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Ultrafast-charging and long cycle-life anode materials of TiO₂-bronze/nitrogen-doped graphene nanocomposites for high-performance lithium-ion batteries†

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Emerging technologies demand a new generation of lithium-ion batteries that are high in power density, fast-charging, safe to use, and have long cycle lives. This work reports charging rates and specific capacities of TiO₂(B)/N-doped graphene (TNG) composites. The TNG composites were prepared by the hydrothermal method in various reaction times (3, 6, 9, 12, and 24 h), while the N-doped graphene was synthesized using the modified Hummer's method followed by a heat-treatment process. The different morphologies of TiO₂ dispersed on the N-doped graphene sheet were confirmed as anatase-nanoparticles (3, 6 h), TiO₂(B)-nanotubes (9 h), and TiO₂(B)-nanorods (12, 24 h) by XRD, TEM, and EELS. In electrochemical studies, the best battery performance was obtained with the nanorods TiO₂(B)/N-doped graphene (TNG-24h) electrode, with a relatively high specific capacity of 500 mA h g⁻¹ at 1C (539.5 mA g⁻¹). In long-term cycling, excellent stability was observed. The capacity retention of 150 mA h g⁻¹ was observed after 7000 cycles, at an ultrahigh current of 50C (27.0 A g⁻¹). The synthesized composites have the potential for fast-charging and have high stability, showing potential as an anode material in advanced power batteries for next-generation applications.

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Introduction

The necessity of exploring renewable and environmentally compatible alternative energy sources and energy storage systems cannot be underemphasized. Lithium-ion batteries (LIBs) are one of the most effective energy storage systems and have attracted an attention from researchers and industries.^{1–3} The LIBs draw an attraction for these applications due to high specific capacity, long cycle life, no memory effect, low self-discharge rate, and long shelf life. But despite the numerous advantages over other batteries, LIBs have limitations which limit their uses in applications such as in electric vehicles. These applications require high energy density, good cycling performance, high rate capability, safety in use, and low production cost.^{4–6} To address these requirements,

research on the anode and cathode – principal parts of LIBs – have focused on enhancing their electrochemical properties.⁷ For anode materials, it is a challenge to obtain a high specific capacity along with the quick charge property, due to the kinetics of Li-ion diffusion in the anode structure. Safety is also an issue that needs to be considered.

Currently, graphite is commonly used as a commercial anode material, although it has some severe disadvantages. Graphite allows only one Li-ion intercalated with six carbon atoms, and thus its theoretical capacity is only 372 mA h g⁻¹. Besides, the low diffusion rate of Li-ions is 10⁻⁹ cm² s⁻¹, which proceed batteries with low power density. When graphite is lithiated at low potential (~0.1 V vs. Li/Li⁺), an additional concern is also present: the lithium electroplating can cause a short circuit and thermal runaway conditions resulting in the combustion of organic electrolyte and catastrophic battery explosion.^{8,9} Therefore, industries are looking for new materials to replace graphite.

Si, Ge, and Sn elements that can alloy with lithium and deliver a high theoretical capacity of 4200, 1600, and 994 mA h g⁻¹, respectively, have been investigated for high capacity anodes.^{10–12} However, problems of high-volume expansion (>400%), poor electron transport, capacity fading, and low coulombic efficiency (CE) have been observed. Nano-structural fabrication of these materials can alleviate these problems.^{13–15}

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Moreover, the effective strategy focuses on using a nano-composite of these elements with various types of carbonaceous materials.¹⁶ Specifically, a 2D-graphene sheet, which provides stable and flexible matrices, acts as a mechanical buffering zone against the large volume change problem of the active materials.^{17,18} The large surface area can contain the active materials and provide the conducting backbone on the 2D-structure.¹⁹ Moreover, doping the graphene with atoms such as nitrogen can enhance the specific capacity of the anode material.

Recently, nitrogen-doped graphene (N-doped graphene) has become the trend of anode materials for next-generation LIBs due to its certain properties.^{20,21} First, electrode/electrolyte wettability is induced by the nitrogen defects on the surface. Second, the types of pyridinic-N and pyrrolic-N, which contained lone electron pairs, can establish a delocalized conjugated system with the sp²-hybridization carbon frameworks to enhance the electrical conductivity. Third, both pyridinic-N and quaternary-N enhance lithium intercalation and extraction, increasing the specific capacity of the anode. In the charged state, a negative field is generated by electron delocalization in pyridinic-N and quaternary-N which offer more active sites and enhance the Li-ion interaction in the graphene structure. Thus, N-doped graphene is a promising anode material for LIBs.^{22–25}

Another alternative anode material is TiO₂, which offers advantages of high Li-ion diffusion rate in lithiation processes (~1.6 V vs. Li/Li⁺), safety in use, and good stability. Choosing an anode with a higher lithiation potential greatly reduces the chance of a battery explosion.²⁶ Among the polymorphs, bronze-TiO₂ (TiO₂(B)) is attractive mainly due to high energy density related to the unit cell contains 8 Ti sites and 10 Li-ion sites, giving a theoretical capacity of 1.25 Li⁺/Ti (~420 mA h g⁻¹). The open structure of TiO₂(B) provides 1D infinite channels, which can accommodate the volume changes.^{27–29} Interestingly, pseudocapacitive behavior was also reported for TiO₂(B), in which the monoclinic crystal structure of TiO₂(B) has a low-density crystal framework. These exhibit large channels and voids related to the size of Li-ion that result in a fast Li-ion diffusion rate during the charging/discharging. As a result, the fast charging performance is led.^{30–33} However, the TiO₂(B) has a relatively low specific capacity and poor electrical conductivity. These drawbacks could be overcome by combining TiO₂(B) with N-doped graphene, which has a high specific capacity and conductivity. The flexibility of the latter would also improve the cycling capability of the LIBs.

In this study, N-doped graphene and TiO₂(B) were synthesized and investigated as composite materials by simple chemical method and heat treatment process. The TiO₂(B)/N-doped graphene (TNG) composites are expected to have excellent anode material characteristics, having high specific capacity at high current density, adaptive fast charging capability, long life cycle, and high level of safety.

Experimental procedure

N-doped graphene synthesis

N-doped graphene, using a method previously developed by our group^{34,35} was synthesized using graphite powder as a precursor.

To reduce the initial powder size, commercial graphite powder (Sigma-Aldrich, <20 microns, synthetic) was ball milled for 20 h. The modified Hummers' method, which provides an oxidation reaction, was then used to prepare graphene oxide (GO). The GO was obtained as a brown powder after it was dried in the oven at 60 °C and followed by thermal reduction at 800 °C for 5 h under an inert atmosphere to obtain reduced graphene oxide (rGO). Then, melamine (Sigma-Aldrich, 99%), used as a nitrogen source, was homogeneously mixed with the prepared rGO and heat-treated at 800 °C for 1 h under N₂ atmosphere to obtain N-doped graphene.

TNG composites preparation

In the procedures for synchronous TNG preparation (Fig. S1†), a hydrothermal method was used to prepare TNG composite products. The precursors of anatase-TiO₂ (Ajax Finechem, 99%) and N-doped graphene powder in the ratio of 1 : 1 were dispersed in 10 M NaOH solution (RCI Labscan, 99%). This suspension was continuously stirred followed by sonicated at room temperature, to obtain a homogeneous suspension. The mixed suspension was transferred into a Teflon autoclave reactor with 70% filling. To study the phase transformation, these suspensions were heated at 180 °C for various reaction times of 3, 6, 9, 12, and 24 h. The hydrothermal products were washed with deionized water to remove the excess NaOH until the pH was ~7. After pH adjustment, these products were soaked in 0.1 M HNO₃ solution (Merck, 65%) for 24 h to promote the ion-exchange process. The soaked products were collected and dried at 80 °C for 12 h. The intermediate samples were heat-treated at 400 °C for 5 h under N₂ atmosphere to obtain the TNG composites. The nomenclatures used for the final products were TNG-xh, where *x* was the hydrothermal reaction time.

Materials characterization

Phases confirmation and crystallinity of the prepared products were analyzed using the powder X-ray diffraction (XRD Rigaku Miniflex II desktop diffractometer with Cu K α radiation) technique. The morphology, distribution, and crystalline phase of the synthesized products were characterized by scanning electron microscopy (SEM, JEOL JSM-7800F-Prime), and high-resolution transmission electron microscopy (HRTEM, JEOL JEM-2200FS). The electron energy loss spectroscopy (EELS) was measured by scanning transmission electron microscopy equipped with an EELS spectrometer (STEM-EELS, JEM-ARM200F). XPS analysis was conducted using an AXIS ultra DLD spectrometer. Raman spectroscopy was performed with a Jobin Yvon Horiba, model T64000. The specific surface area of N-doped graphene was obtained by Brunauer–Emmett–Teller (BET, Quantachrome Autosorb 1-MP) method. A quantitative analysis was performed by the thermogravimetric analyzer (TGA, Mettler Toledo: TGA/DSC 3+ HT1600).

Electrochemical measurement

The half-cells were fabricated in 2032 type coin cells based on the configuration of Li metal(–) | electrolyte | TNG composite(+)



with a liquid electrolyte of 1 M lithium hexafluorophosphate (LiPF_6) solution in 1 : 1 ethylene carbonate (EC)-dimethyl carbonate (DMC) + 10% fluoroethylene carbonate (FEC). Both sides of the electrodes were separated by polypropylene (PP, Celgard 2400) separator. To prepare the electrodes, active materials, and conductive Super-P (conductive carbon black, NCM HERSBIT Chemical Co. Ltd) and sodium alginate (SA) binder (Loba Chemie, 99%) were homogeneously mixed with a weight ratio of 70 : 15 : 15, which were coated onto a copper foil by a doctor blade. Then, the coated copper foil was dried in a vacuum at 100 °C for 12 h. The 2032 type coin cells were assembled in a glove box under argon atmosphere. The charge-discharge curves and cyclic stability were measured using a battery test system (Neware BTS-4000) at room temperature with the current density of 10C ($1\text{C} = 539.5 \text{ mA g}^{-1}$) and voltage cut-off between 0.01 and 3.00 V. The rate cycle performance was measured by varying the applied current density from 1C to 50C. In this work, the charge-discharge rates (C rate) of TNG anode are calculated based on the theoretical specific capacity of TiO_2 (335 mA h g^{-1}) and graphene (744 mA h g^{-1}).

Results and discussion

N-doped graphene characterizations

The XRD patterns (Fig. S2†) of graphite, GO, rGO, and N-doped graphene were measured in a range of 10° to 60° 2θ . The XRD pattern of GO represented the (001) crystal plane of GO. The experimental d -spacing of graphite was expanded after oxidation in the presence of an oxidizing agent in acid condition. This step caused several oxygens containing groups in the oxidized graphite and then generated an oxygenated graphene layer corresponded to the physical appearance of brownish GO.³⁶ The interlayers distance increased because the oxygen-containing groups were formed on each graphene layer. After the thermal reduction of GO, the XRD pattern of rGO was observed as a broadening diffraction peak at $\sim 26^\circ$. This was caused by poorly ordered graphene stacking and the expanded d -spacing to 3.70 Å. Significantly, the (002) plane of rGO was located at a slightly higher diffraction angle than that of the graphite because oxygen functional groups remained on rGO sheets. In the case of N-doped graphene, the diffraction peak of (002) crystal plane was slightly become a sharp peak compared to rGO. This suggested that nitrogen elements were doped in the graphene layers, substituting some oxygen functional groups, which remained on the graphene surface.

Raman spectroscopy was used to study the structure and electronic properties of graphene. The Raman spectra of pristine graphite, GO, rGO, and N-doped graphene (Fig. 1) represented the characteristic spectra of carbon materials including D, G, and 2D bands. Generally, ideal single-layer graphene demonstrates two dominant G and 2D bands. The phonon vibration mode of sp^2 of carbon was designated by G-band while the 2D-band was related to several graphene layers. In synthesized graphene, D-band was not only used to determine the defect of graphene sheets but also offered a clear view of defects due to doped nitrogen in the N-doped graphene. From the Raman spectra, the downshifted G peaks of the synthesized GO,

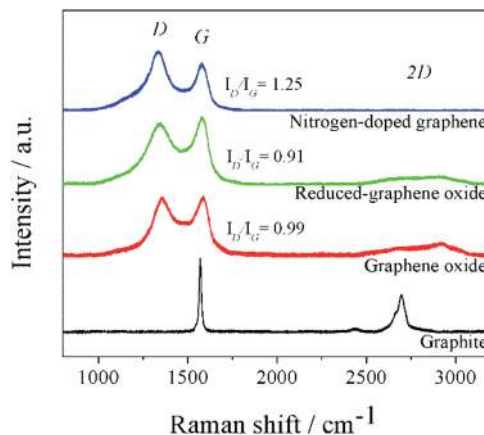


Fig. 1 Raman spectra of pristine graphite, GO, rGO, and N-doped graphene.

rGO, and N-doped graphene products were displayed at 1586, 1579, and 1575 cm^{-1} , respectively. The downshifted G peaks from GO to rGO suggested conjugated structure during the pyrolysis of the oxide functional group. Furthermore, the downshifted G peak of N-doped graphene corresponded to the electron-donating capability of N-heteroatoms.^{37,38} To estimate the disorder and defect of graphene, the intensities of D and G-bands in Raman spectra, denoted as I_D and I_G , were used to assign the I_D/I_G ratio. The I_D/I_G value of GO was 0.99. The value decreased to 0.91 for rGO because the oxygen functional groups which are distributed on GO were considered as defects. Comparing with N-doped graphene, the I_D/I_G of N-doped graphene was 1.25, which increased after the nitrogen-doping process. This suggested that N-doped graphene had more defects than rGO, indicating successful nitrogen doping on the graphene sheets. Also, nitrogen-doping induced a lot of topological defects on the graphene surface and increased disorder degree of graphene structure. These resulted in a further increase in lithium storage sites.³⁹

XPS technique was carried out to explore the binding configuration of the N-doped graphene structure. Three peaks of the XPS survey spectrum (Fig. 2(a)) were present at 284.3, 398.3, and 533.3 eV, corresponding to C 1s, N 1s, and O 1s, respectively. The O 1s peak came from the water absorption on the N-doped graphene surface, besides from oxygen residues in GO which had thermal stability.^{40,41} The high magnified XPS curve fitting of C 1s was displayed in Fig. S3† in the range of 275–300 eV. The C 1s spectrum of N-doped graphene shows fitting peaks. C 1s spectrum of N-doped graphene contained four peaks at the binding energy of 284.0, 285.0, 285.9, 286.7, 287.6, and 288.8 eV, corresponding to carbon (C–C), sp^2 carbon (C=C), N- sp^2 carbon (C=N), C–O bonds, N- sp^3 carbon (C–N) bonds and C=O bonds, respectively.^{36,42} The C–C and C=C bonds were assigned to conjugated carbons in a planar honeycomb lattice of graphene. The C–N and C=N bonds could be generated from the substitution of nitrogen atoms located at the defect or edge of the graphene structure, confirming the formation of N-doped carbon. The schematic structure of



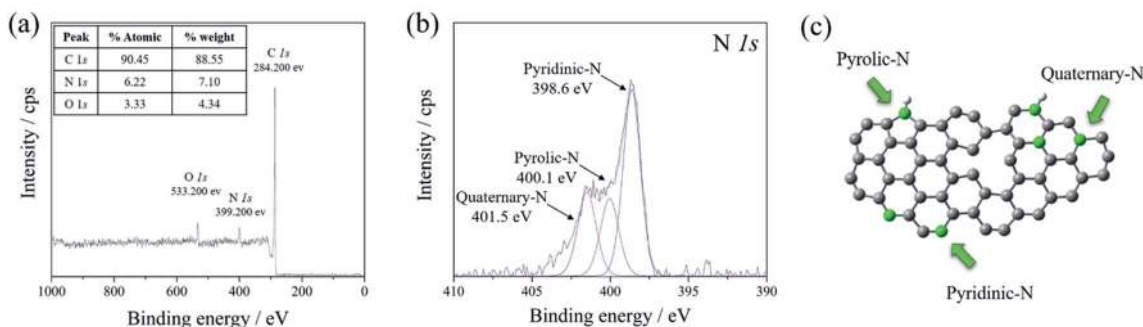


Fig. 2 (a) XPS survey spectra of N-doped graphene, and the high magnified curve fitting of (b) N 1s spectra. (c) The schematic structure of nitrogen bonding configurations of N-doped graphene.

Table 1 The comparison of specific surface areas of synthesized carbon products

	Graphite	rGO	N-doped graphene
Specific surface area, S_{BET} ($\text{m}^2 \text{g}^{-1}$)	120.00	949.10	714.03

nitrogen bonding configurations of N-doped graphene is illustrated in Fig. 2(c). The XPS N 1s spectra in Fig. 2(b) show three types of nitrogen states, including pyridinic-N (398.6 eV), pyrrolic-N (400.1 eV), and quaternary-N (401.5 eV).^{43,44} The high amount of nitrogen-doped was observed in pyridinic-N and pyrrolic-N. Both nitrogen functional groups had high electronegativity comparing to carbon which affected the cumulative escalation of the interaction between the carbon-nitrogen doped position in the carbon structure and Li-ions in the electrolyte. Moreover, the surface area of N-doped graphene was also measured at $714.03 \text{ m}^2 \text{g}^{-1}$, which was relatively high compared to that of the graphite precursor ($45.19 \text{ m}^2 \text{g}^{-1}$). These properties give the increment of the reversible specific capacity of N-doped graphene.³⁹

As the graphene-based materials are intercalation materials, a huge specific surface area is an important characteristic of synthesized graphene that determines the interface interaction between graphene and Li-ion during charge-discharge processes. To determine the specific surface area of synthesized graphene, N_2 adsorption is commonly used as the analytical technique, where the Brunauer-Emmett-Teller (BET) theory is generally used to analyze the adsorption data.⁴⁵ The specific surface areas (S_{BET}) of graphite, rGO, and N-doped graphene are shown in Table 1. The surface areas of rGO and N-doped graphene were larger than that of graphite due to the exfoliation of the graphitic layer.⁴⁶ The results further show that more single-, few-layer structures were manufactured in rGO by thermal exfoliation at high temperature, hence has a higher specific surface area. Besides, heat-treating N-doped graphene with melamine could significantly shrink the graphene interlayers, indicated by the increased $I_{\text{D}}/I_{\text{G}}$ value. The surface area difference between synthesized graphene and individual isolated graphene sheets⁴⁷ which has the-theoretical surface area of $2630 \text{ m}^2 \text{g}^{-1}$, implied that rGO and N-doped graphene were

composed of multilayered graphene sheets. This was consistent with the results of XRD, SEM, and Raman spectroscopy.

TNG nanocomposites characterizations

The physical and structure information of prepared TNG composites materials were further characterized by several techniques, including XRD, SEM, HRTEM, and STEM-EELS. As shown in Fig. 3, the XRD patterns displayed the crystalline phase of TNG in various hydrothermal reaction time conditions. The characteristic peaks at $36\text{--}38^\circ$ of anatase-phase TiO_2 (JCPDS no. 21-1272) matched very well with the TNG-3h and TNG-6h samples, suggesting that the reaction times below 6 h were not enough to transform anatase phase to $\text{TiO}_2(\text{B})$ phase (JCPDS no. 46-1238). However, the XRD patterns of TNG-9h were difficult to identify exactly as the $\text{TiO}_2(\text{B})$ phase due to the weak

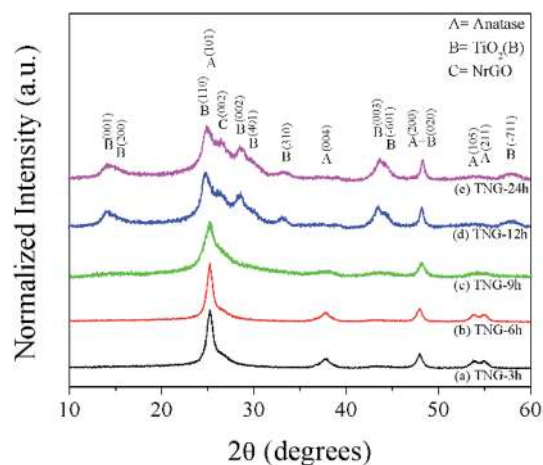


Fig. 3 XRD patterns of the prepared products. The labels of diffraction patterns for anatase, $\text{TiO}_2(\text{B})$, and N-doped graphene are shown for reference.



intensity. Therefore, the exact crystalline phase needed to be identified by other techniques, namely HRTEM and STEM-EELS. In the case of TNG-12h and TNG-24h products, their XRD patterns could be clearly indexed as the $\text{TiO}_2(\text{B})$ phase. Because the most intense peak of $\text{TiO}_2(\text{B})$ (110) and N-doped graphene (002) was at 2θ of 24.5° and 26° , respectively.^{48,49} The XRD patterns of the composites appeared as a broad peak with a shoulder of the higher angle. These indicated the formation of mixed $\text{TiO}_2(\text{B})$ and N-doped graphene in these prepared products.

To observe the morphology of the synthesized samples, SEM images of bare N-doped graphene and TNG in various reaction times are shown in Fig S4.† The surface topography of N-doped graphene (Fig S4(a)†) displays significantly delaminated graphene layers, forming a porous structure, as a result of thermal exfoliation. Normally, the main physical forms of $\text{TiO}_2(\text{B})$ products after hydrothermal synthesis represent the one-dimensional morphologies such as rod, ribbon, tube, and wire. In this experiment, TNG-3h and TNG-6h, illustrated in Fig S4(b) and (c),† respectively, appeared as the small particle agglomeration of TiO_2 , decorating on N-doped graphene layers. Therefore, these particles belong to the anatase phase, in agreement with the XRD results. Fig. S4(d)† shows the SEM image of the TNG-9h product, in which TiO_2 nanowire was well dispersed on the surface of N-doped graphene. Furthermore, the topography of TNG-12h and TNG-24h can also be seen in Fig S4(e)† and 7(f). These SEM images show the TiO_2 nanorods, entangled on the N-doped graphene surface. The nanorod structure of the $\text{TiO}_2(\text{B})$ phase was observed in the diameters of nanometres and length up to a hundred nanometres. However, the size of TiO_2 nanorods in the TNG-24h composites was larger than those in TNG-12h composite because of the increasing degree of particle agglomeration and crystal growth under long hydrothermal reaction time.

So far, the structure of TNG composites was preliminarily characterized by XRD and SEM techniques. To consolidate further insights into the microstructures of the TNG nanocomposite products, TEM techniques, and HR-TEM were performed (Fig. 4). For the TNG-3h and TNG-6h samples (Fig. 4(a) and (c), respectively) the TEM images confirmed that the TiO_2 nanoparticles were well dispersed on N-doped graphene had a nanoellipse shape with a diameter in the range of 10–20 nm. Nonetheless, a few small agglomerations of TiO_2 nanoparticles were presented in the SEM image of TNG-6h. This growth state was induced to encourage an elongation of 2D-structure when increased the hydrothermal reaction time. For this reason, the nanotube structure of TiO_2 decorated on N-doped graphene was presented in the TNG-9h product (Fig. 4(e)), corresponding to the appearance of nanowire structures in the low magnification of SEM image. The TiO_2 nanotubes diameter ranged from 10–15 nm, whereas the lengths of the nanotubes varied from 50–300 nm. The low-magnification TEM images of TNG-12h and TNG-24h (Fig. 4(g) and (i), respectively), showed intermittent nanorods dissipated on N-doped graphene sheets. The length of nanorods in both samples was longer than 1 μm . In contrast, the diameter of nanorods in the TNG-24h sample was larger

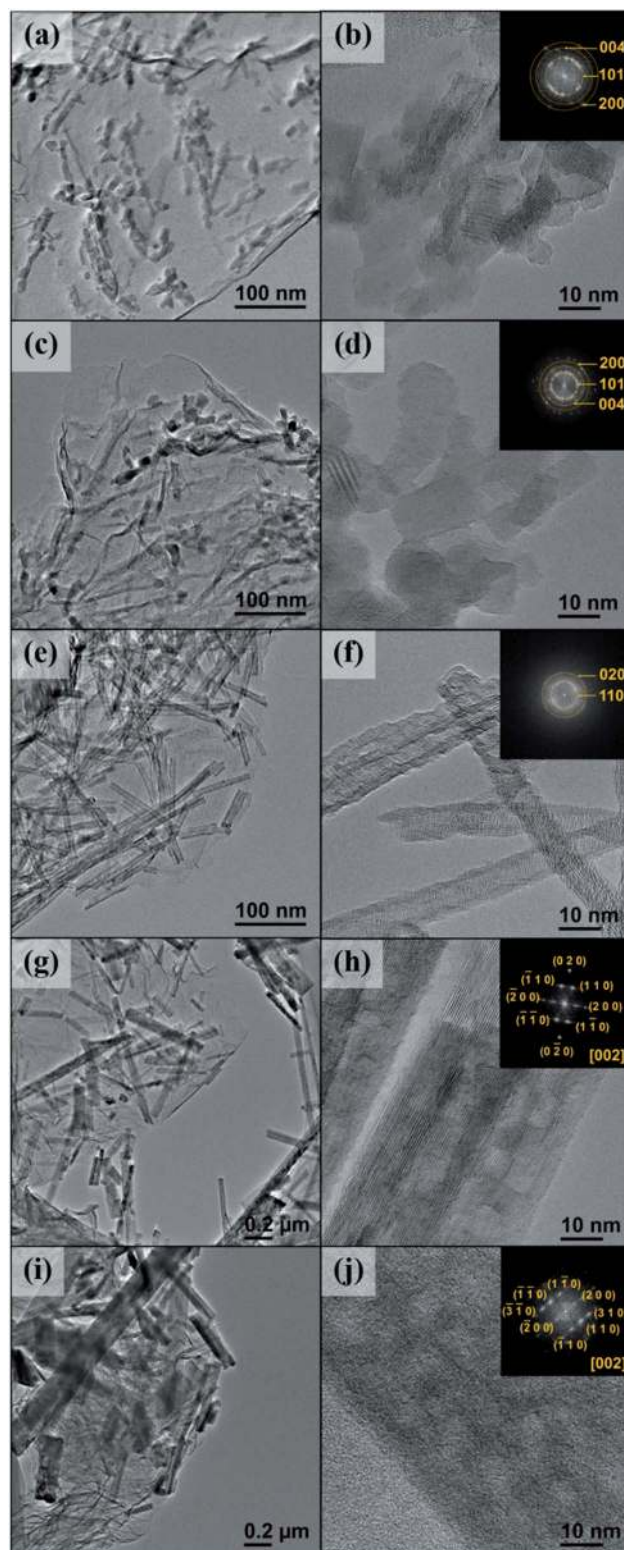


Fig. 4 Bright-field TEM and HRTEM images of TiO_2 composites on N-doped graphene and the corresponding FFT patterns inserted: (a and b) TNG-3h, (c and d) TNG-6h, (e and f) TNG-9h, (g and h) TNG-12h, and (i and j) TNG-24h.



than those in the TNG-12h, thus the reaction time played an important role in nanorods crystal growth.

The HR-TEM images of the TNG composite materials are shown in Fig. 4, where the corresponded fast Fourier transform (FFT) patterns are inserted at the top right corner. The HR-TEM images represent the high-resolution views of TiO₂ and FFT patterns which were obtained from the lattice fringe image of as-products. They could be used to determine the phase of synthesized products. From the calculated FFT patterns, the bright spots were due to the presence of highly crystalline TiO₂ on the N-doped graphene sheets. These patterns are represented by the weak continuous circle patterns.

In agreement with the XRD study, the diffraction ring of FFT patterns in Fig. 4(b) and (d) revealed that nanoparticles were anatase TiO₂ phase. Furthermore, the FFT inset of Fig. 4(f) confirmed that the nanotubes were successfully indexed as the TiO₂(B) phase. In Fig. 4(h) and (j), the presence of FFT patterns taken from the TNG-12h and TNG-24h products provided the spot diffraction, which corresponded to the [002] orientation of TiO₂(B) nanorods. To further investigate the phase transformation from anatase to TiO₂(B), all TNG composites were identified by the STEM-EELS technique, discussed in the next section.

To further investigate the TiO₂ structure of TNG composites, EELS spectra were taken from the area of as-prepared TiO₂. Fig. 5 shows Ti L_{2,3}-edge spectra taken from nanoparticles, nanotube-, and nanorod-TiO₂ of TNG products. There were two major contributions in L₂ and L₃ edges. They could be attributed to the crystal field splitting by the 2p core hole spin-orbit coupling into two sub-bands of t_{2g} orbitals (d_{xy}, d_{xz}, and d_{yz}) and e_g orbitals (d_{z²} and d_{x²-y²}) of Ti 3d state in octahedral coordination. The spectra from TNG-3h and TNG-6h showed a typical spectral feature of anatase⁴⁵ where the e_g peak of L₃-edge had shoulder at 462 eV. On the other hand, the spectra from TNG-9h, TNG-12h, and TNG-24h clearly showed only one broad peak of e_g band at 461 eV with positively skewed distribution, and e_g peaks in L₂-edge slightly shifted toward lower energies by around 1.5 eV. From the overall spectra, both TiO₂ polymorphs

represented a similar overall orbital splitting in t_{2g} which related to the point between the oxygen neighbors and their π-type bond, besides the splitting of e_g peak showed the major difference which related to the point directly toward the oxygen ligands forming p orbitals. Therefore, the e_g band of both polymorphs showed different peaks because the crystal field splitting changed when the configurations of the oxygen ligands were slightly altered in different structures noted as bond lengths and bond angles.⁵⁰ The primitive octahedral TiO₆ unit of anatase and TiO₂(B) are shown in Fig. 6. When focusing on the bond angles of O-Ti-O, the large distortion located out of the plane and main axis affected the splitting of e_g peak in both phases.

Hence, the number of different bond lengths not only contributed to the spectra but also leads to the splitting of e_g state. The Ti-O bond lengths of octahedrons in anatase and TiO₂(B) are shown in Fig. 6(a, b and c), respectively. In anatase local structure, the two-elongated Ti-O bonds along the z-axis (d_{z²}) were about 1.975 Å, while the four equatorial Ti-O bonds along the x- and y-axis (d_{x²-y²}) with a slightly smaller bond length of 1.924 Å.⁵⁴ It suggested that the two peaks of e_g splitting in L₃-edge were obviously displayed in the EELS spectra of anatase. In the case of TiO₂(B), two types of Ti⁴⁺ atom were occupied in the unit cell denoted as Ti1 and Ti2 in Fig. 6(b and c). It is well known that the crystal structure of TiO₂(B) has low symmetry due to the imperfect octahedral unit located inside. Due to the EELS spectra reflect the local symmetry, the several bond length and distorted angle of octahedral TiO₆ corresponded to an asymmetric broadening peak at 461.5 eV. Thus, this work is the first report about the novel information of the TiO₂(B) EELS spectrum which can be used for an advanced TiO₂(B) phase identification.

The model illustration of TNG composite morphology transformation from anatase to TiO₂(B) is shown in Fig. 7. Due to the increase of hydrothermal reaction time, the morphology of investigated materials can be varied as nanoparticles, nanosheet, nanotube, and nanorods. As it was mentioned previously, it can be suggested that Ti-O-Ti bonds in the TiO₂ precursor

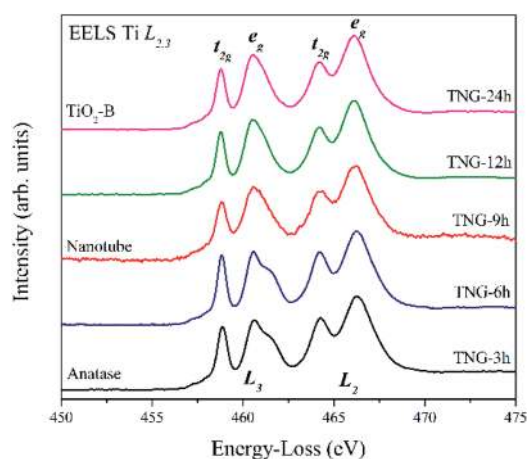


Fig. 5 EELS spectra of Ti L_{2,3}-edges taken from various morphologies of TiO₂ in the TNG samples.

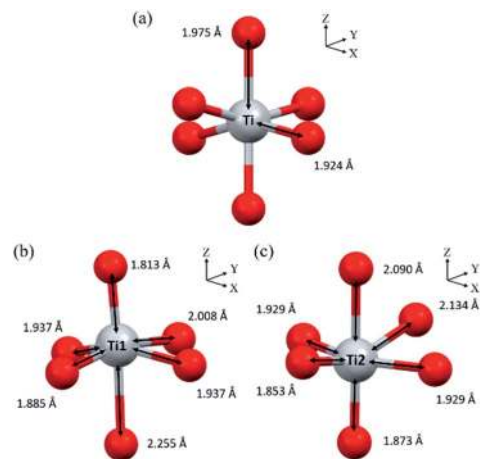


Fig. 6 Primitive octahedral TiO₆ units in (a) anatase and (b and c) TiO₂(B).



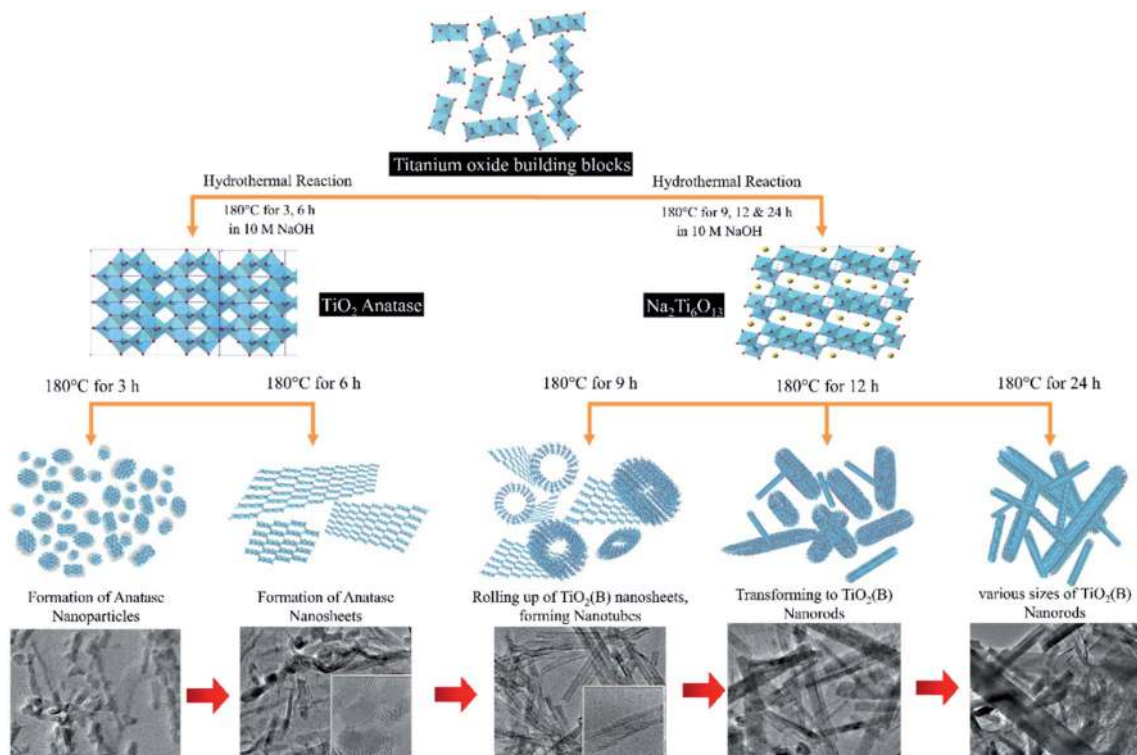
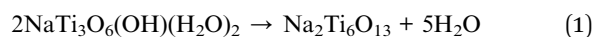


Fig. 7 Schematic diagram of TiO_2 morphology transformation with the corresponding TEM images according to reaction times of the hydrothermal synthesis.

were broken to the TiO_2 building blocks because OH^- ions which came from NaOH solution attacked the former anatase structure. Next, the TiO_2 building blocks grew during the hydrothermal reaction. In the case of short reaction time as 3 h, the anatase nanoparticles were attached to N-doped graphene sheets. When the reaction time was increased to 6 h, the small nanoparticles were closed to form nanosheets corresponding to the epitaxial crystal growth along the c -axis. When the reaction time up above 9 h, the exfoliation of the single-layered of $\text{Na}_2\text{Ti}_6\text{O}_{13}$ sheets could be promoted in the solution. The interlayer space between the sheets was intercalated by Na^+ . It can cause the neutralization of negative charge in multi-layered structures of the TiO_6 octahedral. Thus, these processes could lead to $\text{TiO}_2(\text{B})$ formation. In the presence of TNG-9h, the flat nanosheets as well as the curved ones were eventually rolled up and scroll along the b -axis to strain-induced surface free energy changes of the curvature of the nanotube walls.³⁰ Hence, the $\text{TiO}_2(\text{B})$ nanotubes were observed in TNG-9h.⁵² When the reaction time was extended to 12 h, the nanosheets roll in further to transform into $\text{TiO}_2(\text{B})$ nanorods. Finally, the nanorods were still observed in more various sizes at the reaction time of 24 h because of the $\text{TiO}_2(\text{B})$ crystal growth along with the preferred orientation as well as reducing surface energy of the nanorods.^{53,54} To explain the reaction process and phase formation of the primary intermediate products, the XRD technique was used to characterize as represented in Fig. S5.† The XRD patterns in Fig S5(a)† of the hydrothermal product indexed as a hydrated form of $\text{NaTi}_3\text{O}_6(\text{OH})(\text{H}_2\text{O})_2$. It can be suggested that

OH^- and H_2O molecules were retained inside the crystal structure because the removal of excess ion and drying processes is not enough. However, this product can be dehydrated to be a tropical $\text{Na}_2\text{Ti}_6\text{O}_{13}$ product by DI-water washing and drying in an oven. For the dehydrated mechanism, the water molecules were removed as shown in the eqn (1).⁵⁵



After the ion-exchange process, the product seems to match with the $\text{H}_2\text{Ti}_3\text{O}_7$ reference pattern. Thus, the absence of sodium titanate peaks suggests that all the Na^+ ions were replaced by H^+ ions as confirmed by the XRD results in Fig. S5(b).† The XRD patterns of the final products after heat treatment in Fig. S5(c)† can be indexed to the monoclinic crystal structure of $\text{TiO}_2(\text{B})$. These results show the phase transformation of titanate products during this procedure.

Table 2 The comparison of specific surface areas of synthesized nanocomposites

Sample	Specific surface area, S_{BET} ($\text{m}^2 \text{g}^{-1}$)	BET C -constant
TNG-3h	240.54	68.97
TNG-6h	235.09	75.86
TNG-9h	346.53	53.73
TNG-12h	188.66	78.08
TNG-24h	174.80	62.17



Table 3 The content (wt%) in the TNG nanocomposites

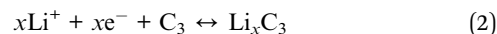
Samples	The weight content (wt%)	
	TiO ₂	N-doped graphene
TNG-3h	42.71	57.29
TNG-6h	43.29	56.71
TNG-9h	48.56	51.44
TNG-12h	48.44	51.56
TNG-24h	49.34	50.66

Besides the remarkable influence of phase composition, the morphology explained above also influenced the electrochemical properties of anode materials.⁵⁶ To realize the active sites of Li-ion interaction *via* the charge–discharge process, this work has reported the specific surface areas of synthesized TNG nanocomposites (Table 2). The values were in good agreement with the individual morphology of TiO₂. These suggested that nanoparticles of anatase in TNG-3h (240.54 m² g⁻¹) showed higher surface area than anatase nanosheet in TNG-6h (235.09 m² g⁻¹) because two-dimensional (2D) was formed by the agglomeration of small anatase nanoparticles. Because of the double-walled surface of nanotube⁵⁷ TNG-9h (346.53 m² g⁻¹) consisted of TiO₂(B) nanotube possesses a higher surface area than nanorods in TNG-12h (188.66 m² g⁻¹) and TNG-24h (174.80 m² g⁻¹) with relatively large size in one dimension of crystal growth. As these results have shown, the effect of high surface areas and unique surface energetics of synthesized materials play a major role in providing the high specific capacity and rates cycle performance.⁵⁸ Hence, the electrochemical properties of all nanocomposites were examined to gain more quantitative insight.

The successful quantitative analysis of TiO₂ on N-dope graphene is also reflected in the thermal decomposition behavior

(TGA curves) as presented in Fig. S6.† The TNG samples were heated in the range of RT–800 °C under air flow (20 ml min⁻¹) with a heating rate of 10 °C min⁻¹. In the initial TGA curves, it is believed to be due to the evaporation of physisorbed water. Then, the remaining oxygen functionalities in N-doped graphene were decomposed in the region of 150–300 °C. Thus, the samples exhibited negligible weight loss in the range of 300–500 °C, which was associated with the removal of more stable nitrogen-doped groups. Obviously, the major weight loss, represented above 500 °C, is most likely due to the pyrolysis of the carbon skeleton. Considering the TGA curved, the weight loss of TNG products was stabilized at temperature upon ~725 °C. This implied that the remaining weight after the high temperature is TiO₂, corresponding to its better thermal stability. From the TGA curves, it can indicate the intrinsic amount of N-doped graphene and titania in the as-prepared TNG products, represented in Table 3.

To reveal the correlation of Li-ion insertion/extraction reaction mechanism into the anode structure, the first three cycles galvanostatic charge–discharge (GCD) profiles were measured from TNG nanocomposites ranging from 0.01–3.00 V (*vs.* Li⁺/Li) at a current density of 1C rate (1C = 539.5 mA g⁻¹), as illustrated in Fig. 8. The GCD curves demonstrated that the prepared electrodes were initially observed small potential plateaus below ~0.5 V (*vs.* Li⁺/Li) corresponding to Li-ion intercalation into the graphene layers.^{59–61} The basic reaction mechanism of lithium-ion intercalation in the graphene electrode⁶² can be written as eqn (2):



In the case of coulombic efficiency, these electrodes exhibited relatively low coulombic efficiency at the initial charging–discharging cycle as represented in Table S2.† It could be due to

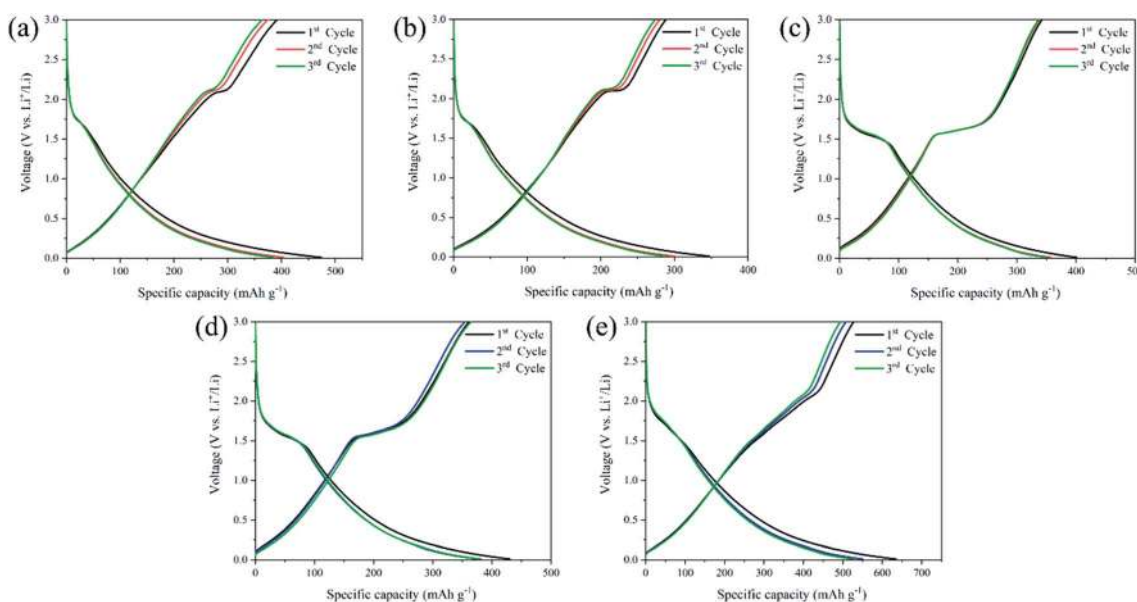
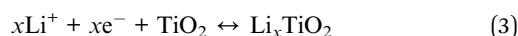


Fig. 8 GCD profiles at the first three cycles of (a) TNG-3h, (b) TNG-6h, (c) TNG-9h, (d) TNG-12h, and (e) TNG-24h nanocomposites.



Li-ion storage at irreversible sites. It can be seen that the presence of the plateau at ~ 0.6 V (vs. Li^+/Li) is usually assigned to the formation of a solid-electrolyte-interphase (SEI) film.^{63,64} Because of the large specific surface areas of the N-doped graphene sheet, this could cause one of the reasons for higher irreversible capacity due to SEI formation at the interface with the electrolyte.^{65,66} Moreover, the decomposition reaction of the electrolytes is suppressed on the TNG surface. This may be attributed to the weak interaction between electrolyte and electrode.⁶⁷ In the previous characterization, two phases of anatase and $\text{TiO}_2(\text{B})$ were mainly obtained corresponding to S-shaped voltage curves of GCD in each TNG nanocomposite sample. For these reasons, the GCD profiles were clearly observed in two forms due to the different Li-ion intercalation mechanisms of anatase and $\text{TiO}_2(\text{B})$. It is well known that the Li-ion insertion/extraction mechanism in TiO_2 electrodes can be represented as eqn (3):



The insertion coefficient (x) is related to different TiO_2 polymorphs, morphology, and crystallographic orientation. The x in anatase is usually close 0.5.⁶⁸ $\text{TiO}_2(\text{B})$ lithiates to $x = 0.7\text{--}0.8$ when lithiation with stable and slow rate cycling.^{69–71}

Obviously, the GCD profiles of TNG-3h and TNG-6h anodes in Fig. 8(a) and (b), respectively, provided the typical characteristic voltage plateaus of anatase with a clear plateau region at ~ 1.75 V (vs. Li^+/Li).³³ This reflects the phase transition from the anatase to the orthorhombic $\text{Li}_{0.5}\text{TiO}_2$ phase.⁷² Fig. 8(c–e) shows the GCD profiles of TNG-9h, TNG-12h, and TNG-24h electrodes, respectively. These electrodes obtained the phase composition of $\text{TiO}_2(\text{B})$. The dominated plateau between 1.4 and 1.6 V (vs. Li^+/Li) are corresponding to Li-ion intercalation of $\text{TiO}_2(\text{B})$.³⁰ It has been reported that the Li-ion intercalation into $\text{TiO}_2(\text{B})$ is significantly different from anatase because of a pseudocapacitive process. The Li-ion diffusion process into $\text{TiO}_2(\text{B})$ is effectively encouraged to store and transport Li-ion along with the structure.⁵⁶ Amongst the titania polymorphs, much of the argument in the $\text{TiO}_2(\text{B})$ monoclinic structure has the lowest density crystal framework (3.73 g cm^{-3}) and larger open channels and pores.⁷³

As illustrated in Fig. 9(a), three possible lithium-ion storage sites (A, B, and C sites) and lithium pathways were indicated in the schematic crystal structure of $\text{TiO}_2(\text{B})$ with a space group of $C2/m$.⁷⁴ The framework of the $\text{TiO}_2(\text{B})$ structure is composed of ReO_3 type blocks that consist of distorted TiO_6 octahedral sharing the corners of other octahedral.⁷⁵ “A” storage site act as a Li-ion diffusion pathway and favourable lithium insertion site.^{76,77} Thus, the pseudo-capacitive channel into $\text{TiO}_2(\text{B})$, paralleling to the b -axis of its perovskite-like layered structure, propose the possibility of fast Li-ion diffusion pathways.⁷⁸ Furthermore, the crystal voids in TiO_2 polymorphs, anatase (Fig. 9(b)), and $\text{TiO}_2(\text{B})$ (Fig. 9(c)) affect the amount of Li-ion storages. Therefore, the higher pores with large open channels in $\text{TiO}_2(\text{B})$ compared to anatase provided the potential to allow Li-ion intercalation into the structure in order to reach an optimum theoretical charge capacity of around 335 mA h g^{-1} . In

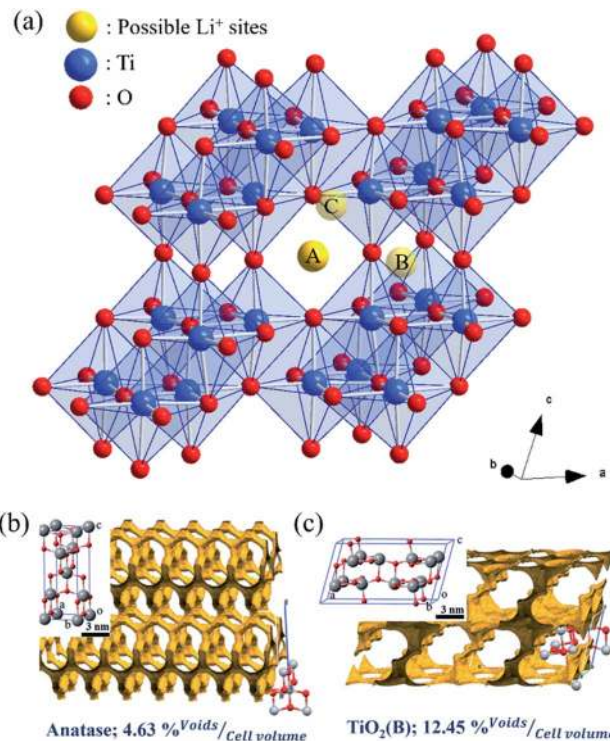


Fig. 9 (a) Three possible Li-ion storage sites (yellow atoms) in the schematic crystal structure of $\text{TiO}_2(\text{B})$ reported by Armstrong *et al.*⁷⁹ The blue atoms are Ti atoms and the red one's O atoms. The crystal voids within the structures of (b) anatase and (c) $\text{TiO}_2(\text{B})$ showing the available voids (yellow shading) and their relative percentage volume. These structures were reproduced from the report of Yothin *et al.*²⁹ with permission from The Royal Society of Chemistry.

addition, the first three discharge-charge capacities and coulombic efficiencies of TNG nanocomposite anodes at a current rate of 1C as shown in Table S1,[†] indicated the diminishing irreversible capacity losses with increasing cycles. Among all preparative TNG composites, the TNG-24h consisted of titania nanorods revealed superior initial specific capacity and delivers excellent first three coulombic efficiencies of 83.12%, 92.30%, and 94.01%, respectively. These could be due to the outstanding pseudocapacitive storage sites of the $\text{TiO}_2(\text{B})$ and the high-conductivity of N-doped graphene, which is encouraged by electronic conductivity with small resistance.

As the interdependent effects of TNG composites that are suitable to individual applications as anode for LIBs, Fig. 10(a) shows the rate capability of the prepared electrodes at various rate current densities within the range of 1C to 50C. It is clearly seen that discharge specific capacity of TNG-24h electrodes at each current rate were observed to have relatively high-capacity retention. Especially, the specific capacity of 200 mA h g^{-1} was obtained in the TNG-24h electrode at an ultrafast charge state of 50C. Then, the specific capacity was rapidly increased to 500 mA h g^{-1} without the capacity lost when the current rate returned to the initial rate at 1C, suggesting excellent rate capability as well as good cycle performance. Compared to TNG-3h, TNG-9h, and TNG-12h based composites, the rate capacities



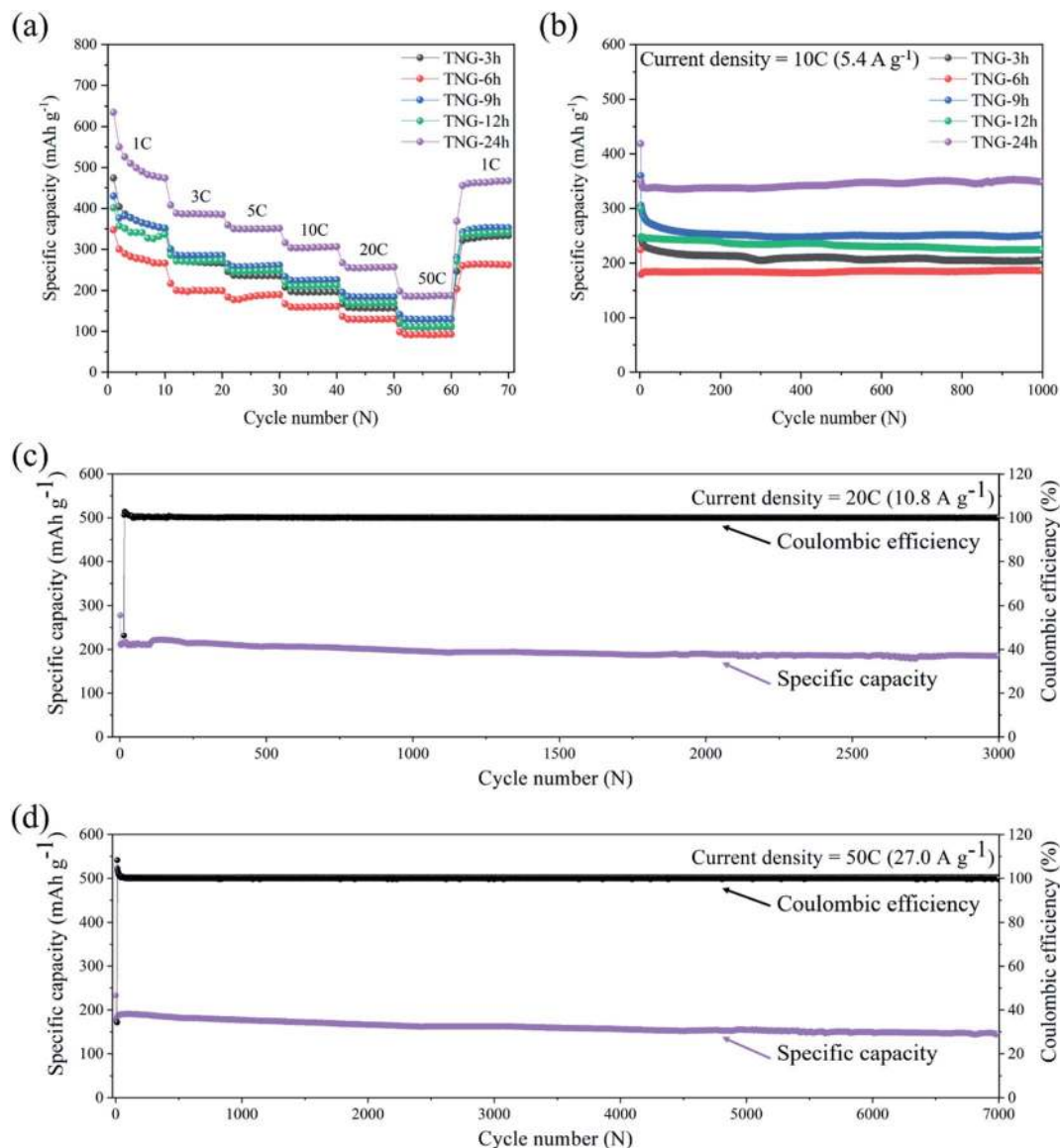


Fig. 10 Electrochemical measurements of TNG electrodes: (a) the rate cycle performance at the current density in the range of 1C to 50C. (b) The cycle stability of TNG composites at a current density of 10C for 1000 cycles and (c and d) long-term cycling of TNG-24h electrode at current densities of 20C and 50C, respectively.

were similarly obtained whereas the lowest capacity was obtained from TNG-6h.

To investigate the effect of the TiO₂ morphology on LIBs performance, the cycling ability of TNG electrodes with different reaction times was conducted at 10C up to 1000 cycles, as shown in Fig. 10(b). The composite electrodes of TNG-3h, TNG-6h, TNG-9h, TNG-12h, and TNG-24h delivered the reversible capacities of 204.87, 186.39, 250.98, 223.72, and 347.55 mA h g⁻¹, respectively. The highest capacity retention was provided by the nanorods-TiO₂(B) in TNG-24h because high crystallinity structure favored the Li-ion intercalation process. The distinctive monoclinic crystal structure of TiO₂(B) and its low-density crystal framework, exhibits larger channels and voids, and also a higher specific Li-ion storage capacity as compared to the tetragonal structures of anatase.³⁰ In the case

of nanotube-TiO₂(B) morphology in the TNG-9h sample, both electronic and Li-ion transport in the composite electrode provided short Li-ion transportation length, facilitating the faster transport of Li-ion during the charge–discharge process. Besides, the double-sided characteristics of the nanotube can promote the active surface for Li-ion intercalation in the charging process. Thus, this behavior affects the specific capacity.

In the case of anatase-TiO₂, the nanosheets-TiO₂ in TNG-6h were generated from the agglomeration of nanoparticles-TiO₂ found in TNG-3h. Because the surface area of nanosheets-TiO₂ is lower than nanoparticles-TiO₂, the Li-ion diffusion was limited by fewer active sites in the lithiation process, resulting in a lower specific capacity. The nanorods-TiO₂(B) was growth on the high surface area of N-doped graphene sheets. Thus, the



Table 4 Comparison of the capacity and cycle performance of TiO₂ and graphene-based anode materials in LIBs

Materials	Year	Specific capacity (mA h g ⁻¹)	Current density (mA g ⁻¹)	Cycle number	Ref.
TiO ₂ (B)/N-doped graphene	2020	347.6	5400 (10C)	1000	This work
		184.4	10 800 (20C)	3000	
		146.2	27 000 (50C)	7000	
TiO ₂ (B) nanowires	2019	180	100 (0.6C)	150	85
		70	10 000 (60C)	10 000	
TiO ₂ (B) hierarchical tubular structure	2015	160	5C	400	86
TiO ₂ (B) nanosheet/rGO	2014	230	1C	1000	87
		175	20C	1000	
TiO ₂ (B) nanowires/N-doped graphene	2015	220.7	10C	1000	88
Flower-like TiO ₂ (B)/Graphene	2018	171	5C	100	89
TiO ₂ (B) nanosheet/N-doped graphene	2017	326.4	0.6C	100	90
TiO ₂ /rGO	2017	79	3400 (20C)	100	91
TiO ₂ nanorods/rGO	2019	353.6	100	100	92
TiO ₂ /rGO	2017	182	500	2000	93
		150	4000	50	
TiO ₂ nanotube/Graphene	2013	80	8000	2000	94
		169	1000 (6C)	—	
N-doped TiO ₂ nanotube/N-doped graphene	2014	80	5000 (30C)	—	95
		180	400	—	
Petal-like TiO ₂ nanosheets	2014	180	400	—	96
TiO ₂ microparticles/rGO	2020	143	1C	100	97
Mesoporous TiO ₂ spheres/graphene	2018	218	0.2C	100	98
Hollow TiO _{2-x} porous microspheres	2016	151	190 (1C)	300	99
TiO ₂ /rGO	2016	200	100	50	100

electron can be transferred easily through the interconnection between TiO₂(B) and N-doped graphene. Besides, N-doped graphene in the TNG composites plays an important role in electrochemical performances. The nanosheets of N-doped graphene, which doped nitrogen element possess high electrical conductivity than bare graphene, which served as the interconnected bridges between TiO₂ particles. Therefore, the conductivity was increased, while the internal resistance was decreased in LIBs.^{80,81}

To study the long-term cycling capability, the superior current densities of 20C and 50C were applied to TNG-24h electrodes for 3000 and 7000 cycles, respectively. The capacity retention was recorded to 184.36 and 146.15 mA h g⁻¹, respectively. These indicate that the TNG-24h electrode has excellent cycling stability in order to use the ultrafast-charge application. However, the small capacity fading was observed after the 3000th cycle. This could be caused by the cracking behavior corresponded to the extension and contraction of the electrode in the charge–discharge processes.⁸²

The TNG composites with remarkable high-rate and long-life performances have a great potential to use as anodes for the next generation LIBs. Besides, the developed materials preparation technique in this work was facile processes and higher specific capacity and exhibited excellent stability at the ultra-high current density compared with some previous work of TiO₂-based materials for LIBs. Although the alloying materials, *i.e.*, Si, Sn, and Ge, have a much higher specific capacity than that commercial graphite and TiO₂. However, the use of these elements is hampered by the pulverization of the particles due to the high volumetric change during lithiation and delithiation

cycles, which leads to the destabilization of SEI film. These problems result in fast capacity fading and low coulombic efficiency. Comparing to other anodes, organic compounds are also desirable for anode materials, but the poor cycle stability and low power density limit their large-scale application.⁸³ Owing to the major weakness, its capacity at a high current density cannot be maintained due to its poor conductivity, and also the time-/energy-consuming synthetic procedure is very complicated.⁸⁴ Our report has claimed that these specific capacities, rate performances, and capacity retention are much higher than recent reports, as shown in Table 4. Obviously, the high-battery performance of the present TNG nanocomposites is among the best ones to be the most promising negative electrode materials in the next-generation of advanced LIBs.

Conclusions

TNG composites were successfully synthesized using the hydrothermal method. The effect of reaction durations on the TiO₂ phase and morphology transformation has been observed. The nanoparticle-anatase to nanorod-TiO₂(B) structures distributed on a highly conductive N-doped graphene network. The as-prepared TNG composites exhibited long cycle stability and high-rate capability when evaluated as anode materials for rechargeable LIBs. Interestingly, the TNG-24h composited with TiO₂(B) nanorods delivered a specific capacity of around 350 mA h g⁻¹ at the high current density of 10C for 1000 cycles, due to the high crystal voids and pseudo-capacitive tunnels of TiO₂(B) favor to diffuse and transport Li-ion along with the structure. This behaviour provided the best performance as



high-capacity anode materials in LIBs. In long term cycling, the TNG-24h displayed excellent cycling stability with a high specific capacity of around 150 mA g⁻¹ during ultrafast-charge of 50C. The synthesized TNG composites are considered promising anode materials for high-performance LIBs.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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