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Unique CdS@MoS₂ Core Shell Heterostructure for Efficient Hydrogen Generation Under Natural Sunlight

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The hierarchical nanostructured CdS@MoS₂ core shell was architected using template free facile solvothermal technique. More significantly, the typical hexagonal phase of core CdS and shell MoS₂ has been obtained. Optical study clearly shows the two steps absorption in the visible region having band gap of 2.4 eV for CdS and 1.77 eV for MoS₂. The FESEM of CdS@MoS₂ reveals the formation of CdS microsphere (as a core) assembled with 40–50 nm nanoparticles and covered with ultrathin nanosheets of MoS₂ (Shell) having size 200–300 nm and the 10–20 nm in thickness. The overall size of the core shell structure is around 8 μm. Initially, there is a formation of CdS microsphere due to high affinity of Cd ions with sulfur and further growth of MoS₂ thin sheets on the surface. Considering band gap ideally in visible region, photocatalytic hydrogen evolution using CdS@MoS₂ core shell was investigated under natural sunlight. The utmost hydrogen evolution rate achieved for core shell is 416.4 μmole h⁻¹ with apparent quantum yield 35.04%. The photocatalytic activity suggest that an intimate interface contact, extended visible light absorption and effective photo generated charge carrier separation contributed to the photocatalytic enhancement of the CdS@MoS₂ core shell. Additional, the enhanced hole trapping process and effective electrons transfer from CdS to MoS₂ in CdS@MoS₂ core shell heterostructures can significantly contribute for photocatalytic activity. Such core shell heterostructure will also have potential in thin film solar cell and other microelectronic devices.

In recent years, The worldwide energy crisis and eco-friendly energy demand call for the generation of advanced photocatalytic materials as one of the major solving strategy¹. In the past few decades various semiconductor photocatalyst have been demonstrated for conversion of solar energy into electrical and chemical energy due to their unique properties of semiconductor catalyst, such as high extinction coefficient of light absorption, tuneable band gap, and charge carrier multiplication effect². Recently, in order to look for clean reproducible energy, considerable attention has been paid to the abundant available natural solar light harvesting for hydrogen generation using semiconductor as a photocatalyst³. The transition metal oxide semiconductor catalyst such as ZnO and TiO₂ reported for best photocatalytic activity. Along with ZnO and TiO₂ some other oxide catalyst such as SnO₂⁴, Ta₂O₅⁵, Nb₂O₅⁶, Co₃O₄⁷, WO₃⁸, V₂O₅⁹, MoO₃¹⁰, Bi₂MoO₄¹¹, and CdMoO₄ also reported¹². But these materials have wide band gap (>3.0 eV) and show good activity only under UV light which is only 5% of the solar spectrum and did not show good photocatalytic activity in the visible light¹³. Furthermore, there are attempts on development of anion doped metal oxide nanomaterials for visible light active photocatalyst¹⁴. However, the stability of these materials limits for its commercial use as an active photocatalyst. Hence, researchers are focused their research on the development of stable and efficient visible light active photocatalyst for H₂ generation¹³.

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The transition metal chalcogenide semiconductor catalyst such as CdS reported as one of the promising semiconductor catalyst for hydrogen generation due to its narrow band gap 2.4 eV, proper valence band position and excellent stability^{15,16}. The conduction band edge of CdS is more negative than the reduction potential of H⁺/H₂, making it more suitable for the H₂ generation^{17,18}. There are reports on the synthesis of CdS/MoS₂ composites via various approaches, which shows the good photocatalytic activity as compared to pristine CdS¹⁹. However, there is still scope to improve photocatalytic activity of CdS toward water splitting by solving the photo corrosion and faster charge recombination problem^{14,20}. Therefore, researcher focused much attention mainly on fabricating nanostructure heterojunction to reduce surface recombination, enhanced visible light absorption and modifying surface defects in order to improve the water splitting activity^{21,22}. Each semiconductor in the heterostructures could be properly chosen and tailored independently to complete the designed functionality. Recently, two dimensional layered molybdenum disulfide (MoS₂) has attracted a great deal of attention in the field of photocatalysis and energy applications, due to its excellent photocatalytic performance and cycling stability²³. Also the MoS₂ and CdS have same hexagonal crystal structure, which ensures the intimate heterojunction formation. Hinnemann B. *et al.* from density functional theory calculation, revealed that free energy of atomic hydrogen bonding to MoS₂ is zero, which is very close to Pt and makes it promising catalyst for hydrogen generation²⁴. The experimental and computational results disclosed that the edges of ultrathin MoS₂ nanosheets act as an active sites for the photocatalytic H₂ generation²⁵. Yan *et al.* demonstrated the CdS/MoS₂ core shell nano rod synthesis for efficient hydrogen production by a facile chemical deposition method¹⁹.

Herein, CdS@MoS₂ core shell fabrication was demonstrated by simple solvothermal route. The hierarchical heterostructures of few layered ultrathin nanosheets of MoS₂ uniformly grown on the surface of CdS microsphere is demonstrated. The as-prepared CdS@MoS₂ core shell heterostructures show much higher photocatalytic activity and better stability toward water splitting compared with pristine CdS and MoS₂ under natural solar light. Our superior photocatalytic hydrogen generation results suggest CdS@MoS₂ core shell is promising cost effective photocatalyst for the clean energy fuel.

Experimental Methods

Material synthesis. All chemicals were purchased from Fisher Scientific with (Purity 99%). All the chemicals were AR grade and used without any further purification. Cadmium nitrate (4 mmol) and ammonium heptamolybdate tetra hydrate (4 mmol) dissolved in 50 ml methanol separately each. Drop by drop addition of nitric acid in ammonium molybdates solution with constant stirring for 10 minutes to see the clear solution. Followed by drop wise addition of dissolved cadmium nitrate solution into ammonium molybdate solution with constant stirring. After complete addition, it was allow to stir for 15 more minutes with drop wise addition of dissolved thiourea, followed by stirring for 15 more minutes. The solution was then packed in Teflon coated hydrothermal reactor and kept for 48 hours at 150 °C. After completion of reaction, reactor was allowed to cool down to room temperature naturally, followed by washing of the product with distilled water and filtered using 0.41 whatman filter paper. Further, products were washed using AR grade ethanol for several times and then kept for drying at 80 °C for 4 hours in heating oven. The dried powder catalyst again annealed in the tubular furnace at 400 °C for 3 hours in inert atmosphere. After completion of the annealing reaction, allow to cool the product naturally and used for further characterization and photocatalytic study.

Photocatalytic hydrogen generation via water splitting. The photocatalytic water splitting was carried out by using 100 ml double distilled water in 250 ml quartz reactor containing 0.1 g CdS@MoS₂ core shell catalyst. The 0.25 M Na₂S and 0.35 M Na₂SO₃ solutions were used as sacrificial reagents. Argon gas was purged through this reaction mixture in order to remove the dissolved gases. The reactor has septum arrangement to collect the evolved gas through the gas tight syringe for analysis using gas chromatography. The all assembly was arranged in an open natural solar light. The quartz reactor has arrangement for water circulation in order to absorb the IR radiation which minimizes the overheating effect. As soon as the reactor was exposed to natural sunlight, the photocatalytic splitting of water was started and generated H₂ collected in a head space of the reactor. The amount of H₂ evolved and its purity was analyzed by gas chromatograph with time. For comparative study the Photocatalytic activity for H₂ generation carried out using Xe-lamp light source (LOT ORIEL GRUPPE, EUROPA, LSH302) of intensity 300 W. The apparent quantum efficiency of the hydrogen generated was calculated using the following equation.

$$\text{AQE (\%)} = \frac{\text{Number of H}_2 \text{ molecule evolved} \times 2}{\text{Number of Incident photon}} \times 100$$

Results and Discussion

The X-Ray Diffraction technique is used to recognize the crystal structure and phase purity of the samples and results as shown in Fig. 1a. The XRD results confirm the formation of typical hexagonal phase of MoS₂. The peaks observed at 2θ = 13.9, 33.0 and 58.7 corresponding to the (002), (101) and (008) facets of hexagonal MoS₂, respectively and matches well with the previous reports of JCPDS card No. 01-075-1539¹⁴. From the XRD results, it is revealed that MoS₂ prior growth in crystallinity is at the (002) facet. The XRD (Fig. 1a) also shows the characteristics peaks of hexagonal CdS structure at 2θ = 23.1, 24.9, 26.5, 43.8, 47.8, 51.9, 66.8, 70.9, 72.4 and 75.5 which matches well with the previous reports having JCPDS card No. 01-077-2306. The Fig. 1a clearly reveals that CdS and MoS₂ are having hexagonal phase. In the XRD (Fig. 1a), MoS₂ peaks are indicated by #, whereas the CdS peaks are indicated by *.

The Raman measurements were conducted using laser wavelength of 532 nm with laser power 10 mW with a laser spot size 1 μm and the results were shown in Fig. 1b. The broad peak at ~300, 595 and 870 cm⁻¹ are assigned

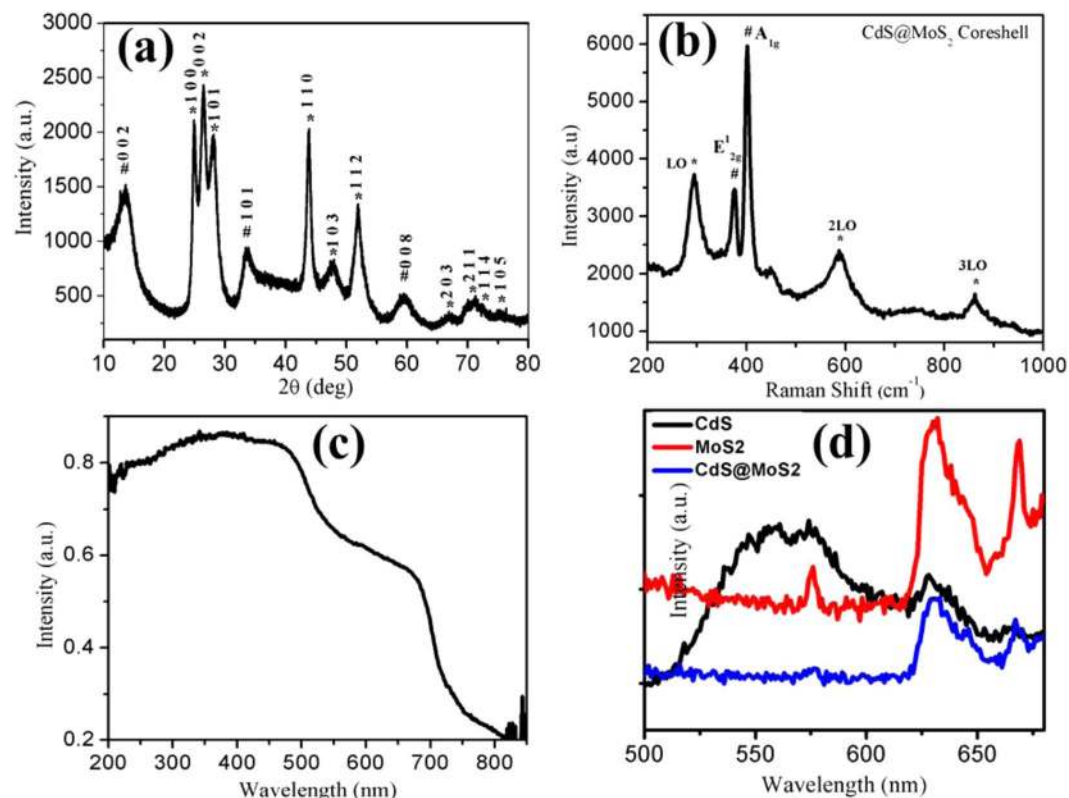


Figure 1. (a) XRD of CdS@MoS₂ core shell. MoS₂ and CdS peaks are indicated by # and *, respectively. (b) Raman spectrum of CdS@MoS₂ core shell (c) UV-Visible absorption spectrum of CdS@MoS₂ core shell (d) Photoluminescence spectrum of CdS@MoS₂ core shell material along with reference, spectrum of CdS and MoS₂.

to fundamental optical phonon mode (LO) with the first overtone mode (2LO), and second overtone (3LO) of CdS, respectively^{14,26}. Raman spectroscopy is very powerful refined technique, used to access phases present and number of layers of MoS₂ material in terms of the position and frequency difference of two characteristic vibrational modes, E₁^{2g} and A_{1g}²⁷. The CdS@MoS₂ core shell structure shows two peaks centered at 378 and 401 cm⁻¹ are assigned as an E₁^{2g} and A_{1g} vibration mode of MoS₂. The E₁^{2g} mode is attributed to the in-plane vibration of Mo and S atoms, while the A_{1g} mode is related to the out-of-plane vibration of S atoms²⁸. The frequency difference between the E₁^{2g} and A_{1g} peak of MoS₂ measured to be 23 cm⁻¹. This value is smaller than that of the bulk MoS₂, indicating presence of few layered shell of MoS₂ nanosheets on CdS core²⁹. Dravid et. al. well discussed that Raman spectroscopy has also been utilized to investigate the lattice strain and van der Waals interaction at the interface of the 2D MoS₂ nanosheet²⁸. The in-plane Raman E₁^{2g} and A_{1g} mode, is sensitive to the built-in strain of 2D MoS₂ and the reflection of interlayer van der Waals interactions, respectively²⁸. Thus, it is reasonable to predict that the shift of E₁^{2g} and A_{1g} modes of the MoS₂ in CdS@MoS₂ core shell structures towards the lower frequency side attributed to the effect of lattice strain due to the curving of MoS₂ shell^{30,31}.

As the photo absorption properties play a crucial role in determining the photocatalytic activity, UV-visible absorption spectrum of synthesized CdS@MoS₂ core shell material was recorded, and the result was depicted in the Fig. 1c. The UV-visible absorption spectrum clearly shows the two steps absorption in the visible region. The first absorption peak observed at 510 nm having band gap 2.4 eV is attributed to CdS³², whereas the second absorption peak observed at 700 nm having band gap 1.77 eV is attributed to MoS₂^{14,33}. The band gap of the material was also confirmed from tauc plot as shown in Fig. S1 (Supporting Information). The two absorption peak observed in UV-visible absorption study confirms the presence of both CdS and MoS₂. The sample absorbing more photons with high energy will produce more electron and hole with a high reduction ability to reduce the water molecules into hydrogen. Thus, the sample may exhibit a high hydrogen evolution rate.

Photoluminescence (PL) investigations are useful for the analysis of migration, transfer, and recombination processes of photo induced electron-hole pairs in a semiconductor material¹⁵. Photoluminescence measurement for CdS@MoS₂ core shell was recorded at room temperature using excitation wavelength 350 nm and results were shown in Fig. 1d. The CdS sample shows the broad emission peak centered at 560 nm attributed to the energy transition corresponding to the band gap energy. In case of MoS₂ and CdS@MoS₂ the two emission peaks are observed at 630 and 670 nm corresponds to band edge transition and defects created at the interface of the core shell and its matches well with the previous reports³⁴⁻³⁶. The more PL intensity for pristine MoS₂ and CdS sample reveals the faster electron and hole pair recombination. Whereas, the significant decline in PL intensity of CdS@MoS₂ core shell is due to the efficient photo-carrier separation by band bending (energetic gradient)

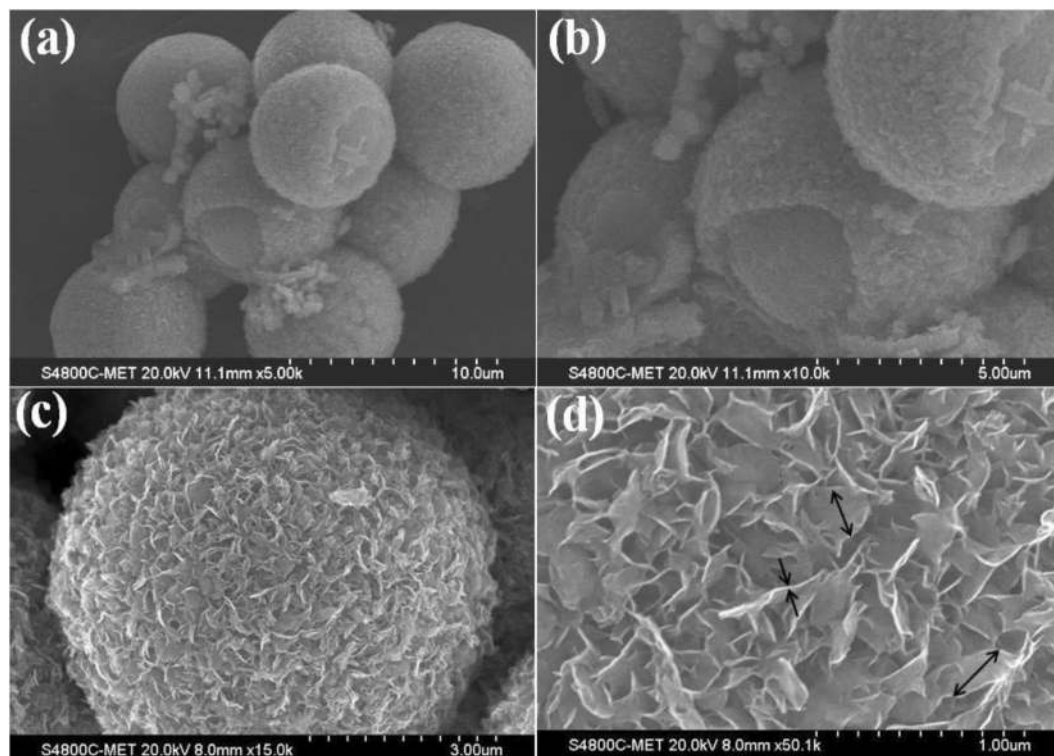


Figure 2. (a–d) FESEM images of CdS@MoS₂ core shell with various magnifications.

at hetero-structure interfaces. The CdS@MoS₂ Core shell structure facilitates the separation of photo induced electron-hole pairs due to the formation of the hetero junction between the CdS and MoS₂, and make it available for the photocatalytic activity. Here, we have optimized the structure of the core-shell to achieve the optimum performance.

The surface morphology of the synthesized CdS@MoS₂ core shell was investigated by FESEM shown in Fig. 2a–d. From FESEM images, it is clearly observed that the formation of the inner core and outer shell like morphology. The average size of the CdS@MoS₂ core shell was found to be around ~8 μm as shown in Fig. 2a. It is observed that the inner core CdS size is measured to be ~7 μm and CdS core is made from nano particles. Whereas, the thickness of the MoS₂ shell is observed to be nearly 500 nm in size as observed in Fig. 2a,b. In Fig. 2c,d, it is clearly observed that the outer shell MoS₂ is composed of small ultrathin nanosheets of size 200–300 nm and average measured thickness is around 10–20 nm as shown in Fig. 2d. It is noteworthy that such unique CdS@MoS₂ core shell structure using solvothermal process is hitherto unattempted.

Field emission transmission microscopy (FETEM) was used to further analyse the crystallinity, morphology, size of the particle and ultrathin nano sheets in the CdS@MoS₂ core shell sample. In a typical TEM image shown in Fig. 3a, it is observed the presence of inner core and outer shell of the CdS@MoS₂ core shell material. The Fig. 3a clearly reveals, CdS core is composed of particles of 40–50 nm average size. The space between the core and shell is accompanied by the ultrathin nanosheets. The outer surface of the shell covered by ultrathin nanosheets is just feather like appearance as shown in Fig. 3b. In a higher magnification image (Fig. 3c), it reveals that outer shell of MoS₂ is composed of few layered nanosheets. The d-spacing measured from the HRTEM image taken at the edge of the sheet is 0.277 nm corresponding to the (100) plane of the hexagonal MoS₂ phase, which in accordance with the XRD results. The selected area electron diffraction taken at the edge of the nanosheet confirms the polycrystalline nature of the MoS₂ shell as shown in Supporting Information Fig. S2.

The CdS@MoS₂ core shell material further examined by Scanning transmission electron microscopy (STEM) and elemental results as displayed in Fig. 4. Both Cd and Mo are observed to be uniformly distributed. The Cd signal is appeared in the inner side i.e. core (Fig. 4b). By overlapping the Cd signal, it is found that MoS₂ seats at outer shell with average thickness of around 500 nm (Fig. 4c), it is consistent of the our FESEM and TEM measurement. The S signal is observed to be uniformly distributed at inner and outer part of the CdS@MoS₂ core shell material.

X-ray photoelectron spectroscopy (XPS) was used to determine the chemical composition and chemical states of the CdS@MoS₂. Figure 5a shows two peaks at 412.9 and 406.2 eV corresponds to the characteristic binding energies of Cd²⁺ 3d_{3/2} and 3d_{5/2} in CdS, respectively¹⁹. Figure 5b displays Mo 3d peaks at 229.0 eV and 232.2 eV energy values, corresponding to the 3d_{5/2} and 3d_{3/2} core level peaks. The peak at 226.2 eV matches well with the binding energy of S 2s in sulfides²². The small shoulder peak at 235.5 eV corresponding to 3d_{3/2} of MoO₃ because of surface oxidation or may be from the starting material³⁷. The S 2p peak can be deconvoluted into two peaks at 163.0 eV and 161.8 eV (Fig. 5c), corresponds to the 2p_{1/2} and 2p_{3/2} orbital¹⁹. These binding energy values are

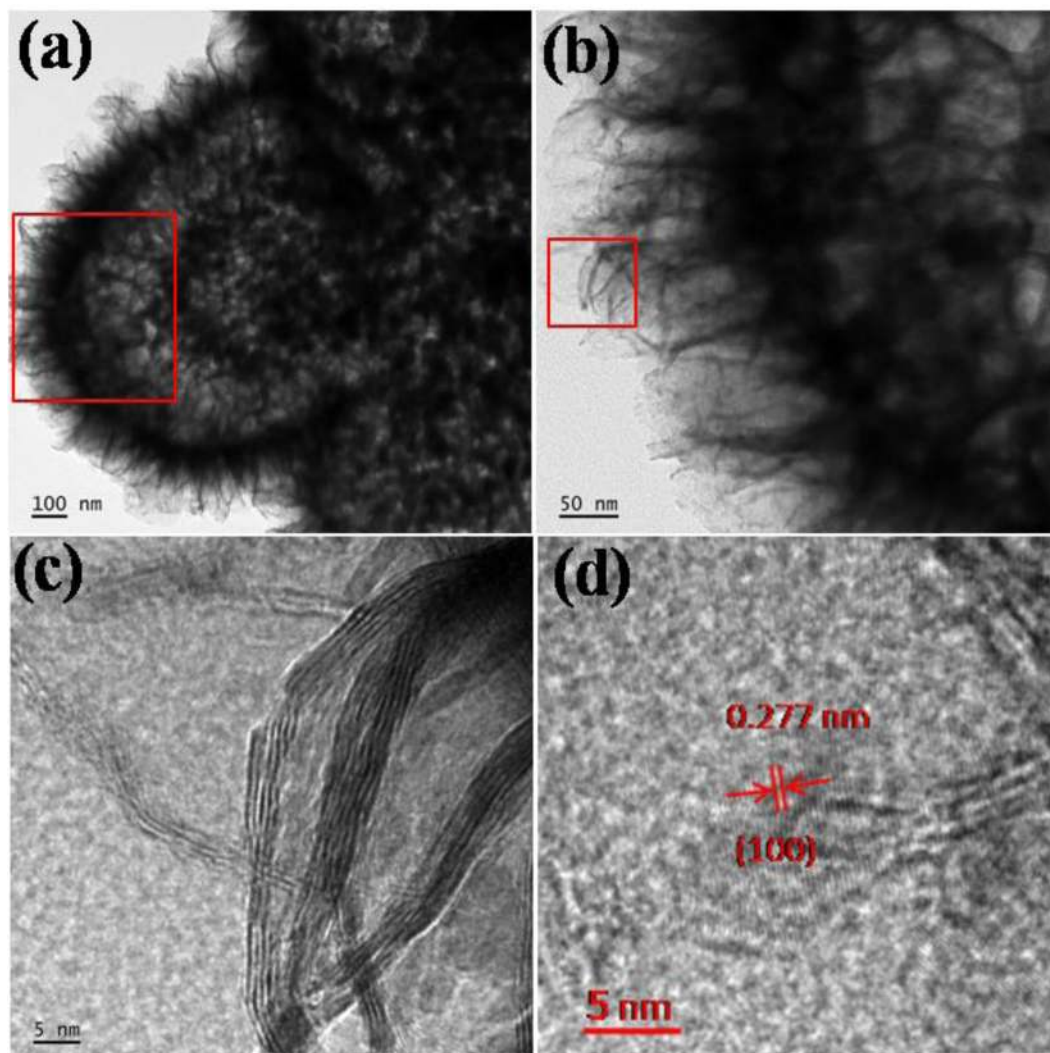


Figure 3. (a,b) FE-TEM images of CdS@MoS₂ core shell and (c,d) High resolution image.

consistent with those reported in previous studies and confirm the expected charge states of Cd²⁺, Mo⁴⁺ and S²⁻ in the CdS@MoS₂ core shell. Together with SEM and TEM observation, these results validate the CdS@MoS₂ core shell.

Formation and growth mechanism of CdS@MoS₂ core shell heterostructures. The schematic representation of growth mechanism for CdS@MoS₂ core shell material is shown in Fig. 6. The CdS@MoS₂ core shell was prepared solvothermally using cadmium nitrate, ammonium molybdates and thiourea in the methanol medium. During the solvothermal process, the cadmium, ammonium, molybdenum and sulfur ions are formed. Initially, in the hydrothermal reactor Cd²⁺ react with the S²⁻ ion to form the CdS nano particles as per the reaction 1 and shown in Fig. 6a. Initially, the CdS formation preferred due to the high reactivity of the Cd²⁺ ion towards the S²⁻ ion is more than Mo⁴⁺. Also, the energy require for the CdS formation is lower compare to MoS₂. The as formed CdS nano particle comes together in order to reduce the surface energy and formation of CdS microsphere takes place (Fig. 6b).



At solvothermal condition, the Mo⁴⁺ ions are reacting with the S²⁻ ion in supersaturated solution to form MoS₂ nuclei. These nuclei further grow by typical crystal growth mechanism and formation of layered ultrathin nanosheets takes place. Due to prolong hydrothermal reaction time, there is a uniform oriental 2D anisotropic growth of these ultrathin nanosheets on the surface of CdS microsphere to form CdS@MoS₂ core shell (Fig. 6c). This is a very distinctive structure demonstrated in this manuscript, not reported so far in the literature for the CdS@MoS₂ core shell.

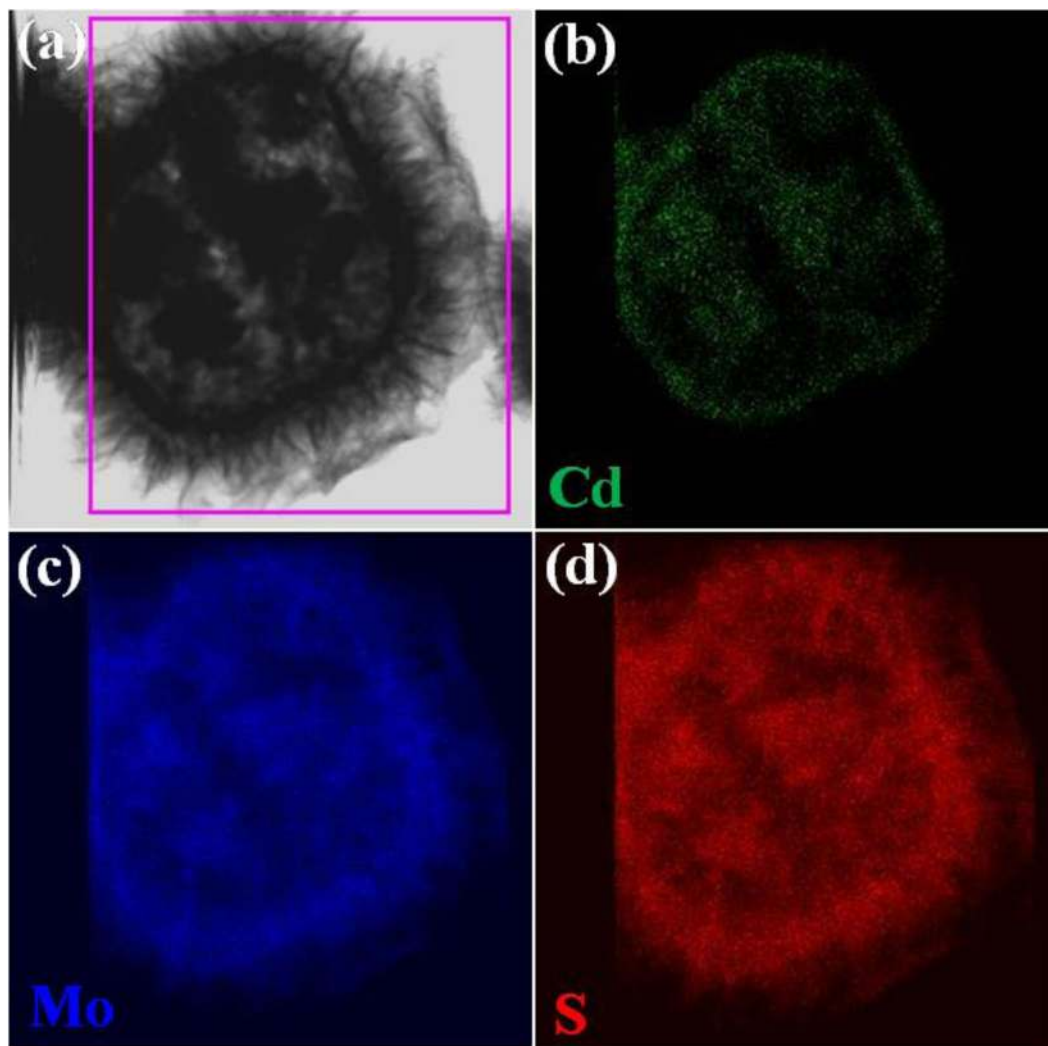
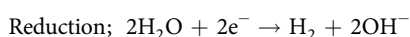
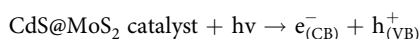


Figure 4. (a) TEM image of CdS@MoS₂ core shell and Corresponding elemental mapping STEM images of (b) Cd, (c) Mo and (d) S element.

Photocatalytic activity. The photocatalytic activity of CdS@MoS₂ core shell material was evaluated by measuring the amount of H₂ generated via photocatalytic water splitting. The photocatalytic H₂ generation activity of synthesized material was carried out in presence of solar light. In the present study, we used Na₂S/Na₂SO₃ as a sacrificial reagent. Yang *et al.* earlier discussed the use of Na₂S/Na₂SO₃ as a sacrificial reagent and its detailed mechanism. It seems that presence of Na₂S/Na₂SO₃ decreases the e⁻ and h⁺ pair recombination³⁸. The Na₂S/Na₂SO₃ suppresses the evolution of oxygen through the formation of free radicals³⁹. As the reactor containing photocatalyst is irradiated with solar light, the e⁻ in conduction band (CB) and h⁺ in valence band (VB) is generated. The photo-generated h⁺ in VB consumed by sulfide S²⁻ and sulfite SO₃²⁻ enabling the effective separation of charge carriers⁴⁰. While, e⁻ in CB reduces the two H₂O molecule into one H₂ gas molecule¹². The advantages and detailed mechanism of S²⁻/SO₃²⁻ as scavenger is well discussed in literature by Bahnemann *et al.*⁴⁰. The possible photocatalytic water splitting mechanism is as shown below.



The results of photocatalytic H₂ generated in μMole with respect to time for CdS@MoS₂ core shell material along with CdS and MoS₂ were shown in Fig. 7 and in Table 1. For comparative study the CdS and MoS₂ were synthesized at identical condition and tested for the photocatalytic H₂ generation activity via water splitting reaction under natural solar light. It is clear that CdS@MoS₂ core shell material shows significantly enhanced H₂ generation activity as compare with the pristine CdS and MoS₂. For comparative study, the cumulative H₂ generation for all the catalysts using solar simulator (xenon lamp) along with H₂ generation activity under sunlight were tested and obtained results shown in Fig. S3 and Table 1. The highest rate for H₂ generation under xenon and natural

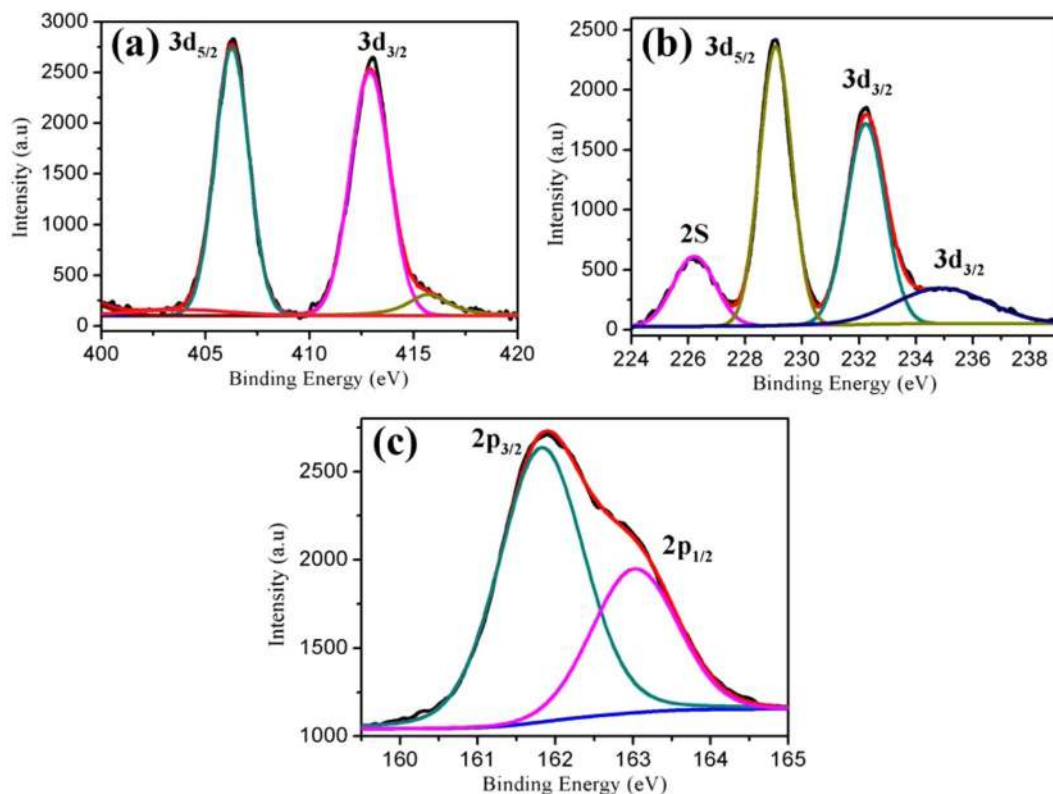


Figure 5. X-ray photoelectron spectra of (a) Cd 3d (b) Mo 3d and (c) S2p of the CdS@MoS₂ core shell sample.

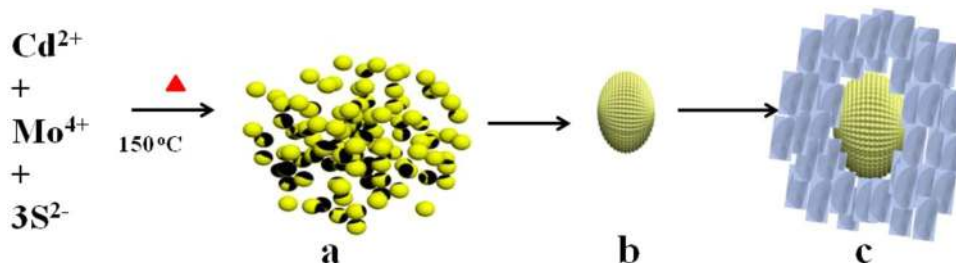


Figure 6. Schematic representation of CdS@MoS₂ core shell formation. (a) CdS nano particles (b) CdS microsphere (c) CdS@MoS₂ core shell.

solar light i.e., 231 and 416.4 $\mu\text{Mole h}^{-1}$, respectively. The photocatalytic activity results shows less H₂ generation using xenon lamp as compare to activity obtained in natural solar light. The H₂ generation achieved in solar light is 416.4, 180.6 and 53.0 $\mu\text{Mole h}^{-1}$ for CdS@MoS₂ core shell, CdS and MoS₂, respectively. It is important to note that all the photocatalysis experiments reported in this manuscript were carried out at an identical condition. The rate of H₂ generation for CdS@MoS₂ is approximately twice and six times more than CdS and MoS₂, respectively. The utmost hydrogen generation rate achieved for core shell is 4,164 $\mu\text{mole h}^{-1}$ with apparent quantum yield 35.04%, which is much better than other reports for similar materials^{22,25,41}. The higher photocatalytic activity obtained for CdS@MoS₂ core shell is due to heterostructures of CdS@MoS₂ core shell, which facilitate the photo-carrier separation and prevents carrier recombination. Consequently, the core shell heterostructures can increase large light harvesting efficiency leading to an enhanced photocatalytic activity. Furthermore, the defects created at the interface of CdS@MoS₂ core shell heterostructures, which ultimately act as a photo-carrier trap and suppresses the recombination significantly. In case of CdS@MoS₂ core shell the e⁻ and h⁺ generated due to irradiation with solar light at CdS and then they transfer to the ultrathin nanosheets of MoS₂. The 2D ultrathin nanosheets allows electron to move in the layered structure and make it available for the photocatalytic activity. The ultrathin nanosheets of MoS₂ have more number of edges, significantly its acts as active sites for the photocatalytic H₂ generation reaction. From the four hours linear H₂ generation plot, it is observed there is continuous H₂ generation in presence of solar light, concludes the good efficiency of the CdS@MoS₂ core shell catalyst. The stability of the CdS@MoS₂ catalyst was evaluated by analyzing the XRD of catalyst before and after photocatalysis

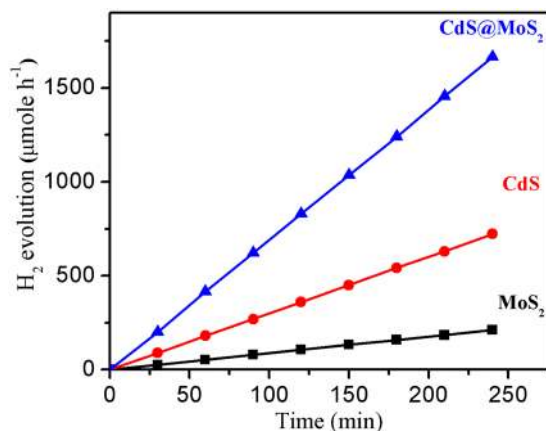


Figure 7. Photocatalytic hydrogen generation activity plot with time for CdS@MoS₂ core shell along with CdS and MoS₂.

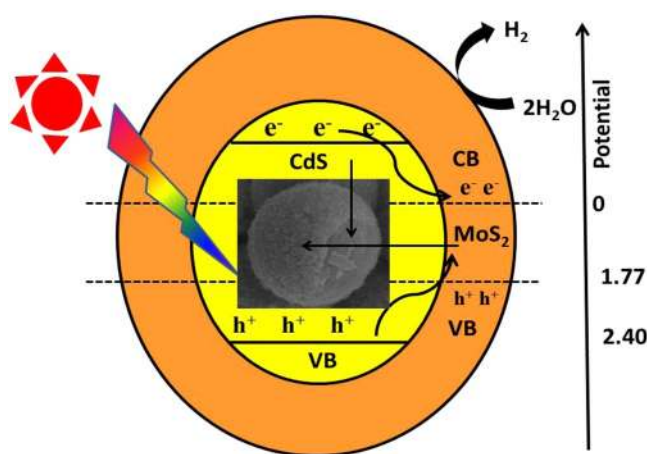


Figure 8. Schematic representation of CdS@MoS₂ core shell as photocatalyst.

Sr. No.	Catalyst used	Average H ₂ generated using xenon lamp (μmole)	Average H ₂ generated in Sunlight (μmole)	% Apparent quantum yield
1	CdS	102	180.6	15.19
2	MoS ₂	29	53.0	4.45
3	CdS@MoS ₂ Core shell	231	416.4	35.04

Table 1. Hydrogen generation via Water splitting.

(Fig. S4, Supporting Information), and the results did not show any significant change in the phase purity of the sample, which reveals the CdS@MoS₂ core shell catalyst have good stability during the photocatalytic activities. The less hydrogen generation for CdS observed is due to faster carrier recombination and photo corrosion issues in the CdS sample¹⁴.

This is the only reports which depicts the photocatalytic H₂ generation under natural sunlight catalyzed by CdS@MoS₂ core shell catalyst. The activity performance of the catalyst was established by performing reusability study in that catalyst retains its activity after the fifth recycle with negligible decrease, which reveals the reproducibility of the results as shown in Fig. S5 in Supporting Information. We did not found any H₂ production in the absence of catalyst and in the dark condition (without light). Overall, it is concluded that the hydrogen generation achieved is due to CdS@MoS₂ core shell catalyst only.

The possible schematic illustration for the photo generated e⁻ and h⁺ pair separation, its transfer process from CdS to MoS₂ and hydrogen generation mechanism for CdS@MoS₂ core shell heterostructures is proposed in Fig. 8. The enhanced photocatalytic activity for CdS@MoS₂ core shell is attributed to the effective electron transfer through the interface formed between the core CdS microsphere and ultrathin nanosheets of MoS₂ shell

as evidenced by the HR-TEM images in Fig. 3, which obviously avert the recombination of the photo generated carriers. The reduction in e^- and h^+ pair recombination of CdS@MoS₂ core shell is also evidenced by photoluminescence study as shown in Fig. 1d. Under solar light irradiation, the electrons in the valence band (VB) of the CdS core are excited to the conduction band (CB), while the holes are in the VB of CdS. The excited electron in the CB of CdS further transferred into the CB of MoS₂ due to its lower band position¹³. Whereas the holes from the VB of CdS are trapped at the interface of the CdS@MoS₂ core shell, which obviously prevents the recombination of photogenerated carriers¹⁰. The photogenerated holes (h^+) in the VB of MoS₂ and at the interface of the CdS@MoS₂ core shell consumed by sulfide S²⁻ and sulfite SO₃²⁻. Whereas, the e^- from the CB of MoS₂ take part in the reduction of two H₂O molecule into one H₂ molecule. The MoS₂ nanosheets shell also act as co-catalyst and obviously enhance the transportation of photogenerated electrons and get it available for photocatalytic activity. This fact also supports for the improvement in the photocatalytic activity of the CdS@MoS₂ core shell.

Conclusion

In the summary, a facile solvothermal approach for ultrathin MoS₂ nanosheets shell on CdS microsphere, resulting in unusual CdS@MoS₂ core shell heterostructures. We observed the photoluminescence quenching phenomenon for CdS@MoS₂. This may be attributed to the heterostructures effect resulted from the realigned band structure in CdS@MoS₂. The CdS@MoS₂ exhibited enhanced photocatalytic hydrogen generation 416.1 $\mu\text{Mole h}^{-1}$, which is much better than that of pristine CdS and MoS₂. The core shell structure of the catalysts played an important key role in the enhancement of H₂ production. The core shell heterostructures allow faster charge transportation and minimizes the recombination, which subsequently enhance the photocatalytic activity. These finding will open up new avenues for developing low cost heterostructures photo catalyst for water splitting. Our superior photocatalytic hydrogen generation results suggest CdS@MoS₂ core shell is promising cost effective photocatalyst for the clean energy fuel.

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Author Contributions

S.R.K., S.W.G. and B.B.K., Planned and designed the experiments. S.R.K., S.W.G., N.S. and B.B.K. performed experiments, collected data and calculations. S.R.K., B.B.K., C.K., A.F., analyzed data and interpreted the experimental results. All authors contributed equally to writing manuscript.

Additional Information

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