

Zircon to monazite phase transition in CeVO₄: X-ray diffraction and Raman-scattering measurements

V. Panchal,¹ S. López-Moreno,² D. Santamaría-Pérez,³ D. Errandonea,^{1,*} F. J. Manjón,⁴ P. Rodríguez-Hernandez,² A. Muñoz,² S. N. Achary,⁵ and A. K. Tyagi⁵

¹*Departamento de Física Aplicada-ICMUV, MALTA Consolider Team, Universidad de Valencia, Edificio de Investigación, C/Dr. Moliner 50, 46100 Burjassot (Valencia), Spain*

²*Departamento de Física Fundamental II, Instituto de Materiales y Nanotecnología, Universidad de La Laguna, La Laguna 38205, Tenerife, Spain*

³*Departamento de Química Física I-MALTA Consolider Team, Universidad Complutense de Madrid, Avda. Complutense s/n, 28040 Madrid, Spain*

⁴*Instituto de Diseño para la Fabricación y Producción Automatizada, MALTA Consolider Team, Universidad Politécnica de Valencia, Camino de Vera s/n, 46022 Valencia, Spain*

⁵*Chemistry Division, Bhabha Atomic Research Centre, Trombay, Mumbai 400085, India*

(Received 14 March 2011; revised manuscript received 2 May 2011; published 7 July 2011)

X-ray diffraction and Raman-scattering measurements on cerium vanadate have been performed up to 12 and 16 GPa, respectively. Experiments reveal at 5.3 GPa the onset of a pressure-induced irreversible phase transition from the zircon to the monazite structure. Beyond this pressure, diffraction peaks and Raman-active modes of the monazite phase are measured. The zircon-to-monazite transition in CeVO₄ is distinctive among the other rare-earth orthovanadates. We also observed softening of external translational T(E_g) and internal ν₂(B_{2g}) bending modes. We attribute it to mechanical instabilities of zircon phase against the pressure-induced distortion. We additionally report lattice-dynamical and total-energy calculations which are in agreement with the experimental results. Finally, the effect of nonhydrostatic stresses on the structural sequence is studied and the equations of state of different phases are reported.

DOI: [10.1103/PhysRevB.84.024111](https://doi.org/10.1103/PhysRevB.84.024111)

PACS number(s): 61.50.Ks, 61.05.cp, 78.30.-j, 71.15.Mb

I. INTRODUCTION

Recently significant research has been carried out in the field of zircon-type orthovanadates (AVO₄, A = trivalent atom) due to their wide practical applications. Orthovanadates are well known for birefringent materials and optical polarizers. These materials are potential candidates for laser-host materials and also find applications in other fields such as cathodoluminescent, thermophosphors, scintillators, phosphors, and nuclear-waste storage materials.¹⁻³ In general, the AVO₄ orthovanadates crystallize in two polymorphs, a tetragonal zircon-type structure⁴ [space group (SG): *I*4₁/*amd*] and a monoclinic monazite type⁵ (SG: *P*2₁/*n*). The zircon structure is composed of alternating edge-sharing AO₈ dodecahedra and VO₄ tetrahedra forming chains parallel to the *c* axis, while in the monazite structure AO₉ polyhedra are edge shared with VO₄ tetrahedra along the *c* axis (see Fig. 1).

The phase stability of lanthanide-based ABO₄ compounds depends on the A/B cation size ratio, and those with a large ionic radius, like LaVO₄, crystallize in the monazite structure. The compound CeVO₄, with Ce having a smaller ionic radius than La, crystallizes in the zircon structure, although it is located close to the boundary of the zircon and monazite structures. Hence, the zircon phase is expected to undergo a structural phase transition at a relatively low pressure compared with that of other orthovanadates. In view of the nature of the structural phase transitions in these compounds, it is quite imperative to understand the behavior of the zircon-structured orthovanadates under hydrostatic compression and verify whether the cation A radius plays any vital role in the structural phase transitions. As a matter of fact, most of the rare-earth orthovanadates have been observed to

undergo an irreversible zircon-to-scheelite phase transition.⁶ However, rare-earth-based zircon compounds with relatively large ionic radii, like CeVO₄, NdVO₄, and PrVO₄, have not been investigated so far.

Recent high-pressure (HP) measurements performed in orthophosphates evidence the importance of the cation A ionic radius in the sequence of structural phase transitions. Raman spectroscopy measurements and lattice-dynamic calculations supported by x-ray diffraction (XRD) in TbPO₄ indicate a zircon-to-monoclinic phase transition at 9.5 GPa.^{7,8} Similarly an XRD investigation in orthophosphates, viz. YPO₄ and ErPO₄, reports a zircon-to-monazite phase transition in both compounds.⁹ On the other hand, ScPO₄, YbPO₄ and LuPO₄ undergo a reversible zircon-to-scheelite phase transition.^{10,11} Thus it is a key issue to find what factor governs both zircon-to-scheelite and zircon-to-monazite phase transitions in orthovanadates and orthophosphates. In this paper we report XRD and Raman scattering measurements in CeVO₄, the mineral wakefieldite (Ce), up to 12.0 and 15.9 GPa, respectively, together with *ab initio* calculations.

II. EXPERIMENTAL DETAILS

CeVO₄ samples were prepared by a solid-state reaction of appropriate amounts of predried Ce₂O₃ (Indian Rare Earth Ltd., 99%) and V₂O₅ (Alfa-Aesar, 99%). Homogeneous mixtures of the reactants were pelletized and heated at 800 °C for 24 h and then cooled to ambient temperature. Further, the pellets were reground and heated again at 1100 °C for 24 h. The sample obtained was characterized by XRD as a single phase of CeVO₄ of a zircon-type structure.

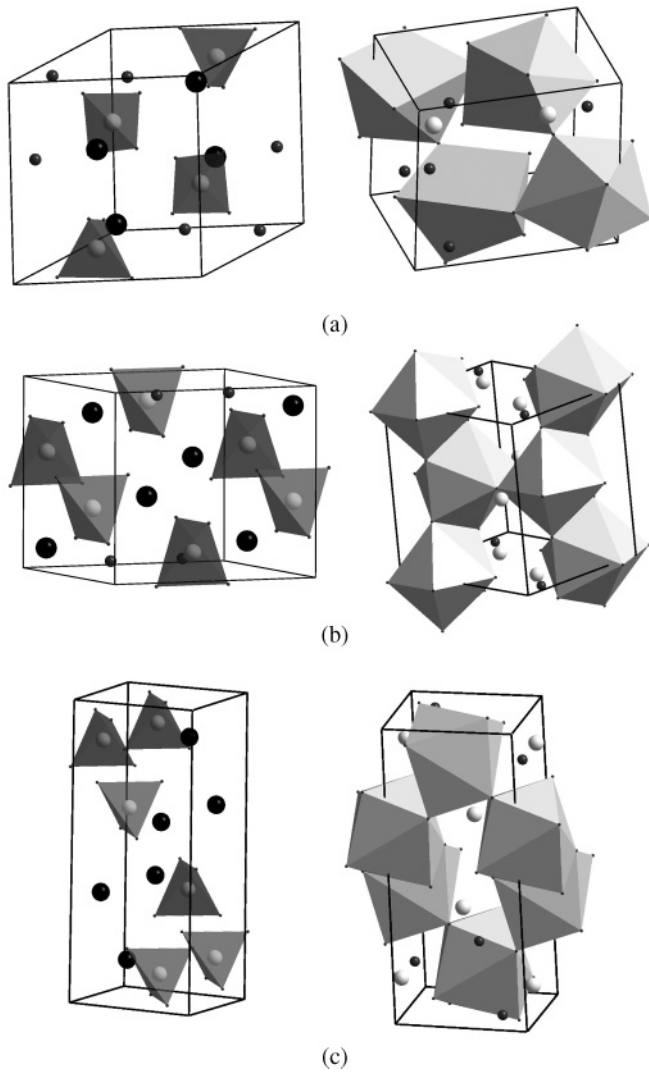


FIG. 1. (a) Schematic view of the zircon structure, (b) schematic view of the monazite structure, (c) schematic view of the scheelite structure. Black solid spheres correspond to Ce atoms, dark-gray solid spheres correspond to O atoms, and white solid spheres correspond to V atoms. The different polyhedra are illustrated in all figures (left VO_4 octahedra, right CeO_8 or CeO_9 polyhedra).

Angle-dispersive powder XRD measurements both at ambient pressure and at HP were recorded on the 135-mm Atlas change-coupled device (CCD) detector placed at 110 mm from the sample of an Xcalibur diffractometer (Oxford Diffraction Limited) using $K_{\alpha 1}$: $K_{\alpha 2}$ molybdenum radiation. The x-ray beam was collimated to a diameter of $300 \mu\text{m}$. The same setup was used previously to successfully characterize the HP phases in ABO_4 oxides up to 20 GPa.¹² On the other hand, Raman measurements both at ambient pressure and HP were performed in the backscattering geometry using a 632.8-nm He-Ne laser and a Horiba Jobin-Yvon LabRAM high-resolution microspectrometer in combination with a thermoelectric-cooled multichannel CCD detector with spectral resolution below 2 cm^{-1} .

For HP measurements on CeVO_4 , a finely grounded powder sample of CeVO_4 , along with $2\text{-}\mu\text{m}$ -diameter ruby balls, was loaded in a preindented steel gasket with a

$200\text{-}\mu\text{m}$ -diameter hole inside a diamond-anvil cell (DAC). A 4:1 methanol-ethanol mixture was used as the pressure-transmitting medium.^{13,14} The pressure was determined using the ruby-fluorescence technique.¹⁵ A modified Merrill-Bassett DAC was used for XRD measurements and a membrane-type DAC was used for Raman measurements.

We also performed compression measurements in CeVO_4 at ambient and high temperatures using steel-belted Bridgman-type opposed tungsten-carbide anvils with a tip 15 mm in diameter.¹⁶ The sample was contained in a pyrophyllite chamber which consists of two pyrophyllite gaskets of optimized thickness (0.5 mm each) in a split gasket geometry. Cubic boron nitride was the pressure medium used in these experiments. To increase the temperature we used a graphite heater.¹⁷ Temperature was measured using a steel-shielded K-type thermocouple and pressure by the calibration of the load applied to the anvils against HP resistivity transitions in Bi, Yb, CdTe, and *n*-type InSe.¹⁸ Effects of pressure in the thermocouple were neglected. Recovered samples from different experimental runs were analyzed by XRD and Raman spectroscopy.

III. THEORETICAL METHOD AND COMPUTATIONAL DETAILS

It is well known that *ab initio* methods have allowed detailed studies of the energetics of materials under HPs.¹⁹ Total-energy calculations were done within the framework of the density-functional theory (DFT), and the Kohn-Sham equations were solved using the projector-augmented wave (PAW)^{20,21} method as implemented in the Vienna *ab initio* simulation package (VASP).²² We used a plane-wave energy cutoff of 520 eV to ensure a high precision in the calculations. It is known that DFT within the local-density approximation (LDA) or the generalized-gradient approximation (GGA) often yields incorrect results for systems with *f* electrons with a small *f* orbital overlap and narrow *f* bands. The implementation of the DFT + *U* method has been found to improve the results in the study of cerium compounds.²³ The exchange and correlation energies were described within the GGA in the PBEsol prescription.²⁴ The GGA + *U* method was used to account for the strong correlation between the electrons in the Ce 4*f* shell on the basis of Dudarev's method.²⁵ In this method the on-site Coulomb interaction, *U* (Hubbard term), and the on-site exchange interaction J_H are treated together as $U_{\text{eff}} = U - J_H$. For our GGA + *U* calculations we chose a value $U = 6 \text{ eV}$ and $J_H = 1 \text{ eV}$ for the Ce atom. These values were chosen by comparison with the electronic structure study of CeVO_4 done by Da Silva *et al.*²⁶ The Monkhorst-Pack scheme was employed for the Brillouin-zone (BZ) integrations²⁷ with the grids $4 \times 4 \times 4$, $4 \times 4 \times 2$, and $4 \times 4 \times 3$ for the zircon, scheelite, and monazite phases, respectively. In the relaxed-equilibrium configuration, the forces are less than $4 \text{ meV}/\text{\AA}$ per atom in each of the Cartesian directions. Lattice-dynamics calculations of phonon modes were performed at the zone center (Γ point) of the BZ. The calculations provided information about the frequency, symmetry, and polarization vector of the vibrational modes in each structure. Highly converged results on forces are required for the calculation of a dynamical matrix of

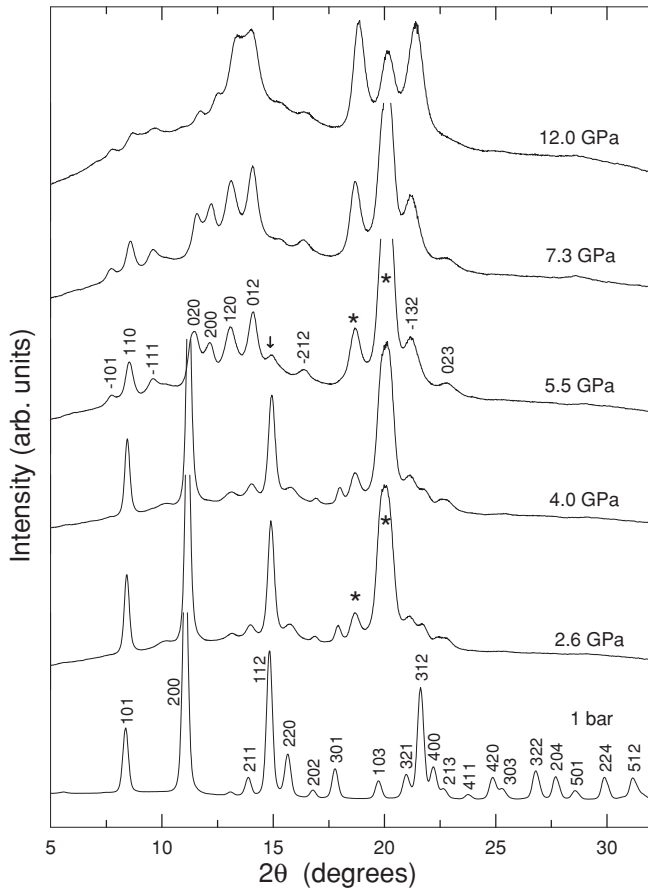


FIG. 2. Evolution of the XRD patterns of CeVO₄ as a function of pressure. Asterisks represent the diffraction peaks due to gasket. Arrow indicates the remnant (112) diffraction peak of the zircon phase.

lattice-dynamics calculations. We use the direct-force-constant approach (or supercell method).²⁸ The construction of the dynamical matrix at the Γ point of the BZ is particularly simple and involves separate calculations of the forces in which a fixed displacement from the equilibrium configuration of the atoms within the primitive unit cell is considered. Symmetry further reduces the computational efforts by reducing the number of such independent displacements in the analyzed structures. Diagonalization of the dynamical matrix provides both the frequencies of the normal modes and their polarization vectors. It allows us to identify the irreducible representations and the character of phonon modes at the Γ point.

IV. RESULTS

A. High-pressure XRD

At ambient conditions CeVO₄ crystallizes in the zircon phase (SG: $I4_1/amd$, $Z = 4$). The zircon structure can be described by CeO₈ dodecahedrons with eight similar Ce-O distances and isolated VO₄ tetrahedrons (see Fig. 1). Figure 2 shows the selective XRD diffraction patterns of CeVO₄ at representative pressures. There were no noticeable changes in the diffraction pattern up to 4 GPa, and the diffraction peaks could be indexed to the zircon phase. At 5.5 GPa the appearance of many extra diffraction peaks was observed along

TABLE I. Lattice parameters and fractional coordinates of CeVO₄ in the zircon phase at ambient conditions. SG $I4_1/amd$, $Z = 4$, $a = 7.399$ Å, $c = 6.482$ Å, and residuals $R_p = 11.66$, $R_{wp} = 11.24$.

Atoms	Sites	x	y	z
Ce	4a	0	0.7500	0.1250
V	4b	0	0.2500	0.3750
O	16h	0	0.4264	0.2134

with the weak remnant (112) diffraction peak of the zircon phase, as shown by the arrow in Fig. 2. These changes in the diffraction patterns are indicative of a structural phase transition in CeVO₄ at this pressure. The new diffraction peaks of CeVO₄ could be assigned to the monoclinic monazite phase ($P2_1/n$). On further increase of pressure the monazite phase was found to be stable up to 12 GPa, which is the highest pressure reached in our XRD measurements. On release of pressure the monazite phase was quenched, thus indicating the irreversible nature of the zircon-to-monazite phase transition. These results are in agreement with Raman measurements and calculations presented in the next sections.

The Rietveld refinement of all the background-corrected diffraction patterns was carried out using Powdercell software.²⁹ The diffraction patterns up to $2\theta = 18^\circ$ were used to avoid any kind of gasket interference in the structural refinements. At ambient pressure the lattice parameters for zircon-structured CeVO₄ are refined as $a = 7.399$ Å and $c = 6.482$ Å, which are consistent with the Inorganic Crystal Structure Database (ICSD) card No. 66033. The lattice parameters and fractional coordinates of the zircon phase of CeVO₄ are given in Table I and are also in good agreement with those reported in previous HP and high-temperature XRD measurements by Range *et al.*,³⁰ $a = 7.383$ Å, $c = 6.485$ Å.

The pressure evolution of the lattice parameters and equation of state of CeVO₄ in the zircon and monazite phases is shown in Fig. 3. It can be seen that in the zircon phase the a axis is more compressible than the c axis, as is evident from the increase of the c/a ratio from 0.878 at ambient pressure to 0.887 at 4 GPa. The linear compressibility of the c axis ($K_c = 1.05 \times 10^{-3}$ GPa⁻¹) is smaller than that of the other two axes ($K_a = K_b = 3.48 \times 10^{-3}$ GPa⁻¹). These values are similar to those obtained in other zircon-type vanadates⁶ and related oxides.³¹ Due to deterioration of diffraction patterns in the monazite phase beyond 5.5 GPa, the value of β angle refined at 5.5 GPa was kept constant for further refinements. A typical Rietveld refinement fit of diffraction data at 7.2 GPa is shown in Fig. 4, and the lattice parameters and fractional coordinates are presented in Table II. These parameters are in close agreement with lattice parameters reported in recent XRD measurements in monazite CePO₄.³² In the monazite phase we observed an anomalous behavior of the axial compression. From Fig. 4 it can be seen that the a and b axes decrease with pressure while the c axis increases with pressure. A similar behavior was observed in pressure-induced monazite structured for orthophosphates.^{8,9} However, this behavior is in contrast with results in monazite-type CePO₄ and LaPO₄.^{9,32} The different axial compression of CeVO₄ and CePO₄ could be attributed to the more distorted monazite structure of CeVO₄ than that of CePO₄. The reason is that the monazite phase of CeVO₄

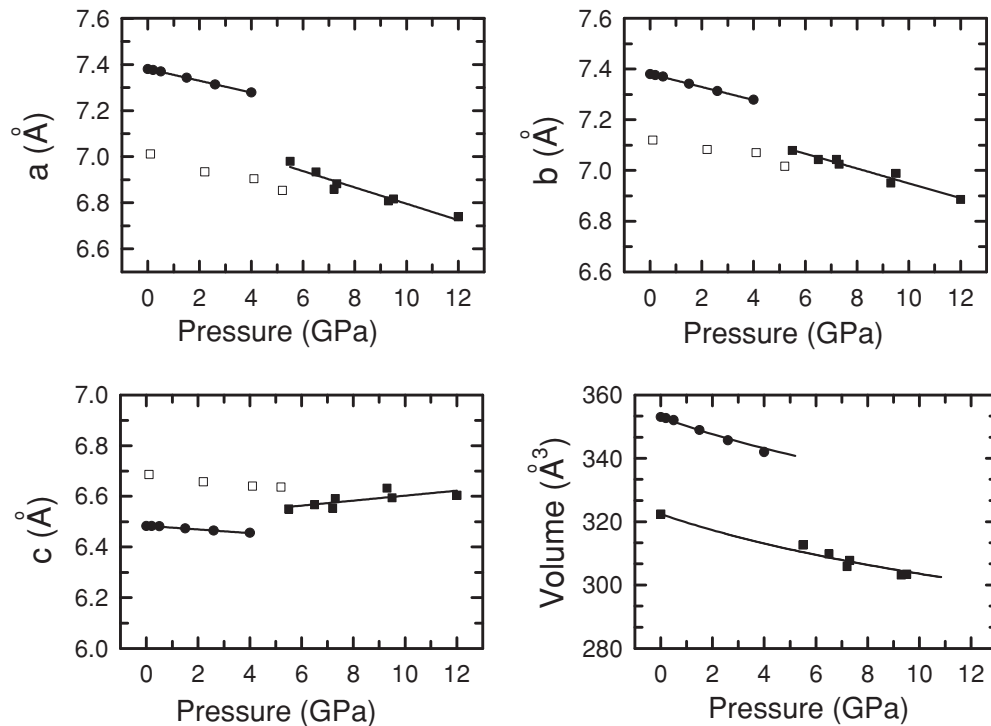


FIG. 3. Pressure dependence of the lattice parameters and volume. Filled circles (squares) correspond to the zircon (monazite) phase of CeVO_4 . The solid lines in the lattice-parameter plots are linear fits of the data, and the solid lines in the volume data correspond to the third-order Birch–Murnaghan equation of state.

is induced by pressure while that of CePO_4 occurs already at ambient pressure. Note that difference is found always when a HP monazite is compared with an ambient-pressure monazite. The zircon-to-monazite phase transition in CeVO_4 exhibits a volume collapse of 8.6% at 4 GPa which is quite large compared with those observed in orthophosphates.^{8,10} A fit of the volume vs. pressure data of the zircon phase to a second-order Birch–Murnaghan equation of state ($B'_0 = 4$) gives a bulk modulus of 118.9 GPa. This value is the smallest one found in orthovanadates,⁶ thus indicating that CeVO_4 is the most compressible orthovanadate to date. In the monazite phase, volume vs. pressure data fitted to a

third-order Birch–Murnaghan equation of state give bulk moduli $B_0 = 142$ GPa and $B'_0 = 4.4$, which are relatively close to the bulk modulus of monazite-type CePO_4 ($B'_0 = 122$ GPa).³² During the fitting we excluded the pressure-volume data point collected at 12 GPa because it deviated from the systematic behavior of the rest of the data. A possible reason for this might be the assumption that the β angle does not change under compression, which could lead to a unit-cell volume miscalculation.

The occurrence of the zircon-to-monazite transition instead of the zircon-to-scheelite transition is apparently in contradiction to the results reported by Range *et al.*³⁰ These authors, who used a large-volume press, found that, upon compression at room temperature zircon-type CeVO_4 transforms to the scheelite phase. They observed the occurrence of the monazite phase only under the combined effect of pressure and

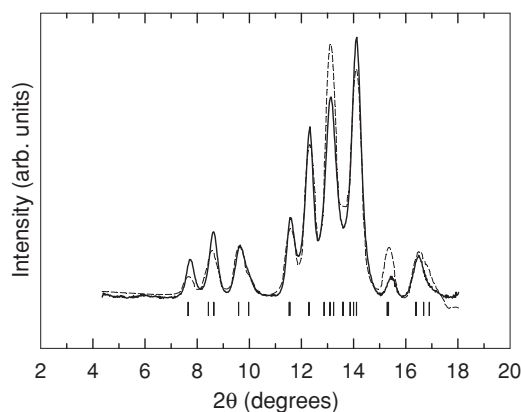


FIG. 4. Observed and calculated XRD diffraction patterns for the monazite ($P2_1/n$, $Z = 4$) phase of CeVO_4 at 7.3 GPa. Bars indicate the expected positions of diffraction peaks.

TABLE II. Lattice parameters and fractional coordinates of CeVO_4 in the monazite phase at 5.5 GPa. SG $P2_1/n$, $Z = 4$, $a = 6.980$ Å, $b = 7.079$ Å, $c = 6.550$ Å, $\beta = 105.3^\circ$, and residuals $R_p = 32.46$, $R_{wp} = 37.99$.

Atoms	Sites	x	y	z
Ce	4e	0.2818	0.1591	0.1000
V	4e	0.3047	0.1635	0.6124
O ₁	4e	0.2508	0.0055	0.4458
O ₂	4e	0.3811	0.3320	0.4982
O ₃	4e	0.4745	0.1054	0.8042
O ₄	4e	0.1268	0.2164	0.7108

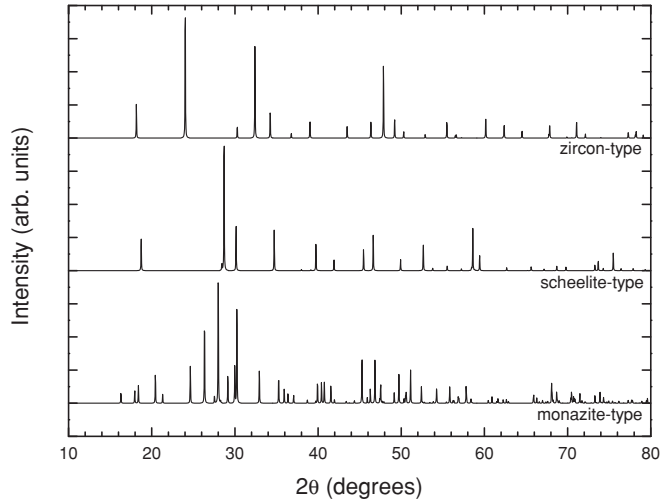


FIG. 5. XRD patterns collected from different phases of CeVO₄ recovered after experiments in the Bridgman cell.

temperature. To corroborate these results we performed similar experiments using a large-volume press equipped with Bridgman anvils. The obtained results from XRD measurements performed from samples quenched from different pressures are shown in Fig. 5 and tabulated in Table III. Our results show that zircon-type CeVO₄ transforms to the scheelite phase at 4 GPa and to the monazite phase at the same pressure but at temperatures close to 600 °C. These results confirm the results reported by Range *et al.*³⁰ The main difference between these experiments and the experiments performed using a DAC is that in the Bridgman cell a solid-pressure medium is used; i.e., the nonuniaxial stresses are more important. Therefore, our results evidence that zircon-type CeVO₄ undergoes a transformation to the scheelite structure at ambient temperature under nonhydrostatic compression. At high temperatures, stresses are relaxed and therefore the monazite phase can be obtained as in the DAC experiments. It is interesting to note that, in the Bridgman-cell experiments, both the monazite and the scheelite phases are quenchable to ambient conditions, which confirms the nonreversibility of both transitions. The unit-cell parameters and atomic position of monazite-type and scheelite-type CeVO₄ at ambient conditions are summarized in Table IV. For the sake of comparison the lattice parameters of the monazite^{30,33} and scheelite phases³⁰ reported in earlier measurements are also given.

TABLE III. Results obtained from Bridgman-cell experiments.

Experiment number	Temperature (°C)	Pressure (GPa)	Recovered sample phase
Run 1	25	2	Zircon
Run 2	600	2	Zircon
Run 3	25	4	Scheelite
Run 4	600	4	Monazite
Run 5	25	6	Scheelite
Run 6	600	6	Scheelite

B. Structural calculations

The calculated total energy (E) as a function of volume of three different phases of CeVO₄ is shown in Fig. 6(a). According to the calculations, the most stable phase at ambient pressure is zircon. The calculated values of lattice parameters, fractional coordinates, and bulk moduli for all three phases are listed in Table V. These values are in good agreement with the experimental ones. The bulk modulus of the monazite phase is slightly underestimated by the calculations. In particular, the calculated B₀ for monazite is 2% smaller than in low-pressure zircon, which is unusual for a HP phase. The thermodynamic phase transition between two structures occurs when the Gibbs free energy (G) is the same for both phases. To obtain the Gibbs

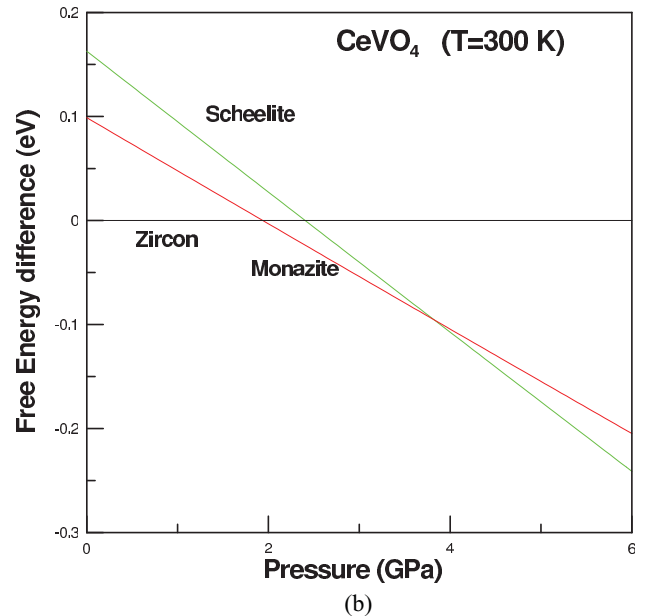
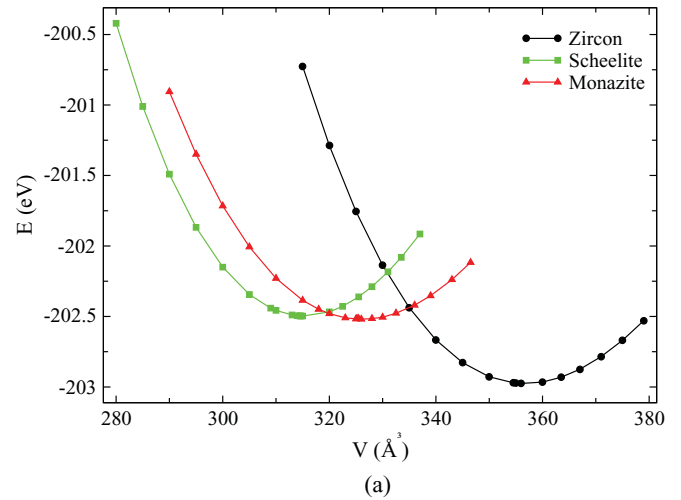


FIG. 6. (Color online) (a) Variation of total energy, at T = 0 K, as a function of volume for zircon-, scheelite-, and monazite-type CeVO₄. Filled circles, squares, and triangles correspond to zircon, monazite, and scheelite phases, respectively. (b) Plot of the free energy, at T = 300 K, versus pressure for the scheelite and the monazite phases. The free energy of the zircon phase has been taken as a reference.

TABLE IV. Lattice parameters of CeVO_4 in the scheelite phase and monazite phase obtained from Bridgman-cell compression with temperature variations. For comparison purpose lattice parameters reported by Range *et al.*³⁰ and Yoshimura and Sata³³ are also given.

Monazite phase ($P2_1/n$)			Scheelite phase ($I4_1/a$)	
This paper	Ref. 30	Ref. 33	This paper	Ref. 30
$a = 7.001 \text{ \AA}$	$a = 7.003 \text{ \AA}$	$a = 6.98 \text{ \AA}$	$a = 5.163 \text{ \AA}$	$a = 5.165 \text{ \AA}$
$b = 7.221 \text{ \AA}$	$b = 7.227 \text{ \AA}$	$b = 7.22 \text{ \AA}$	$c = 11.849 \text{ \AA}$	$c = 11.848 \text{ \AA}$
$c = 6.703 \text{ \AA}$	$c = 6.685 \text{ \AA}$	$c = 6.76 \text{ \AA}$		
$\beta = 105.07$	$\beta = 105.13$	$\beta = 105.02$		

free energy we use a quasi-harmonic Debye model that allows obtaining G at room temperature from calculations performed for $T = 0 \text{ K}$.³⁴ Figure 6(b) shows the pressure dependence of the Gibbs free energy difference at $T = 300 \text{ K}$ for the scheelite and the monazite phases with respect to the zircon phase, which is taken as reference. In calculations, at nonzero temperatures all vibrational contributions are included, in contrast with 0 K calculations which neglect the vibrational contribution: zero point energy. The calculations show a phase-transition sequence with the formation of the monazite phase being prior to that of the scheelite phase, in good agreement with our experimental findings. Calculations also indicate that the monazite phase may transform to the scheelite phase at HPs. At 300 K , the calculated phase-transition pressure for the zircon-to-monazite transition is 2 GPa , and the theoretical phase-transition pressure for the monazite-to-scheelite transition is 3.9 GPa . Calculations done at 0 K and 600 K gave the same structural sequence, being the zircon-to-monazite transition pressure not affected by temperature. However, the monazite-to-scheelite transition pressure increases with temperature, in agreement with the experimental results reported by Range *et al.*³⁰ This positive slope can be rationalized by using the Clausius–Clapeyron equation that relates differences in volumes (ΔV) and entropy (ΔS) across a phase boundary to the slope of the phase boundary in P-T space: $dP/dT = \Delta S/\Delta V$.³⁵ Since the volume per formula unit is smaller in scheelite than in monazite CeVO_4 , the positive slope of the monazite-to-scheelite phase boundary indicates that the transition is also associated with an entropy decrease. To conclude this section we would like to mention that, at 300 K , the calculated zircon-to-scheelite transition pressure agrees with the DAC experiments; however, in our experiment, we recorded data till 16 GPa and we have found no evidence for the monazite-to-scheelite phase transition. A possible cause for the nonoccurrence of the monazite-to-scheelite transition is the presence of kinetics barriers, as observed in other ABO_4 compounds.³⁶ Consequently, the atomic displacements needed to trigger the monazite-to-scheelite transition can take place only if the equilibrium transition pressure is sufficiently overstepped. New XRD measurements at pressures beyond 15 GPa and at different temperatures are needed to check the possibility of the second phase transition. Note that under nonhydrostatic compression the scheelite phase is found at 4 GPa , which supports the theoretical results. Apparently, nonhydrostatic stresses reduce kinetic barriers in CeVO_4 . However, this subject is still open to new studies.

C. Raman scattering

1. Zircon-structured CeVO_4

At ambient conditions, CeVO_4 exists in the zircon structure (space group $I4_1/amd$ and point group D_{4h}) with two formula units per primitive cell. Group-theoretical analysis predicts 12 Raman-active modes $2A_{1g} + 4B_{1g} + B_{2g} + 5E_g$.³⁷ These modes can be further classified into internal (ν_1 - ν_4) and external (translational T and rotational R) modes of VO_4 units as follows:

$$\Gamma = A_{1g}(\nu_1, \nu_2) + B_{1g}(2T, \nu_3, \nu_4) + B_{2g}(\nu_2) + E_g(2T, R, \nu_3, \nu_4). \quad (1)$$

Figure 7(a) shows the Raman spectra of CeVO_4 in the zircon phase at different pressures up to 4.5 GPa . Eight Raman modes have been observed at ambient condition out of 12 Raman peaks predicted for CeVO_4 in the zircon phase. The symmetry assignment for the Raman modes has been performed in accordance with our calculations and the comparison with previous results in other vanadates, and it is summarized in Table VI.

TABLE V. Calculated lattice parameters, bulk moduli, and fractional coordinates of CeVO_4 in the zircon phase at ambient conditions (SG $I4_1/amd$, $Z = 4$, $a = 7.423 \text{ \AA}$, $c = 6.461 \text{ \AA}$, and the bulk modulus $B_0 = 115.24 \text{ GPa}$, $B'_0 = 4.83$); in the monazite phase at 5.77 GPa (SG $P2_1/n$, $Z = 4$, $a = 6.869 \text{ \AA}$, $b = 7.119 \text{ \AA}$, $c = 6.56 \text{ \AA}$, $\beta = 104.88^\circ$, and the bulk moduli $B_0 = 109.63 \text{ GPa}$, $B'_0 = 3.24$); and in the scheelite phase at ambient conditions. (SG $I4_1/a$, $Z = 4$, $a = 5.165 \text{ \AA}$, and $c = 11.795 \text{ \AA}$, and the bulk moduli $B_0 = 138.65 \text{ GPa}$, $B'_0 = 3.84$.)

Phase	Atoms	Sites	x	y	z
Zircon	Ce	4a	0	0.7500	0.1250
	V	4b	0	0.2500	0.3750
	O	16h	0	0.4289	0.2053
Monazite	Ce	4e	0.2865	0.1581	0.1086
	V	4e	0.3002	0.1681	0.6184
	O ₁	4e	0.2393	0.9971	0.4264
	O ₂	4e	0.3883	0.3489	0.4949
	O ₃	4e	0.4902	0.1113	0.8327
Scheelite	O ₄	4e	0.1158	0.2269	0.7386
	Ce	4b	0	0.2500	0.6250
	V	4a	0	0.2500	0.1250
	O	16f	0.2500	0.1182	0.0463

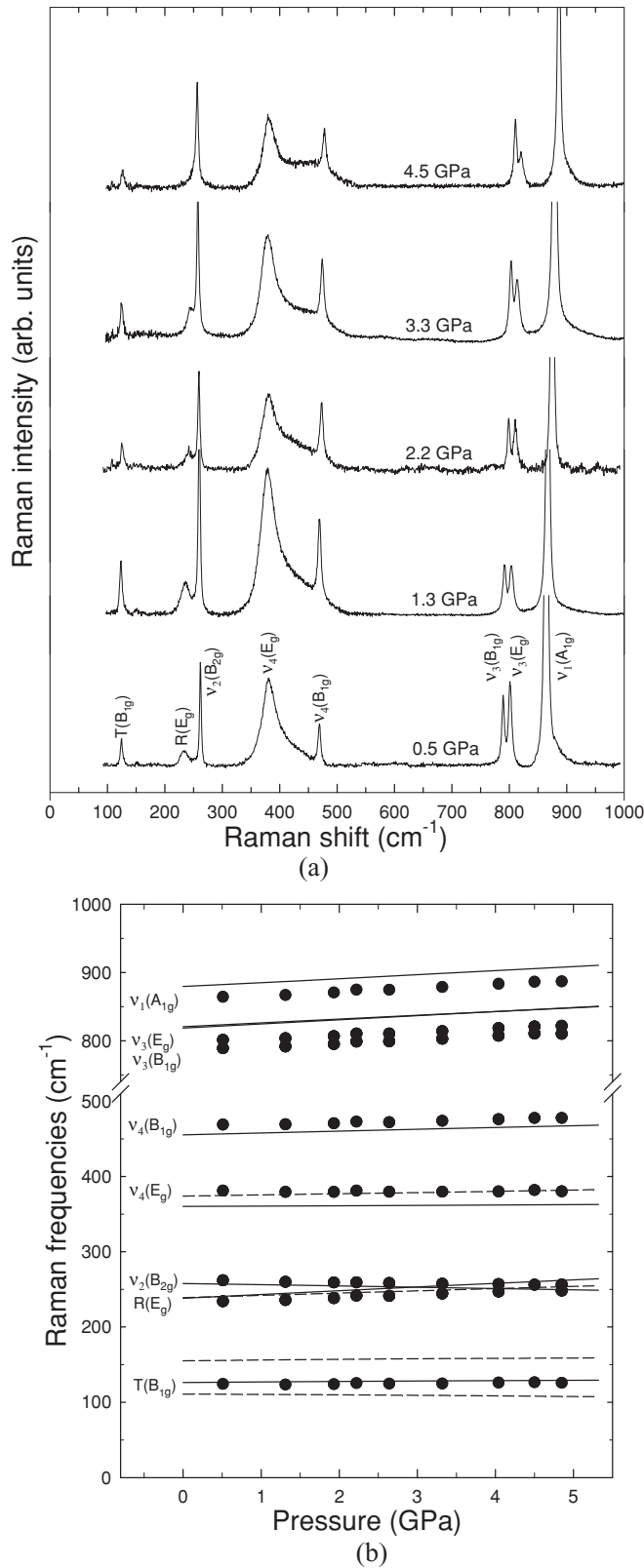


FIG. 7. (a) Raman spectra of CeVO₄ in the zircon phase between 0.5 and 4.5 GPa, (b) experimental pressure dependence of the Raman-mode frequencies in zircon-type CeVO₄. The solid lines are the calculated modes. The dashed lines represent Raman modes not observed in the experiments.

As can be seen from Fig. 7(a) and Table VI, the intense symmetric-stretching internal mode $\nu_1(A_{1g})$, observed at 864 cm^{-1} near ambient pressure exhibits the lowest frequency for this mode when compared with that of other rare-earth orthovanadates.³⁸ This result suggests that CeVO₄ exhibits the weakest intratetrahedral V-O bonds of all orthovanadates. The two asymmetric-stretching modes $\nu_3(E_g)$ and $\nu_3(B_{1g})$ have been observed at 801 and 789 cm^{-1} , respectively. Apart from these phonons of CeVO₄, the four bending modes of the VO₄ unit could be observed in the frequency range from 260 to 500 cm^{-1} . However, we have observed only three bending modes [$\nu_2(B_{2g})$, $\nu_2(A_{1g})$, and $\nu_4(B_{1g})$], whose frequencies at ambient pressures are 262 , 381 , and 469 cm^{-1} , respectively. The asymmetric-bending mode $\nu_4(E_g)$ could not be detected. In most of the orthovanadates this mode remains undetectable.^{39–43} Similarly, out of the five external modes, two $T(E_g)$ modes and one $T(B_{1g})$ mode have not been detected. These modes are absent probably due to their weak Raman-scattering cross sections.

Figure 7(b) shows the pressure dependence of Raman modes of CeVO₄ in the zircon phase. The symmetry assignment for the Raman modes along with their experimental and calculated frequencies, pressure coefficients, and mode Grüneisen parameters (γ) are shown in Table VI. A very good agreement is found between experimental and theoretical results. The frequencies of almost all Raman-active modes of the zircon phase exhibit a positive pressure coefficient. The rotational mode has the highest Grüneisen parameter, being the mode most sensitive to changes of volume. In addition, the internal bending mode of $\nu_2(B_{2g})$ symmetry shows a negative pressure coefficient. A similar soft behavior was observed for the $\nu_2(B_{2g})$ mode in other orthovanadates, e.g., YVO₄,³⁹ YbVO₄,⁴¹ LuVO₄,⁴² ScVO₄,⁴⁰ and also in ScPO₄.¹⁰ Our calculations also predict the softening of the $T(E_g)$ mode with lowest frequency in CeVO₄. This softening has been also observed in other orthovanadates and it could be a characteristic behavior of zircon-type compounds.^{39–42} It is important to note that the softening of both $T(E_g)$ and $\nu_2(B_{2g})$ modes has been observed in compounds exhibiting the zircon-to-scheelite transition^{39–44} and also in those showing the zircon-to-monazite transition.^{7,10} Therefore, it indicates that the softening of both modes is a consequence of the instability of the zircon phase. We will discuss this issue in great detail in our discussion section.

As regards the values of the pressure coefficients of the Raman-active modes of the zircon-type orthovanadates, it is worth mentioning that the pressure coefficients of the internal stretching modes in CeVO₄ are of the order of 4.9 – $5.5\text{ cm}^{-1}/\text{GPa}$, which are in close comparison with those of other orthovanadates.^{39–44} Similarly, the pressure coefficients of the $\nu_2(B_{2g})$ bending mode are of the order of -1.2 as it is among, those of other orthovanadates. Finally, it is interesting to note that the pressure coefficient of the rotational $R(E_g)$ mode of CeVO₄ has the smallest pressure coefficient of all known orthovanadates, thus indicating weak Ce – VO₄ bonds compared with those of other orthovanadates.^{39–44}

TABLE VI. *Ab initio* calculated and experimental frequencies at ambient conditions (0.5 GPa), pressure coefficients, and mode Grüneisen parameters (γ) of the CeVO₄ in the zircon phase.

Raman-mode Symmetry	ω_0^a (cm ⁻¹)	$d\omega/dP^a$ (cm ⁻¹ /GPa)	γ^a	ω_0^b (cm ⁻¹)	$d\omega/dP^b$ (cm ⁻¹ /GPa)	γ^b
T(E _g)	110.5	-0.52	-0.63	-	-	-
T(B _{1g})	123.4	0.59	0.54	124.4	0.56	0.54
T(E _g)	153.0	0.66	0.57	-	-	-
R(E _g)	232.5	4.92	2.61	234.1	3.52	1.79
T(B _{1g})	236.7	2.92	1.58	-	-	-
ν_2 (B _{2g})	254.1	-1.18	-0.97	261.9	-1.25	-0.57
ν_4 (E _g)	354.6	0.42	0.20	381.1	0.12	0.04
ν_2 (A _{1g})	368.7	1.46	0.52	-	-	-
ν_4 (B _{1g})	451.4	2.27	0.65	468.9	2.26	0.57
ν_3 (B _{1g})	828.3	5.08	0.80	789.1	5.26	0.79
ν_3 (E _g)	828.7	5.41	0.85	801.3	4.95	0.73
ν_1 (A _{1g})	872.3	4.37	0.66	864.3	5.52	0.76

^aTheoretical calculations.^bExperimental data.

2. Monazite-structured CeVO₄

Raman-scattering measurements in CeVO₄ evidence that this compound undergoes a zircon-to-monazite phase transition above 5 GPa. Group-theoretical calculations for CeVO₄ in the monazite phase (SG: $P2_1/n$, point group C_{2h}^5) predict 36 vibrational Raman modes at the BZ center with the following symmetries $\Gamma = 18A_g + 18B_g$.⁴⁵ These modes can be further classified into internal (ν_1 - ν_4) and external (translational T and rotational R) modes of VO₄ units as follows:

$$\Gamma = A_g(6T, 3R, \nu_1, 2\nu_2, 3\nu_3, 3\nu_4) + B_g(6T, 3R, \nu_1, 2\nu_2, 3\nu_3, 3\nu_4) \quad (2)$$

Raman spectra of CeVO₄ in the monazite phase at selected pressures are shown in Fig. 8(a). At 5.3 GPa we have observed the appearance of several Raman bands, both in the range between 0 and 500 cm⁻¹ and between 700 and 1000 cm⁻¹, accompanied by the broadening of many Raman modes. These changes in the Raman spectra are indicative of a structural phase transition toward the lower-symmetry monoclinic monazite phase. This result is consistent with our XRD investigations and with previous results in orthophosphates.^{7,10} Out of 36 Raman-active modes in the monazite phase, we observed only 27 modes with measurable intensities up to 12 GPa. The absence of nine Raman modes could be due to a very small Raman-scattering cross section and also possibly due to overlapping of many A_g and B_g modes due to their small splitting. Similarly, in Raman-scattering measurements of several monazite-type orthophosphates, 22 Raman modes were recorded out of 36 Raman active modes.⁴⁶ Above 12 GPa only the Raman-active modes of the high-frequency region above 700 cm⁻¹ were noticeable, and on further increase of pressure at 15.9 GPa all the Raman modes are completely diminished. This could have happened due to an extremely weak Raman signal at this pressure, or possibly due to a second phase transition to a Raman nonactive phase, or because of pressure-induced amorphization. Future studies will be done to clarify this point, which is beyond the scope of this paper.

The mode assignment of the experimentally observed Raman modes in CeVO₄ in the monazite phase was done

by comparing the experimental and calculated values of frequencies and pressure coefficients and is shown in Table VII together with the mode Grüneisen parameters (γ), calculated using the bulk modulus of the monazite phase, $B_0 = 131.5$ GPa, obtained from XRD measurements. The assignment of modes can be done broadly in two regions. The internal modes of the vibration of VO₄ tetrahedra cover the frequency range of 290–950 cm⁻¹, while the external modes of VO₄ unit span the frequency range of 50–270 cm⁻¹. This assignment is consistent in comparison with orthophosphates.⁴⁶ In general, the calculated Raman frequencies and their pressure coefficients are in good agreement with our experimental ones. The only mode whose assignment is doubtful is the soft mode observed at 89.6 cm⁻¹ at ambient pressure. This mode shows a frequency close to that of the calculated A_g mode at 92.5 cm⁻¹ but its negative pressure coefficient is more similar to that of the calculated B_g mode at 95.3 cm⁻¹.

Figure 8(b) shows the pressure dependence of the Raman-active modes of CeVO₄ in the monazite phase. All the experimental Raman-active modes show a positive pressure coefficient, except the two Raman modes at 89.6 and 140.4 cm⁻¹, which is in good agreement with our calculations that predict three soft modes. The pressure evolution observed for the Raman modes is similar to that reported for monazite CePO₄.³² Only the soft mode of B_g symmetry calculated at 66.2 cm⁻¹ at ambient pressure was not measured. The softening of the same Raman modes was previously observed in TbVO₄.⁸ On release of pressure from 15.9 GPa to ambient pressure the monazite phase was recovered, thus showing the irreversible nature of the zircon-to-monazite phase transition in good agreement with our XRD results. It must be noted that a similar behavior was observed for TbPO₄.⁷

3. Scheelite-structured CeVO₄

Group-theoretical calculations for CeVO₄ in the scheelite phase (SG: $I4_1/a$, point group C_{4h}^6) predict 13 vibrational Raman modes at the BZ center with the following symmetries: $\Gamma = 3A_g + 5B_g + 5E_g$.⁴⁷ As already commented on, our theoretical calculations predict the possibility of the scheelite phase

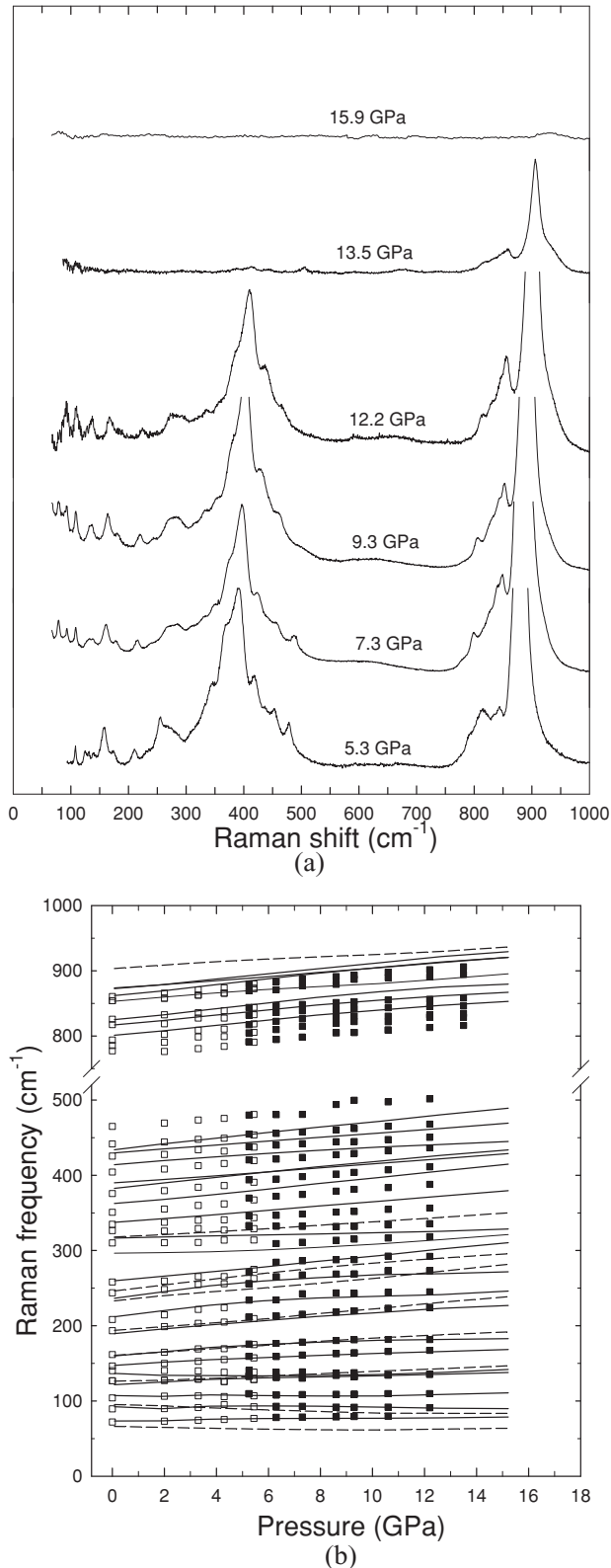


FIG. 8. (a) Raman spectra of the monazite phase of CeVO_4 at pressures between 5.3 and 15.9 GPa, (b) experimental pressure dependence of the Raman-mode frequencies in monazite phase CeVO_4 . Filled and empty squares correspond to Raman modes in monazite CeVO_4 on upstroke and downstroke, respectively. The solid lines are the calculated modes. The dashed lines represent Raman modes not observed in the experiments.

for CeVO_4 which has been observed in Bridgman-cell experiments. Table VIII shows the comparison of theoretically calculated and experimental Raman modes of CeVO_4 in the scheelite phase along with pressure coefficients and mode Grüneisen parameters (γ); to evaluate the mode Grüneisen parameters (γ) the calculated bulk modulus $B_0 = 138.65$ GPa is used. The calculated Raman frequencies match quite well with the experimental ones at ambient conditions. The pressure coefficients are similar to those observed in the scheelite phase in other vanadates.^{39,43} In Table VIII, it can be seen that the strongest mode of scheelite, the symmetric-stretching A_g mode, is observed at 826.7 cm^{-1} , which is consistent with other scheelite orthovanadates.^{39,40} This implies a collapse of the frequency of the strongest mode in comparison with the zircon structure, a typical feature of the zircon-to-scheelite transition. Finally, the frequencies of other Raman modes are also in quite good agreement with those of other scheelite orthovanadates.^{39,40}

V. DISCUSSION

Zircon-structured orthovanadates under hydrostatic compression were observed to undergo the zircon-to-scheelite and then the scheelite-to-fergusonite phase transitions.^{39–42} In all cases, the scheelite phase was found to be quenchable on release of pressure. The zircon-to-scheelite phase transition is a rather sluggish first-order and reconstructive transition in nature⁴⁸ while the scheelite-to-fergusonite transition is a ferroelastic second-order and displacive transition.⁴⁹ In our case, the zircon-to-scheelite transition involves a volume collapse of 11%, which is consistent with the first-order character of the transformation. On the other hand, the zircon-to-monazite phase transition is a rather sudden transition with a volume collapse of 8%, with an increase of the Ce^{3+} coordination and without the coexistence of phases over a wide pressure range (typical of the zircon-to-scheelite transition). All these facts indicate the first-order and reconstructive nature of this phase transition. The structural relationships between monazite and zircon were elucidated by Ni *et al.*⁵ These relationships can be used to understand the reconstructive nature of the zircon-to-monazite phase transition with the help of Fig. 1. In monazite, CeO_9 polyhedra share edges and corners with an isolated VO_4 tetrahedra, whereas in Zircon it is CeO_8 dodecahedra that share edges and corners. The edger sharing VO_4 and CeO_8 polyhedral chains along the [001] direction of the zircon structure appears to be like polyhedral chains in monazite, although the chains are twisted to accommodate the ninth atom. However, due to the extra ninth oxygen atom, the packing efficiency in the monazite phase is better compared with that of the zircon phase. This result supports the observed higher bulk modulus value for monazite CeVO_4 . This new atomic arrangement in the monazite phase can be realized by considering a slight shift of the [001] planes and a rotation of VO_4 polyhedra in the a - b plane. The fact that this transition (as well as the zircon-to-scheelite transition) is not reversible is a notable aspect. We think that a large kinetic barrier is the cause of nonreversibility, which is consistent with the reconstructive mechanism of this transition and its first-order character. Another important issue to remark is that the presence of AO_9 polyhedra (a cationic pentagonal interpenetrating tetrahedral polyhedral) allows monazite to accommodate chemically

TABLE VII. Theoretical and experimental frequencies, pressure coefficients, and mode Grüneisen parameters of monazite CeVO_4 . Grüneisen parameters, $\gamma = (B_0/\omega_0)d\omega/dP$, calculated assuming $B_0 = 142$ GPa.

Raman-mode Symmetry	ω^a (cm^{-1})	$d\omega/dP^a$ ($\text{cm}^{-1}/\text{GPa}$)	γ^a	ω^b (cm^{-1})	$d\omega/dP^b$ ($\text{cm}^{-1}/\text{GPa}$)	γ^b
B_g	66.2	-0.58	-1.11	-	-	-
A_g	73.5	0.51	0.82	72.1	0.35	0.69
A_g	92.5	0.27	0.35	89.6	-0.31	-0.49
B_g	95.3	-1.20	-1.61	-	-	-
A_g	107.3	0.01	0.01	104.2	0.13	0.18
B_g	121.6	1.35	1.30	126.3	0.55	0.62
B_g	126.4	1.28	1.18	-	-	-
A_g	136.6	-0.30	-0.27	140.4	-0.69	-0.69
A_g	147.4	1.67	1.32	146.5	1.26	1.22
A_g	159.9	2.34	1.69	161.8	0.95	0.83
B_g	160.7	2.35	1.68	-	-	-
B_g	189.6	2.89	1.75	193.4	1.97	1.45
A_g	193.8	2.81	1.66	-	-	-
B_g	211.9	3.24	1.76	208.3	1.68	1.15
B_g	233.0	2.90	1.44	-	-	-
A_g	236.6	3.53	1.72	243.8	2.16	1.26
B_g	246.2	3.94	1.83	-	-	-
A_g	259.7	3.29	1.46	258.2	2.56	1.41
B_g	296.5	0.96	0.39	310.1	1.58	0.72
A_g	316.5	0.69	0.27	326.5	0.37	0.16
B_g	317.9	1.98	0.74	-	-	-
A_g	337.4	2.78	0.97	334.8	1.57	0.67
A_g	362.7	3.37	1.09	350.9	2.67	1.08
A_g	382.7	3.67	1.12	375.8	2.60	0.98
B_g	390.1	2.50	0.76	404.4	2.40	0.84
B_g	414.1	2.42	0.70	425.6	1.72	0.57
A_g	429.8	2.63	0.72	441.7	2.02	0.65
B_g	433.9	3.81	1.03	465.0	3.68	1.12
A_g	801.3	3.99	0.59	771.2	3.04	0.56
B_g	817.2	3.93	0.57	784.5	2.90	0.53
A_g	825.3	4.41	0.64	794.1	1.86	0.33
A_g	853.3	2.90	0.41	817.5	1.59	0.28
B_g	862.1	4.44	0.61	825.2	1.72	0.29
B_g	872.8	3.94	0.54	854.4	3.17	0.53
A_g	873.8	3.03	0.41	860.2	3.15	0.52
B_g	903.7	2.29	0.31	-	-	-

^aTheoretical calculations.

^bExperimental data.

diverse cations. The irregular coordination around the A cation does not place severe symmetry, size, or charge constraints and allows large domains of chemical composition. Therefore, we predict that, in addition to CeVO_4 , other zircon-structured vanadates with large A cations could also take the monazite structure under pressure, becoming isostructural to LaCrO_4 .

An important issue related to the fact that under compression the zircon structure becomes mechanically unstable is the softening of external translational $T(E_g)$ and internal $\nu_2(B_{2g})$ bending modes. In zircon-structured compounds, the softening of the $T(E_g)$ mode is related to a softening of the C_{44} elastic constant. This fact is caused by monoclinic distortions in the a - c or b - c planes.⁵⁰ Similarly, the softening of the $\nu_2(B_{2g})$ mode is related to a softening of the C_{66} elastic constant, which is caused by orthorhombic distortions in the basal plane.⁵⁰ Low-temperature studies of zircon-structured DyVO_4 , DyAsO_4 , and TbVO_4 indicate a similar phonon

softening which can be attributed to the distortion of the zircon structure along the $[110]$ or $[100]$ direction.^{51,52} By analogy, the pressure-induced contraction would lead to similar distortions in the zircon structure, making this structure unstable.

To contribute to the systematic understanding of structural properties of zircon-type oxides, here we will make the attempt to understand the factors which govern the phase stability at ambient pressure and the sequence of pressure-induced structural phase transitions in orthovanadates and orthophosphates. As commented on earlier, the phase stability of zircon-type orthovanadates and orthophosphates and their phase transitions seem to greatly depend on the ionic radii. This is consistent with the recent updated version of Bastide's diagram for ABX_4 compounds, elaborated on and discussed by Errandonea and Manjón.⁵³ In Bastides diagram, the phase stability and phase transitions of ABO_4 orthovanadates and orthophosphates can be understood by considering the role

TABLE VIII. *Ab initio* calculated and experimental frequencies at ambient pressure for the scheelite phase.

Raman mode symmetry	ω_0^a (cm ⁻¹)	$d\omega/dP^a$ (cm ⁻¹ /GPa)	γ^a	ω_0^b (cm ⁻¹)
T(E _g)	106.9	-0.22	-0.35	-
T(B _g)	132.0	-0.77	-1.04	-
T(E _g)	165.4	2.65	2.42	-
T(B _g)	176.9	2.80	2.37	-
R(A _g)	206.4	0.52	0.42	230.2
R(E _g)	273.1	2.57	1.48	313.5
ν_2 (A _g)	331.5	2.67	1.28	346.6
ν_2 (B _g)	349.1	1.30	0.61	367.7
ν_4 (B _g)	370.9	2.96	1.29	403.4
ν_4 (E _g)	393.6	2.70	1.11	427.8
ν_3 (E _g)	772.6	3.61	0.76	742.3
ν_3 (B _g)	817.1	3.09	0.62	799.8
ν_1 (A _g)	840.26	2.71	0.54	826.7

^aTheoretical calculations.^bExperimental data.

played by cationic radii, r_A and r_B , with respect to anion radius, r_O .⁵⁴ In this diagram, those compounds which have r_A/r_O and r_B/r_O cation-to-anion ratios well inside the stability region of the zircon structure are observed to undergo the zircon-to-scheelite phase transition and follow the traditional north-east (NE) rule in Bastide's diagram.⁵⁴ Such is the case of many zircon-structured orthovanadates, like YVO₄,³⁹ ScVO₄,⁴⁰ YbVO₄,⁴¹ LuVO₄,⁴² and also in ScPO₄.¹⁰ On the other hand, those compounds whose r_A/r_O and r_B/r_O cation-to-anion ratios fall near the border of the stability region of the zircon and monazite structures could crystallize in both phases, like CeVO₄, TbPO₄,^{7,8} and YPO₄.⁹ In this case, the compounds crystallizing in the zircon phase, like CeVO₄, are prone to undergo the zircon-to-monazite phase transition. A few orthophosphates, like YbPO₄ and LuPO₄,¹¹ are exceptions to this rule and deserve further investigation. Finally, the last case would correspond to compounds whose r_A/r_O and r_B/r_O cation-to-anion ratios fall well inside the stability region of the monazite phase. This is the case for most orthophosphates, which already crystallize in the monazite phase. To conclude we would like to add that the crystal chemistry arguments here used to discuss the structural behavior of phosphates and vanadates has been recently satisfactorily used to describe arsenates and chromates.⁵⁵

Finally, we would like to comment on additional consequences of the observed pressure-driven structural changes. CeVO₄ is a large-bandgap material with a bandgap energy of 3.1–4.2 eV.^{56,57} It is usually assumed that the electronic structure near the Fermi level is dominated by V 3*d* and O 2*p* states. However, recently a bandgap of 1.8 eV was measured⁵⁸ and attributed to the presence of localized 4*f* levels of Ce between the valence and conduction bands. Clearly there is a

contradiction between both pictures of the electronic band structure which could be solved by HP optical-absorption studies like those already preformed in PbWO₄.⁵⁹ It is known that pressure affects differently localized and delocalized electronic states. In particular, if 4*f* states are present near the Fermi level, we expect E_g to be considerable reduced by pressure within the stability range of the zircon structure. In addition, both the zircon-to-scheelite and zircon-to-monazite transitions will probably cause a large collapse of the bandgap as a consequence of the atomic rearrangement after the phase transitions. Finally, another issue interesting to explore in the future are the effects caused by HP *f* electron delocalization of lanthanides,⁶⁰ which, among other things, should modify the magnetic properties of compounds like CeVO₄.

VI. CONCLUSION

XRD, Raman-scattering, and theoretical studies of CeVO₄ up to 16 GPa suggests that the low-pressure zircon phase undergoes an irreversible zircon-to-monazite phase transition at 5.3 GPa. XRD and Raman signals weaken considerably beyond 12 and 16 GPa, respectively, evading any further measurements. The symmetries of the Raman modes in the zircon and monazite phases of CeVO₄ have been assigned in accordance with our lattice-dynamics calculations. In general, good agreement is found for the zircon-to-monazite phase transition by our experimental and theoretical data. Our calculations predict the monazite-to-scheelite phase transition, which could not be detected experimentally under hydrostatic conditions. However, the scheelite structure is observed upon compression under nonhydrostatic conditions. The softening of external translational T(E_g) and internal ν_2 (B_{2g}) bending modes can be attributed to distortion in the zircon structure along the [110] or [100] direction associated with the phase transition. We have also discussed the dependence of the phase stability of the zircon phase and its pressure-induced phase transitions either to the scheelite or monazite phases in orthovanadates and orthophosphates on their r_A/r_O and r_B/r_O cation-to-anion ratios. Finally, the effects of nonhydrostaticity in the structural sequence are discussed. We found that different compression methods induced a transition to different structures.

ACKNOWLEDGMENTS

Research financed by the Spanish MCINN under Grant Nos. MAT2010-21270-C04-01/02/04, CSD2007-00045, and CTQ2009-14596-C02-01. Some of the authors are members of the MALTA Consolider Team. S. L.-M. acknowledges the financial support of the CONACyT México under a postdoctoral fellowship. Supercomputer time has been provided by the Red Española de Supercomputación (RES) and the MALTA cluster.

*daniel.errandonea@uv.es

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