,4(2'HPDA)}] ₂ 2H ₂ O	14.44	06.00
6.	ZrO[{1,4(2'HPDA)}] ₂	11.44	05.44
7.	VO [{1,4(2'HPDA)}] ₂	14.66	07.00
8.	Zn [{1,4(2'HPDA)}] ₂	11.66	05.00
9	Cd [{1,4(2'HPDA)}] ₂	12.66	06.00
10.	Hg [{1,4(2'HPDA)}] ₂	29.44	21.66
	Standard Nistatin	20.11	21.33

5.00 to 9.00 mm are less active, above 09.00 to 14.00 mm are moderately active, 15.00 to 29.00 are significantly active

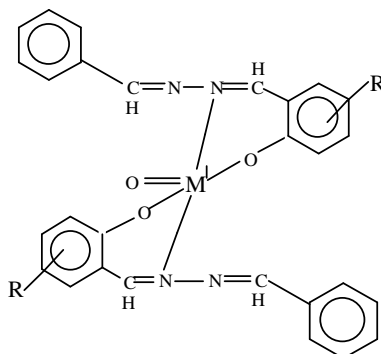
A possible explanation for the high toxicity of metal complexes can be explained as follows. The increase in the activity of metal complexes may be due to effect of metal ions on the normal cell process. The polarity of metal ion is considerably reduced on chelation which is mainly because of partial sharing of its positive charge with a donor groups and possibly π -electron delocalization over the whole molecule. Such molecule increases the lipophilic character of the metal complexes which probably leads to break down of permeability barrier of the cells resulting in interference with normal cell process³⁸. Better activities of the metal complexes as compared to free ligand could also be understood in terms of chelation theory, which explains that a decrease in polarizability of the metal could enhance the lipophilicity of the complexes.

Conclusions

Based on the above observations of IR, ESR, electronic spectral data, magnetic, conductance measurements and insolubility in common organic solvents and high melting points, we propose the octahedral structure to Mn(II), Fe(III), Co(II), Ni(II) and Cu(II), square pyramidal for ZrO(II) and VO(II), tetrahedral geometry of Zn(II), Cd(II) and Hg(II).



M = Mn(II), Fe(III), Co(II), Ni(II) and Cu(II)



M = ZrO (II), VO(II), Zn(II), Cd(II) and Hg(II)

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